



Cite this: *Mater. Adv.*, 2023,
4, 3822

The thermodynamic and mechanical properties of Earth-abundant metal ternary boride $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions for impact- and wear-resistant alloys

Pavlo Prysyazhnyuk ^{*a} and Devis Di Tommaso ^b

The tetragonal ternary borides Mo_2MB_2 ($\text{M} = \text{Fe}, \text{Mn}$) and their solid solutions are promising candidates for tungsten-free wear-resistant alloys. In this study, we investigated the thermodynamic, mechanical, vibrational, structural, and electronic properties of Mo_2MB_2 compounds and their solid solutions using density functional theory (DFT) calculations. To model $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ as disordered substitutional solid solutions, we employed the DFT-based virtual crystal approximation (VCA) and the cluster expansion (CE) approaches. The calculation results showed that both Mo_2MB_2 compounds are mechanically and thermodynamically stable, exhibiting similar features in terms of electronic structure and chemical bonding types. The temperature-dependent interaction parameter for the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ phase in the compound energy formalism (CEF) notation was accessed using a combination of the ATAT and the Thermo-Calc software. To experimentally validate the calculated thermodynamic parameters, an isopleth was constructed by intersecting the Fe–Mn–Mo–B–C alloy system within suitable concentration ranges for hardfacing alloy development, and the key alloys were analyzed to compare their phase composition and constitution with the calculation results. The Vickers hardness (H_V) of the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions was determined as an average value from different models, including elastic constants calculated using DFT. Both VCA and CE approaches showed an H_V increase from 22.9 to 24.7 GPa within the composition range of $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$ to $\text{Mo}_2(\text{Fe}_{0.25}\text{Mn}_{0.75})\text{B}_2$. Based on the calculated properties, $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solutions can serve as a reinforcement phase in high-manganese hardfacing Fe-based alloys.

Received 19th June 2023,
Accepted 30th July 2023

DOI: 10.1039/d3ma00313b

rsc.li/materials-advances

1. Introduction

Tungsten carbide (WC) is the typical material used in applications that require wear resistance. However, the price of raw materials for tungsten carbide has increased significantly in recent decades. WC-based materials are widely applied in diverse fields, such as machining, oil and gas equipment, mining machines, cermets, wear-resistant alloys, and coatings, owing to their versatility. Recent advancements in tungsten-free materials with adequate mechanical properties now offer opportunities for partial replacement of WC-based materials in specific areas, including hardfacing alloys, where their use is less critical than in the energy industry. This approach provides a more rational redistribution of available tungsten resources. Transition-metal ternary borides such as Mo_2FeB_2 are materials

with unique physical, chemical, and mechanical properties: high microhardness, stability in contact with liquid alloys and relatively low price compared to WC.¹ Mo_2FeB_2 is used as a component of metal matrix composites and coatings with the composite structure using Fe-based alloys alloyed with other metals such as Mn, Cr, and Ni as the metal matrix.^{2–4} For example, the effect of Mn in such materials and coatings is critical to achieving complex impact abrasion wear resistance due to the formation of a TWinning Induced Plasticity (TWIP) steel⁵ matrix. Hence, the intentional development of hardfacing alloys with such a combination of characteristics can be achieved by predicting the effect of Mn on the composition and properties of key (both ceramic and metal) phases using atomistic and thermodynamic models of the corresponding solid solutions. Investigations of the alloying with Mn on the microstructure and the mechanical properties of the sintered Fe/ Mo_2FeB_2 composites showed that the Mn additions lead to an increase in plasticity and transverse rupture strength and a decrease in hardness and size of the Mo_2FeB_2 grains.^{6,7} Remarkable improvements were also observed in the

^a Department of Welding, Ivano-Frankivsk National Technical University of Oil and Gas, 076019, Ivano-Frankivsk, Ukraine. E-mail: pavlo.prysyazhnyuk@nunge.edu.ua^b Department of Chemistry, School of Physical and Chemical Sciences, Queen Mary University of London, Mile End Road, London E14NS, UK

tribological properties measured in contact with the ceramic (Al_2O_3) counter body.⁸ Using diffraction patterns analysis, Yu *et al.*⁶ showed that $(\text{Mo,Fe,Mn})_3\text{B}_2$ substitution solutions are possible, but the detailed solubility ranges of Mn in Mo_2FeB_2 were not determined.

The Mo_2FeB_2 materials has also been the focus of theoretical investigations. Wang *et al.*⁹ used density functional theory (DFT) to calculate the enthalpies of formation, electronic properties, and elastic constants for Mo_2FeB_2 for different magnetism (nonmagnetic, ferromagnetic, and antiferromagnetic) and it was found that Mo_2FeB_2 is a ground state structure. In further DFT investigations,¹⁰ elastic constants and thermodynamic properties were calculated for Mo_2FeB_2 at constant volume and different pressures in the range from 0 to 100 GPa. Shi *et al.*¹¹ proposed a model based on chemical bond strength analysis and Mulliken overlap population calculation to predict the hardness of ternary borides. Their model was in good agreement with experimental results for the Vickers hardness (H_V) of Mo_2FeB_2 (23.06 GPa), with a calculated value of 21.7 GPa compared to the experimental value of 23.1 GPa. He and co-workers¹² used another approach to calculate the hardness of Mo_2FeB_2 based on elastic constants, according to the empirical Chen model.¹³ The reported H_V value of 24.6 GPa was also close to the experimental value.

In contrast to Mo_2FeB_2 , the experimental data for Mo_2MnB_2 are essentially limited by the isothermal section constructed in ref. 14, where a stable phase with a U_3Si_2 -type structure was observed and described as solid solution $(\text{Mo,Mn})_3\text{B}_2$. However, DFT calculations performed by Yang *et al.*¹⁵ show that Mo_2MnB_2 is a stable phase with a predicted H_V value of 22.8 GPa. Atomistic models of the $(\text{Mo,Mn})_2(\text{Fe,Mn})\text{B}_2$ of substitutional solid solutions were also generated by partially replacing Fe or Mo in the Mo_2FeB_2 parent lattice with Mn avoiding the use of supercell structures.¹⁶ The resulting structures after geometry optimization were more stable than Mo_2FeB_2 and had increased Young's modulus and H_V values. This approach is, however, not sufficiently robust to predict the properties of the solution in the ground state due to the limited number of atomic configurations. To overcome this limitation, it can be extended with techniques that allow the analysis of a large number of clusters, such as supercells with different atomic arrangements. Two such techniques are cluster expansion (CE) and special quasirandom structures (SQS) formalism.¹⁷ These techniques enable a systematic *ab initio* investigation of ground-state structures, which represent the lowest energy set of supercells corresponding to disordered structures with various composition in real alloy systems. This is accomplished through stepwise concentration-dependent DFT lattice stability calculations, performing the screening of energy in composition space within the given crystal structure symmetry. The main distinction between CE and SQS lies in their approaches to supercell generation. CE involves DFT calculations of a wide range of supercells with different compositions and sizes, whereas SQS utilizes a limited number of relatively small supercells (typically containing 30–50 atoms per cell) that comprise the most important

interactions between components. Despite the wide integration of the SQS approach into software tools for high-throughput computing¹⁸ and its successful testing for various phase models, the tetragonal (space group $P4/mbm$) structure is not included in the current SQS database of the ATAT software.¹⁹ Therefore, additional SQS model generation is needed using corresponding algorithms as implemented in the mcsqs code,²⁰ along with further validation of the data using other calculation approaches and experimental data. Another suitable approach for modelling disordered substitutional solid solutions *via* DFT is the Virtual Crystal Approximation (VCA) technique.²¹ This method involves the concentration-dependent averaging of pseudopotentials for the crystal lattice sites occupied by alloy components. However, it is important to note that the applicability of the VCA is limited to systems that exhibit insignificant deviations from Vegard's law.²²

A comprehensive analysis of the experimental investigations^{3,4,6,7} and theoretical studies^{9–12,15,16} of Mn_2FeB_2 and similar compounds shows that current investigations mainly focus on these compounds in their pure state. However, there is a lack of data regarding the properties of their based solid solutions. Therefore, the motivation behind this study was to investigate Mo_2FeB_2 -based solid solutions using DFT-based methods supported by key experiments. As Mn is a typical component of impact- and wear-resistant alloys, the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions were the main focus, being promising candidates for use as a reinforcement phase in the design of new alloys.

In the present study, DFT calculations were employed to investigate the structural, mechanical, thermodynamic and electronic properties of Mo_2FeB_2 and Mo_2MnB_2 ternary borides as well as of $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions which were modeled using CE and VCA approaches. Finally, the thermodynamic parameters were determined according to the compound energy formalism (CEF)²³ representation in order to extend existing thermodynamic databases for the calculation of the equilibrium phase composition of impact- and wear-resistant alloys.

2. Methodology

2.1 Atomistic models

The crystal structure of the ternary borides Mo_2FeB_2 and Mo_2MnB_2 has a tetragonal lattice unit cell belonging to the $P4/mbm$ space group and containing two Fe(Mn), four Mo, and four B atoms at the Wyckoff positions 2a, 4g, and 4h, respectively, as shown in Fig. 1(a). Such a structure arises through cooperative growth of the cubic intermetallic MoFe (MoMn) phase (space group $Pm3m$) and the hexagonal diboride MoB_2 (space group $P6/mmm$).²⁴ The experimental lattice parameters are $a = b = 5.782 \text{ \AA}$ and $c = 3.148 \text{ \AA}$ for Mo_2FeB_2 ²⁵ and $a = b = 5.781\text{--}5.812 \text{ \AA}$ and $c = 3.127\text{--}3.136 \text{ \AA}$ ²⁶ for Mo_2MnB_2 . According to the CEF and considering the data presented in ref. 27, the $\text{Mo}_2(\text{M}^I, \text{M}^{II})\text{B}_2$ solid solutions can be modelled as a phase with three sublattices filled with Fe(Mn), Mo and B and described



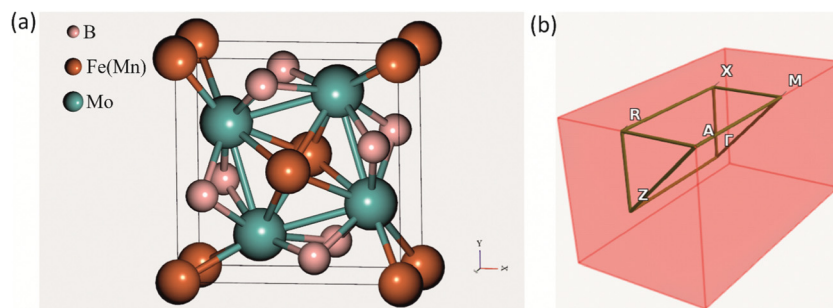


Fig. 1 Crystal structure of a tetragonal Mo_2FeB_2 and Mo_2MnB_2 unit cell (a) and corresponding BZ with high symmetry points (b).

with a formula unit $(\text{Fe,Mn})_{0.2}^{\text{I}}(\text{Mo})_{0.4}^{\text{II}}(\text{B})_{0.4}^{\text{III}}$, where the symbol “:” denotes sublattices, the subscript the stoichiometric coefficients, and superscript the number of the sublattice. Such a phase representation was used to model the disordered substitutional solid solutions within the VCA approach, in which the average weight of Fe and Mn pseudopotentials at a given composition was set to 2a site. To apply the CE approach to the $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions, the Fe or Mn was placed at 2a sites and supercells (clusters) with different Mn concentration were generated for energy analysis to determine effective cluster interactions (ECI) and ground state properties as implemented in the ATAT toolkit.¹⁹ The Monte Carlo simulation, combined with the semi-grand canonical potential of lattice models’ energetics resulting from the CE analysis as implemented in the Easy Monte Carlo code (emc2)²⁸ integrated into the ATAT toolkit, was utilized to extrapolate composition-dependent configurational mixing thermodynamic quantities (free energy, enthalpy, and entropy) to non-zero temperatures (in the range from 1000 to 2000 K). This approach provided valuable data for constructing phase diagrams that reflect real conditions. The calculated excess properties were converted into a format suitable for thermodynamic assessment as suggested by Gautam Ghosh in ref. 19. Next, the thermodynamic parameters of the $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions were accessed via the PARROT module of the Thermo-Calc 2022 software.²⁹

