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Primary trifluoroborate-iminiums enable facile access to chiral α -aminoboronic acids *via* Ru-catalyzed asymmetric hydrogenation and simple hydrolysis of the trifluoroborate moiety†

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This work describes the first preparation and application of primary trifluoroborate-iminiums (pTIMs) as a new, easily accessible and valuable class of organoboron derivatives. An array of structurally diverse pTIMs was prepared from potassium acyltrifluoroborates in excellent yields. Highly efficient and enantioselective [(*R,R*)-TethTsDpen-RuCl] complex-catalyzed hydrogenation of pTIMs provided direct access to chiral primary trifluoroborate-ammoniums (pTAMs). Moreover, facile synthesis of a series of structurally diverse chiral α -aminoboronic acids from chiral pTAMs was accomplished through novel, operationally simple and efficient conversion using hexamethyldisiloxane/aqueous HCl. Using no chromatography at any point, this work allowed easy access to chiral α -aminoboronic acids, as exemplified by the synthesis of optically pure anti-cancer drugs bortezomib and ixazomib.

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Introduction

Organoboron compounds¹ play a pivotal role in synthetic organic chemistry, where they have enabled a myriad of useful transformations^{2–4} and have also been used in the field of materials science.⁵ The most advantageous application of organoboron compounds was achieved in medicinal chemistry⁶ where several boron-containing drugs reached the market in the last two decades. Among those the key role is attributed to proteasome-inhibiting⁷ anti-cancer drugs bortezomib⁸ and ixazomib.⁹ Furthermore, a structural analog delanzomib is in the development phase.¹⁰ These drugs possess a unique chemical structure in the sense that they all contain an α -aminoboronic acid moiety (Scheme 1). Beside their application in medicinal chemistry, α -aminoboronic acid derivatives found numerous applications in organic synthesis.¹¹ However, preparation of chiral α -aminoboronates is more challenging compared to their natural counterparts and methods for their preparation are limited. A seminal work on the synthesis of chiral α -aminoboronates was conducted by Matteson, who developed a powerful homologation reaction which provided chiral (α -chloroalkyl)boronic esters that could be transformed into α -

