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SrO-layer insertion in Ruddlesden–Popper Sn-based perovskite enables efficient CO₂ electroreduction towards formate†

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Tin (Sn)-based oxides have been proved to be promising catalysts for the electrochemical CO₂ reduction reaction (CO₂RR) to formate (HCOO⁻). However, their performance is limited by their reductive transformation into metallic derivatives during the cathodic reaction. This paper describes the catalytic chemistry of a Sr₂SnO₄ electrocatalyst with a Ruddlesden–Popper (RP) perovskite structure for the CO₂RR. The Sr₂SnO₄ electrocatalyst exhibits a faradaic efficiency of 83.7% for HCOO⁻ at -1.08 V vs. the reversible hydrogen electrode with stability for over 24 h. The insertion of the SrO-layer in the RP structure of Sr₂SnO₄ leads to a change in the filling status of the anti-bonding orbitals of the Sn active sites, which optimizes the binding energy of *OCHO and results in high selectivity for HCOO⁻. At the same time, the interlayer interaction between interfacial octahedral layers and the SrO-layers makes the crystalline structure stable during the CO₂RR. This study would provide fundamental guidelines for the exploration of perovskite-based electrocatalysts to achieve consistently high selectivity in the CO₂RR.

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Introduction

The electrochemical CO₂ reduction reaction (CO₂RR) is an attractive and sustainable approach to mitigate the greenhouse effect by storing renewable energy (*e.g.*, solar and wind energy) in the form of chemical bonds.^{1–3} The CO₂RR can yield a variety of valuable products, among which formate (HCOO⁻) shows the advantages of environmental compatibility and transportability as an important raw material.^{4,5} It is also a promising material for liquid hydrogen storage for fuel cell applications.^{5,6} HCOO⁻ production by the CO₂RR has been proved to be predominately realized by tin(Sn)-based electrocatalysts.^{7–12} The emerging evidence suggested that Sn oxides exhibit high activity.^{7,13} Besides, Sn-based electrocatalysts have the advantages of low cost, abundance in the earth's crust and environmental friendliness.^{7–9} Nevertheless, their selectivity to HCOO⁻ would drop with the proceeding of the CO₂RR due to their reductive transformation into metallic derivatives.^{14,15} Thence, there is an

urgent need to design Sn-based oxides with consistently high selectivity.

The electrocatalytic performance of Sn-based oxides can be promoted through compositional and structural engineering, which can tune the electrical properties and coordination stability of the active sites.^{16–19} For instance, N doping had been conducted to regulate the charge density of the Sn active sites in SnO₂ to enhance the selectivity towards HCOO⁻.¹⁹ Besides, SrSnO₃ has been proved to show high performance for the CO₂RR to HCOO⁻ as a result of the perovskite structure.¹⁶ However, the stability of SrSnO₃ is still not satisfactory. The deactivation mechanism of such ABO₃ perovskite electrocatalysts was proposed to be the surface amorphization or reconstruction during the electrocatalytic processes.^{20–22} Therefore, effective strategies to stabilize the crystal structures of perovskites would help further enhance the performance of Sn-based electrocatalysts. The insertion of rock-salt AO-layers in Ruddlesden–Popper (RP) perovskite not only induces interlayer interaction to stabilize the crystal structure, but also optimizes the electrical properties of the active B-site.^{23–26}

This paper describes the design and synthesis of two-dimensional layered Sr₂SnO₄ with a RP structure for the selective conversion of CO₂ to HCOO⁻. Compared with the SnO₂ and SrSnO₃ samples, Sr₂SnO₄ shows enhanced selectivity with the Faradaic Efficiency (FE) for HCOO⁻ reaching 83.7% (-1.08 V vs. the reversible hydrogen electrode). Besides, Sr₂SnO₄ exhibits long-term stability for over 24 h. As revealed by the experiments and density functional theory (DFT) calculations, the SrO-layer in the RP structure is the key to achieving good performance,

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which could optimize the binding energy of *OCHO and promote structural stability.

