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rsc.li/catalysisA structure-sensitive descriptor for the design of active sites on MoS₂ catalysts†Hai-Yan Su,^a Federico Calle-Vallejo^{*bc} and Keju Sun^{*d}

MoS₂ catalysts hold great promise for numerous reactions of industrial and technological interest. However, general guidelines for the design of their active sites remain elusive. We hypothesize that this is because the link between their geometric structure and reactivity is yet to be established at the atomic scale. Here we show that *cn*, a metric based on the number of sulfur atoms coordinated to Mo atoms, captures the trends in reactivity of MoS₂ catalysts with various sulfur vacancy contents. This is illustrated for the adsorption energies of numerous monatomic and polyatomic species. More importantly, *cn* can be used to predict the reaction and activation energies of common formation and dissociation reactions in catalysis. Finally, *cn* is used to outline the optimal configuration of MoS₂ active sites for the electrocatalytic hydrogen evolution reaction: the highest exchange current density corresponds to terrace sites with adjacent S vacancies with *cn* in the range of 4.33 to 4.67.

1. Introduction

The computational description of catalytic reactions on transition metals has achieved great advances in the past three decades.^{1–3} The atomic-scale surface chemistry that determines the observable macroscopic kinetics has been established in numerous cases, and key descriptors such as adsorption energies of intermediates,^{1,2,4,5} band centers,¹ and coordination numbers,^{6,7} among others,^{8–10} have been shown to correlate with catalytic activities. These correlations are the groundwork for the rational design of metal catalysts. However, catalysts based on transition metal compounds

wherein the metal atoms have nonzero oxidation states are still insufficiently understood in view of their inherent complexity.

Indeed, the surface reactivity of oxidized transition metal compounds depends on the interplay between the geometric and electronic structures, the local stoichiometry of the surface, and the oxidation states of the components.^{11,12} This has traditionally limited the generality of analyses for those catalysts and motivated the appearance of various *ad hoc* descriptors. For example, the number of outer electrons and the occupancy of *e_g* orbitals correlate with adsorption energies and the oxygen evolution activity on monoxides, perovskites and spinels.^{13–15} The average oxygen 2p-state energy, oxygen p-band center and vacancy formation energies capture the activity trends for the oxygen reduction reaction (ORR) of solid oxide fuel cell cathodes on rutiles and perovskites.^{16,17} In addition, the oxygen vacancy formation energy is a descriptor for the trends in C–O and N–O bond scission of rutile oxides.¹⁸ The oxygen vacancies of TiO₂ electroreduce formic acid to methanol and the activities depend linearly on the formation energies of such vacancies.¹⁹ The “adjusted coordination number”, can be used to predict the *H adsorption energy and C–H activation energies at the oxygen sites on V₂O₃, Cr₂O₃, Co₃O₄, and NiO.²⁰ Finally, the activity of Ni₃S₂ catalysts is determined by the Ni–S coordination number: Ni surface sites with three S nearest neighbors provide nearly optimal energetic conditions for ORR electrocatalysis.²¹

Molybdenum disulfide (MoS₂) is widely studied as a catalyst in various processes including hydrodesulfurization,²² hydrogen evolution reaction

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† Electronic supplementary information (ESI) available: Examples for *cn* calculations, correlations between d-band center and S vacancy formation energies and adsorption energies, correlations between cohesive energies and vacancy formation energies and number of S atoms coordinated to Mo atoms, adsorption-energy scaling relations, differential charge density maps for atomic adsorption, transition-state configurations, adsorption configurations on Mo edges, adjacent and far S vacancy configurations, tabulated data, and details of the electrocatalytic modelling are provided in the ESI. See DOI: <https://doi.org/10.1039/d3cy00575e>



(HER),^{23,24} and catalytic hydrogenation.^{25–27} While MoS₂ edges are generally accepted as the catalytic active sites,^{22,23,25} recent studies show that the MoS₂(001) basal plane modified by doping transition metals or creating S vacancies displays enhanced catalytic performance for the HER,^{28–30} CO₂ hydrogenation,^{31,32} and hydrodeoxygenation reactions.³³ In particular, a Co-doped MoS₂(001) basal plane shows long-term durability with >5000 cycles and an overpotential of 156 mV at a current density of 10 mA cm^{−2} for the HER, comparable to the most active MoS₂-based electrocatalysts in acidic media.³⁰ These findings open up new avenues for improved MoS₂ catalyst design, as the (001) basal plane has a great density of sites compared to edges. However, the nanoscale factors that enhance the activity remain elusive. Shedding light into this exciting subject requires a direct link between the catalytic activity and the atomic-scale properties of the material. In fact, the design and implementation of MoS₂ catalysts could be greatly streamlined by linking the geometric structure of their active sites to the catalytic activity.

To quantitatively outline the active sites of MoS₂ catalysts, here we study the adsorption of various monatomic (H, B, C, N, O, and F) and polyatomic species (CO, NO, CH, OH, SH, CH₂, CH₃, NH₂, CNH₂, and CCH₃) at undercoordinated Mo atoms (namely, S vacancies) on the MoS₂(001) basal plane for a wide range of sulfur vacancy coverages ($\theta = 1/15$ – $7/15$ ML). We find that the total number of S atoms coordinated to Mo atoms at sulfur vacancies (denoted by *cn*) smoothly captures the trends in adsorption energies. In addition, *cn* is well correlated with the enthalpy and activation energy of

elementary reactions such as OH, H₂O, CH₄ and CH₃OH formation and NO dissociation.

For completeness, we benchmark *cn* against common descriptors such as the formation energies of sulfur vacancies (ΔE_{vac}), cohesive energy (E_{coh}), work function (ϕ), d-band center of the Mo atoms (ϵ_d), charge transferred to the adsorbates, integrated crystal orbital overlap population (*coop*) and crystal orbital Hamilton population (*cohp*). We find that *cn* matches or even surpasses some of those descriptors, while being more easily assessed. Finally, we use *cn* to outline the optimal coordination environment of MoS₂ terrace sites for the HER.

