

Cite this: *Mater. Adv.*, 2023,
4, 2365

Sn²⁺ doping-induced large extra vibrational energy of an excited state for efficient blue emission in Cs₂SnCl₆:Bi[†]

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Luminescent metal halides doped with ns²-metal ions, such as 6s²-metal Bi³⁺, show excellent optoelectronic properties. However, the origin of the high photoluminescence quantum yield (PLQY) of Cs₂SnCl₆:Bi remains controversial and unclear. In this study, a series of Cs₂SnCl₆:Bi were synthesized by adopting different tin precursors of SnCl₂/SnCl₄. The samples from SnCl₂ exhibited much higher PLQY than those from SnCl₄, and the doping concentrations of Bi³⁺ had little effect on the high PLQY. When H₃PO₂ is added to the precursor of SnCl₄ to reduce Sn⁴⁺ to Sn²⁺, the prepared Cs₂SnCl₆:Bi shows bright emission light and high PLQY. The high PLQY of Cs₂SnCl₆:Bi is attributed to the co-doping of Bi³⁺ and Sn²⁺. [SnCl₄]²⁻ and two V_{Cl} induce a large extra vibrational energy of the heavy localized excitons from [BiCl₅]²⁻ to enhance the efficiency of radiative recombination for high PLQY. This work deepens the understanding of the luminescence mechanism of Cs₂SnCl₆:Bi and supplies significant references for developing more perovskite materials with outstanding luminescence.

Received 6th March 2023,
Accepted 26th April 2023

DOI: 10.1039/d3ma00107e

rsc.li/materials-advances

Introduction

Luminescent metal halides have attracted the attention of many researchers. Due to their superior optical properties, such as tunable emissions, high photoluminescence quantum yield (PLQY) and long diffraction length, Pb-based halides show promising perspective in LEDs, solar cells, photodetectors and X-ray imaging.^{1–11} However, the toxicity of the heavy metal Pb limits the development of practical applications.^{12,13} Accumulated Pb in human body can cause numerous brain-related symptoms, such as intellectual disability. Lead-free metal halides as substitutions have become a research tendency in recent years. Tin (Sn), which also has the characteristics of ns² electrons and octahedral coordination, is a potential replacement for Pb. CsSnX₃ exhibits tunable emission light through changing the ratio of halogens, which is similar to the properties of CsPbX₃.^{14–16} However, CsSnX₃ is unstable in air conditions, in which Sn²⁺ is easily oxidized to Sn⁴⁺, causing severe decomposition. Therefore, nontoxic and stable perovskite alternatives must be urgently developed.

Recently, many lead-free metal halides have been explored, such as Cs₂NaInCl₆, Cs₂AgInCl₆, Cs₃Bi₂I₉, Cs₃Cu₂I₅ and Cs₂ZrCl₆, owing to their excellent optoelectronic properties and stable structures.^{17–23} Among them, Cs₂SnCl₆ with its stable tetravalent cation is considered as a promising candidate for photoelectric applications.^{24,25} The Cs₂SnCl₆ crystal is derived from the three-dimensional CsSnCl₃ perovskite, which periodically removes half of the Sn atoms at the center of each [SnCl₆]²⁻ octahedron, forming a vacant-ordered double perovskite structure (space group *Fm* $\bar{3}$ *m*). The [SnCl₆]²⁻ octahedral structure is isolated by a Cs⁺ cation. Pure Cs₂SnCl₆ exhibits poor photoluminescence property, but it is an excellent host material. Cs₂SnCl₆ is a zero-dimensional structure, and it can localize the excitons for the luminescent centers to achieve high effective emissions. Through doping with a luminescent center, such as Bi³⁺, Sb³⁺ and Te⁴⁺, Cs₂SnCl₆ can exhibit excellent luminescence with tunable emission wavelength.^{26–29} Sb³⁺-doped Cs₂SnCl₆ has been reported to emit orange light with a photoluminescence quantum yield (PLQY) of 37%, and Bi³⁺ doping resulted in blue light emission with a PLQY of 79%.^{30–32} The Bi³⁺/Te⁴⁺ co-doped perovskite derivative Cs₂SnCl₆ shows highly efficient and dual-band-tunable white-light emission owing to the strong electron-phonon coupling and efficient energy transfer.^{33–35}

Interestingly, when Cs₂SnCl₆:Bi is synthesized by adopting SnCl₂ as a precursor of the Sn source, Cs₂SnCl₆:Bi can get a high PLQY of ~78%. However, if using SnCl₄ as a precursor to synthesize Cs₂SnCl₆:Bi, the PLQY of the prepared Cs₂SnCl₆:Bi is

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[†] Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3ma00107e>

