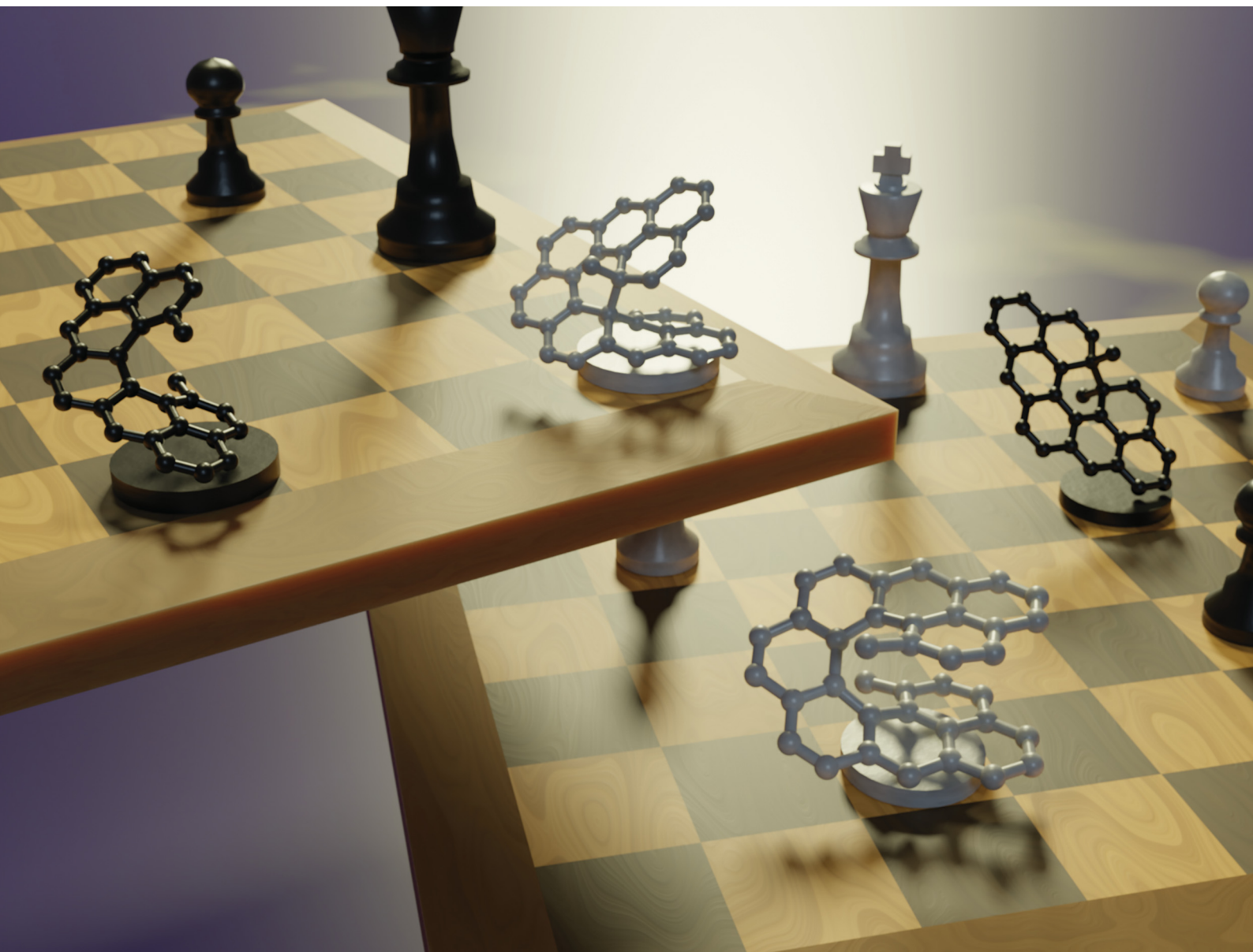


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Dimethylnonacethrene – *en route* to a magnetic switch†

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Dimethylnonacethrene is the first derivative of the cethrene family that is energetically more stable than the product of its electrocyclic ring closure. Compared to the shorter homologue dimethylcethrene, the new system is EPR-active, because of a significantly lowered singlet–triplet gap, and displays remarkable stability. Our results suggest that adjustment of the steric bulk in the fjord region can enable realisation of diradicaloid-based magnetic photoswitches.