2. Results

We created S vacancies in the range of $\theta = 1/15$ – $7/15$ ML, by sequentially removing one to seven surface S atoms in a clockwise manner, as shown in Fig. 1. The total number of S atoms coordinated to Mo atoms (*cn*) at a S vacancy (dashed circle, Fig. 1) is defined as $cn = N_S/N_{\text{Mo}}$, where N_S is the number of S atoms coordinated to the three Mo atoms at the vacancy, and N_{Mo} is the number of Mo atoms which the S atoms in N_S are coordinated to. Ample details, equations and numerical examples are provided in section S1.† In the following, we will illustrate the relationships between common descriptors used to capture adsorption-energy trends on MoS₂(001). Although the descriptor-based analyses aim to predict catalytic activities using the properties of clean surfaces only, previous works have pointed out that adsorbate features often need to be

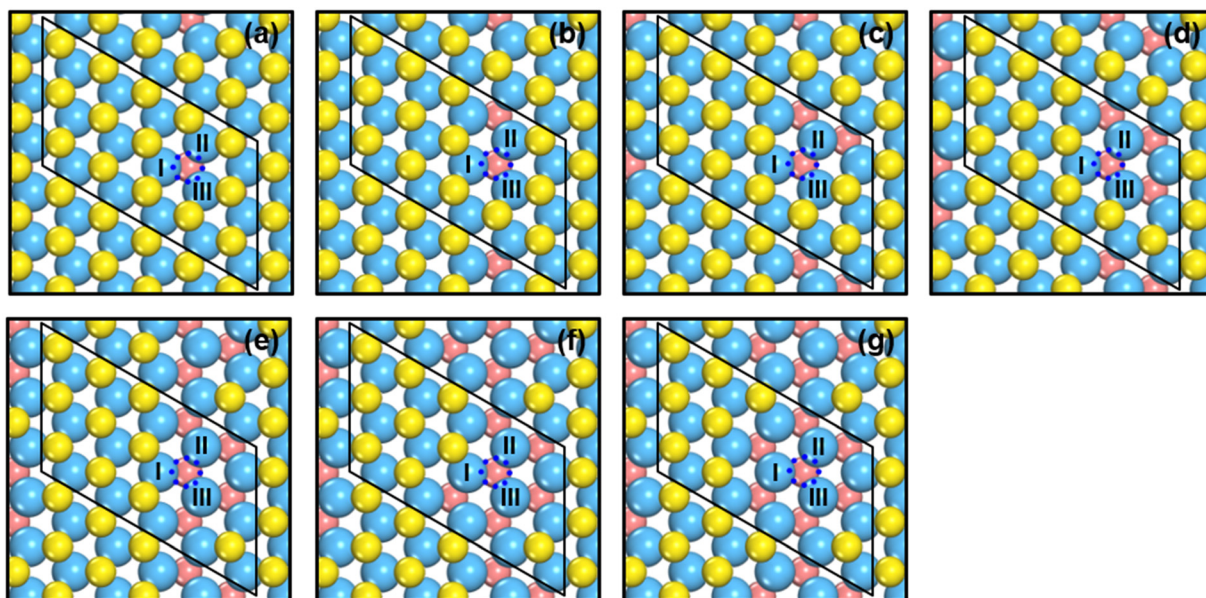


Fig. 1 Top view of single-layer MoS₂(001) at (a) 1/15 ML (*cn* = 5.00, ΔE_{vac} = 5.98 eV); (b) 2/15 ML (*cn* = 4.67, ΔE_{vac} = 6.00 eV); (c) 1/5 ML (*cn* = 4.33, ΔE_{vac} = 6.25 eV); (d) 4/15 ML (*cn* = 4.00, ΔE_{vac} = 6.52 eV); (e) 1/3 ML (*cn* = 3.67, ΔE_{vac} = 6.54 eV); (f) 2/5 ML (*cn* = 3.33, ΔE_{vac} = 6.67 eV); (g) 7/15 ML (*cn* = 3.00, ΔE_{vac} = 7.02 eV) S vacancy coverages (θ). Blue, yellow and vermilion balls represent Mo, upper S and lower S atoms, respectively. The *x* and *y* coordinates of upper and lower S atoms coincide. If vermilion balls are visible in the top view, S atoms of the upper layer were removed. The blue dashed circle marks the adsorption site under study. I, II, III denote the Mo atoms at the adsorption sites. More details on the calculation of *cn* are in section S1 in the ESI.†



incorporated into the descriptors to capture covalent adsorbate–surface interactions.⁸ Thus, in addition to cn , work function (ϕ), ΔE_{vac} and ε_d , which are calculated on clean surfaces, we consider as descriptors the excess charge on the adsorbates, and the integrated coop and coh, which are calculated on the surfaces with adsorbates.

As shown in Fig. 1 and 2 and Table S1,[†] cn progressively decreases as the S vacancy coverage goes from 1/15 to 7/15 ML, whereas the reverse trend is observed for S vacancy formation energy (ΔE_{vac}). This is because with decreasing cn , more electrons are available on Mo atoms, which lifts up the Fermi level, leading to smaller ϕ . In turn, a smaller ϕ implies more facile electron donation to S atoms (stronger S adsorption) at the vacancy. Besides, creating a S vacancy is the reverse process of S adsorption at such a vacancy, so: $\Delta E_{\text{Ads}}(\text{S}) = -\Delta E_{\text{vac}}$ (see eqn (2) and (3) in the Methods section). Therefore, the linear relations between cn , ϕ and ΔE_{vac} are expected to have a negative slope, as shown in Fig. 2a and Table S2.[†] We note that Rahman *et al.*³⁴ used an (8×8) MoS_2 supercell to show that the first ΔE_{vac} is 5.85 eV, which is comparable to our first ΔE_{vac} (5.98 eV). The slight discrepancy probably stems from the different plane-wave cutoffs (500 vs. 400 eV) and vacancy coverages (1/64 vs. 1/15 ML). In addition, we found a good correlation between ε_d and ΔE_{vac} (Fig. S1 and Table S2[†]).