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below 30%. The underlying reason for this phenomenon remains unclear. Motivated by this, we synthesized a series of Bi^{3+} -doped Cs_2SnCl_6 samples using SnCl_2 or SnCl_4 as the precursor under different conditions. Through comparing the photophysical properties of the Cs_2SnCl_6 samples with different Bi^{3+} doping content, we find that the samples from SnCl_2 have much higher PLQY values than those from SnCl_4 , and the doping content of Bi^{3+} has little effect on the high PLQY. When some reductive H_3PO_2 is added in the precursor solution with SnCl_4 , the prepared $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ shows brighter emission light and higher PLQY than the sample without H_3PO_2 . H_3PO_2 provides a reducing environment and reduces Sn^{4+} to Sn^{2+} during the reaction, and XPS spectra proved the existence of Sn^{2+} in $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ synthesized with H_3PO_2 . $[\text{SnCl}_4]^{2-}$ and two V_{Cl} induce the local lattice distortion, which can cause heavy localized excitons of $[\text{BiCl}_5]^{2-}$. The localized excitons make for a large extra vibrational energy to reduce the non-radiative transition and enhance the efficiency of radiative recombination. This work deepens the understanding of the luminescence mechanism of $\text{Cs}_2\text{SnCl}_6\text{:Bi}$, but also introduces novel ideas and approaches to develop more perovskite materials with outstanding luminescent properties.

Results and discussion

As shown in the experimental processes in Fig. 1a, SnCl_2 or SnCl_4 , CsCl and BiCl_3 were added to the autoclave containing hydrochloric acid. The autoclave was heated in the oven at 180°C for 10 hours to obtain $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ single crystals. As shown in Fig. 1b, the single crystals show regular hexagonal shape, and the size is about several to tens of micrometers,

indicating good crystallinity. EDS mapping images show that Cs, Sn, Cl, and Bi are evenly distributed in Fig. 1d, meaning Bi was successfully doped into the lattice of Cs_2SnCl_6 . Through changing the doping content of Bi and the Sn source from SnCl_2 or SnCl_4 , a series of Cs_2SnCl_6 samples were prepared. The actual doping amount of Bi was tested by inductively coupled plasma optical emission spectrometry (ICP-OES), and the element contents were given in Table 1. The samples are named as $\text{Sn}^{x+}\text{:yBi}$ (Sn^{x+} represents that the sample is synthesized by SnCl_2 or SnCl_4 in the raw materials; y represents the actual doping concentration of Bi^{3+} replacing the site of Sn^{4+}). As shown in Fig. 1c, all of the samples are confirmed to be pure Cs_2SnCl_6 with space group (PDF #70-2413). The position of the diffraction peak gradually shifted to a low angle as the concentration of doped Bi^{3+} increased, which indicates that Bi^{3+} was doped into the lattice of the Cs_2SnCl_6 crystal. Rietveld refinement of the samples of $\text{Sn}^{2+}\text{:}0.08\% \text{Bi}$ and $\text{Sn}^{4+}\text{:}0.04\% \text{Bi}$ were performed to acquire information on the crystal structure, as shown in Fig. S1 (ESI †). $\text{Sn}^{2+}\text{:}0.08\% \text{Bi}$ and $\text{Sn}^{4+}\text{:}0.04\% \text{Bi}$ are in the $Fm\bar{3}m$ space group. The lattice parameters of $\text{Sn}^{2+}\text{:}0.08\% \text{Bi}$ are $a = b = c = 10.388 \text{ \AA}$, $\alpha = \beta = \gamma = 90^\circ$, as the lattice parameters of $\text{Sn}^{2+}\text{:}0.08\% \text{Bi}$ are $a = b = c = 10.386 \text{ \AA}$, $\alpha = \beta = \gamma = 90^\circ$. Because the radius of the Bi^{3+} ion (1.03 \AA for coordination number (CN) = 6) is larger than that of the Sn^{4+} ion (0.69 \AA for CN = 6), $\text{Sn}^{2+}\text{:}0.08\% \text{Bi}$ has larger lattice parameters than $\text{Sn}^{4+}\text{:}0.04\% \text{Bi}$. Therefore, these results further verified that the single phase was successfully obtained and the crystal structure was unchanged with the introduction of Bi^{3+} ions. Although the feeding amount of BiCl_3 was from 1% to 8%, the actual doping contents of Bi^{3+} for all samples are at a very low level, less than 2%. In Table 1, all

