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Screening of electrode materials for ammonium ion batteries by high throughput calculation†

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Ammonium-ion batteries (AIBs) have attracted intense interest lately as promising energy storage systems due to their light weight, safe, inexpensive, and widely available advantages. It is of great significance to find a fast ammonium ion conductor for the electrode of AIBs that directly affects the electrochemical performance of the battery. Using high-throughput bond-valence calculation, we screened the electrode materials of AIBs with a low diffusion barrier from more than 8000 compounds in the ICSD database. Twenty-seven candidate materials were finally identified by the bond-valence sum method and density functional theory. Their electrochemical properties were further analyzed. Our results, which give the relationship between the structure and electrochemical properties of various important electrode materials which are suitable for AIBs development, may pave the way for next-generation energy storage systems.

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1. Introduction

With the worldwide emphasis on renewable energy sources such as solar and wind, battery energy storage has become an essential solution for grid stability and reliability. Traditional metal ion batteries (Li^+ ,¹ Na^+ ,² K^+ ,³ *etc.*) are relatively expensive in terms of abundance in Earth's crust and mining cost. Emerging non-metallic batteries (proton,⁴ NH_4^+ ,⁵ Cl^- ,⁶ *etc.*) have developed rapidly in recent years. Their cost, capacity, operating voltage, rate performance, and cycle stability are better than metal batteries. Among them, ammonium-ion batteries (AIBs) are promising candidates for large-scale energy storage systems. Compared with metal batteries, ammonium ions have the merit of low molar mass (18 g mol^{-1}), rich reserves and being renewable (only containing N and H essential elements), and having smaller hydrated ion sizes.^{7,8} Compared with acidic electrolytes, electrolytes containing NH_4^+ show weak acidity and are less corrosive, will not lead to the dissolution of electrode materials and have higher HER (hydrogen evolution reaction) overpotential than electrolytes containing H^+ .^{9,10}

The progress of ammonium-ion batteries (AIBs)^{5,9,11–15} is still in its infancy. Ji's team systematically studied the insertion of NH_4^+ into layered V_2O_5 . It was found that ammonium-ion forms strong hydrogen bonds with layered V_2O_5 , and the diffusion

mode of NH_4^+ is similar to monkey swing.¹¹ Wang also developed NH_4^+ storage based on supercapacitor. The research indicates that the ammoniation/deamination mechanism is dominated by non-diffusion controlled pseudo capacitance behavior.¹⁴ So far, the reported ammonium storage materials are basically MXene,¹⁶ Prussian blue analogues,¹⁷ conductive polymers (PANI)¹⁸ and metal oxides (MoO_3 , V_2O_5 , MnO_x , $\text{Fe}_5\text{-V}_{15}\text{O}_{39}(\text{OH})_9 \cdot 9\text{H}_2\text{O}$).^{12,19,20} However, the performances achieved to date remain inferior to what lithium-ion batteries can do, showing limited capacity and cycling stability. Therefore, to achieve rechargeable AIBs with excellent performance, it is necessary to explore new electrodes for NH_4^+ storage.

At the current stage, the discovery of these energy materials relies predominantly on experimental serendipity and the try-and-error experimental process that are inefficient and time-consuming. Nevertheless, much evidence has shown that the discovery process of new energy materials could be significantly accelerated by the data mining process and high throughput calculation.^{21,22} Now, various experimental and computational repositories such as Inorganic Crystal Structure Database (ICSD),²³ Crystallography Open Database (COD),²⁴ the Materials Project,²⁵ and others store thousands of experimental and calculated structures, which already serve as potential batteries electrodes or still wait for such an opportunity.

Meanwhile, A. O. Boev *et al.* extracted 2800 initial structures from the material database and identified 33 potential compounds for cathode materials of lithium-oxygen batteries with low normalized surface energy.²⁶ Max Avdeev used high-throughput screening to conduct a detailed analysis of the 13 000 materials in ICSD, and identified the crystal structures featuring infinite networks of pathways suitable for Li^+ , Na^+ , K^+ , Ag^+ , and Cu^+ ionic transport.²⁷ However, there is no report on

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the high-throughput screening method to determine the candidate electrode materials for ammonium-ion batteries.

As inspired by the ion exchange experiments, the compounds from ICSD (inorganic crystal structures database)²³ which contain K⁺ ion were chosen. We replaced the K⁺ ions in these compounds with NH₄⁺ to work as the input of our following screening process, since the K⁺ and NH₄⁺ ions have the same valence and almost the same ionic radii (149 pm for K⁺ and 148 pm for NH₄⁺).²⁸ Further, we used the bond valence energy landscape method (BVEL),^{29–33} density functional theory (DFT) method and high-throughput screening to find new electrode materials for AIBs. 27 candidate materials were identified. Our theoretical results may open new opportunities for the development of high-performance AIBs.

2. Methods

2.1 Screening algorithm

The approach can be considered as multi-level screening, where at each higher level the complexity of computational methods increases, while the number of candidates reduces. A total of 120 481 database entries were screened. We selected the world's largest database ICSD for completely identified inorganic crystal structures as the input to the screening framework. The main steps of the filtering algorithm are as follows: (I) Screening compounds containing K⁺ in the database and replacing K⁺ with NH₄⁺, about 8221 compounds were obtained. (II) From bond valence energy landscape (BVEL) calculation, 441 compounds with diffusion barriers less than 0.3 eV were screened. (III) Considering that structures which are suitable for chemical ion exchange route should have robust framework. 166 compounds with general formula of (NH₄)_aM_b(XY)_c_d (M – metal, in which inactive metals such as Zn, Zr, Ga, Ge, Mg, Al, *etc.*, for electrolyte and active transition metal such as Fe, Mn, Ti, Cr, V, Co, Ni, Nb, Mo, *etc.* for electrode material; XY_c – polyanion groups, X – metalloids such as B, Si, Ge, As and Mo, Nb, *etc.*, Y – O, F, S, *etc.*) were screened.³⁴ (IV) 108 compounds were obtained by a filter of specific capacity greater than 100 mA h g⁻¹. (V) After the first principle calculation, the 72 compounds were screened using the condition that the open circuit voltage (OCV) was higher than 3 V. (VI) The final screening condition is that the volume expansion rate of the optimized structure is less than 10% compared with the K-contained analog, and 27 compounds are selected (Fig. 1).