Upon atomic adsorption (see structures in Fig. 3), the charge transferred to the adsorbates gradually increases with increasing ΔE_{vac} (Fig. 2b), which is a result of the

smaller ϕ of the $\text{MoS}_2(001)$ surface at higher vacancy content. The slope becomes more negative as a function of the adsorbate valency, in agreement with bond order conservation theory.³⁵ Indeed, the slopes get increasingly steeper as species with single (F, H), double (O) and triple or higher (B, N and C) valency are considered. Compared to the cn vs. ΔE_{vac} and ϕ vs. ΔE_{vac} relations, the one between excess charge and ΔE_{vac} has less negative slopes (see Table S2[†]). This can be attributed to the smaller range spanned by the charge transferred ($0.86 e^-$) on $\text{MoS}_2(001)$ compared to the corresponding value ($2.24 e^-$) on the close-packed surfaces of 4d and 5d transition metals in previous works.⁸ These results reflect a more modest variation in electronic structure of $\text{MoS}_2(001)$ as vacancies are introduced, such that the materials properties (*i.e.*, descriptors) need larger variations to reach a given effect compared to pure transition metal surfaces. This is confirmed by the trends in integrated coop for atomic adsorption, which hardly change as a function of ΔE_{vac} in Fig. 2c. Similarly, the trends in integrated coh are nearly flat as a function of ΔE_{vac} in Fig. 2d. In addition, ΔE_{vac} and cn are linearly related to the cohesive energy of the materials (E_{coh}), as shown in Fig. S2.[†] The anticorrelation between ΔE_{vac} and E_{coh} is physically meaningful, as it reflects the fact that as MoS_2 slabs contain less bonds, breaking Mo–S bonds is increasingly difficult. In turn, the correlation between cn and E_{coh} reflects the fact that as Mo atoms have more S neighbors, there is more energy stored in the solid.

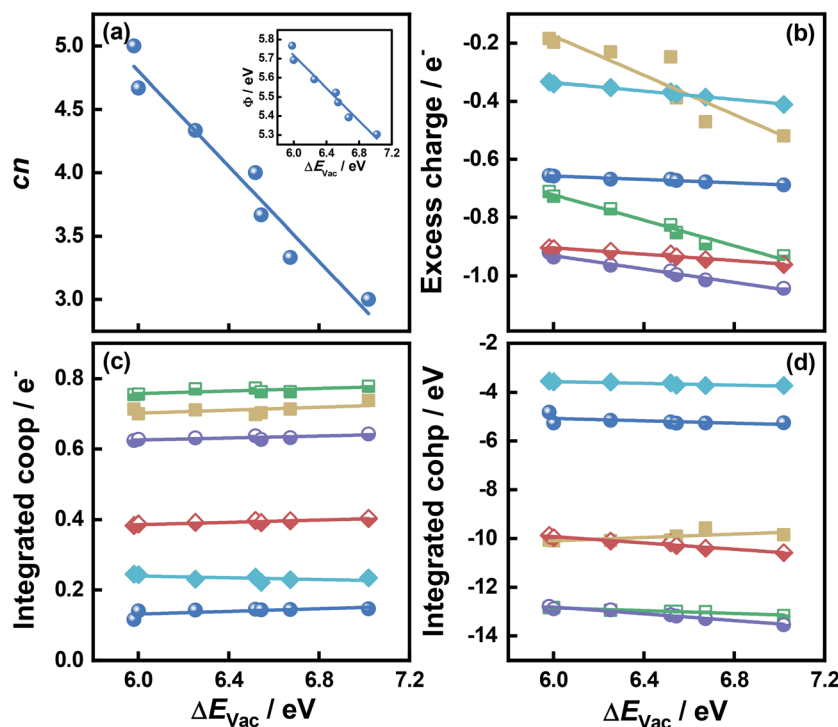


Fig. 2 Relationships between the formation energy of S vacancies (ΔE_{vac}) on $\text{MoS}_2(001)$ and: (a) the number of S atoms coordinated to Mo atoms (cn). Inset: Correlation between ΔE_{vac} and work function (ϕ). (b) The excess charge on the adsorbates. (c) Integrated crystal orbital overlap population (coop). (d) Integrated crystal orbital Hamilton population (coh) upon atom adsorption. The data in this figure are in Tables S1 and S2[†] (section S9).