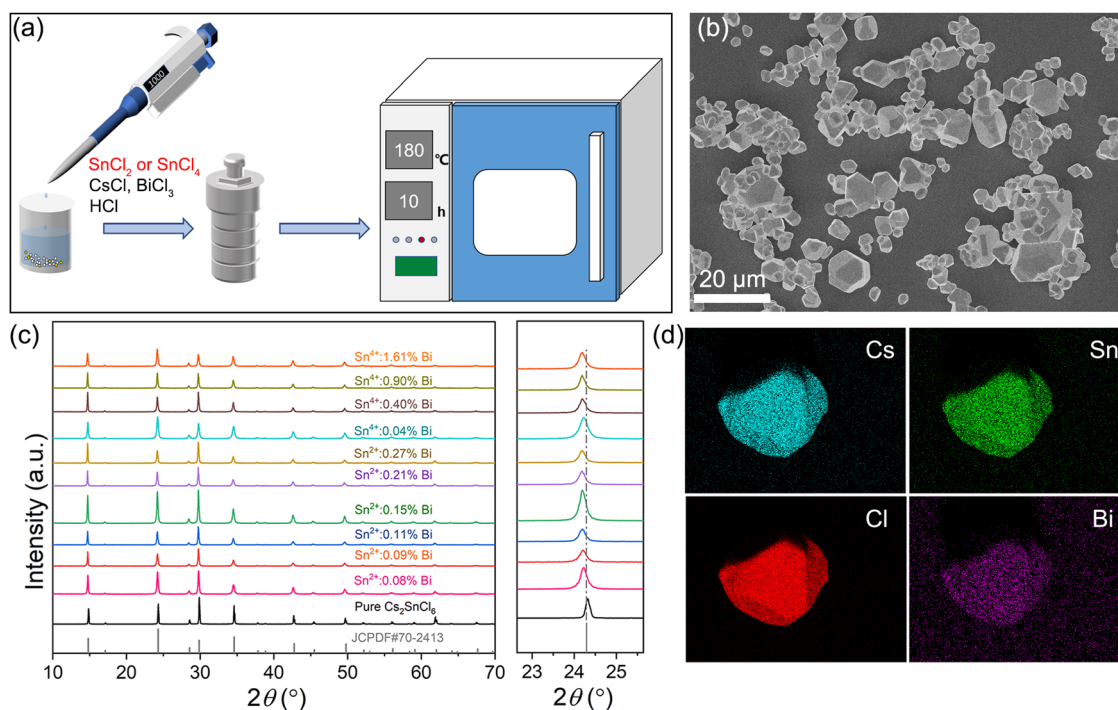


Fig. 1 (a) Schematic diagram of the synthesis process. (b) SEM image of the $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ particles. (c) XRD patterns of $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ with varied doping content and different tin precursors. (d) EDS mapping for showing the distribution of individual elements Cs, Sn, Cl and Bi.



Table 1 Doping content of Bi and PLQY for the samples synthesized by the precursor of SnCl₂ or SnCl₄

Sn precursor	SnCl ₄				SnCl ₂					
	Feeding Bi (%)	Incorporated Bi (%)	PLQY (%)		Feeding Bi (%)	Incorporated Bi (%)	PLQY (%)			
Feeding Bi (%)	0.99	2.91	4.76	7.41	0.99	2.91	4.76	7.41	9.09	13.04
Incorporated Bi (%)	0.04	0.41	0.90	1.61	0.08	0.09	0.11	0.15	0.20	0.27
PLQY (%)	31.8	28.39	6.46	7.9	57.29	53.79	52.39	51.44	51.78	52.98

samples synthesized by SnCl₂ exhibit a higher PLQY than those by SnCl₄, as shown in Fig. 2a (the values of PLQY were calculated according to the PL spectra in Fig. S2, ESI[†]). PLQY of the samples synthesized by the SnCl₂ precursor solution were above 50%, and did not change with the increase of the Bi³⁺ doping contents. Conversely, the PLQY of the samples synthesized by the SnCl₄ precursor was varied with Bi³⁺ doping concentration, and could not reach 50%. Thus, the kind of tin halides precursor was the major factor for the high PLQY, and the Bi³⁺ doping content was a minor factor and had little effect on the high PLQY. In Fig. 2b and Fig. S3 (ESI[†]), all of the samples of Cs₂SnCl₆:Bi have a broad emission peak at 452 nm, and the optimal excitation wavelength of all samples is 349 nm. Moreover, when the mixture precursors of SnCl₂ and SnCl₄ were used, the PLQY of the prepared samples was enhanced from 12.53 to 49.38% with the increase of the SnCl₂ proportion in the precursor, as shown in Fig. S4 (ESI[†]). In order to research the reason for the different PLQY, we chose two samples of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi, which have almost the same amount of

