

Cite this: *Chem. Sci.*, 2023, 14, 12637

All publication charges for this article have been paid for by the Royal Society of Chemistry

Template-assisted synthesis of isomeric copper(I) clusters with tunable structures showing photophysical and electrochemical properties†

Jun-Jie Fang, Zheng Liu, Yang-Lin Shen, Yun-Peng Xie * and Xing Lu *

A comparative study of structure–property relationships in isomeric and isostructural atomically precise clusters is an ideal approach to unravel their fundamental properties. Herein, seven high-nuclearity copper(I) alkynyl clusters utilizing template-assisted strategies were synthesized. Spherical Cu₃₆ and Cu₅₆ clusters are formed with a [M@(*V*/PO₄)₆] (M: Cu²⁺, Na⁺, K⁺) skeleton motif, while peanut-shaped Cu₅₆ clusters feature four separate PO₄ templates. Experiments and theoretical calculations suggested that the photophysical properties of these clusters are dependent on both the inner templates and outer phosphonate ligands. Phenyl and 1-naphthyl phosphate-protected clusters exhibited enhanced emission features attributed to numerous well-arranged intermolecular C–H⋯π interactions between the ligands. Moreover, the electrocatalytic CO₂ reduction properties suggested that internal PO₄ templates and external naphthyl groups could promote an increase in C₂ products (C₂H₄ and C₂H₅OH). Our research provides new insight into the design and synthesis of multifunctional copper(I) clusters, and highlights the significance of atomic-level comparative studies of structure–property relationships.

Received 5th September 2023

Accepted 12th October 2023

DOI: 10.1039/d3sc04682f

rsc.li/chemical-science

Introduction

During the past few decades, atomically precise metal nano-clusters have contributed significantly to the rapid development of nanomaterials by elucidating structure–property relationships.^{1–6} Alkynyl ligands have more abundant coordination modes than the commonly used thiolate and phosphine ligands, and therefore alkynyl-protected coinage metal clusters have remarkable structures.^{7–11} However, the reaction between alkynyl ligands and Cu⁺ often produces insoluble polymers, which increases the difficulty of synthesis and isolation.^{12,13} To address this issue, template agents have been introduced into the synthesis process to balance the charge distribution of the clusters and prevent or interrupt the formation of insoluble polymers resulting from the reaction between alkynyl and Cu⁺. As a result, the synthesis of high-nuclearity copper(I) alkynyl clusters with larger size and unique properties has been facilitated.^{14–17}

Polyoxometalates (POMs) have attracted considerable attention due to their structural diversity and desirable properties in materials science, medicinal chemistry, and catalysis.^{18–20}

Previous studies have reported silver alkynyl clusters with richer structures and more unique properties constructed with POMs.^{21–24} Polyoxovanadates and polyoxomolybdates, both members of the POM family, constitute a fascinating family of polynuclear oxo-anions that exhibit variable coordination geometries, but have a strong tendency to exhibit mixed-valent states,^{25–27} which leads to the easy oxidation of Cu⁺. This oxidizability significantly increases the difficulty in synthesizing copper(I) alkynyl clusters constructed with POMs.

Metal phosphonates not only exhibit various structures ranging from discrete molecules to multi-dimensional coordination polymers that provide a bridge for the construction of high-nuclearity metal clusters,^{28–30} but also hold potential applications in biotechnology, water treatment, photochemistry, and magnetism.^{31–34} We previously reported the use of (ⁿBu/^lBu)PO₃H₂ as precursors to construct a series of silver clusters. (ⁿBu/^lBu)PO₃H₂ units serve as tripod pillars that have a variety of coordination modes with metals and can be used as structure-directing templates to form enlarged composite clusters.^{23,24,35–37} However, the reactions of polyoxometalates and phosphonates in this system are complicated, and V/P elements tend to exhibit multiple valent states that are prone to oxidize Cu⁺. Consequently, it is more difficult to synthesize copper(I) clusters with POM or phosphonate templates than silver(I) clusters. So far, only a few successfully synthesized copper(I) alkynyl clusters constructed with POM or phosphonate templates have been reported, including Mak's previously reported Cu₃₃,³⁸ Cu₄₆,³⁹ Cu₄₇,³⁹ Cu₆₂ (ref. ³⁸) and our reported Cu₂₅.⁴⁰