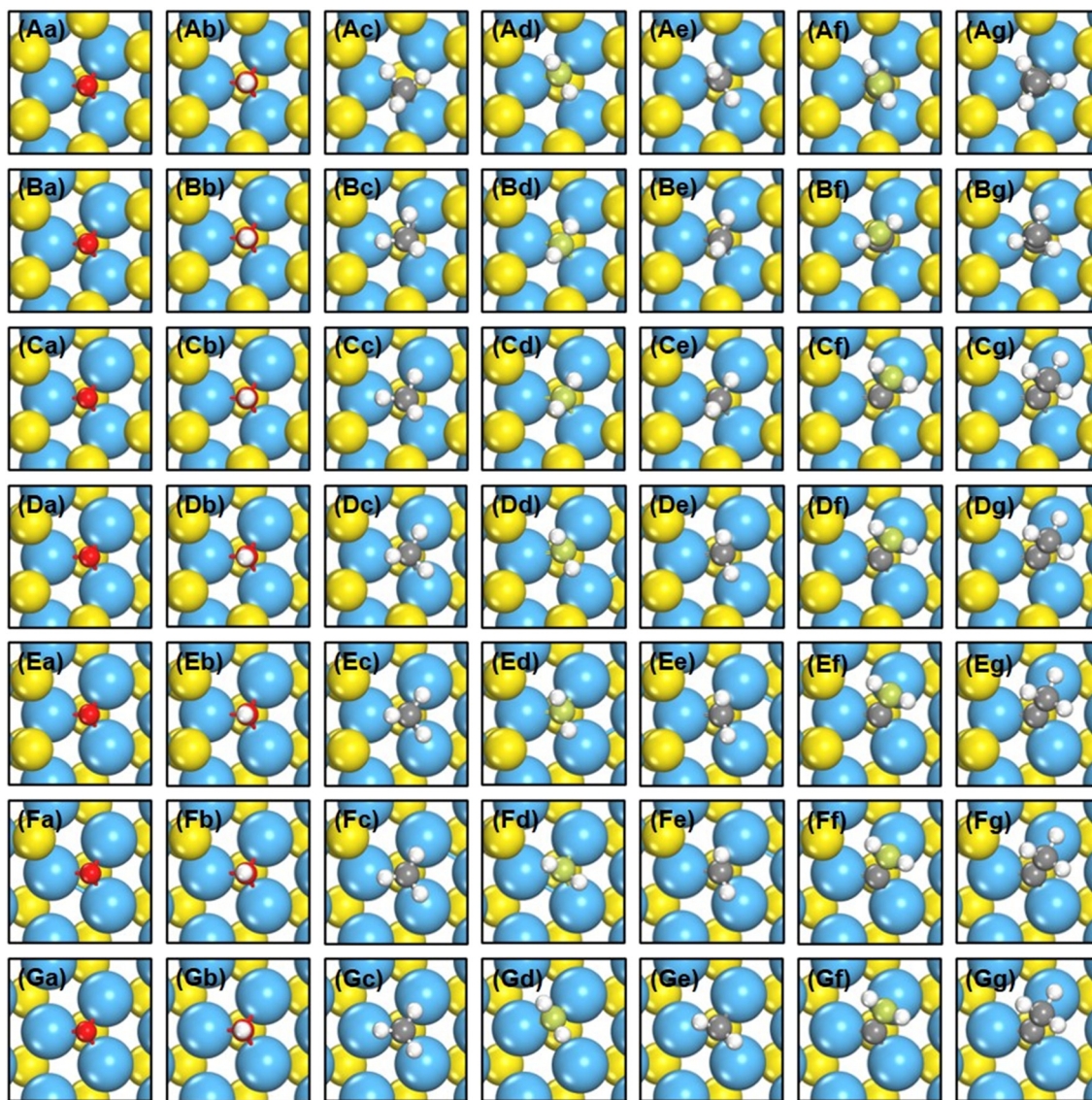


Fig. 3 Top views of the optimized structures of (a) *O, (b) *OH, (c) *CH₃, (d) *NH₂, (e) *CH₂, (f) *CNH₂ and (g) *CCH₃ adsorption on MoS₂(001) surface with S vacancy coverages of (A) 1/15, (B) 2/15, (C) 1/5, (D) 4/15, (E) 1/3, (F) 2/5 and (G) 7/15 ML. Blue, yellow, gray, white, red and green balls represent Mo, S, C, H, O and N atoms, respectively. Because of the similarity in the adsorption structure of monatomic and diatomic species, only the adsorption structures of *O and *OH are shown.

So far, the relationships between a variety of descriptors and ΔE_{vac} on MoS₂(001) were presented. Because ΔE_{vac} is in this case equivalent to the additive inverse of $\Delta E_{\text{ads}}(\text{S})$, it is reasonable to expect that ΔE_{vac} should be able to describe the trends in ΔE_{ads} of other adsorbates.^{5,8,36} We probed the adsorption of H, B, C, N, O and F on MoS₂(001) at various S vacancy contents, and the calculated ΔE_{ads} was correlated with ΔE_{vac} in Fig. 4a (see Tables S3–S9†). We note that scaling relations also exist between ΔE_{ads} of these atoms, as seen from the correlation between ΔE_{N} and ΔE_{ads} of other atomic species in Fig. S3†.

In Fig. 4a, we observe that all atomic species bind more strongly as ΔE_{vac} is increased, which indicates that they all bind to MoS₂ similarly, despite their different chemical natures. Since ΔE_{vac} scales with cn , ϕ , ϵ_{d} and excess charge

(Fig. 2 and S1†), these descriptors capture the adsorption-energy trends similarly, as confirmed in Fig. 4b–d and S1†. However, the integrated coop and cohp, which are almost independent on ΔE_{vac} , scale with ΔE_{ads} of the atoms with steep slopes, see Fig. 4e and f. Hence, in this case a high accuracy of these two descriptors would be required in order to suitably predict ΔE_{ads} . Averaging the mean absolute errors (MAEs) for the correlations between ΔE_{ads} and ΔE_{vac} , cn , ϕ , ϵ_{d} , excess charge, integrated coop and cohp we obtain 0.06, 0.07, 0.07, 0.07, 0.07, 0.16 and 0.10 eV, respectively (see Table S10†). Hence, cn is as predictive as ΔE_{vac} , ϕ , ϵ_{d} and excess charge. However, cn is arithmetically calculated, unlike the other parameters. This simple fact helps save computational time, as the assessment of ΔE_{vac} , ϕ , ϵ_{d} and excess charge requires electronic-structure calculations.