2.2 Bond valence energy landscape (BVEL) method

The bond-valence method³⁵ is the development of the principle of local charge neutrality (first proposed by Pauling). The valence of an atom is the sum of the individual bond valences surrounding the atom. The results show that the dependence of s_{ij} on the distance between the i^{th} and j^{th} ions R_{ij} can be well described by exponential function.

$$s_{ij} = \exp\left(\frac{R_0 - R_{ij}}{b}\right)$$

where R_0 and b are empirical parameters.³⁶

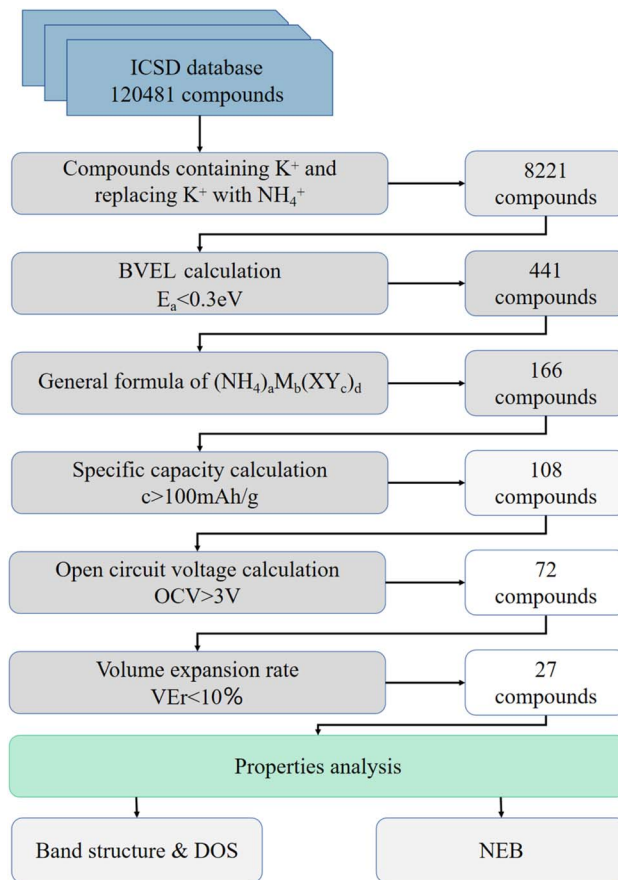


Fig. 1 The schema diagram of the screening process. All data relating to screening process are included in the ESI.

As shown in S. Adams's work,³¹ the bond valence site energy of i can then be interpreted as a result of a Morse-type interaction with the adjacent anions j , it can be combined with the Coulomb repulsion $E_{\text{repulsion}}$ of i with other immobile cations.

$$E = \sum_j D_0 \left\{ \sum_{i=1}^N \left[\left(\frac{S_{ij}}{S_{\text{min},ij}} \right)^2 - \frac{2S_{ij}}{S_{\text{min},ij}} \right] \right\} + E_{\text{repulsion}}$$

$$E_{\text{repulsion}}(ij) = \frac{q_i q_j}{R_{ij}} \operatorname{erfc} \left(\frac{R_{ij}}{f(r_i + r_j)} \right)$$

where r_i and r_j are the covalent radii of the respective cation (or anion) pair i, j and f is a screening factor (typically of the order $f \approx 0.75$).³⁷ It may be noted that for the purpose of analyzing E landscapes, Coulomb repulsions are considered only between mobile and immobile cations to reveal continuous pathways, while coulomb attraction terms are generally integrated into the Morse attraction term. The charge screening is based on an error function complement term $E_{\text{repulsion}}$.

The console utilities CrysFML (Crystallographic Fortran 95 Modules Library) which are developed by N. A. Katcho *et al.*³⁸ are used for bond-valence energy landscapes calculations. Small Python scripts have been written to bond the input and output between programs to reduce human intervention and facilitate



data management. Within the theoretical framework of bond-valence sum, the diffusion barriers (E_a) in three directions of a , b and c of the cell are obtained. The fundamental process that enables the functioning of most of today's rechargeable batteries is the reversible intercalation of ions into electrode. Ionic conductivity emerges as a paramount criterion. We abandon the structures with diffusion barrier higher than 0.3 eV.

2.2.2 Structure and specific capacity selection. Most of the cathode materials for Li-ion or Na-ion batteries have more than 3 elements in their chemical formulas.^{39,40} For AIBs, we focus on compounds with the general formula of $(\text{NH}_4)_a\text{M}_b(\text{XY})_c\text{O}_d$. And we also eliminate the compounds which have elements with low abundance and high price such as Sc, Y, In or rare-earth metals.

Specific capacity is an indicator to the amount of electric charge stored by the electroactive materials in a unit mass. The formula of the specific capacity is as follows:

$$C = nF/M$$

where n is the number of transferred electrons, M is the molar mass of the selected compounds, and F is Faraday constant (26 801 mA h mol⁻¹). We selected compounds with theoretical specific capacity more than 100 mA h g⁻¹ for the next round of screening. 108 kinds of materials were screened from 8221 types of materials.

2.2.3 Density functional theory (DFT) calculations. The density functional theory method has emerged as an integral and

important part of the screening process for electrode materials design. All calculations were performed with the Vienna *Ab initio* Simulation Package (VASP),⁴¹ within the projector augmented wave (PAW) approach.⁴² The Python Materials Genomics (pymatgen) materials analysis library was used for all analyses.⁴³ All structural relaxations and total energy calculations were carried out using parameters similar to those used in the Materials Project.⁴⁴ The key parameters are the use of the Perdew–Burke–Ernzerhof (PBE) generalized-gradient approximation (GGA)⁴⁵ exchange correlation functional, an energy cutoff of 520 eV and a k -point mesh of at least 4377 per atom. The relaxation process is divided into two steps. First, keep the cell volume unchanged, optimize the atomic position, and then optimize the cell fully until the convergence criterion is reached. For structures with fractional occupancy in the database, VCA (virtual crystal approximation)⁴⁶ method is used. VCA method is to scale the local potential, the augmentation charges, and the non-local pseudopotential strength parameters of atoms occupied by fractions from the values provided. We also carried out the static calculations self-consistent and non-self-consistent calcul to explore the density of states and electronic band structure of compounds.

Open circle voltage (OCV)^{47,48} is determined by the chemical potential difference between cathode and anode, and can be calculated as

$$\text{OCV} = \frac{E_{\text{cathode}(x)} - E_{\text{anode}}}{Z_e}$$

Table 1 27 potential electrode materials for ammonium-ion batteries after the screening. ID is the ICSD collection code of the compound containing K⁺ before replacement, and E (eV) is the minimum potential barrier value calculated by BVFL. The dimension of the diffusion path (D) is judged by barrier energies along the three coordinate axes. Open circuit voltage OCV (V), B_g (eV) is the band gap. The c (mA h g⁻¹) is the specific capacity. V (Å³) is the unit cell volume of compounds containing K⁺, V' (Å³) is the unit cell volume of compounds containing NH₄⁺ after replacement and optimization, and Ver (%) is the volume expansion rate

