

Cite this: *Nanoscale Adv.*, 2023, 5, 1405

Revealing the electronic, optical and photocatalytic properties of PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) heterostructures

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Using DFT, the electronic structure, optical, and photocatalytic properties of PN (P = Ga, Al) and M₂CO₂ (M = Ti, Zr, Hf) monolayers and their PN-M₂CO₂ van der Waals heterostructures (vdWHs) are investigated. Optimized lattice parameters, bond length, bandgap, conduction and valence band edges show the potential of PN (P = Ga, Al) and M₂CO₂ (M = Ti, Zr, Hf) monolayers in photocatalytic applications, and the application of the present approach to combine these monolayers and form vdWHs for efficient electronic, optoelectronic and photocatalytic applications is shown. Based on the same hexagonal symmetry and experimentally achievable lattice mismatch of PN (P = Ga, Al) with M₂CO₂ (M = Ti, Zr, Hf) monolayers, we have fabricated PN-M₂CO₂ vdWHs. Binding energies, interlayer distance and AIMD calculations show the stability of PN-M₂CO₂ vdWHs and demonstrate that these materials can be easily fabricated experimentally. The calculated electronic band structures show that all the PN-M₂CO₂ vdWHs are indirect bandgap semiconductors. Type-II[-I] band alignment is obtained for GaN(AIN)-Ti₂CO₂[GaN(AIN)-Zr₂CO₂ and GaN(AIN)-Hf₂CO₂] vdWHs. PN-Ti₂CO₂ (PN-Zr₂CO₂) vdWHs with a PN(Zr₂CO₂) monolayer have greater potential than a Ti₂CO₂(PN) monolayer, indicating that charge is transfer from the Ti₂CO₂(PN) to PN(Zr₂CO₂) monolayer, while the potential drop separates charge carriers (electron and holes) at the interface. The work function and effective mass of the carriers of PN-M₂CO₂ vdWHs are also calculated and presented. A red (blue) shift is observed in the position of excitonic peaks from AlN to GaN in PN-Ti₂CO₂ and PN-Hf₂CO₂ (PN-Zr₂CO₂) vdWHs, while significant absorption for photon energies above 2 eV for AlN-Zr₂CO₂, GaN-Ti₂CO₂ and PN-Hf₂CO₂, give them good optical profiles. The calculated photocatalytic properties demonstrate that PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs are the best candidates for photocatalytic water splitting.

Received 9th January 2023
Accepted 2nd February 2023

DOI: 10.1039/d3na00017f

rsc.li/nanoscale-advances

1. Introduction

Photocatalytic water splitting is considered as the most suitable way to convert sunlight into chemical energy in the form of hydrogen.¹ Therefore, understanding the mechanism of photocatalytic water splitting and exploring new photocatalysts are of interest for both fundamental research and industrial applications.² Light harvesting, separation, and diffusion of photogenerated carriers (electrons and holes) to the surface, and hydrogen and oxygen evolution reactions are the major steps in photocatalysis. A suitable bandgap³ to strengthen (weaken) electron-hole separation (recombination) time,⁴ and band edge positions (straddling the redox potential of water) of a potential semiconductor photocatalyst, are key in order to achieve the overall water splitting process.⁵

In the case of two-dimensional (2D) materials, a large surface area provides a shorter route for photogenerated electron-hole pairs to flow to the surface and facilitate a redox reaction, hence increasing the lifetime of carriers. Therefore, 2D materials are considered as pioneer candidates for photocatalytic water splitting as compared to their bulk counterparts.⁶ Performance of 2D materials is not limited to photocatalytic water splitting, and is also gaining much attention in both optoelectronics and thermoelectric device applications.^{7,8} In the family of 2D materials, MXenes produced by etching the A layer from the MAX phase,⁹ have a rich surface with functional groups (O, OH, F), responsible for excellent electronic,¹⁰ optical,¹¹ electrochemical,¹² and mechanical properties^{13,14} in catalysis,¹⁵ energy storage,¹⁶ batteries,¹⁷ photocatalytic water splitting¹⁸ and other related fields.¹⁹ Another group of the 2D family, PN (P = Al, Ga) monolayers, are excellent candidates as compared to their bulk counterparts, due to their thin structure, high active surface, high stability, large energy bandgap, blue-shifted photoluminescence peaks, high internal quantum efficiency and transfer of charge from Al/Ga to N.^{20–23} Experimentally, AlN is prepared by CVD, MBE and PVT techniques,²⁴ while GaN is

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prepared by a migration enhanced encapsulated growth technique.²⁵ Structural stability of the PN (P = Al, Ga) monolayer are further confirmed by phonon spectrum calculations.^{26,27} Strain engineering can also be performed to tune the bandgap of PN monolayers without structure distortion.^{28,29} PN (P = Al, Ga) nanowire,^{30,31} nanosheets^{32,33} and nanoribbons,^{34,35} have also been reported with fascinating properties.

The emergence of vdWHs in the 2D field offers a new way to combine different monolayers for novel electronic, optoelectronic, and photocatalytic applications.^{36,37} Some vdWHs based on MXene, AlN, and GaN like MoS₂/MXene,³⁸ TiO₂/MXenes,³⁹ MXene/graphene,⁴⁰ TMDCs/MXenes,⁴¹ SiS/MXenes,⁴² BSe/MXenes,⁴³ AlGaN/GaN,⁴⁴ MoSe₂/GaN,⁴⁵ AlN/GaN,⁴⁶ WS₂/GaN,⁴⁷ ZnO/GaN,⁴⁸ graphene/AlN,⁴⁹ BP/AlN,⁵⁰ and AlN/InSe,⁵¹ have already been fabricated for practical device applications. Electronic structure, optical properties and overall water splitting performance of GaN/Hf₂CO₂ and AlN/Hf₂CO₂ vdWHs have been investigated in ref. 52 and 53. A type-II band alignment for full water splitting is also observed in boron phosphide-blue phosphorene,⁵⁴ C₂N-based type-II heterojunctions⁵⁵ and other 2D materials.⁵⁶

From this aspect, surprisingly no investigation has addressed PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr) vdWHs. Therefore,

based on the same hexagonal symmetry, small lattice mismatch and unique performance of all MXenes of group M₂CO₂ (M = Ti, Zr, Hf) and PN (P = Al, Ga) monolayers, we combine these monolayers, in the form PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs. A detailed study is gained to explore the electronic structure, optical and photocatalytic properties of PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs. Our investigation revealed that all these six heterostructures are promising candidates for photocatalytic and optoelectronic devices.