State Key Laboratory of Materials Processing and Die & Mould Technology, School of Materials Science and Engineering, Huazhong University of Science and Technology, Wuhan 430074, China. E-mail: xieyp@hust.edu.cn; lux@hust.edu.cn

† Electronic supplementary information (ESI) available. CCDC 2219441, 2116536, 2285409, 2219449, 2219459, 2219463, 2219464 and 2219447. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d3sc04682f>



Recently, atomically precise copper clusters and their correlations between structures and properties have received a lot of attention, particularly with regard to photoluminescence^{41,42} and CO₂ reduction reaction (CO₂RR).^{43–45} For example, Jin *et al.* found that the phosphorescence quantum efficiency of M@Cu₁₄ (M: Au, Cl) nanoclusters can be tuned by a single-atom kernel.⁴⁶ Zang *et al.* observed that Au doping in the innermost shell of Au₁₂(AgCu)₃₈ significantly enhances near-infrared photoluminescence intensity.⁴⁷ Bakr *et al.* reported that extended C–H⋯π and π⋯π intermolecular ligand interactions significantly enhance photoluminescence.⁴⁸ Atomically precise clusters with well-defined architectures and chemical compositions have attracted attention in the field of heterogeneous catalysis.^{49–53} For instance, Liu *et al.* synthesized Cu₃₂H₂₀L₁₂ (L: S₂P(OⁱPr)₂) exhibiting high selectivity (FE_{HCOOH} = 89% at –0.3 V, 83% at –0.4 V) for the CO₂RR.⁵⁴ Zang *et al.* observed that ditetrahedron-shaped Cu₈ exhibited a higher FE_{HCOOH} (≈92%) at –1.0 V than the cube-shaped Cu₈.⁵⁵ Tang *et al.* reported that M₁₅ (Au₇Ag₈, Ag₉Cu₆, and Au₂Ag₈Cu₅) exhibits drastically different CO₂RR performances in a wide voltage range.⁵⁶ These atomically precise clusters facilitate systematic comparative studies of structure–property relationships.

In this study, we designed a synthesis strategy to address the aforementioned challenges by adding weak reducing agents and phosphonate ligands, successfully synthesizing seven copper(i) alkynyl clusters. Single crystal structure analysis revealed that the structural modulation of these clusters can be achieved through the bonding of metal cations to templates. One Cu²⁺, Na⁺, or K⁺ ion captured by six VO₄ or PO₄ tetrahedrons constructs a [M@(V/PO₄)₆] (M: Cu²⁺, Na⁺, K⁺) skeleton, which tends to form structurally similar spherical Cu₃₆ and Cu₅₆ clusters. Meanwhile, in the reaction system where Na⁺ and K⁺ ions are replaced with NH₄⁺, peanut-shaped Cu₅₆ clusters can be constructed from four PO₄ tetrahedrons. The crystalline-state emissions of these clusters depend on the inner templates and the outer phosphonate ligands. The surfaces of PhOPO₃^{2–} and 1-NaphOPO₃^{2–} protected clusters are interconnected through numerous well-arranged intermolecular C–H⋯π interactions between the ligands of adjacent clusters, leading to enhanced emission. Electrochemical measurements of CO₂ reduction reaction showed that PO₄-encapsulated Cu₅₆ possesses higher catalytic activity for the CO₂RR than VO₄-encapsulated Cu₅₆, and the PO₄ template and external naphthyl group can improve the yield of C₂ products (C₂H₄ and C₂H₅OH).

Results and discussion

Synthesis and general characterization

The synthesis of clusters, [Cu^{II}(VO₄)₆@Cu^I₅₆(^tBuPO₃)₆(V₃O₃OH)(V₃O₆OH)(^tBuC≡C)₁₈] (**1**), [Cu^{II}(VO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₆]F₂ (**2**), [Na^I(PO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₆]F (**3-Na**), [K^I(PO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₆]F (**3-K**), [Na^I(PO₄)₆@Cu^I₅₆(PhOPO₃)₆(^tBuC≡C)₂₆]F (**3-PhOPO₃**, PhOPO₃ = phenyl phosphate), [Na^I(PO₄)₆@Cu^I₅₆(1-NaphOPO₃)₆(^tBuC≡C)₂₆]F (**3-NaphOPO₃**, 1-NaphOPO₃ = 1-naphthyl

phosphate), [(PO₄)₄@Cu^I₅₆(1-NaphOPO₃)₆(^tBuC≡C)₃₂] (**4**), requires meticulous control over various phosphonates and vanadates.