ID	Formula	E	D	c	OCV	B_g	V	V'	Ver
16 557	NH ₄ TiNbO ₅	0.09	2	121.22	5.24	2.88	452.43	483.36	6.84
18 022	NH ₄ Zn(BeF ₃) ₃	0.30	3	101.86	6.85	6.14	361.15	390.72	8.19
30 910	NaNH ₄ GeO ₃	0.17	2	186.03	4.37	3.31	353.52	378.93	7.19
39 440	NH ₄ CrPO ₄ F	0.29	3	161.38	4.80	2.57	855.76	910.87	6.44
39 697	(NH ₄) ₂ (Zr _{0.86} Ti _{0.14})(Si ₃ O ₉)(H ₂ O)	0.23	3	171.19	3.76	0.04	969.57	1060.22	9.35
39 735	NH ₄ GeOPO ₄	0.26	3	145.59	5.04	3.15	794.77	853.99	7.45
48 177	NH ₄ Li ₂ (BO ₃)	0.26	3	336.97	3.39	4.34	331.05	356.73	7.76
59 281	(NH ₄) ₂ MgWO ₂ (PO ₄) ₂	0.29	3	124.60	5.06	3.55	898.89	968.91	7.79
59 285	NH ₄ (TiO)(PO ₄)	0.24	3	168.48	5.16	3.07	866.17	919.61	6.17
65 260	NH ₄ Li ₄ (AlO ₄)	0.15	1	273.36	3.95	4.16	889.14	968.35	8.91
69 429	NH ₄ (SbO)(SiO ₄)	0.17	3	116.47	5.33	2.67	893.75	954.77	6.83
74 591	NH ₄ (TiO)(AsO ₄)	0.19	3	131.97	5.13	3.13	928.52	989.23	6.54
79 650	NH ₄ SnO(PO ₄)	0.22	3	116.47	5.30	3.02	925.30	983.87	6.33
79 651	NH ₄ (VO)(PO ₄)	0.24	3	165.36	3.98	1.55	864.21	909.35	5.22
79 702	(NH ₄) ₂ Ni(WO ₂ (PO ₄) ₂)	0.19	3	123.45	4.93	3.38	896.14	964.34	7.61
80 023	NH ₄ (TiO)((P _{0.56} As _{0.44})O ₄)	0.22	3	150.20	5.15	3.12	894.19	960.44	7.41
80 893	NH ₄ GaF(PO ₄)	0.30	3	145.59	5.18	3.98	835.98	895.06	7.07
82 398	(NH ₄) ₂ (Be ₂ Si ₃ O ₉)	0.28	3	217.80	4.92	5.03	829.99	871.74	5.03
82 457	NH ₄ Li(Si ₂ O ₅)	0.28	2	187.34	4.99	4.98	233.78	255.75	9.40
91 534	NH ₄ (Ti _{0.936} Sn _{0.064})O(PO ₄)	0.24	3	163.80	4.90	—	870.50	913.43	4.93
200 310	(NH ₄) ₂ Ni(MoO ₄) ₂	0.27	2	141.37	5.16	3.85	1747.49	955.38	9.34
200 613	NH ₄ Fe(PO ₄)F	0.28	3	157.58	5.17	2.33	874.62	925.15	5.78
203 214	(NH ₄) ₂ (Mn ₃ (OH) ₂ (VO ₄) ₂)	0.30	2	124.89	3.48	1.96	487.51	261.09	7.11
203 218	NH ₄ (MoO ₃)(AsO ₄)	0.28	1	100.33	5.24	2.86	493.16	542.13	9.93
280 327	(NH ₄) ₂ (TiSi ₃ O ₉)(H ₂ O)	0.16	3	194.12	5.19	3.46	920.71	985.68	7.06
400 850	(NH ₄) ₂ (TiO) ₂ (As _{0.43} P _{0.57} O ₄) ₂	0.21	3	150.57	5.15	3.16	895.69	959.36	7.11
411 501	(NH ₄) ₃ (BSb ₄ O ₁₃)	0.21	3	113.67	5.22	1.78	598.67	642.27	7.28



where E_{cathode} is defined as the free energy of the intercalated cathode. E_{anode} is defined as the free energy of each atom of the anode (NH_4^+). Z_e represents the transferred charge. For NH_4^+ , we choose a condensed phase similar to body centered cubic lattice, and Z is the valence number ($Z = 1$ for NH_4^+). The widely used anode in AIBs is conductive polymer materials such as polyaniline (PANI) which has good electronic conductivity, high accommodation capabilities for large ions and light weight.¹⁸ This anode has higher chemical potential compared with our “fake” pure NH_4^+ lattice. Thus, our OCV values are overestimated can only work as upper limit of the expected OCV values.

To further evaluate the ionic conductivities, the nudged elastic band (NEB)^{49,50} method based on DFT is adopted to identify ion migration paths and compute their corresponding energy barriers. Despite the success of the NEB method in characterizing the diffusion process in electrode, the significant computational expense of NEB relative to standard DFT calculations, *e.g.*, geometry relaxations and static energy calculations, has hampered its application to cover all the filtered compounds. Only the final screening compounds are calculated to verify the diffusion barrier calculated by the BVLE method.

3. Results and discussion

3.1 Screening overview

In the screening algorithm, only the first screening method BVLE diffusion barrier calculation is empirical method. The rest of the filtering is based on first principle calculations. The diffusion

barrier, open circuit voltage, and specific capacity are designed for better battery performance. While the general formula of $(\text{NH}_4)_a\text{M}_b(\text{XY}_c)_d$ and smaller volume expansion rate compounds are selected to obtain more stable structures. Moreover, the smaller volume expansion rate means that the compound can be more likely synthesized by the ion exchange experiment.

From 8221 compounds containing K^+ in the ICSD database, we obtained 27 potential electrode materials for AIBs (Table 1). The calculated OCV is in the range of 3.39 V to 6.85 V. The maximum OCV is obtained from $\text{NH}_4\text{Zn}(\text{BeF}_3)_3$ which contains fluorine atoms. The OCV of $(\text{NH}_4)_a\text{M}_b(\text{PO}_4)_d$ derivatives are about 5.0 V depending on the different metal ion involved. The specific capacity of selected compounds is in the range of 100.33 to 336.97 mA h g^{-1} . The maximum specific capacity is calculated from $\text{NH}_4\text{Li}_2(\text{BO}_3)$ which contains only light elements. We also explored their electronic structures within the theoretical framework of density functional theory. Band structures of the screened compounds have been obtained, and the band gaps are shown in Table S6 in ESI.† Further, we have also performed DFT-based NEB calculations for 27 candidate compounds to analyze the feasibility of the screening algorithm and the diffusion properties of candidate materials (Table 2).