Results and discussion

Structures of catalysts

Fig. 1a shows the X-ray diffraction (XRD) patterns of SnO₂, SrSnO₃, and Sr₂SnO₄, the diffraction peaks of which are readily indexed to tetragonal SnO₂ (*P42mm*, JCPDS No. 41-1445), cubic SrSnO₃ (*Pm3m*, JCPDS No. 74-1298) and tetragonal Sr₂SnO₄ (*I4/mmm*, JCPDS No. 97-002-7114), respectively. No peak corresponding to other phases was detected, demonstrating the high phase purity of SnO₂, perovskite SrSnO₃, and RP perovskite Sr₂SnO₄ catalysts. The scanning electron microscopy (SEM) images reveal the porous structures of the samples (Fig. S1†). In addition, the electron dispersive X-ray spectroscopy (EDS) mappings exhibit uniform distributions of Sr, Sn, and O elements (Fig. S2 and S3†). To further confirm the formation of the RP phase, the high-resolution transmission electron microscopy (HRTEM) image of Sr₂SnO₄ was collected (Fig. 1b). The observed lattice fringes with a spacing of 0.28 nm and 0.20 nm correspond to the (110) and (200) planes of Sr₂SnO₄. The inset shows the result of fast Fourier transform (FFT) analysis of the HRTEM image, which further confirms the tetragonal structure of Sr₂SnO₄. Moreover, the selected area electron diffraction (SAED) pattern of Sr₂SnO₄ clearly reflects the tetragonally arranged diffraction spots of the [001] zone axes (Fig. 1c), consistent with the result of the HRTEM characterization.

Raman spectroscopy was used to investigate the differences in structures between the three Sn-based oxides (Fig. 1d). The bands centered at 472, 631 and 771 cm⁻¹ of SnO₂ correspond to the vibration modes of E_g, A_{1g}, and B_{2g}, respectively.²⁷ The band of SrSnO₃ at 399 cm⁻¹ corresponds to SnO₃ torsional mode, and the peak at 576 cm⁻¹ is associated with Sn–O stretching vibrations, indicating the formation of [SnO₆] octahedra.^{28,29} The band of Sr₂SnO₄ located at 445 cm⁻¹ is related to the vibration

of the bridge Sr–O–Sn oxygen bond axis in A_{1g} symmetry along the *c* axis. This band shows the interactions between the perovskite layers and the rock-salt SrO-layers in the RP structure.³⁰ The differences in the crystal structures of the three oxides are illustrated in Fig. 1e. For SrSnO₃, Sr²⁺ ions occupy the A-sites, whereas Sn⁴⁺ ions form the corner-linked [SnO₆] octahedra at the B-sites. Sr₂SnO₄ is composed of rock-salt SrO-layers and SrSnO₃ perovskite layers along the crystallographic *c*-axis direction. This alternating layered RP structure would lead to unique structural and electrical properties, which could be beneficial for promoting electrocatalytic performance.^{31,32}

CO₂RR performance

The CO₂RR performance was measured in a three-electrode H-type cell, with CO₂-saturated 0.5 M KHCO₃ as electrolyte. The products were analysed at potentials between –0.88 and –1.28 V vs. the reversible hydrogen electrode (RHE). All the following potentials are referred to the RHE. Sr₂SnO₄ exhibits a stable current density of up to 30.0 mA cm⁻² at –1.28 V, surpassing those of the SrSnO₃ (22.9 mA cm⁻²) and the SnO₂ (17.8 mA cm⁻²) electrocatalysts (Fig. S4a, S5a, and S6†). The FE for HCOO⁻ against potential over Sr₂SnO₄ shows a volcano trend, which first increases from 46.5% at –0.88 V to a peak value of 83.7% at –1.08 V. It then decreases to 62.7% at –1.28 V (Fig. 2a). The FEs for HCOO⁻ over SnO₂ and SrSnO₃ also show similar volcano trends (Fig. S4b and S5b†), but with much lower peak values (64.6% for SrSnO₃ and 58.9% for SnO₂). Besides, Sr₂SnO₄ also exhibits the highest HCOO⁻ partial current density of 12.4 mA cm⁻² at –1.08 V compared to SrSnO₃ (9.0 mA cm⁻²) and SnO₂ (5.9 mA cm⁻²) (Fig. S7†). The stability tests were performed at a consistent potential of –1.08 V (Fig. 2c). Sr₂SnO₄ displays no obvious decay in the current density with the FE for HCOO⁻ maintained at about 79.0%. However, there is an obvious increase in the current density for SrSnO₃ and SnO₂, with a decrease in HCOO⁻ FE at the same time. Although the

