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Enhancing the FRET by tuning the bandgap of acceptor ternary ZnCdS quantum dots

 Muhammad Mubeen,^a Noor ul Ain,^a Muhammad Adnan Khalid,^a Maria Mukhtar,^a Bushra Naz,^a Zumaira Siddique,^a Anwar Ul-Hamid ^b and Azhar Iqbal ^{*a}

In this article, we report the band gap tuning of ternary ZnCdS quantum dots (QDs) by varying the concentration of the capping ligand, mercaptoacetic acid (MAA). The functionalization of QDs leads to the control of their size and band gap due to the quantum confinement effect, causing blue shift in the absorption and photoluminescence (PL) spectra with a gradual change in the concentration of the capping ligand from 0.5 to 2.5 M. Ensulizole (2-phenylbenzimidazole-5-sulfonic acid) is an important organic ultraviolet (UV) filter that is frequently used in sunscreen cosmetics. An effective overlapping of the PL spectrum of ensulizole and the absorption spectrum of QDs with 2.5 M MAA is achieved. A formidable decrease in the PL intensity and the PL lifetime of ensulizole promotes an efficient Förster resonance energy transfer (FRET) from sunscreen ensulizole to the QDs. The magnitude of the FRET efficiency (E) is $\sim 70\%$. This very high value of E is the signature of the existence of a very fast energy transfer process from ensulizole to the MAA functionalized ZnCdS QDs. The dyad system consisting of ZnCdS QDs and ensulizole sunscreen can serve as a prototype model to develop a better understanding of the photochemistry of ensulizole and consequently the formulation of more efficient sunscreen cosmetics.

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1. Introduction

QDs exhibit fascinating optical and electronic properties in contrast to their bulk counterparts¹ due to the quantum confinement effect.^{2,3} The novel properties of QDs can be concocted by regulating their size and composition *via* alloying and doping techniques.^{4,5} The QDs based on metal chalcogenides of the main group II–VI have gained much importance in the fields of optoelectronics,⁶ QD lasers,⁷ heterojunction solar cells,⁸ photodetectors,^{9,10} photovoltaics,¹¹ biosensing,¹² biolabeling¹³ and for linear and non-linear optical limiting properties^{14,15} due to their tunable optical properties and size-tunable fluorescence emission.¹⁶

Recently, much attention has been devoted to the synthesis of binary,¹⁷ ternary,¹⁸ and core-shell type QDs.¹⁹ Binary QDs of group II–VI have been proven to be potential candidates for solar cells, light-emitting diodes, and lasers, but ternary semiconductor QDs have compositional dependent optoelectronic properties *i.e.*, for instance, the PL emission and photoluminescence quantum yield (PLQY) can be tuned by varying their composition.²⁰ QDs with a binary composition have discrete bandgaps which limit their applications whereas

ternary QDs have continuously tunable bandgap energy.²¹ Moreover, binary QDs mostly encompass toxic metals which also restrict their expediency. Nonetheless, ternary QDs are considered potential substitutes for binary QDs as they have high PLQY²² with a low level of toxicity. The commonly used biocompatible ternary QDs involve AgInS, CuInS, and AgCuS.^{23,24}

Among the ternary QDs, ZnCdS QDs have earned too much attention in recent decades due to the incredible dependency of bandgap (2.4–3.7 eV) on their size and composition.^{25,26} The novel properties of ZnCdS QDs lie in between ZnS and CdS QDs. The formation of these alloyed QDs follows the Hume–Rothery rule.²⁷ The electronegativity of Zn and Cd is 1.6 and 1.7 with corresponding atomic sizes 142 pm and 161 pm respectively.²⁶ This slight difference in electronegativities (0.1) and atomic radii (12%) advocates a feasible combination between Zn and Cd in the formation of ZnCdS QDs.²⁸ By changing Zn/Cd ratio in ZnCdS QDs, the absorption and emission spectra can be tuned. An increase in Zn content decreases the size of ZnCdS QDs, which leads to a remarkable quantum confinement effect. The ZnCdS QDs occupy an important position in the application of photoconductive devices, photocatalysis, and heterojunction solar cells due to their broad bandgap.²⁹ The PLQY of ZnCdS nanocrystals can be optimized through the suitable adjustment of the composition of Zn in ZnCdS.³⁰ It is foreseeable that such alloyed QDs provide multiple possibilities for bandgap tailoring and find their extensive use in a variety of applications.^{31,32}

^aDepartment of Chemistry, Quaid-I-Azam University, Islamabad-45320, Pakistan.
 E-mail: aiqbal@qau.edu.pk

^bCore Research Facilities, King Fahd University of Petroleum & Minerals, Dhahran 31261, Saudi Arabia



Moreover, QDs are generally considered to be toxic but their toxicity depends on their shape, size, and surface functionalization.³³ Also, the introduction of Zn reduces the toxicity of CdS and renders it more biocompatible and safe to use in cosmetic products. The experiments have revealed that the human skin is not labile for the penetration of commercially available QDs.³⁴ Therefore, the combination of these ternary QDs with ensulizole sunscreen can improve the efficiency of commercial sunscreens.