II. Computational details

We employ DFT⁵⁷ in VASP,⁵⁸ with PBE⁵⁹ and HSE06⁶⁰ functionals for both atomic relaxation and electronic bandstructure calculations. These calculations are performed with a conjugate gradient algorithm to minimize the atomic forces (energy) with a tolerance of 0.01 eV Å⁻¹ (10⁻⁴ eV). A cut-off energy of 500 eV, Γ -centered k -point mesh of $8 \times 8 \times 1$, and a vacuum of 25 Å along the z -direction were used to ensure a negligible interaction between adjacent layers.

AIMD⁶¹ simulations were used to investigate the thermal stabilities of PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs

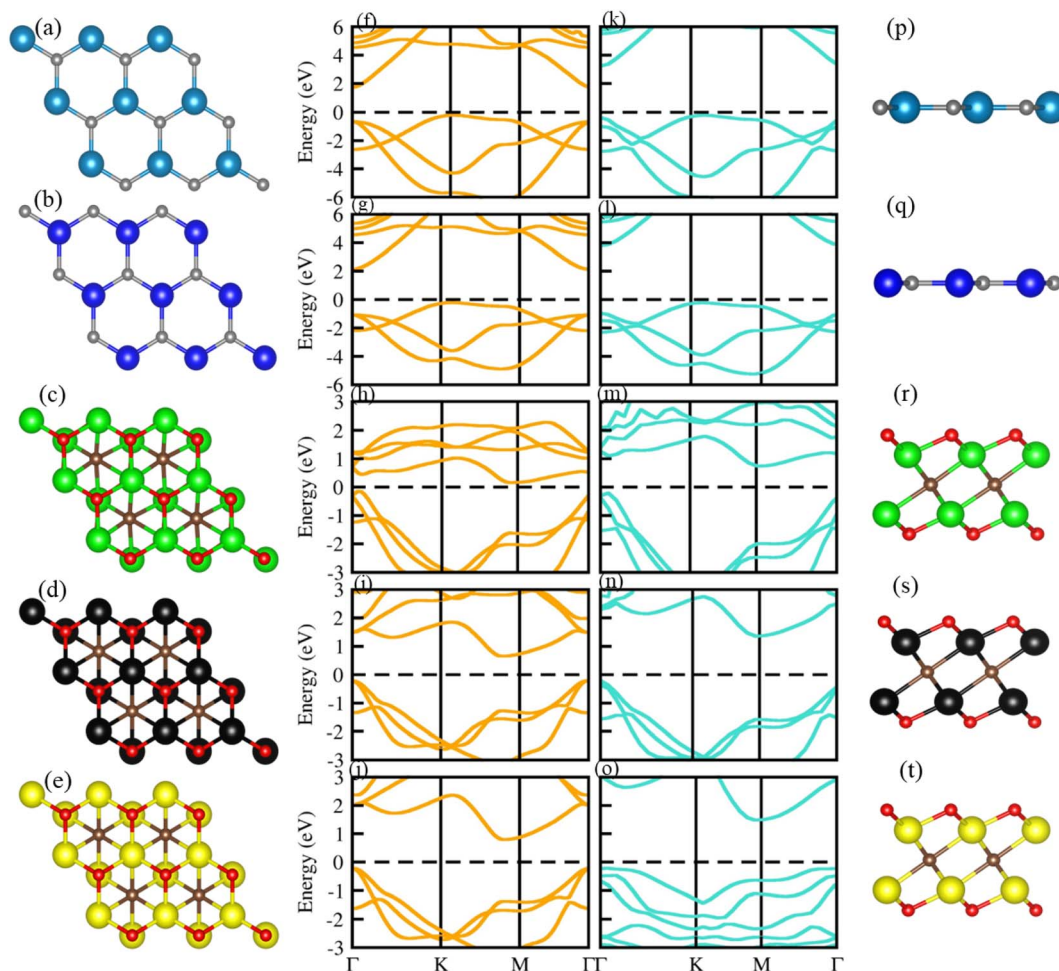


Fig. 1 Geometrical structure (top view in first column), electronic bandstructure (PBE in second column; HSE06 in third column) and side view (fourth column) of GaN (a), (f), (k), (p)), AlN (b), (g), (i), (q)), Ti₂CO₂ (c), (h), (m), (r)), Zr₂CO₂ (d), (j), (n), (s)) and Hf₂CO₂ (e), (l), (o), (t)).



through the nose thermostat algorithm at 300 K. The system was allowed to relax for 12 picoseconds with a time step of 3 femtoseconds.^{62,63}

The Bethe–Salpeter equation (BSE) is also solved in the GW₀ calculation⁶⁴ using the epsilon package⁶⁵ in Quantum Espresso⁶⁶ to obtain $\epsilon_2(\omega)$. Furthermore, using the turbo Lanczos algorithm,⁶⁷ we investigated the absorption coefficient (α) to understand the optical behaviour of these systems in detail.⁶⁸

III. Results and discussion

Optimized geometry (top and side view) and electronic bandstructure in Fig. 1, show that PN (P = Al, N) and M₂CO₂ (M = Ti, Zr, Hf) monolayers are indirect bandgap semiconductors. In the case of the PN (P = Ga, Al) monolayer, the CBM (VBM) lies at the $\Gamma(K)$ -point, while for M₂CO₂ the VBM lies at the Γ -point and the CBM is between the K and M point of the first BZ, consistent with ref. 43 and 69. The photocatalytic response in Fig. 2, shows that Ti₂CO₂ and Zr₂CO₂ cross (fail to cross) the valence (conduction) band edge, while both the PN and Hf₂CO₂ monolayers cross both the conduction and valence band edges (see also Table 1), in agreement with ref. 43. These results show the potential of PN (M = Ga, Al) and M₂CO₂ (M = Ti, Zr, Hf) monolayers in photocatalytic applications and the authenticity of the present approach to combine these monolayers to form vdWHs for efficient electronic, optoelectronic and photocatalytic applications.