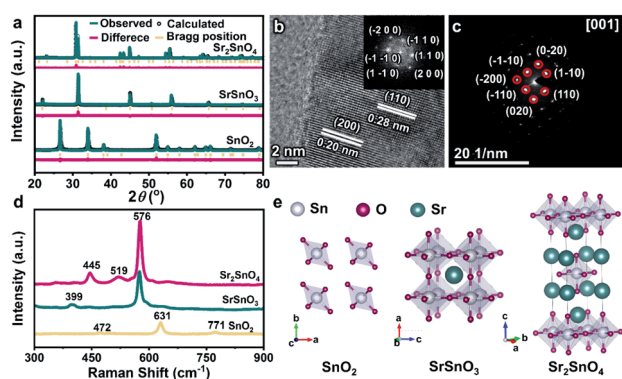


Fig. 1 (a) The XRD patterns of SnO₂, SrSnO₃, and Sr₂SnO₄ with their Rietveld refinement results. (b) HRTEM and FFT (inset) images of Sr₂SnO₄. (c) The SAED patterns along [001] of Sr₂SnO₄. (d) The Raman spectra of SnO₂, SrSnO₃, and Sr₂SnO₄. (e) Illustrations of the crystal structures of SnO₂, SrSnO₃, and Sr₂SnO₄. Color code: O (purple), Sr (turquoise), and Sn (gray).

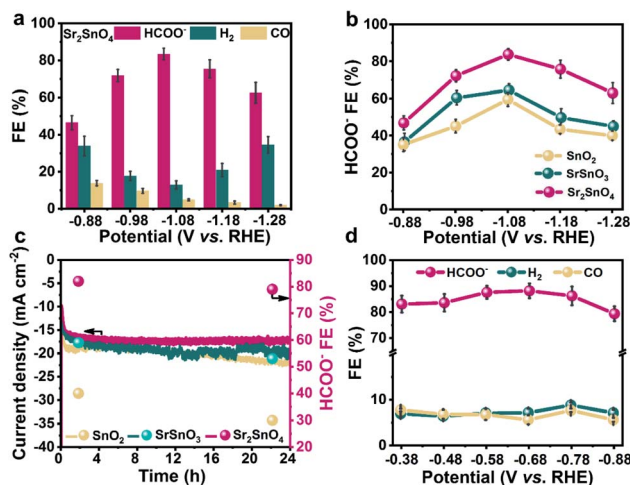


Fig. 2 (a) FEs for various products over Sr₂SnO₄. (b) FEs for HCOO⁻ over SnO₂, SrSnO₃ and Sr₂SnO₄ at different potentials. (c) Stability tests of SnO₂, SrSnO₃ and Sr₂SnO₄ at a potential of –1.08 V vs. RHE. (d) FEs for various products over Sr₂SnO₄ in a flow cell.



selectivity of these Sn-based electrocatalysts may not reach the high value reported in some representative studies,^{33,34} the comparison of their performance could reflect the structural advantages of the RP perovskite. The electrocatalytic performance of Sr₂SnO₄ was also evaluated in a flow cell with the catalysts dispersed on a gas diffusion electrode (GDE) to enhance the mass transfer of CO₂. The peak FE towards HCOO⁻ over Sr₂SnO₄ is 84.5% at -0.68 V (Fig. 2d), and the current density for HCOO⁻ reaches about 79.2 mA cm⁻² at -0.88 V (Fig. S8†). Moreover, Sr₂SnO₄ displays long-term stability for over 50 h with the HCOO⁻ partial current density maintained at about 59.1 mA cm⁻² (Fig. S9†). The electrolyte was replaced during the chronoamperometric test to avoid the formation of (bi)carbonate salts on the diffusion layer, which caused the strong fluctuation of current density. The FEs towards H₂ and CO are below 10.0% with negligible change. According to the results of the activity evaluation, the Sr₂SnO₄ with the RP perovskite structure achieved consistently high selectivity towards HCOO⁻. The performance of Sr₂SnO₄ is comparable to some of the state-of-the-art electrocatalysts for HCOO⁻ production through the CO₂RR (Table S1†).