To achieve this, we employed three different phosphonate ligands with varying steric hindrances to regulate their structures (Fig. 1). Reaction conditions were optimized to obtain the crystalline products, including the ratio of template precursors and (^tBu/PhO/1-NaphO)PO₃H₂ ligands, time of ultrasonic conditions, amount of Et₃N used, amount of PhMe₂SiH added, and solvothermal temperature. Adjusting the addition ratio of NH₄VO₃ and ^tBuPO₃H₂ allowed us to obtain single crystals of **1** and **2**. It should be noted that precise control of the amount of PhMe₂SiH added was crucial to overcome the oxidizability of template precursors. Under natural light, the appearance of single crystals of **1–4** is dark red blocky and yellow blocky, respectively (Fig. S1†).

We elucidated the molecular structures of **1–4** through single crystal X-ray diffraction (SCXRD), and we have listed their crystal data, structure refinement, and selected bond lengths in ESI Tables 1, 2 and 5–12.† SCXRD analysis revealed that clusters **1** and **4** are both neutral clusters. While crystals of **1** and **4** are insoluble in common solvents, other crystals dissolve easily in CH₂Cl₂, CH₃OH, and CH₃CH₂OH, providing opportunities to examine their solution stabilities and ionic valences using electrospray ionization mass spectrometry (ESI-MS). For example, in the positive ion mode ESI-MS of **2** dissolved in CH₃OH and CH₂Cl₂, the two-charge prominent peaks centered at *m/z* = 3598.27 (**2-a**), 3608.24 (**2-b**), and 3617.23 (**2-c**) are assigned formulas of [Cu^{II}@(VO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₅F₂ + H⁺]²⁺ (calcd *m/z* = 3598.54), [Cu^{II}@(VO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₅F₂ + H₂O + H⁺]²⁺ (calcd *m/z* = 3608.05), and [Cu^{II}@(VO₄)₆@Cu^I₅₆(^tBuPO₃)₆(^tBuC≡C)₂₅F₂ + (H₂O)₂ + H⁺]²⁺ (calcd *m/z* = 3617.05), respectively (Fig. 2). We accurately identified other species produced by ionization of one or two ^tBuC≡C[–] ligands from the cations of the pristine clusters or by

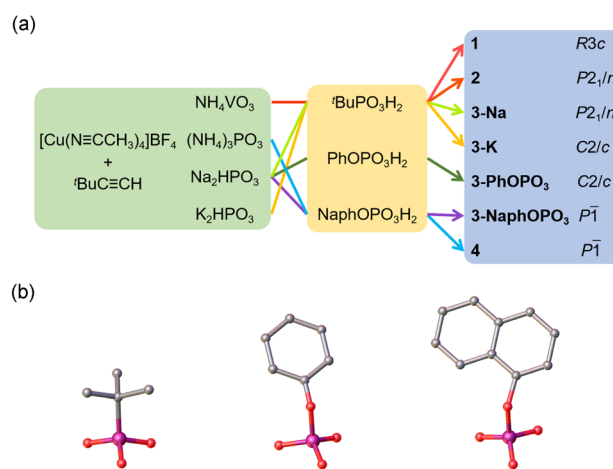


Fig. 1 Synthetic route and structures of phosphonate ligands. (a) Synthetic route for **1–4**. (b) Molecular structures of ^tBuPO₃H₂, PhOPO₃H₂, and 1-NaphOPO₃H₂ ligands. Color codes: P, purple; O, red; C, gray. Hydrogen atoms are omitted for clarity.