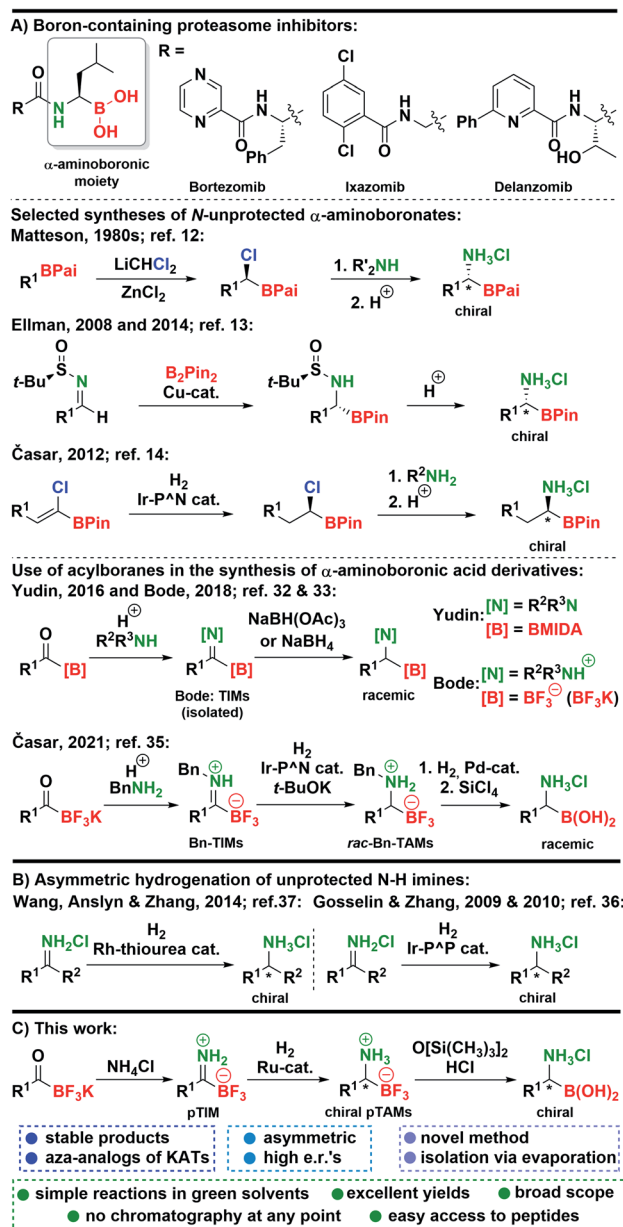
aminoboronic acids (Scheme 1).¹² In 2008, Ellman reported a novel approach to obtain chiral α -aminoboronic esters using highly diastereoselective Cu-catalyzed addition of bis(pinacolato)diboron to *N*-*tert*-butanesulfinyl aldimines (Scheme 1).¹³ Next, we reported an alternative asymmetric hydrogenation approach to Matteson's methodology to prepare chiral (α -chloroalkyl)boronic- and α -aminoboronic esters (Scheme 1).¹⁴ Subsequently, novel methodologies for the synthesis of α -aminoboronates were reported by groups of Yudin,¹⁵ Sawamura,¹⁶ Morken,¹⁷ and others.^{11,18} In recent years, this field has continued to flourish with many exciting methods being developed.^{19–33} Among recent methodologies, discoveries of Yudin³² and Bode³³ caught our attention. Namely, in these reports trifluoroborate-iminiums (TIMs) and MIDA-protected iminoboronates were prepared from acylboranes,³⁴ and converted to racemic *N*-substituted α -aminoboronic acid derivatives with borohydride reagents (Scheme 1). This stimulated us to explore Ir-catalyzed hydrogenation of benzyl-protected TIMs (Bn-TIMs) to benzyl-protected trifluoroborate-ammoniums (Bn-TAMs). Although a series of racemic Bn-TAMs was prepared and their conversion to α -aminoboronic acids was demonstrated (Scheme 1), an asymmetric approach was not successful as a consequence of *E/Z*-isomerism that was observed in Bn-TIMs.³⁵

Despite significant advances in the area, a need for new synthetic routes that would allow the preparation of chiral *N*-unprotected α -aminoboronic acids exists. Many of the existing methods are either racemic, allow limited scope of substrates, or do not enable the synthesis of *N*-unprotected derivatives.^{11,18–33,35}

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† Electronic supplementary information (ESI) available: Preparation of pTIMs, enantioenriched pTAMs, and α -aminoboronic acids, details regarding reaction screening and optimization of reactions, determination of optical purities, and ¹H, ¹³C, ¹¹B, and ¹⁹F NMR spectra. See DOI: 10.1039/d1sc07065g



Scheme 1 (A) Boron-containing proteasome inhibitors and key syntheses of α -aminoboronates. (B) Asymmetric hydrogenation of unprotected *N*-H imines. (C) This work.

Based on our recent research,³⁵ we envisioned that unsubstituted primary TIMs (pTIMs) could provide access to chiral α -aminoboronic acids using asymmetric hydrogenation. Moreover, the use of pTIMs would eliminate the *N*-deprotection step to obtain the free amino group, thereby shortening the overall synthetic sequence. However, the unprotected *N*-H imine functionality is considered challenging and thus underexplored in the area of asymmetric hydrogenation (Scheme 1),^{36–38} which suggests that asymmetric hydrogenation of pTIMs might be a demanding task. Herein, we present the results of our investigation that started with the preparation of a structurally diverse array of a novel subtype of TIM, *i.e.* pTIMs, which underwent Ru-catalyzed asymmetric hydrogenation to chiral

pTIMs, followed by mild and efficient conversion of the tri-fluoroborate moiety to boronic acid functionality (Scheme 1).