Ensulizole is a prominent UV-B (280–315 nm) type filter³⁵ frequently used in sunscreen cosmetic products. As QDs have a large absorption cross-section that can be tuned over a broad spectral region, therefore, the association between ensulizole and QDs may act as an imperative route of absorbed energy dissipation. This photo-absorbed energy transfer dynamics can be assessed *via* Förster resonance energy transfer (FRET) model.³⁶ The FRET is a non-radiative mode of energy transfer from an excited state donor to the ground state of proximate acceptor system.³⁷ The FRET is frequently used as a molecular ruler to gauge the distance in the biological systems.^{37,38} The other applications of FRET include visualization of molecular interactions,³⁹ single molecule conformational dynamics,⁴⁰ and monitoring binding/unbinding events in biological systems.⁴¹ Recently, the effect of size of CdS QDs as a FRET acceptor antagonist on a sunscreen donor has been reported.⁴² Very recently, the work of Cowden *et al.*⁴³ reveals that the combination of two sunscreen molecules into one composite sunscreen forms a product with good stability towards UV light absorption and its dissipation in ethanol and bestows advantages over single-component of sunscreens.⁴³

In this article, we investigate the surface modification of ZnCdS QDs by varying the concentration of capping ligand that tunes the photophysical properties of QDs. The concentration of capping ligand MAA is varied from 0.5 to 2.5 M to achieve the maximum overlapping between the PL emission spectrum of ensulizole and the absorption spectrum of QDs. These QDs act as FRET acceptor in conjunction with ensulizole donor as a model system for the study of dissipation of absorbed UV energy. The UV-absorbed energy transfer efficiency from

ensulizole sunscreen to the QDs is ~70%, which is significantly higher than the efficiencies of 67.3% (ref. 42) and 64.05% (ref. 44) reported in the previous works.

2. Experimental

2.1. Materials

Zinc sulfate dihydrate ($\text{ZnSO}_4 \cdot 2\text{H}_2\text{O}$, $\geq 98\%$), cadmium chloride pentahydrate ($\text{CdCl}_2 \cdot 5\text{H}_2\text{O}$, 99%), mercaptoacetic acid (HSCH_2COOH , 98%), sodium sulfide nonahydrate ($\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$, 99.9%), and sodium hydroxide (NaOH , $\geq 99\%$) were obtained from Sigma-Aldrich. Ensulizole (98%) was procured from the Tokyo Chemical Industry (TCI). All chemicals were of analytical grade and were used without any further treatment. Deionized water was utilized as a solvent in all the measurements.

2.2. Synthesis

The ZnCdS QDs were synthesized by employing the colloidal method.^{45,46} Briefly, 0.1 M cationic precursors $\text{ZnSO}_4 \cdot 2\text{H}_2\text{O}$ and $\text{CdCl}_2 \cdot 5\text{H}_2\text{O}$ were prepared separately in deionized water and then taken into a three-necked round bottom flask along with 0.5 M MAA. The reaction mixture was stirred at 300 rounds per minute (RPM) and heated at 80 °C temperature under an inert atmosphere of nitrogen gas. To hold the pH of the solution between 9 and 11, NaOH solution (1 M) was added drop by drop. When the solution turned transparent, anionic precursor $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$ (0.1 M) was added. The whole reaction mixture was heated and stirred for about one hour. The final solution was precipitated with ethanol and then isolated through ultracentrifugation. The obtained crystals of QDs were washed repeatedly with a mixture of ethanol and water to remove unreacted ingredients and finally dried at 50 °C in the oven for 6–7 hours. Similarly, the whole reaction procedure was repeated with other concentrations of MAA *i.e.* 1.0 M, 1.5 M, 2.0 M, and 2.5 M. Moreover, 5 mL of 85 μM ensulizole was mixed with 1 mL of 133 μM of ZnCdS QDs in a round bottom flask and stirred for 30 minutes at 60 °C for the effective association of the QDs with the ensulizole. The resultant solution was used as such for the FRET measurements. The molar concentration of QDs was measured