doping Bi, for photophysical performance testing and analysis. In Fig. 2b, comparing the PL spectra of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi, the shape and center position of the emission peak are the same, whereas the intensity is different. The inset photograph shows that Sn²⁺:0.08%Bi emits brighter blue light under UV light (365 nm) than Sn⁴⁺:0.04%Bi. The PL decay curves of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi are well fitted by a single exponential, and the lifetime values of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi are close at 622.6 and 700.6 ns respectively (Fig. 2c). The PL spectra and lifetime values indicated that the blue emission comes from the same radiation transition. In Fig. 2d, comparing the absorption spectra of pure Cs₂SnCl₆, Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi, the absorption edge at 315 nm is from the host Cs₂SnCl₆, but Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi have a small absorption peak at 340 nm, which is attributed to the absorption of doping Bi ions. Thus, according to the above results, it can be concluded that the improvement of the radiative transition efficiency is the main reason for the improvement of PLQY in Sn²⁺:0.08%Bi.³⁰ We speculate that (1) using SnCl₂ as the

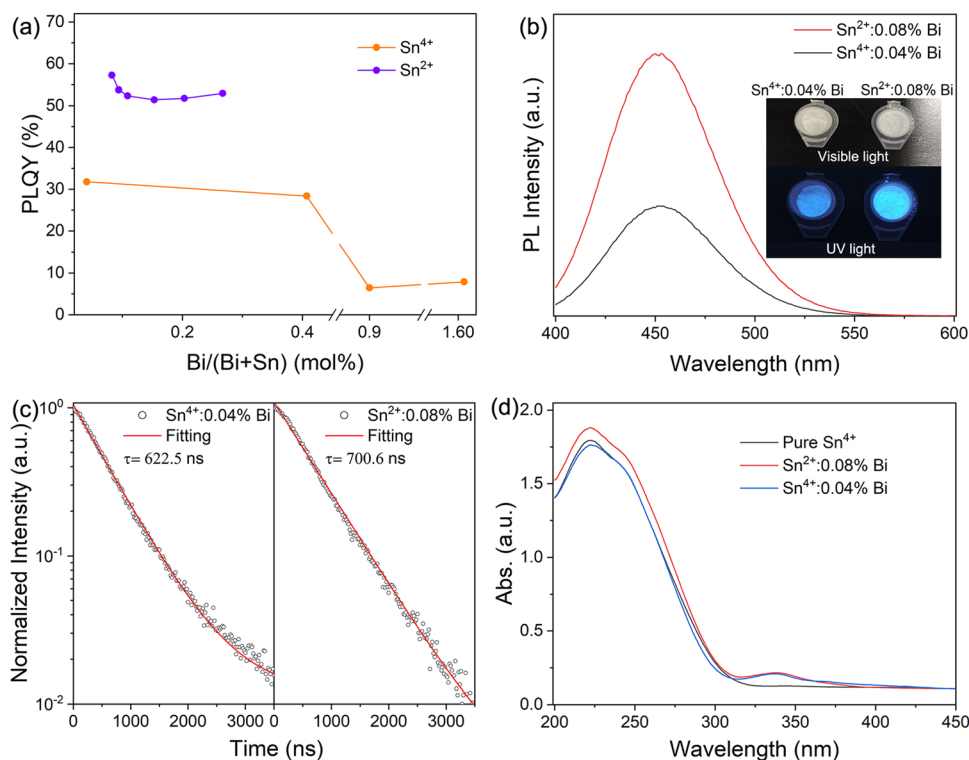


Fig. 2 (a) PLQY for the samples with different doping contents of Bi³⁺ using SnCl₂ or SnCl₄ as the precursor. (b) PL intensities of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi, which have almost the same amount of doping Bi³⁺; the inset of (b) is the photograph of these two samples under visible light and UV light. (c) Time-resolved PL and respective fitting curves of Sn²⁺:0.08%Bi and Sn⁴⁺:0.04%Bi. (d) UV-Vis absorption spectra of Sn²⁺:0.08%Bi, Sn⁴⁺:0.04%Bi and Pure Sn⁴⁺.



raw material to synthesize Cs_2SnCl_6 may introduce some Sn^{2+} into the lattice because both Sn^{2+} and Sn^{4+} ions have octahedral coordination; (2) Sn^{2+} does not transfer energy to Bi, but only improves the radiative transition efficiency of Bi^{3+} due to the lattice distortion.

