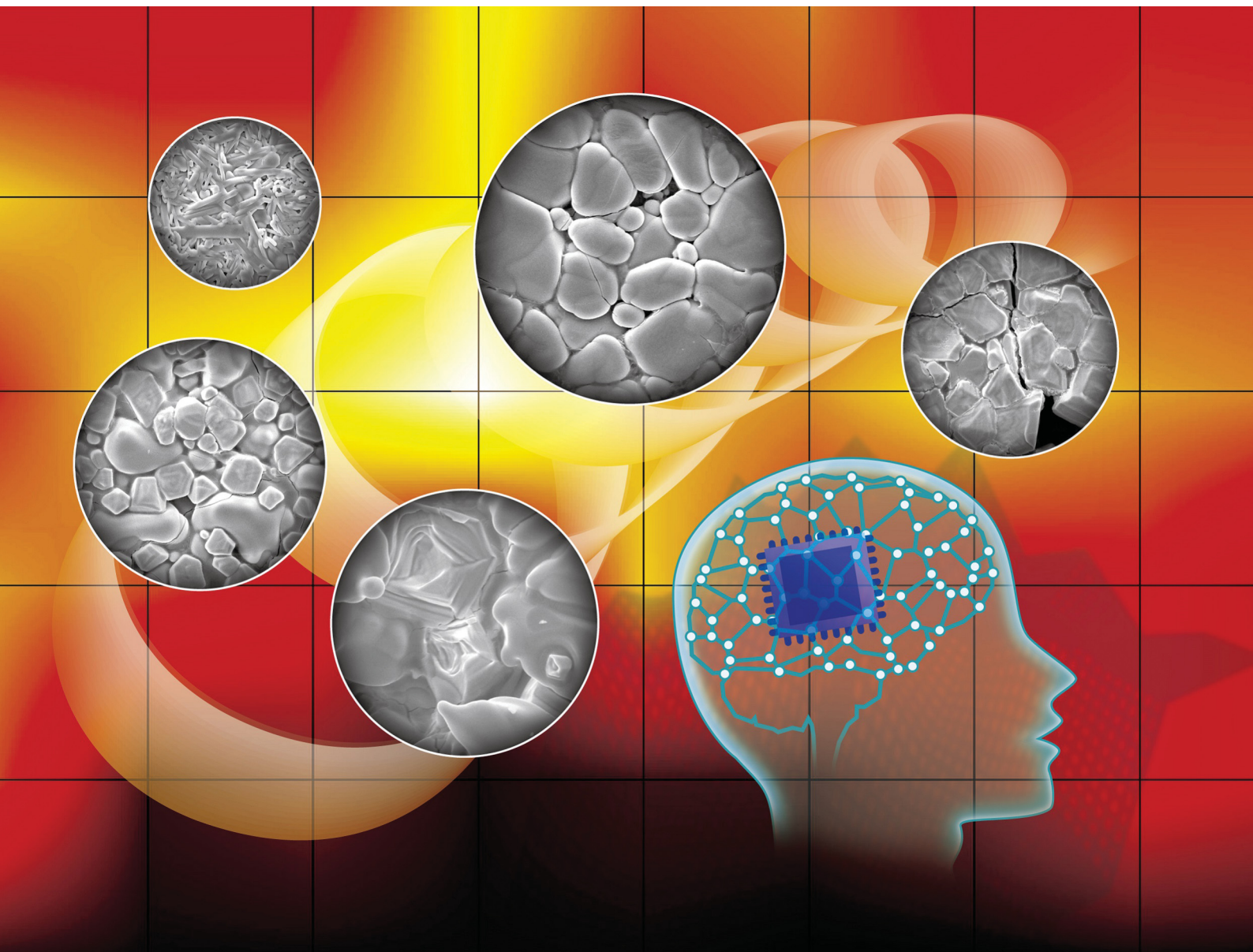


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PAPER

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Process optimisation for NASICON-type solid electrolyte synthesis using a combination of experiments and bayesian optimisation†