Open-shell molecular fragments of graphene,^{1–4} a subclass of polycyclic aromatic hydrocarbons (PAHs) that display magnetic properties, hold prospects of making light-weight and flexible materials, and eliminating the use of rare earth elements in electronic devices.^{5–9} An attractive opportunity to modulate magnetic properties of a device on demand are PAH-based magnetic photoswitches that operate under ambient conditions.^{10–13} There are different ways of tuning the magnetic properties in PAHs, for example, by variation of substituents^{14,15} or chemical treatment,^{16,17} but the ability to reversibly turn magnetic properties off by using light remains a design challenge.^{13,18,19}

To control magnetic properties in such systems by means of secondary orbital, or through-space, interactions,^{20–22} Juriček and coworkers developed a family of C-shaped helical PAHs named cethrenes (Fig. 1).^{21,23,24} This class of compounds display either a singlet diradicaloid or a triplet diradical ground

state, and a varying degree of cofacial overlap between the terminal phenalenyl units.²⁵ Using a diphenyl derivative of the first homologue (Fig. 1A), we demonstrated²¹ that through-space interactions (bonding within the HOMO, antibonding within the LUMO) lower the energy gap between the singlet ground state and the triplet excited state ($\Delta E_{ST} \sim 6 \text{ kcal mol}^{-1}$) compared to the planar Z-shaped isomer heptazethrene ($\Delta E_{ST} \sim 9 \text{ kcal mol}^{-1}$). As a result, an electron paramagnetic resonance (EPR) signal originating from the thermally populated triplet state of cethrene is observed, while heptazethrene is EPR silent.²⁶

During these investigations, we found that cethrene undergoes a formally symmetry-forbidden thermal 6π electrocyclic ring closure^{27–31} with a surprisingly low activation barrier,³² followed by an oxidation to a flat hydrocarbon. To prevent the irreversible oxidation step and benefit from the photochemical electrocyclic process, methyl groups were installed in the fjord region, yielding a reversible chiral photoswitch^{33,34} (Fig. 1C). Unfortunately, the improved stability came at the cost of the EPR response, as the installment of the methyl substituents led to an increase of ΔE_{ST} by $\sim 4 \text{ kcal mol}^{-1}$ compared to the system without methyl groups. In addition to the electronic effect, the methyl groups increase ΔE_{ST} by increasing the helical twist which causes weakening of the through-space orbital overlap along with a significantly improved configurational stability.^{35,36} Despite the absence of the EPR response, we postulated that the working principle of dimethylcethrene could be applied in the design of an all-organic magnetic photoswitch, a concept that was successfully realised by Dumele and coworkers in a system operating at cryogenic temperatures.³⁷

To achieve a chiral magnetic photoswitch that would function at ambient temperature, we pursued the third homologue in the series, nonacethrene³⁸ (Fig. 1B), featuring a helicene backbone extended by two rings compared to cethrene. Even though the necessity to functionalise the fjord region was thought to be alleviated, we could only observe what appears to be the product of a 6π electrocycloisolation of nonacethrene followed by a 1,5-hydrogen shift. In this work, we pursued the next step toward