To further confirm the effect of Sn^{2+} in the precursor, we added a little H_3PO_2 solution into the precursor solution with SnCl_4 to reduce some Sn^{4+} to Sn^{2+} . The prepared $\text{Cs}_2\text{SnCl}_6\text{:Bi}$ samples with/without H_3PO_2 were named as $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$. As shown in Fig. 3a, the addition of H_3PO_2 does not change the XRD patterns. $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$ are the pure Cs_2SnCl_6 phase without impure phases. In Fig. 3b, the normalized PLE spectra can well coincide, meaning that H_3PO_2 did not change the transition level. In Fig. 3c, the PL intensity of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ is higher than that of $\text{Sn}^{4+}\text{:Bi}$. In the inset image, the $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ powder shows brighter blue light under UV light. Meanwhile, the lifetime values are 730.4 and 744 ns for $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$, respectively. The PLQY of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ is up to 35.4%, which is more than twice that of $\text{Sn}^{4+}\text{:Bi}$ (13.7%). These results of PL, PLE and lifetime are consistent with those of $\text{Sn}^{4+}\text{:yBi}$ and $\text{Sn}^{2+}\text{:yBi}$. H_3PO_2 is supposed to reduce some Sn^{4+} to Sn^{2+} in the precursor solution, and Sn^{2+} is then doped into the Sn^{4+} lattice of Cs_2SnCl_6 together with Bi during the process of hydrothermal synthesis, resulting in a large increase of PLQY. Therefore, the $\text{Sn}^{2+}\text{:yBi}$ samples using SnCl_2 as the precursor are Sn^{2+} and Bi co-doped Cs_2SnCl_6 .

To verify the presence of divalent tin ions in the prepared $\text{Cs}_2\text{SnCl}_6\text{:Bi}$, we performed XPS tests. Since Sn^{2+} on the surface

of particles is easily oxidized to the tetravalent form in under air conditions, to ensure the reliability of the XPS data, we cut the single crystal in an inert atmosphere and tested the fresh cross-section. In the inert atmosphere, Sn^{2+} on the fresh cross-section is not oxidized and can be accurately detected. The full-scan XPS spectra in Fig. S5 (ESI[†]) show that all elements were detected at the cross-section. In Fig. 4a and b, the Sn 3d spectra of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ shows the obvious peaks of Sn^{2+} at binding energies of 493.65 and 485.25 eV, while that of $\text{Sn}^{4+}\text{:Bi}$ do not exhibit any peaks for Sn^{2+} . This result proves that the samples prepared by using SnCl_2 or SnCl_4 and H_3PO_2 is Bi and Sn^{2+} co-doped Cs_2SnCl_6 , and the improved PLQY may come from the distortion of the local structure caused by the divalent tin. The crystal diagram of the co-doped structure is shown in Fig. 5a. The Cs_2SnCl_6 crystal is a vacant-ordered double perovskite (space group $Fm\bar{3}m$). $[\text{SnCl}_6]^{2-}$ octahedral is isolated by the Cs^+ cation. When SnCl_2 is used as a precursor and co-doped with Bi, Sn^{2+} partially replaces Sn^{4+} to form a 4-coordinated octahedral structure $[\text{SnCl}_4]^{2-}$ with two Cl vacancies; Bi replaces Sn^{4+} to form a 5-coordinated octahedral structure with a Cl vacancy. Bi belongs to the isolated luminescence center, which serves to localize the excitons and enhance radiative recombination.³⁰ The spatial distortion of the $[\text{BiCl}_5]^{2-}$ polyhedron will cause the change of luminescence properties, while the spatial distortion is affected by its surrounding environment. The incorporated Sn^{2+} in the state of $[\text{SnCl}_4]^{2-}$ and two V_{Cl} can provide a large space to cause the local lattice distortion. When some $[\text{SnCl}_4]^{2-}$ are adjacent with $[\text{BiCl}_5]^{2-}$, the heavy distortion of $[\text{BiCl}_5]^{2-}$ occurs. This distortion