Although the lattice mismatch of GaN(AlN) with Ti₂CO₂ is 7.52 (8.45)%, with Zr₂CO₂ is 1.66 (0.90)% and with Hf₂CO₂ is 0.45 (0.30)%, these are experimentally achievable⁷⁰ in the

Table 1 Lattice constant (a in Å), bond length (Ga–N, Al–N, M–O and M–C in Å), band gap (E_g in eV), and conduction and valence band edge potentials (E_{CB} and E_{VB} in eV) for PN (P = Ga, Al) and M₂CO₂ (M = Ti, Zr, Hf) monolayers

Monolayers	GaN	AlN	Ti ₂ CO ₂	Zr ₂ CO ₂	Hf ₂ CO ₂
a	3.255	3.288	3.01	3.31	3.27
Ga–N	1.894	—	—	—	—
Al–N	—	1.807	—	—	—
M–O	—	—	1.97	2.119	2.091
M–C	—	—	2.21	2.359	2.332
E_g (PBE)	1.98	2.35	0.3	0.865	0.99
E_g (HSE06)	3.23	3.98	0.95	1.57	1.69
E_{CB}	−1.0022	−1.6823	0.3398	0.0798	−0.002
E_{VB}	2.1977	2.3176	1.2898	1.6498	1.6923

fabrication of PN-M₂CO₂ vdWHs. But controlling the orientation of layers while fabricating vdWHs using mechanical exfoliation is quite difficult. Moreover, the electronic bandstructure is very sensitive to the stacking of layers. Therefore, we fabricated four different stacking patterns of PN-M₂CO₂ (P = Ga, Al; M = Ti, Zr, Hf) vdWHs, as presented in Fig. 3. In stacking (a), the M(O) atom of the M₂CO₂ layer is placed on the top of Ga/Al(N) atom of the PN layer; in stacking (b), the M(O) atom of the M₂CO₂ layer is placed on top of the N(Ga/Al) atom of the PN layer; in stacking (c), the M(C) atom of the M₂CO₂ layer is placed on top of the N(Ga/Al) of the PN layer; and stacking (d) is the reciprocal of stacking (c). The magnitude of the binding energy ($E_b = E_{PN-M_2CO_2} - E_{M_2CO_2} - E_{PN}$, where $E_{PN-M_2CO_2}$ is the total energy of the vdWHs, $E_{M_2CO_2}$ is the total energy of the isolated M₂CO₂ monolayer, and E_{PN} is the total energy of the isolated PN monolayer) and the interlayer distance, show the disparity between different stacking patterns. The shorter the interlayer distance and the smaller the binding energy, corresponds to a more stable configuration, see Table 2. Therefore, stacking (d) of PN-M₂CO₂ vdWHs is the most favorable stacking configuration. Negative binding energy confirms that formation of PN-M₂CO₂ vdWHs are exothermic.^{71–73} The optimized lattice constants and bond lengths for the most stable stacking configuration are given in Table 3.

Further, the thermal stability of stacking (d) of PN-M₂CO₂ vdWHs, is verified using AMID simulation with $3 \times 3 \times 1$ supercell, (see top view of the structures, before and after heating in Fig. 4). It is clear that after heating for 5 ps at 1 fs time stops at 300 K, all six of the most stable configuration patterns of PN-M₂CO₂ vdWHs show no broken bonds (Fig. 4, column 3), while free energy oscillates slightly (Fig. 4, column 2), and hence confirms the thermal stability of understudy vdWHs. Therefore, stacking (d) is the most stable stacking configuration and is further investigated in detail.

The behavior of the bandstructure and bandgap values are dependent on exchange correlation functionals.⁷⁴ Therefore, we have calculated the bandstructure of PN-M₂CO₂ vdWHs using both PBE and HSE06 functionals, see Fig. 5, with the bandgap values listed in Table 3. One can see that HSE06 band gap values are larger than those of the PBE method, as the PBE approach underestimates the band gap values of semiconductors. For PN-

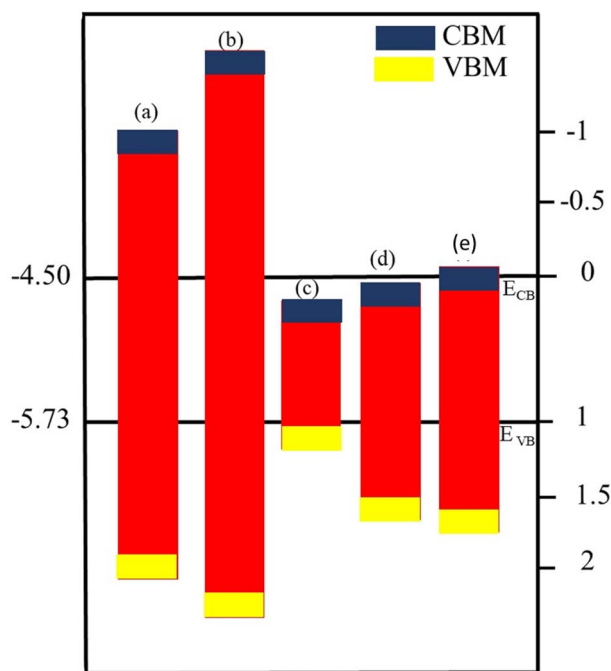


Fig. 2 Band alignment for the valence band (VB) and conduction band (CB) edge of (a) GaN, (b) AlN, (c) Ti₂CO₂, (d) Zr₂CO₂ and (e) Hf₂CO₂ monolayers.