3.1.1 Layered type. After screening, only two compounds have an obvious layered structure, namely $\text{NH}_4\text{TiNbO}_5$ and $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$. As shown in Table S6.† A recent study⁵¹ performed ion exchange from KTiNbO_5 to $\text{NH}_4\text{TiNbO}_5$. The cell parameters of $\text{NH}_4\text{TiNbO}_5$ are consistent with those in this paper, which demonstrate *Pnmm* orthorhombic structure with

Table 2 Diffusion barrier from NEB method E_{neb} (eV) and BVLE method three axis $E_{\text{a/b/c}}$ (eV) of 27 potential ammonium-ion battery electrode materials, and their open circuit voltage OCV (V) is also listed

ID	Formula	E_{a}	E_{b}	E_{c}	E_{neb}	OCV
16 557	$\text{NH}_4\text{TiNbO}_5$	0.09	0.65	—	0.11	5.24
18 022	$\text{NH}_4\text{Zn}(\text{BeF}_3)_3$	3.06	3.06	0.30	1.18	6.85
30 910	$\text{NaNH}_4\text{GeO}_3$	—	0.17	2.00	0.67	4.37
39 440	$\text{NH}_4\text{CrPO}_4\text{F}$	1.00	0.29	1.00	0.69	4.80
39 697	$(\text{NH}_4)_2(\text{Zr}_{0.86}\text{Ti}_{0.14})(\text{Si}_3\text{O}_9)(\text{H}_2\text{O})$	0.85	0.85	0.23	0.26	3.76
39 735	$\text{NH}_4\text{GeOPO}_4$	0.91	0.91	0.26	1.24	5.04
48 177	$\text{NH}_4\text{Li}_2(\text{BO}_3)$	1.72	0.26	1.72	0.12	3.39
59 281	$(\text{NH}_4)_2\text{MgWO}_2(\text{PO}_4)_2$	0.53	0.49	0.29	0.46	5.06
59 285	$\text{NH}_4(\text{TiO})(\text{PO}_4)$	0.67	0.67	0.24	0.79	5.16
65 260	$\text{NH}_4\text{Li}_4(\text{AlO}_4)$	0.15	—	—	0.60	3.95
69 429	$\text{NH}_4(\text{SbO})(\text{SiO}_4)$	0.72	0.72	0.17	0.83	5.33
74 591	$\text{NH}_4(\text{TiO})(\text{AsO}_4)$	0.51	0.51	0.19	0.14	5.13
79 650	$\text{NH}_4\text{SnO}(\text{PO}_4)$	0.67	0.67	0.22	1.67	5.30
79 651	$\text{NH}_4(\text{VO})(\text{PO}_4)$	0.65	0.65	0.24	0.14	3.98
79 702	$(\text{NH}_4)_2\text{Ni}(\text{WO}_2(\text{PO}_4)_2)$	0.44	0.44	0.19	0.48	4.93
80 023	$\text{NH}_4(\text{TiO})(\text{P}_{0.56}\text{As}_{0.44}\text{O}_4)$	0.58	0.58	0.22	0.62	5.15
80 893	$\text{NH}_4\text{GaF}(\text{PO}_4)$	1.17	1.17	0.30	0.21	5.18
82 398	$(\text{NH}_4)_2(\text{Be}_2\text{Si}_3\text{O}_9)$	0.77	0.28	0.77	0.49	4.92
82 457	$\text{NH}_4\text{Li}(\text{Si}_2\text{O}_5)$	3.17	0.28	—	0.50	4.99
91 534	$\text{NH}_4(\text{Ti}_{0.936}\text{Sn}_{0.064}\text{O})(\text{PO}_4)$	0.67	0.67	0.24	0.14	4.90
200 310	$(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$	0.27	—	0.47	0.54	5.16
200 613	$\text{NH}_4\text{Fe}(\text{PO}_4)\text{F}$	0.28	1.03	1.03	0.12	5.17
203 214	$(\text{NH}_4)_2(\text{Mn}_3(\text{OH})_2(\text{VO}_4)_2)$	—	0.30	2.31	0.12	3.48
203 218	$\text{NH}_4(\text{MoO}_2)(\text{AsO}_4)$	—	0.28	—	0.69	5.24
280 327	$(\text{NH}_4)_2(\text{TiSi}_3\text{O}_9)(\text{H}_2\text{O})$	0.16	0.50	0.49	0.63	5.19
400 850	$(\text{NH}_4)_2(\text{TiO})_2(\text{As}_{0.43}\text{P}_{0.57}\text{O}_4)_2$	0.57	0.57	0.21	0.63	5.15
411 501	$(\text{NH}_4)_3(\text{BSb}_4\text{O}_{13})$	0.21	0.23	0.67	0.23	5.22



$a = 6.47 \text{ \AA}$, $b = 3.80 \text{ \AA}$, $c = 19.29 \text{ \AA}$, and $V = 474.06 \text{ \AA}^3$. The calculations in this paper show basically consistent results of $a = 6.61 \text{ \AA}$, $b = 3.83 \text{ \AA}$, $c = 19.09 \text{ \AA}$, and $V = 483.36 \text{ \AA}^3$. And it has a theoretical open circuit voltage of 5.24 V and a low diffusion barrier of 0.11 eV calculated by NEB (and 0.09 eV calculated by BVEL), which means that it may be a high-performance electrode material for AIBs. Meanwhile, from our screening results, the number of compounds containing Mo is large, which is consistent with the types of electrode materials reported for AIBs. Wherein $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$ has 5.16 V open circuit voltage and 0.47 eV diffusion barrier calculated by BVEL.

We take $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$ as an example. $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$ (Fig. 2(b)) is a compound obtained by replacing K^+ from $\text{K}_2\text{Ni}(\text{MoO}_4)_2$ (Fig. 2(a)). The space group is P1, and NH_4^+ can migrate along the a -axis and c -axis. Therefore, after BVEL calculation, as shown in Fig. 2(c), the low-valued isosurface identifies areas with a low energy barrier and therefore represents the most likely pathway for ion transport. In addition, Fig. 2(d) shows the charge density after the DFT calculation. When we replace K^+ with NH_4^+ , and after structural relaxation, it is found that the unit cell has a small expansion (Fig. 2(a) and (b)), the unit cell volume of $\text{K}_2\text{Ni}(\text{MoO}_4)_2$ expanded from 1747.49 \AA^3 to 1910.76 \AA^3 (Tables S4 and S5 in ESI[†]). At the same time, the symmetry of the structure also changed because of the difference between the tetrahedral structure of NH_4^+ and the spherical structure of K^+ . This expansion distortion will not cause the structure of the compound to collapse. From NEB calculation, it is revealed that the strong hydrogen bond makes the diffusion pattern of NH_4^+ ions similar to the monkey swing (Fig. 2(f)), which is consistent with the study of Ji's team.¹¹ And this diffusion mode is accompanied by the continuous formation and breaking of hydrogen bonds, which can show rapid electrode dynamics and further improve the electrochemical performance of AIBs.

