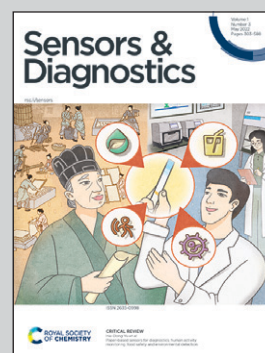


**Showcasing research from Professor Jobin Cyriac's laboratory, Department of Chemistry, Indian Institute of Space Science and Technology, Thiruvananthapuram, Kerala, India.**

Detection and screening of basic amino acids using the luminescence switching of WS<sub>2</sub> nanosheets-Ag<sub>2</sub>O nanoparticles composite

A 0D-2D composite material was prepared by combining WS<sub>2</sub> nanosheets and Ag<sub>2</sub>O nanoparticles. The fluorescence from WS<sub>2</sub> nanosheets was quenched in the presence of Ag<sub>2</sub>O nanoparticles. The spectral overlap of the nanosheets' emission with surface plasmon resonance absorption of Ag<sub>2</sub>O nanoparticles explains the nonradiative energy transfer. The quenched fluorescence was selectively recovered in the presence of three basic amino acids, His, Lys and Arg, at a pH below the isoelectric point of each amino acid.

**As featured in:**



See Jobin Cyriac *et al.*,  
*Sens. Diagn.*, 2022, 1, 485.

## PAPER

[View Article Online](#)  
[View Journal](#) | [View Issue](#)
Cite this: *Sens. Diagn.*, 2022, 1, 485

# Detection and screening of basic amino acids using the luminescence switching of a WS<sub>2</sub> nanosheet–Ag<sub>2</sub>O nanoparticle composite†

 Neema Pallikarathodi Mani,<sup>a</sup> Karthika S. Sunil,<sup>b</sup> Ann Mary Tomy,<sup>a</sup>  
 Bhasha Sathyan<sup>a</sup> and Jobin Cyriac  <sup>✉</sup>

The Förster resonance energy transfer (FRET) pair-based detection has gained considerable attention due to its promising sensitivity and selectivity. Herein, we report a turn-on sensor for detecting basic amino acids utilizing the fluorescence emission properties of WS<sub>2</sub> nanosheets (WS<sub>2</sub> NSs). The sensor functions based on a FRET dyad. The addition of AgNO<sub>3</sub> into an alkaline solution of WS<sub>2</sub> NSs leads to the *in situ* formation of silver oxide nanoparticles (Ag<sub>2</sub>O NPs). The negative surface charge and the presence of sulphur vacancies, together with the possibility for a multiplexed adsorption platform of WS<sub>2</sub> NSs, account for the formation Ag<sub>2</sub>O NPs on their surface. In the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP system (termed as a 0D–2D composite), the fluorescence emission of the nanosheets became subdued due to the FRET. This 0D–2D composite is proven to be an excellent turn-on sensor for basic amino acids. Histidine, lysine and arginine can induce the aggregation of the Ag<sub>2</sub>O NPs, which shuts the FRET pathway, along with the regeneration of the fluorescence of WS<sub>2</sub> NSs. The aggregation of Ag<sub>2</sub>O NPs occurs in a sensor solution with pH below the isoelectric points of the amino acids to help to discriminate them. The dynamic range and limit of detection of the sensor have been evaluated. The protein detection capability of the sensor was verified using a lysine containing protein, ubiquitin, and by real sample analysis, using biological fluids.

 Received 18th January 2022,  
 Accepted 9th March 2022

DOI: 10.1039/d2sd00009a

[rsc.li/sensors](https://rsc.li/sensors)

## 1. Introduction

One of the salient features of atomically thin, 2D layered materials is the interrelation of their dimensionality and fundamental properties.<sup>1,2</sup> The quantum confinement effect that emerged upon scaling down the dimensions is manifested as photoluminescence (PL) with an improved quantum yield, enabling these materials to be used as fluorescent-based chemical sensors.<sup>3,4</sup> The natural abundance of transition metal dichalcogenides (TMDs) such as MoS<sub>2</sub>, WS<sub>2</sub>, *etc.*, makes them attractive in the research realm of 2D layered materials. The increased demand in the market for miniaturization of electronic devices and their semiconducting nature with a sizeable bandgap catalyzes such studies. Layered TMDs of formula MX<sub>2</sub> typically have a thickness of 6–9 Å and consist of a single layer of hexagonally packed metal atoms sandwiched between two layers of chalcogen atoms.<sup>5,6</sup> The atoms within the layer are bonded by strong covalent bonds, whereas

different layers are held together by weak van der Waals interactions, which enable them to exfoliate into monolayers with the emergence of exotic properties.<sup>7</sup> In layered materials, the van der Waals interactions are not limited to interplanar interactions. Generally, any passivated, dangling bond-free surface interacts with another moiety *via* van der Waals forces. Consequently, any layered 2D material can be integrated with various materials of different dimensionalities to form mixed-dimensional structures. 0D–2D materials are such heterostructures produced by integrating 2D materials with 0D materials, primarily through noncovalent interactions. The 2D components may be graphene or transition metal dichalcogenides (TMDCs) and the 0D components may be small organic molecules or quantum dots (QDs). In contrast to conventional, epitaxially grown heterostructures, the interface in a mixed-dimensional structure is relatively complex and less constrained due to the absence of the need for lattice matching. The usual discontinuity in the band structure and resulting potential energy barriers, the density of states (DOS), also undergoes an abrupt transition, which has several observable implications, such as additional resistance at the junction that results from the change in the number of conductance channels.

Amino acids (AAs), being the building units of proteins, are associated with various physiological processes and are vital

<sup>a</sup> Department of Chemistry, Indian Institute of Space Science & Technology, Thiruvananthapuram – 695 547, India. E-mail: [jobincyriac@iist.ac.in](mailto:jobincyriac@iist.ac.in)

<sup>b</sup> Department of Chemistry, Indian Institute of Science Education and Research, Berhampur – 760010, India

† Electronic supplementary information (ESI) available. See DOI: 10.1039/d2sd00009a



components of life processes.<sup>8,9</sup> For example, one of the basic amino acids, lysine (Lys), is associated with weight gain in animals, polyamine synthesis and the Krebs–Henseleit cycle.<sup>10</sup> Histidine (His), another basic AA, is closely related to the growth and repair of tissues, the control of transmission of metal elements in biological bases, *etc.*<sup>11</sup> Arginine (Arg), the third and most basic AA, is known for its association with cell division, immune function, the healing of wounds, the release of hormones, *etc.*<sup>12</sup> The AA detection, in general, is thus significant in terms of valuable information it can render in the fields of nutritional analysis and diagnosis of diseases such as Alzheimer's, pancreatitis, *etc.* Among AAs, the detection of thiol-containing AAs is relatively easy owing to their unique nucleophilicity,<sup>13</sup> whereas discrimination of non-thiol-containing AAs is still challenging. Electrochemical and chromatographic approaches are widely used strategies for the detection and characterization of AAs.<sup>14–19</sup> Yet, the demand for sophisticated instruments, operational inconveniences, economic viabilities, choice of various parameters such as response time, detection limit, *etc.*, pave the way for the quest for more desirable and convenient methods. Optical chemosensors are marked with their high sensitivity with notable selectivity along with the possibility of visual discrimination.<sup>8</sup> Hence, a potential fluorometric probe for recognizing specific AAs susceptible to high-throughput assays holds a breakthrough in the current scenario.

Several nanomaterials are demonstrated as signal transducers in colourimetric/fluorescence chemical sensors. The ease of surveilling the light signal has made them a potential candidate in surface plasmon resonance (SPR)-based calorimetric sensing. Many reports that exploit the changes in the SPR peak of noble metal nanoparticles upon interaction with AAs are available. For example, *p*-sulfonatocalix[4]arene thiol modified Au nanoparticles have been used as a colourimetric sensor for His, Lys and Arg in water, based on the principle of broadening and shifting of SPR bands to the red region upon interaction, with linear detection limits of 1 mM, 2 mM and 4 mM, respectively.<sup>20</sup> Chiral detection of L-His was reported based on L-arginine-sulfonated-substituted zinc tetraphenylporphyrin modified Ag nanoparticles (L-Arg–ZnTPPS–Ag NPs).<sup>21</sup> Many groups reported fluorescence-based detection of AAs with improved sensitivity. Lu *et al.* utilized the fluorescence emission of a water-soluble fluorescein-bis-acrylate carrying pyridinium moiety having high sensitivity towards biothiols, which can discriminate simultaneously basic amino acids and thiol-containing amino acids from other amino acids.<sup>13</sup> A simple fluorescent turn-on sensor, employing two complementary cucurbit[*n*]uril microarrays, has been demonstrated for the sensitive detection of basic AAs in water.<sup>22</sup> A close observation of the literature reveals that the detection strategy depends on either the agglomeration of nanoparticles followed by a change in the SPR band or the change in the fluorescence emission of the probe. A combined approach with the ease of agglomeration of nanoparticles and the sensitivity of the fluorescence turn-on methodology would be attractive in chemical sensor applications.