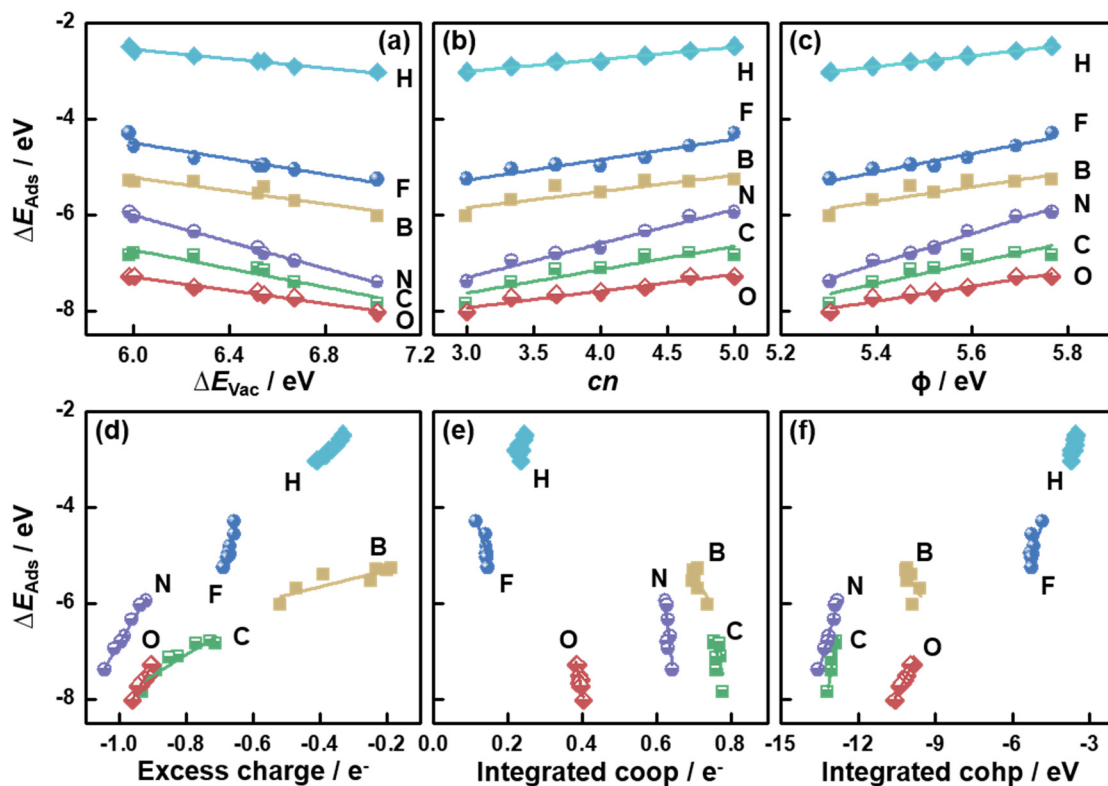


Fig. 4 Adsorption energies (ΔE_{Ads}) of atomic species on $\text{MoS}_2(001)$ as a function of (a) the formation energies of S vacancies (ΔE_{Vac}); (b) the number of S atoms coordinated to Mo atoms (cn); (c) work function (ϕ); (d) the excess charge on the adsorbates; (e) integrated crystal orbital overlap population (coop); and (f) crystal orbital Hamilton population (cohφ). The data in this figure can be found in Tables S3–S9.†

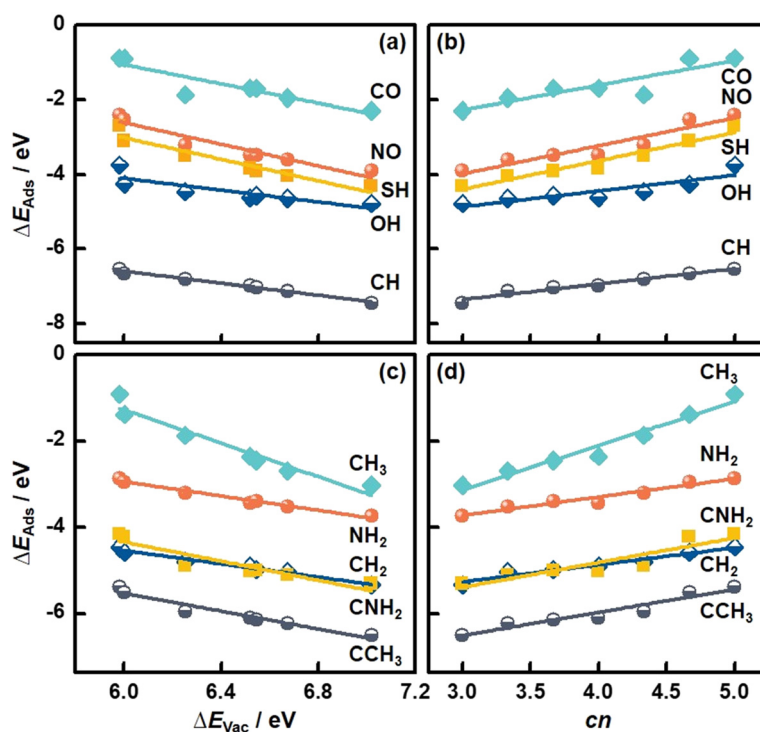


Fig. 5 Adsorption energies (ΔE_{Ads}) of molecules on $\text{MoS}_2(001)$ as a function of (a and c) the formation energies of S vacancies (ΔE_{Vac}), and (b and d) the number of S atoms coordinated to Mo atoms (cn). See the data in this figure in Tables S11–S14.†



To understand the nature of the bonds between the atomic species and MoS₂(001), we performed the differential charge density analysis in Fig. S4.† The maps show considerable charge withdrawal from the MoS₂(001) surfaces by F and O, indicating the creation of mostly ionic bonds. In contrast, considerable charge builds up between the MoS₂(001) surfaces and B and C, implying that their bonds are largely covalent. With increasing S vacancy content, less charge withdrawal from B and C is found, which corresponds to the larger excess charge seen in the Bader analysis (Fig. 4d). For N, both ionic and covalent bonding contribute significantly. Despite the different bonding nature, the good correlation between ΔE_{Ads} and the charge transferred to the atoms (Fig. 4d) indicates that the ionicity of the bonds is key for the activity trends on MoS₂(001) with vacancies.

We also studied the adsorption of diatomic (CO, NO, CH, OH, SH) and multiatomic species (CH₂, CH₃, NH₂, CNH₂, CCH₃) on MoS₂(001). The structural and energetic data appear in Fig. 3 and Table S3.† As shown in Fig. 5a, ΔE_{Ads} of diatomic species becomes more negative with increasing ΔE_{Vac} (see Table S11†), in line with the monatomic adsorption in Fig. 4. Averaging the MAEs for the diatomic adsorption we obtain 0.11 eV (see Table S11†), which is only slightly larger than the value for monatomic adsorption.