2.2 Computational details

All DFT-based first-principles calculations for structure relaxation during VCA and CE analysis were carried out with Vienna *Ab initio* Simulation Package (VASP v. 6.1.1) plane wave *ab initio* code³⁰ using general gradient approximation (GGA)³¹ with Perdew–Burke–Ernzerhof (PBE)³² exchange–correlation functional. For the integration over the first Brillouin zone (BZ) (Fig. 1(b)) the Monkhorst–Pack³³ meshing scheme of 1000 k -points per reciprocal atom (KPPRA), corresponding to the $4 \times 4 \times 8$ special k -points sampling and ferromagnetic state for the Mo_2MeB_2 was set. During the self-consistent field (SCF) calculations and the cutoff energy for the plane-wave-basis set was 520 eV at the convergence criterion for the total energy of 10^{-7} eV. Molybdenum borides with different stoichiometries were chosen as reference phases for the adjustment of parameters during lattice stability calculations, using the following k -point meshes: $12 \times 12 \times 7$ for MoB_2 , $8 \times 8 \times 1$ for $\alpha\text{-MoB}$, $8 \times 3 \times 8$

for $\beta\text{-MoB}$, and $4 \times 4 \times 5$ for Mo_2B . For the relaxed structures of interest, the elastic constants and related properties were calculated by deriving the elastic constants from the strain–stress relationship³⁴ with further post-processing calculation using the VASPKIT utility.³⁵

To obtain the Vickers hardness (H_V), which is one of the key mechanical properties for wear-resistant alloys, calculations were performed using different formalisms based on approximations using the elastic moduli values as proposed by Chen *et al.* (H_V^{Chen}),¹³ Tian *et al.* (H_V^{Tian}),³⁶ Teter (H_V^{Teter}),³⁷ Mazhnik and Oganov (H_V^{Mazhnik}),³⁸ Miao *et al.* (H_V^{Miao}).³⁹ To validate the calculated data through experimental investigations, the approach proposed by Kabir *et al.*⁴⁰ was used, which includes obtaining H_V as an average value from those calculated with different models (eqn (1)–(5)):

$$H_V^{\text{Chen}} = 2[(G/B)^2 G]^{0.585} - 3 \quad (1)$$

$$H_V^{\text{Tian}} = 0.92(G/B)^{1.137} G^{0.708} \quad (2)$$

$$H_V^{\text{Teter}} = 0.151G \quad (3)$$

$$H_V^{\text{Miao}} = \frac{(1 - 2\nu)E}{6(1 + \nu)} \quad (4)$$

$$H_V^{\text{Mazhnik}} = 0.096 \frac{(1 - 8.5\nu + 19.5\nu^2)E}{(1 - 7.5\nu + 12.2\nu^2 + 19.6\nu^3)} \quad (5)$$

where B , G , and E are the bulk, shear, and Young’s modulus, respectively, and ν is the Poisson’s ratio.

The phonon calculations were performed using the PHONOPY code⁴¹ in combination with density-functional-perturbation theory (DFPT)⁴² calculations of the force-constants as implemented in VASP. For this purpose, $2 \times 2 \times 2$ supercells with 80 atoms per cell were constructed from the optimized Mo_2FeB_2 and Mo_2MnB_2 structures. To obtain thermodynamic properties at constant pressure, the Quasi-Harmonic Approximation (QHA) approach was combined with DFPT and applied to the supercells with different volumes to obtain the Vinet⁴³ equations of state (EOSs) at given temperature ranges. Using EOSs for volume minimization in the given temperature ranges, the Gibbs free energy and thermal properties at constant pressure were obtained.

To obtain accurate electronic properties, all-electron calculations were performed using the Amsterdam Modelling Suite (AMS) BAND⁴⁴ module with the linear combination of atomic



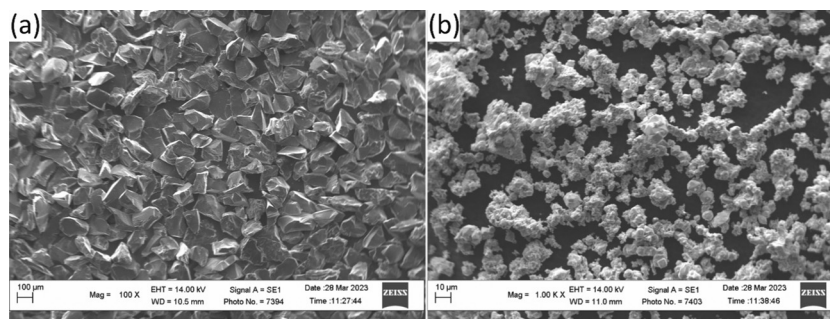


Fig. 2 The morphology of the initial boron carbide (a) and Mo (b) powders.

orbitals (LCAO) and the Slater-type basis functions with triple- ζ polarized orbital basis set (TZP). The regular k -point sampling grid of the first BZ was $5 \times 5 \times 9$ and other settings were set according to the «good» numerical quality for spin-unrestricted calculations.

2.3 Experimental details

For the validation of the calculations results, the experimental alloys were prepared by arc-melting of the initial components in the form of powder mixtures drawn into low-carbon steel sheath to reproduce the compositions close to the typical Fe-based hardfacing alloys with nominal compositions $\text{Fe}_{0.6975-x}\text{Mn}_x\text{Mo}_{0.135}\text{B}_{0.134}\text{C}_{0.0335}$.⁴⁵ As initial components, powders of boron carbide with an angular particles morphology (Fig. 2(a)) (fraction F120, PJSC “Zaporozhabrasive”, Ukraine), Mo powder with a spherical (Fig. 2(b)) (PM grade, 1–5 μm , GrandLada Co., Ltd) and Mn powder (grade MN997 (GOST 6008-90), JSC “Nikopol Ferroalloy Plant”, Ukraine). The key alloys were investigated using X-ray diffraction (XRD) (Shimadzu XRD-7000, Japan) and further analyzed with Match! Software⁴⁶ Microstructure investigations were performed using scanning electron microscopy (SEM) techniques (Zeiss EVO-40XVP, UK), including secondary electrons (SE) and backscattered electron diffraction (BSD). The microhardness measurements were performed using the Vickers method with an indentation load of 0.2 kg (PMT-3M hardness tester, LLC NPF “Standard-M”, Ukraine).

3. Results and discussions

3.1 Energy and Stability

Accurately predicting the formation energies from the DFT calculations of compounds such as Mo_2FeB_2 and Mo_2MnB_2 that involve bonding between metallic and non-metallic components of a phase requires corrections due to the systematic error caused by chemical bonding features.⁴⁷ According to Stevanović *et al.*⁴⁸ the corrected formation energy $\Delta H_f(A_{n1}A_{n2} \dots)$ can be written as follows:

$$\Delta H_f(A_{n1}A_{n2} \dots) = \Delta H_f^{\text{GGA}+U}(A_{n1}A_{n2} \dots) - \sum_i n_i \delta \mu_i \quad (6)$$

where A_{n1} and A_{n2} are the components of a phase, $\Delta H_f^{\text{GGA}+U}(A_{n1}A_{n2} \dots)$ is the formation energy for the compound and components in

their reference states calculated at the DFT-GGA level with the Hubbard U correction (3 eV for the transition metals), $\delta \mu_i$ is the correction term for the component i , and n_i is the stoichiometric coefficient for that component. In a given case, to calculate the formation energies for Mo_2FeB_2 and Mo_2MnB_2 , the correction terms $\delta \mu_i$ for Fe and Mn are equal to -0.03 and -0.15 eV, respectively,⁴⁸ while the parameter $\delta \mu_{\text{Mo}}$ has to be determined. In order to obtain value of $\delta \mu_{\text{Mo}}$, the formation energies of the four molybdenum borides with different structure and composition (Fig. 3) were calculated and fitted using least-square method with respect to the experimental values.²⁷ The values of $\Delta H_f^{\text{GGA}+U}$ of Mo_xB_y computed using the GGA+ U method show a similar concentration dependence to the experimental values. In both cases, the formation energies of α -MoB and β -MoB are the highest and are close to each other, while $\Delta H_f^{\text{GGA}+U}$ for MoB_2 and Mo_2B is relatively higher (Fig. 4). After solving the least-squares problem with respect to $\delta \mu_{\text{Mo}}$ according to eqn (6) the corresponding value was $+0.237$ eV. As a result of using the obtained correction term for Mo in molybdenum borides, the accuracy of the prediction of the energies of formation was significantly improved, since the calculated values for α -MoB, β -MoB and Mo_2B represent experimental data within the chemical accuracy (4184 J mol^{-1}). Considering the values for Fe and Mn from⁴⁸ and the fitted value for Mo, the resulting formation energies for the Mo_2FeB_2 and Mo_2MnB_2 target phases were calculated according to eqn (6) as follows:

$$\Delta H_f(\text{Fe}(\text{Mo}, \text{B})_2) = -47\,562 - 0.4 \times 22\,905 + 0.2 \times 14\,472 = -53\,830 \text{ J mole}^{-1} \quad (7)$$

$$\Delta H_f(\text{Mn}(\text{Mo}, \text{B})_2) = -36\,114 - 0.4 \times 32\,468 + 0.2 \times 2894 = -44\,698 \text{ J mole}^{-1}. \quad (8)$$

3.2 Electronic properties

To compare the electronic structure of Mo_2FeB_2 and Mo_2MnB_2 , the total density of states (TDOS) and the partial density of states (pDOS) were performed (Fig. 5). Both compounds are metallic because there is no gap at the Fermi energy level (E_F). This is due to the formation of peaks near the E_F in the 0 to -2 eV range, which are mainly composed of d-states of the metal atoms. These d-states provide metallic bonding between Fe–Mo and Mn–Mo pairs. At the same time, the DOS on E_F is



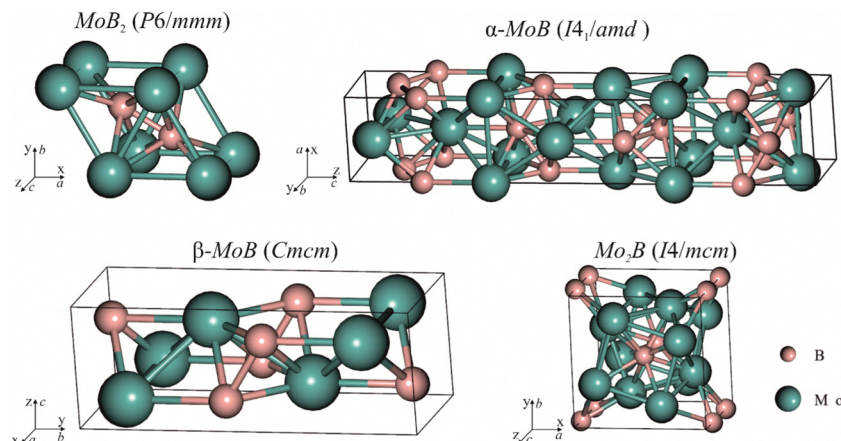


Fig. 3 Crystal structures of molybdenum borides selected for energy of formation calculations and fitting with experimental data.