Here, we report a fluorescence sensor for basic amino acids based on WS<sub>2</sub> nanosheet–silver oxide nanoparticle (WS<sub>2</sub> NS–Ag<sub>2</sub>O NP) composite materials. The composite was prepared *via* the *in situ* formation of silver oxide nanoparticles (Ag<sub>2</sub>O NPs) by the combined action of fluorescent WS<sub>2</sub> nanosheets (WS<sub>2</sub> NSs) in the presence of NaOH. The luminescent WS<sub>2</sub> NSs were synthesized using a facile hydrothermal method. The fluorescence of WS<sub>2</sub> NSs was quenched upon the formation of Ag<sub>2</sub>O NPs due to resonance energy transfer. The obvious spectral overlap of WS<sub>2</sub> NS emission with that of surface plasmon resonance (SPR) absorption of Ag<sub>2</sub>O NPs explains the nonradiative energy transfer. The quenched fluorescence was selectively recovered in the presence of three basic amino acids: His, Lys and Arg, at a pH below the isoelectric point of each amino acid. The fluorescence recovery was attributed to the aggregation of Ag<sub>2</sub>O NPs and thereby the lack of spectral overlap between WS<sub>2</sub> NSs and Ag<sub>2</sub>O NPs. The sensor has an appreciable dynamic range and limit of detection (LOD) for all three AA analytes. The formation of Ag<sub>2</sub>O NPs, the comprehensive mechanism of quenching and the recovery of fluorescence of WS<sub>2</sub> NSs were proposed with experimental support. The real sample analysis was also demonstrated using biological fluids.

## 2. Experimental

### 2.1 Synthesis of WS<sub>2</sub> nanosheets

The WS<sub>2</sub> nanosheets were prepared by hydrothermal reaction of commercially available WS<sub>2</sub> powder and NaOH in an aqueous medium by taking a molar ratio of 1:1. Briefly, WS<sub>2</sub> powder (100 mg) was mixed with NaOH (16.0 mg) in 10 mL of water. The resultant solution was taken in a Teflon-lined autoclave and kept for 24 h at 220 °C. The solution after the reaction was centrifuged at 3000 rpm for 30 min; the supernatant was collected, filtered and dialyzed for 3 h using a dialysis bag (molecular weight cut off 500). These purified, light-yellow solutions were used for all further studies.

### 2.2 Synthesis of the WS<sub>2</sub> nanosheet–Ag<sub>2</sub>O NP nanocomposite

To a 2.5 mL of purified WS<sub>2</sub> nanosheet solution, aliquots of AgNO<sub>3</sub> solution (1, 2, 3 or 4 mL of 0.02 M) were added dropwise by stirring. The colour of the solution turned yellow *in situ* due to the formation of Ag<sub>2</sub>O NPs on the surface of the WS<sub>2</sub> nanosheets.

## 3. Results and discussion

### 3.1 Synthesis and characterization of WS<sub>2</sub> nanosheets

In parallel with the steadily growing interest in 2D layered materials, different synthetic strategies have been reported for these materials. We propose a simple approach to prepare 2D layered materials in a nanosize regime employing a hydrothermal reaction. The vigorous reaction of WS<sub>2</sub> powder with a 1:1 molar ratio of NaOH at 220 °C yielded exfoliated WS<sub>2</sub> nanosheets. The optimized reaction time for the best



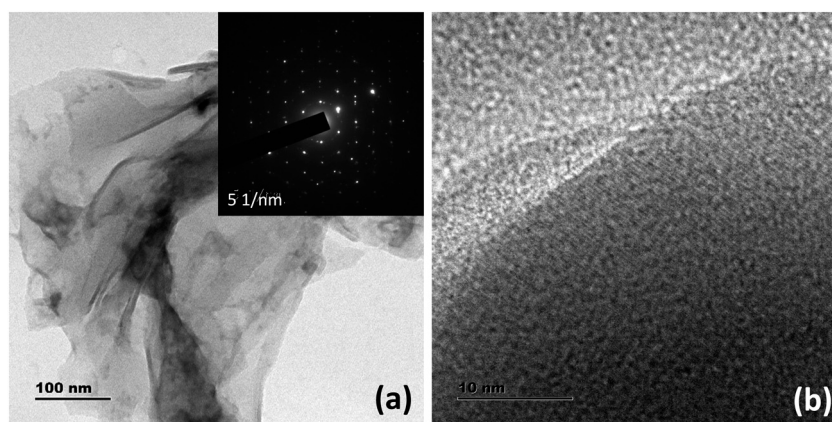


photoluminescence (PL) properties was 24 h. Detailed microscopy and photophysical characterization techniques have been carried out to understand the structural and optical attributes of the material. The transmission electron microscopy (TEM) images (see Fig. 1) show the formation of NSs with a lateral dimension in the range of a few hundreds of nanometers. Notably, the crystallinity of the sheets is maintained even after the vigorous hydrothermal reaction, as evident in the selected area electron diffraction (SAED) pattern. The AFM micrographs confirm the formation of nanosheets with a few layer thicknesses (Fig. S1, ESI†). The thickness measurement from the height profiles drawn at different regions of the AFM micrographs confirms the layer thickness of 2.8 nm, indicating the presence of 2–3 layers of sheets.

Raman spectroscopy can effectively determine the number of layers in the exfoliated TMDs. Generally, there are four major Raman-active modes in bulk WS<sub>2</sub>, namely A<sub>1g</sub>, E<sub>1g</sub>, E<sub>2g</sub><sup>1</sup>, and E<sub>2g</sub><sup>2</sup> modes. Due to the forbidden selection rule in the back-scattering geometry and the limited rejection against Rayleigh scattering, it is challenging to detect the E<sub>1g</sub> and E<sub>2g</sub><sup>2</sup> modes, respectively.<sup>23–26</sup> However, the in-plane vibrational E<sub>1g</sub><sup>1</sup> mode and the out-of-plane vibrational A<sub>1g</sub> mode are more significant as they can be used to identify the number of layers present in the sheets. A thorough study regarding the frequency trends of the A<sub>1g</sub> and E<sub>2g</sub> modes with varying layer thicknesses was performed by Molina-Sánchez and Wirtz.<sup>27</sup> As the number of layers increases from monolayer to bulk, the A<sub>1g</sub> mode tends to blue-shift, whereas the E<sub>1g</sub><sup>1</sup> mode shows a red-shift. The weak interlayer interactions and strong dielectric screening of the long-range Coulomb interactions are reasons for this behavior. Therefore, a direct correlation between the frequency difference ( $\Delta\omega$ ) of the E<sub>1g</sub><sup>1</sup> mode and A<sub>1g</sub> mode and the sample thickness can be established.<sup>28</sup> In the present case, for WS<sub>2</sub> nanosheets, the value of  $\Delta\omega$  is <65 cm<sup>-1</sup>, which corresponds to monolayer thickness, invariably denoting the exfoliation of the bulk material into nanosheets of a few layers (Fig. 2a).

The crystalline structure and phase of WS<sub>2</sub> NSs were investigated by X-ray diffraction (XRD), using WS<sub>2</sub> powder as a

reference (Fig. S2, ESI†). The diffraction peaks of the powder sample were indexed using JCPDS Card No. 08-0237, corresponding to the pure hexagonal (P63/mmc space group) WS<sub>2</sub> phase. The (002) peak indicates the neat stacking of layers in the powder sample, along with the peaks corresponding to the (004), (100), (101), (102), (103), (006), (105), (106), (110), (008), (112), (107), (114) and (116) planes.<sup>29,30</sup> It can be observed that all the diffraction peaks of WS<sub>2</sub> NSs can be indexed to the hexagonal 2H phase of bulk WS<sub>2</sub>, demonstrating the well resolved crystalline structure of WS<sub>2</sub> NSs exfoliated under the hydrothermal conditions. The absence of any additional peak indicates neither any noticeable impurity nor any phase change to the 1T phase during the synthesis. Besides, the weak signal from the (002) plane, compared to the WS<sub>2</sub> powder, denotes the few-layer structure of WS<sub>2</sub> NSs.<sup>31</sup> The chemical state and elemental composition of WS<sub>2</sub> NSs were examined by X-ray photoelectron spectroscopy (XPS) analysis (Fig. S3, ESI†). The high-resolution core-level XPS peaks corresponding to W<sup>4+</sup> 4f<sub>7/2</sub> and W<sup>4+</sup> 4f<sub>5/2</sub> were observed at binding energy values of 33.3 and 35.5 eV, respectively, and a small peak at 39.3 eV for W 5p<sub>3/2</sub> indicates the dominance of the (IV) oxidation state for W in WS<sub>2</sub> NSs.<sup>31–33</sup> The absence of peaks corresponding to the 1T phase (in the range of 31.6–32.2 eV)<sup>34</sup> indicates no phase transformation during the hydrothermal synthesis, corroborating the observation in the XRD spectrum. The peaks at 162.0 and 163.2 eV correspond to the S 2p<sub>3/2</sub> and 2p<sub>1/2</sub> orbitals of divalent S, bonded to W<sup>4+</sup>. The peaks corresponding to the oxidized form of S are present at binding energy values of 168.0 and 169.2 eV, corresponding to S 2p<sub>3/2</sub> and 2p<sub>1/2</sub>, respectively. We surmise that, during the hydrothermal reaction, the edge S may be oxidized, leading to the formation of S and O containing functionalities. These functional groups describe the substantial aqueous solubility and phenomenal temporal stability. A single peak in the high-resolution XPS spectrum of oxygen at 531.6 eV indicates the S–O bonds (Fig. S3c, ESI†). The absence of a W<sup>6+</sup> oxidation state peak and multiple peaks for oxygen in the XPS spectrum depicts that W did not undergo oxidation during the synthesis.



