




 Cite this: *RSC Adv.*, 2023, **13**, 17914

Vertically aligned MoS₂ nanosheets on monodisperse MXene as electrolyte-philic cathodes for zinc ion batteries with enhanced capacity†

 Wanting Su,[‡] Man Lang,[‡] Qingxiao Zhang, Yanan Yang, Huili Li * and Fang Zhang *

Zinc ion batteries (ZIBs) have attracted extensive attention for their high safety and environmentally friendly nature, and considerable theoretical capacities. Due to its unique two-dimensional layered structure and high theoretical specific capacities, molybdenum disulfide (MoS₂) presents as a promising cathode material for ZIBs. Nevertheless, the low electrical conductivity and poor hydrophilicity of MoS₂ limits its wide application in ZIBs. In this work, MoS₂/Ti₃C₂T_x composites are effectively constructed using a one-step hydrothermal method, where two-dimensional MoS₂ nanosheets are vertically grown on monodisperse Ti₃C₂T_x MXene layers. Contributing to the high ionic conductivity and good hydrophilicity of Ti₃C₂T_x, MoS₂/Ti₃C₂T_x composites possess improved electrolyte-philic and conductive properties, leading to a reduced volume expansion effect of MoS₂ and accelerated Zn²⁺ reaction kinetics. As a result, MoS₂/Ti₃C₂T_x composites exhibit high voltage (1.6 V) and excellent discharge specific capacity of 277.8 mA h g⁻¹ at 0.1 A g⁻¹, as well as cycle stability as cathode materials for ZIBs. This work provides an effective strategy for developing cathode materials with high specific capacity and stable structure.

 Received 10th April 2023
 Accepted 29th May 2023

DOI: 10.1039/d3ra02352d

rsc.li/rsc-advances

Introduction

In recent years, with the rapid development of the economy and increasing energy consumption, high-capacity batteries are highly desired for energy storage purposes.^{1–3} In these respects, lithium-ion batteries have been widely used in electronic products and transportation industries due to their high energy density.^{1,4,5} However, lithium-ion batteries suffer from the problems of high production cost and intrinsic safety issues. Extensive research has been conducted on metal ion batteries (such as Mg²⁺, Zn²⁺, Na⁺, etc.), which are abundant in resources and low cost.^{1,6–10} Aqueous zinc ion batteries (ZIBs) are considered to be ideal candidates over lithium-ion batteries due to their low redox potential (−0.763 V vs. standard hydrogen electrode (SHE)),^{6,11–13} high theoretical specific capacity (820 mA h g⁻¹ and 5855 mA h cm⁻³),^{14,15} non-toxic properties⁶ and abundant Zn resources.⁷ A typical ZIB generally compose of Zn metal anode, electrolyte, and cathode electrode for accommodating Zn ions.¹⁶ Except for the most common anode material of zinc metal in ZIBs, battery capacities are more limited

by cathodes than anodes. With regards to the cathode materials, extensively investigated active materials include manganese-based materials, vanadium-based materials, Prussian blue analogs, and organic electrode materials and so on.^{17–21} However, the electrochemical and physical properties of these active materials are still inadequate to achieve high-performance ZIBs.²² It is urgent to develop cathode materials for ZIBs with high specific capacity and stable structures for large-scale commercial applications.^{23,24}

Molybdenum disulfide (MoS₂), a representative transition metal dichalcogenide, possesses the two-dimensional (2D) layered structure similar to graphite.^{25–27} The MoS₂ is considered as promising host material for ion intercalation due to the unique layered structure bonded by weak van der Waals' forces.^{2,28,29} These characteristics indicate that the layered materials MoS₂ demonstrate extraordinary potential as electrode materials for ZIBs.^{30,31} However, the low intrinsic conductivity, poor structural stability, and severe self-stacking of MoS₂ hinder its wide application of MoS₂ in zinc storage materials.² In addition, the larger hydrated ionic diameter of Zn²⁺ (approximately 0.86 nm) and strong electrostatic interactions with MoS₂ result in sluggish reaction kinetics and difficultly reversible Zn²⁺ intercalation.^{15,32–34} To alleviate these problems, several modification methods have been proposed to improve the electrochemical performance of ZIBs using MoS₂, including the introduction of sulfur defects,³⁵ partial replacement of sulfur atoms by oxygen atoms,³⁶ modulation of

Key Laboratory of Resource Chemistry, Joint International Research Laboratory of Resource Chemistry, Ministry of Education, College of Chemistry and Materials Science, Shanghai Normal University, Shanghai 200234, China. E-mail: li_huili@shnu.edu.cn; zhangfang@shnu.edu.cn

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3ra02352d>

‡ Wanting Su and Man Lang contributed equally to this work.



phases,³³ and expansion of layer spacing.³⁷ For instance, Wang preferentially inserted Zn ions into sulfur vacancies building defect engineered MoS_{2-x} nanosheets.³⁵ The ZIBs based on MoS_{2-x} nanosheets exhibited discharge specific capacity of 138.6 mA h g⁻¹ at 0.1 A g⁻¹. In another work, molecular engineering in terms of structure defects manufacturing and O-doping on MoS₂ was reported by Li to expand the interlayer spacing (from 6.2 to 9.6 Å).²⁶ Therefore, the tailored MoS₂-O cathode respectively displayed specific capacity of 261 mA h g⁻¹ at 0.1 A g⁻¹ with 102.4 mA h g⁻¹ at 10 A g⁻¹, indicating unprecedentedly high rate-capability. In addition, due to metallic 1 T MoS₂ with better electronic conductivity comparable than 2H MoS₂, vertically aligned 1 T phase MoS₂ nanosheet array on carbon cloth were fabricated to improve capacities.³³ As a result, the free-standing cathode exhibited excellent specific capacity of 198 mA h g⁻¹ at current density of 0.1 A g⁻¹. Most recently, MoS₂/polyaniline (PANI) hybrid with inter-overlapped heterostructure was explored with expanded interlayer spacing as large as 1.03 nm of (002) plane in the MoS₂/PANI hybrid.³⁷ The MoS₂/PANI hybrid delivered high reversible capacity of 106.5 mA h g⁻¹ at a current density of 1.0 A g⁻¹. Nonetheless, the conductivity, swelling effect and self-aggregation phenomenon still need to be further improved to obtain high capacity and stable cycling ZIBs.