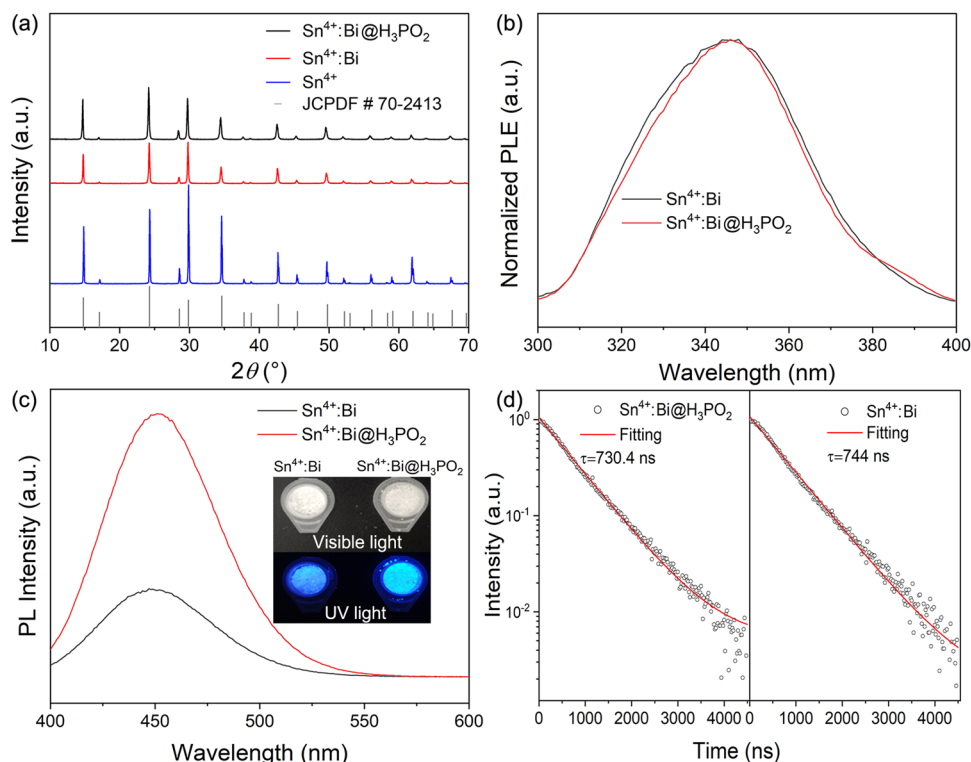


Fig. 3 (a) XRD patterns of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$, $\text{Sn}^{4+}\text{:Bi}$ and pure Sn^{4+} . (b) PLE, (c) PL spectra of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$; the inset of (c) is the photograph of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$ powders under visible light and UV light. (d) Time-resolved PL and respective fitting curves of $\text{Sn}^{4+}\text{:Bi@H}_3\text{PO}_2$ and $\text{Sn}^{4+}\text{:Bi}$.



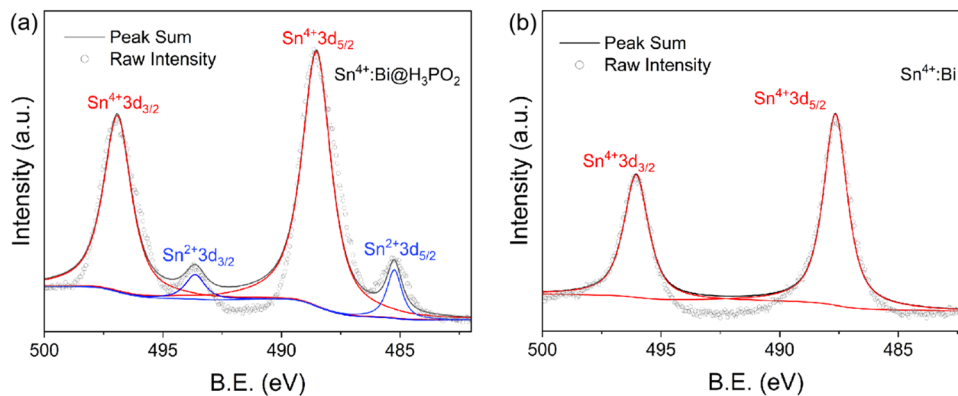


Fig. 4 High-resolution XPS spectra of Sn 3d for (a) $\text{Sn}^{4+}:\text{Bi}@\text{H}_3\text{PO}_2$ and (b) $\text{Sn}^{4+}:\text{Bi}$.