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Na superionic conductor (NASICON)-type $\text{LiZr}_2(\text{PO}_4)_3$ (LZP) is an oxide-based solid electrolyte candidate for use in all-solid-state Li-ion batteries. However, as the ionic conductivity is insufficient, doping with aliovalent cations has been carried out to improve the Li-ion conductivity by controlling the composition and crystal structure. Li-ion conductivity is also affected by the microstructural properties of a sintered body, such as density, morphology, and elemental distribution, and thus, controlling process parameters, such as heating conditions during the solid-state reaction, improves conductivity. Using an exhaustive experimental approach, Ca and Si co-doped Li-rich NASICON-type LZP was synthesised *via* solid-state reactions under various two-step heating conditions to yield the highest Li-ion conductivity by optimising the conditions. The highest total Li-ion conductivity of $3.3 \times 10^{-5} \text{ S cm}^{-1}$ was obtained when the sample was first heated at 1050 °C and then heated at 1250 °C. The crystal structures, relative densities, micromorphologies, and Li-ion conductivities of the materials were characterised, and their relationships were investigated. These relationships were complex, and intuitively determining the optimal conditions was challenging with only a few experiments. Instead, as a proof-of-concept study, the collected data were used to demonstrate that Bayesian optimisation (BO) efficiently improved the experimental determination of the optimal heating conditions. The BO-guided experimental investigation determined the optimal conditions more rapidly compared to conventional trial-and-error approaches employed in the materials industry. The efficiency factor was approximately double that of the exhaustive search.

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Introduction

Research and development of large-scale rechargeable Li-ion batteries are critical, as they are essential for use in electric vehicles and aircraft¹ and the realisation of new social infrastructure, such as smart grids. However, safety concerns due to the usage of flammable organic electrolytes in current Li-ion batteries should be addressed. Replacement of the liquid electrolyte with a non-flammable inorganic solid electrolyte in fabricating an all-solid-state Li-ion battery is an attractive option, and thus, several Li-ion conductors have been intensively studied.^{2,3} Typical inorganic solid electrolytes include

sulfides and oxides.^{4–7} Sulfide electrolytes, such as $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$, exhibit high Li-ion conductivities,^{8–11} but sulfides are unstable in the atmosphere and generally react with water, forming toxic H_2S .¹² Conversely, oxide-based solid electrolytes should display excellent chemical stabilities, although their Li-ion conductivities are generally lower than those of sulfides, partly due to strong interactions between Li^+ and O^{2-} .¹³ Oxide-based solid electrolytes with Na superionic conductor (NASICON),^{14,15} garnet,^{16,17} or perovskite structures^{18,19} are representative fast Li-ion conductors. Among these, NASICON-type $\text{LiZr}_2(\text{PO}_4)_3$ (LZP) gained particular interest, as stable charge/discharge cycles were reported using Li metal/LZP/LiFePO₄ all-solid-state Li metal batteries.²⁰ LZP exhibits four crystalline phases (polymorphs), denoted as α -, α' -, β -, and β' -phases,^{21–23} which depend on the sintering conditions.^{15,24} The Li-ion conductivity, including bulk and grain-boundary contributions, of the α -phase is the highest at $\sim 10^{-6} \text{ S cm}^{-1}$ at room temperature.^{25,26} However, as the ionic conductivity is insufficient for use in all-solid-state Li-ion (metal) batteries, substitution with aliovalent cations was attempted to improve the Li-ion conductivity,^{21,27–31} e.g. the replacement of the Zr^{4+} ions in LZP with lower-valence

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† Electronic supplementary information (ESI) available: EDS images of the sintered body, histograms of the number of observations required to determine the heating conditions affording the highest ionic conductivity and the lowest activation energy, and selected properties experimentally determined for all 54 samples (PDF). See DOI: <https://doi.org/10.1039/d2ma00731b>



cations, such as Ca^{2+} , led to the formation of interstitial Li ions, which enhanced the Li push-out mechanism due to double-Li occupancies close to the 6b sites.¹⁴ However, further replacement with dopant cations reduced the ionic conductivity due to the trapping effect between the dopant and Li ions.³² Thus, carefully controlling the dopant ions, *e.g.* *via* co-doping, is necessary to optimise the Li-ion conductivity.^{24,33} Besides compositional control, Li-ion conductivity is also affected by density, morphology, and elemental distribution,²¹ and thus, controlling process parameters, such as heating conditions during the solid-state reaction, is effective in improving conduction performance.³⁴

Optimising the composition and production process is critical for use in the materials industry, but the traditional trial-and-error approach often requires economic cost and time. Previously, machine learning was successful in improving materials,³⁵ and we successfully demonstrated the efficient compositional optimisation of experimental Li-ion conductivity using Bayesian optimisation (BO) of $\text{Li}_{(1+2x+y)}\text{Ca}_x\text{Y}_y\text{Zr}_{(2-x-y)}(\text{PO}_4)_3$ and $(\text{Li}_y\text{La}_{(1-y)/3})_{(1-x)}\text{Sr}_{0.5x}\text{NbO}_3$.^{24,36} Additionally, BO should be a suitable approach in analysing production processes for which the optimal parameters are challenging to determine manually or heuristically.³⁷ BO is applied to establish parameters in polymer fibre synthesis,³⁸ TiO_2 film deposition,³⁹ gas atomisation of alloy powders,⁴⁰ and permanent magnet synthesis.⁴¹ However, reports of process optimisation in terms of Li-ion conductors are limited.