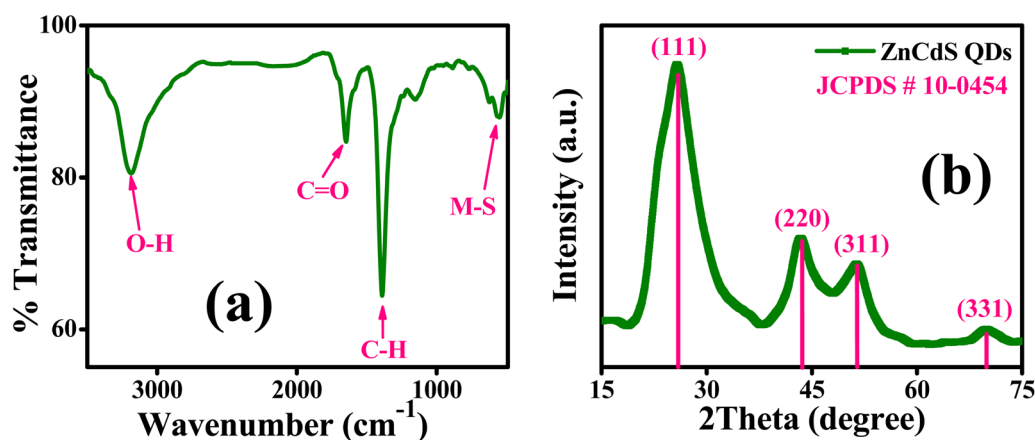


Fig. 1 (a) FTIR spectrum, and (b) XRD pattern of MAA (2.5 M) functionalized ZnCdS QDs.

by dividing the total volume per liter of the QDs by the volume of single QD, which was assessed from TEM micrograph (Fig. 2(b)). For the determination of total volume per liter of QDs, it was believed that both the nano and bulk materials had same density.⁴⁷ Afterward, the molar absorption coefficient of QDs was estimated from the slope of the line by plotting absorption verses concentrations as mentioned in our previously reported work.⁴⁴

2.3. Characterization techniques

To confirm the functionalization of the synthesized ZnCdS QDs, Fourier transform infrared (FTIR) spectroscopic analysis is performed in the region between 4000 and 500 cm^{-1} by using a BRUKER TENSOR-II FTIR spectrometer. The crystal structure of ZnCdS QDs is determined by using PANalytical X-ray diffractometer (XRD) model 3040/60 X'Pert PRO operated at 45 kV and 40 mA, having CuK_α ($\lambda = 1.54 \text{ \AA}$) as the radiation source over an angle range of 10–80° at a step size of 0.02°. The UV-Vis absorption measurements are performed with the help of the Shimadzu-1601 spectrophotometer in the wavelength range of 200–800 nm. The elemental analysis and particle size investigation of synthesized ZnCdS QDs are performed by JSM-6460 LV energy dispersive X-ray (EDX) and high-resolution transmission electron microscope (TEM), Philips CM 30, JEOL, JEM-2100F. Steady-state photoluminescence (SSPL) and time-resolved photoluminescence (TRPL) measurements are recorded by employing the PicoQuant, Germany, FluoTime 300 (FT-300) time-correlated single-photon counting (TCSPC) setup. The PL measurements of the samples are performed at room temperature after pulsed excitation at 306 nm with a pulsed

LED laser coupled to a TCSPC setup. The pulse duration of PLS-300 is 416 ps and pulse energy is 0.077 pJ. The long-pass color filter FGL400 is applied to filter the residual light of the excitation source and to elude the second-order of 306 nm.

3. Results and discussion

FTIR analysis is performed to detect different functional groups present in the sample of MAA capped ZnCdS QDs as depicted in Fig. 1(a). No characteristic peak is observed between the regions 2550–2620 cm^{-1} for S–H stretching mode that indicates the existence of strong binding interaction between the sulfur of the thiol group (S–H) of capping ligand to the surface of ZnCdS QDs. Hence, it can be inferred that ZnCdS QDs are successfully functionalized by MAA. The characteristic peaks that existed at 3190 cm^{-1} and 1648 cm^{-1} are attributed to the O–H and C=O bond stretching vibrations of the carboxylic group of capping ligand, respectively. The strong peak centered at 1388 cm^{-1} corresponds to the H–C–H bending mode. The typical peak observed at 560 cm^{-1} can be assigned to the stretching vibration modes of the metal–sulfur bond of ZnCdS.

The powder XRD technique is used to determine the crystal structure and crystallinity along with the crystallite size of the synthesized material. The XRD pattern of 2.5 M ZnCdS QDs along with the reference card is displayed in Fig. 1(b). This pattern matches with the crystal structure of the cubic zinc blende and is confirmed by JCPDS-10-0454. The four distinct peaks appeared at 2θ values of 25.8°, 43.6°, 51.4°, and 70.0° are associated with (111), (220), (311), and (331) crystal planes, respectively. The plane indexed with (111) is the most intense