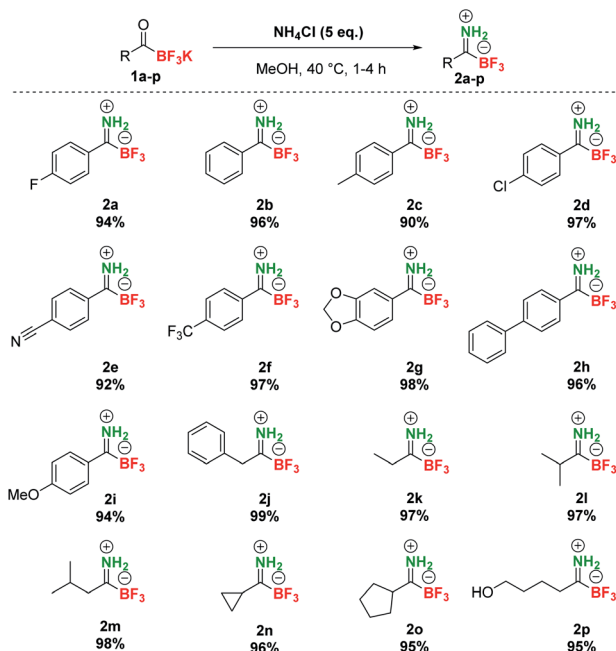
Results and discussion

At the outset of our studies, we investigated the preparation of pTIMs. Although secondary and tertiary TIMs were prepared before by Bode³³ and Bn-TIMs in our group,³⁵ the formation of pTIMs was not an obvious task. We began our research using potassium acyltrifluoroborate (KAT) **1a** as a model substrate for the formation of the corresponding pTIM **2a** upon reaction with NH₄Cl as a simple nitrogen precursor (Scheme S2†).

We initially screened various reaction conditions to establish the viability of the envisioned approach, using NH₄Cl as an environmentally benign and non-toxic nitrogen source. Interestingly, no reaction took place between **1a** and NH₄Cl (2 eq.) in MeCN and DMF even at 60 °C in the presence of molecular sieves (3 Å) (Table S1, entries 1–3†), although these solvents enabled the preparation of secondary and tertiary TIMs.^{33,35} More encouraging results with 38% yield of **2a** in 20 h (0.2 M concentration) were observed in MeOH at room temperature (Table S1, entry 4†). Interestingly, the reaction conducted at 0.05 M concentration afforded 62% conversion due to increased solubility of NH₄Cl (Table S1, entry 5†). Increasing the temperature to 40 °C gave full conversion both at 0.2 M and 0.05 M concentration, albeit at the expense of 5–10% of unknown side product formation (Table S1, entries 6 and 7†) and a similar outcome was observed with 1 eq. NH₄Cl (Table S1, entry 8†). By increasing the amount of NH₄Cl to 5 eq., a clean reaction was obtained with 100% conversion in 4 h (Table S1, entry 9†). Finally, even without molecular sieves, a clean reaction was observed between **1a** (0.08 M) and NH₄Cl (5 eq.) in MeOH at 40 °C and afforded pTIM **2a** after 4 h in 94% yield (Table S1, entry 10†). With optimal reaction conditions in hand, we converted a series of aromatic and aliphatic KATs **1a–2p** to pTIMs **2a–2p**. Overall, these reactions proceeded smoothly and provided aromatic and aliphatic pTIMs **2** in excellent 90–99% yields and high purity (Scheme 2).