Ti₃C₂T_x MXene, typical 2D transition metal carbonitride, exhibits excellent properties, such as abundant surface groups,³⁸ tunable interlayer spacing,³⁹ excellent electrical conductivity⁴⁰ and good hydrophilicity.⁴¹ Ti₃C₂T_x MXene has been widely developed and investigated for energy storage. Herein, the MoS₂/Ti₃C₂T_x composites were constructed through the 2D MoS₂ nanosheets vertically aligned on 2D Ti₃C₂T_x MXene by a one-step hydrothermal method. The resulting 3D open network architecture of the MoS₂/Ti₃C₂T_x composites increased the interfacial contact with the electrolyte and shortened the diffusion pathways for Zn²⁺. Due to the high electrical conductivity and hydrophilicity of monodisperse Ti₃C₂T_x MXene, the introducing of Ti₃C₂T_x suppressed self-stacking of MoS₂ nanosheets, accelerated the migration rate of Zn²⁺ and enhanced the reaction kinetic rate. Impressively, the as-fabricated MoS₂/Ti₃C₂T_x composites exhibited high voltage to 1.6 V, high discharge specific capacity of 277.8 mA h g⁻¹ at 0.1 A g⁻¹ and 80% capacity retention after 5000 cycles when the current density was 10.0 A g⁻¹. In addition, the intercalation/deintercalation mechanism and kinetic diffusion process of Zn ion diffusion were demonstrated using *ex situ* X-ray diffraction (ex-XRD), X-ray photoelectron spectroscopy (XPS) and galvanostatic intermittent titration technique (GITT). This work presented new insights on the application of layered materials as high-performance cathodes for ZIBs.

Experimental

Synthesis of Ti₃C₂T_x MXene

As previously reported, Ti₃C₂T_x was obtained after selective etching of the Al layer from Ti₃AlC₂ feedstock.^{42,43} First, 3.2 g LiF was dissolved in 40 mL HCl (9 mol L⁻¹), stirred for five minutes and then 2 g Ti₃AlC₂ was slowly added to the above solution and

the mixed solution was stirred at 35 °C for 24 h. The obtained mixed solution was washed with deionized water by centrifugation to a supernatant with pH of approximately 6, and the precipitate was sonicated with an appropriate amount of deionized water under argon for 1 h. Finally, the Ti₃C₂T_x dispersion was centrifuged at 3500 rpm for 20 min to obtain Ti₃C₂T_x dispersion.

Synthesis of MoS₂/Ti₃C₂T_x MXene composites

In brief, 0.3 mmol ammonium molybdate and 13.1 mmol thiourea were dissolved in 50 mL water, and then 10 mg Ti₃C₂T_x (5 mg mL⁻¹) was added to the mixed solution, which was stirred for 15 min at room temperature and sonicated for 5 min. The mixed solution was transferred to Teflon stainless steel kettles and reacted at 200 °C for 24 h. The resulting product was then washed three times with deionized water and freeze-dried for two days. In contrast, MoS₂ nanosheets were synthesized without the addition of Ti₃C₂T_x. All the materials were used directly for electrochemical experiments without further purification.

Characterization methods

XRD were obtained using a Rigaku D/Max B diffractometer with Cu-K α radiation. The morphology and microstructure of the samples were observed by field emission scanning electron microscopy (SEM, HITACHI, S-4800) and transmission electron microscopy (TEM, JEOL JEM-2100), respectively. XPS were measured by using Thermo Scientific Nexsa system to analysis the chemical composition of the material surface. The atomic force microscope (AFM) tested the sheet thickness of Ti₃C₂T_x MXene using Bruker Dimension Icon. The specific surface area was tested with Tri Star II 3020 by Michael Corporation, USA. Raman spectroscopy were performed with a DXRZXi model (Thermo, USA).

Electrochemical characterizations

Electrochemical performance tests were performed in a CR2032 coin cell with MoS₂/Ti₃C₂T_x as the cathode, 0.1 mm zinc foil as the anode, 21 mol L⁻¹ lithium bis(trifluoromethane sulfonyl) imide salt (LiTFSI, Innochem, 98%) and 1 mol L⁻¹ Zn(CF₃SO₃)₂ (TCL, >98%) as the electrolyte, and a glass fiber filter membrane (Whatman, GF/D) as the septum. The MoS₂/Ti₃C₂T_x cathode was prepared by first grinding the active material, carbon black, and polyvinylidene fluoride (PVDF, Kejing Co. Shenzhen) in a 7 : 2 : 1 ratio in a mortar and pestle for thirty minutes, then pouring the solid powder into a small glass vial and adding an appropriate amount of *N*-methylpyrrolidone (NMP, Aladdin, >99%) solvent and stirring for 6 h. Finally, the mixed slurry was evenly coated on the titanium foil substrate and dried under vacuum at 80 °C for 12 h. The dried poles were cut into $\Phi = 12$ mm discs, and the average active substance loading of the prepared cathodes was 1.0–1.5 mg cm⁻². The cell multiplicity performance, cycling performance and GITT tests were carried out with the LAND 2001A cell test system in the voltage range of 0–1.6 V. Cyclic voltammetry (CV), electrochemical impedance spectrogram (EIS), and galvanostatic charge/discharge (GCD) were tested



using the same electrochemical workstation (CHI 760E, C&H Co., Shanghai, China), and the frequency range of EIS spectrogram was 10^{-2} – 10^5 . The above tests were performed at room temperature.

Results and discussion

Synthesis and characterization of MoS₂/Ti₃C₂T_x composites

The MoS₂/Ti₃C₂T_x MXene composites were synthesized *via* hydro-thermal reaction of (NH₄)₆Mo₇O₂₄·4H₂O and NH₂CSNH₂ in presence of Ti₃C₂T_x MXene as illustrated in Fig. 1a. Prior to the hydro-thermal reaction, homogeneous single-layer Ti₃C₂T_x MXene dispersions with thickness of 2.2 nm (Fig. S1 and S2a†) were obtained through LiF and HCl etching of Ti₃AlC₂ MAX (Fig. S2b†), followed by ultrasonic process. The Raman spectra (Fig. 1b) and XRD patterns (Fig. S3–S4†) were recorded to understand the characteristic of MoS₂/Ti₃C₂T_x MXene composites. According to the XRD pattern, the characteristic peaks of the (002) plane at 2θ values of 9.6° for Ti₃AlC₂ MAX shifted to 7.2° for Ti₃C₂T_x MXene, which was remarkably shifted to 10.0° of MoS₂/Ti₃C₂T_x MXene (Fig. S3†). In comparison to pure MoS₂, the diffraction peak at the (002) plane of MoS₂/Ti₃C₂T_x MXene was slightly shifted from 13.6° to 12.6°. Based on the Bragg's equation ($n\lambda = 2d \sin \theta$), the interlayer spacing of MoS₂ was modified 0.65 nm of pure MoS₂ from to 0.70 nm in the MoS₂/Ti₃C₂T_x MXene composites. Furthermore, the characteristic peaks at 143.1 cm⁻¹ (*J*₁), 232.6 cm⁻¹ (*J*₂), and 332.5 cm⁻¹ (*J*₃) demonstrated the successful construction of dominating 1T MoS₂ in heterostructure MoS₂/Ti₃C₂T_x MXene (Fig. 1b). The peak located at 278.2 cm⁻¹ was assigned to the E_{1g}

characteristic vibration of Mo in octahedral coordination, where the surface metal Mo atoms were connected to S atoms.