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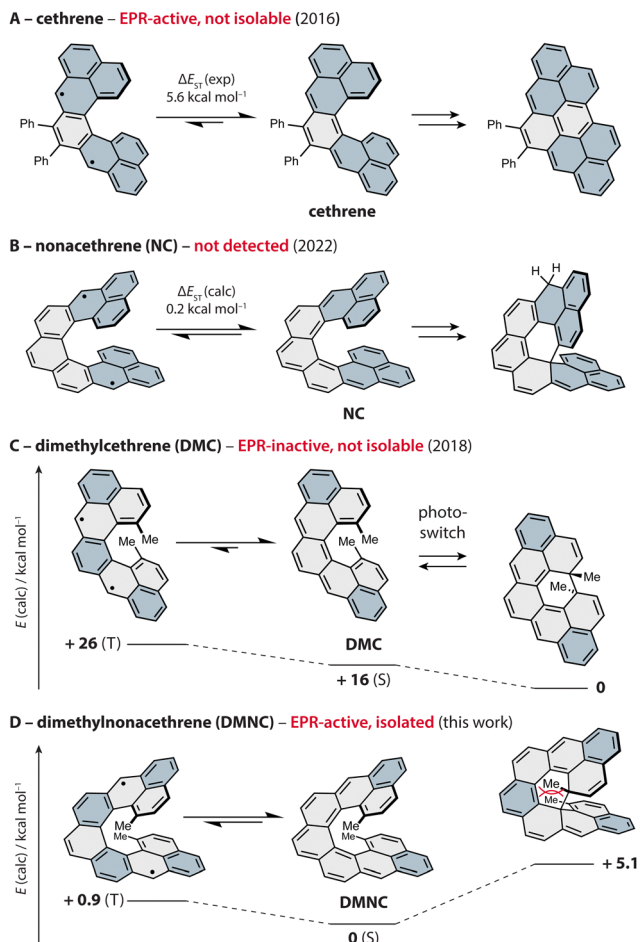
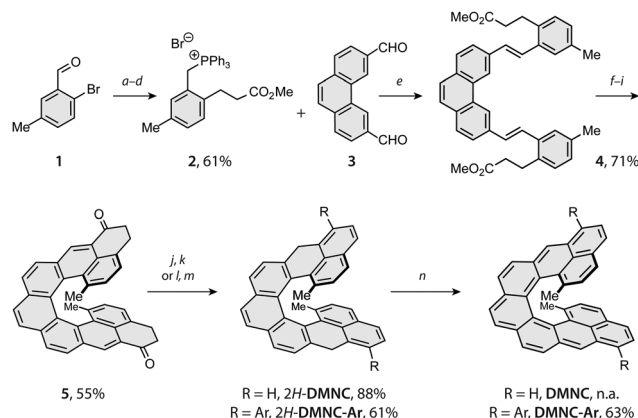


Fig. 1 Progress in the development of magnetic photoswitches based on cethrenes, shown as thermal equilibrium between the singlet ground state (S, middle) and triplet excited state (T, left); ΔE_{ST} from ref. 21 (A), 38 (B) and 34 (C). The isolated products of electrocyclisation are shown on the right, except for **DMNC** (product not observed). The phenalenyl units (A, B) and Clar sextets (C, D) are highlighted in blue. Relative energies for **DMC** (0, +16) and **DMNC** (all) were obtained via DFT (ω B97DX/Def2SVP) and CAS(2,2) and the energies of the singlet ground states via broken-symmetry calculations.

realisation of a cethrene-based magnetic photoswitch and present dimethylnonacethrene (**DMNC**, Fig. 1D, left), the first isolable cethrene derivative. In contrast to all other reported cethrenes, **DMNC** is lower in energy compared to the product of its 6π electrocyclisation, which was not observed even upon light irradiation, and has the lowest experimentally estimated ΔE_{ST} in the series ($\sim 1 \text{ kcal mol}^{-1}$).

The synthetic strategy toward **DMNC** capitalises on the route developed for nonacethrene³⁸ (**NC**) and employs two photochemical steps to build the [7]helicene backbone (Scheme 1). Starting from a commercial 2-bromo-5-methylbenzaldehyde (**1**), intermediate **2** was prepared in a series of four steps, including Heck coupling with methyl acrylate (*a*), reduction of the double bond and the aldehyde moiety with sodium borohydride and nickel(II) chloride (*b*), substitution of the benzylic hydroxyl group by a bromide with phosphorus tribromide (*c*) and treatment with triphenylphosphine (*d*) to afford the desired phosphonium salt **2**