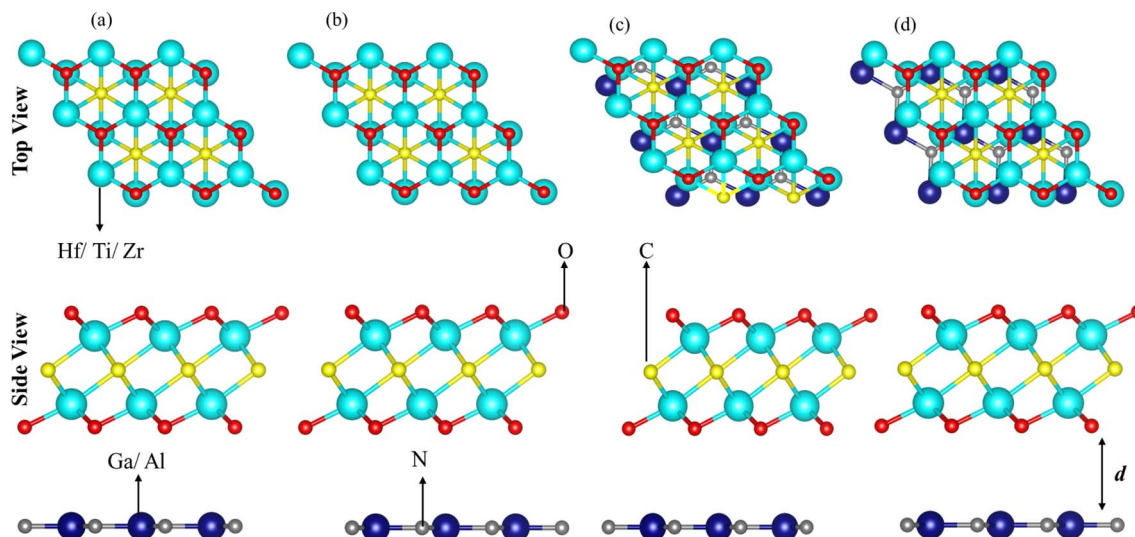


Fig. 3 Possible stacking configurations of PN (P = Ga, Al) and M_2CO_2 (M = Ti, Zr, Hf) vdWHs.

Table 2 Binding energy (E_b in eV) and interlayer distance (d in Å) of the PN (P = Ga, Al)- M_2CO_2 (M = Ti, Zr, Hf) vdWHs in different stacking configurations

Stacking	GaN-Ti ₂ CO ₂	GaN-Zr ₂ CO ₂	GaN-Hf ₂ CO ₂	AlN-Ti ₂ CO ₂	AlN-Zr ₂ CO ₂	AlN-Hf ₂ CO ₂
E_b (a)	-0.808	-0.116	-0.141	-0.739	-0.962	-0.064
d	3.26	3.33	3.37	4.305	3.28	3.289
E_b (b)	-0.818	-0.157	-0.182	-0.324	-0.359	-0.064
d	3.24	3.246	3.24	3.31	3.3	3.39
E_b (c)	-0.813	-0.117	-0.003	-0.055	-0.939	-0.064
d	3.27	3.28	3.29	3.28	3.25	3.29
E_b (d)	-0.833	-0.217	-0.382	-0.884	-1.062	-0.104
d	3.05	3.21	3.19	3.24	3.22	3.23

Table 3 Lattice constant (a in Å), bandgap (E_g in eV), potential difference (ΔV in eV), work function (ϕ in eV), effective mass (m_e^* and m_h^*), conduction and valence band edges (E_{CB} and E_{VB} in eV) of PN- M_2CO_2 (P = Ga, Al; M = Ti, Zr, Hf) vdWHs

Heterostructure	a	E_g -PBE	E_g -HSE	ΔV	ϕ	m_e^*	m_h^*	E_{CB}	E_{VB}
GaN-Ti ₂ CO ₂	3.31	0.14	0.65	0.71	6.46	1.07	2.59	0.35	1.0
GaN-Zr ₂ CO ₂	3.28	0.81	1.62	10.44	7.60	1.30	2.43	-0.11	1.51
GaN-Hf ₂ CO ₂	3.26	0.91	1.69	2.44	5.67	1.20	1.40	-0.15	1.53
AlN-Ti ₂ CO ₂	3.15	0.60	1.15	11.01	6.17	3.44	3.01	0.12	1.27
AlN-Zr ₂ CO ₂	3.30	0.84	1.85	0.41	7.31	1.47	2.31	-0.03	1.46
AlN-Hf ₂ CO ₂	3.28	0.88	1.79	8.02	5.50	1.24	1.54	-0.183	1.60

Ti₂CO₂(GaN-Zr₂CO₂ and PN-Hf₂CO₂) vdWHs, the VBM lies at the K - M (Γ)-point and the CBM at the Γ (K - M)-point of the first BZ. For AlN-Zr₂CO₂, both the VBM and CBM lie at the K - M point of the first BZ. These results show that all the vdWHs are indirect bandgap semiconductors, where electrons must undergo a significant change in momentum for a photon of energy (E_g) to create an electron-hole pair.⁷⁵ In the case of indirect bandgap semiconductors, the electron interacts with the phonon (phonon) to gain (gain or lose) energy (momentum). Hence, the position of the CBM and VBM can be aligned using suitable light according to the bandgap, and a transition is possible at

lower energy using phonons, making them useful for laser applications.⁷⁶⁻⁷⁸

To check the atomic states in the CBM/VBM at the Fermi level and band alignment, we plotted the partial density of states (PDOS) for PN- M_2CO_2 (P = Ga, Al; M = Hf, Ti and Zr) vdWHs, see Fig. 6. In the case of PN-Ti₂CO₂, the VBM (CBM) is from the N_p (Ti_d) state of the PN(Ti₂CO₂) monolayer confirming type-II band alignment. The relative band position allows the mobility of electron (holes) from the conduction (valence) band of Ti₂CO₂(PN) to the conduction (valence) band of the PN(Ti₂CO₂) monolayer, offering a potential technique for charge separation. Hence, PN-Ti₂CO₂ vdWHs are promising candidates