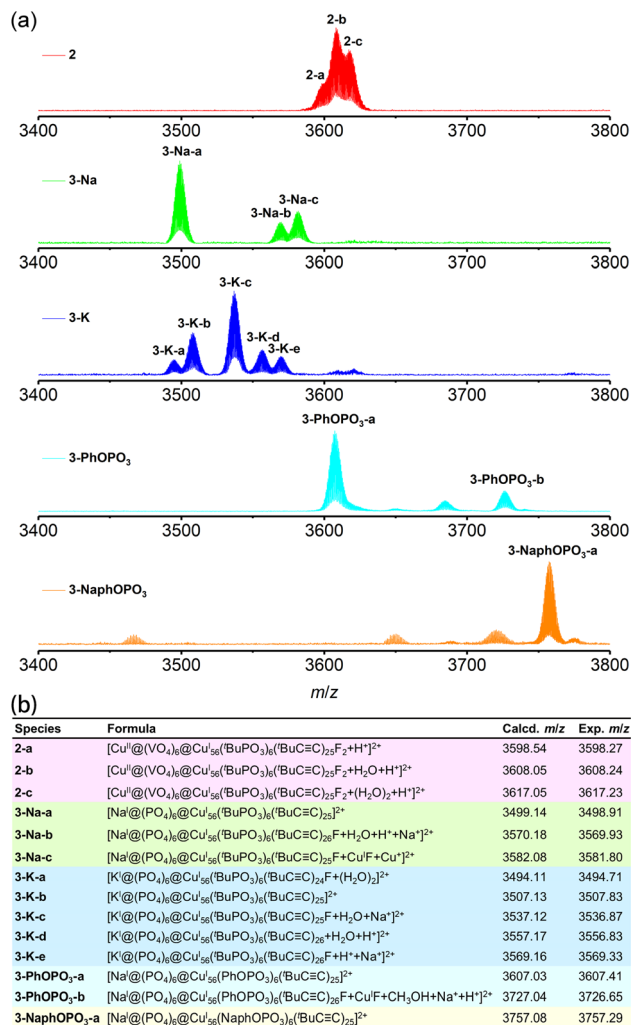


Fig. 2 Electrospray ionization mass spectra. (a) Experimental isotope distributions of selected peaks for 2–3-NaphOPO₃ (positive-ion mode, dissolved in CH₂Cl₂ and CH₃OH mixed solution). (b) The table is the formula assignment for these peaks.

further addition of H⁺, Na⁺, or H₂O, indicating that these clusters are well stabilized in solution (Fig. S10–S15[†]).

Powder X-ray diffraction (PXRD) patterns of them *versus* the calculated patterns from single-crystal X-ray diffraction revealed their high phase purity, as shown in Fig. S16[†]. Their good thermal stability under the crystalline-state below 220 °C is exhibited in their thermogravimetric analysis (TGA) diagrams (Fig. S17[†]). Taking sample 2 as an example, the major weight loss (56.9%) in the range of 300–440 °C corresponds to a large proportion of ligand (^tBuC≡C[−] and ^tBuPO₃^{2−}) decomposition, with a second weight loss (1.5%) between 440 and 800 °C that may be attributed to a fraction of ligand (^tBuC≡C[−] and ^tBuPO₃^{2−}), Cu^I, and [VO₄]^{3−} volatilization. The Fourier transform infrared (FT-IR) spectra confirmed the existence of alkynyl (2090 cm^{−1}) in these clusters (Fig. S18[†]).

Crystal structures of 1–4

Metal clusters can be synthesized using a variety of building blocks, and as previously reported, cap-shaped