Scheme 1 Reaction conditions: (a) $\text{Pd}(\text{OAc})_2$, TBAB, PPh_3 , K_2CO_3 , methyl acrylate, DMF, 90°C , 20 h; (b) NiCl_2 , NaBH_4 , THF/MeOH 1 : 1, 0°C to rt, 1 h; (c) PBr_3 , Et_2O , rt, 18 h; (d) PPh_3 , toluene, 100°C , 16 h; (e) NaOMe , THF, rt, 20 h; (f) $h\nu$, PO, toluene, rt, 2 h; (g) LiOH , THF/ H_2O 10 : 1, 80°C , 16 h; (h) $(\text{COCl})_2$, 65°C , 2 h; (i) TiCl_4 , CH_2Cl_2 , -50 to -20°C , 6 h; (j) NaBH_4 , $\text{CH}_2\text{Cl}_2/\text{EtOH}$ 2 : 1, rt, 5 h; (k) *p*-TSA, toluene, 90°C , 5 min; (l) ArMgBr , THF, -78°C to rt, 18 h; (m) TFA, CH_2Cl_2 , rt, 5 min; (n) *p*-chloranil, benzene- d_6 , 50°C , 10 min; TBAB = tetrabutylammonium bromide, DMF = *N,N'*-dimethylformamide; THF = tetrahydrofuran; rt = room temperature; PO = propylene oxide; *p*-TSA = *para*-toluenesulfonic acid; TFA = trifluoroacetic acid. Ar = 3,5-di-*tert*-butylphenyl.

in an overall 61% yield. Next, a Wittig reaction of **2** with dialdehyde **3**³⁸ in the presence of sodium methoxide (*e*) gave diester **4** in a 71% yield. Subsequent photocyclodehydrogenation of **4** using a medium-pressure mercury lamp and iodine as an oxidant (*f*) afforded [7]helicene intermediate (**9**; see the ESI[†]) in a 75% yield, the structure of which was confirmed by single-crystal X-ray diffraction (SC-XRD; ESI[†]). Upon hydrolysis with lithium hydroxide (*g*), the diacid intermediate was transformed into the corresponding diacyl chloride (*h*), followed by a Friedel-Crafts acylation mediated by titanium tetrachloride (*i*) to afford the key intermediate **5** in a 73% yield over the three steps (for SC-XRD of **5**, see the ESI[†]).

From **5**, first the synthesis of parent **DMNC** was attempted, where a reduction with sodium borohydride followed by a dehydration with a catalytic amount of *p*-toluenesulfonic acid afforded the dihydro precursor **2H-DMNC** in an 88% yield over the two steps. The final oxidation step was performed in argon-saturated benzene with *p*-chloranil, yielding an insoluble, EPR-active solid (Fig. S49, ESI[†]). Because structure identification was hampered on account of limited solubility and the fact that presence of monoradical impurities could not be excluded, solubilising 3,5-di-*tert*-butylphenyl groups were installed via a nucleophilic 1,2-addition to the carbonyl groups of **5**. These substituents provided also kinetic stabilisation, as they were installed next to the positions that display the highest spin density in the triplet excited state (Table S4, ESI[†]). After the addition, dehydration occurred either directly or was promoted by adding a drop of trifluoroacetic acid to afford the dihydro precursor **2H-DMNC-Ar** in form of the most stable isomer featuring two naphthalene and one phenanthrene units. Its structure was supported by SC-XRD (Fig. 2, left). The final step