The chemical composition and valence states of MoS₂/Ti₃C₂T_x MXene composite were further analyzed using XPS spectra. The presence of Mo, S, Ti, C and O elements were confirmed in MoS₂/Ti₃C₂T_x MXene composites (Fig. S5a†). In Ti 2p spectrum (Fig. S5b†), the signals at 458.3 and 463.7 eV were identified as the 2p_{3/2} and 2p_{1/2} orbits of Ti³⁺ in Ti₃C₂T_x MXene, respectively. The absence of 2p_{3/2} and 2p_{1/2} doublet at 458.7 and 464.6 eV indicated the absence of the Ti–O bonds in Ti₃C₂T_x MXene and surface-attached oxygen containing groups. Deposited in high-resolution S 2p spectrum (Fig. 1c), the S 2p_{3/2} and S 2p_{1/2} spectra were also deconvoluted into 1T and 2H phases. The peaks at binding energies of 162.6 eV and 163.9 eV were severally assigned to S 2p_{1/2} and S 2p_{3/2} of 1T phases, indicating the presence of divalent sulfur ions. The Mo⁶⁺ 3d_{3/2}, Mo⁴⁺ 3d_{3/2} and Mo⁴⁺ 3d_{5/2} of 1T phases were characterized at 236.5 eV, 232.3 eV and 229.2 eV, respectively (Fig. 1d). The Mo in hierarchical structure of MoS₂/Ti₃C₂T_x MXene composites were dominant in Mo⁴⁺ and minimal in the oxidation of Mo⁶⁺. Furthermore, different from the commercial MoS₂ dominant in 2H phases (Fig. S6†), the MoS₂/Ti₃C₂T_x MXene composites mainly consisted of 1T phase (about 66%), which was consistent with Raman spectra (Fig. 1b).

The morphology and structure of MoS₂/Ti₃C₂T_x MXene composites were characterized using Scanning Electron Microscope (SEM), Energy Dispersive X-ray Spectroscopy (EDX) and Transmission Electron Microscope (TEM). The MoS₂/Ti₃C₂T_x MXene composites exhibited flower-like structures with evident wrinkles and corrugations (Fig. 2a–c). Due to the

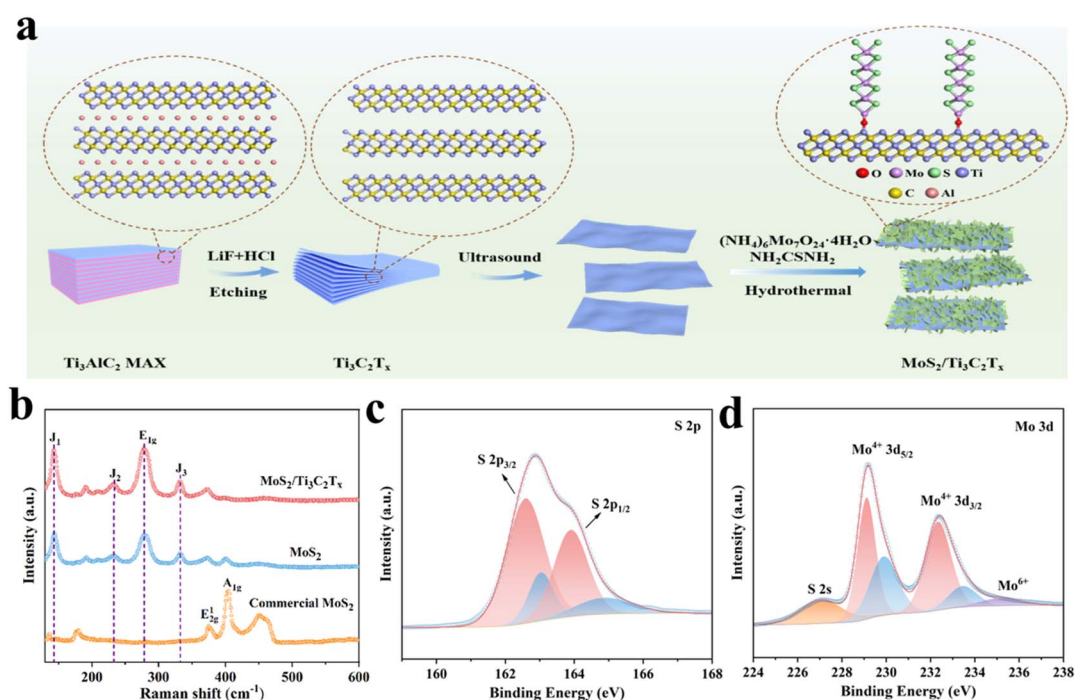


Fig. 1 (a) Schematic synthesis of the MoS₂/Ti₃C₂T_x composites. (b) Raman spectra of commercial MoS₂, pure MoS₂ and MoS₂/Ti₃C₂T_x composites. Detailed XPS spectra of (c) S 2p and (d) Mo 3d.



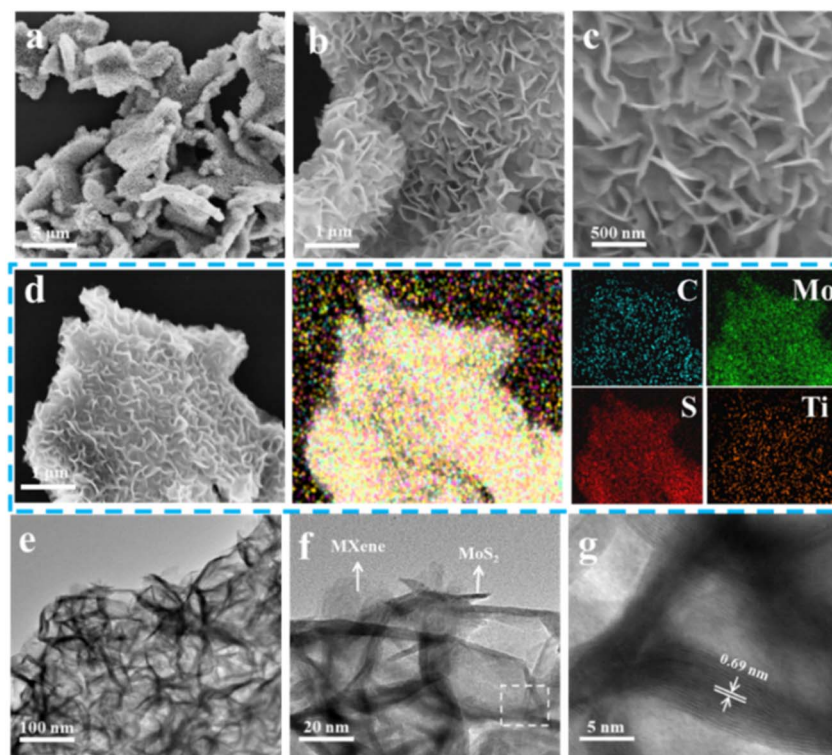


Fig. 2 (a–c) SEM images, (d) element mapping and (e–g) TEM images of $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene composites.