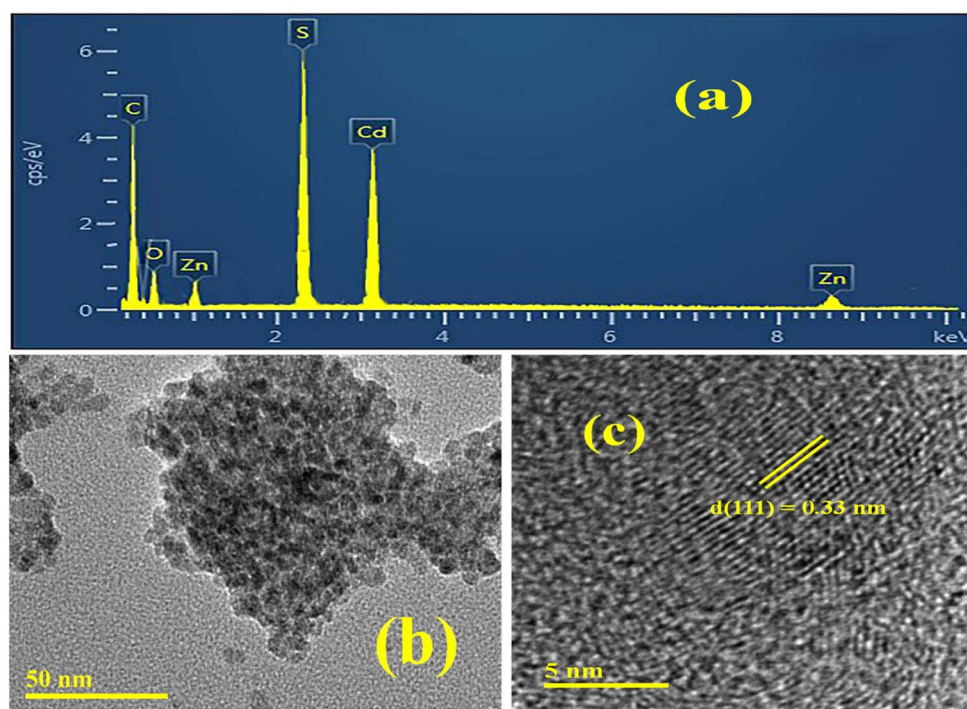


Fig. 2 (a) EDX analysis, (b) TEM micrograph, and (c) HRTEM micrograph of MAA (2.5 M) functionalized ZnCdS QDs.



peak and the manifestation of two closely spaced reflection planes (220) and (311) verify the crystalline nature of synthesized ZnCdS QDs. The crystallite size of 2.5 M MAA functionalized ZnCdS QDs is determined ~ 4.9 nm with the help of Scherrer's eqn (1).

$$D = \frac{K\lambda}{\beta \cos \theta} \quad (1)$$

where D , k , λ , β and θ are crystallite size, Scherrer's constant (0.94 for cubic crystal structure), the wavelength of X-rays used, full width at half maximum, and Bragg angle respectively. The purity of the synthesized material is confirmed by EDX analysis as depicted in Fig. 2(a). To further verify the particle size of synthesized QDs, TEM measurements are performed and the images are displayed in Fig. 2(b). The particle size estimated from the TEM micrograph using ImageJ software is ~ 5 nm which is in good concurrence with the crystallite size evaluated from the aforementioned eqn (1). The high-resolution TEM micrograph (Fig. 2(c)) unveils the crystal lattice parameters of the synthesized ZnCdS QDs, which corresponds to a d -spacing of 0.33 nm of the (111) crystal plane.

Fig. 3(a) displays UV-Vis absorption spectra of ZnCdS QDs with different concentrations of the functionalizing ligand (MAA). The ZnCdS QDs having a wide range of absorption edge lies in between bulk ZnS (3.62 eV)⁴⁸ and CdS (2.42 eV).⁴⁹ By increasing the concentration of capping ligand MAA, the blue shift in the first absorption exciton peak of QDs is a clear manifestation of the quantum confinement effect. Moreover, when the concentration of the capping ligand increases, more sites of QDs are engaged that decrease the nucleation and the size decreases. A blue shift is observed in absorption spectra that corresponds to an increase in bandgap energy that is estimated from Tauc's method and is displayed in Fig. 3(b).

As reflected in Fig. 3(b), the value of bandgap energy for ternary ZnCdS QDs comprising different concentrations of MAA varies linearly *i.e.* with the increase in the concentration of capping ligand from 0.5 M to 2.5 M, the value of the bandgap energy of QDs also increases from 2.88 eV to 3.37 eV, respectively. This again can be ascribed to the quantum confinement prodigy. The observation of this wide range in tunability of

bandgap energy of ternary ZnCdS QDs in contrast to binary ZnS and CdS can also be attributed to the composition of alloyed QDs along with the quantum size effect.⁵⁰

From bandgap data, the average particle size of all the synthesized ZnCdS QDs is determined by applying Brus equation⁵¹ eqn (2).

$$E_g^* = E_g + \frac{h^2}{8r^2} \left(\frac{1}{m_o m_e} + \frac{1}{m_o m_h} \right) - \frac{1.8e^2}{4\pi\epsilon_r\epsilon_o r} \quad (2)$$

where, E_g^* and E_g are the bandgap energies of QDs and bulk material, respectively. r is the radius of QDs, h is Planck's constant, m_e is the effective mass of the electron ($0.2 m_o$), m_h is the effective mass of hole ($0.8 m_o$), m_o is the rest mass of the free electron (9.11×10^{-31} kg), ϵ_r is relative dielectric constant ($5.7 \epsilon_o$), ϵ_o is the dielectric constant of the vacuum (8.85×10^{-12} C² N⁻¹ m⁻²) and e is the standard charge on an electron (1.6022×10^{-19} C). The particle sizes of QDs in terms of diameter estimated from eqn (2) are 7.8 nm, 7.0 nm, 6.0 nm, 5.4 nm, and 5.1 nm for 0.5 M, 1.0 M, 1.5 M, 2.0 M, and 2.5 M MAA functionalized ZnCdS QDs, respectively.