3.1.2 Tunnel type. The tunnel electrode materials account for 25 of the 27 selected materials. To take a further look at $\text{NH}_4\text{Fe}(\text{PO}_4)\text{F}$, the minimum BVS mismatch path of NH_4^+ shows obvious 3 dimension pathways. From the BVEL data map (isosurface = $0.58 e \text{ Bohr}^{-3}$), ammonium-ion diffuses in a zig-zag-shaped path (Fig. 3(c)), which is the same as reported in ref. 52 which the electrode material NaFeSO_4F of Na-ion batteries was shown.⁵² In these tunnel type electrodes, there is 1 dimension diffusing type electrode. The BVEL map of $\text{NH}_4\text{-Zn}(\text{BeF}_3)_3$ shows a 1 dimension diffusion along the z -axis (Fig. 3(g)), which is similar to the reported $\alpha\text{-MoO}_3$ electrode.⁵³ It has a simple tunnel structure, and the formation and fracture of hydrogen bonds are still the primary way of its diffusion, which also guarantees ultrafast kinetics in the reversible insertion/deinsertion process during cycling.⁵⁴ $\text{NH}_4\text{CrPO}_4\text{F}$, $\text{NH}_4(\text{TiO})\text{AsO}_4$ and $(\text{NH}_4)_3(\text{BSb}_4\text{O}_{13})$ also have very small diffusion barriers and large specific capacity, which is potential candidate for fast ammonium ion storage electrode.

3.2 Electronic structure

The preliminary idea to calculate the electronic band structure of a potential electrode material is to determine its metallic, semiconducting or insulating character. The structures and

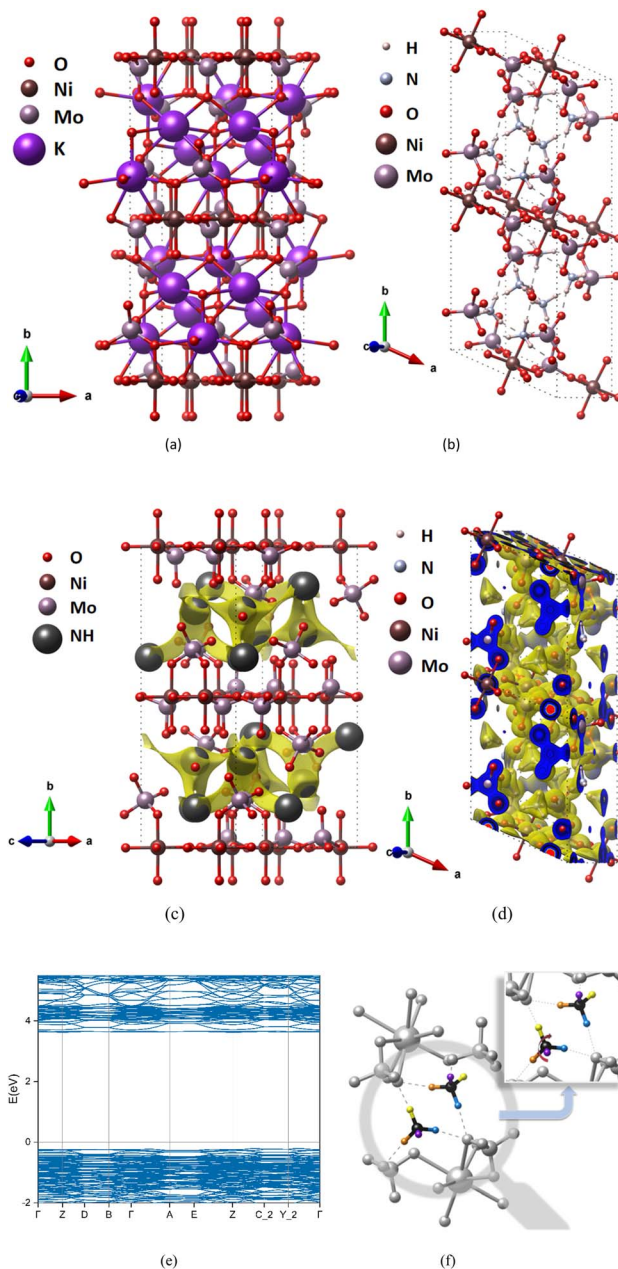


Fig. 2 Structure and charge distribution of $(\text{K}/\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$, the atomic type is marked in the upper left corner of the figures, where "NH" represents ammonium-ion. Unoptimized K precursor $\text{K}_2\text{Ni}(\text{MoO}_4)_2$ (a) and optimized structure after replacing K^+ with NH_4^+ $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$ (b). BVEL data (c), charge density diagram (d) and band structure calculated by DFT (e). The diffusion mode of ammonium ion is revealed by NEB calculation (f), in which ammonium is a colored atom, and the same color represents the same atom before and after diffusion.

density of states diagrams are listed in Fig. S1–S27 and Tables S7–S33 of the ESI.[†] Almost all the compounds have band gaps, suggesting poor electronic conductivity. Therefore, further work is needed to improve electronic conductivity.

Among the 27 selected compound, there are eleven $(\text{NH}_4)_a\text{-M}_b(\text{PO}_4)_a$ derivatives. The band gap of $\text{NH}_4\text{CrPO}_4\text{F}$,