After the preparation of a series of pTIMs **2**, we focused our attention on their asymmetric hydrogenation to chiral pTIMs **3** (Table 1). Initially, we decided to explore a diverse set of catalysts based on Ir, Ru and Rh metal precursors in combination with various P^P and P^N-ligands (Fig. S3†), as well as commercially available Noyori–Ikariya-type catalysts (Fig. S4†). The initial hydrogenation screening was done on substrate **2a** at 20 bar H₂ using Ir, Ru and Rh metal precursors (2 mol%) in combination with (*S*)-DTBM-MeO-BIPHEP (**L1**) and (*R,R*)-Ph₂P-ThrePHOX (**L2**) ligands which provided good conversion in hydrogenation of Bn-TIMs to Bn-TAMs in the presence of *t*-BuOK base,³⁵ albeit at moderate e.r.s. In contrast to Bn-TIMs, hydrolysis of **2a** to **1a** or decomposition of **2a** was observed with selected catalytic systems when *t*-BuOK was used as an additive (Table 1, entries 1–4; Table S2, entries 1–6†) or no conversion of pTIM **2a** to pTAM **3a** was noticed in the absence of *t*-BuOK (Table 1, entries 5 and 6; Table S2, entries 7–10†). Encouraging first hit was obtained with the catalyst based on [Ru(*p*-cymene)Cl₂]₂ and (*S*)-DTBM-MeO-BIPHEP (**L1**) which





Scheme 2 Synthesis of pTIMs 2 from KATs 1.

enabled 23% conversion to **3a** (racemic) in the absence of *t*-BuOK (Table 1, entry 7; Table S2, entry 11[†]). Furthermore, reaction using **RuBF₄(*p*-cymene)(pyr)[(R,R)-TsDpen]** provided 89% conversion to **3a** and a promising e.r. of 79 : 21 when *t*-BuOK was not used (Table 1, entry 8; Table S2, entry 12[†]), which is in sharp contrast with a similar reaction using *t*-BuOK (Table 1, entry 4, Table S2; entry 6[†]) where only hydrolysis of **2a** to **1a** was observed. These results indicated that the base additive is detrimental to hydrogenation of pTIMs **2** and that only Ru-based catalysts are feasible for this reaction (Scheme 3). Currently, the price of Ru metal is more than seven times lower than that of Ir metal and thirtyfold lower compared to Rh metal.³⁹ This implies that hydrogenation of pTIMs is not only a more atom economic approach, but also requires a less expensive metal catalyst than in the case of Bn-TIMs. For these reasons, we focused further on Ru-based catalysis.

Therefore, additional extensive catalyst screening was conducted using Ru-metal precursors in combination with various P[^]P and P[^]N-ligands and an extended list of Noyori–Ikariya-type catalysts⁴⁰ (Table S3[†]). The study of ruthenium catalysts revealed that Ru-(*p*-cymene)-phosphine ligand complexes showed very low activity and no enantioselectivity in the hydrogenation of substrate **2a** (Table S3, entries 2–14[†]). The RuOAc₂-phosphine complexes showed limited enantioselectivity and also low conversions (Table S3, entries 15 and 16[†]), while the RuCl-diphosphine-diamine complexes performed slightly better in terms of enantioselectivity, but with similarly meagre conversions (Table S3, entries 17–24[†]). The dimethylamine adducts of RuCl-phosphine complexes were slightly more active but gave low enantioselectivity (Table S3, entries 25–27[†]). In contrast to these experiments, the best results for conversion of **2a** to **3a** were achieved with Noyori–Ikariya-type