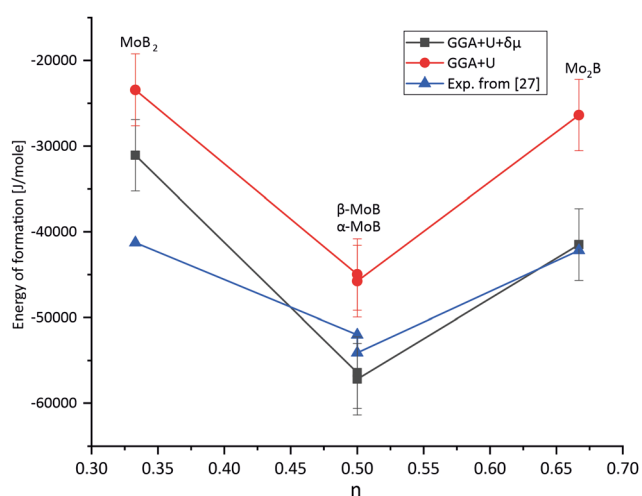


Fig. 4 Results of fitting the formation energies for the molybdenum borides with different stoichiometry using the $\delta\mu_{\text{Mo}}$ correction term.

relatively higher for Mo_2MnB_2 , showing the higher metallic properties and electrical conductivity for this compound. In both cases, there are high DOS peaks below the E_F , located between -2 and -7 eV and composed of p-states of boron and d-states of metal atoms, consistent with pd hybridization and formation of covalent bonds between Fe–B, Mn–B, and Mo–B pairs. At the lower level, corresponding to the energy of about -10 eV, there is a DOS peak resulting from the overlap of s and p-states, leading to the sp hybridization, typical for the covalent bonds between boron atoms. For both Mo_2FeB_2 and Mo_2MnB_2 , peaks due to antibonding p-states of B and Mo in antiphase interaction have been observed above the Fermi energy (E_F) in the energy range of 12 to 20 eV. These peaks are located near the Fermi level and are composed of conductivity-enhancing d-states of metal atoms.

Further insight in the electronic properties of Mo_2MnB_2 and Mo_2FeB_2 , their Fermi surfaces and band structures were calculated (Fig. 6). The calculated Fermi surface (FS) topologies show

that both compounds have similar elongated FS along the ΓZ direction, indicating charge transport anisotropy in both cases. The minor topology differences related to the more closed shape of the FS for Mo_2MnB_2 compared to Mo_2FeB_2 may indicate that Mo_2MnB_2 exhibits stronger electron correlation effects and higher electrical conductivity.⁴⁹ Analysis of the FS curvatures shows that in Mo_2FeB_2 the electron-like and hole-like conductivity coexist, while in the case of Mo_2MnB_2 the electron-like conductivity is more dominant.⁵⁰ The calculated band structures of Mo_2MnB_2 and Mo_2FeB_2 show that in both cases the valence and conduction bands with dominant d-state character overlap at the Fermi level. At the energies below (from -2 eV to -7 eV) there are bands contributed by combinations of dominant p and d states and s and p states in the range -9 eV to -11 eV. The energy dispersion curves in the conduction band just above the Fermi level are contributed by the d states, and at the higher energies (3 eV to 20 eV) the dominant states are s and p with a slight contribution from d electrons. Based on the calculations of TDOS, pDOS, and band structure, it can be concluded that the ternary borides Mo_2MnB_2 and Mo_2FeB_2 exhibit homeopolar B–B bonding, heteropolar B–Mo, B–Fe or B–Mn covalent bonds and metallic bonding between Mo, Fe and Mn.

To study the interatomic bonding properties of Mo_2MnB_2 and Mo_2FeB_2 in more detail, the total electron density (ρ), the electron localization function (ELF), and the reduced density gradient (RDG) were mapped with contour plots along the (001) planes containing Fe or Mn and B atoms (Fig. 7). The mapped fields for both borides appear to be very similar, with the ρ maps showing increased charge density values between B–B and Fe(Mn)–B pairs, corresponding to relatively high ELF values. The highest ELF values of 0.6 to 0.7 (typical for the covalent bond) are between the nearest boron atoms and around the Fe cores, pointing towards the boron atoms. The charge density and ELF values are weaker in the directions between metal atoms, with typical values ranging from 0.2 to 0.3, consistent with metallic bonding. The RDG parameter is informative in the presence of covalent bonds as it indicates



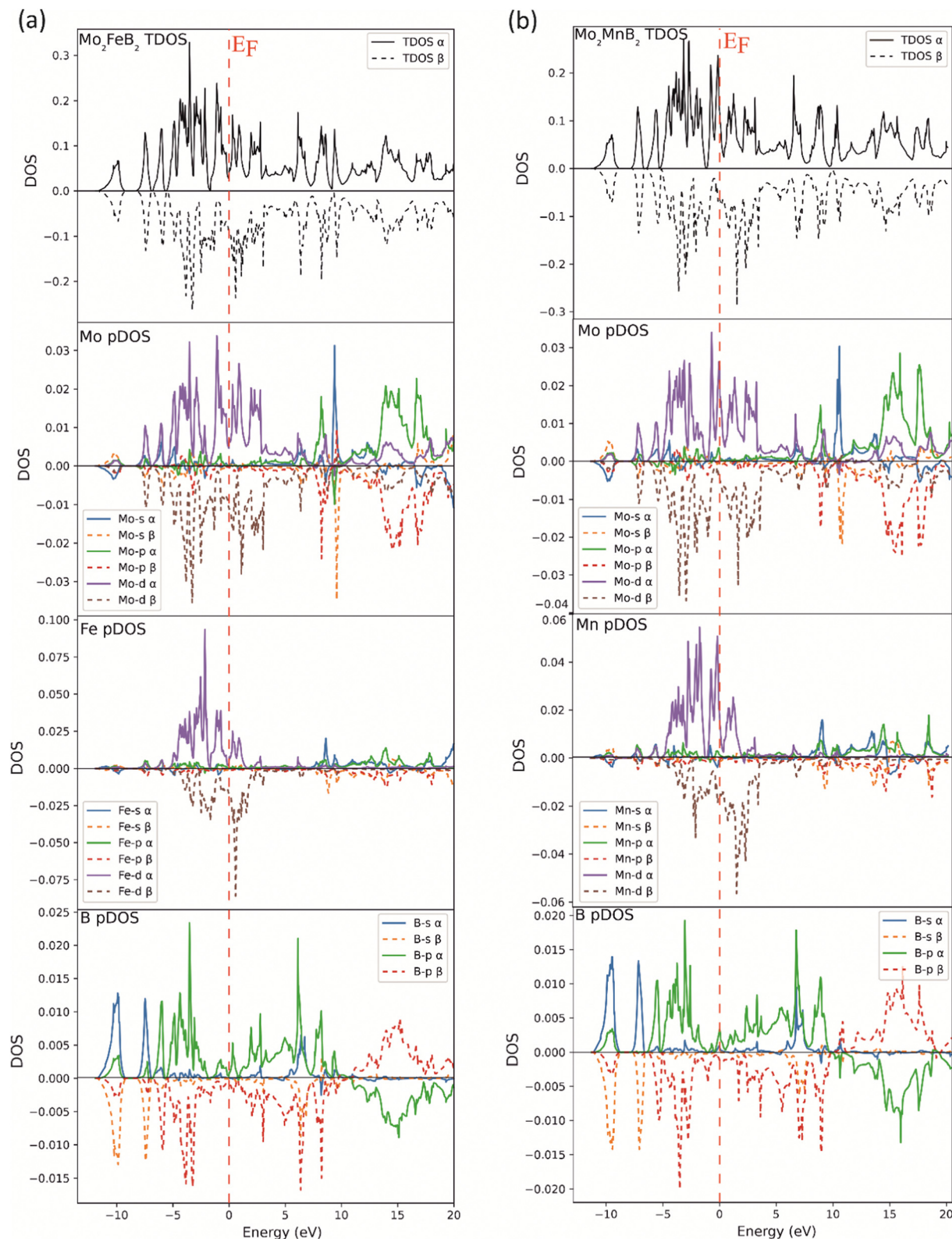


Fig. 5 Total and partial DOS of the Mo_2FeB_2 (a) and Mo_2MnB_2 (b) borides.

regions of high electron cloud stability with low RDG values. In the case of Mo_2MnB_2 and Mo_2FeB_2 , it was found that the regions with the lowest RDG values are located between pairs

of B–B, Fe(Mn)–B, and Fe(Mn)–Fe(Mn), corresponding to all types of bonding identified by the analysis of pDOS and band structure.



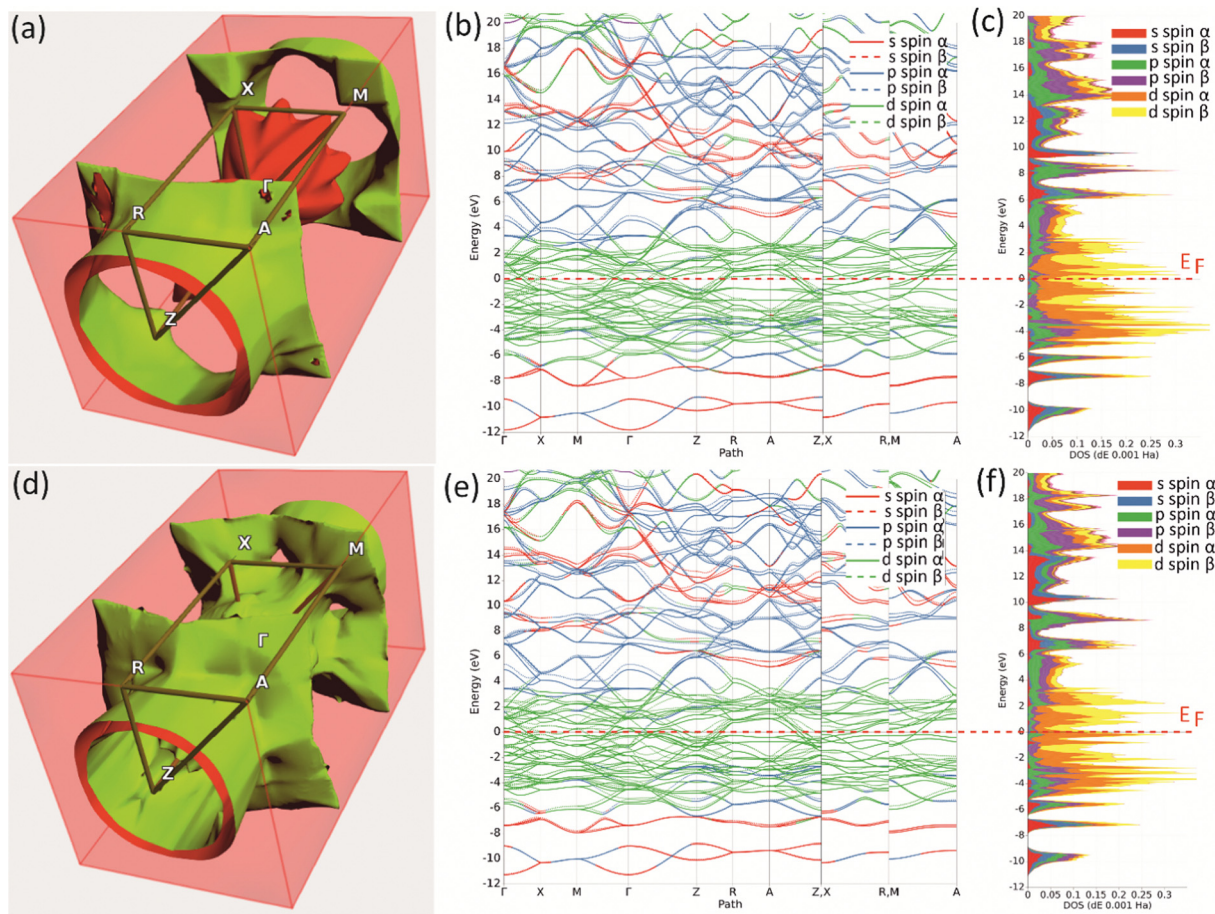


Fig. 6 The Fermi surfaces (a), (d), band structure (b), (e) and projected density of states (c), (f) of the Mo_2FeB_2 (a)–(c) and Mo_2MnB_2 (d)–(f) borides.

3.3 Thermodynamic properties

To evaluate the vibrational properties for the Mo_2FeB_2 , Mo_2MnB_2 compounds and hypothetical solid solution $\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$, phonon calculations were performed. The phonon dispersion curves (ω_{qj}) were calculated for along the high symmetry points $G \rightarrow X \rightarrow M \rightarrow G \rightarrow Z \rightarrow R \rightarrow A \rightarrow Z \rightarrow X \rightarrow R \rightarrow M \rightarrow A$ path (Fig. 8(a), (c) and (e)). No imaginary phonon frequencies were present in all cases, indicating that pure ternary borides and $\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$ solid solution are dynamically stable structures. This result agrees with the previously calculated lattice stability values. The phonon spectra in all cases consist of three branches assigned to the acoustic spectra and the rest to the optical. A comparison of the partial density of phonon states (pPHDOS) (Fig. 8(b), (d) and (f)) between investigated structure shows that the segments contributed mainly by metallic atoms are very similar. However, the segments contributed by boron are wider and have higher frequencies indicating higher interatomic bonding strength.