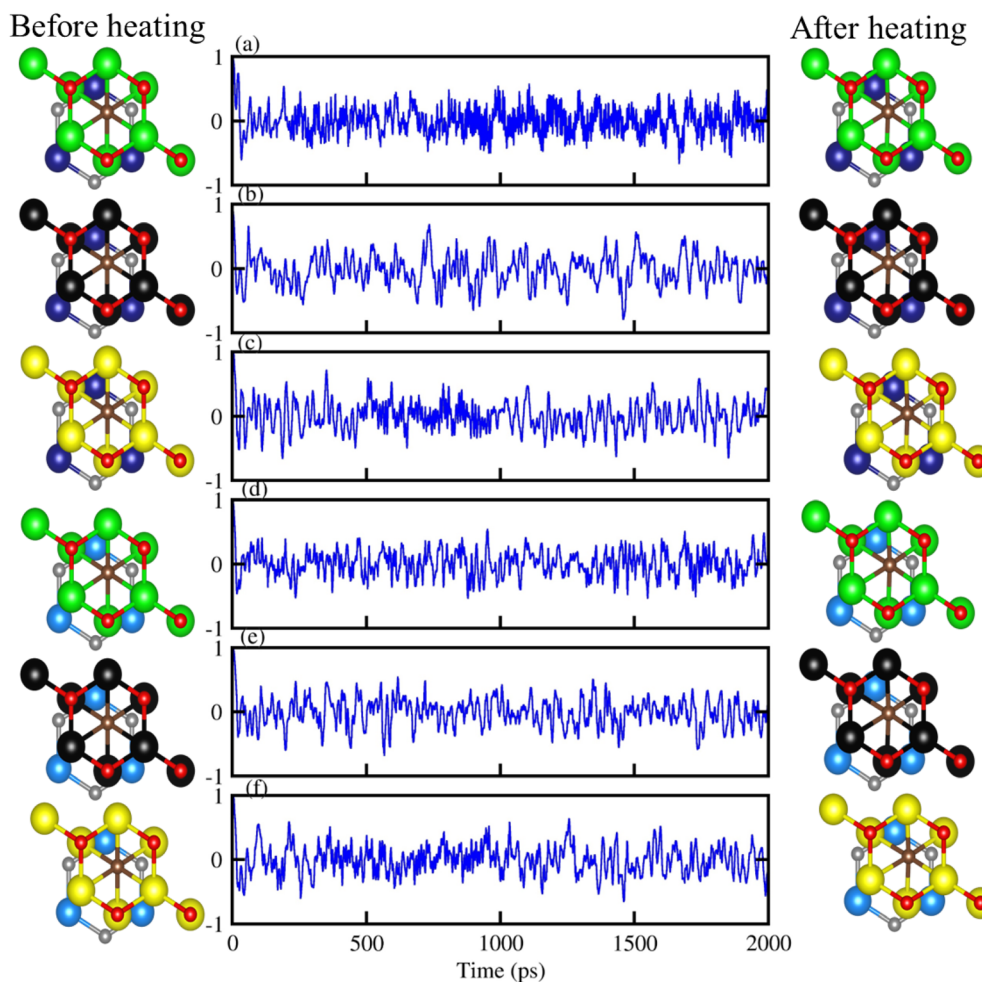


Fig. 4 Geometrical structure before heating (first column), with fluctuating energy (second column) and after heating (third column) of: (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂ vdWHs using AIMD simulation.

for solar energy conversion.⁷⁹ In the case of PN-Zr₂CO₂(Hf₂CO₂), the VBM and CBM is due to the C_p state and the Zr/Hf_d state, indicating type-I band alignment, as the CBM and VBM is from the same monolayer. Both the electron and hole from the PN is transported to Zr₂CO₂(Hf₂CO₂), so would not provide good charge separation. However, since the Fermi level of semiconductors tune the band alignment (for non-intrinsic semiconductors), the carrier (electrons and holes) movement can be hindered in a certain direction (depending on the internal electric field direction); that is, electrons may be transported to Zr₂CO₂(Hf₂CO₂), but hole movement to Zr₂CO₂(Hf₂CO₂) is hindered. Hence, the resulting charge separation makes the materials the best candidates for light emission applications.⁸⁰

The average electrostatic potential shown in Fig. 7 shows that in the case of the PN-Ti₂CO₂ (PN-Zr₂CO₂) vdWH, the PN(Zr₂CO₂) monolayer has deeper potential than the Ti₂CO₂(-PN) monolayer, indicating that charge is transferred from the Ti₂CO₂(PN) monolayer to the PN(Zr₂CO₂) monolayer. Meanwhile, in the case of GaN-Hf₂CO₂ (AlN-Hf₂CO₂), the Hf₂CO₂(AlN) monolayer has deeper potential than GaN(Hf₂CO₂), showing that charge is transferred from GaN(Hf₂CO₂) to Hf₂CO₂(AlN).

The potential drop across the vdWHs (see Table 3) separates charge carriers (electron and holes) at the interface.^{81,82} Fabrication of the vdWHs enhances the electronic properties of the corresponding monolayers, hence effectively modulating the work function (ϕ). It also changes with strain and electric field engineering and doping.⁸³ So, we have calculated the work function ($\phi = E_{\text{vac}} - E_{\text{F}}$, where E_{vac} (E_{F}) represent the vacuum potential which is derived from the electrostatic potential (Fermi energy)) for all PN-M₂CO₂ (P = Ga, Al; M = Ti, Zr, Hf) vdWHs, see Fig. 7 and Table 3. The calculated work function of PN-M₂CO₂ vdWHs is almost the average of the corresponding monolayers (PN (P = Ga, Al) and M₂CO₂ (M = Ti, Zr, Hf)), which is efficient for the transfer of charge. The effective mass of the electron and holes for PN-M₂CO₂ (P = Ga, Al; M = Ti, Zr, Hf) vdWHs are also calculated and are presented in Table 3. Materials with lower effective mass and high carrier mobility, are strongly preferred for high performance electronic devices,⁸⁴ while a large effective mass greatly suppresses the quantum tunnelling, hence these materials are more appealing for logic devices.⁸⁵ For AlN-Ti₂CO₂ the effective mass of electrons and holes is greater and is suggested for logic devices, while for the