**Fig. 1** (a) TEM micrographs of WS<sub>2</sub> nanosheets. The inset of (a) shows the SAED pattern presenting the crystalline nature of the WS<sub>2</sub> nanosheets. (b) Magnified TEM image of the WS<sub>2</sub> NSs depicting the Moiré pattern.



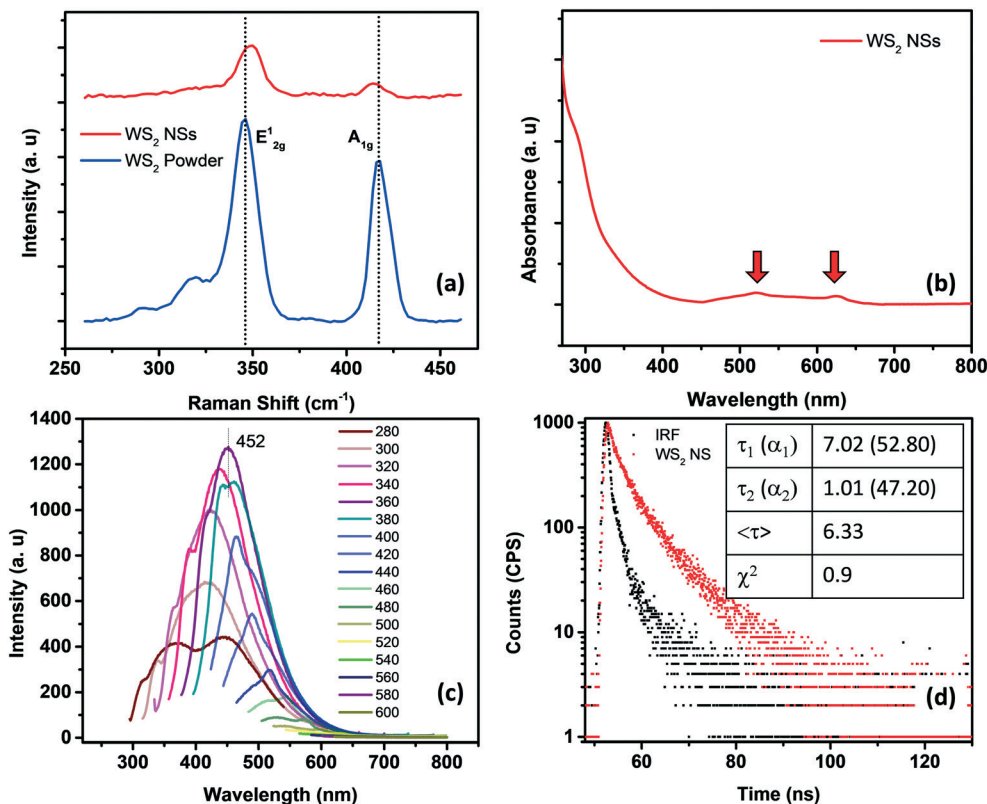


Fig. 2 (a) Raman spectra of WS<sub>2</sub> powder (blue) and WS<sub>2</sub> nanosheets (red). (b) The UV-visible absorption spectrum of WS<sub>2</sub> nanosheets, showing the signature peaks of the sheets corresponding to A and B excitons. (c) Excitation dependent emission spectra of WS<sub>2</sub> NSs. (d) Lifetime spectrum of WS<sub>2</sub> NSs fitted with a bi-exponential function.

The optical signatures of WS<sub>2</sub> NSs have been evaluated using UV-visible and fluorescence spectroscopy. In general, WS<sub>2</sub> powder possesses four major absorption bands at around 630, 528, 456 and 417 nm, which are attributed to A, B, C and D excitons, respectively.<sup>35</sup> The transitions from the spin-splitting valence band to the conduction band at the *K* point of the Brillouin zone are responsible for the emergence of peaks A and B. In contrast, the optical transitions from the deep valence band to the conduction band are denoted as C and D peaks.<sup>36</sup> As the lateral size of the sheets decreases due to enhanced quantum confinement effects and edge effects, these excitonic positions can be blue-shifted.<sup>37</sup> It is known that a continuum of absorption in the excitonic position confirms the phase change of a semiconducting 2H form to a metallic 1T form.<sup>38</sup> The A and B excitons present at 625 and 520 nm, respectively, for WS<sub>2</sub> NSs confirm the retention of the semiconducting nature of the bulk powder after the hydrothermal reaction in the presence of alkali (Fig. 2b). The quantum confinement effects due to the reduction in lateral dimensions of the sheets to the nano regime are manifested as high PL (see the fluorescence spectra given in Fig. 2c). A typical excitation dependent emission was observed for WS<sub>2</sub> NSs, with the best emission centered at 452 nm at an excitation wavelength of 360 nm. The size heterogeneity of the sheets ranging from a few nanometers to micrometers can be one of the reasons behind this behavior. The quantum

yield (QY) of the sample calculated using quinine sulphate as a standard has a value of 2.2%, which is comparable with the reported QY value of WS<sub>2</sub> nanosheets.<sup>37,39</sup> Also, the material shows exceptional temporal stability as well as photostability. This remarkable PL from the nanosheets and its inherent stability open the opportunity to exploit these materials in the field of fluorescence-based chemical sensors. Time-resolved fluorescence spectroscopic analysis of WS<sub>2</sub> NSs was performed to unveil the excited state phenomenon occurring within the fluorophore material. The WS<sub>2</sub> NSs are reported to have a very short lifetime of 800 ps.<sup>40</sup> By shrinking the thickness of the bulk materials, the photo-excited carrier lifetime is generally increased due to the enhanced stability of the excitons. In the present case, the decay curves for WS<sub>2</sub> NSs were collected and a bi-exponential fitting was obtained with carrier lifetimes of 2.75 and 9.07 ns by keeping the numerical fitting parameter  $\chi^2$  near 1 (Fig. 2d).

### 3.2 Fabrication of the WS<sub>2</sub> NS-Ag<sub>2</sub>O nanoparticle composite

One of the remarkable properties of the present WS<sub>2</sub> NSs is their stable PL for several months under laboratory conditions. It was observed that the PL of the nanosheets was quenched instantaneously by the addition of dil. AgNO<sub>3</sub>. Upon incremental addition of AgNO<sub>3</sub> solution (100  $\mu$ L each of 0.02 M) to a solution of 2.5 mL of WS<sub>2</sub> nanosheets, at



room temperature, the PL emission is found to plummet accordingly with a change in color of the solution from very light yellow to deep yellow (Fig. 3). The color change, as well as the change in PL, was rapid. This color change was also reflected in the UV-vis spectrum. It has been observed that along with the addition of  $\text{AgNO}_3$ , a new absorption peak below 400 nm emerged (Fig. 3b). The appearance of the yellow color and the new peak at  $\sim 400$  nm prompted us to surmise that the *in situ* formation of nanoparticles, having surface plasmon resonance (SPR) properties, might have taken place at the surface of  $\text{WS}_2$  NSs.