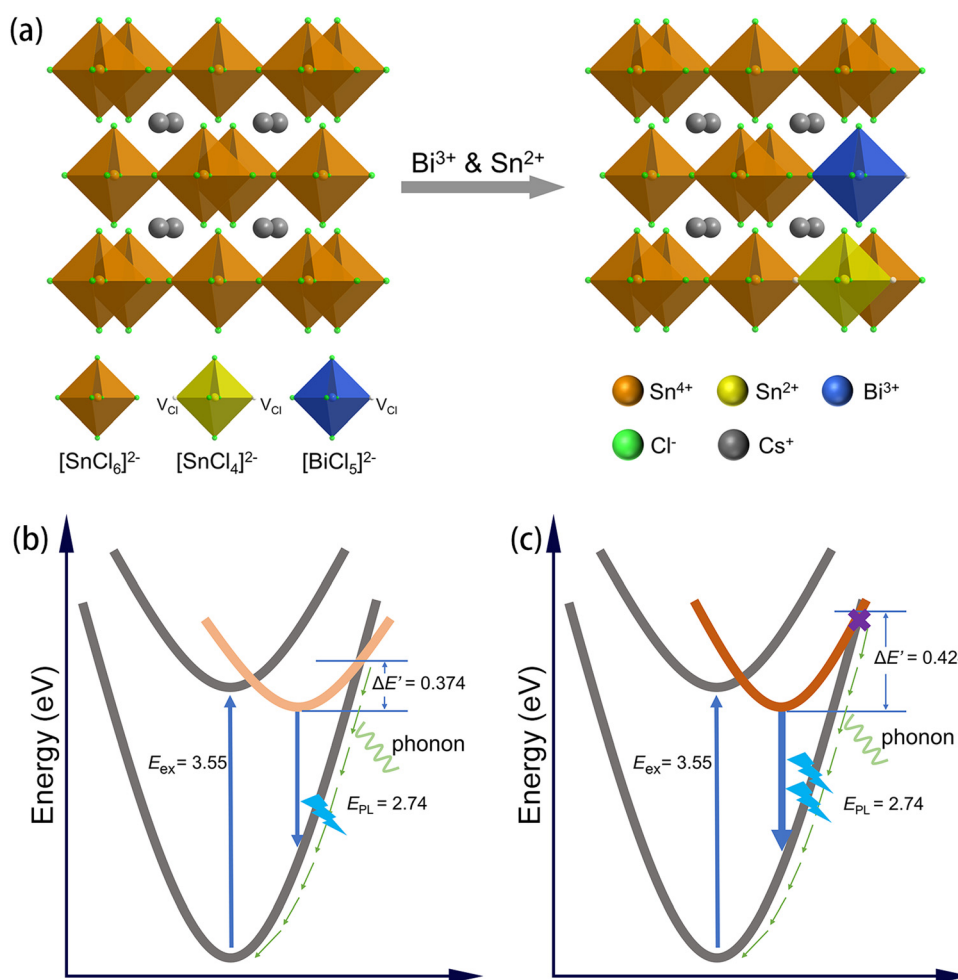


Fig. 5 (a) Schematic diagram of the crystal structure of Bi^{3+} and Sn^{2+} -doped Cs_2SnCl_6 . Schematic of the energy level structure for (b) $\text{Sn}^{4+}:\text{Bi}$ and (c) $\text{Sn}^{4+}:\text{Bi}@\text{H}_3\text{PO}_2$.

induces more localized excitons by the Jahn-Teller effect,³⁰ which can cause a large additional vibrational energy to prohibit the non-radiative transitions and enhance the efficiency of radiative recombination. To prove this inference, the extra vibrational energy is the direct evidence. Fig. S6 (ESI[†]) shows the change

in the temperature-dependent PL intensity. Above 125 °C, $\text{Sn}^{4+}:\text{Bi}@\text{H}_3\text{PO}_2$ still emitted light, whereas $\text{Sn}^{4+}:\text{Bi}$ no longer emitted light. The emission intensity of $\text{Sn}^{4+}:\text{Bi}@\text{H}_3\text{PO}_2$ decreased more slowly than that of $\text{Sn}^{4+}:\text{Bi}$ with the increase of temperature. According to eqn (1),



$$I = I_0 / (1 + e^{-\Delta E' / k_B T}) \quad (1)$$

where I is the PL intensity, $\Delta E'$ is the extra vibrational energy required to produce a non-radiative transition, and k_B is the Boltzmann constant.³⁶ It can be determined that $\text{Sn}^{4+}:\text{Bi}@H_3\text{PO}_2$ should have a higher $\Delta E'$ than $\text{Sn}^{4+}:\text{Bi}$. In order to obtain the value of the extra vibrational energy, the PL spectra of the samples $\text{Sn}^{2+}:0.08\% \text{ Bi}$ and $\text{Sn}^{4+}:0.04\% \text{ Bi}$ were measured at different temperatures. Through fitting the PL intensity as a function of temperature as shown in Fig. S7 (ESI[†]), the values of $\Delta E'$ for the samples $\text{Sn}^{2+}:0.08\% \text{ Bi}$ and $\text{Sn}^{4+}:0.04\% \text{ Bi}$ are 424 and 374 meV, respectively. $\text{Sn}^{2+}:0.08\%$ had a larger value of the extra vibrational energy than $\text{Sn}^{4+}:0.04\% \text{ Bi}$, which means that the non-radiative transitions of $\text{Sn}^{2+}:0.08\% \text{ Bi}$ can be further prohibited to improve the efficiency of radiative transitions. As shown in Fig. 5b and c, the incorporated Sn^{2+} in the state of $[\text{SnCl}_4]^{2-}$ and two V_{Cl} increased the energy value of $\Delta E'$ of $[\text{BiCl}_5]^{2-}$ due to the more localized excitons, which reduces the probability of a non-radiative transition, and increases the probability of a radiative transition. Thus, $\text{Cs}_2\text{SnCl}_6:\text{Bi}^{3+}$ and Sn^{2+} can finally achieve a significant improvement of PLQY.