Operando Raman and ATR-SEIRAS spectra

To reveal the origin of the high stability of Sr₂SnO₄, *operando* Raman spectra of different samples were collected using a homemade electrochemical cell (Fig. S10†). With the decrease of the potential from -0.58 V to -1.28 V, the A_{1g} peak of crystalline SnO₂ at 631 cm⁻¹ declines in intensity, which finally disappears after -1.08 V (Fig. 3a). This phenomenon would be caused by the reduction of SnO₂ to metallic Sn at high cathodic potentials, as proved by the formation of the Sn phase after the CO₂RR stability test (Fig. S11a†).¹⁴ In the *operando* Raman spectra of SrSnO₃, the appearance of peaks at 631 cm⁻¹ below -1.08 V suggests the surface reconstruction of SrSnO₃ to form SnO₂ during the CO₂RR, which would then be partially reduced to metallic Sn (Fig. 3b). This result could be supported by the XRD pattern of SrSnO₃ after the stability test, which clearly shows the presence of SnO₂ and Sn phases (Fig. S11b†).¹⁶ In contrast, the Raman bands of Sr₂SnO₄ at 445 cm⁻¹ arising from the interactions between the perovskite layer and rock-salt SrO-layer remain unchanged at different potentials, same as the bands corresponding to the [SnO₆] octahedra at 576 cm⁻¹

(Fig. 3c). These phenomena are in agreement with the end-of-test XRD pattern, showing diffraction peaks of only Sr₂SnO₄ (Fig. S11c†). The HRTEM and SEM images of the three Sn-based oxide catalysts after the CO₂RR stability test have been characterized. There was a significant change in the morphology of SnO₂ (Fig. S12a†). Such a change could be ascribed to the reduction of SnO₂ to metallic Sn (Fig. S13†). Similar phenomena could also be found for SrSnO₃, which could be reduced to SnO₂ and Sn (Fig. S12b and S14†). However, the morphology of the Sr₂SnO₄ electrocatalyst was maintained after the CO₂RR stability test (Fig. S12c†). Besides, the HRTEM image showed that the crystalline nature of Sr₂SnO₄ was maintained (Fig. S15†). EDS mappings of Sr₂SnO₄ after the stability test exhibit uniform distributions of Sr, Sn, and O elements (Fig. S16†). The results above demonstrate the superior stability of Sr₂SnO₄ during the CO₂RR. The high durability of Sr₂SnO₄ would be induced by the SrO-layer insertion in the RP structure, which leads to the strong interaction between perovskite layers and rock-salt SrO-layers to prohibit structural deformation and reconstruction.²⁶

In addition to good stability, the RP perovskite Sr₂SnO₄ electrocatalyst shows high selectivity towards HCOO⁻. To find the underlining mechanism of the promoted electrocatalytic performance, *operando* attenuated-total-reflection surface-enhanced infrared absorption spectroscopy (ATR-SEIRAS) characterization was conducted. The peaks at around 1630 cm⁻¹ observed in all three samples can be assigned to the interfacial H₂O (Fig. 4a-c).^{9,35} The bands at 1367–1358 cm⁻¹, 1397–1390 cm⁻¹ and 1390–1386 cm⁻¹ for SnO₂, SrSnO₃ and Sr₂SnO₄ with an obvious Stark effect (peak position shifted to a lower wavenumber at more negative potentials) could be assigned to *OCHO adsorbed on the surface.^{36,37} The *OCHO peaks of SnO₂ can be observed from -0.48 V to -1.28 V. For SrSnO₃, the signal of *OCHO appears at -0.68 V and disappears at -1.18 V. The signal of *OCHO was detected from -0.58 V to -0.88 V for Sr₂SnO₄. Since *OCHO is the essential intermediate for the generation of HCOO⁻,³⁸ its rapid desorption on Sr₂SnO₄ would be the result of the weak adsorption of *OCHO, considering its high selectivity towards HCOO⁻.^{39,40} Therefore, the high selectivity of Sr₂SnO₄ would originate from the optimized binding energy of *OCHO on the surface, which might be realized by the SrO-layer insertion in the RP perovskite structure.

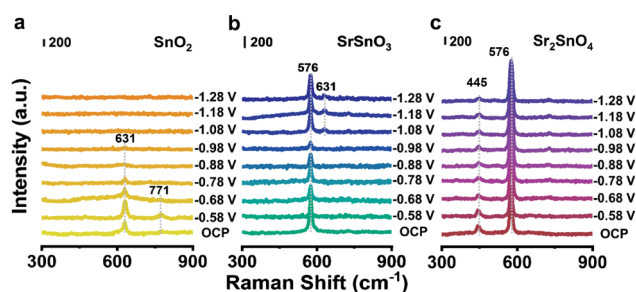


Fig. 3 *Operando* Raman spectra of (a) SnO₂, (b) SrSnO₃ and (c) Sr₂SnO₄ recorded at various potentials vs. RHE in a CO₂-saturated 0.5 M KHCO₃ solution, respectively.