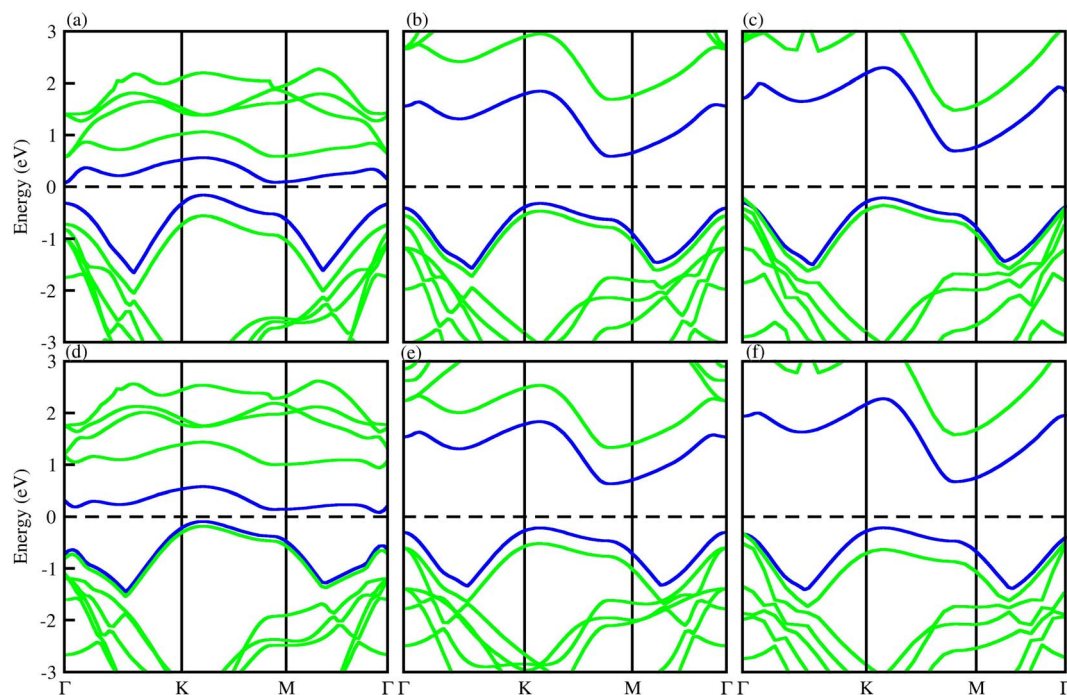


Fig. 5 PBE (blue) and HSE06 (green) band structures of (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂ vdWHs.

rest of the vdWHs the effective mass of electrons and holes is less than AlN-Ti₂CO₂, and is considered to be the best for high performance electronic devices.

The optical properties of materials define how materials react to incident electromagnetic radiation. We used DFT-PBE simulations to calculate optical features of PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs such as the imaginary part of the dielectric function ($\epsilon_2(\omega)$ in Fig. 8) and absorption coefficients ($\alpha(\omega)$ in Fig. 9). The observed excitonic peaks appear at 2.0 (1.9) eV and 2.8 (2.2) eV for AlN-Ti₂CO₂ (GaN-Ti₂CO₂) vdWHs; at 2.2 (2.2) eV and 3.7 (3.8) eV for AlN-Zr₂CO₂ (GaN-Zr₂CO₂) vdWHs; at 2.7 (2.2) eV and 3.7 (2.3) eV for AlN-Hf₂CO₂ (GaN-

Hf₂CO₂) vdWHs. A red (blue) shift is observed in the position of excitonic peaks from AlN to GaN in PN-Ti₂CO₂ and PN-Hf₂CO₂ (PN-Zr₂CO₂) vdWHs. Furthermore, absorption coefficient α (μm^{-1}) provides the photon power attenuation when it passes through the material. It also tells us how far light of specific energy and wavelength can spike the surface of the material before being absorbed.^{86,87} Although absorption is mainly determined by $\epsilon_2(\omega)$, this generalization is obviously not valid if the medium has very large $\alpha(\omega)$.⁸⁸ Therefore, we have further calculated the $\alpha(\omega)$ of PN-M₂CO₂ (P = Ga, Al; M = Ti, Zr, Hf) vdWHs, as presented in Fig. 9, using the Lanczos program to solve the recursively quantum Liouville equation in the

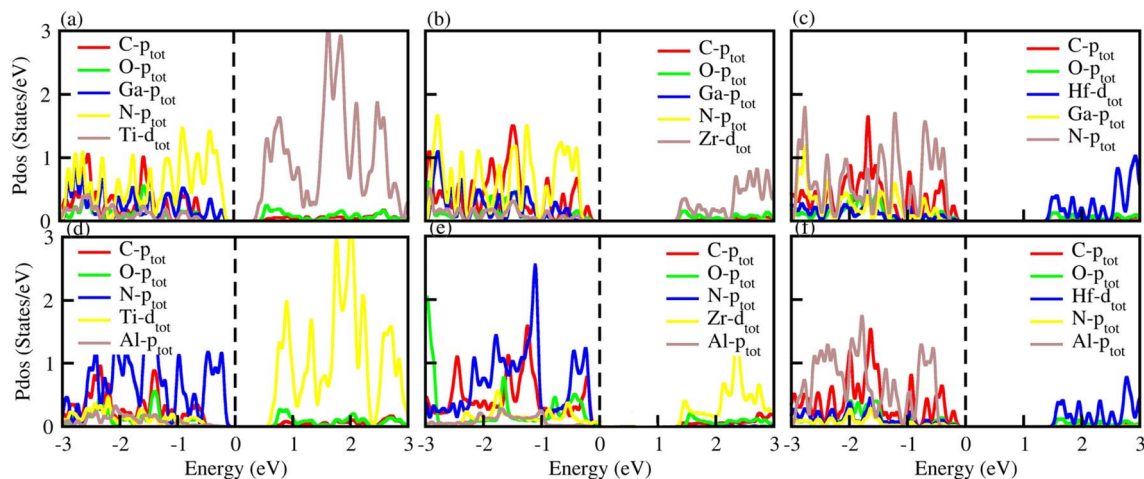


Fig. 6 PDOS of (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂ vdWHs.



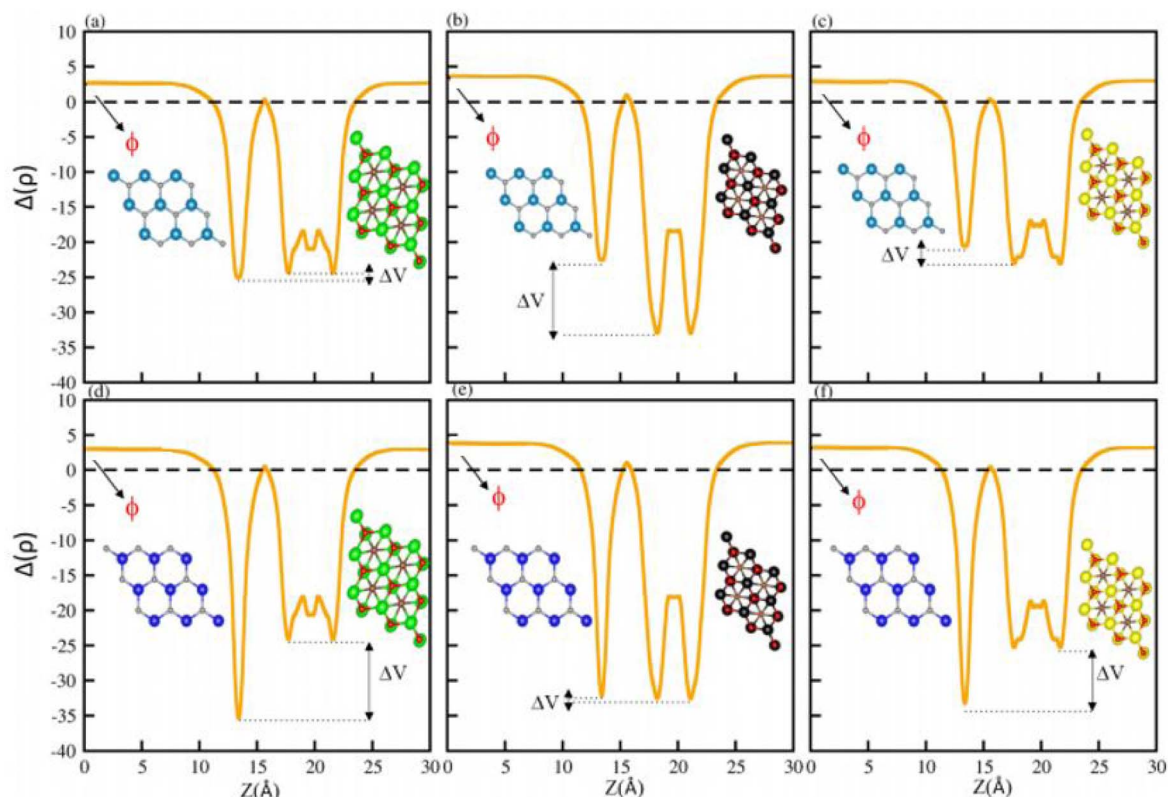


Fig. 7 Average electrostatic potential of (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂ vdWHs. The work function (ϕ) and potential drop (ΔV) are highlighted.

standard batch representation, which allows us to avoid interactions and multiplications of large matrices.⁸⁹ The obtained values for $\alpha(\omega)$ show that there is significant absorption for photon energies above 2 eV for AlN-Zr₂CO₂, GaN-Ti₂CO₂ and PN-Hf₂CO₂, making them candidates with a good optical profile. While in the case of AlN-Ti₂CO₂ and GaN-Zr₂CO₂, the absorption is for photon energies above 1.8 eV. It is straight forward to understand the difference in spectra resulting from the bandstructure.⁹⁰

To utilize solar energy for environmentally friendly applications, it is necessary to design novel materials with visible light

absorption. Therefore, to explore the ability of PN-M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs for water splitting, we calculated the photocatalytic properties by using Mulliken electronegativity.^{91–93} The conduction band (CB) and valence band (VB) edge position of the semiconductors are the two most important parameters to determine photocatalytic activity using under-stimulated sunlight irradiation.⁹⁴ We denote it as H⁺/H₂ for reduction and O₂/H₂O for oxidation, as shown in Fig. 10. If E_{CB} is positioned more negatively than the energy of the H₂/H⁺ potential and the E_{VB} is more positive than the energy of the O₂/H₂O potential, then water molecules can be easily split into H₂

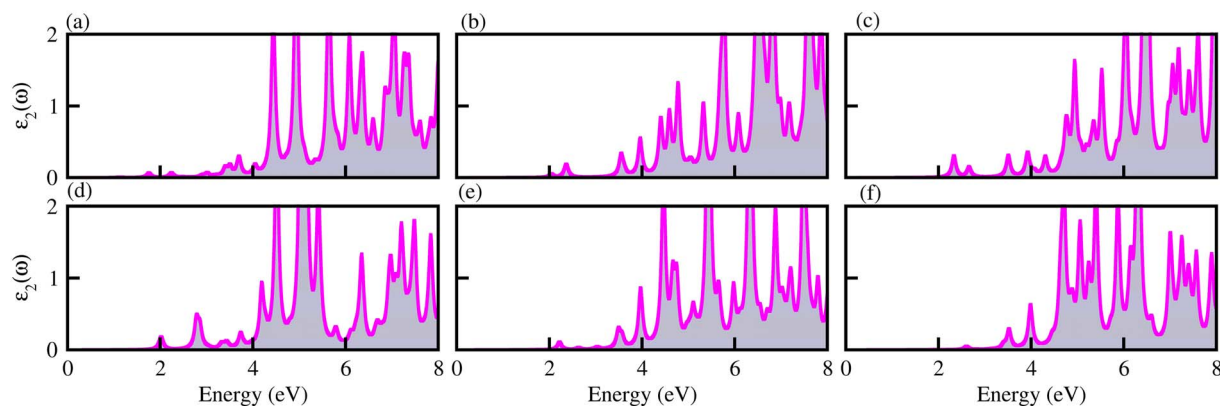


Fig. 8 $\epsilon_2(\omega)$ of (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂ vdWHs.



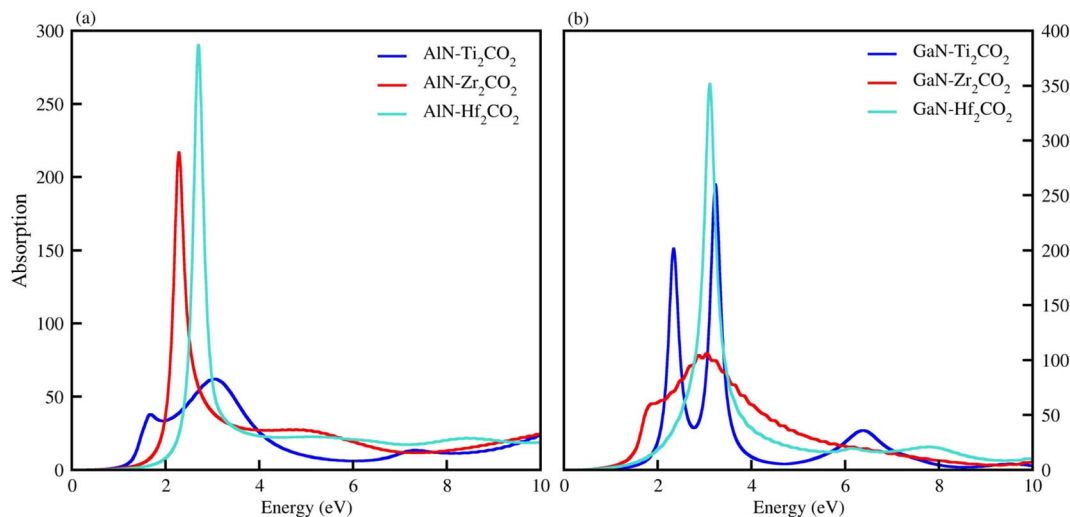


Fig. 9 $\alpha(\omega)$ of (a) AlN-Ti₂CO₂, AlN-Zr₂CO₂ and AlN-Hf₂CO₂ and (b) GaN-Ti₂CO₂, GaN-Zr₂CO₂, GaN-Hf₂CO₂, vdWHs.