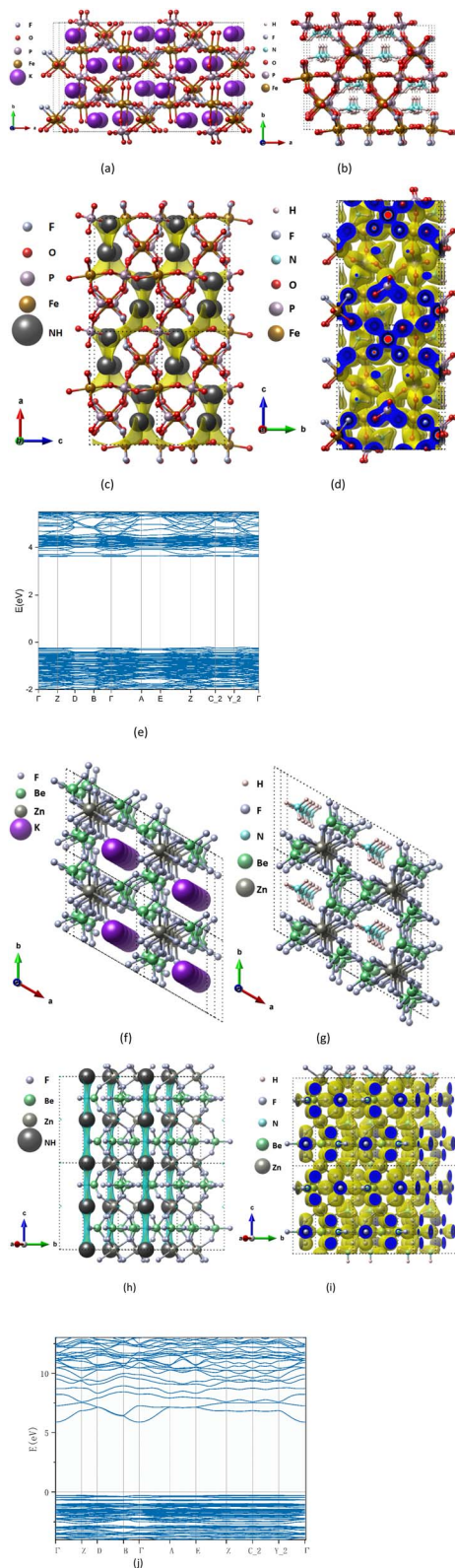


Fig. 3 Structure and charge distribution of $(\text{K}/\text{NH}_4)\text{Fe}(\text{PO}_4)\text{F}$ and $(\text{K}/\text{NH}_4)\text{Zn}(\text{BeF}_3)_3$ ($2 \times 2 \times 2$ supercell), the atomic type is marked in the upper left corner of the figures, where "NH" represents ammonium-ion. Unoptimized K precursor $\text{KFe}(\text{PO}_4)\text{F}$ (a), $\text{KZn}(\text{BeF}_3)_3$ (f) and optimized structure after replacing K^+ with NH_4^+ $\text{NH}_4\text{Fe}(\text{PO}_4)\text{F}$ (b) $\text{NH}_4\text{-Zn}(\text{BeF}_3)_3$ (g). BVLE data (c and h) and charge density diagram (d and i) and band structure calculated by DFT (e and j).

$\text{NH}_4\text{GeOPO}_4$, $(\text{NH}_4)_2\text{MgWO}_2(\text{PO}_4)_2$, $\text{NH}_4(\text{TiO})(\text{PO}_4)$, $\text{NH}_4\text{-SnO}(\text{PO}_4)$, $\text{NH}_4(\text{VO})(\text{PO}_4)$, $(\text{NH}_4)_2\text{Ni}(\text{WO}_2(\text{PO}_4)_2)$, $\text{NH}_4\text{-TiO}((\text{P}_{0.56}\text{As}_{0.44})\text{O}_4)$, $\text{NH}_4\text{GaF}(\text{PO}_4)$, $\text{NH}_4\text{Fe}(\text{PO}_4)\text{F}$ and $(\text{NH}_4)_2(\text{TiO})_2(\text{As}_{0.43}\text{P}_{0.57}\text{O}_4)_2$ is 2.57, 3.15, 3.55, 3.07, 3.02, 1.55, 3.38, 3.12, 3.98, 2.33 and 3.16 eV respectively. The band gap of LiFePO_4 determined with the *ab initio* GGA+U method is 3.8 eV.⁵⁵ Incorporating of metal ion can efficiently modulate the band gap of $(\text{NH}_4)_a\text{M}_b(\text{PO}_4)_d$ derivatives. The only exception is $\text{NH}_4(\text{VO})(\text{PO}_4)$ which has much smaller band gap (1.55 eV). Also, there are five $(\text{NH}_4)_a\text{M}_b(\text{Si}_c\text{O}_d)_e$ derivatives. The band gap of these compound is in the range 0.04 eV to 5.03 eV. The maximum band gap is 6.14 eV calculated from fluorid $\text{NH}_4\text{Zn}(\text{BeF}_3)_3$. The theoretical band gap of $(\text{NH}_4)_2(\text{Zr}_{0.86}\text{Ti}_{0.14})(\text{Si}_3\text{O}_9)(\text{H}_2\text{O})$ calculated by virtual crystal approximation is very small (0.04 eV) and $\text{NH}_4(\text{Ti}_{0.936}\text{Sn}_{0.064})\text{O}(\text{PO}_4)$ does not show band gap. These two values are abnormal requiring further calculation under more complex methodological architectures.

3.3 Diffusion barrier calculation

Further, we have also carried out more accurate NEB calculations of the screened 27 compounds. The layered structure of $\text{NH}_4\text{TiNbO}_5$ and $(\text{NH}_4)_2\text{Ni}(\text{MoO}_4)_2$ has small diffusion barrier (0.11 and 0.54 eV) which indicates fast NH_4^+ kinetics. From a structural point of view, the compounds with small diffusion barrier have obvious diffusing channel. The selected bulk materials shown in ESI,[†] have diverse diffusion pattern.

The results (Table 2) show that there is small error between the NEB method based on DFT and the BVLE calculation results for most cases. The main cause of the deviation is that a "static" structure is used during the BVLE calculation. And the cell parameters of the compound remain unchanged after NH_4^+ substitution, which further leads to the difference between BVLE and the experimental (or DFT) barrier. For instance, the diffusion barrier of $\text{NH}_4\text{Zn}(\text{BeF}_3)_3$ and $\text{NH}_4(\text{SnO})(\text{PO}_4)$ using NEB based on DFT optimization shows ~ 1 eV higher than the BVLE barrier, which is due to the significant changes in structural parameters compared with its K precursor (specific data are shown in Tables S4 and S5 of ESI[†]). Despite deviations in the computed values, the high-throughput screening is capable to predict trends in structure–property relationships. The magnitude of deviations can differ substantially for different materials, environments, and thermodynamic factors. The results from the NEB method are more trustable and should have the quantitative agreement of dynamical properties with experimental results. Nevertheless, the BVLE method is still reliable as a rough selection of candidate battery electrode materials. Considering the complexity of the diffusion process of these compounds and the NEB paths considered may be unexhaustive, all the 27 compounds are reserved for further comprehensive study.

4. Summary

In this work, we have scanned 8221 K^+ -contained compounds from ICSD to find the proper electrode material for ammonium-ion batteries. The compounds with the exchange of K^+ ions with NH_4^+ ions are used as the input of the following screening



process. After the BVEL diffusion barrier filter, general formula filter, OCV screening, specific capacity screening, and volume expansion rate screening, we have successfully predicted 27 potential fast ammonium ion conductors for AIBs, all of which can stimulate further experimental studies and thorough theoretical investigations.

Our research is not exhaustive due to the rich structural chemistry of these filtered compounds, it gives an idea of the research interest in electrode materials for ammonium-ion batteries applications. Due to the simplicity of the BVEL method, its straightforward application to high-throughput schemes has some shortcomings. The empirical parameters of the BVEL method for NH_4^+ ions are incomplete. This is the main factor that causes the miss. Although the ionic radii and valence of K^+ and NH_4^+ ions are similar, the stability analysis after the replacement would probably be insufficient, and additional phonon dispersion calculation would be required. Structures with partial occupations need to be properly treated, which requires more complex methodological architectures. Continued effort is undertaken to expand the databases used for screening and solving associated algorithmic and computational problems.