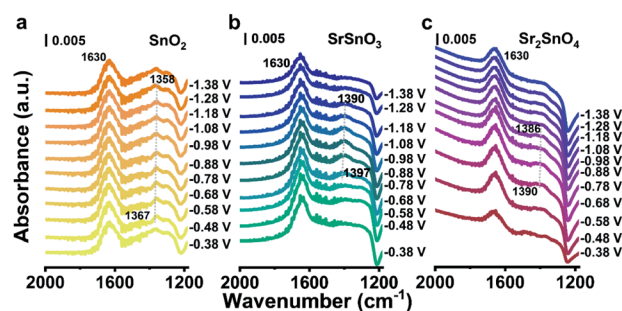


Fig. 4 *Operando* ATR-SEIRAS spectra of (a) SnO₂ (b) SrSnO₃ and (c) Sr₂SnO₄ recorded at various potentials vs. RHE in a CO₂-saturated 0.5 M KHCO₃ solution, respectively.



DFT simulations

To further reveal the intrinsic role of the RP structure of Sr_2SnO_4 in promoting selectivity, DFT simulations were performed. Thermodynamically stable $\text{SnO}_2(110)$,⁴¹ $\text{SrSnO}_3(001)$ (Table S2†) and SrO-terminated $\text{Sr}_2\text{SnO}_4(002)$ surfaces²⁴ were chosen as the starting models for the computational study. The influence of surface vacancies, which significantly affect the surface catalytic reactions, was considered due to the relatively large energy required for the formation of bulk oxygen vacancies in these samples (Tables S3 and S4†). According to their surface phase diagrams (Fig. S17†), $\text{SnO}_2(110)/4\text{O}_{\text{br}}\text{H}^*$,⁴¹ $\text{SrSnO}_3(001)/7\text{Ov}$ and $\text{Sr}_2\text{SnO}_4(002)/8\text{Ov}$ were found to be stable in the CO_2RR process. These models were selected for the understanding of the reaction mechanism, which would represent the real active sites of the catalysts during the reaction. As shown in the well-discussed reaction pathway of the CO_2RR to HCOO^- (Fig. S18†), the $^*\text{OCHO}$ intermediate is formed by the hydrogenation of CO_2 , which subsequently produces HCOOH by the transfer of electrons and protons.^{41,42} The results of DFT simulations show that the formation energy of $^*\text{OCHO}$ over the $\text{Sr}_2\text{SnO}_4(002)/8\text{Ov}$ surface (0.20 eV) is higher than those over the $\text{SrSnO}_3(001)/7\text{Ov}$ and the $\text{SnO}_2(110)$ (Fig. 5a). The process of $^*\text{OCHO}$ desorption to form HCOOH on $\text{Sr}_2\text{SnO}_4(002)/8\text{Ov}$ is spontaneous. However, the energy barriers for $^*\text{OCHO}$ desorption over $\text{SnO}_2(110)$ and $\text{SrSnO}_3(001)/7\text{Ov}$ are around 0.70 eV and 0.39 eV. Therefore, Sr_2SnO_4 undergoes easy desorption of $^*\text{OCHO}$ for the CO_2RR to HCOO^- , which is in accordance with the results of the *operando* ATR-SEIRAS characterization. Besides, the ΔG_{H^*} on $\text{Sr}_2\text{SnO}_4(002)/8\text{Ov}$ is higher than those on the $\text{SrSnO}_3(001)/7\text{Ov}$ and the $\text{SnO}_2(110)$ (Fig. 5b), as reflected by the relatively low selectivity towards H_2 on Sr_2SnO_4 (Fig. 2a, S4b and S5b†). Since the thermodynamically stable SrO-layer in $\text{Sr}_2\text{SnO}_4(002)$ may be involved in the reaction (Fig. S19†), its role in the CO_2RR was investigated, which would also promote $^*\text{OCHO}$ generation

and inhibit H_2 production (Fig. S20†). These results indicate that Sr_2SnO_4 exhibits high selectivity due to the optimized binding strength of $^*\text{OCHO}$.