Fig. 4(a) exhibits the SSPL spectra of ternary ZnCdS QDs with PL emission peaks centered at 590 nm, 573 nm, 552 nm, 516 nm, and 468 nm for 0.5 M, 1.0 M, 1.5 M, 2 M, and 2.5 M concentrations of capping ligand MAA, respectively. As the concentration of MAA increases, a blue shift in PL emission peak due to the quantum confinement effect is noticed which is in accordance with the UV-Vis absorption measurements of ZnCdS QDs as displayed in aforementioned Fig. 3(a). This gradual shift in PL emission peak with change in concentration of capping ligand MAA is also an indication of successful synthesis of assorted sized ZnCdS QDs. The extended PL tail towards the longer wavelengths of QDs containing the higher conc. of MAA can be attributed to the presence of trap states, which are accountable to cause defects such as dangling bonds and vacancies on the surface of the QDs.

The TRPL measurements are performed to get information about the PL decay kinetics of ZnCdS QDs. The PL decay kinetics demonstrate an excellent trend in the lifetime for differently sized functionalized ZnCdS QDs as evident from Fig. 4(b). Moreover, the shape of all the curves of Fig. 4(b)

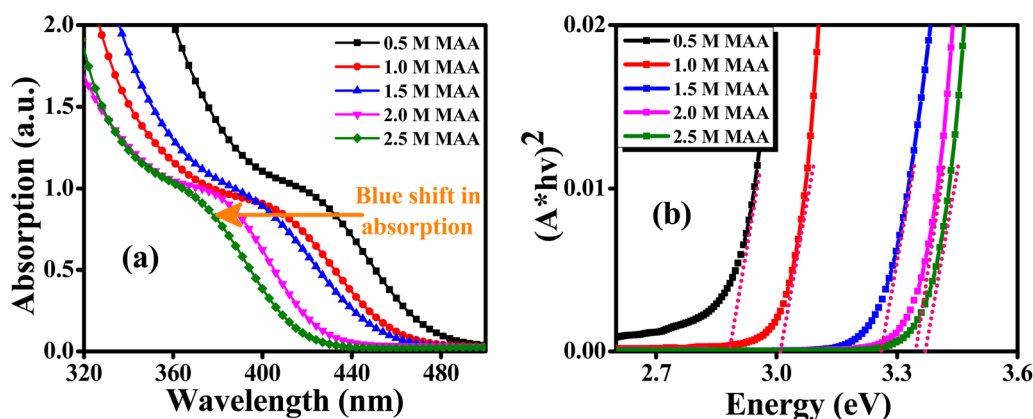


Fig. 3 (a) UV-Visible absorption spectra and (b) Tauc's plot of ZnCdS QDs functionalized with different concentrations of MAA.



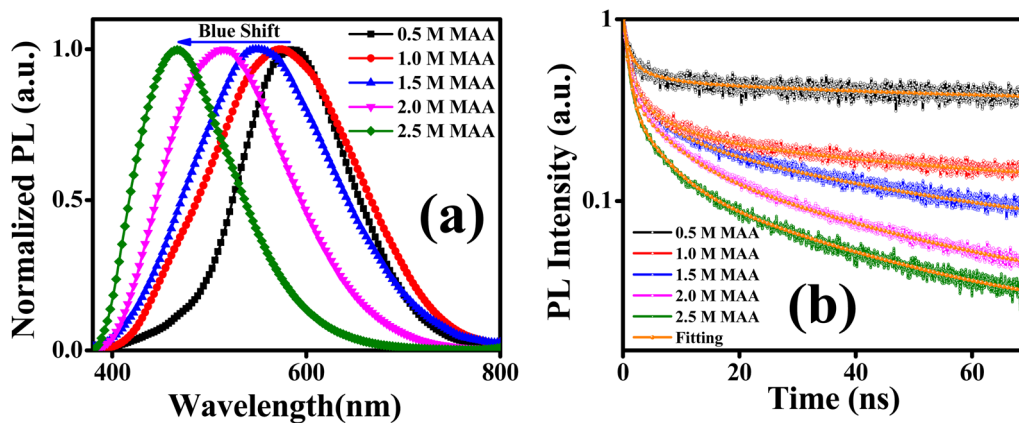


Fig. 4 (a) Normalized SSPL spectra, and (b) PL decay kinetics of ZnCdS QDs functionalized with different concentrations of MAA.

reveals that the 1st order decay kinetics is operative. The measured data is best fitted by tri-exponential decay function for the PL kinetics of different sized ZnCdS QDs. The various fitting parameters extracted from the best fits are shown in Table 1. The average PL decay lifetimes (τ_{average}) are calculated by employing eqn (4) using the time constants and the coefficients extracted from eqn (3).

$$Y = Y_0 + \sum_{i=1}^n A_i e^{-\frac{x}{t_i}} \quad (3)$$

$$\tau_{\text{average}} = \frac{A_1 \tau_1 + A_2 \tau_2 + A_3 \tau_3}{A_1 + A_2 + A_3} \quad (4)$$

here A_i and t_i are the associated coefficients and the time constants, respectively.