The phonon dispersion curves for all structures exhibit the highest frequencies (22 THz for Mo_2FeB_2 ($\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$) and 20 THz for Mo_2MnB_2) between the G and Z symmetry points. Both materials have two intervals where no vibrations occur and as a result, the vibrational spectra can be divided into a low-frequency segment, mainly contributed by partial spectra

of Mo with Fe and/or Mn, and a high-frequency segment dominated by B spectra. Based on the relationship between frequencies, force constants, and interatomic bond strength, the low-frequency segment can be attributed to weak metallic bonding, while the high-frequency segments correspond to strong covalent bonds of B–B and B-metal pairs. When comparing Mo_2FeB_2 and Mo_2MnB_2 , the pPHDOS for segments mostly contributed by metal atoms are similar. However, Mo_2FeB_2 shows broader ranges with higher frequencies in segments primarily contributed by boron, indicating stronger interatomic bonding. In $\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$, the pPHDOS of Fe and Mn remain similar in shape to those in the end members, while the pPHDOS of B and Mo exhibit a tendency to overlap. The analysis of the total densities of phonon states (tPHDOS) (Fig. 8(g)) shows that the tPHDOS of $\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$, resulting from the overlapping tPHDOSs of the end members, exhibits values close to averaged. This causes the formation of more uniform peak shapes in a broad high-frequency region corresponding to the covalent bonding, which may have a positive effect on the elastic constants values for solid solutions.

In order to obtain the thermodynamic properties of Mo_2FeB_2 and Mo_2MnB_2 at constant pressure for the purpose of analysis within the CALPHAD approach, phonon spectra and energies were calculated for $2 \times 2 \times 2$ supercells. The volumes of these



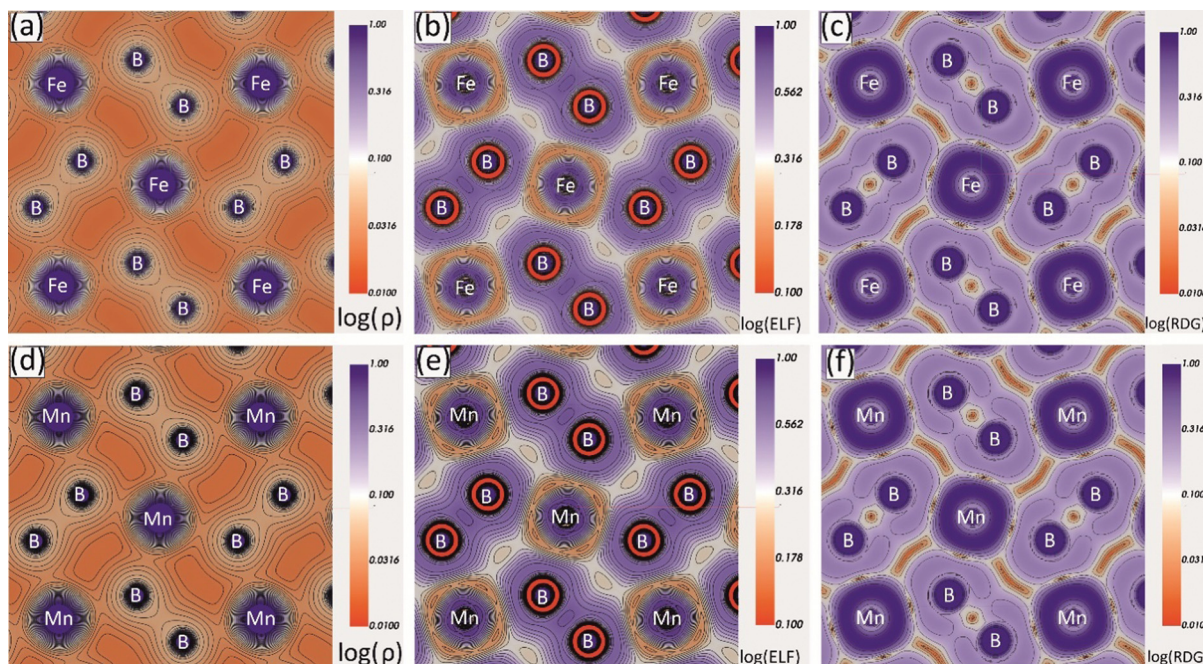


Fig. 7 Charge density (ρ), electron localization function (ELF) and reduced density gradient (RDG) contour maps along the (001) plane for the Mo_2FeB_2 (a)–(c) and Mo_2MnB_2 (d)–(f) borides.

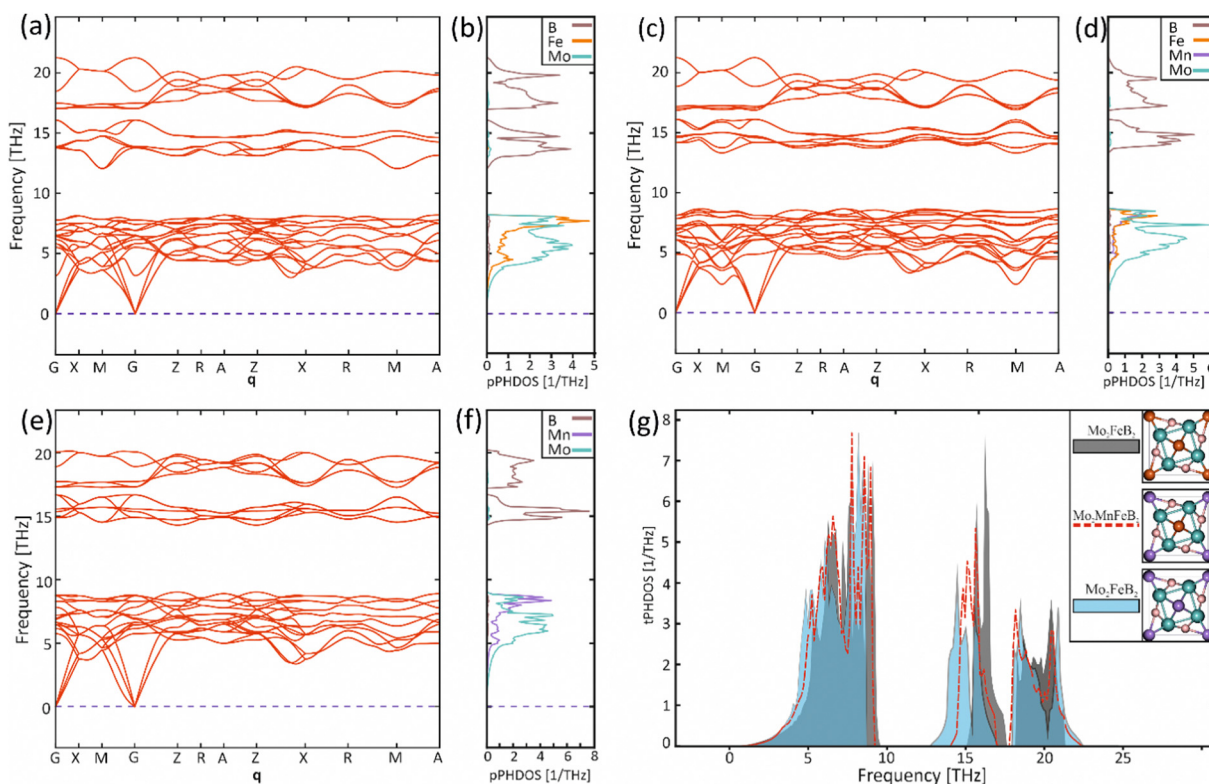


Fig. 8 Phonon dispersion curves (a), (c), (e) along with the partial phonon density of states (b), (d), (f) for the Mo_2FeB_2 (a), (b), $\text{Mo}_2\text{Fe}_{0.5}\text{Mn}_{0.5}\text{B}_2$ (c), (d), Mo_2MnB_2 (e), (f) and the corresponding total phonon densities of states (g).

supercells were varied in steps of 2% by creating 5 supercells each in the decreasing and increasing directions relative to the

equilibrium volume. Using the Phonopy code, temperature-dependent constant-volume thermodynamic properties and

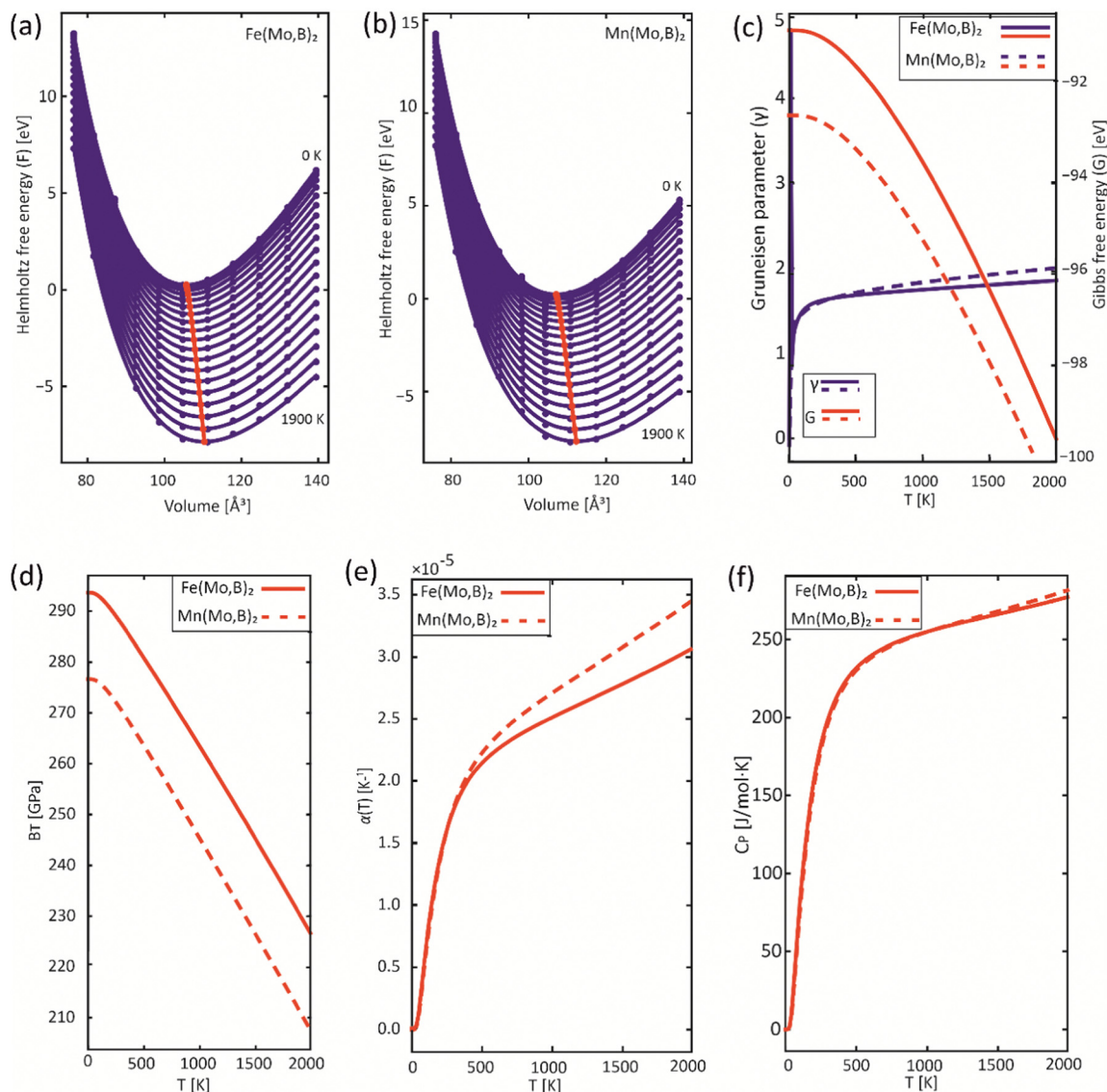


Fig. 9 Results of the QHA fitting of the EOSs (a), (b) and the resulting temperature dependencies of the Gruneisen parameter and Gibbs energy (c), bulk modulus (d), thermal expansion (e) and heat capacity (f).

equations of state through the quasi-harmonic approximation (QHA) were derived (Fig. 9(a) and (b)) and the temperature-dependent values for the Gruneisen parameter (γ) (Fig. 9(c)), Gibbs energy (G) (Fig. 9(c)), bulk modulus (B_T) (Fig. 9(d)), thermal expansion coefficient (α) (Fig. 9(e)), and specific heat capacity at constant pressure (C_p) (Fig. 9(f)) were determined. Upon analysis of the obtained data, it was revealed that Mo₂FeB₂ exhibits a higher bulk modulus compared to Mo₂MnB₂, with a difference of approximately 20 GPa, while Mo₂FeB₂ demonstrates a lower coefficient of thermal expansion (α) within the studied temperature range of 0 to 2000 K. However, the specific heat capacity values were found to be very similar for both compounds, which can be attributed to their similar electronic properties and chemical bonding nature (Fig. 7).