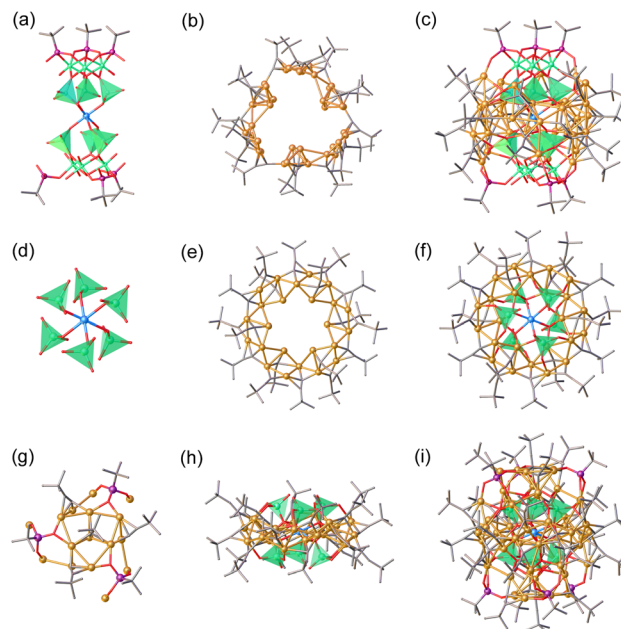


Fig. 3 The detailed structure analysis of 1 and 2. VO₄ tetrahedrons are highlighted in green. (a) [Cu^{II}(VO₄)₆@(^tBuPO₃)₆(V₃O₃OH)(V₃O₆OH)]^{18−} skeleton. (b) [Cu₃₆(^tBuC≡C)₁₈]¹⁸⁺ ring. (c) Overview of the total structure of [Cu^{II}(VO₄)₆@Cu₃₆(^tBuPO₃)₆(V₃O₃OH)(V₃O₆OH)(^tBuC≡C)₁₈]. (d) [Cu^{II}@(VO₄)₆]^{16−} kernel. (e) Wreath-shaped [Cu₂₄(^tBuC≡C)₁₈]⁶⁺ ring. (f) Wheel-shaped [Cu^{II}@(VO₄)₆@Cu₂₄(^tBuC≡C)₁₈]^{10−} unit viewed from the *c*-axis. (g) Cap-shaped [Cu₁₆(^tBuC≡C)₄(^tBuPO₃)₃]⁶⁺ peripheral structural unit. (h) Wheel-shaped [Cu^{II}@(VO₄)₆@Cu₂₄(^tBuC≡C)₁₈]^{10−} unit viewed from the *b*-axis. (i) Overview of the total structure of [Cu^{II}@(VO₄)₆@Cu₃₆(^tBuPO₃)₆(^tBuC≡C)₂₆]²⁺. Color codes: Cu^I, yellow; Cu^{II}, light blue; V, green; P, purple; O, red; C, gray. Hydrogen atoms are omitted for clarity.

[VO₂(OH)(^tBuPO₃)₂]^{2−}, [(VO₂)(^tBuPO₃)₂]^{3−}, [V₃O₆(-OH)(^tBuPO₃)₃]^{4−}, and [V₄O₈(^tBuPO₃)₄]^{4−} units have been found to be versatile for the anion-templated synthesis of high-nuclearity silver(i) clusters.^{24,37} Herein we report the successful formation of copper clusters using [VO₄]^{3−}/[PO₄]^{3−} tetrahedrons and [(^tBu/PhO/1-NaphO)PO₃]^{2−} ligands as demonstrated in Fig. 3. The resulting cluster 1 crystallizes in the *R3c* space group, being assembled from one penetrating [Cu^{II}(VO₄)₆@(^tBuPO₃)₆(V₃O₃OH)(V₃O₆OH)]^{18−} skeleton and one [Cu₃₆(^tBuC≡C)₁₈]¹⁸⁺ ring at the waist with argentophilic Cu⋯Cu bond distances in the range 2.511–3.016 Å. The [Cu₃₆(^tBuC≡C)₁₈]¹⁸⁺ ring has a van der Waals diameter of approximately 19.8 Å (Fig. S2[†]). The oxidizability of vanadate *in situ* generated a Cu²⁺ cation,³⁹ which is bridged by six [VO₄]^{3−} tetrahedrons, forming an elongated octahedral coordination geometry Cu^{II}O₆.

By reducing the addition of vanadates and reaction time, cluster 2 was successfully synthesized. Furthermore, clusters 3-Na, 3-K, 3-PhOPO₃, and 3-NaphOPO₃ were obtained by replacing vanadates with phosphonates resulting in crystal structures similar to cluster 2. For instance, SCXRD analysis and ESI-MS revealed that the overall charge balance of cluster 2 against the cationic cluster is provided by two F[−] counteranions. As