existence of $\text{Ti}_3\text{C}_2\text{T}_x$ MXene skeleton, heterogeneous nucleation avoided the agglomeration of MoS_2 nanosheets. Moreover, MoS_2 nanosheets were vertically aligned on the $\text{Ti}_3\text{C}_2\text{T}_x$ MXene surface, displaying uniform distribution of C, Mo, S and Ti elements (Fig. 2d). The hierarchical structure of MoS_2 nanosheets on the highly conductive $\text{Ti}_3\text{C}_2\text{T}_x$ MXene substrate facilitated fast electron transfer and easy penetration of the electrolyte, resulting in considerable electrochemical properties. Furthermore, the staggered MoS_2 nanosheets with low stacking height, providing lowered the Zn^{2+} migration potential barrier, provided more active sites for the intercalation of zinc ions.⁴⁴ As elucidated in Fig. 2e–g, the fabricated MoS_2 nanosheets exhibited the slightly modified interlayer spacing of 0.7 nm compared to commercial 2H MoS_2 (0.65 nm, Fig. S7†), which was well in accordance with XRD data (Fig. 1b). Due to the reduced ion diffusion resistance and energy barrier, the enlarged interlayer spacing accelerated the reaction kinetics of Zn^{2+} for $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene composites.

Electrochemical performance of $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ composites for ZIBs

$\text{Zn}/\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene batteries were constructed utilizing $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene as cathode and zinc as anode with electrolyte consisted of 21 mol L^{-1} LiTFSI and 1 mol L^{-1} $\text{Zn}(\text{CF}_3\text{SO}_3)_2$, respectively (Fig. 3a). The electrochemical performance of $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene composites was investigated through GCD and rate capability of $\text{Zn}/\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene batteries between 0 V and 1.6 V as demonstrated in Fig. 3b and c. $\text{Zn}/\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ batteries displayed specific discharge capacities of 277.8, 228.5, 175.6, 138.9 and 106.4 mA h g^{-1} at current

densities of 0.1, 0.2, 0.5, 1.0, and 2.0 A g^{-1} , respectively. The specific discharge capacities were found to be higher than that of MoS_2 (Fig. 3c and S8†), MoS_2 with 1 mol L^{-1} $\text{Zn}(\text{CF}_3\text{SO}_3)_2$ electrolyte (Fig. S9†) and other MoS_2 -based ZIBs as presented in Fig. 3d and Table S1.†^{31–35,37,45–53} This was mainly attributed to the increased specific surface area of $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene composites in Fig. S10.† When the current density returned to 0.1 A g^{-1} , the corresponding discharge capacity recovered to approximately 233.0 mA h g^{-1} maintaining about 84% of the original discharge capacity. This indicated the good electrochemical reversibility and fast reaction kinetics of the prepared ZIBs. Furthermore, the prepared ZIBs using $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ MXene composites exhibited a power density of 37.7 W kg^{-1} and an energy density of 108.3 W h kg^{-1} . Ascribing to the introduction of $\text{Ti}_3\text{C}_2\text{T}_x$ MXene, the vertically aligned heterostructure with expanded interlayer spacing and good electrolyte-philic properties (Fig. S11†) of MoS_2 generated the considerable electrochemical properties.

From the EIS spectra in Fig. 3e, Nyquist diagram possessed a semicircle in the high-frequency region and a diagonal line in the low-frequency region. The intercept on the real axis and diameter of the semicircle reflected the equivalent series resistance (R_e) and the charge-transfer resistance (R_{ct}), respectively.⁵⁴ The slope of diagonal line in the low-frequency region indicates Warburg resistance (Z_w) caused by ion diffusion in the electrode. The $\text{Zn}/\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ batteries displayed smaller R_{ct} (31.1 Ω) than that of Zn/MoS_2 batteries (114.7 Ω), attributed to the high conductivity and reduced the zinc ions intercalation energy barrier of the 2D/2D $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ heterostructure. Furthermore, $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ composites exhibited decreased Z_w