In turn, *cn* provides trends with positive slopes in Fig. 5b. This means that ΔE_{Ads} of diatomic species becomes less negative as *cn* is increased. This trend is also found for monatomic adsorption and stems from the tendency of less coordinated surface sites to bind more strongly as a means of compensating their undercoordination. It is worth noting that *cn* is correlated with ΔE_{Ads} of the diatomic species with an average MAE of 0.11 eV (see Table S12†), which is as low as that of the correlations based on ΔE_{Vac} . In addition, good correlations between ΔE_{Ads} of multiatomic adsorbates and ΔE_{Vac} and *cn* are observed, see Fig. 5c and d. Averaging the MAEs for the two descriptors we obtain 0.09 eV in both cases, which are slightly smaller than the corresponding values for diatomic adsorption. Further details can be found in Tables S13 and S14.† We note that various S vacancy arrangements with *cn* in the range of 3.67 to 4.67 and their relation to the adsorption of species were studied before.³⁷ The results suggest that the correlation between *cn* and ΔE_{Ads} also holds for other S vacancy arrangements, such that *cn* might be a general adsorption-energy descriptor on MoS₂(001).

Moreover, we investigated the ability of *cn* to describe trends in elementary reactions. As shown in a previous work, ΔE_{Ads} of atomic species is closely related to the reaction energy (ΔH) of an elementary reaction step.³⁷ Indeed, for a surface recombination reaction $*A + *B \rightarrow C + 2*$, ΔH is given by eqn (1):

$$\Delta H = \Delta H_{\text{g}} - \Delta E_{\text{Ads}}(\text{A}) - \Delta E_{\text{Ads}}(\text{B}) \quad (1)$$

where $\Delta H_{\text{g}} = E_{\text{C}} - E_{\text{A}} - E_{\text{B}}$ is the corresponding gas-phase reaction energy, and E_{A} , E_{B} and E_{C} are the total energies of free A, B and C. Since ΔH_{g} is constant for a given choice of A,

B and C, ΔH is primarily determined by the adsorption energies of the reactants $*A$ and $*B$. Because ΔE_{Ads} is shown here to depend on *cn*, it is likely that ΔH be linearly correlated with *cn*. For instance, for the reactions $*\text{CH}_3 + *H \rightarrow \text{CH}_4 + 2*$, $*\text{OH} + *H \rightarrow *H_2\text{O} + *$, and $*\text{CH}_3\text{O} + *H \rightarrow *CH_3\text{OH} + *$, the adsorption of reactants ($*\text{CH}_3$, $*\text{OH}$, $*\text{CH}_3\text{O}$ and $*H$) depends on *cn*, while the products are stable molecules which either do not adsorb (CH₄) or adsorb weak enough not to be structure sensitive (H₂O and CH₃OH).³⁷

Fig. 6a and Tables S15 and S16† confirm that ΔH of elementary reactions is linearly related to *cn*. In addition, although the adsorption of reactants and products depends on *cn* in a similar way, ΔH is well described by *cn* for $*O + *H \rightarrow *OH + *$, and $*NO + * \rightarrow *N + *O$. We note that $*NO$ dissociation has a positive slope, unlike the recombination reactions in which the slopes are negative. This positive slope appears because ΔH for dissociations is mostly determined by the products instead of reactants, as opposed to recombination reactions.

Since saddle points are often hard to find, assessing activation energies is an arduous task. Brønsted–Evans–Polanyi (BEP) relations are extremely useful in catalysis, as they linearly

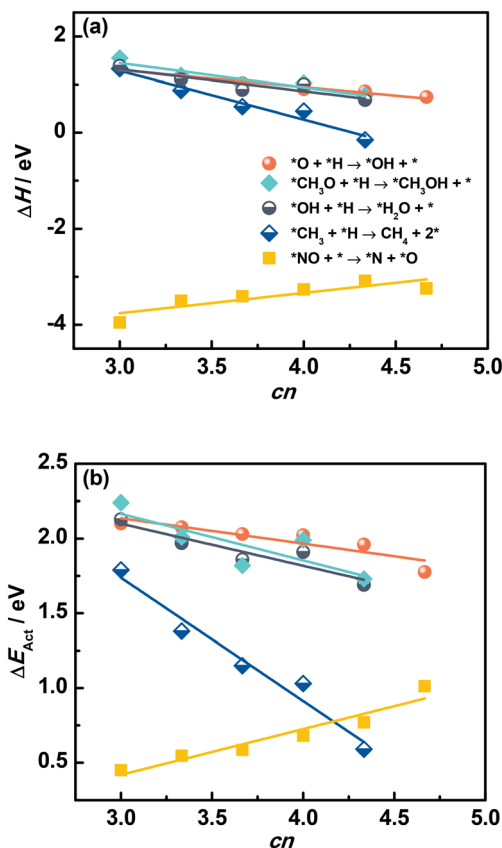


Fig. 6 Reaction energies ΔH (a) and activation energies ΔE_{Act} (b) of various elementary reactions as a function of the number of S atoms coordinated to Mo atoms (*cn*) on MoS₂(001). The slopes for $*NO$ dissociation are positive because ΔH and ΔE_{Act} are mostly determined by the products instead of reactants, as opposed to the recombination reactions. The data in this figure appear in Tables S15 and S16.†



connect the activation energy ΔE_{Act} of an elementary reaction to its reaction energy ΔH .^{38–41} With this in mind, we examined the trends in ΔE_{Act} of elementary reactions as a function of *cn* in Fig. 6b (see also Fig. S5†). Since the formation of OH, H₂O and CH₄ has been shown to largely determine the rate of CO and CO₂ hydrogenation reactions on MoS₂ catalysts,^{37,42,43} the correlation between ΔE_{Act} and *cn* is encouraging, as it indicates that *cn* captures not just the trends in reaction thermodynamics but also the reaction kinetics. Similar to ΔH , the trend of ΔE_{Act} of dissociation reactions is different from that of recombination reactions.