shown in Fig. 3, spherical cluster **2** has a multishelled core architecture of $\text{Cu}^{\text{II}}@(\text{VO}_4)_6@(\text{Cu}^{\text{I}})_{56}(\text{BuC}\equiv\text{C})_{26}$, with argentophilic $\text{Cu}\cdots\text{Cu}$ bond distances in the range 2.423–3.055 Å. Furthermore, the *in situ* generated Cu^{2+} cation is encapsulated in six $[\text{VO}_4]^{3-}$ templates. The wreath-shaped $[\text{Cu}_{24}(\text{BuC}\equiv\text{C})_{18}]^{6+}$ ring has a hexagonal inner Cu_{12} unit, which is connected with the $[\text{Cu}^{\text{II}}@(\text{VO}_4)_6]^{16-}$ polyoxoanion template to form the wheel-shaped $[\text{Cu}^{\text{II}}@(\text{VO}_4)_6@(\text{Cu}_{24}(\text{BuC}\equiv\text{C})_{18})]^{10-}$ unit with a van der Waals diameter of about 20.4 Å (Fig. S2†). Cluster **2** is formed by fusing two cap-like $[\text{Cu}_{16}(\text{BuC}\equiv\text{C})_4(\text{BuPO}_3)_3]^{6+}$ peripheral structural units staggered exactly by 60° on the upper and lower sides of the wheel-shaped unit. The packing pattern of crystal **2** shows that these molecules are well-arranged through intermolecular $\text{C}\cdots\text{H}\cdots\text{F}$ interactions (4.079, 4.119, and 4.244 Å) (Fig. S3†).

Both spherical copper clusters, Cu_{36} and Cu_{56} , have similar internal distorted octahedron kernels, which are often observed in POMs.⁵⁷ As shown in Fig. 4, internal elongated MO_6 octahedron skeletons are regulated by external distinct cap-shaped units. Specifically, the $\text{Cu}^{\text{II}}\text{O}_6$ octahedron of cluster **1** has $\text{Cu}^{\text{II}}\cdots\text{O}_v$ (O_v : oxygen atom of the $[\text{VO}_4]^{3-}$ templates) bond distances ranging between 2.071 and 2.083 Å, whereas the $\text{Cu}^{\text{II}}\cdots\text{O}_v$ bond distances in cluster **2**, capped by $[\text{Cu}_{16}(\text{BuC}\equiv\text{C})_4(\text{BuPO}_3)_3]^{6+}$ units, are longer and range between 2.265, 2.270 and 2.310 Å. Additionally, cluster **3-K** has a K^+ ion core with a larger ion radius than the Na^+ ion, causing a larger KO_6 octahedron ($\text{K}\cdots\text{O}_p$ (O_p : oxygen atom of the $[\text{PO}_4]^{3-}$ templates) bond distances: 2.634, 2.635, and 2.669 Å) than the NaO_6 octahedron. The steric hindrances of outer phosphonate ligands increase from BuPO_3^{2-} , PhOPO_3^{2-} to 1-NaphOPO_3^{2-} , resulting

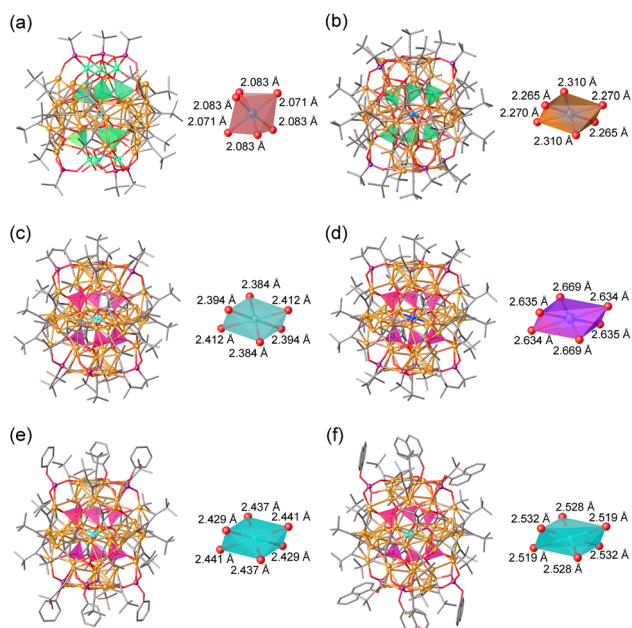


Fig. 4 Overview of the total structures of spherical Cu_{36} and Cu_{56} clusters, and their internal distorted MO_6 (M : Cu^{2+} , Na^+ , K^+) octahedron kernels. (a) **1**, (b) **2**, (c) **3-Na**, (d) **3-K**, (e) **3-PhOPO₃**, and (f) **3-1-NaphOPO₃**. Color codes: Na, light blue; K, dark blue; Cu^{I} , yellow; Cu^{II} , sky blue; V, green; P, purple; O, red; C, gray. Hydrogen atoms are omitted for clarity.

in a gradual increase of the internal NaO_6 octagon size and the change of the space group from $P2_1/n$, $C2/c$, to $P\bar{1}$.