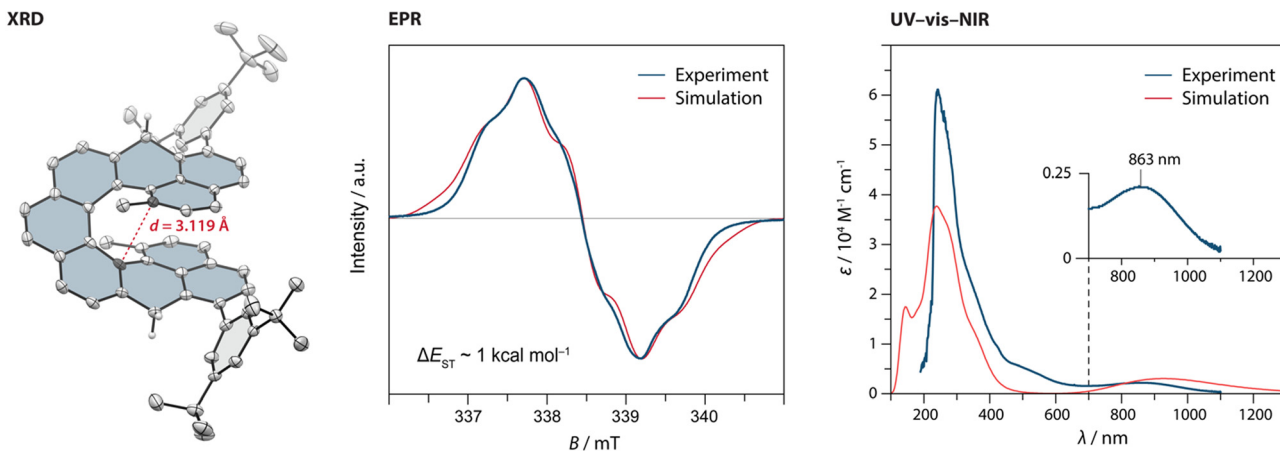


Fig. 2 (left) The solid-state structure of **2H-DMNC-Ar** with the thermal ellipsoids shown at a 50% probability level. The nonacethrene core is highlighted in blue. All hydrogen atoms, except for those of the CH₂ groups, are omitted for clarity. (centre) Comparison of the experimental and simulated electron paramagnetic resonance (EPR) spectra of a solution of **DMNC-Ar** (8.7 mM in benzene-*d*₆) at 298 K. Experimental parameters are shown in Table S8 (ESI[†]) (right) Comparison of the experimental (at 298 K) and simulated UV-vis-NIR absorption spectra of a solution of **DMNC-Ar** (CH₂Cl₂; TD-DFT/B3LYP/6-31(*d,p*)).

was performed in argon-saturated benzene-*d*₆ with *p*-chloranil as the oxidant, and the conversion was followed by ¹H NMR spectroscopy. At room temperature, the conversion was relatively slow but heating to 50 °C led to a full conversion within 10 min, as indicated by a complete disappearance of the signals that belong to the starting material (Fig. S1, ESI[†]) and a matching *m/z* peak in HRMS. No new peaks could be observed, apart from a very broad signal spread over the aromatic region. The desired **DMNC-Ar** was isolated as a deep-purple solid ($\lambda_{\text{max}} = 863$ nm, Fig. 2, right) after column chromatography in a 63% yield.

DMNC-Ar was studied by EPR spectroscopy as a solid powder, in a solid *o*-terphenyl matrix and in a benzene solution (Fig. 2, centre and Fig. S48, ESI[†]). In all three cases, a strong but broadened signal was detected without the observation of a forbidden half-field signal. The broadening is presumably the result of several unresolved small hyperfine couplings. In the benzene solution (8.7 mM), a discernible quintet multiplicity is observed, which is in a good qualitative agreement with the simulation (Fig. 2, centre) performed using DFT-calculated hyperfine coupling constants (Tables S4 and S8, ESI[†]). The shape and the intensity of the signal remained almost unchanged after the benzene sample was kept for a year in a sealed EPR tube, demonstrating that **DMNC-Ar** is stable in the absence of oxygen and does not undergo 6 π electrocyclic ring closure, in contrast to its shorter homologue **DMC**. Notably, the signal could be detected even after two weeks of the sample being kept in a non-deoxygenated CH₂Cl₂ solution as well as using a solution prepared from a solid powder that was stored under ambient conditions for several months. For a comparison, kinetically stabilised nonazethrene has a half-life of 16 h³⁹ and nonacene is reported to be stable at room temperature only under anhydrous argon atmosphere.⁴⁰