Thus, in this study, we focused on double-doped LZP $\text{Li}_{1.45}\text{Ca}_{0.15}\text{Zr}_{1.85}\text{Si}_{0.15}\text{P}_{2.85}\text{O}_{12}$ (LZP-CS), because the compound shows improved Li-ion conductivity. Ca and Si ions partially occupy octahedral Zr and tetrahedral P sites in LZP-CS.^{15,24} The sintered bodies of LZP-CS samples were synthesised *via* conventional solid-state reactions with 54 different two-step heating conditions, and the resultant 54 sintered bodies were investigated exhaustively in terms of phase stability, sintered structure, and ionic conductivity. Furthermore, BO was adopted for these 54 datasets to reveal efficient processing parameter optimisations. This study mainly focused on the optimisation of the experimental process.

Experimental

$\text{Li}_{1.45}\text{Ca}_{0.15}\text{Zr}_{1.85}\text{Si}_{0.15}\text{P}_{2.85}\text{O}_{12}$ was synthesised *via* conventional solid-state reactions at 54 different heating conditions. Li_2CO_3 (99.9%; Kojundo Chemical Laboratory, Sakado, Japan), ZrO_2 (99.9%; Kishida Chemicals, Osaka, Japan), CaCO_3 (99.5%; Kishida Chemicals), SiO_2 (99.9%; Kojundo Chemical Laboratory), and $\text{Zr}(\text{HPO}_4) \cdot 1.5\text{H}_2\text{O}$ (Toagosei, Tokyo, Japan) were used as raw materials. Stoichiometric amounts of these raw materials were mixed in alumina mortars with distilled water in a 1 : 2 mass ratio. The mixed powders were dried at 100 °C and pelletised under uniaxial pressing at 127 MPa. The pellets were calcined for 4 h at nine different temperatures (1st heating): 900, 950, 1000, 1050, 1100, 1150, 1200, 1250, or 1300 °C. After the 1st heating, samples were ball-milled using a ZrO_2 pot and balls (diameters of 10 mm) at 400 rpm for 30 min using a planetary ball mill (Pulverisette 7,

Fritsch, Idar-Oberstein, Germany). Thereafter, samples were re-pelletised and sintered for 4 h at six different temperatures (2nd heating): 1100, 1150, 1200, 1250, 1300, or 1350 °C. Hereafter, heating conditions are denoted as 'HC-X-Y', where X and Y correspond to the 1st and 2nd heating temperatures (°C). The sample size was approximately 9 mm in diameter and 1.5 mm in thickness. The crystal phase of the sample was determined by X-ray diffraction (XRD) using a MiniFlex 600 diffractometer (Rigaku, Tokyo, Japan) with $\text{CuK}\alpha$ radiation. The micromorphology of the sample was evaluated using scanning electron microscopy (SEM, JSM-6360LV, JEOL, Tokyo, Japan), and elemental analysis was performed using energy-dispersive X-ray spectroscopy (EDS, JSM-6010LA, JEOL). The ionic conductivities of the bulk structures of the pellets after the 2nd heating were measured *via* AC impedance. Both faces were polished with 2000-grit abrasive papers and Au was sputtered to produce the electrodes. The complex impedance was measured using an impedance analyser (VMP-300, BioLogic, Seyssinet-Pariset, France) at 30, 60, and 90 °C in the frequency range 1–10⁶ Hz with a potential of 1 V. The activation energy of the Li conductivity was estimated using Arrhenius plots of the measured Li-ion conductivity. The experimentally determined bulk densities, ionic conductivities, and activation energies are presented in Table S1 (ESI[†]).

BO was performed using the COMmon BO library (COMBO)⁴² to evaluate the efficiency of the optimised heating condition to obtain the highest ion conductivity or minimum activation energy among the experimentally evaluated samples. The conductivities at 30 °C and activation energies were used as the objective variables, and the calcination and sintering temperatures were used as descriptors in BO. Gaussian kernels were used and hyperparameters were automatically optimised at every step by maximising the type II likelihood, as implemented in the COMBO library.⁴² The random and BO searches were performed 1000 times each, and the average numbers of observations required to reach the optimal solution were compared. The expected improvement strategy was used to obtain the acquisition function.³⁶