To obtain the necessary thermodynamic data for thermodynamic assessment, such as entropy, we performed numerical

differentiation of the Gibbs energy values with respect to the temperature. According to the SGTE (Scientific Group Thermo-data Europe) notation. The temperature dependence for the molar Gibbs energy of Mo₂MB₂ compounds (where M = Fe or Mn) can be represented by the following equation, assuming a CEF model with three sublattices filled with M, Mo, and B:

$$G_{M:Mo:B}^{Mo_2MB_2} = a + bT + cT \ln T + 0.2G_M^{SER} + 0.4G_{Mo}^{SER} + 0.4G_B^{SER} \quad (9)$$

where a , b and c are the model coefficients to be fitted and G^{SER} is the Gibbs energy of the pure elements, referred to the enthalpy for its stable state at 298.15 K. After fitting the results obtained by the QHA approach and the formation energies using the PARROT module the result representations of the



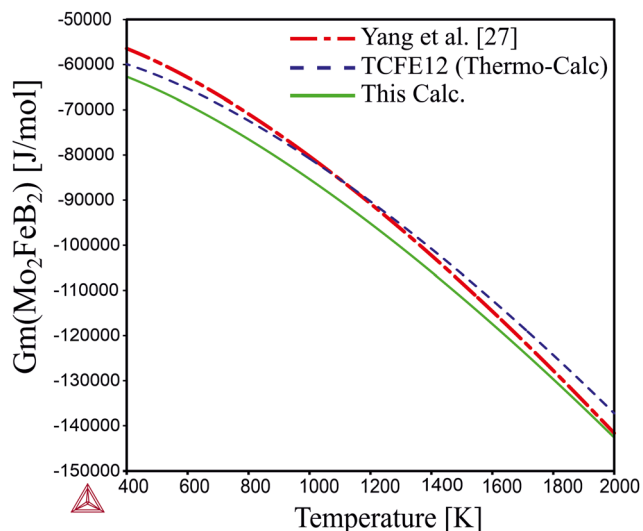


Fig. 10 Temperature dependence of the Gibbs energy calculated using optimized parameters and parameters given in TCFE12 thermodynamic database and the work.²⁷

Gibbs energies of Mo_2FeB_2 and Mo_2MnB_2 were as follows:

$$G_{\text{M:Mo:B}}^{\text{Mo}_2\text{MB}_2} = -54850.033 - 19.202T + 2.814T \ln T + 0.2G_{\text{Fe}}^{\text{SER}} + 0.4G_{\text{Mo}}^{\text{SER}} + 0.4G_{\text{B}}^{\text{SER}} \quad (10)$$

$$G_{\text{M:Mo:B}}^{\text{Mo}_2\text{MB}_2} = -43794.911 - 20.621T + 3.13T \ln T + 0.2G_{\text{Mn}}^{\text{SER}} + 0.4G_{\text{Mo}}^{\text{SER}} + 0.4G_{\text{B}}^{\text{SER}} \quad (11)$$

Comparing the calculated data with those obtained from experimental alloys in ref. 22 and from the TCFE12

Thermo-Calc database (Fig. 10), it can be concluded that the Gibbs energy values agree well with both references. Therefore, it is expected that the free energy values for Mo_2MnB_2 obtained by the same methodology would accurately reflect experimental data.

The formation energy of $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions was determined by assuming that Mn atoms replace Fe atoms in the sublattice of Mo_2FeB_2 . This assumption was based on the similarities in crystal structures, electronic, and phonon properties of the end-member ternary borides. Previous studies on coatings based on Mo_2FeB_2 have also suggested that Mn has a significant influence on the size, shape, and mechanical properties of boride phase grains. The hypothetical solid solution can be represented by the generalized CEF formula $(\text{Fe,Mn})_{0.2}^{\text{I}}(\text{Mo})_{0.4}^{\text{II}}(\text{B})_{0.4}^{\text{III}}$, and its analysis was carried out using the CE method with the ATAT code.²⁶

For the analysis, 121 clusters generated by the MIT *Ab initio* phase stability (MAPS) code⁵¹ were used to simulate different configurations of disordered solid solutions with varying Fe and Mn ratios. Energy calculations were performed for a ferromagnetic state, which was initially set by the spin polarization of electrons. Analysis of the calculated energies (Fig. 11) shows that certain structures with low Mn content exhibit positive mixing energy values, indicating their instability under the given conditions (0 K). In contrast, the most stable configurations are characterized by high concentrations of dissolved Mn (above 50 at%). The data approximation using parameters describing effective cluster interactions (ECI) demonstrates sufficient accuracy to extrapolate the data to practically significant temperature ranges.

Using the Monte Carlo simulations of lattice models as implemented in emc2 code at the constant temperatures of

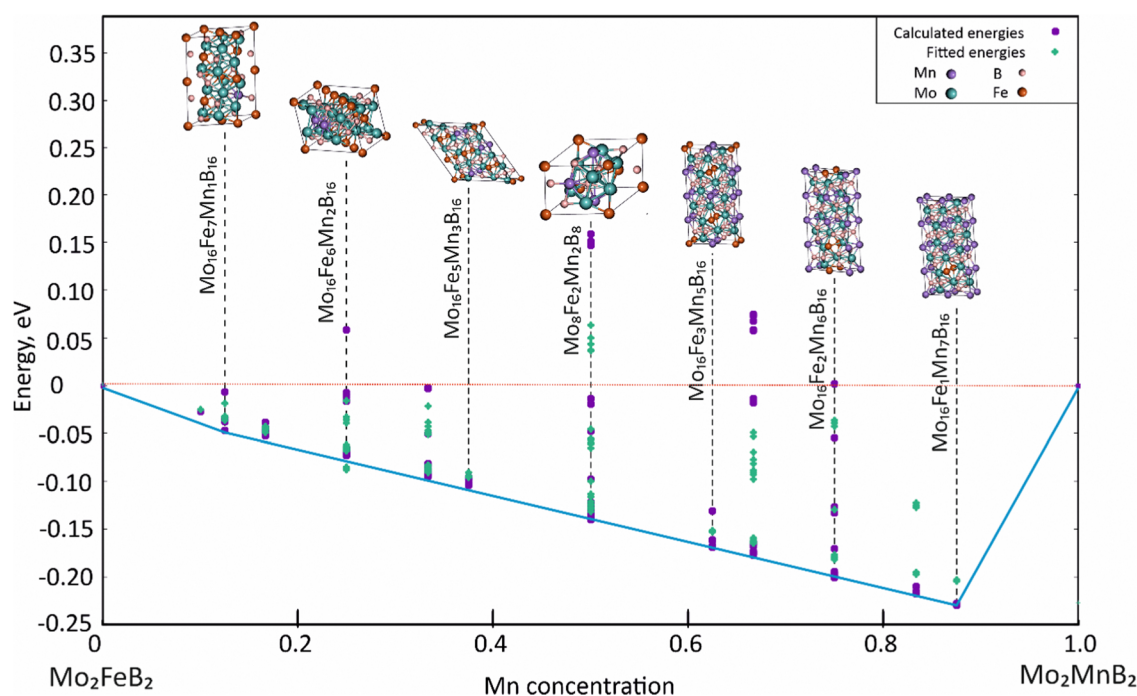


Fig. 11 Concentration dependence of pseudo-binary mixing energy for $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions calculated using CE with ATAT and VASP.

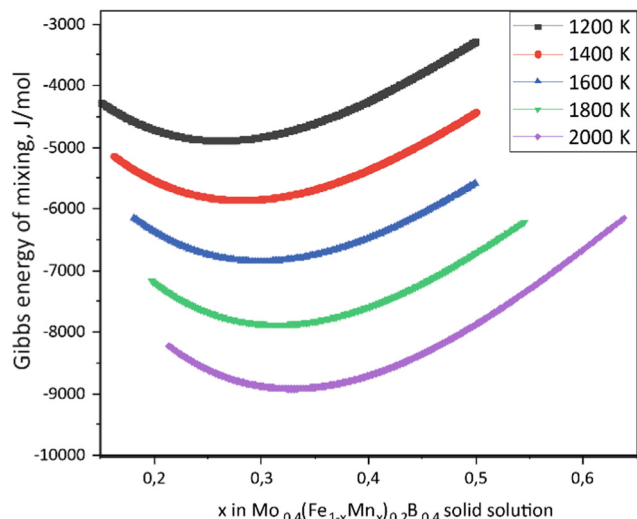


Fig. 12 Concentration dependencies of the Gibbs energy of mixing for $\text{Mo}_{0.4}(\text{Fe}_{1-x}\text{Mn}_x)_{0.2}\text{B}_{0.4}$ solid solutions calculated using the emc2 code with ATAT for different temperatures.

1200, 1400, 1600, 1800 and 2000 K the concentration dependencies of the Gibbs energy of mixing were evaluated (Fig. 12). Next, considering the description of excess Gibbs energy in the form of (eqn (12)) with the representation of the interaction parameter $L = V_1 + V_2 \times T$ as temperature dependent with two optimizing variables (V) and performing optimization procedure with PARROT module the resulting interaction parameter can be described by (eqn (13)). The negative values of optimized variables V_1 and V_2 indicate high negative value of excess energy of mixing, which increase while temperature increases.

$${}^{ex}G^{\text{Mo}_2\text{MeB}_2} = x_{\text{Fe}}^I x_{\text{Mn}}^I x_{\text{Mo}}^{II} L^{\text{Mo}_2\text{MeB}_2}_{\text{Fe,Mn:Mo:B}} \quad (12)$$

$$L^{\text{Mo}_2\text{MeB}_2}_{\text{Fe,Mn:Mo:B}} = -11710.385 - 13.203 \times T \quad (13)$$

To validate the obtained thermodynamic parameters described by eqn (10), (11) and (13), they were integrated with thermodynamic parameters from,^{52,53} and others.⁵⁴ These