The formation of silver oxide nanoparticles ( $\text{Ag}_2\text{O}$  NPs) from  $\text{AgNO}_3$  in the presence of  $\text{NaOH}$  at elevated temperatures is well documented.<sup>41,42</sup> Our  $\text{WS}_2$  NS solution contains  $\text{NaOH}$ , which was left over from the hydrothermal reactions that anticipated the formation of silver oxide nanoparticles. X-ray diffraction studies unveil the nature of the nanoparticles, as the peak positions for silver nanoparticles (which usually form a face centred cubic lattice) will be different from silver oxide nanoparticles (which form cubic crystals upon reaction between  $\text{AgNO}_3$  and  $\text{NaOH}$ ). The XRD pattern of the  $\text{WS}_2$  NS-NP nanocomposite shows remarkable peaks at  $2\theta$  values of  $27.57^\circ$ ,  $31.86^\circ$ ,  $38.18^\circ$ ,  $46.98^\circ$ ,  $54.88^\circ$ ,  $64.47^\circ$  and  $69.46^\circ$ , apart from the peaks corresponding to  $\text{WS}_2$  NSs (Fig. S4, ESI†). These peaks were indexed to the (110), (111), (200), (211), (220), (311) and (222) crystal planes of the  $\text{Ag}_2\text{O}$  NPs, having a cubic phase of silver oxide [JCPDS no. 41-1104].<sup>43,44</sup> Therefore, we surmise the formation of  $\text{Ag}_2\text{O}$  NPs. We further carried out XPS analysis to confirm the result obtained from the XRD analysis. The high-resolution XPS spectrum corresponding to Ag 3d shows peaks at 367.5 and 373.5 eV, which can be assigned to  $3d_{5/2}$  and  $3d_{3/2}$ , respectively, of  $\text{Ag}^+$  ions in the crystal lattice of  $\text{Ag}_2\text{O}$  (Fig. S5a, ESI†).<sup>43,44</sup> The formation of Ag–O was further confirmed from the XPS spectrum of oxygen 1s. We can observe the emergence of a second peak in the lower energy range of the

deconvoluted spectrum of oxygen, corresponding to the Ag–O bond apart from the peak corresponding to the O–S bond (Fig. S5b, ESI†).

Further, TEM analysis confirms the formation of spherical nanoparticles on the surface of  $\text{WS}_2$  NSs, forming a typical 0D–2D nanocomposite composed of  $\text{Ag}_2\text{O}$  NPs on the surface of  $\text{WS}_2$  NSs (Fig. 4). Optimization of the concentration of  $\text{WS}_2$  to  $\text{AgNO}_3$  was carried out to obtain the homogeneous size distribution of  $\text{Ag}_2\text{O}$  NPs. From TEM, it was evident that, for a volume ratio of 2.5 mL of  $\text{WS}_2$  solution, a uniform size distribution of NPs could be seen for 3 mL of 0.02 M  $\text{AgNO}_3$ , and this nanocomposite material is used for further studies. The size distribution of these NPs at various volume ratios along with average diameters is given in Fig. S6†

One of the significant observations is the quenching of the PL of  $\text{WS}_2$  nanosheets upon the formation of the  $\text{WS}_2$ – $\text{Ag}_2\text{O}$  NP composite. The general mechanism of quenching involves electron transfer and energy transfer. The spectral overlap of donor emission and acceptor absorption is one of the prerequisites for a celebrated excited-state quenching mechanism, Förster resonance energy transfer (FRET), to materialize. As shown in Fig. 5, there is a remarkable spectral overlap of the absorption of the nanocomposite with the emission of  $\text{WS}_2$  NSs. Since the  $\text{Ag}_2\text{O}$  NPs are dispersed on the surface of the nanosheets, the proximity of donor and acceptor species, the second criterion to realize FRET, is satisfied. Therefore, we ascribe the PL quenching to the FRET mechanism because of the presence of  $\text{Ag}_2\text{O}$  NPs with SPR absorption at  $\sim 400$  nm.

### 3.3 The nanocomposite as a turn-on sensor for basic amino acids

The quenched fluorescence of  $\text{WS}_2$  NSs could be recovered by annihilating the nanosheet–nanoparticle dyad, which opens the possibility of creating a turn-on sensor. We have seen that the quenching occurs in the presence of  $\text{Ag}_2\text{O}$  NPs. Disruption of these  $\text{Ag}_2\text{O}$  NPs can shut the energy transfer

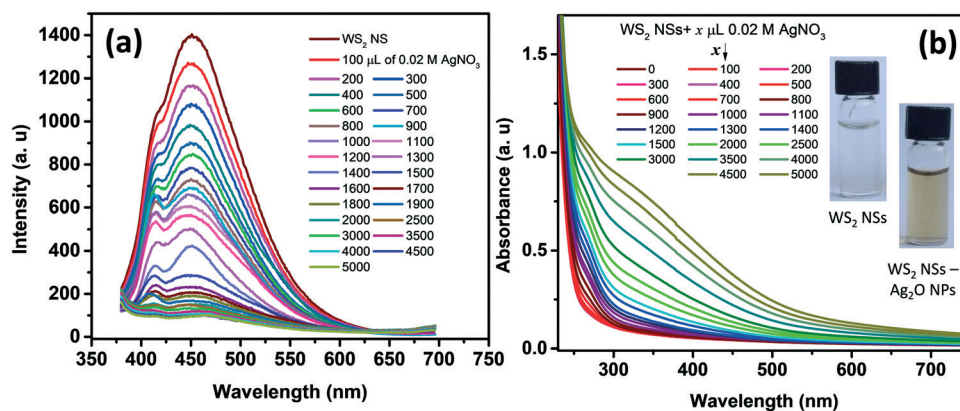


Fig. 3 (a) The PL emission spectra of  $\text{WS}_2$  NSs after adding various concentrations of the  $\text{AgNO}_3$  solution. (b) Absorption spectra of  $\text{WS}_2$  nanosheets after adding various concentrations of the  $\text{AgNO}_3$  solution. The inset of (b) shows the photographs of the  $\text{WS}_2$  NS solution and  $\text{WS}_2$  NS– $\text{Ag}_2\text{O}$  NP nanocomposite under visible light.





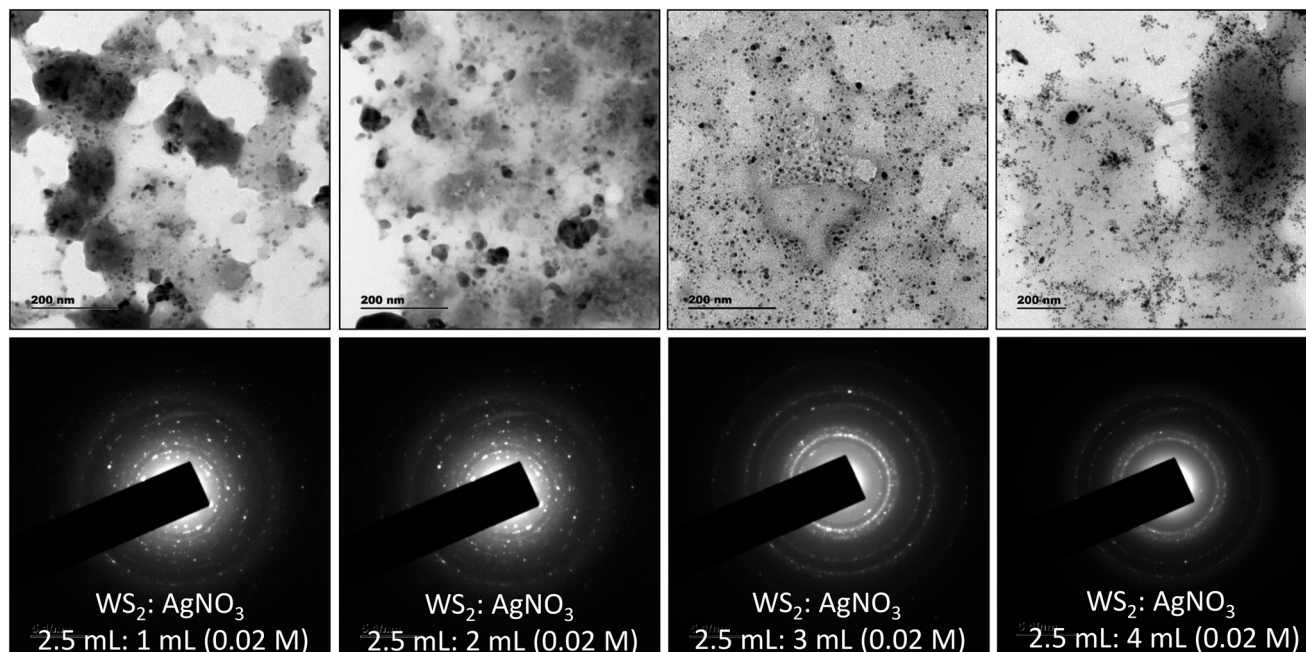


Fig. 4 TEM images and the SAED patterns of  $\text{WS}_2$  NSs solution after adding different volumes of  $\text{AgNO}_3$  solution. The upper pane shows the TEM images and the lower pane shows the SAED patterns. The volume ratio is written in each image.

pathways, and hence, the PL of  $\text{WS}_2$  NSs can be recovered. The stability of any colloid is due to the presence of charges on its surface, and in the present case, it can be inferred by zeta potential measurements (Fig. S7, ESI†). Any moiety with a counter charge can destroy the colloidal system and lead to nanoparticle aggregation. Such aggregation of  $\text{Ag}_2\text{O}$  NPs leads to the disappearance of SPR peaks, which is responsible for FRET. The biocompatibility and notable aqueous solubility of  $\text{Ag}_2\text{O}$  NPs and  $\text{WS}_2$  NSs prompted us to choose any biologically relevant molecules as analytes. Interestingly, the quenched PL was found to recover with the addition of basic amino acids at a pH value lower than their isoelectric points. We now formulated the  $\text{WS}_2$  NS- $\text{Ag}_2\text{O}$  NP

composite material as a turn-on sensor to detect the three basic amino acids His, Lys and Arg.