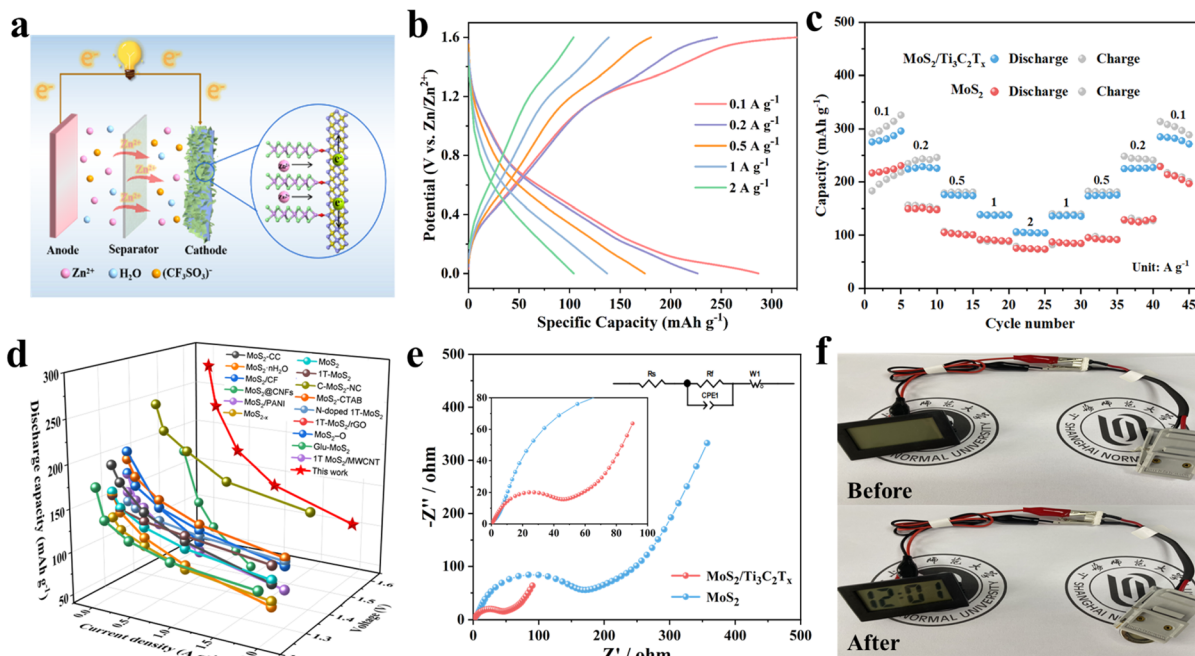


Fig. 3 Electrochemical performances of Zn//MoS₂/Ti₃C₂T_x MXene batteries in 21 mol L⁻¹ LiTFSI and 1 mol L⁻¹ Zn(CF₃SO₃)₂ aqueous electrolyte. (a) Schematics of the mechanism of Zn//MoS₂/Ti₃C₂T_x MXene battery. (b) The galvanostatic charge/discharge curves and (c) rate capability of Zn//MoS₂/Ti₃C₂T_x MXene batteries under various current densities. (d) Discharge capacity of ZIBs compared with the values reported of other MoS₂ based ZIBs. (e) Electrochemical impedance spectroscopy spectra of Zn//MoS₂/Ti₃C₂T_x and Zn//MoS₂ batteries. Inset: the partial enlargement EIS at the high frequency region. (f) Digital photographs of Zn//MoS₂/Ti₃C₂T_x MXene batteries to power a lock.

compared to MoS₂ and commercial MoS₂, signifying the faster diffusion rate of Zn²⁺ and improved ion conductivity with introduction of conductive Ti₃C₂T_x in Fig. S12.† Remarkably, an exceptionally long lifetime of 5000 cycles was obtained with stable capacity retention of over 80% when cycled at high current density of 10.0 A g⁻¹ (Fig. S13†), better than that of Zn//MoS₂ batteries using 1 mol L⁻¹ Zn(CF₃SO₃)₂ electrolyte (Fig. S14†). The MoS₂/Ti₃C₂T_x MXene heterostructure structure with expanded layer spacing was conducive to intercalation/delamination of zinc ion. The morphologies of MoS₂/Ti₃C₂T_x MXene nanosheets were well preserved without amount of aggregation after cycling measurements (Fig. S15†). Furthermore, the interlayer spacing of MoS₂/Ti₃C₂T_x MXene composites well maintained with obvious lattice fringes after long-term cycle test (Fig. S16†). Moreover, at low current density of 1.0 A g⁻¹, capacity can be maintained about 80% of initial capacity after 800 charge/discharge cycles (Fig. S17†), indicating excellent structural stability for ZIB. In order to explore the practical application value of ZIBs, one Zn//MoS₂/Ti₃C₂T_x MXene ZIB can illuminate a clock (Fig. 3f), revealing that the prepared ZIBs can be used as electrochemical storage device to applied other electronic products.

Kinetics of Zn//MoS₂/Ti₃C₂T_x ZIBs

The CV curve was obtained at different sweep rates within the voltage window of 0–1.6 V vs. Zn/Zn²⁺ in Fig. 4a. With the increase of scanning rate from 0.4 mV s⁻¹ to 1.2 mV s⁻¹, the area of the CV curve gradually increased along with obviously enhancement in current intensity of the redox peak. This was

because the diffusion resistance increased with the increase of sweep rate. At scanning rate of 0.1 mV s⁻¹, the couple cathode peak and anode peak severally represented at 0.27 V and 1.25 V, in accordance with the discharge and charge plateaus in GCD curves (Fig. 3b), corresponding to insertion/desertion of zinc ions. Based on the CV curves obtained at various rates, electrochemical kinetic processes of Zn//MoS₂/Ti₃C₂T_x batteries can be determined by the following relation:

$$i = av^b$$

where a and b were variable parameters, i and v represented peak current and the corresponding scan rate, respectively. In general, while the co-efficient b value equaled to 1 or 0.5, the electrochemical process was dominated by capacitive-limited controlled behavior or diffusion-limited processes, respectively. From the slope fitting curves in Fig. 4b, the b values of peaks 1 and 2 were separately calculated as 0.80 and 0.88, revealing that capacitive-limited process and diffusion-controlled behavior synergistically dominated the charge storage process of MoS₂/Ti₃C₂T_x composites. Moreover, the capacitive (k_1v) contribution and diffusion ($k_2v^{1/2}$) contribution of MoS₂/Ti₃C₂T_x composites can be further rationally quantified by the following equation:

$$i = k_1v + k_2v^{1/2}$$

where k_1 and k_2 were parameters changed with the peak current (i). Fig. 4c illustrated the ratio of the capacitive contribution (blue) to the diffusion contribution (red) as a function of scan



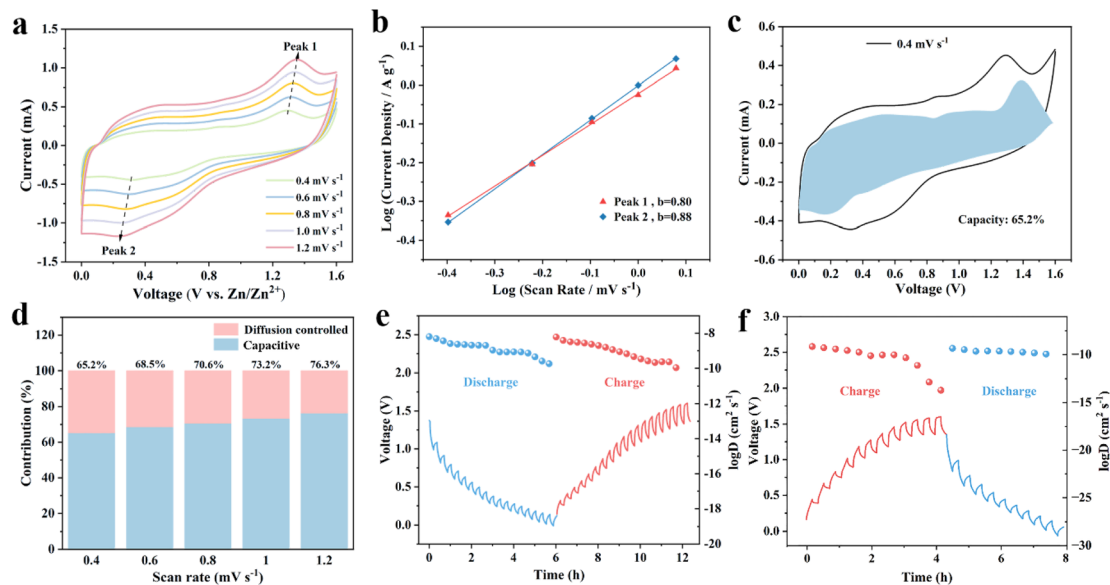


Fig. 4 (a) Cyclic voltammogram profiles of Zn//MoS₂/Ti₃C₂T_x batteries at various scan rates. (b) The fitting plots between log(*i*) and log(*v*) at various peak currents. (c) Quantification of the capacitive (blue area) and diffusion charge storage in the MoS₂/Ti₃C₂T_x electrode at a scan rate of 0.4 mV s⁻¹ and (d) the ratio of the capacitive contribution (blue) to the diffusion contribution (red) as a function of scan rate. Charge/discharge GITT curves and corresponding Zn²⁺ diffusion coefficients for (e) MoS₂/Ti₃C₂T_x and (f) MoS₂ at 0.5 A g⁻¹.

rate. The value of capacitive contribution exhibited nearly 65.2% at the scan rate of 0.4 mV s⁻¹ (Fig. 4d). This value steadily increased with increasing scan rates due to the slower lower capacitive-limited process compared to diffusion-controlled behavior.