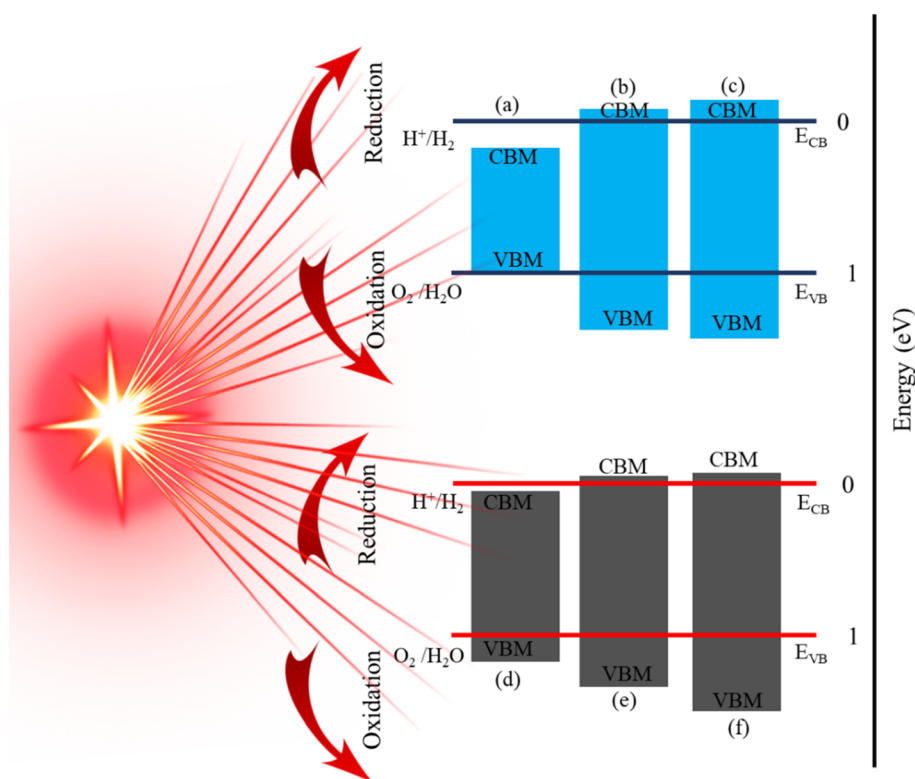


Fig. 10 Band alignment for the valence band (VB) and conduction band (CB) edge of (a) GaN-Ti₂CO₂, (b) GaN-Zr₂CO₂, (c) GaN-Hf₂CO₂, (d) AlN-Ti₂CO₂, (e) AlN-Zr₂CO₂ and (f) AlN-Hf₂CO₂, vdWHs.

and O₂.⁹⁵ In GaN-Zr₂CO₂, GaN-Hf₂CO₂, AlN-Zr₂CO₂, and AlN-Hf₂CO₂, the E_{CB} is positioned more negatively than the energy of the H₂/H⁺ potential, and can split water into H₂. Similarly, if the E_{VB} is more positive than the energy of the O₂/H₂O potential, then the water molecule can easily split into O₂. Similarly, GaN-Ti₂CO₂ and AlN-Ti₂CO₂ have the ability to produce O₂ while E_{CB} has a positive value and is unable to produce H₂ due to the narrow bandgap nature. All these findings demonstrate that PN-

M₂CO₂ (P = Al, Ga; M = Ti, Zr, Hf) vdWHs are the best candidates for photocatalytic water splitting and provide a guideline for designing high performance nanoelectronic, optoelectronic and photocatalytic device applications.

IV. Conclusions

Using DFT, the calculated electronic band structure and photocatalytic performance of PN (P= Ga, Al) and M₂CO₂ (M = Ti,



Zr, Hf) monolayers show the potential of such materials in photocatalytic applications. Based on the same hexagonal symmetry and experimentally achievable lattice mismatch of PN (P = Ga, Al) with M_2CO_2 (M = Ti, Zr, Hf) monolayers, we have fabricated PN- M_2CO_2 vdWHs. The thermal stability of PN- M_2CO_2 vdWHs are confirmed by calculating the binding energies, interlayer distance and AIMD, which demonstrate that these materials can easily be fabricated experimentally. All the PN- M_2CO_2 vdWHs are indirect bandgap semiconductors, while Type-II[-I] band alignment is obtained for GaN(AlN)- Ti_2CO_2 [-GaN(AlN)- Zr_2CO_2 and GaN(AlN)- Hf_2CO_2] vdWHs. The PN- Ti_2CO_2 (PN- Zr_2CO_2) vdWH, PN(Zr_2CO_2) monolayer, have greater potential than the Ti_2CO_2 (PN) monolayer, indicating that charge is transferred from the Ti_2CO_2 (PN) to PN(Zr_2CO_2) monolayer, hence, the potential drop at the interface separates charge carriers (electron and holes). The work function and effective mass of the carriers of PN- M_2CO_2 vdWHs are also calculated and presented. For the AlN- Ti_2CO_2 effective mass of electrons and holes is greater, and is hence suggested for logic devices, while for the rest of the vdWHs, the effective mass of electrons and holes are lower. Based on $\epsilon_2(\omega)$ $\alpha(\omega)$, a red (blue) shift is observed in the position of excitonic peaks from AlN to GaN in PN- Ti_2CO_2 and PN- Hf_2CO_2 (PN- Zr_2CO_2) vdWHs, while a significant absorption for photon energies above 2 eV for AlN- Zr_2CO_2 , GaN- Ti_2CO_2 and PN- Hf_2CO_2 , make them candidates with a good optical profile. The calculated photocatalytic properties demonstrate that PN- M_2CO_2 (P = Al, Ga; M = Ti, Zr, Hf) vdWHs are the best candidates for photocatalytic water splitting and provide a guideline for experimentalist to design high performance nanoelectronics, optoelectronic and photocatalytic device applications.

Conflicts of interest

There is no conflicts to declare.

Acknowledgements

This research was funded by the Princess Nourah bint Abdulrahman University Researchers, Riyadh, Saudi Arabia, with supporting project number (PNURSP2023R71). The authors extend their sincere appreciation to Princess Nourah bint Abdulrahman University.

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