### 3.4 Sensor action

Amino acids, though zwitterionic, tend to be positively charged at a pH below their isoelectric points ( $p^I$ ).<sup>45</sup> The charges on the surface of AAs can be varied accordingly by changing the pH. The  $p^I$  values of His, Lys and Arg are 7.59, 9.74 and 10.76, respectively. Hence at a pH of 7, all three AAs will be positively charged, and at pH 9, Lys and Arg will be positively charged, whereas, at pH 10, only Arg will be positively charged. The PL response of the composite towards the three AAs at pH 7 is given in Fig. 6 and S8, ESI†.

It was noted that the PL responses towards these three AAs (His, Lys and Arg) were distinct from each other. We conducted the experiments at three different pH values: at pH 7, 9 and 10. As evident from Fig. 7, at pH 7, all three AAs could turn-on the PL, whereas the recovery effect is negligible for other AAs and potential interfering biomolecules. Thus, the composite is selective towards the detection of all the basic AAs at pH 7. The selectivity is further evaluated by the selectivity coefficient ( $K_{\text{arg}} = S_{\text{arg}}/S_0$ ; where  $S_{\text{arg}}$  is the response of the sensor towards Arg and  $S_0$  is the response of the sensor to other molecules/ions)<sup>46,47</sup> using Arg as the standard at different pH values, and is given in Table S1†. The AA structure with a net positive charge below its isoelectric point provides us room for distinction between the basic AAs (Fig. S9, ESI†). For example, by changing pH 7 to pH 9 or 10, the PL response of the nanocomposite has changed. At pH 9 and 10, His cannot turn on the PL, whereas both Lys and Arg

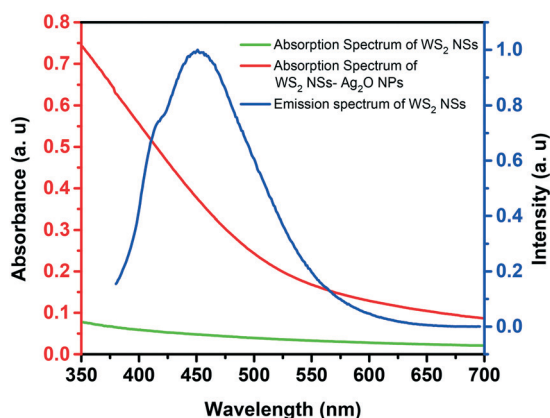


Fig. 5 Spectral overlap of the absorption spectrum of the  $\text{WS}_2$  NS- $\text{Ag}_2\text{O}$  NP composite (red trace) (by the addition of 3.5 mL  $\text{AgNO}_3$ ) and the emission spectrum of  $\text{WS}_2$  nanosheets (blue trace). The absorption spectrum of  $\text{WS}_2$  NSs alone (green) is given for comparison.



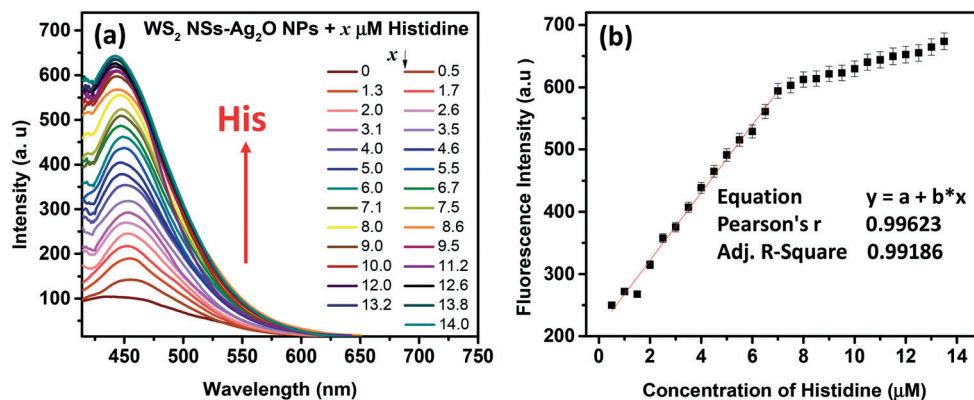


Fig. 6 (a) Response of His towards fluorescence emission of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite at pH 7. The concentration of His in micromolar is given in the figure legend. The excitation wavelength is 360 nm. (b) The calibration curve shows the dynamic range of His. The linear region is considered as the dynamic range of the sensor towards His detection.

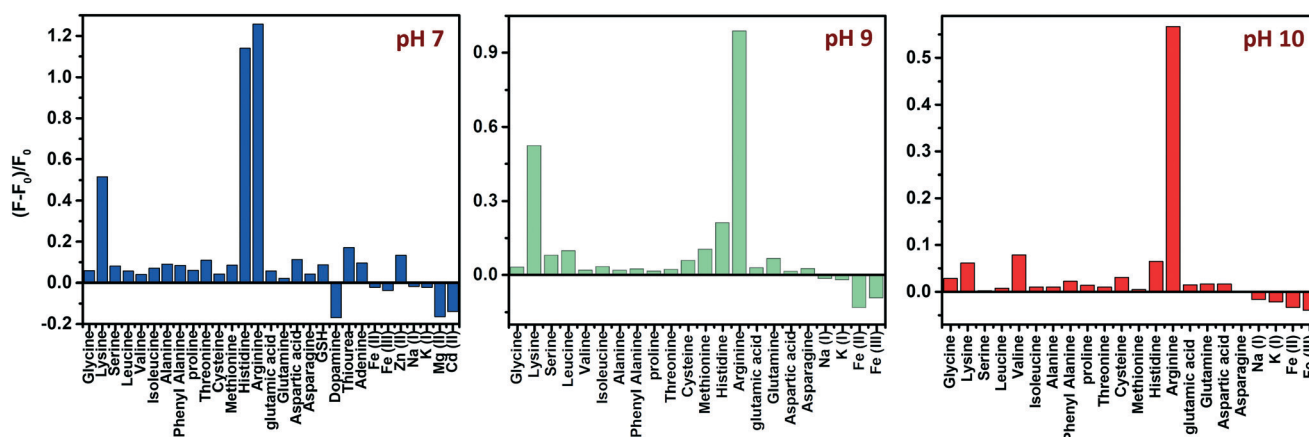


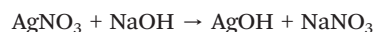
Fig. 7 pH-dependent selectivity: effect of different AAs, other biomolecules and metal ions on the fluorescence intensity of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite at three different pH values.

could still turn-on the emission of the nanocomposite. And at pH 10, Lys also fails to recover the PL of the composite; only Arg can turn-on the emission. Thus, screening within the basic AAs can be accomplished by the 0D-2D nanocomposite sensor solution, merely by altering the pH of the sensor solution. The fluorescence response of the sensor solution towards Lys and Arg at pH 9 and 10, respectively, is given in Fig. S10, ESI†. The dynamic range of His (at pH 7), Lys (at pH 7 and 9) and Arg (at pH 7, 9 and 10) is summarized in Table S2†. The photographs of the WS<sub>2</sub> NSs, WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite, and nanocomposite solution containing His, Lys and Arg under visible and UV light show the possibility of a visual screening (Fig. S11, ESI†). The zeta potential values of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite at various pH values are given in Table S3† exhibiting its stability at various pH values.

### 3.5 Mechanism of formation of Ag<sub>2</sub>O NPs and fluorescence turn-on sensing

AgNO<sub>3</sub> in the presence of NaOH is known to form Ag<sub>2</sub>O according to the equation given below.<sup>41</sup> Thus we propose

that in the presence of NaOH, which was used for the hydrothermal reaction, Ag<sub>2</sub>O particles might have formed.