Results and discussion

Crystal phase identification

Fig. 1(a) shows the XRD patterns of the samples after the 1st heating and simulated patterns based on inorganic crystal structure database (ICSD)-derived structure inputs of $\alpha\text{-LiZr}_2(\text{PO}_4)_3$ (#201935) and $\beta\text{-LiZr}_2(\text{PO}_4)_3$ (#91113). The primary products after the 1st heating at 900–1050 °C and 1100–1300 °C are β - and α -type, respectively. The α - and β -phases coexist in the samples heated at >1100 °C but <1250 °C, as indicated by the diffraction peaks at $2\theta \approx 15.8^\circ$, 25.8° , and 26.5° , which are characteristic of the β -phase. Fig. 1(b) shows the XRD patterns of HC-1050-Y (1100 ≤ Y ≤ 1350) and the ICSD-derived simulated pattern of monoclinic ZrO_2 (#68782), in addition to the simulated patterns of α - and $\beta\text{-LiZr}_2(\text{PO}_4)_3$ shown in Fig. 1(a). When the 1st heating temperature is <1200 °C, the β -phase is the main phase. However, when the 2nd heating temperature is >1250 °C, the characteristic peaks



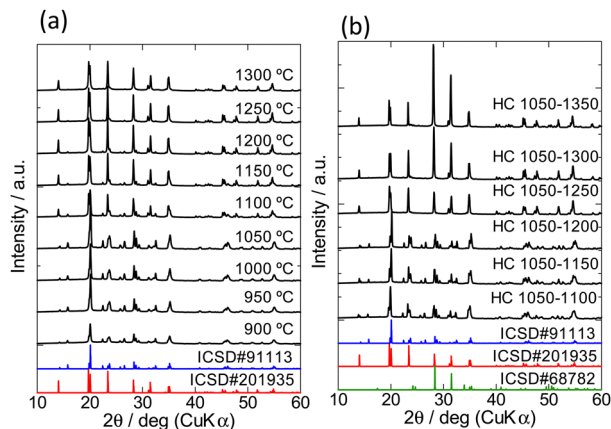


Fig. 1 X-Ray diffraction (XRD) patterns of the samples. (a) Samples after the 1st heating at temperatures ranging from 900 to 1300 °C and simulated patterns of α - and β -LiZr₂(PO₄)₃, which are #201935 and #91113, respectively, in the inorganic crystal structure database (ICSD) (b) XRD patterns of HC-1050-Y (1100 ≤ Y ≤ 1350) samples, the same simulated patterns shown in (a), and the ICSD-derived simulated pattern of monoclinic ZrO₂ (#68782).

of the β -phase are not observed, indicating that the sample exhibits a single α -phase. Ca-doped LZP requires heating at 1200 °C to yield the single α -phase,¹⁵ whereas LZP co-doped with Ca and Si requires a higher temperature to yield the single α -phase. Temperatures of 2nd heating of >1300 °C yield samples with higher intensities at $2\theta \approx 28.4^\circ$ and 31.5° compared to those of the α -phase simulated pattern (#201935). These peaks are consistent with the characteristic peaks of monoclinic ZrO₂ (#68782), suggesting the formation of ZrO₂. Similar results are observed when the temperature of the 1st heating is 900, 950, 1000, 1100, 1150, 1200, 1250, or 1300 °C.

Relative density and micromorphology

The bulk densities of all samples are shown as a heatmap in Fig. 2(a). The black dots in the mesh correspond to the sample treatments. When the 1st heating temperature is higher and the 2nd heating temperature is lower, the bulk density generally increases. Conversely, the opposite trend is observed when the 2nd heating temperature increases, *e.g.*, Fig. 2(b) shows the mass losses of the samples as a function of the 2nd heating temperature after the samples are preheated at 900 °C (1st heating temperature). A gradual decrease in the bulk density is observed upon increasing the 2nd heating temperature. A particularly large mass loss is observed when the heating temperature is increased from 1300 to 1350 °C. Therefore, the mass loss is ascribed to the evaporation of oxides with high vapor pressures, such as Li₂O and P₂O₅.²⁴

SEM images of the sample surfaces are shown in Fig. 3. The HC-X-1100 samples (Fig. 3(a–c)) form entangled structures with micrometre-sized needle-like crystals. These crystal morphologies are mainly due to the crystal structures, as the main phases of these samples consist of the β -type polymorphs. EDS of these samples reveals homogeneous distributions of constituent elements, except Li, in the sintered bodies