The average PL lifetime of ZnCdS QDs with 0.5 M MAA is 12.13 ns and as the concentration of MAA increases, PL lifetime decreases accordingly as evident from Table 1. The longest PL lifetime is observed with the 0.5 M concentration of MAA, this is because, for large-sized QDs, energy levels are more closely spaced in which electron and hole can easily be trapped, consequently, exciton in large-sized QDs exists for a longer period than in small-sized QDs.⁵² Therefore, an increase in the concentration of MAA capping ligand results in the decrease in particle size of QDs, which leads to the creation of shallow trap states and subsequently, shortening of PL lifetime as displayed in Fig. 5.

Table 1 Fitting parameters extracted from the best fits of the measured PL decay kinetics for various MAA functionalized ZnCdS QDs

Conc. of MAA	A_1	τ_1 (ns)	A_2	τ_2 (ns)	A_3	τ_3 (ns)	τ_{average} (ns)
0.5 M	0.63	1.26	0.15	6.68	0.22	46.99	12.13
1.0 M	0.62	1.19	0.18	6.16	0.19	42.20	9.96
1.5 M	0.65	1.09	0.20	5.73	0.15	37.92	7.54
2.0 M	0.67	0.89	0.21	5.94	0.11	31.10	5.32
2.5 M	0.71	0.56	0.19	4.11	0.09	24.22	3.40

The Fig. 5 displays a relation between the concentration of capping ligand and average PL lifetime, suggesting PL lifetime varies linearly with the concentration of the functionalizing ligand MAA.

The energy transfer from the photoexcited ensulizole to the MAA functionalized ZnCdS QDs can be envisioned by applying the FRET model. The FRET process arises due to non-radiative energy transfer absorbed by the donor fluorophore in the presence of an acceptor moiety in its proximity. As a result, the energy transfer befalls without the emission of photons due to long-range dipole-dipole interactions of transition dipoles of donor and the acceptor molecules. To study the energy transfer between ensulizole and QDs, a FRET model is employed. The overlapping between the emission spectrum of the donor and the absorption spectrum of the acceptor (Fig. 6(a)) is the basic requirement for occurrence of effective FRET process. A drop in the PL lifetime of a donor in the presence of an acceptor (Fig. 7(a)) is a quantitative measure of FRET efficiency. In the absence of the ZnCdS QDs counterpart, the average PL decay lifetime displayed by sole ensulizole is 5.09 ns. This value is reduced from 5.09 ns to 1.54 ns for an ensulizole-ZnCdS QDs dyad as exhibited by Fig. 7(b).

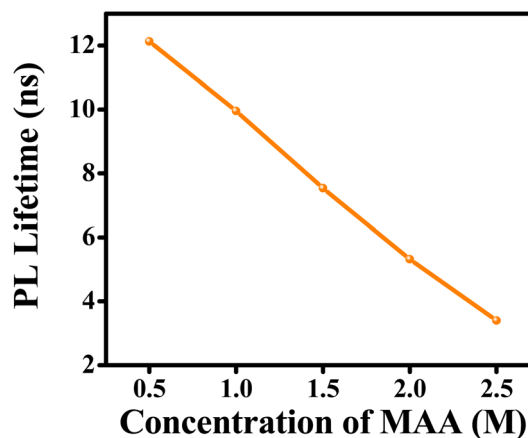


Fig. 5 Effect of concentration of capping ligand (MAA) on the average PL lifetime of ZnCdS QDs.