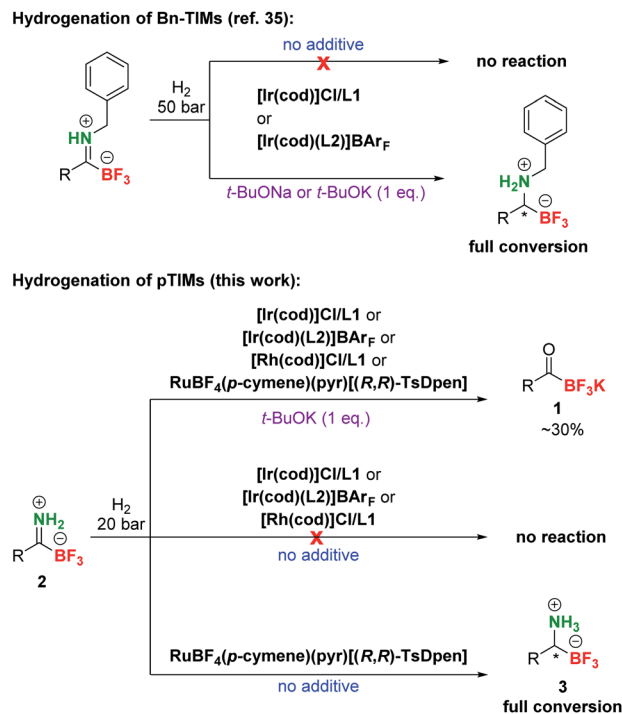
Table 1 Asymmetric hydrogenation of model pTIM **2a**^a

Entry ^b	Catalyst	3a [%]	e.r.
1 ^c	[Ir(cod)]Cl/L1	0 ^d	n.a.
2 ^c	[Ir(cod)(L2)]BAR _F	0 ^d	n.a.
3 ^c	RuCl[(<i>p</i> -cymene)]Cl/L1	0 ^d	n.a.
4 ^c	RuBF₄(<i>p</i>-cymene)(pyr)[(R,R)-TsDpen]	0 ^d	n.a.
5	[Ir(cod)]Cl/L1	0 ^e	n.a.
6	[Ir(cod)(L2)]BAR _F	0 ^e	n.a.
7	RuCl[(<i>p</i> -cymene)]Cl/L1	23	50 : 50
8	RuBF₄(<i>p</i>-cymene)(pyr)[(R,R)-TsDpen]	89	79 : 21
9	RuCl(<i>p</i>-cymene)[(R,R)-TsDpen]	81	89 : 11
10	RuBF₄(<i>p</i>-cymene)[(R,R)-TsDpen]	100	89 : 11
11	[(R,R)-TethTsDpen-RuCl]	100	97 : 3
12	(R,R)-TsDENEB	88	92 : 8
13	RuCl(<i>p</i>-cymene)[(R,R)-FsDpen]	36	91 : 9
14	RuCl(mesitylene)[(R,R)-TsDpen]	40	88 : 12
15 ^f	RuBF₄(<i>p</i>-cymene)(pyr)[(R,R)-TsDpen]	100	77 : 23
16 ^g	RuBF₄(<i>p</i>-cymene)(pyr)[(R,R)-TsDpen]	100	80 : 20
17 ^g	RuCl(<i>p</i>-cymene)[(R,R)-TsDpen]	100	88 : 12

^a Cod = cyclooctadiene, BAR_F = tetrakis(3,5-bis(trifluoromethyl)phenyl) borate. ^b Reactions were conducted by using **2a** (0.1 mmol) and catalyst (2 mol%) in MeOH (0.05 M) with 20 bar H₂ at room temperature for 20 h. The yields were determined by ¹H-NMR and the e.r. (*S*/*R*) by chiral HPLC. ^c *t*-BuOK (1 eq.) as additive. ^d ~30% conversion to **1a** observed. ^e No reaction observed. ^f With 50 bar H₂. ^g Reaction time 44 h.

catalysts (Table 1). **RuCl(*p*-cymene)[(R,R)-TsDpen]** performed with slightly lower conversion compared to **RuBF₄(*p*-cymene)(pyr)[(R,R)-TsDpen]**, but gave higher e.r. (Table 1, entry 8 vs. entry 9). Furthermore, a full conversion was achieved with **RuBF₄(*p*-cymene)[(R,R)-TsDpen]** along with a good e.r. (89 : 11, Table 1, entry 10). Gratifyingly, Wills's tethered catalyst **[(R,R)-TethTsDpen-RuCl]**⁴¹ provided full conversion and an excellent