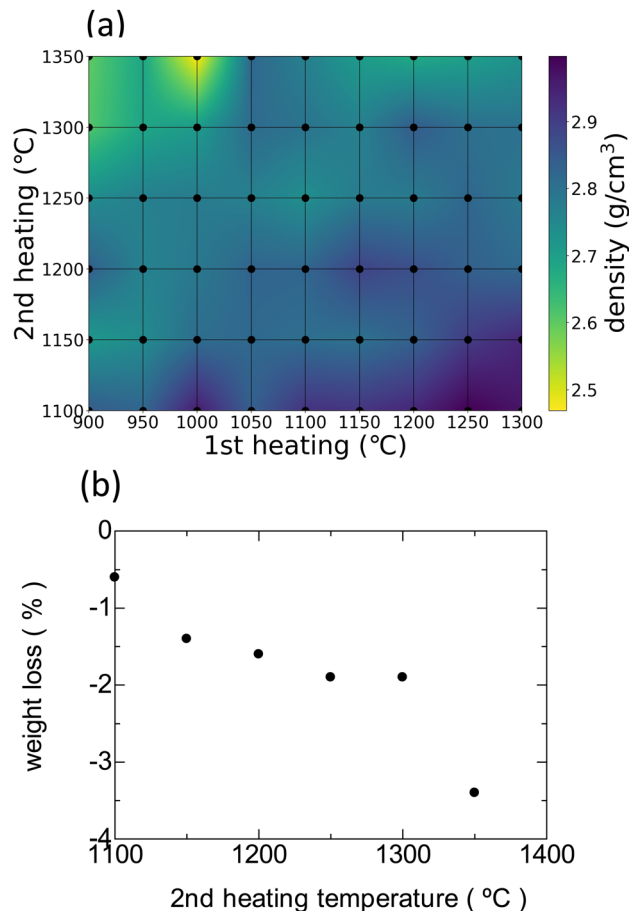


Fig. 2 (a) Bulk densities of the HC-X-Y (900 ≤ X ≤ 1300, 1100 ≤ Y ≤ 1350) samples. (b) Mass losses of the HC-900-Y (1100 ≤ Y ≤ 1350) samples.

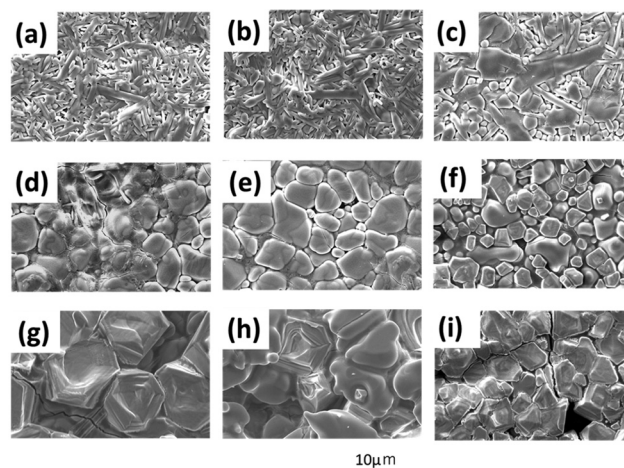


Fig. 3 Scanning electron microscopy images of the sample surfaces. (a) HC-900–1100, (b) HC-1050–1100, (c) HC-1300–1100, (d) HC-900–1250, (e) HC-1050–1250, (f) HC-1300–1250, (g) HC-900–1350, (h) HC-1050–1350, and (i) HC-1300–1350.

(Fig. S1, ESI†). Fig. 3(d, h, and i) show the SEM images of samples after 2nd heating at ≥ 1250 °C, wherein the main



phases are α -type polymorphs. Unlike the needle-like crystals observed in the β -type polymorph phase, the particles in the sintered bodies do not display specific crystal habits, and their aspect ratios are small. EDS mapping indicates the segregation of Ca-containing phases in the grain-boundary regions of the samples, HC-900–1250 and HC-1050–1250, as shown in Fig. S2 (ESI[†]). A Ca-rich boundary phase was observed in Ca-doped LZP by Xie *et al.*¹⁵ In HC-1300–1250, several angular particles and voids are observed, which may be due to thermal evaporation of volatile components. In the HC-X–1350 samples (Fig. 3(g–i)), particle growth is observed, and the particle size gradually decreases with an increase in the 1st heating temperature. HC-1300–1350 displays larger voids than those of HC-1300–1250 in their sintered bodies, indicating further evaporation of volatile components. This should be the evaporation-derived loss of sintering aids *via* the high-temperature 1st heating, as shown in Fig. 2(b), *i.e.* suppression of the 1st heating temperature prevents the evaporation of volatile components. Conversely, an increase in the 2nd heating temperature leads to particle growth, which reduces the density of the sintered body.