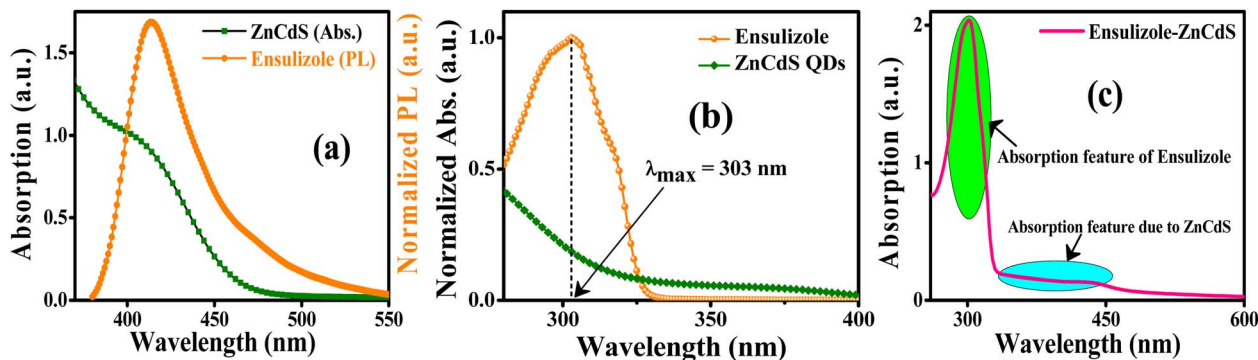


Fig. 6 (a) Spectral overlapping between the absorption spectrum of ZnCdS QDs and PL spectrum of ensulizole. (b) Normalized absorption spectrum of ensulizole and 2.5 M MAA capped ZnCdS QDs. (c) Absorption spectrum of ensulizole–ZnCdS QDs dyad containing MAA concentration 2.5 M.

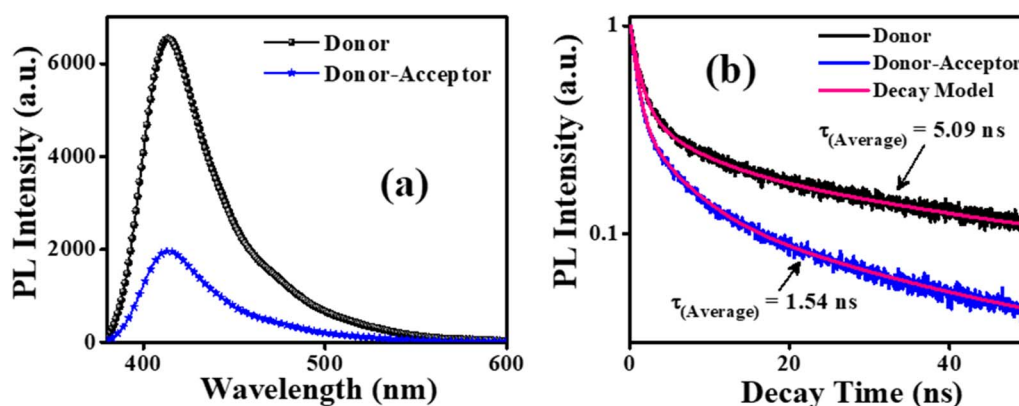


Fig. 7 (a) SSPL spectra and (b) PL decay kinetics of donor and the donor–acceptor.

The normalized absorption spectrum of ensulizole measured in the aqueous medium is displayed in Fig. 6(b) that exhibits maximum absorption at 303 nm. Fig. 6(c) displays as measured absorption spectrum of ensulizole in association with ZnCdS QDs. The presence of absorption features of both the ensulizole and the QDs is a discernible indication of the association of ensulizole with QDs. It can be assessed from Fig. 6(c) that the absorption of ensulizole at 303 nm is 2.04 and that of ZnCdS QDs is 0.15. By applying Beer–Lambert's law, the average number of ensulizole molecules associated with the 2.5 M MAA functionalized ZnCdS QDs are determined. This correspond to

$\sim 2.56 \times 10^{17}$ ensulizole molecules and $\sim 8.01 \times 10^{16}$ number of ZnCdS QDs. Thus, the average number of experimentally determined ensulizole molecules that surround one QDs is ~ 3.2 . The rate of FRET is dependent on the dipole moments of the donor and the acceptor, the refractive index of the medium, quantum yield of the donor, spectral overlap integral, orientation factor, and the distance between donor and acceptor species.⁵³ All the different FRET parameters are assessed by utilizing equations from (5) to (9) and are given in Table 2.

$$E = 1 - \frac{\tau_{\text{DA}}}{\tau_{\text{D}}} \quad (5)$$

$$E = \frac{R_0^6}{R_0^6 + r^6} \quad (6)$$

$$k_{T(r)} = \frac{1}{\tau_{\text{D}}} \left(\frac{R_0}{r} \right)^6 \quad (7)$$

$$R_0(\text{\AA}) = 0.211 \sqrt[6]{k^2 n^4 \phi_{\text{D}} J(\lambda)} \quad (8)$$

$$J(\lambda) = \int_0^{\infty} F_{\text{D}}(\lambda) \epsilon_{\text{A}}(\lambda) \lambda^4 d\lambda \quad (9)$$