To determine ΔE_{ST} , variable-temperature EPR measurements were performed. The intensity of the signal arising from the thermally populated triplet excited state decreased with lowering the temperature, and the fitting of these data using the Bleaney–Bowers equation⁴¹ (Fig. S51, ESI[†]) gave a ΔE_{ST}

estimate of ~ 1 kcal mol⁻¹, matching well the values (0.3–2.0 kcal mol⁻¹) calculated using different methods (Table S5, ESI[†]). Compared to the flat isomer nonazethrene (5.6 kcal mol⁻¹), linear nonacene (5.8 kcal mol⁻¹)³⁹ as well as non-planar tetrabenzoperylene (3.4 kcal mol⁻¹),⁴² ΔE_{ST} of **DMNC-Ar** is significantly smaller, which illustrates the impact of through-space interactions. In addition, this value is smaller than that of the shorter homologue cethrene (5.6 kcal mol⁻¹),³² which reflects the gain of two (**DMNC**) versus one (cethrene) Clar sextet in the diradical resonance structure (see Fig. 1A, B, and Fig. S58 and Table S9 for discussion of NICS and ACID studies, ESI[†]). This trend is similar to that for zethrene and its higher homologues.^{43,44} Compared to **NC** ($\Delta E_{\text{ST}} \sim 0.2$ kcal mol⁻¹), ΔE_{ST} of **DMNC-Ar** is larger, which manifests the steric effect of the methyl groups leading to the increase of helical twist. For example, the distance between the carbon atoms that display the largest through-space orbital overlap (highlighted in Fig. 2, left) increases from 2.90 (**NC**; DFT) to 3.09 Å (**DMNC-Ar**; DFT), or from 2.974(0)³⁸ to 3.118(6) Å (XRD; Fig. 2, left) in the case of dihydro precursors, as observed³⁴ previously for cethrene and **DMC**.

To photochemically induce 6 π electrocyclicalisation, UV-vis-near-infrared (NIR) absorption studies were performed in CH₂Cl₂. **DMNC-Ar** displays a broad absorption band in the NIR region (Fig. 2, right; $\lambda_{\text{max}} = 863$ nm) characteristic of diradicaloid compounds with small HOMO–LUMO gaps. Irradiation of the sample with LED light sources (254, 365, 505, 740 and 800 nm) for several minutes at room or cryogenic (77 K, 2-methyl tetrahydrofuran matrix) temperature did not result in any change of the absorption spectrum. DFT calculations were performed to calculate the relative energies of **DMNC** (the “open” form) and its electrocyclic product (the “closed” form) where three different functionals (PBE0, B3LYP-D3 and ω B97DX) show the same trend: the open form is lower in energy by 5–12 kcal mol⁻¹, supporting all experimental observations. This feature is in stark contrast to all other reported cethrenes and results from a destabilising steric effect in the fjord region caused by the methyl groups. This effect is more pronounced in the closed form (~ 9 kcal mol⁻¹) than in the



open form ($\sim 4 \text{ kcal mol}^{-1}$), which leads to inversion of the relative energies (for NC, the closure is only slightly exoergic; see Table S8, ESI†). Considering the relatively small energy difference between the open and the closed form of DMNC, these results suggest that either the activation energy of the ring closure is too large or, more likely, the activation energy of the ring opening is too small,⁴⁵ which would result in a short-lived closed form that is undetectable even at cryogenic temperatures.

Our findings illustrate that similarly to planar zethrene analogues, π -extension is an effective tool to decrease ΔE_{ST} in helical cethrenes. Methyl substituents in the fjord region prevent reactivity previously observed for nonacethrene and, notably, invert the typical trend of relative energies for cethrenes and their electrocyclic products. Two aryl groups were installed in the periphery of DMNC to provide solubility and to promote kinetic stabilisation. The results of our spectroscopic measurements and computational studies all point toward inverted equilibrium between the open and the closed form, in contrast to all previously reported cethrenes. In addition, an extraordinary stability of the open form without significant decomposition was observed in solution and in the solid state over several months. The closed form was not observed even upon photoirradiation under ambient or cryogenic conditions. The adjustment of the steric bulk in the fjord region represents an opportunity for further optimisation to achieve bistability and a viable strategy to realise a magnetic photoswitch operating at ambient temperature.