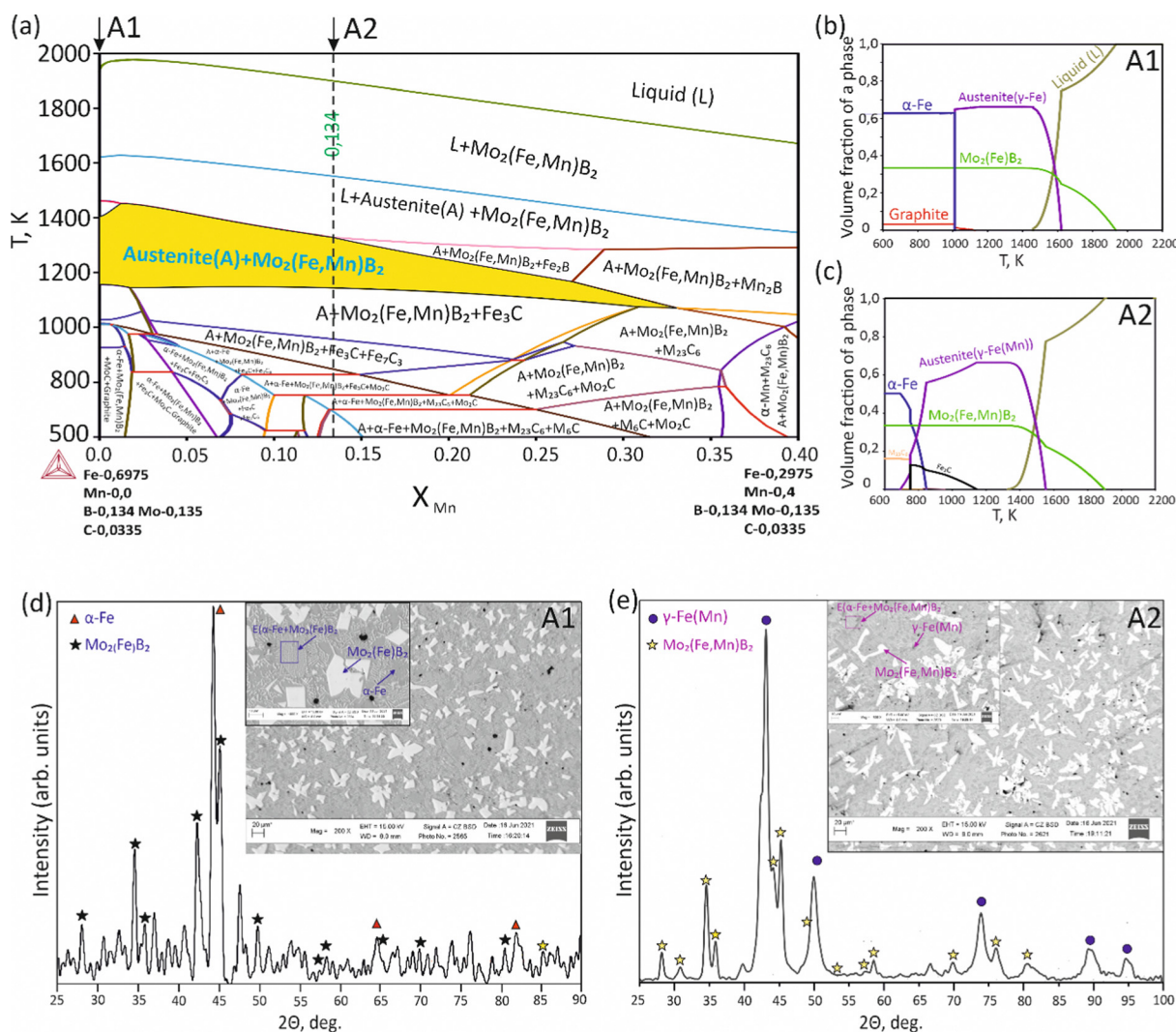


Fig. 13 Isoleth section of the Fe-Mn-Mo-B-C system corresponding to the $\text{Fe}_{0.6975-x}\text{Mn}_x\text{Mo}_{0.135}\text{B}_{0.134}\text{C}_{0.0335}$ alloys (a), solidification paths of $\text{Fe}_{0.6975}\text{Mo}_{0.135}\text{B}_{0.134}\text{C}_{0.0335}$ (A1) (b) and $\text{Fe}_{0.5635}\text{Mn}_{0.134}\text{Mo}_{0.135}\text{B}_{0.134}\text{C}_{0.0335}$ (c) and XRD and SEM results for the A1 (d) and A2 (e) alloys.

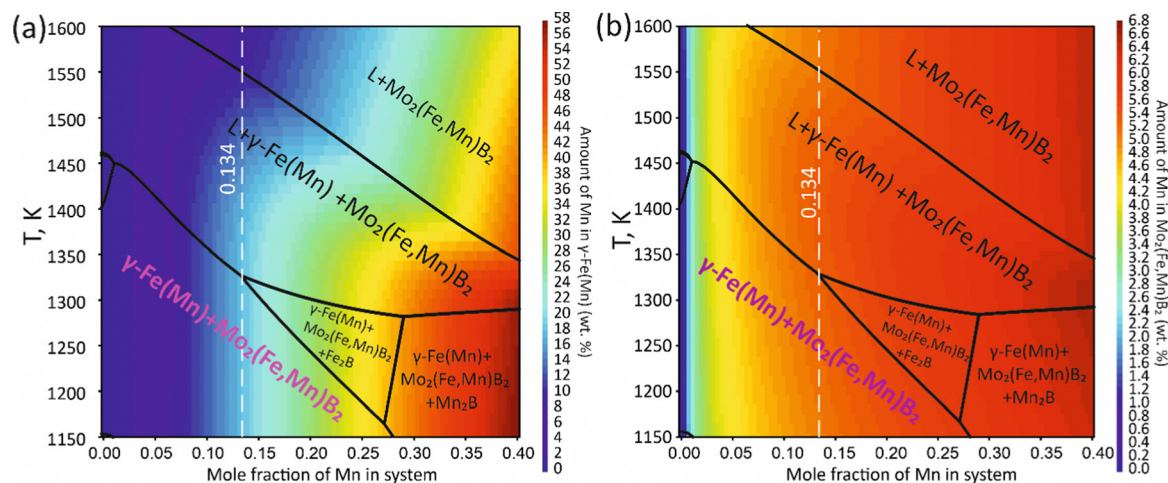


Fig. 14 Mapping of the calculated equilibrium Mn concentration in the γ -Fe(Mn) (a) and $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ phases (b).

parameters cover the entire Fe–Mn–Mo–B–C system within typical ranges for Fe-based hardfacing alloys and were compiled into a TDB file. The resulting CALPHAD calculations allow for the construction of an isopleth illustrating the effect of Mn on the phase composition of the $\text{Fe}_{0.6975-x}\text{Mn}_x\text{Mo}_{0.135}\text{B}_{0.134}\text{C}_{0.0335}$ alloys (Fig. 13(a)). As is shown in the figure, increasing the Mn content in the alloys up to 1.2 at% leads to a slight (40 K) increase in the melting temperature. However, within the concentration range of 1.2 to 40 at%, a nearly linear decrease (ranging from ~ 1975 to ~ 1675 K) in the liquidus temperature is observed. Meanwhile, the primary crystallization temperature interval remains practically unchanged at around 350 K and is represented by a two-phase region: liquid (L) + $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$, which is adjacent to a three-phase region: L + Mo_2FeB_2 + austenite (γ -Fe(Mn)), indicating eutectic formation. Below the solidus line, three regions are identified

where γ -Fe(Mn) coexists with boride phases. Among these regions, the wide two-phase region: γ -Fe(Mn) + Mo_2FeB_2 is particularly important from a practical standpoint and exists within the concentration range up to approximately 32 at%. Regions involving the primary crystallization of Fe_2B and Mn_2B are undesirable due to the high brittleness of such borides. Therefore, an optimal concentration is approximately 13.4 at% Mn. Below the temperature of ~ 1000 K, γ -Fe(Mn) and α -Fe coexist with cementite (Fe_3C) and metastable carbide phases of M_xC_y type. The solidification process of the Mn-free alloy (A1) (Fig. 13(b)) comprises three significant stages: primary crystallization of Mo_2FeB_2 , formation of the eutectic (γ -Fe + Mo_2FeB_2) and subsequent γ -Fe \rightarrow α -Fe transformation. The microstructural analysis of the corresponding alloy (Fig. 13(d)) reveals primary Mo_2FeB_2 as uniformly distributed grains with a cuboidal shape, which corresponds to the tetragonal structure of this

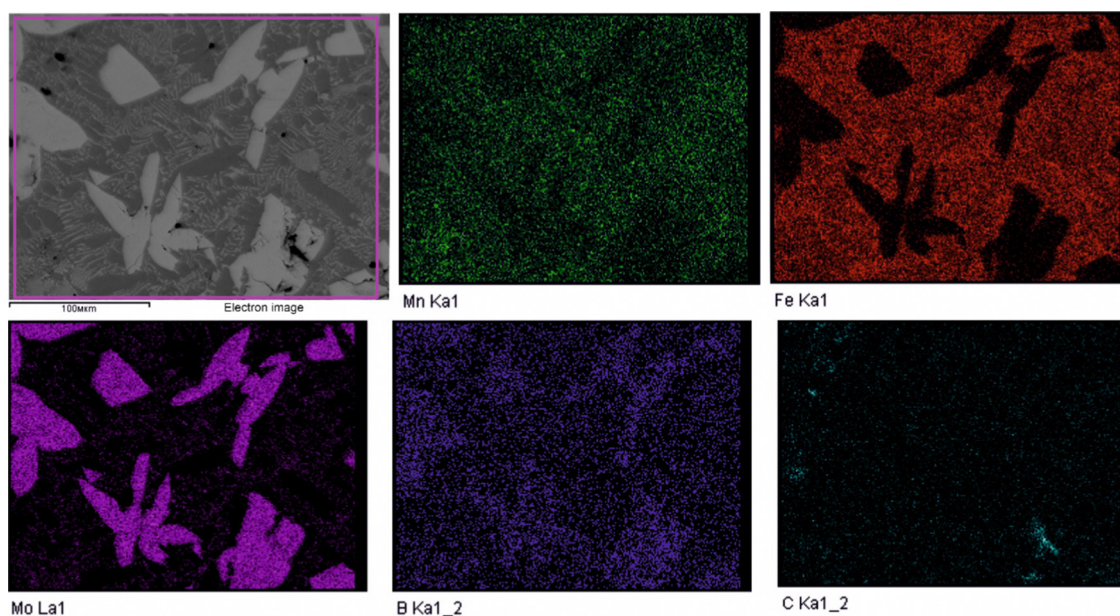


Fig. 15 EDS elemental maps of the experimental alloy A2 showing distribution of the Mn, Fe, Mo, B and C.



phase. In addition, a rod-like eutectic (α -Fe + Mo_2FeB_2) and wide regions of α -Fe are observed, showing agreement with the predictions derived from thermodynamic calculations. The solidification behavior of the Mn-enriched alloy (A2) (Fig. 13(c)) is similar up to ~ 1200 K but exhibits a broader temperature range for the existence of γ -Fe(Mn) due to its stabilization by the effect of Mn, as commonly observed in austenitic high-Mn (Hadfield) steels.⁵⁵ The microstructural investigations (Fig. 13(e)) accurately reproduce the calculated phase composition, revealing $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ grains with a faceted angular shape and a lamellar eutectic structure (γ -Fe(Mn) + $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$). Consequently, the presence of Mn significantly impacts the phase morphology, leading to more pronounced symmetric branching of the boride crystals. This phenomenon can be attributed to the preferential arrangement of Mn along specific crystallographic planes within the $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solution, which results in a reduction in its surface energy and promotes growth in the appropriate directions. The XRD results show that alloy A1 consists of α -Fe and Mo_2FeB_2 (Fig. 13(d)), with lattice parameters $a = 5.807$ Å and $c = 3.142$ Å, which are close to those of the pure compound,²⁴ while alloy A2 consists of γ -Fe and Mo_2FeB_2 (Fig. 13(e)), with changed lattice parameters ($a = 3.783$ Å and $c = 3.152$ Å), indicating the partial substitution of the Fe by Mn. The presence of γ -Fe at room temperature in alloy A2 suggests that the diffusion processes responsible for the formation of α -Fe cementite and other metastable carbide phases are inhibited below ~ 1150 K. As a result, the phase composition at room temperature is identical to that calculated within the temperature range of 1150–1325 K. To investigate the equilibrium Mn contents in the key phases (γ -Fe and $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$), concentration mapping was performed within the high-temperature regions (Fig. 14). According to the calculation results, the equilibrium concentration of Mn in Mo_2FeB_2 is approximately 5 wt%, corresponding to a solid solution with the formula unit $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$. The Mn concentration in γ -Fe is around 18 wt%, which is a typical content for impact-resistant TWIP steels.⁵ The lattice parameters of the relaxed $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$ solid solution models, calculated using VCA ($a = 5.7712$ Å and $c = 3.1517$ Å) and for the cluster with the lowest energy determined by CE ($a = 5.7707$ Å and $c = 3.1563$ Å), are in good agreement with those obtained experimentally by XRD. Elemental EDS mapping (Fig. 15) for the alloy A2 reveals that Mo and B are primarily concentrated in regions corresponding to the $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$ phase, while Fe is generally found in the intergranular space corresponding to the eutectic. Mn is evenly redistributed between the eutectic and boride phases, indicating its solubility in both the metal and ceramic boride phases. The local EDS analysis in the central regions of the boride grains shows the presence of Mn in amounts of approximately 3–5 wt%. The primary locations of carbon are the pores and other insignificant microstructural defects.