The survey scan XPS spectra of WS<sub>2</sub> NSs and the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP composite show peaks corresponding to W, S, and O, along with the peaks of Ag in the case of the composite (Fig. 8a and b). The deconvoluted spectra of W of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP composite show peaks corresponding to W 4f<sub>7/2</sub>, 4f<sub>5/2</sub> and W 5p<sub>3/2</sub> at slightly lower binding energy values (33.23, 35.33 and 39.13 eV, respectively) comparable to WS<sub>2</sub> NSs, due to the weaker W-S bond (Fig. 8c). These had binding energy values of 33.3, 35.5 and 39.3 eV for the case of WS<sub>2</sub> NSs (see Fig. S3, ESI† and the discussion in section 3.1). The high-resolution XPS peak corresponding to S indicates the chemical reaction possibilities at the nanosheet surface. In both cases, *viz.* NSs alone and the 0D-2D composite, the XPS peak split to W-S and S-O bonds (Fig. 8d and S3†). The peaks assigned to the metal-S bond (162.0 and



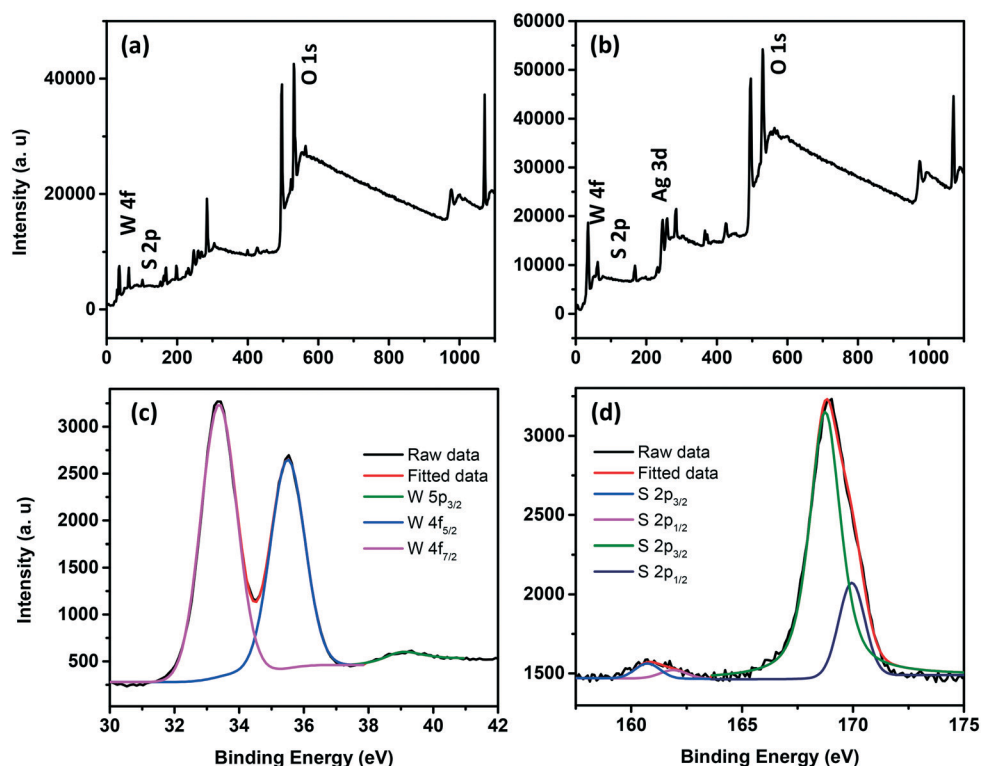


Fig. 8 Survey scan XPS spectra of  $\text{WS}_2$  NSs (a) and the  $\text{WS}_2$  NS- $\text{Ag}_2\text{O}$  NP nanocomposite (b). High resolution XPS spectra of the W (c) and S (d) regions of the  $\text{WS}_2$  NS- $\text{Ag}_2\text{O}$  NP nanocomposite.

163.2 eV) for the case of  $\text{WS}_2$  NSs have shifted to 160.7 eV and 161.9 eV in the composite material. This peak shift is attributed to the contribution from the Ag-S bond and subsequent peak broadening.

The crucial part of the turn-on sensing mechanism, on the other hand, was mainly determined by the protonated form of the basic AAs below their  $p^I$ . In the presence of AAs, the negative charges on the surface of  $\text{Ag}_2\text{O}$  NPs are getting nullified, leading to the aggregation of  $\text{Ag}_2\text{O}$  NPs and hence the reduction of the SPR peak intensity. This decrease in SPR absorption leads to the shutting off of the FRET pathway and the PL is subsequently recovered. The TEM micrographs, UV-

visible absorption studies and lifetime analysis corroborate with these assumptions. As shown in the TEM images given in Fig. 9 and S12, ESI<sup>†</sup>, His, Lys and Arg induce the aggregation of  $\text{Ag}_2\text{O}$  NPs at pH 7. At pH 9, the intact  $\text{Ag}_2\text{O}$  NPs are visible in the nanocomposite samples after adding His. At pH 10, addition of His or Lys does not lead to aggregation. This observation further authenticates the involvement of aggregation for the PL recovery (Fig. S13 and S14, ESI<sup>†</sup>). The absence of SPR peaks in the UV-visible absorption spectra (Fig. S15, ESI<sup>†</sup>) validates the argument.

Another vital observation of the interaction of AAs with the nanocomposite is the precipitation of  $\text{Ag}_2\text{O}$  NPs upon the

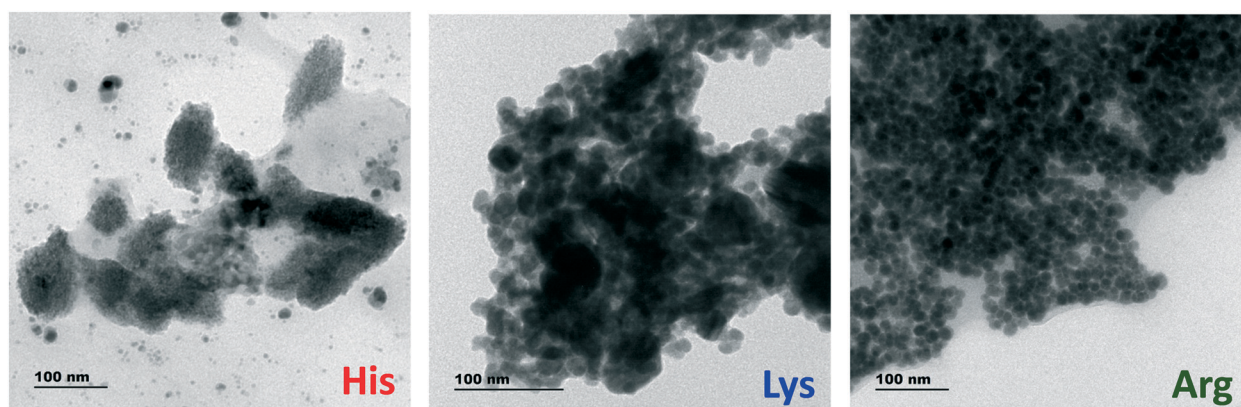


Fig. 9 TEM images of the  $\text{WS}_2$  NS- $\text{Ag}_2\text{O}$  NP nanocomposite upon interaction with His, Lys and Arg at pH 7.



addition of excess AAs (concentration well above the dynamic range). The yellow color of the solution disappeared with the precipitation of Ag<sub>2</sub>O as a brown precipitate. The TEM analysis of this colorless supernatant solution shows that Ag<sub>2</sub>O NPs have been completely precipitated along with some WS<sub>2</sub> NSs and the solution contains only the remaining WS<sub>2</sub> NSs (Fig. S16, ESI†). The crystallinity and the neat morphology of the sheets were lost after the precipitation of Ag<sub>2</sub>O NPs. We surmise that the locations with inscribed markings (in Fig. S16, ESI†) are the spots where the Ag<sub>2</sub>O NPs adsorbed onto the nanosheets. The Ag<sub>2</sub>O NPs are formed onto the surface of WS<sub>2</sub> NSs and may be stabilized by the groups containing S and O. Also, the dangling S atoms of the nanosheets take part in the reaction, eventually forming compounds like Ag<sub>2</sub>S as well. When the NPs detach from the sheets by aggregation followed by precipitation, the remaining sheets contain several etched areas due to the removal of adsorbed Ag<sub>2</sub>O NPs along with dangling S groups (as Ag<sub>2</sub>S). The etched regions are primarily defects due to S or its functional group vacancies. High-resolution XPS of the supernatant solution shows weak S peaks, which are proof of the removal of S from the sensor material by precipitation of nanoparticles. The oxygen region of the XPS spectrum, nonetheless, shows a single peak at 531.5 eV, as in the case of WS<sub>2</sub> NSs, which invariably shows the S–O bond. The high-resolution XPS analysis of W indicates no significant peak shift from the nanocomposite values (33.22, 35.32, and 39.02 eV), implying that the oxidation state of W was not affected (Fig. S17, ESI†). The absence of peaks corresponding to W–O and the W<sup>6+</sup> state indicate that the dangling bonds created by the removal of S was not satisfied by oxygen. Notably, a full recovery of PL was not attained even after the addition of AAs well above the dynamic range. This could be attributed to the defects created during the removal of Ag<sub>2</sub>O NPs by the precipitation, which can act as seats for nonradiative transfers (Fig. S18, ESI†).

The excited-state lifetime analysis depicts the complexity of the interactions between the AAs and the nanocomposite (Fig. S19, ESI†). It has been observed that, upon the addition of AgNO<sub>3</sub> into WS<sub>2</sub> NSs, one of the lifetime components becomes nonradiative with the highest contribution (0.13 ns and 80.38% contribution). We attribute this shift to the energy transfer from WS<sub>2</sub> NSs to Ag<sub>2</sub>O NPs formed on their surface. The absence of FRET upon the addition of AAs is substantiated by the recovery of their lifetime values near that of bare/free WS<sub>2</sub> NSs. The emergence of a third lifetime component with a nonradiative pathway for the solutions with His, Lys and Arg might have arisen from the defect created in the sheets during the Ag<sub>2</sub>O NP formation.