Li-ion conductivity

Fig. 4 shows the typical AC impedance plots at 30 °C of the (a) HC-1050–1200 and (b) HC-1050–1250 sintered bodies. The AC impedance plots consist of semicircles and spikes in the high- and low-frequency regions, respectively. The low-frequency spike corresponds to the capacitance at the interface between the LZP and Au electrode. A single distorted semicircle and two semicircles are observed for (a) HC-1050–1200 and (b) HC-1050–1250, respectively, and thus, the semicircles contain bulk and grain-boundary resistance contributions.⁴³ In HC-1050–1250, the ionic conductivities of the bulk and grain boundary are 7.0×10^{-5} and 5.6×10^{-5} S cm⁻¹ (at 30 °C), respectively, with a total ionic conductivity of 3.3×10^{-5} S cm⁻¹ based on fitting using an equivalent circuit model. However, the separation of the two resistance components is technically challenging in the case of a single distorted semicircle, *e.g.*, HC-1050–1200. Thus, hereafter, the Li-ion conductivity reported in this study contains both resistance components (total ionic conductivity).

Fig. 4(c) shows the distribution of the total ionic conductivity at 30 °C as a heatmap, as functions of 1st and 2nd heating conditions (dots in the mesh correspond to the preparation conditions). Compared to Fig. 2(a), no clear correlation between conductivity and bulk density is observed. The samples with lower sintering temperatures during the 2nd heating exhibit lower conductivities due to the formation of the low-conducting β -phase.^{23,24} The conductivities of the samples with 1st heating temperatures of >1100 °C are lower than those of the samples with 1st heating temperatures of <1050 °C. Similarly, decreased ionic conductivities are observed after the 2nd heating at temperatures of >1250 °C. This may be due to the vaporisation of volatile compounds, such as Li₂O and P₂O₅, as indicated in Fig. 2(b), and segregation of Ca-containing compounds at the grain boundary. Therefore, control of the 1st and 2nd heating conditions is required to optimise Li-ion conductivity. HC-1050–1250 exhibits the highest total conductivity of 3.3×10^{-5} S cm⁻¹ at 30 °C,

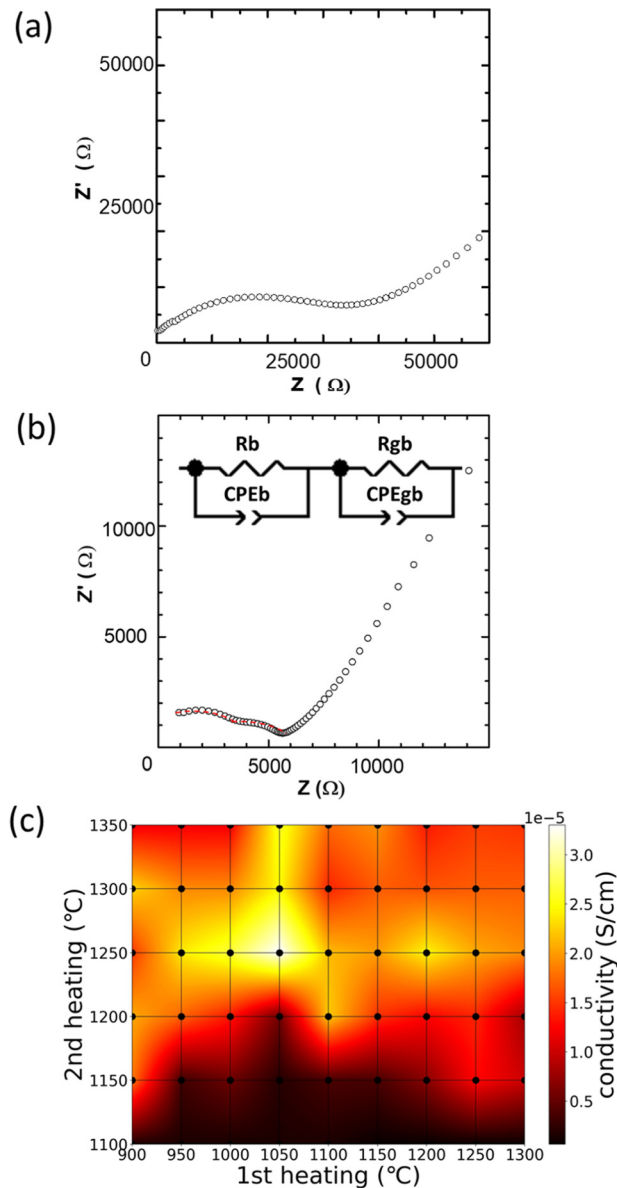


Fig. 4 (a) AC impedance plot of HC-1050–1200 at 30 °C, (b) AC impedance plot of HC-1050–1250 at 30 °C with a fitted curve (red solid line), and (c) heatmap of the ion conductivities of all samples at 30 °C.

which may be attributed to less segregation of Ca components at the grain boundary and little component loss *via* evaporation.