Data availability

All data generated or analyzed during this study are included in this article and its ESI.†

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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References

- R. Xiao, H. Li and L. Chen, High-throughput design and optimization of fast lithium ion conductors by the combination of bond-valence method and density functional theory, *Sci. Rep.*, 2015, **5**, 14227.
- V. Palomares, P. Serras, I. Villaluenga, K. B. Hueso, J. Carretero-González and T. Rojo, Na-ion batteries, recent advances and present challenges to become low cost energy storage systems, *Energy Environ. Sci.*, 2012, **5**, 5884–5901.
- X. Wu, D. P. Leonard and X. Ji, Emerging Non-Aqueous Potassium-Ion Batteries: Challenges and Opportunities, *Chem. Mater.*, 2017, **29**, 5031–5042.
- R. Emanuelsson, M. Sterby, M. Strømme and M. Sjödin, An All-Organic Proton Battery, *J. Am. Chem. Soc.*, 2017, **139**, 4828–4834.
- Y. Wang and S. F. Kuchena, Recent Progress in Aqueous Ammonium-Ion Batteries, *ACS Omega*, 2022, **7**, 33732–33748.
- X. Zhao, S. Ren, M. Bruns and M. Fichtner, Chloride ion battery: a new member in the rechargeable battery family, *J. Power Sources*, 2014, **245**, 706–711.
- V. Sidey, On the effective ionic radii for ammonium, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2016, **72**, 626–633.
- C. D. Wessells, S. V. Peddada, M. T. McDowell, R. A. Huggins and Y. Cui, The Effect of Insertion Species on Nanostructured Open Framework Hexacyanoferrate Battery Electrodes, *J. Electrochem. Soc.*, 2011, **159**, A98–A103.
- X. Wu, Y. Qi, J. J. Hong, Z. Li, A. S. Hernandez and X. Ji, Rocking-Chair Ammonium-Ion Battery: A Highly Reversible Aqueous Energy Storage System, *Angew. Chem., Int. Ed.*, 2017, **56**, 13026–13030.
- G. Liang, F. Mo, X. Ji and C. Zhi, Non-metallic charge carriers for aqueous batteries, *Nat. Rev. Mater.*, 2020, **6**, 109–123.
- S. Dong, W. Shin, H. Jiang, X. Wu, Z. Li, J. Holoubek, W. F. Stickle, B. Key, C. Liu, J. Lu, P. A. Greaney, X. Zhang and X. Ji, Ultra-fast NH_4^+ Storage: Strong H Bonding between NH_4^+ and Bi-layered V_2O_5 , *Chem*, 2019, **5**, 1537–1551.
- J. Han, A. Varzi and S. Passerini, The Emergence of Aqueous Ammonium-Ion Batteries, *Angew. Chem., Int. Ed.*, 2022, **61**, e202115046.
- H. Li, J. Yang, J. Cheng, T. He and B. Wang, Flexible aqueous ammonium-ion full cell with high rate capability and long cycle life, *Nano Energy*, 2020, **68**, 104369.
- P. Wang, Y. Zhang, Z. Feng, Y. Liu and C. Meng, A dual-polymer strategy boosts hydrated vanadium oxide for ammonium-ion storage, *J. Colloid Interface Sci.*, 2022, **606**, 1322–1332.
- X. Wu, Y. Xu, H. Jiang, Z. Wei, J. J. Hong, A. S. Hernandez, F. Du and X. Ji, NH_4^+ Topotactic Insertion in Berlin Green: An Exceptionally Long-Cycling Cathode in Aqueous Ammonium-Ion Batteries, *ACS Appl. Energy Mater.*, 2018, **1**, 3077–3083.
- M. R. Lukatskaya, O. Mashtalir, C. E. Ren, Y. Dall'Agnese, P. Rozier, P. L. Taberna, M. Naguib, P. Simon, M. W. Barsoum and Y. Gogotsi, Cation Intercalation and High Volumetric Capacitance of Two-Dimensional Titanium Carbide, *Science*, 2013, **341**, 1502–1505.
- K. Itaya, T. Ataka and S. Toshima, Spectroelectrochemistry and electrochemical preparation method of Prussian blue modified electrodes, *J. Am. Chem. Soc.*, 1982, **104**, 4767–4772.
- S. F. Kuchena and Y. Wang, Superior Polyaniline Cathode Material with Enhanced Capacity for Ammonium Ion Storage, *ACS Appl. Energy Mater.*, 2020, **3**, 11690–11698.
- Y. Song, Q. Pan, H. Lv, D. Yang, Z. Qin, M. Y. Zhang, X. Sun and X. X. Liu, Ammonium-Ion Storage Using Electrodeposited Manganese Oxides, *Angew. Chem., Int. Ed.*, 2021, **60**, 5718–5722.