Conflicts of interest

There are no conflicts to declare.

Notes and references

- W. Zeng and J. Wu, *Chemistry*, 2021, 7, 358–386.
- Y. Morita, S. Suzuki, K. Sato and T. Takui, *Nat. Chem.*, 2011, 3, 197–204.
- T. Kubo, *Bull. Chem. Soc. Jpn.*, 2021, 94, 2235–2244.
- S. Dong and Z. Li, *J. Mater. Chem. C*, 2022, 10, 2431–2449.
- G. Schweicher, G. Garbay, R. Jouclas, F. Vibert, F. Devaux and Y. H. Geerts, *Adv. Mater.*, 2020, 32, 1905909.
- Q. Li, Y. Zhang, Z. Xie, Y. Zhen, W. Hu and H. Dong, *J. Mater. Chem. C*, 2022, 10, 2411–2430.
- X. Hu, W. Wang, D. Wang and Y. Zheng, *J. Mater. Chem. C*, 2018, 6, 11232–11242.
- K. Liu, B. Ouyang, X. Guo, Y. Guo and Y. Liu, *npj Flexible Electron.*, 2022, 6, 1.
- P. Murto and H. Bronstein, *J. Mater. Chem. C*, 2022, 10, 7368–7403.
- S. Nakatsuji, *Chem. Soc. Rev.*, 2004, 33, 348–353.
- M. B. S. Wonink, B. P. Corbet, A. A. Kulago, G. B. Boursalian, B. de Bruin, E. Otten, W. R. Browne and B. L. Feringa, *J. Am. Chem. Soc.*, 2021, 143, 18020–18028.
- T. Nishiuchi, R. Ito, E. Stratmann and T. Kubo, *J. Org. Chem.*, 2020, 85, 179–186.
- H. Kurata, S. Kim, K. Matsumoto, T. Kawase and M. Oda, *Chem. Lett.*, 2007, 36, 386–387.
- Z. Lin, L. Chen, Q. Xu, G. Shao, Z. Zeng, D. Wu and J. Xia, *Org. Lett.*, 2020, 22, 2553–2558.
- W. Wang, L. Ge, G. Xue, F. Miao, P. Chen, H. Chen, Y. Lin, Y. Ni, J. Xiong, Y. Hu, J. Wu and Y. Zheng, *Chem. Commun.*, 2020, 56, 1405–1408.
- B. Wu, T. Wang, Y. Feng, Z. Zhang, L. Jiang and C. Wang, *Nat. Commun.*, 2015, 6, 6468.
- J. Guo, Y. Yang, C. Dou and Y. Wang, *J. Am. Chem. Soc.*, 2021, 143, 18272–18279.
- D. Čavlović and M. Juriček, *Chimia*, 2019, 73, 313–316.
- S. Hayami, S. M. Holmes and M. A. Halcrow, *J. Mater. Chem. C*, 2015, 3, 7775–7778.
- F. Saal, F. Zhang, M. Holzapfel, M. Stolte, E. Michail, M. Moos, A. Schmiedel, A.-M. Krause, C. Lambert, F. Würthner and P. Ravat, *J. Am. Chem. Soc.*, 2020, 142, 21298–21303.
- P. Ravat, T. Šolomek, M. Rickhaus, D. Häussinger, M. Neuburger, M. Baumgarten and M. Juriček, *Angew. Chem., Int. Ed.*, 2016, 55, 1183–1186.
- J. Huang, B. G. Sumpter, V. Meunier, Y.-H. Tian and M. Kertesz, *Chem. Mater.*, 2011, 23, 874–885.
- M. Juriček, *Chimia*, 2018, 72, 322–327.
- P. Ravat, T. Šolomek, P. Ribar and M. Juriček, *Synlett*, 2016, 1613–1617.
- S. Takamuku, M. Nakano and M. Kertesz, *Chem. – Eur. J.*, 2017, 23, 7474–7482.