Fig. 5(a) shows the Arrhenius plots of the total ionic conductivities of all samples. The minimum activation energy is 0.31 eV for HC-1250–1250 (orange line in Fig. 5(a)). The heating condition dependences of the activation energies calculated using the Arrhenius plots are shown in Fig. 5(b). The grey, orange, and blue dots in Fig. 5(a) are the calculated points. Generally, a decrease in activation energy leads to an increase in ionic conductivity, and this trend is roughly observed in the heatmap of activation energies, as shown in Fig. 5(b), in comparison with that of ionic conductivity, as shown in Fig. 4(c). However, the optimised heating condition differs slightly in terms of the activation energy (HC-1250–1250) and



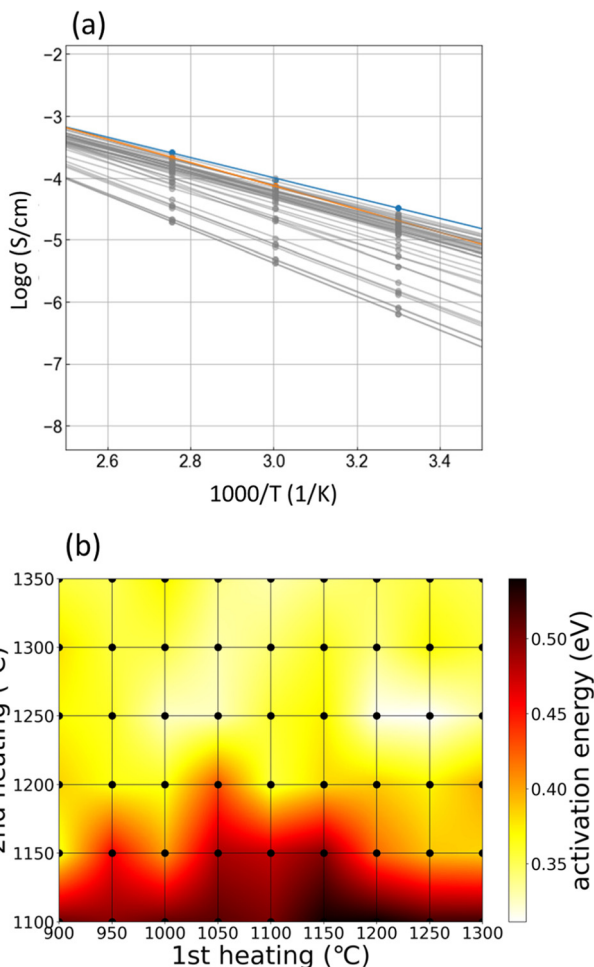


Fig. 5 (a) Arrhenius plots of all samples. The orange line indicates HC-1250–1250, which exhibits the minimum activation energy, and the blue line indicates HC-1050–1250, which exhibits the maximum Li-ion conductivity at 30 °C. (b) Heatmap of the activation energies of all samples.

Li-ion conductivity (HC-1050–1250). Therefore, the ionic conductivity also depends on parameters such as carrier concentration and grain-boundary structures. Hence, the relatively low ionic conductivity of HC-1250–1250, despite the lowest activation energy, is due to the vaporisation of Li during the 1st high-temperature heating. To summarise the results shown in Fig. 2–5, the heating conditions affect various parameters, such as the crystal phase, density, microstructure, elemental distribution, and composition deviation *via* vaporisation, and thus, the ionic conductivity-related properties. The relationship among these observations is quite complex, and therefore, intuitive determination of the optimised heating conditions is challenging with only a few experiments.

Determining the optimal conditions using BO

BO is performed to reduce the optimisation cost. As this study was a demonstration, *i.e.*, proof-of-concept, all our experimental datasets containing various 1st and 2nd heating conditions were used. These datasets were invisible to the Bayesian optimiser prior to optimisation, and the hidden experimental data were

opened in line with the instructions by the Bayesian optimiser, except at the first random selection step. Similarly, random sample selection was also applied for comparison. One thousand independent search cycles were performed using each selection scheme. Fig. 6 shows changes in the probability of establishing the highest ionic conductivity and the lowest activation energy as functions of the number of experiments (observations). Fig. S3 and S4 (ESI[†]) show the histograms of the number of observations required to determine the heating conditions affording the highest ionic conductivity and the lowest activation energy, respectively. In both cases, a mountainous distribution is observed. As shown in Fig. 6(a and b), during the initial 10–15 steps, the probability of discovery ratio of the BO search is slightly lower than or almost equal to that of the random search. This may indicate that exploration is prioritised to form a reliable prediction function. After the initial 10–15 steps, the probability of discovery ratio of the BO search increases