3.4 Mechanical properties

To evaluate the mechanical properties of the hypothetical $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions resulting from CE analysis, calculations were performed for the most stable clusters with an increasing step of $x = 0.125$ (Fig. 11) using the finite

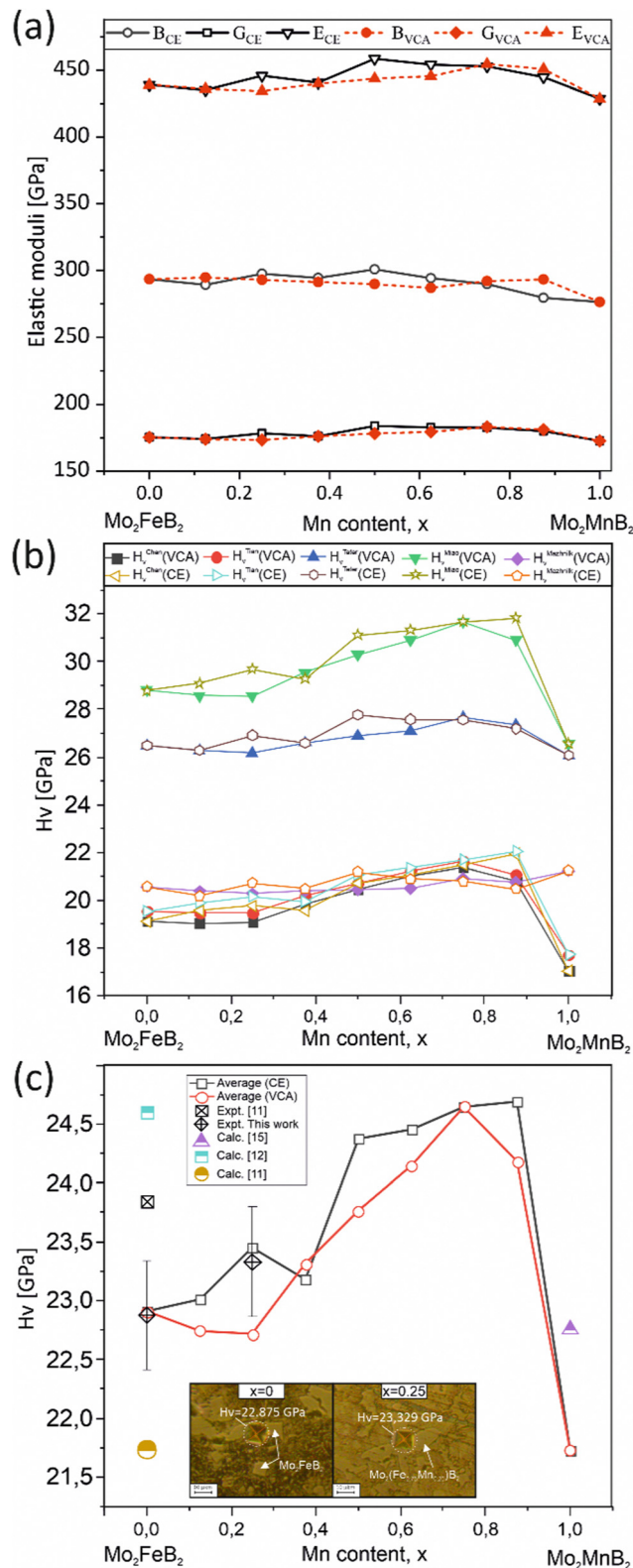


Fig. 16 Mechanical properties of the most stable structures obtained using CE, along with the corresponding structures calculated using the VCA: elastic moduli (a), H_v calculated using different formalisms (eqn (1)–(5) (b), and the averaged H_v value compared to measured values and data from ref. 11, 12 and 15 (c).



differences approach as implemented in VASP to determine the elastic tensor. Additionally, the elastic tensors for the corresponding virtual crystal approximation (VCA) models were computed, assuming averaged Mn and Fe pseudopotentials with different weights in the Fe site of Mo_2FeB_2 . The derived elastic moduli, namely B , G , and E , were then compared between the CE and VCA models (Fig. 16(a)). The results demonstrate sufficient agreement between the two techniques, showing a consistent trend of increasing moduli with increasing Mn content up to 50–75%, followed by a subsequent decrease. Considering that the VCA approach²¹ is suitable for modeling disordered alloys with similar structural and electronic properties and small deviations from Vegard's law, the agreement between CE and VCA results indicates the existence of stable substitutional solid solutions in the Mo_2FeB_2 – Mo_2MnB_2 system. Furthermore, all analyzed structures, including clusters

and VCA models, meet the mechanical stability criteria derived from elastic constants according to ref. 56. Next, the Mn concentration-dependent H_V values were calculated for the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions using different formalisms represented by eqn (1)–(5) (Fig. 16(b)). As can be seen from the figure, the H_V values calculated for the clusters generated by CE are close to those modeled by VCA. The main differences observed between H_V calculated using different models are that H_V^{Teter} and H_V^{Miao} exhibit the highest values in the range of 26–32 GPa, while the values of H_V^{Chen} , H_V^{Tian} , and H_V^{Mazhnik} are close to each other and are in a range of 17–22 GPa. It is worth noting that all calculated H_V values, except for H_V^{Mazhnik} , show similar trends of increasing as the Mn content increases in the range of $x = 0$ to $x = 0.75$, with a further decrease at $x > 0.75$. The averaged H_V values (Fig. 16(c)) show small differences with respect to the method (CE or VCA) used for the solid solutions

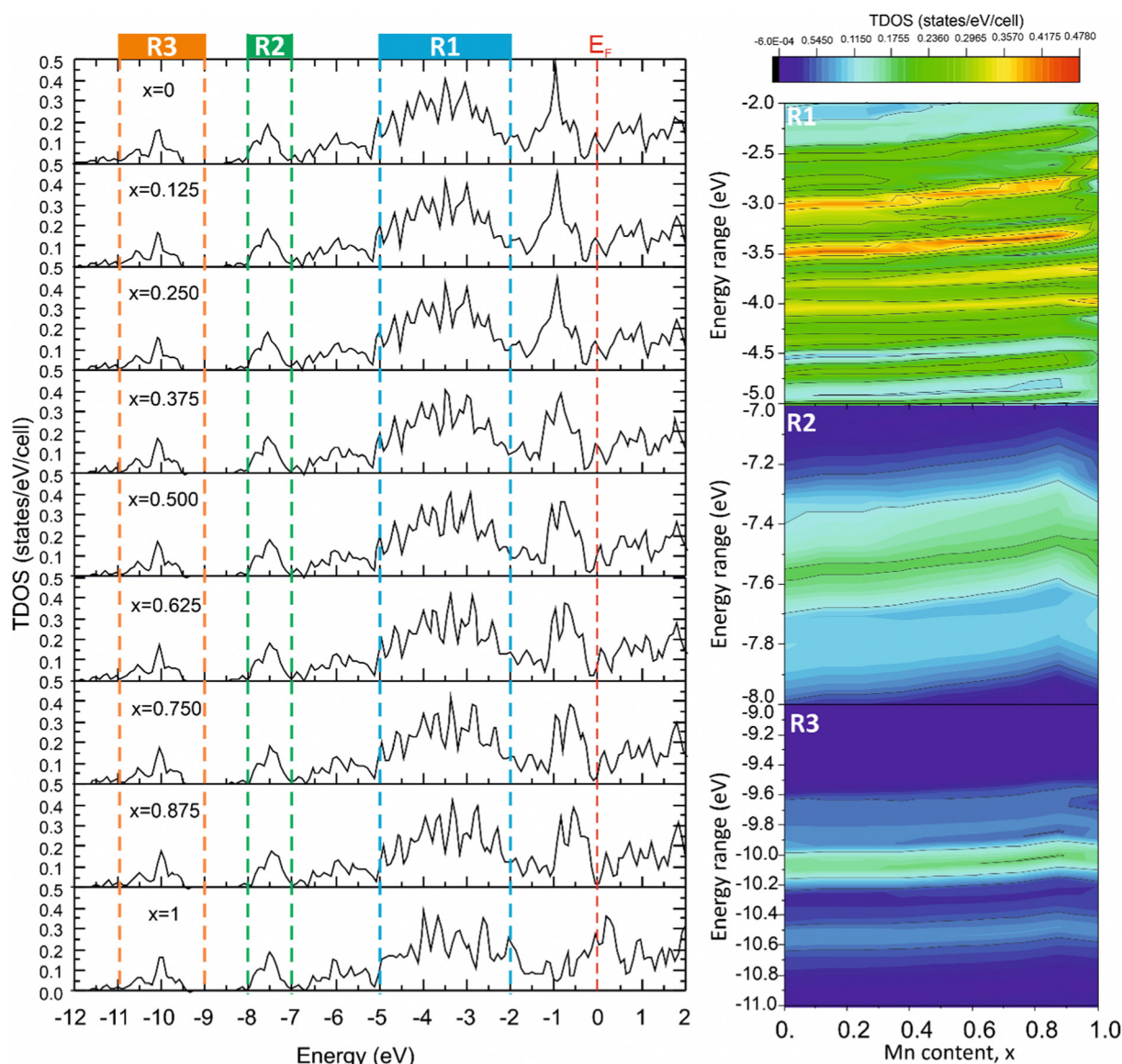


Fig. 17 The total DOS of the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions calculated using VCA approach for different Mn contents (x).



models, and the calculated values are in good agreement with the results of the performed experimental measurements for the key alloys ($x = 0$ and $x = 0.25$), experimental data from,¹¹ and also closely match the calculated data from^{11,12,15} when considering mean values of H_V .

The significant correlation effects between mechanical properties and the TDOS peak values, corresponding to the energy ranges associated with different types of chemical bonding, were observed for the solid solutions modeled using the VCA approach in ref. 57,58. To reveal such effects in the $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions, the TDOSs for different Mn contents were calculated using the VCA approach (Fig. 17) and analyzed for the energy ranges corresponding to different bonding types (Fig. 5). This analysis was performed to include contributions resulting from pd hybridization for the Fe(Mn)–B (R1, Fig. 17) and Mo–B (R2, Fig. 17) pairs, as well as sp hybridization for B–B pairs (R3, Fig. 17), by mapping TDOS values as dependent on Mn content and energy range. The tendency of the TDOS values to change in relation to the calculated H_V values with increasing Mn content was observed for all energy ranges, especially for the energy range corresponding to the Mo–B bonding. In this range, the TDOS values exhibited an increase up to $x = 0.75$ and then decreased, similar to the observed behavior in the calculated H_V values.

4. Conclusions

This study reports a DFT investigation of the thermodynamic, electronic, vibrational, and mechanical properties of the ternary borides Mo_2FeB_2 and Mo_2MnB_2 , as well as the thermodynamic and mechanical properties of the $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions. The formation energies of Mo_2FeB_2 and Mo_2MnB_2 , determined using correction terms based on fitting of the experimental and DFT formation energies for typical molybdenum borides with varying stoichiometry, demonstrate the significant stability of these compounds over a wide temperature range. Phonon calculations were performed to assess the vibrational properties, confirming the absence of imaginary frequencies in both compounds. Furthermore, analysis of the elastic tensors revealed that the relationships between their components satisfied the structure-dependent criteria for mechanical stability, suggesting that both Mo_2FeB_2 and Mo_2MnB_2 are stable compounds. A noticeable difference between Mo_2FeB_2 and Mo_2MnB_2 was observed in their electronic structure, with Mo_2MnB_2 exhibiting slightly stronger metallic characteristics and a predominant electron-like conductivity. This finding is supported by the comparative analysis of the DOS and the topology of the Fermi surfaces.