- 20 W. Xu, L. Zhang, K. Zhao, X. Sun and Q. Wu, Layered ferric vanadate nanosheets as a high-rate NH_4^+ storage electrode, *Electrochim. Acta*, 2020, **360**, 137008.
- 21 A. Gandomi and M. Haider, Beyond the hype: big data concepts, methods, and analytics, *J. Int. Inf. Manag.*, 2015, **35**, 137–144.
- 22 M. J. Shaw, C. Subramaniam, G. W. Tan and M. E. Welge, Knowledge management and data mining for marketing, *Decis Support Syst*, 2001, **31**, 127–137.
- 23 A. Belsky, M. Hellenbrandt, V. L. Karen and P. Luksch, New developments in the Inorganic Crystal Structure Database (ICSD): accessibility in support of materials research and design, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2002, **58**, 364–369.
- 24 S. Grazulis, D. Chateigner, R. T. Downs, A. F. T. Yokochi, M. Quiros, L. Lutterotti, E. Manakova, J. Butkus, P. Moeck and A. Le Bail, Crystallography Open Database – an open-access collection of crystal structures, *J. Appl. Crystallogr.*, 2009, **42**, 726–729.
- 25 B. Hourahine, B. Aradi, V. Blum, F. Bonafe, A. Buccheri, C. Camacho, C. Cevallos, M. Y. Deshayé, T. Dumitrica, A. Dominguez, S. Ehlert, M. Elstner, T. van der Heide, J. Hermann, S. Irle, J. J. Kranz, C. Kohler, T. Kowalczyk, T. Kubar, I. S. Lee, V. Lutsker, R. J. Maurer, S. K. Min, I. Mitchell, C. Negre, T. A. Niehaus, A. M. N. Niklasson, A. J. Page, A. Pecchia, G. Penazzi, M. P. Persson, J. Rezac, C. G. Sanchez, M. Sternberg, M. Stohr, F. Stuckenberg, A. Tkatchenko, V. W. Yu and T. Frauenheim, DFTB+, a software package for efficient approximate density functional theory based atomistic simulations, *J. Chem. Phys.*, 2020, **152**, 124101.
- 26 A. O. Boev, S. S. Fedotov, K. J. Stevenson and D. A. Aksyonov, High-throughput computational screening of cathode materials for Li–O₂ battery, *Comput. Mater. Sci.*, 2021, **197**, 110592.
- 27 M. Avdeev, M. Sale, S. Adams and R. P. Rao, Screening of the alkali-metal ion containing materials from the Inorganic Crystal Structure Database (ICSD) for high ionic conductivity pathways using the bond valence method, *Solid State Ionics*, 2012, **225**, 43–46.
- 28 A. G. Volkov, S. Paula and D. W. Deamer, Two mechanisms of permeation of small neutral molecules and hydrated ions across phospholipid bilayers, *Bioelectrochem. Bioenerg.*, 1997, **42**, 153–160.
- 29 I. D. Brown, Recent developments in the methods and applications of the bond valence model, *Chem. Rev.*, 2009, **109**, 6858–6919.
- 30 L. L. Wong, K. C. Phuah, R. Dai, H. Chen, W. S. Chew and S. Adams, Bond Valence Pathway Analyzer—An Automatic Rapid Screening Tool for Fast Ion Conductors within softBV, *Chem. Mater.*, 2021, **33**, 625–641.
- 31 S. Adams and R. P. Rao, Understanding Ionic Conduction and Energy Storage Materials with Bond-Valence-Based Methods, *Bond Valences*, 2014, pp. 129–159.
- 32 S. Adams and R. P. Rao, High power lithium ion battery materials by computational design, *Phys. Status Solidi A*, 2011, **208**, 1746–1753.
- 33 S. Adams, From bond valence maps to energy landscapes for mobile ions in ion-conducting solids, *Solid State Ionics*, 2006, **177**, 1625–1630.
- 34 S. S. Fedotov, N. A. Kabanova, A. A. Kabanov, V. A. Blatov, N. R. Khasanova and E. V. Antipov, Crystallochemical tools in the search for cathode materials of rechargeable Na-ion batteries and analysis of their transport properties, *Solid State Ionics*, 2018, **314**, 129–140.
- 35 I. D. Brown, Recent developments in the methods and applications of the bond valence model, *Chem. Rev.*, 2009, **109**(12), 6858–6919.
- 36 S. Adams, *softBV ver. 0.96*, 2004.
- 37 S. Adams, *Practical Considerations in Determining Bond Valence Parameters*, 2013, pp. 91–128.
- 38 N. A. Katcho, J. Carrete, M. Reynaud, G. Rousse, M. Casas-Cabanas, N. Mingo, J. Rodríguez-Carvajal and J. Carrasco, An investigation of the structural properties of Li and Na fast ion conductors using high-throughput bond-valence calculations and machine learning, *J. Appl. Crystallogr.*, 2019, **52**, 148–157.
- 39 W.-J. Zhang, Structure and performance of LiFePO₄ cathode materials: a review, *J. Power Sources*, 2011, **196**, 2962–2970.
- 40 C. Li, H. P. Zhang, L. J. Fu, H. Liu, Y. P. Wu, E. Rahm, R. Holze and H. Q. Wu, Cathode materials modified by surface coating for lithium ion batteries, *Electrochim. Acta*, 2006, **51**, 3872–3883.
- 41 G. Kresse and J. Furthmüller, Efficient iterative schemes for *ab initio* total-energy calculations using a plane-wave basis set, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 42 P. E. Blöchl, Projector augmented-wave method, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.
- 43 S. P. Ong, W. D. Richards, A. Jain, G. Hautier, M. Kocher, S. Cholia, D. Gunter, V. L. Chevrier, K. A. Persson and G. Ceder, Python Materials Genomics (pymatgen): A robust, open-source python library for materials analysis, *Comput. Mater. Sci.*, 2013, **68**, 314–319.
- 44 A. Jain, S. Ong, G. Hautier, W. Chen, W. Richards, S. Dacek, S. Cholia, D. Gunter, D. Skinner, G. Ceder and K. Persson, Commentary: The Materials Project: A materials genome approach to accelerating materials innovation, *APL Mater.*, 2013, **1**, 011002.
- 45 J. P. Perdew, K. Burke and M. Ernzerhof, Generalized Gradient Approximation Made Simple, *Phys. Rev. Lett.*, 1996, **77**, 3865; *Phys. Rev. Lett.*, 1997, **78**, 1396.
- 46 D. A. Blom and T. Vogt, Probing Compositional Order in Atomic Columns: STEM Simulations Beyond the Virtual Crystal Approximation, *Microsc. Microanal.*, 2020, **26**, 46–52.
- 47 M. K. Aydinol, A. F. Kohan, G. Ceder, K. Cho and J. Joannopoulos, *Ab initio* study of lithium intercalation in metal oxides and metal dichalcogenides, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1997, **56**, 1354–1365.
- 48 X. Rocquefelte, F. Boucher, P. Gressier and G. Ouvrard, First-Principle Study of the Intercalation Process in the Li_xV₂O₅ System, *Chem. Mater.*, 2003, **15**, 1812–1819.
- 49 Z. Rong, R. Malik, P. Canepa, G. Sai Gautam, M. Liu, A. Jain, K. Persson and G. Ceder, Materials Design Rules for



- Multivalent Ion Mobility in Intercalation Structures, *Chem. Mater.*, 2015, **27**, 6016–6021.
- 50 Y. Mo, S. P. Ong and G. Ceder, Insights into Diffusion Mechanisms in P2 Layered Oxide Materials by First-Principles Calculations, *Chem. Mater.*, 2014, **26**, 5208–5214.
- 51 O. A. Drozhzhin, V. V. Grigoryev, A. M. Alekseeva, R. R. Samigullin, D. A. Aksyonov, O. V. Boytsova, D. Chernyshov, V. V. Shapovalov, A. A. Guda, A. V. Soldatov, K. J. Stevenson, A. M. Abakumov and E. V. Antipov, Revisited $\text{Ti}_2\text{Nb}_2\text{O}_9$ as an Anode Material for Advanced Li-Ion Batteries, *ACS Appl. Mater. Interfaces*, 2021, **13**, 56366–56374.
- 52 A. D. Dembitskiy, D. A. Aksyonov, A. M. Abakumov and S. S. Fedotov, NH_4^+ -based frameworks as a platform for designing electrodes and solid electrolytes for Na-ion batteries: a screening approach, *Solid State Ionics*, 2022, **374**, 115810.
- 53 L. Zheng, Y. Xu, D. Jin and Y. Xie, Novel Metastable Hexagonal MoO_3 Nanobelts: Synthesis, Photochromic, and Electrochromic Properties, *Chem. Mater.*, 2009, **21**, 5681–5690.
- 54 G. Liang, Y. Wang, Z. Huang, F. Mo, X. Li, Q. Yang, D. Wang, H. Li, S. Chen and C. Zhi, Initiating Hexagonal MoO_3 for Superb-Stable and Fast NH_4^+ Storage Based on Hydrogen Bond Chemistry, *Adv. Mater.*, 2020, **32**, e1907802.
- 55 F. Zhou, K. Kang, T. Maxisch, G. Ceder and D. Morgan, The electronic structure and band gap of LiFePO_4 and LiMnPO_4 , *Solid State Commun.*, 2004, **132**, 181–186.