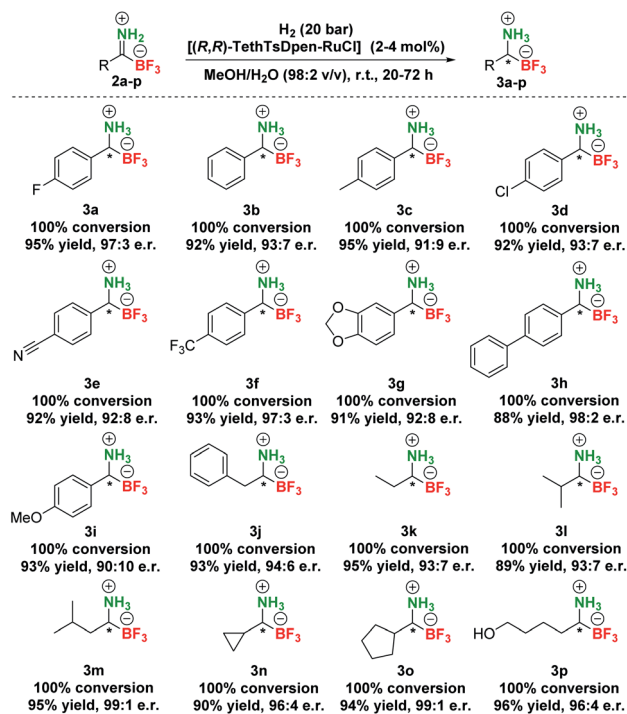




Scheme 3 Comparison of different reactivities of Bn-TiMs and pTiMs. Under the conditions that were successful in hydrogenation of the former, only partial hydrolysis to KAT was observed with model pTiM and an altogether different catalytic system was required. Catalyst and ligand structures are presented in Table 1.

97 : 3 e.r. (Table 1, entry 11). Catalysts **(R,R)-TsDENEB®**, **RuCl(p-cymene)[(R,R)-FsDpen]**, and **RuCl(mesitylene)[(R,R)-TsDpen]** also afforded good enantioselectivities, albeit at incomplete conversions (Table 1, entries 12–14). Reactions with incomplete conversion could be completed by using higher hydrogen pressure (50 bar, Table 1, entry 15 vs. entry 8) with marginal loss of enantioselectivity, or by prolonging the reaction time to 44 h with retained e.r. (Table 1, entries 16 and 17 vs. entries 8 and 9).

Based on these screening results, **[(R,R)-TethTsDpen-RuCl]** was chosen as the optimal catalyst for the asymmetric reduction of TiMs 2 to TAMs 3. Next, the effect of the solvent was carefully examined (Table S4†), and MeOH was confirmed as the optimal choice. In addition, a mixture of MeOH/H₂O (1–10% v/v) performed similarly at full conversion and nearly the same enantioselectivity, while the MeOH/H₂O = 1 : 1 mixture or pure H₂O promoted notable hydrolysis of **2a** to **1a** (Table S4, entries 1–9†). Furthermore, we established that concentration had no effect on conversion in the range of 0.025–0.1 M. Of note, the highest enantioselectivities were achieved at 0.05–0.10 M concentration (Table S4, entries 10–12†). Variations in **[(R,R)-TethTsDpen-RuCl]** loading were also examined (0.1–2.0 mol%) and while the reaction proceeded to completion also at 1 mol% catalyst (Table S4, entries 13–15†), 2 mol% catalyst loading was chosen in hydrogenation of other compounds to ensure full conversions across the whole substrate scope. Next, we proceeded with preparative hydrogenation of pTiMs **2a–p**. Hydrogenation of aromatic derivatives **2a–h** proceeded smoothly with full



Scheme 4 Scope of asymmetric hydrogenation of pTiMs 2.