The kinetics of Zn²⁺ intercalation during the charge/discharge reaction of MoS₂/Ti₃C₂T_x cathode material was studied in depth using the GITT. The diffusion coefficient of Zn²⁺ (*D*_{Zn²⁺}) was calculated using the following equation:

$$D_{\text{Zn}^{2+}} = \frac{4}{\pi\tau} \left(\frac{mV_m}{MA} \right)^2 \left(\frac{\Delta E_s}{\Delta E_\tau} \right)^2$$

where the variables τ , M , m , V_m and A severally represented the relaxation time, molar mass of active material, mass of active material, molar volume, and the contact area between electrode and electrolyte. Additionally, ΔE_s and ΔE_τ represented the pulse voltage change and voltage change of constant-current charge/discharge, respectively. In the voltage range of 0–1.6 V, the (*D*_{Zn²⁺}) of MoS₂/Ti₃C₂T_x MXene ZIBs were calculated to be in the range of 10⁻⁸ to 10⁻¹⁰ cm² s⁻¹ (Fig. 4e), which was superior to that of MoS₂ (10⁻⁹ to 10⁻¹⁴ cm² s⁻¹, Fig. 4f), and other previously reported modified-MoS₂ electrode materials (10⁻¹⁰ to 10⁻¹⁴ cm² s⁻¹).^{32,45} The improved performance of the Zn//MoS₂/Ti₃C₂T_x MXene ZIBs can be attributed to the introduction of Ti₃C₂T_x into MoS₂, which increased the contact area, shortened the transport distance of Zn²⁺ and generated the accelerated migration rate of Zn²⁺. The increase of charge transfer rate and the rapid diffusion of Zn²⁺ ensured the excellent rate performance and cycle stability of the Zn//MoS₂/Ti₃C₂T_x MXene ZIBs.

Mechanism of Zn//MoS₂/Ti₃C₂T_x ZIBs

Ex situ XRD characterization and XPS measurements were performed to investigate the Zn storage mechanism of MoS₂/

Ti₃C₂T_x electrode. Specifically, the charge–discharge marked states (points A–G) were selected for *ex situ* XRD testing (Fig. 5a), focusing on the structural evolution of MoS₂/Ti₃C₂T_x electrode. During the discharge process (A–D), the (002) characteristic peak of MoS₂/Ti₃C₂T_x electrode shifted to the lower angle, demonstrating that the insertion of Zn²⁺ ions with expanding the layer spacing of MoS₂/Ti₃C₂T_x electrode. After charging (D–G), the (002) peak position returned to the original point A, which was caused by the desertion of Zn²⁺ ions from the MoS₂/Ti₃C₂T_x electrode. The charge/discharge process was consistent with the CV curves (Fig. 4a). The results indicated that the structure of MoS₂/Ti₃C₂T_x possessed good reversibility without collapsing during the ion embedding-detachment process.

The reversible insertion/extraction mechanism of Zn²⁺ was also confirmed through XPS measurements (Fig. 5b and c). As expected, no diffraction signal of Zn 2p was detected in the pristine condition. Nevertheless, two remarkable diffraction peaks were observed at 1045.8 eV and 1022.7 eV while discharged, corresponding to Zn 2p_{1/2} and Zn 2p_{3/2}. The diffraction peak indicated the insertion of Zn²⁺ to the MoS₂/Ti₃C₂T_x electrode and the adsorption of Zn²⁺ in the electrolyte due to the capacitive behavior in consistent with the *b* value in CV measurement (Fig. 4d). Upon reverse charging, the inserted portion of Zn²⁺ disappeared from the diffraction peak but the adsorbent portion remained, indicating the reversible insertion/extraction of Zn²⁺. The effect of Zn²⁺ on the valence changes of Mo elements can be evidenced in the *ex situ* XPS patterns (Fig. 5c). With the insertion of Zn²⁺ at discharged state, the S 2s peak shifted by about 0.2 V to the lower binding energy and the intensity of Mo⁶⁺ was aggressively enhanced. Conversely, the S 2s peak returned to its original state and the intensity of Mo⁶⁺ significantly decreased at the discharged state. These results confirmed the zinc storage mechanism of MoS₂/



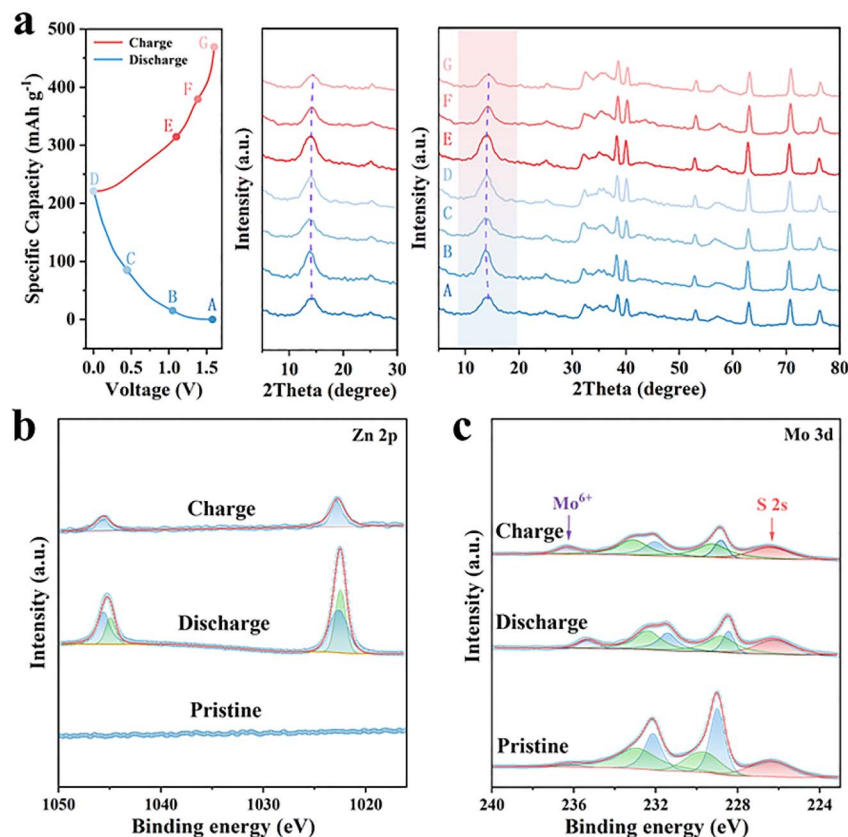
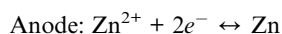
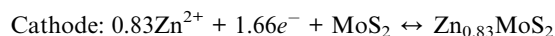