We were interested in knowing whether the aggregation of Ag<sub>2</sub>O nanoparticles (in the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP nanocomposite) by basic AAs can be modulated by changing the pH and the fluorescence can be quenched reversibly or not. Interestingly, we observed that the increase in the pH of the sensor solution from 7 to 9 for the analyte His led to the fluorescence quenching. This observation is attributed to the separation of

aggregated particles due to increased pH (above the isoelectric value of His). We examined the resultant solution using the UV-visible absorption spectra and TEM (see Fig. S20†). After the pH increase, the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP composite shows the presence of a SPR peak at 360 nm, depicting separated Ag<sub>2</sub>O NPs. The re-appearance of colour was also observed (see the photographs given in Fig. S20†). The same experiment was done for the analyte Lys at pH 7 and 9 and then the pH was increased to 10. Similar results were obtained. These experiments show that the sensor action is reversible and can be modulated by changing the pH.

### 3.6 Protein sensing and real sample detection

Protein-specific detection is of fundamental importance for diagnosing and detecting protein biomarkers and deciphering biological processes at cellular levels. Most fluorescent probes work under the principle of changes in their photophysical characteristics upon interaction with the protein through chemical reactions or mere physical interactions. Yet, the fluorescence labelling and subsequent detection are invariably laborious, and the development of a mix-and-detect strategy that employs turn-on-based fluorescence detection will be highly appealing.

To test our fluorescent probe for protein detection, ubiquitin, a regulatory protein found in eukaryotes tissues, was selected as a target protein, as it contains seven lysine residues in its structure. Two significant functions of ubiquitin are marking proteins to be eliminated from the system and modulating the functions of the substrate protein with which it can conjugate without metabolically destabilizing this acceptor protein.<sup>48</sup> The abnormal concentration of ubiquitin is associated with many physiological aberrations. For example, high ubiquitin levels can be observed in the cerebrospinal fluid of Alzheimer's disease patients.<sup>49</sup> Therefore, facile detection and precise quantification of ubiquitin are of utmost importance.

We demonstrated a label-free detection of ubiquitin using the present fluorescent probe, the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP composite. The PL intensity is found to be enhanced (6.5 fold) with the ubiquitin concentration from 0 to 61 µg mL<sup>-1</sup> linearly ( $R^2 = 0.99$ ) at room temperature without any waiting period (Fig. S21, ESI†). From the UV-vis absorption, the SPR peak of Ag NPs was decreased, suggesting a similar mechanism as we have observed in the case of basic AAs (Fig. S22, ESI†). The lifetime analysis further corroborates these results. As is evident from Fig. S23, ESI†, the lifetime components of WS<sub>2</sub> NSs were regenerated in the presence of ubiquitin with similar contributions [9.1 ns (54.4) and 2.8 ns (45.6) for WS<sub>2</sub> NSs and 8.5 ns (56.4) and 2.3 ns (43.6) for the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP nanocomposite with ubiquitin]. The zeta potential values of the WS<sub>2</sub> NS–Ag<sub>2</sub>O NP nanocomposite with ubiquitin are also commensurate with the value of WS<sub>2</sub> NSs (Table S3, ESI†). The excitation spectrum (Fig. S24, ESI†) of WS<sub>2</sub> NSs at an emission position of 450 nm shows two peaks at ca. 265 and 360 nm, referring to two electronic states



**Table 1** The recovery and relative standard deviation of the measurements obtained from real sample analysis

| Sample       | Spiked ( $\mu\text{M}$ ) | Observed (mean; $\mu\text{M}$ ) | Recovery % | % RSD |
|--------------|--------------------------|---------------------------------|------------|-------|
| Blood sample | 1.5                      | 1.5188 $\pm$ 0.044              | 101.25     | 1.67  |
|              | 2                        | 2.0038 $\pm$ 0.027              | 100.19     | 0.81  |
|              | 2.5                      | 2.5037 $\pm$ 0.019              | 100.15     | 0.97  |
| Urine sample | 1.75                     | 1.7716 $\pm$ 0.072              | 101.24     | 1.75  |
|              | 2                        | 2.0031 $\pm$ 0.093              | 100.16     | 2.03  |
|              | 2.25                     | 2.2772 $\pm$ 0.108              | 101.21     | 2.14  |
|              | 2.5                      | 2.5232 $\pm$ 0.092              | 100.93     | 1.66  |

available for the system. The WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite, on the other hand, has only one major peak at around 310 nm. The disappearance of the longer wavelength peak is attributed to the effective energy transfer due to changes in surface states. However, the re-appearance of the longer wavelength excitation peak ( $\sim$ 360 nm) in the case of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite with ubiquitin at 360 nm verifies the shutting off of FRET between WS<sub>2</sub> NSs and Ag<sub>2</sub>O NPs in the presence of protein and the regeneration of the fluorescence intensity. The mass spectra of pure protein and protein in the presence of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP composite are given in Fig. S25a and b, ESI†. The characteristic protein peaks are evident in the mass spectrum of pure protein. In contrast, these peaks are absent in the case of the analyte-containing sensor solution (ubiquitin + WS<sub>2</sub> NS-Ag<sub>2</sub>O NP solution) (see Fig. S25a, ESI†). This implies that the protein was bound to the Ag<sub>2</sub>O NPs and did not appear in the mass spectrum.

Evaluation of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite utility as a sensor for basic AAs in real sample detection was also performed using blood and urine samples. Detection and quantification of Lys are significant, as they give a straightforward diagnosis of various disorders and diseases with the presence of Lys in urine and plasma.<sup>50</sup> We chose Lys as our analyte for real sample analysis. The samples were prepared by spiking known concentrations of Lys and the fluorescence spectra were recorded after addition to the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP nanocomposite at room temperature in buffer solutions. The results obtained are shown in Table 1. The recovery and relative standard deviation of the measurements (% RSD) were evaluated to check the reliability of the present sensor. For both the blood sample and the urine sample, excellent recovery within 2% RSD was obtained, which is superior to those of already reported methods.<sup>51–53</sup> The result indicates the reliability of the sensor for analytical applications.

## 4. Conclusion

In the present study, we demonstrated a luminescent 2D material, exfoliated WS<sub>2</sub> nanosheets, for sensing biologically relevant molecules. Highly fluorescent and stable WS<sub>2</sub> nanosheets were synthesized using a facile hydrothermal synthetic route. Interestingly, Ag<sub>2</sub>O NPs were formed on the WS<sub>2</sub> nanosheets instantaneously by the addition of the AgNO<sub>3</sub>

solution. The formation of Ag<sub>2</sub>O NPs leads to the reduction of PL from WS<sub>2</sub> NSs via FRET. We investigated the effect of positively charged basic nanoparticles on the PL of the WS<sub>2</sub> NS-Ag<sub>2</sub>O NP composite and fabricated a turn-on sensor for three basic amino acids viz. His, Lys and Arg. The easy, straightforward synthetic method for the formation of stable and highly fluorescent WS<sub>2</sub> NSs without using any hazardous solvents, the *in situ* formation of stable Ag<sub>2</sub>O NPs without an extra stabilizing agent, and devising a simple aggregation strategy to shut the FRET pathway, thereby ensuring a turn-on detection and screening of all the basic amino acids, are the highlights of the present sensor system. The practical application of the sensor was demonstrated for the detection of protein: ubiquitin and Lys from real samples – blood and urine.

## Live subject statement

All experiments were performed in accordance with the guidelines “National ethical guidelines for biomedical and health research involving human participants, Indian Council for Medical Research (ICMR), Government of India”, and approved by the ethics committee at “Institutional Ethics Committee (IEC), Indian Institute of Space Science and Technology, Thiruvananthapuram”. Informed consents were obtained from the human participants of this study.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors acknowledge IIST Thiruvananthapuram for funding and the PSG Institute of Advanced Studies, Coimbatore and The National Institute for Interdisciplinary Science and Technology, Thiruvananthapuram for analysis.