conversion at room temperature and 20 bar H₂ with 2 mol% of **[(R,R)-TethTsDpen-RuCl]** in MeOH containing 2% water (v/v) giving the corresponding pTAMs **3a–i** in excellent yields (88–95%) and high enantioselectivities with e.r.s ranging from 98 : 2 for **3h** to 90 : 10 for **3i** (Scheme 4). Interestingly, when aliphatic pTiM **2j** was subjected to hydrogenation under the same conditions, low conversion (45%) was obtained (Table S5, entry 1†). Further optimization of the reaction conditions by increasing the catalyst loading to 4 mol% and prolongation of the reaction time to 72 hours provided **3j** with a 94 : 6 e.r. at full conversion (Table S5, entry 4†). By applying these conditions, aliphatic pTiMs **2j–p** were successfully hydrogenated and provided pTAMs **3j–p** in 89–96% yields and excellent e.r.s that were in the range of 99 : 1 to 93 : 7, which is remarkable for such small compounds.^{36,37} Overall, aliphatic substrates gave slightly higher enantioselectivities compared to aromatic counterparts⁴² and provided the opposite enantiomer⁴³ (Scheme 4).

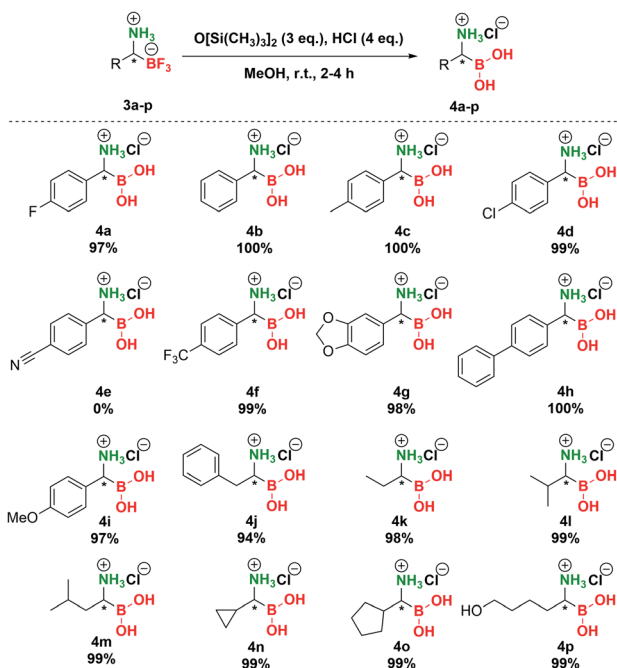
In the final stage of our investigation, we studied the conversion of the trifluoroborate moiety in pTAMs 3 to the boronic acid functionality. Although hydrolysis of organic trifluoroborates to boronic acids in the presence of Cs₂CO₃ was extensively studied in connection with Suzuki–Miyaura coupling,⁴⁴ general methods for the preparative hydrolysis of organic trifluoroborates to boronic acids are scarce and involve TMSCl/H₂O,⁴⁵ TMSCl/K₂CO₃/MeCN,⁴⁶ silica gel/H₂O,⁴⁷ FeCl₃/THF/H₂O,⁴⁸ alumina/H₂O,⁴⁹ and SiCl₄/MeOH + HCl_(aq.).³³ This is associated with the fact that structurally diverse organic trifluoroborates exhibit notably different hydrolysis rates.^{50,51} As we have shown recently, only Bode's SiCl₄ method³³ enabled the conversion of pTAMs to the aminoboronic acids while other methods failed.³⁵ However, huge excess (10 eq.) of difficult-to-



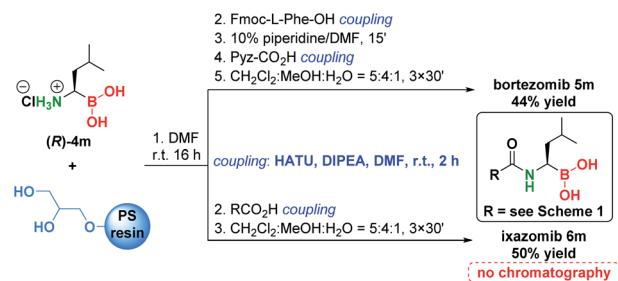
handle SiCl_4 was needed to transform pTAMs to aminoboronic acids.