Fig. 5 Zn-storage mechanism of MoS₂/Ti₃C₂T_x composites. (a) Charge/discharge profiles of Zn//MoS₂/Ti₃C₂T_x battery at a current density of 0.1 A g⁻¹. The *ex situ* XRD patterns collected at different charge/discharge depths. The *ex situ* high-resolution XPS spectroscopy of (b) Zn 2p and (c) Mo 3d at different charge/discharge depths.

Ti₃C₂T_x electrode which agreed well with the *ex situ* XRD results (Fig. 5a).

As mentioned above, it can be calculated that 0.83 Zn²⁺ can be inserted into per molecule of MoS₂ at the current density of 0.1 A g⁻¹. Consequently, the electrochemical mechanism of the MoS₂/Ti₃C₂T_x battery can be illustrated and formulated as following:



During the discharged process, Zn anode lost the electron to be reduced to Zn²⁺. Simultaneously, Zn²⁺ ions were intercalated into the MoS₂ to establish the phase of Zn_{0.83}MoS₂ at the cathode. During the reverse charging, Zn²⁺ ions were gradually extracted from Zn_{0.83}MoS₂ to re-established the MoS₂ phase, while the Zn²⁺ gained the electron to deposit on the anode.

Conclusions

In summary, MoS₂ nanosheets were successfully vertically grown on monodisperse MXene composites through a one-step hydrothermal method. The electrical conductivity and

hydrophilicity in electrolyte of the composite were effectively improved compared with unmodified MoS₂. Meanwhile, the self-aggregation of MoS₂ and volume expansion during the charge/discharge reaction were significantly suppressed, resulting in excellent electrochemical properties with fast reaction kinetics. As a result, the MoS₂/Ti₃C₂T_x cathode demonstrated discharge specific capacity of up to 277.8 mA h g⁻¹ at 0.1 A g⁻¹ and the capacity retention of 80% after 5000 cycles when the current density was 10.0 A g⁻¹. This work provided more possibilities for improving the electrochemical performance of energy storage layered materials.

Author contributions

W. Su contributed to most experimental work. M. Lang completed the required experiments in the revised manuscript. Q. Zhang and Y. Yang assisted in characterization of some materials. H. Li conceived this idea and wrote the manuscript. F. Zhang conceived this idea and supervised the work. All the authors assisted in experiments and provided constructive comments on the text.

Conflicts of interest

There are no conflicts of interest.



Acknowledgements

This work is supported by The National Natural Science Foundation of China (22076122 and 52203269), The Shanghai Sailing Program (No. 20YF1435100), and Shanghai Government (22ZR1480200, 19SG42).