Table 2 The FRET parameters of donor ensulizole and ensulizole–ZnCdS QDs dyad

Parameters	Values
τ_{D}	5.09 ns
τ_{DA}	1.54 ns
E	69.74%
$J(\lambda)$	$1.29 \times 10^{13} \text{ M}^{-1} \text{ cm}^{-1} \text{ nm}^4$
R_0	20 \AA
R	17.37 \AA
$k_{T(r)}$	$4.57 \times 10^8 \text{ s}^{-1}$



where, τ_D , τ_{DA} , R_0 , r , ϕ_D , n , k , λ , ϵ_A , $J(\lambda)$ and $k_{T(r)}$ designate average PL lifetime of the donor only, PL lifetime of donor in the presence of acceptor, Förster distance (the distance at which FRET efficiency is 50%) and the actual distance between the donor and the acceptor, fluorescence quantum yield of the donor, the refractive index of the medium, orientation factor, wavelength, absorption coefficient, spectral overlapping integral,⁵⁴ and non-radiative energy transfer rate, respectively.⁵⁵ For any randomly oriented system containing donor and acceptor species, $k^2 = 2/3$. The fluorescence quantum yield ($\phi_D = 0.315$) of ensulizole is calculated by relative method⁵⁶ making use of fluorescein (in 0.1 M aqueous NaOH) as a reference fluorophore.⁵⁷ It is evident from eqn (5) that energy transfer (E) exclusively depends on the difference in the PL decay lifetime of the donor–acceptor system and the sole donor.

After pulsed excitation at 306 nm, the singlet $\pi\pi^*$ state of ensulizole is excited. The bare photoexcited ensulizole molecules dissipate absorbed UV energy through various radiative and non-radiative conduits. When an acceptor molecule in the ground state exists near the excited donor molecule, then the transition dipoles of donor and the acceptor moieties can resonate with one another. Consequently, the excited-state energy of the donor is transferred to the acceptor *via* the FRET process. The distance between the donor and the acceptor molecules must be in the range of 10–100 Å for a permissible FRET process to occur⁵⁸ and the donor–acceptor pair must fulfill this necessary condition. According to quantum mechanical treatment of FRET, it is an established fact that the efficacy of FRET apart from molecular structure also depends on the density of overlapping states of the donor (emissive state) and acceptor (absorption state).^{59,60}

As ensulizole associated with 2.5 M MAA functionalized ZnCdS QDs displays a reasonable overlapping of the PL spectrum with the first excitonic peak of QDs, Fig. 6(a), suggesting the large magnitude of FRET efficiency (69.74%), Table 2, which is significantly higher than the previously reported literature values 67.3% (ref. 42) and 64.05%.⁴⁴ It is evident from Table 2 that MAA functionalized ZnCdS QDs coupled with ensulizole have a high value of energy dissipation through the non-radiative FRET pathway. The FRET mechanism is exceedingly distance-dependent, so it facilitates the assessment of the distance between donor and acceptor moieties. For the given FRET pair estimated values of E , r , $k_{T(r)}$ and R_0 , from eqn (5) to (8), respectively are summarized in Table 2. The assessed value of R_0 (20 Å) is within permissible limits of 10 Å to 100 Å.⁶¹ The calculated values of E (69.74%) and $k_{T(r)}$ ($4.57 \times 10^8 \text{ s}^{-1}$) specify an efficient transfer of absorbed UV energy from ensulizole to ZnCdS QDs non-radiatively. The actual donor–acceptor distance for this particular pair is ~ 17.37 Å. The evaluation of all the aforementioned parameters confirms the existence of an effective FRET between ensulizole and 2.5 M MAA functionalized ZnCdS QDs.

4. Conclusions

To design an efficient FRET system, the spectral tuning of ternary ZnCdS QDs was performed by gradually increasing the

concentration of MAA from 0.5 to 2.5 M that led to change the bandgap from 2.88 eV to 3.37 eV in contrast to binary bare ZnS (3.62 eV) and CdS (2.42 eV) QDs. The PL spectra of these QDs revealed a shift in emission wavelength from 468 to 590 nm with a successive change in concentration of capping ligand from 0.5 to 2.5 M and this was in consistent with the results of UV-Vis measurements. A decrease in the average PL lifetime of QDs was observed with a gradual rise in the concentration of capping ligand that suggested in large-sized QDs the electron–hole pair could stay longer as compared to small-sized QDs. The effective spectral overlapping between the PL emission spectrum of ensulizole and absorption spectrum of 2.5 M MAA functionalized ZnCdS QDs, decreased the PL intensity of donor in the presence of acceptor in the dyad. The shortening of PL lifetime of the donor from 5.09 ns to 1.54 ns in the presence of acceptor, confirmed the existence of FRET between ensulizole and QDs moieties. The FRET efficiency, 69.74% was exceedingly high and confirmed an efficient FRET process was operative between ensulizole and 2.5 M MAA functionalized ZnCdS QDs. Overall, an increase of 2.44% in FRET was measured for such systems compared to previously reported literature. In future, this work can be extended to the synthesis of more biocompatible QDs and their coupling with various recommended sunscreen components for the formulation of “hybrid sunscreen” cosmetic products.

Author contributions

All authors contributed to the writing of the manuscript and have approved the final version of the manuscript.

Conflicts of interest

The authors declare no competing financial interest.

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