Conclusion

In summary, a series of $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ with different precursors have been synthesized, and the photophysical properties were investigated. Comparing $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ synthesized by SnCl_4 , $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ from SnCl_2 shows much higher PLQY, and the doping content of Bi^{3+} has little effect on the high PLQY. The reductive H_3PO_2 in the precursor solution of SnCl_4 makes the prepared $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ show brighter emission light and higher PLQY (35.4%) than the sample without H_3PO_2 (PLQY of 13.7%). This is because H_3PO_2 reduces Sn^{4+} to Sn^{2+} during the reaction, and the XPS spectra proved the existence of Sn^{2+} in the $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ sample synthesized with H_3PO_2 . The sample is actually Bi^{3+} and Sn^{2+} co-doped Cs_2SnCl_6 . $[\text{SnCl}_4]^{2-}$ and two V_{Cl} induce the local lattice distortion, which can cause the heavy localized excitons of $[\text{BiCl}_5]^{2-}$ to enhance the efficiency of radiative recombination. This work improves the deep understanding of the luminescence mechanism of $\text{Cs}_2\text{SnCl}_6:\text{Bi}$, and provides a new idea for developing efficient perovskite materials.

Experimental section

Materials and chemicals: cesium chloride (CsCl , Macklin, 99.9%), bismuth chloride (BiCl_3 , Macklin, 99.99%), tin chloride dihydrate ($\text{SnCl}_2 \cdot 2\text{H}_2\text{O}$, Aladdin, 99.99%), tin tetrachloride (SnCl_4 , Energy Chemical, 99%), hypophosphorous acid (H_3PO_2 , Aladdin, 50 wt% in water), and hydrochloric acid (HCl , Guangzhou Chemical Reagent Factory, 37 wt% in water) were the starting reagents. All of the chemicals were used without further purification, unless otherwise stated.

Preparation of precursors: 4 mL SnCl_4 was dissolved in 17 mL HCl to prepare 2 M SnCl_4 solution. A total of 12 mmol (3784.08 mg) BiCl_3 was dissolved in 6 mL HCl to prepare 2 M BiCl_3 solution.

Synthesis of $\text{Cs}_2\text{SnCl}_6:\text{Bi}$: an amount of 336.72 mg (2 mmol) CsCl , 225.65 mg (1 mmol) $\text{SnCl}_2 \cdot 2\text{H}_2\text{O}$ or 0.5 mL (1 mmol) SnCl_4 precursor, and 0–50 μL (0–0.1 mmol) BiCl_3 precursor were added in a polytetrafluoroethylene (PTFE) container with 15 mL of 37% HCl . The container was placed in a muffle furnace and kept at 180 °C for 10 h. Crystals were obtained by slowly cooling the solution down to room temperature over the course of 20 h. White crystals of $\text{Cs}_2\text{SnCl}_6:\text{Bi}$ could be separated by immediate centrifugation, and were washed three times with ethanol.

Synthesis of $\text{Cs}_2\text{SnCl}_6:\text{Bi}@H_3\text{PO}_2$: an amount of 336.72 mg (2 mmol) CsCl , 500 μL (1 mmol) SnCl_4 precursor, 50 μL (0.1 mmol) BiCl_3 precursor and 0.05 mL H_3PO_2 were added in a polytetrafluoroethylene (PTFE) container with 15 mL of 37% HCl . The container was placed in a muffle furnace and kept at 180 °C for 10 h. Crystals were obtained by slowly cooling the solution down to room temperature over the course of 20 h. White crystals of $\text{Cs}_2\text{SnCl}_6:\text{Bi}@H_3\text{PO}_2$ could be separated by immediate centrifugation and were washed three times with ethanol.

Measurement and characterization: the phase determination of the obtained samples was carried out by powder X-ray diffraction (XRD) (D2, Bruker, using a $\text{Cu K}\alpha$ rotating anode). Depth-profile XPS spectra were recorded on an AXIS Supra, using an X-ray beam focused to an area of 2 mm² on the sample, and with an etching at the rate of 1 nm for 35 s. The PL, PLE, PLQY spectra and decay curve collection were recorded on an Edinburgh Instruments FS5 spectrometer equipped with a Xenon lamp and an integrating sphere. Inductively coupled plasma optical emission spectrometry (ICP-OES) was performed by using an atomic emission spectrometer (Agilent 5110). UV-vis absorption spectra were collected with a UV-3600i Plus spectrometer.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was financially supported by the National Natural Science Foundation of China (Grant No. 61874074), Science and Technology Project of Shenzhen JCYJ20220531100815034, Technology and Innovation Commission of Shenzhen (20200810164814001), Natural Science Foundation of Guangdong Province (General Program, Grant No. 2022A1515012055 and 2023A1515030014), and Shenzhen Fundamental Research Projects (Grant No. JCYJ20220531103411026).

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