Furthermore, crystal orbital Hamilton population (COHP) analysis was conducted to explore how the SrO-layer insertion in Sr_2SnO_4 would affect the selectivity (Fig. 5c). The anti-bonding orbitals of Sn active sites with the O atom of $^*\text{OCHO}$ below the Fermi level (E_f) imply the weakening adsorption of $^*\text{OCHO}$ on Sn, which is most obvious on the $\text{Sr}_2\text{SnO}_4(002)$ surface. Therefore, the filling of the anti-bonding orbitals leads to a reduction in the bonding strength of Sn–O, corresponding to the relatively weak adsorption strength of $^*\text{OCHO}$ on Sr_2SnO_4 . Besides, the linear relationship between the integrated COHP (iCOHP) and $\Delta G_{^*\text{OCHO}}$ quantitatively confirms that the population of bonding/antibonding orbitals below E_f is the key factor in regulating the $\Delta G_{^*\text{OCHO}}$ (Fig. 5d).⁴¹ Therefore, the tuned electrical properties of active sites in the RP perovskite Sr_2SnO_4 electrocatalyst optimize the binding energy of $^*\text{OCHO}$ over $\text{Sr}_2\text{SnO}_4(002)/8\text{Ov}$, leading to the high selectivity for the CO_2RR to HCOO^- .

Conclusions

In summary, a Sr_2SnO_4 perovskite electrocatalyst with a RP structure was fabricated *via* a simple solid-state reaction. Sr_2SnO_4 exhibits high performance for the CO_2RR with an FE of 83.7% towards HCOO^- and long-term durability of over 24 h. The SrO-layer insertion in the RP structure leads to strong interlayer interactions, which could prohibit structural deformation and reconstruction. Thus, Sr_2SnO_4 possesses good stability. Besides, the anti-bonding orbital filling of the Sn active site in the RP structure results in the favorable binding strength of $^*\text{OCHO}$ on Sr_2SnO_4 , so that it shows higher selectivity towards HCOO^- than the SnO_2 and the SrSnO_3 control samples. Our work demonstrates that Sr_2SnO_4 can serve as an efficient electrocatalyst for HCOO^- production through the CO_2RR . The findings would guide the rational design of perovskite-based electrocatalysts and open up an avenue for tuning the catalytic selectivity and stability by structure engineering.

Data availability

The data that supports the findings of this study is available from the corresponding author upon reasonable request.

Author contributions

J. L. G supervised the project. P. Z. and J. Z. conceived the idea. J. Z. synthesized the catalysts and conducted the CO_2RR performance tests. L. L. L conducted DFT simulations. Y. T. H., G. H., and G. Z. conducted related characterization studies. All the authors participated in the writing of the manuscript.

Conflicts of interest

There are no conflicts to declare.

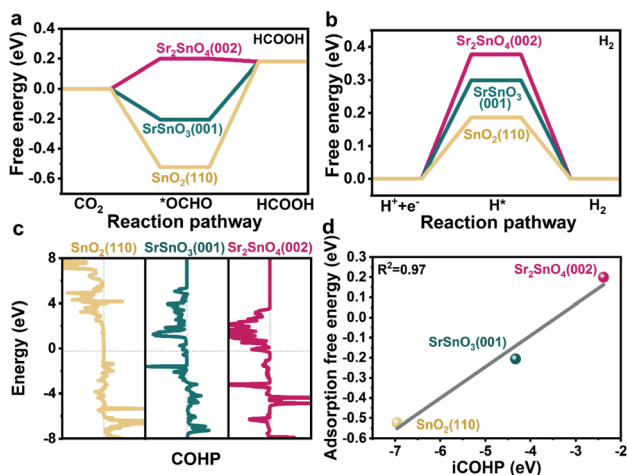


Fig. 5 The calculated free energy for (a) HCOOH and (b) H_2 . (c) COHP curves of Sn active sites on $\text{SnO}_2(110)$, $\text{SrSnO}_3(001)$, and $\text{Sr}_2\text{SnO}_4(002)$ with O of $^*\text{OCHO}$ interaction. (d) The correlation between the iCOHP and adsorption free energy of $^*\text{OCHO}$ on $\text{SnO}_2(110)$, $\text{SrSnO}_3(001)$, and $\text{Sr}_2\text{SnO}_4(002)$.



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