## References

- 1 X. Gu, B. Li and R. Yang, *J. Appl. Phys.*, 2016, **119**, 085106.
- 2 M.-W. Lin, I. I. Kravchenko, J. Fowlkes, X. Li, A. A. Puretzky, C. M. Rouleau, D. B. Geohegan and K. Xiao, *Nanotechnology*, 2016, **27**, 165203.
- 3 G. Eda, H. Yamaguchi, D. Voiry, T. Fujita, M. Chen and M. Chhowalla, *Nano Lett.*, 2011, **11**, 5111–5116.
- 4 P. Neema, A. M. Tomy and J. Cyriac, *TrAC, Trends Anal. Chem.*, 2020, **124**, 115797.
- 5 D. Jariwala, V. K. Sangwan, L. J. Lauhon, T. J. Marks and M. C. Hersam, *ACS Nano*, 2014, **8**, 1102–1120.
- 6 Q. H. Wang, K. Kalantar-Zadeh, A. Kis, J. N. Coleman and M. S. Strano, *Nat. Nanotechnol.*, 2012, **7**, 699–712.
- 7 A. Splendiani, L. Sun, Y. Zhang, T. Li, J. Kim, C. Chim, G. Galli and F. Wang, *Nano Lett.*, 2010, **10**, 1271–1275.
- 8 Y. Zhou and J. Yoon, *Chem. Soc. Rev.*, 2012, **41**, 52–67.
- 9 S. S. Sharma and K.-J. Dietz, *J. Exp. Bot.*, 2006, **57**, 711–726.
- 10 H. Yoshida, Y. Nakano, K. Koiso, H. Nohta, J. Ishida and M. Yamaguchi, *Anal. Sci.*, 2001, **17**, 107–112.





- 11 D. Xiong, M. Chen and H. Li, *Chem. Commun.*, 2008, 880–882.
- 12 T. Liu, N. Li, J. X. Dong, Y. Zhang, Y. Z. Fan, S. M. Lin, H. Q. Luo and N. B. Li, *Biosens. Bioelectron.*, 2017, **87**, 772–778.
- 13 X. Lu, W. Wang, Q. Dong, X. Bao, X. Lin, W. Zhang, X. Dong and W. Zhao, *Chem. Commun.*, 2015, **51**, 1498–1501.
- 14 S. T. R. Naqvi, T. Rasheed, M. N. Ashiq, M. N. ul Haq, S. Majeed, B. Fatima, R. Nawaz, D. Hussain and S. Shafi, *Polyhedron*, 2020, 114426.
- 15 H. Liu, J. Shao, L. Shi, W. Ke, F. Zheng and Y. Zhao, *Sens. Actuators, B*, 2020, **304**, 127333.
- 16 E. V. Suprun, E. V. Karpova, S. P. Radko and A. A. Karyakin, *Electrochim. Acta*, 2020, **331**, 135289.
- 17 M. Hasanzadeh, A. Karimzadeh, N. Shadjou, A. Mokhtarzadeh, L. Bageri, S. Sadeghi and S. Mahboob, *Mater. Sci. Eng., C*, 2016, **68**, 814–830.
- 18 J. L. da Silva, M. A. Beluomini, G. C. Sedenho and N. R. Stradiotto, *Microchem. J.*, 2017, **134**, 374–382.
- 19 B. R. Kolanu, V. Boddula, S. Vadakedath and V. Kandi, *Cureus*, 2017, **9**(3), e1091.
- 20 G. Patel and S. Menon, *Chem. Commun.*, 2009, 3563–3565.
- 21 Y. Sun, L. Zhang and H. Li, *New J. Chem.*, 2012, **36**, 1442–1444.
- 22 T. Minami, N. A. Esipenko, B. Zhang, L. Isaacs and P. Anzenbacher, *Chem. Commun.*, 2014, **50**, 61–63.
- 23 H. Zeng, G.-B. Liu, J. Dai, Y. Yan, B. Zhu, R. He, L. Xie, S. Xu, X. Chen and W. Yao, *Sci. Rep.*, 2013, **3**, 1–5.
- 24 X. Zhang, X.-F. Qiao, W. Shi, J.-B. Wu, D.-S. Jiang and P.-H. Tan, *Chem. Soc. Rev.*, 2015, **44**, 2757–2785.
- 25 Y. Zhao, X. Luo, H. Li, J. Zhang, P. T. Araujo, C. K. Gan, J. Wu, H. Zhang, S. Y. Quek and M. S. Dresselhaus, *Nano Lett.*, 2013, **13**, 1007–1015.
- 26 H. Li, J. Wu, Z. Yin and H. Zhang, *Acc. Chem. Res.*, 2014, **47**, 1067–1075.
- 27 A. Molina-Sanchez and L. Wirtz, *Phys. Rev. B*, 2011, **84**, 155413.
- 28 H. Zeng, G.-B. Liu, J. Dai, Y. Yan, B. Zhu, R. He, L. Xie, S. Xu, X. Chen and W. Yao, *Sci. Rep.*, 2013, **3**, 1608.
- 29 Y. Cheng, J. Peng, B. Xu, H. Yang, Z. Luo, H. Xu, Z. Cai and J. Weng, *IEEE Photonics J.*, 2016, **8**, 1.
- 30 X. Zhang, W. Lei, X. Ye, C. Wang, B. Lin, H. Tang and C. Li, *Mater. Lett.*, 2015, **159**, 399–402.
- 31 G.-Q. Han, Y.-R. Liu, W.-H. Hu, B. Dong, X. Li, Y.-M. Chai, Y.-Q. Liu and C.-G. Liu, *Mater. Chem. Phys.*, 2015, **167**, 271–277.
- 32 K. Zhang, L. Fu, W. Zhang, H. Pan, Y. Sun, C. Ge, Y. Du and N. Tang, *ACS Omega*, 2018, **3**, 12188–12194.
- 33 X. Mao, Y. Xu, Q. Xue, W. Wang and D. Gao, *Nanoscale Res. Lett.*, 2013, **8**, 430.
- 34 S. X. Leong, C. C. Mayorga-Martinez, X. Chia, J. Luxa, Z. Sofer and M. Pumera, *ACS Appl. Mater. Interfaces*, 2017, **9**, 26350–26356.
- 35 A. Bayat and E. Saievar-Iranizad, *J. Lumin.*, 2017, **185**, 236–240.
- 36 A. K. Mishra, K. V. Lakshmi and L. Huang, *Sci. Rep.*, 2015, **5**, 15718.
- 37 W. Zhao, Z. Ghorannevis, L. Chu, M. Toh, C. Kloc, P.-H. Tan and G. Eda, *ACS Nano*, 2013, **7**, 791–797.
- 38 B. Mahler, V. Hoepfner, K. Liao and G. A. Ozin, *J. Am. Chem. Soc.*, 2014, **136**, 14121–14127.
- 39 B. Zhu, X. Chen and X. Cui, *Sci. Rep.*, 2015, **5**, 9218.
- 40 L. Yuan and L. Huang, *Nanoscale*, 2015, **7**, 7402–7408.
- 41 K. T. Sullivan, C. Wu, N. W. Piekkel, K. Gaskell and M. R. Zachariah, *Combust. Flame*, 2013, **160**, 438–446.
- 42 M. Padilla Villavicencio, A. Escobedo Morales, M. Ruiz Peralta, M. Sánchez-Cantú, L. Rojas Blanco, E. Chigo Anota, J. H. Camacho Garcia and F. Tzompantzi, *Catal. Lett.*, 2020, **150**, 2385–2399.
- 43 Z. Lin, Y. Lu and J. Huang, *Cellulose*, 2019, **26**, 6683–6700.
- 44 L. Xu, B. Wei, W. Liu, H. Zhang, C. Su and J. Che, *Nanoscale Res. Lett.*, 2013, **8**, 1–7.
- 45 T. Tsuru, T. Shutou, S.-I. Nakao and S. Kimura, *Sep. Sci. Technol.*, 1994, **29**, 971–984.
- 46 Z. Liu, T. Fan, D. Zhang, X. Gong and J. Xu, *Sens. Actuators, B*, 2009, **136**, 499–509.
- 47 W. Lin, L. Long, L. Yuan, Z. Cao and J. Feng, *Anal. Chim. Acta*, 2009, **634**, 262–266.
- 48 V. Chau, J. W. Tobias, A. Bachmair, D. Marriott, D. J. Ecker, D. K. Gonda and A. Varshavsky, *Science*, 1989, **243**, 1576–1583.
- 49 C. Wloka, V. Van Meervelt, D. van Gelder, N. Danda, N. Jager, C. P. Williams and G. Maglia, *ACS Nano*, 2017, **11**, 4387–4394.
- 50 W. Song, W. Duan, Y. Liu, Z. Ye, Y. Chen, H. Chen, S. Qi, J. Wu, D. Liu and L. Xiao, *Anal. Chem.*, 2017, **89**, 13626–13633.
- 51 V. D. O. Torres, R. C. Piva, W. F. A. Junior and C. A. L. Cardoso, *Orbital: Electron. J. Chem.*, 2018, **10**, 1–8.
- 52 V. I. Chalova, I. B. Zabala-Díaz, C. L. Woodward and S. C. Ricke, *World J. Microbiol. Biotechnol.*, 2008, **24**, 353–359.
- 53 A. Hayat, T. M. Jahangir, M. Y. Khuhawar, M. Alamgir, A. J. Siddiqui and S. G. Musharraf, *J. Cereal Sci.*, 2014, **60**, 356–360.