We also tested the applicability of *cn* on MoS₂ edges, as they display high catalytic activity for various reactions. We employed Mo edges with S vacancies to study *O and *OH adsorption, see Table S17 and Fig. S6†. As the S vacancy concentration varies from 100% to 62.5%, *cn* increases from 2.67 to 3.67 and so do ΔE_{Ads} of *O and *OH. Hence, *cn* seems able to describe adsorption-energy trends on edge sites. However, MoS₂ edges, particularly the S edge, are prone to reconstruction. Indeed, recent works found that Mo atoms on the S edge with 50% S vacancies have S–Mo–S angles of $\sim 100^\circ$ and 130° , unlike those on the Mo edge with 62.5% S vacancies, which keep a trigonal prismatic coordination similar to MoS₂(001) with S–Mo–S angles of $\sim 80^\circ$ and 130° .⁴² Furthermore, Mo edges display S–Mo bond lengths of 2.44, 2.37 and 2.24 Å, while the S–Mo bond lengths on S edges are ~ 2.30 Å. The large variations in the structure of Mo atoms on S edges lead to a dissimilar hybridization compared to Mo edges. In such case, *cn* fails to describe the trends in ΔE_{Ads} : Mo atoms at the S vacancy of the S edge with 50% S vacancies have *cn* = 3.33, which is smaller than the value of 3.67 on the Mo edge with 62.5% S vacancies. However, the species at the S vacancy on the S edge tend to bind more weakly than those on the Mo edge (see Table S18†).⁴² In sum, *cn* holds promise as a descriptor for adsorption-energy trends including edges with Mo atoms in similar geometric configurations compared to terrace sites, but if the edge sites undergo important geometric reconstructions, *cn* is likely to fail.

3. Discussion

Coordination numbers are well-known descriptors for adsorption-energy trends on metals and some oxides. For instance, conventional and “generalized” coordination numbers capture the trends in adsorption energy on extended surfaces and nanoparticles of Pt, Au, Cu and Zn.^{6,44–50} In fact, there is an analytic relationship between coordination numbers and ϵ_{d} ,⁶ and a linear proportionality between ϵ_{d} and ΔE_{Ads} by virtue of the d-band model.¹

The free energy difference between *O and *OH ($\Delta G_{\text{O}} - \Delta G_{\text{OH}}$) is probably the most widely used computational descriptor for the oxygen evolution reaction (OER) activity.^{51–53} Previous works found that the number of oxygen atoms coordinated to surface Mn ions on various sites of Mn₂O₃(110) and MnO₂(110) surfaces are linearly related to $\Delta G_{\text{O}} - \Delta G_{\text{OH}}$, such that the OER activities of Mn oxides are sensitive to oxygen coordination at the surface.⁵⁴ In addition,

Jiang *et al.* reported that “adjusted” coordination numbers exhibit a linear correlation with the *H adsorption energy and C–H activation barrier over various facets and defects on V₂O₃, Cr₂O₃, Co₃O₄, and NiO.²⁰ Viswanathan *et al.* found that ΔG_{OH} and the ORR activities of surface reconstructed Ni₃S₂ catalysts are determined by the Ni–S coordination numbers, such that counting S neighbors around Ni surface sites can be used to anticipate the ORR activity.²¹

Finally, in the following we will illustrate the use of *cn* to design active sites at MoS₂(001) for electrocatalytic hydrogen evolution in two ways. First, as shown in the inset of Fig. 7, a thermodynamic analysis based on the computational hydrogen electrode⁵⁵ predicts the most active terrace sites to be found for *cn* between 4.33 and 4.67, as ΔG_{H} goes from negative to positive in that range and it should ideally be zero, as predicted by the Sabatier principle.^{56,57} In the main panel of Fig. 7, we use the microkinetic model of Nørskov *et al.*⁵⁷ to calculate the exchange current density of various MoS₂ electrodes. An implicit assumption of this model is that all sites are equivalent, such that each datapoint in Fig. 7 represents an electrode with all active sites of that specific *cn* (see full details in section S10†). Fig. 7 shows that *cn* = 4.51 provides the highest HER activity in terms of exchange current density in the range of coordination inspected. Hence, in addition to increasing the coverage of vacancies and straining them,²⁸ an alternative to enhance the HER activity of MoS₂ is to maximize the number of terrace sites with neighboring S vacancies having *cn* in the range of 4.33–4.67. This is a simple quantitative guideline to obtain highly active sites for the HER at sites other than MoS₂ edges.

4. Conclusions

The number of sulfur atoms coordinated to Mo atoms (*cn*) is an intuitive, simple and accurate descriptor for the trends in adsorption energies, the thermodynamics and kinetics of several elementary reactions on MoS₂. Interestingly, *cn* is arithmetically assessed, which helps lower the considerable computational cost typically invested in calculating electronic-structure descriptors. Numerous examples in the literature and the present work highlight the potential of coordination-based descriptors in catalysis. Devising similar descriptors for other oxides, sulfides, nitrides, *etc.* should be possible and may bring a sense of generality to the catalysis theory of oxidized transition metal compounds.

Furthermore, taking electrocatalytic hydrogen evolution as a case study, *cn* was able to outline the geometric configuration of enhanced active sites on MoS₂(001). Specifically, MoS₂ terrace sites with *cn* = 4.51, which correspond to moderately undercoordinated sites with neighboring vacancies, should display maximal HER activities.

5. Methods

Spin-unrestricted density functional theory (DFT) calculations were performed with the Vienna *ab initio* simulation package



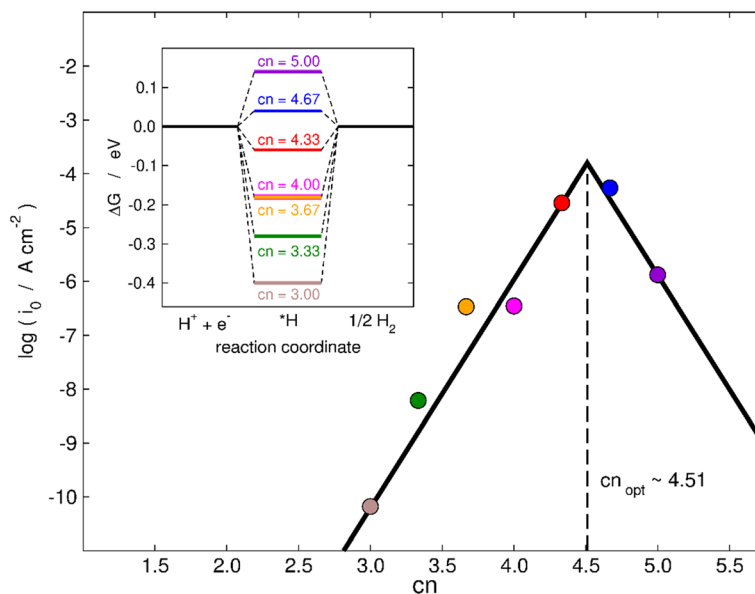


Fig. 7 Trends in exchange current density (i_0) for hydrogen evolution for various sites on $\text{MoS}_2(001)$ with S vacancies as a function of cn . The values of i_0 were calculated using a microkinetic model,⁵⁷ see full details in section S10†. The largest HER activity is found for $cn = 4.51$. Inset: Thermodynamic modelling of the HER on MoS_2 . The thermodynamically ideal active sites with $\Delta G_{\text{H}} = 0$ at 0 V vs. RHE correspond to cn in the range of 4.33–4.67.