References

- 1 X. Jia, C. Liu, Z. G. Neale, J. Yang and G. Cao, *Chem. Rev.*, 2020, **120**, 7795–7866.
- 2 W. S. V. Lee, T. Xiong, X. Wang and J. Xue, *Small Methods*, 2021, **5**, 2000815.
- 3 L. Cao, D. Li, T. Deng, Q. Li and C. Wang, *Angew. Chem., Int. Ed.*, 2020, **132**, 19454–19458.
- 4 Y. Li, Z. Wang, Y. Cai, M. Pam, Y. Yang, D. Zhang, Y. Wang and S. Huang, *Energy Environ. Mater.*, 2022, **5**, 823–851.
- 5 T. Wang, C. Li, X. Xie, B. Lu, Z. He, S. Liang and J. Zhou, *ACS Nano*, 2020, **14**, 16321–16347.
- 6 D. Kundu, B. D. Adams, V. Duffort, S. H. Vajargah and L. F. Nazar, *Nat. Energy*, 2016, **1**, 1–8.
- 7 H. Li, L. Ma, C. Han, Z. Wang, Z. Liu, Z. Tang and C. Zhi, *Nano Energy*, 2019, **62**, 550–587.
- 8 A. Kanwade, S. Gupta, A. Kankane, M. K. Tiwari, A. Srivastava, J. A. K. Satrughna, S. C. Yadav and P. M. Shirage, *RSC Adv.*, 2022, **12**, 23284–23310.
- 9 S. Luo, L. Xie, F. Han, W. Wei, Y. Huang, H. Zhang, M. Zhu, O. G. Schmidt and L. Wang, *Adv. Funct. Mater.*, 2019, **29**, 1901336.
- 10 Z. Yuan, L. Wang, D. Li, J. Cao and W. Han, *ACS Nano*, 2021, **15**, 7439–7450.
- 11 N. Zhang, X. Chen, M. Yu, Z. Niu, F. Cheng and J. Chen, *Chem. Soc. Rev.*, 2020, **49**, 4203–4219.
- 12 N. Guo, W. Huo, X. Dong, Z. Sun, Y. Lu, X. Wu, L. Dai, L. Wang, H. Lin, H. Liu, H. Liang, Z. He and Q. Zhang, *Small Methods*, 2022, **6**, 2200597.
- 13 Y. Shi, Y. Chen, L. Shi, K. Wang, B. Wang, L. Li, Y. Ma, Y. Li, Z. Sun and W. Ali, *Small*, 2020, **16**, 2000730.
- 14 S. Li, S. Zhang, T. Feng, H. Zhou and M. Wu, *RSC Adv.*, 2022, **12**, 18466–18474.
- 15 S. Li, Y. Liu, X. Zhao, K. Cui, Q. Shen, P. Li, X. Qu and L. Jiao, *Angew. Chem., Int. Ed.*, 2021, **60**, 20286–20293.
- 16 M. Wu, G. Zhang, H. Yang, X. Liu, M. Dubois, M. A. Gauthier and S. Sun, *InfoMat*, 2022, **4**, e12265.
- 17 N. Subjalearndee, N. He, H. Cheng, P. Tesatchabut, P. Eiamlamai, P. Limthongkul, V. Intasanta, W. Gao and X. Zhang, *Adv. Fiber Mater.*, 2022, **4**, 457–474.
- 18 J. Xu, Y. Liu, P. Chen, A. Wang, K.-j. Huang, L. Fang and X. Wu, *J. Colloid Interface Sci.*, 2022, **620**, 119–126.
- 19 J. Xu, S. Zhang, Z. Wei, W. Yan, X. Wei and K. Huang, *J. Colloid Interface Sci.*, 2021, **585**, 12–19.
- 20 F. Zhang, M. Du, Z. Miao, H. Li, W. Dong, Y. Sang, H. Jiang, W. Li, H. Liu and S. Wang, *InfoMat*, 2022, **4**, e12346.
- 21 L. Wang, X. Tan, Q. Zhu, Z. Dong, X. Wu, K. Huang and J. Xu, *J. Power Sources*, 2022, **518**, 230747.
- 22 H. Jia, K. Liu, Y. Lam, B. Tawiah, J. H. Xin, W. Nie and S.-X. Jiang, *Adv. Fiber Mater.*, 2023, **5**, 36–58.
- 23 J. Song, K. Xu, N. Liu, D. Reed and X. Li, *Mater. Today*, 2021, **45**, 191–212.
- 24 L. Ma, M. A. Schroeder, O. Borodin, T. P. Pollard, M. S. Ding, C. Wang and K. Xu, *Nat. Energy*, 2020, **5**, 743–749.
- 25 X. Yao, C. Li, R. Xiao, J. Li, H. Yang, J. Deng and M. S. Balogun, *Small*, 2022, **18**, 2204534.
- 26 X. Wu, X. Xie, H. Zhang and K.-J. Huang, *J. Colloid Interface Sci.*, 2021, **595**, 43–50.
- 27 J. Xu, Q. Liu, Z. Dong, L. Wang, X. Xie, Y. Jiang, Z. Wei, Y. Gao, Y. Zhang and K. Huang, *ACS Appl. Mater. Interfaces*, 2021, **13**, 54974–54980.
- 28 X. Geng, Y. Jiao, Y. Han, A. Mukhopadhyay, L. Yang and H. Zhu, *Adv. Funct. Mater.*, 2017, **27**, 1702998.
- 29 H. Sun, H. Liu, Z. Hou, R. Zhou, X. Liu and J.-G. Wang, *Chem. Eng. J.*, 2020, **387**, 124204.
- 30 J. Xu, Z. Dong, K. Huang, L. Wang, Z. Wei, L. Yu and X. Wu, *Scr. Mater.*, 2022, **209**, 114368.
- 31 M. Huang, Y. Mai, G. Fan, X. Liang, Z. Fang and X. Jie, *J. Alloys Compd.*, 2021, **871**, 159541.
- 32 H. Liu, J.-G. Wang, W. Hua, Z. You, Z. Hou, J. Yang, C. Wei and F. Kang, *Energy Storage Mater.*, 2021, **35**, 731–738.
- 33 J. Liu, N. Gong, W. Peng, Y. Li, F. Zhang and X. Fan, *Chem. Eng. J.*, 2022, **428**, 130981.
- 34 C. Li, C. Liu, Y. Wang, Y. Lu, L. Zhu and T. Sun, *Energy Storage Mater.*, 2022, **49**, 144–152.
- 35 W. Xu, C. Sun, K. Zhao, X. Cheng, S. Rawal, Y. Xu and Y. Wang, *Energy Storage Mater.*, 2019, **16**, 527–534.
- 36 S. Li, Y. Liu, X. Zhao, K. Cui, Q. Shen, P. Li, X. Qu and L. Jiao, *Angew. Chem., Int. Ed.*, 2021, **133**, 20448–20455.
- 37 M. Huang, Y. Mai, L. Zhao, X. Liang, Z. Fang and X. Jie, *Electrochim. Acta*, 2021, **388**, 138624.
- 38 Y. Dong, H. Shi and Z. S. Wu, *Adv. Funct. Mater.*, 2020, **30**, 2000706.
- 39 R. Zhao, A. Elzatahry, D. Chao and D. Zhao, *Matter*, 2022, **5**, 8–10.
- 40 X. Tang, X. Guo, W. Wu and G. Wang, *Adv. Energy Mater.*, 2018, **8**, 1801897.
- 41 F. Ming, H. Liang, G. Huang, Z. Bayhan and H. N. Alshareef, *Adv. Mater.*, 2021, **33**, 2004039.
- 42 Y. An, Y. Tian, J. Feng and Y. Qian, *Mater. Today*, 2022, **52**, 146–179.
- 43 Q. Zhang, H. Lai, R. Fan, P. Ji, X. Fu and H. Li, *ACS Nano*, 2021, **15**, 5249–5262.
- 44 S. Li, Y. Liu, X. Zhao, Q. Shen, W. Zhao, Q. Tan, N. Zhang, P. Li, L. Jiao and X. Qu, *Adv. Mater.*, 2021, **33**, 2007480.
- 45 M. Huang, Y. Mai, L. Zhao, X. Liang, Z. Fang and X. Jie, *ChemElectroChem*, 2020, **7**, 4218–4223.
- 46 J. Liu, P. Xu, J. Liang, H. Liu, W. Peng, Y. Li, F. Zhang and X. Fan, *Chem. Eng. J.*, 2020, **389**, 124405.
- 47 Z. Zhang, W. Li, R. Wang, H. Li, J. Yan, Q. Jin, P. Feng, K. Wang and K. Jiang, *J. Alloys Compd.*, 2021, **872**, 159599.
- 48 Y.-T. Wang, Z.-Z. Zhang and M.-X. Li, *Nano Futures*, 2022, **6**, 025001.
- 49 L. Liu, W. Yang, H. Chen, X. Chen, K. Zhang, Q. Zeng, S. Lei, J. Huang, S. Li and S. Peng, *Electrochim. Acta*, 2022, **410**, 140016.



- 50 P. Cao, N. Chen, W. Tang, Y. Liu, Y. Xia, Z. Wu, F. Li, Y. Liu and A. Sun, *J. Alloy and Compd.*, 2022, **898**, 162854.
- 51 Z. Sheng, P. Qi, Y. Lu, G. Liu, M. Chen, X. Gan, Y. Qin, K. Hao and Y. Tang, *ACS Appl. Mater. Interfaces*, 2021, **13**, 34495–34506.
- 52 Y. Q. Jin, H. Chen, L. Peng, Z. Chen, L. Cheng, J. Song, H. Zhang, J. Chen, F. Xie, Y. Jin, J. Shi and H. Meng, *Chem. Eng. J.*, 2021, **416**, 127704.
- 53 H. Jia, M. Qiu, B. Tawiah, H. Liu and S. Fu, *Compos. Commun.*, 2021, **27**, 100841.
- 54 Z. Yao, W. Zhang, X. Ren, Y. Yin, Y. Zhao, Z. Ren, Y. Sun, Q. Lei, J. Wang, L. Wang, T. Ji, P. Huai, W. Wen, X. Li, D. Zhu and R. Tai, *ACS Nano*, 2022, **16**, 12095–12106.