- Y. Li, W.-K. Heng, B. S. Lee, N. Aratani, J. L. Zafra, N. Bao, R. Lee, Y. M. Sung, Z. Sun, K.-W. Huang, R. D. Webster, J. T. López Navarrete, D. Kim, A. Osuka, J. Casado, J. Ding and J. Wu, *J. Am. Chem. Soc.*, 2012, 134, 14913–14922.
- S. Pogodin and I. Agranat, *J. Am. Chem. Soc.*, 2003, 125, 12829–12835.
- T. Šolomek, P. Ravat and M. Juriček, *Trends Chem.*, 2019, 1, 705–706.
- K. Uchida, S. Ito, M. Nakano, M. Abe and T. Kubo, *J. Am. Chem. Soc.*, 2016, 138, 2399–2410.
- J. Downing, V. Dvořák, J. Kolc, A. Manzara and J. Michl, *Chem. Phys. Lett.*, 1972, 17, 70–73.
- J. Kolc and J. Michl, *J. Am. Chem. Soc.*, 1970, 92, 4147–4148.
- T. Šolomek, P. Ravat, Z. Mou, M. Kertesz and M. Juriček, *J. Org. Chem.*, 2018, 83, 4769–4774.
- P. Ravat, T. Šolomek and M. Juriček, *ChemPhotoChem*, 2019, 3, 180–186.
- P. Ravat, T. Šolomek, D. Häussinger, O. Blacque and M. Juriček, *J. Am. Chem. Soc.*, 2018, 140, 10839–10847.
- P. Ravat, R. Hinkelmann, D. Steinebrunner, A. Prescimone, I. Bodoky and M. Juriček, *Org. Lett.*, 2017, 19, 3707–3710.
- P. Ravat, *Chem. – Eur. J.*, 2021, 27, 3957–3967.
- K. Günther, N. Grabicki, B. Battistella, L. Grubert and O. Dumele, *J. Am. Chem. Soc.*, 2022, 144, 8707–8716.
- D. Čavlović, D. Häussinger, O. Blacque, P. Ravat and M. Juriček, *JACS Au*, 2022, 2, 1616–1626.
- R. Huang, H. Phan, T. S. Herng, P. Hu, W. Zeng, S. Dong, S. Das, Y. Shen, J. Ding, D. Casanova and J. Wu, *J. Am. Chem. Soc.*, 2016, 138, 10323–10330.
- A. Jančařík, J. Holec, Y. Nagata, M. Šámal and A. Gourdon, *Nat. Commun.*, 2022, 13, 223.
- B. Bleaney and K. D. Bowers, *Proc. R. Soc. London, Ser. A*, 1952, 214, 451–465.
- J. Liu, P. Ravat, M. Wagner, M. Baumgarten, X. Feng and K. Müllen, *Angew. Chem., Int. Ed.*, 2015, 54, 12442–12446.
- R. Umeda, D. Hibi, K. Miki and Y. Tobe, *Org. Lett.*, 2009, 11, 4104–4106.
- P. Hu and J. Wu, *Can. J. Chem.*, 2017, 95, 223–233.
- The small HOMO–LUMO gaps of 1.69 and 1.86 eV (ref. 34) in cethrene and DMC, respectively, facilitate the formally forbidden thermal ring closure (E_a of ~ 14 and $\sim 20 \text{ kcal mol}^{-1}$, respectively). In DMNC, the HOMO–LUMO gap of 0.94 eV (DFT/B3LYP/6-311+G(d) on ω B97DX/Def2SVP geometry, gas phase) is significantly smaller than that of cethrene, which suggests that E_a of the thermal ring opening is well below 14 kcal mol^{-1} .