(VASP).⁵⁸ The interaction between ionic cores and valence electrons was described by the projector-augmented wave (PAW) method,⁵⁹ and the Kohn–Sham valence electronic wavefunction was expanded using a plane-wave basis set with a kinetic energy cutoff of 400 eV. Exchange and correlation effects on the total energies were calculated within the generalized gradient approximation using the Perdew–Burke–Ernzerhof (PBE) exchange–correlation functional.⁶⁰ All slab atoms and adsorbates were fully relaxed, and the total energies were converged below 10^{-4} eV, while the forces on the atoms were converged below $0.03 \text{ eV } \text{\AA}^{-1}$.

Bulk 2H-MoS_2 is a layered material, and the lattice constants are calculated to be $a = b = 3.19 \text{ \AA}$, $c = 12.43 \text{ \AA}$, which are close to the experimental values ($a = b = 3.16 \text{ \AA}$, $c = 12.29 \text{ \AA}$).⁶¹ Single-layer MoS_2 (denoted MoS_2 hereafter) consists of an S–Mo–S sandwich, where the Mo atoms are arranged in a hexagonal lattice with a trigonal prismatic coordination with respect to the two S layers. The pristine $\text{MoS}_2(001)$ basal plane, with each Mo atom coordinated to six S atoms, was modeled by an S–Mo–S trilayer with (5×3) periodicity. To simulate S vacancies in the range of $\theta = 1/15$ – $7/15$ ML, one to seven surface S atoms were removed in a clockwise manner, as shown in Fig. 1. Some S vacancy arrangements in Fig. 1 may not be the most stable configuration for a given coverage. However, the main purpose of this work is to elucidate how adsorption energies at a given S vacancy (dashed circle, Fig. 1) respond to the coordination environment. For completeness, we compared the relative stability of adjacent and far S vacancies at $2/15$ ML vacancy coverage and found that the former (Fig. S7a†) is slightly more stable than the latter (Fig. S7b†) by 0.06 eV. The higher

stability of adjacent S vacancy compared to far S vacancy was also observed in a previous DFT study.³⁴ The surface Brillouin zone was sampled with a $(2 \times 3 \times 1)$ Monkhorst–Pack k -point grid for the $\text{MoS}_2(001)$ surface.⁶² A vacuum region of at least 15 \AA sufficed to avoid interactions between periodically repeated slabs along the z -direction for all systems. More details about the models can be found in previous works.^{37,42,43}

The adsorption energy (ΔE_{Ads}) was calculated using the adsorption configurations ($E_{\text{MoS}_x+\text{ad}}$) relative to the clean surfaces (E_{MoS_x}) and the isolated adsorbates (E_{ad}):

$$\Delta E_{\text{Ads}} = E_{\text{MoS}_x+\text{ad}} - E_{\text{MoS}_x} - E_{\text{ad}} \quad (2)$$

where x varies with the S vacancy coverage. In eqn (2), a more negative ΔE_{Ads} implies stronger binding, while a more positive ΔE_{Ads} implies weaker binding. The formation energy of the surface sulfur vacancy (ΔE_{Vac}) under study (blue dashed circle in Fig. 1) is:

$$\Delta E_{\text{Vac}} = E_{\text{MoS}_{x-1}} + E_{\text{S}} - E_{\text{MoS}_x} \quad (3)$$

where E_{S} and $E_{\text{MoS}_{x-1}}$ are the energies of atomic sulfur and the surface with one less sulfur atom compared to E_{MoS_x} , respectively. A more positive ΔE_{Vac} means that the sulfur vacancy is more difficult to form.

All transition states (TSs) were located by the force reversed method.⁶³ The relaxation proceeds until the residual forces in each atom are smaller than $0.03 \text{ eV } \text{\AA}^{-1}$. The reaction energy (ΔH) of elementary reactions is calculated as:

$$\Delta H = E_{\text{FS}} - E_{\text{IS}} \quad (4)$$



where E_{FS} and E_{IS} are the energies of the most stable configurations for the separate adsorption states of the intermediates in the final state and initial state, respectively. The elementary activation barrier (ΔE_{Act}) was calculated using the TS (E_{TS}) with respect to the most stable configurations for the separate adsorption states of species at the IS (E_{IS}):

$$\Delta E_{\text{Act}} = E_{\text{TS}} - E_{\text{IS}} \quad (5)$$

Bader charge analyses were performed using a grid-based weight method in which the expression for the fraction of space neighboring each grid point that flows to its neighbors is used as a weight for the discrete integration of functions over the Bader volume.⁶⁴ A positive or negative charge implies charge depletion or charge accumulation, respectively. The coop and cohpc were analyzed using the LOBSTER program.⁶⁵ The coop is an overlap-population-weighted density of states that provides information about electron distribution. In turn, the cohpc is described as an “energy weighted” density of states. When integrated up to the Fermi level, the coop and cohpc hint toward the strength of chemical bonds.⁸ The HER modelling is described in detail in section S10.†

Author contributions

HYS performed the DFT calculations and wrote the first draft of the paper. FCV carried out the electrocatalysis modelling. FCV and KS supervised the project. All authors analyzed the data, discussed the results and edited the manuscript.

Conflicts of interest

There are no conflicts to declare.

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