Therefore, we decided to investigate other options for conversion of pTAMs **3** to α -aminoboronic acids **4** (Table S6†). Reactions using Cs_2CO_3 or silica gel/ Cs_2CO_3 in MeOH or MeOH/ H_2O (95 : 5 v/v) gave no conversion of **3** to **4** (Table S6, entries 1–4†). The same was observed with $\text{TMSCl}/\text{Cs}_2\text{CO}_3$ in MeOH (Table S6, entry 5†). Surprisingly, $\text{TMSCl}/\text{Cs}_2\text{CO}_3$ in MeOH/ H_2O (95 : 5 v/v) provided 100% conversion of **3a**, but the product **4a** was contaminated with a side product (Table S6, entry 6†). Finally, replacement of the silicon source with hexamethyldisiloxane (HMDSO, $\text{O}[\text{Si}(\text{CH}_3)_3]_2$) provided unexpected results (Table S6, entries 7–9†) and to our delight reaction of **3a** with HMDSO (3 eq.) in the presence of aqueous HCl (4 eq.) in MeOH provided complete conversion to **4a** (Table S6, entry 8†). It should be noted that HMDSO is non-toxic and has been studied as an MRI probe, where no toxic effects were observed when it was injected in mammals.⁵² Subsequent preparative transformation of pTAMs **3** to the corresponding α -aminoboronic acids **4** proceeded smoothly in excellent 94–100% yields (Scheme 5) without the loss of stereochemistry⁴² except for **3e** where decomposition of the corresponding boronic acid to the benzyl amine was observed (Scheme S7 and Fig. S9†). Importantly, these highly polar and potentially labile compounds were isolated in pure form only by evaporation of volatiles.⁵³

To demonstrate the significance of the presented methodology, we converted α -aminoboronic acid (**R**)-**4m** to bortezomib and ixazomib using a modified solid phase coupling approach (Scheme 6).⁵⁴ Pure bortezomib and ixazomib were obtained in 44% and 50% yield respectively, which represents overall one of the shortest, simplest and most efficient approaches to these drugs (see Schemes S10–S15†).



Scheme 5 Conversion of trifluoroborates **3** to boronic acids **4**.



Scheme 6 Conversion of (**R**)-**4m** to bortezomib and ixazomib.

Conclusions

In summary, we established a novel three-step synthetic strategy to a diverse array of chiral α -aminoboronic acids. The presented synthetic methodology is based on the conversion of KATs to their direct aza-analogs pTIMs **2**, which are disclosed for the first time and could prove to be valuable organoboron reagents in their own right. Afterwards, highly stereoselective asymmetric hydrogenation of pTIMs **2** to chiral pTAMs **3** using the [(**R,R**)-TethTsDpen-RuCl] catalyst was successfully accomplished, followed by a novel and mild conversion of trifluoroborate to the boronic acid functionality in the presence of HMDSO and aqueous HCl. The represented transformations feature various aliphatic and aromatic substrates, proceed efficiently in high to excellent yields, are conducted in MeOH and water as sustainable solvents, and enable protecting-group free access to chiral α -aminoboronic acids. Our methodology towards these valuable building blocks will enable their broad deployment in access to libraries of novel proteasome inhibitors and beyond,¹¹ as demonstrated by the synthesis of stereochemically pure drugs bortezomib and ixazomib.

Data availability

Data for this paper, including reactions screening and optimization experiments, detailed experimental procedures, overview of key synthetic procedures to bortezomib and characterization data, are provided in the ESI.†

Author contributions

A. Š. and Z. Č. conceived the study. A. Š. performed all experiments and collected the data. I. S. and Z. Č. guided the research. Z. Č. wrote the manuscript. A. Š. and I. S. performed editing of the manuscript. All authors discussed the results and contributed to the finalization of the paper.

Conflicts of interest

There are no conflicts to declare.

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