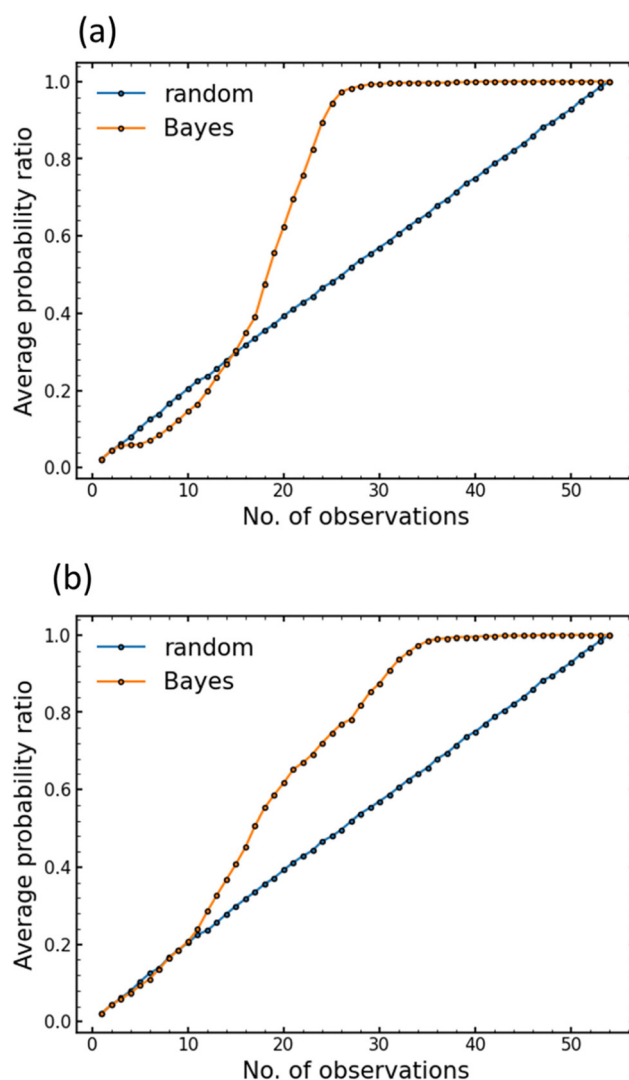


Fig. 6 Performance comparison between Bayesian optimisation (red line) and random search (blue line). (a) Target is the highest ion conductivity at 30 °C. (b) Target is the lowest activation energy.



rapidly, which corresponds to an exploitation search because of less uncertainty for the prediction function in the search space.

As shown in Fig. 6(a), BO may determine the optimal heating condition for ionic conductivity with a 90% probability in 25 experiments, with an efficiency factor of approximately double that of the exhaustive search (54 conditions). Similarly, 31 experiments were conducted to optimise the activation energy with a probability of 90%, as shown in Fig. 6(b). Accordingly, BO is significantly efficient in optimising the heating condition, in addition to the optimisation of chemical composition, as reported in our previous studies.^{24,36}

Conclusions

In this study, we synthesised LZP-CS under 54 different two-step heating conditions. At 30 °C, the HC-1050–1250 sample exhibited the highest Li-ion conductivity of $3.3 \times 10^{-5} \text{ S cm}^{-1}$, including bulk and grain-boundary contributions. Investigation of the sintered samples revealed that the heating conditions affected the crystalline phase, density, microstructure, and elemental distribution. Depending on the 1st and 2nd heating temperatures, the main phase was the α - and/or β -phase polymorph of LZP, accompanied by changes in particle morphology. Furthermore, the sintering conditions affected the necking between particles, pore size, and elemental distribution. The relationships between ionic conductivity and these observations were quite complex, and thus, intuitively determining the optimal heating conditions with a few experiments was challenging.

To reduce the number of experiments required to comprehensively evaluate 54 different two-step heating conditions, BO, which was suitable for the optimal condition search, was used. We used the temperatures of the 1st and 2nd heating as descriptors. BO could establish the optimum heating conditions for the highest Li-ion conductivity with a 90% probability in 25 experiments, which was approximately double the efficiency of the exhaustive search. Thus, BO is effective in efficiently optimising synthetic conditions that affect material properties, such as heating conditions. Note that when process parameters do not affect material properties, this trend is clarified using a smaller number of samples.

Author contributions

H. T., M. N., and Y. O. conceived and directed the project. H. T., H. F., N. T., and T. N. conducted the experimental work and analysed the results. H. T., K. N., and S. H. conducted the computational studies and analysed the results. All authors approved the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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