Through the systematic examination of energy levels in clusters representing different configurations of disordered $\text{Mo}_2(\text{Fe}_{1-x}\text{Mn}_x)\text{B}_2$ solid solutions using the CE method, it was possible to assess the concentration-dependent excess enthalpy of mixing. Subsequently, these results were extrapolated to cover a temperature range from 0 to 2000 K, allowing optimization of thermodynamic parameters. The resulting

temperature-dependent interaction parameter, denoted as $L(\text{Mo}_2\text{MeB}_2, \text{Fe,Mn:Mo:B}; 0)$ according to CALPHAD notation, showed negative values, indicating a favorable tendency for the formation of $\text{Mo}_2(\text{Fe,Mn})\text{B}_2$ solid solutions in Fe-based alloying systems. The evaluation of mechanical properties using both the CE and VCA approaches revealed the positive influence of Mn on the G -modulus and H_V , which are key factors contributing to enhanced resistance against abrasive wear by micro-cutting. The calculated mechanical properties indicated that the most favorable solid solutions within the Mo_2FeB_2 – Mo_2MnB_2 system, were found in the compositional range of $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$ to $\text{Mo}_2(\text{Fe}_{0.25}\text{Mn}_{0.75})\text{B}_2$. The results of thermodynamic modeling of the alloy system corresponding to the high-manganese steel alloyed with equimolar amounts of Mo and B show that a stable phase within a concentration range sufficient for alloy development is close to $\text{Mo}_2(\text{Fe}_{0.75}\text{Mn}_{0.25})\text{B}_2$ solid solution. These findings are consistent with the experimental observations obtained using SEM, XRD, and EDS techniques for the corresponding alloy. Earth-abundant metal ternary borides solid solutions have the potential to be developed into advanced wear-resistant materials that could help reduce the reliance on tungsten-based materials.

Author contributions

Pavlo Prysyazhnyuk: conceptualization, data curation, formal analysis, methodology, software, validation, visualization, writing – original draft. Devis Di Tommaso: methodology, project administration, resources, software, supervision, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research utilized Queen Mary's Apocrita HPC facility, supported by QMUL Research-IT. <https://doi.org/10.5281/zenodo.438045>. We acknowledge the assistance of the Research Software Engineering team in ITS Research at Queen Mary University of London and we extend our sincere appreciation to Tom Bradford for his exceptional support and expert guidance. The authors would like to express their gratitude to Thermo-Calc Software AB for their generous support and provision of high-quality software and special thanks are extended to Therese Gustafsson for her valuable consultations. The authors thank the SCM (Software for Chemistry and Materials) team for their support and provision of a relevant set of high-quality software tools for this study. Special appreciation is extended to Mrs. Sorana Burcusel for her valuable technical assistance in maintaining the modules of the Amsterdam Modeling Suite (AMS). The first author acknowledges the Ministry of Science and Education of Ukraine for support (project "The development of materials for depositing



impact- and wear-resistant coatings in a high-manganese steel – refractory compounds system through electric arc hard-facing” for 2023–2025, state reg. no. 0123U101858) and the Armed Forces of Ukraine for providing the necessary security measures to carry out this work.

References

- 1 L. Tang, P. Wang, Z. Ma, S. Pauliuk, W.-Q. Chen, T. Dai and Z. Lin, *J. Ind. Ecol.*, 2023, **27**, 323–335.
- 2 M. Bembenek, P. Prysyazhnyuk, T. Shihab, R. Machnik, O. Ivanov and L. Ropyak, *Materials*, 2022, **15**.
- 3 Y. Shen, W. Xie, B. Sun, Y. Liu, Y. Li, Z. Cao, Y. Jian and Z. Huang, *Materials*, 2022, **15**, 6729.
- 4 H. Xu, J. Sun, J. Jin, J. Song and C. Wang, *Materials*, 2020, **14**, 46.
- 5 A. Saeed-Akbari, L. Mosecker, A. Schwedt and W. Bleck, *Metall. Mater. Trans. A*, 2011, **43**, 1688–1704.
- 6 P. Xuming, Z. Yong, W. Shaogang and W. Qiuhong, *Int. J. Refract. Met. Hard Mater.*, 2009, **27**, 777–780.
- 7 H. Yu, Y. Zheng, W. Liu, J. Zheng and W. Xiong, *Int. J. Refract. Met. Hard Mater.*, 2010, **28**, 286–290.
- 8 F. Yang, Y. Wu, J. Han and J. Meng, *J. Alloys Compd.*, 2016, **665**, 373–380.
- 9 B. Wang, Y. Liu, J.-W. Ye and J. Wang, *Comput. Mater. Sci.*, 2013, **70**, 133–139.
- 10 B. Wang, B. Ma, W. Song, Z. Fu and Z. Lu, *R. Soc. Open Sci.*, 2018, **5**, 172247.
- 11 Z. Shi, H. Yin, Z. Xu, T. Zhang, G. Yang, Q. Zheng, R. S. Rao, J. Yang, F. Gao, M. Wu and others, *Intermetallics*, 2019, **114**, 106573.
- 12 T. He, Y. Jiang, R. Zhou and J. Feng, *J. Appl. Phys.*, 2015, **118**, 075902.
- 13 X.-Q. Chen, H. Niu, D. Li and Y. Li, *Intermetallics*, 2011, **19**, 1275–1281.
- 14 V. S. Telegus and Y. B. Kuz'ma, *Sov. Powder Metall. Met. Ceram.*, 1971, **10**, 52–56.
- 15 G. Yang, H. Yin, Z. Xu, T. Zhang, J. Yang, F. Gao, Q. Zheng and X. Qu, *J. Alloys Compd.*, 2019, **791**, 761–772.
- 16 Y. H. Lin, C. C. Tong, Y. Pan, W. Y. Liu and A. Singh, *Mod. Phys. Lett. B*, 2017, **31**, 1750138.
- 17 A. Zunger, S.-H. Wei, L. G. Ferreira and J. E. Bernard, *Phys. Rev. Lett.*, 1990, **65**, 353–356.
- 18 A. van de Walle, R. Sun, Q.-J. Hong and S. Kadkhodaei, *Calphad*, 2017, **58**, 70.
- 19 A. van de Walle, M. D. Asta and G. Ceder, *Calphad*, 2002, **26**, 539–553.
- 20 A. van de Walle, P. Tiwary, M. de Jong, D. L. Olmsted, M. Asta, A. Dick, D. Shin, Y. Wang, L.-Q. Chen and Z.-K. Liu, *Calphad*, 2013, **42**, 13–18.
- 21 L. Bellaiche and D. Vanderbilt, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2000, **61**, 7877–7882.
- 22 L. Vegard, *Z. Phys.*, 1921, **5**, 17–26.
- 23 M. Hillert, *J. Alloys Compd.*, 2001, **320**, 161–176.
- 24 M. Lukachuk and R. Pöttgen, *Z. Kristallogr. – Cryst. Mater.*, 2003, **218**, 767–787.
- 25 W. Rieger, H. Nowotny and F. Benesovsky, *Monatsh. Chem.*, 1964, **95**, 1502–1503.
- 26 V. Telegus, *Khim. Khim. Tekhnol.*, 1970, **1**, 96–100.
- 27 X. O. Yang, F. Yin, J. Hu, M. Zhao and Y. Liu, *Calphad*, 2017, **59**, 189–198.
- 28 A. Van De Walle and M. Asta, *Modell. Simul. Mater. Sci. Eng.*, 2002, **10**, 521.
- 29 J.-O. Andersson, T. Helander, L. Höglund, P. Shi and B. Sundman, *Calphad*, 2002, **26**, 273–312.
- 30 J. Hafner and G. Kresse, *Properties of Complex Inorganic Solids*, Springer, US, 1997, pp.69–82.
- 31 J. P. Perdew, J. Chevary, S. Vosko, K. A. Jackson, M. R. Pederson, D. Singh and C. Fiolhais, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1993, **48**, 4978.
- 32 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 33 H. J. Monkhorst and J. D. Pack, *Phys. Rev. B: Solid State*, 1976, **13**, 5188.
- 34 Y. Le Page and P. Saxe, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2002, **65**, 104104.
- 35 V. Wang, N. Xu, J.-C. Liu, G. Tang and W.-T. Geng, *Comput. Phys. Commun.*, 2021, **267**, 108033.
- 36 Y. Tian, B. Xu and Z. Zhao, *Int. J. Refract. Met. Hard Mater.*, 2012, **33**, 93–106.
- 37 D. M. Teter, *MRS Bull.*, 1998, **23**, 22–27.
- 38 E. Mazhnik and A. R. Oganov, *J. Appl. Phys.*, 2019, **126**, 125109.
- 39 N. Miao, B. Sa, J. Zhou and Z. Sun, *Comput. Mater. Sci.*, 2011, **50**, 1559–1566.
- 40 M. Kabir, M. Hossain, M. Ali, M. Uddin, M. Ali, M. Hasan, A. A. Islam and S. Naqib, *Results Phys.*, 2023, 106701.
- 41 A. Togo and I. Tanaka, *Scr. Mater.*, 2015, **108**, 1–5.
- 42 P. Giannozzi, S. de Gironcoli, P. Pavone and S. Baroni, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1991, **43**, 7231–7242.
- 43 P. Vinet, J. H. Rose, J. Ferrante and J. R. Smith, *J. Phys.: Condens. Matter*, 1989, **1**, 1941–1963.
- 44 G. te Velde and E. J. Baerends, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1991, **44**, 7888–7903.
- 45 P. Prysyazhnyuk, L. Shlapak, O. Ivanov, S. Korniy, L. Lutsak, M. Burda, I. Hnatenko and V. Yurkiv, *East.-Eur. J. Enterp. Technol.*, 2020, **4**, 46–51.
- 46 H. Putz and K. Brandenburg, Match! Crystal Impact, Kreuzherrenstr. 102, 53227 Bonn, Germany.
- 47 G. G. C. Peterson and J. Brögh, *J. Phys. Energy*, 2021, **3**, 022002.
- 48 V. Stevanović, S. Lany, X. Zhang and A. Zunger, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2012, **85**, 115104.
- 49 E. Viswanathan, M. Sundareswari, D. S. Jayalakshmi and M. Manjula, *Comput. Mater. Sci.*, 2019, **157**, 107–120.
- 50 W. Wu, M. S. Scheurer, S. Chatterjee, S. Sachdev, A. Georges and M. Ferrero, *Phys. Rev. X*, 2018, **8**, 021048.
- 51 A. van de Walle and G. Ceder, *J. Phase Equilib.*, 2002, **23**, 348–359.
- 52 B. Hallstedt, A. V. Khvan, B. B. Lindahl, M. Selleby and S. Liu, *Calphad*, 2017, **56**, 49–57.



- 53 X. O. Yang, F. Yin, J. Hu, M. Zhao and Y. Liu, *Calphad*, 2017, **59**, 189–198.
- 54 A. van de Walle, C. Nataraj and Z.-K. Liu, *Calphad*, 2018, **61**, 173–178.
- 55 Y. A. Kryl and P. M. Prysyzhnyuk, *J. Superhard Mater.*, 2013, **35**, 292–297.
- 56 F. Mouhat and F.-X. Coudert, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **90**, 224104.
- 57 M. A. Ali and S. H. Naqib, *RSC Adv.*, 2020, **10**, 31535–31546.
- 58 M. A. Ali, M. M. Hossain, M. A. Hossain, M. T. Nasir, M. M. Uddin, M. Z. Hasan, A. K. M. A. Islam and S. H. Naqib, *J. Alloys Compd.*, 2018, **743**, 146–154.