In the reaction system where Na_2HPO_3 and K_2HPO_3 are replaced by $(\text{NH}_4)_3\text{PO}_3$, cluster **4** can be constructed from PO_4 tetrahedrons without the presence of Na^+ or K^+ ions. As shown in Fig. 5, cluster **4** possesses a peanut-shaped structure that consists of four separate $[\text{PO}_4]^{3-}$ tetrahedrons, thirty-two $\text{BuC}\equiv\text{C}^-$ ligands, six 1-NaphOPO_3^{2-} ligands, and fifty-six Cu^+ ions with argentophilic $\text{Cu}\cdots\text{Cu}$ bond distances ranging from 2.401 to 3.056 Å. The absence of metal cations inside the cluster allows for the aggregation of $[\text{PO}_4]^{3-}$ tetrahedrons into the cluster. Cluster substructures $[(\text{PO}_4)_2@(\text{Cu}^{\text{I}})_{28}(1\text{-NaphOPO}_3)_3(\text{BuC}\equiv\text{C})_{16}]^{4+}$ are bridged by four 1-NaphOPO_3^{2-} ligands, with each oxygen atom coordinated to one copper atom.

It is noteworthy that the replacement of BuPO_3^{2-} ligands with PhOPO_3^{2-} and 1-NaphOPO_3^{2-} ligands leads to the

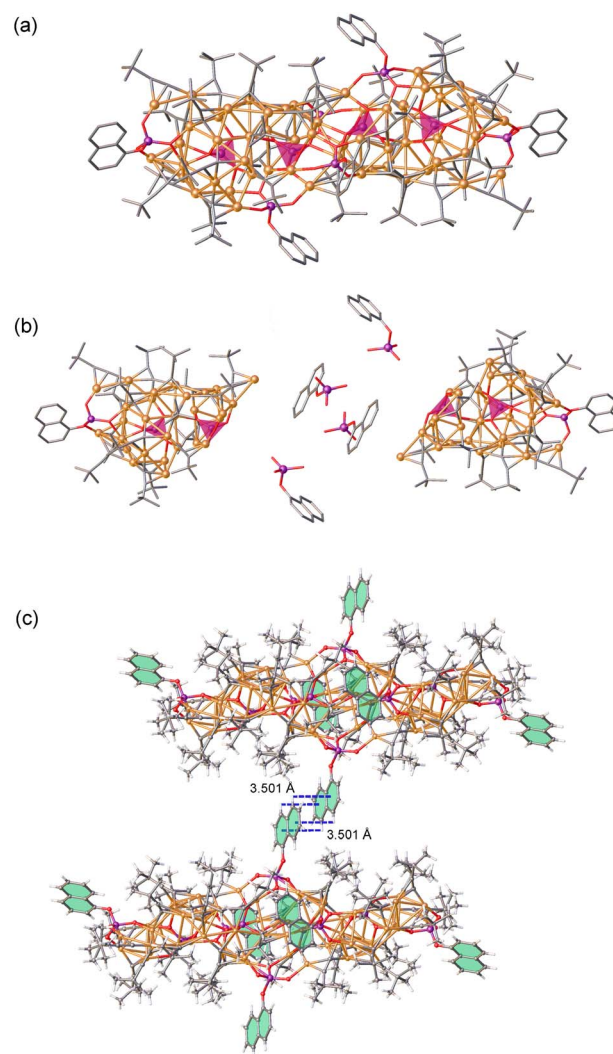


Fig. 5 The detailed structure analysis of **4**. PO_4 tetrahedrons are highlighted in pink. (a) Peanut-shaped $[(\text{PO}_4)_4@(\text{Cu}^{\text{I}})_{56}(1\text{-NaphOPO}_3)_6(\text{BuC}\equiv\text{C})_{52}]$. (b) Two asymmetric units of **4** bridged by four 1-NaphOPO_3^{2-} ligands. (c) The intermolecular $\text{C}\cdots\text{H}\cdots\pi$ interactions between 1-NaphOPO_3^{2-} ligands of adjacent clusters in **4**. Color codes: Cu, yellow; P, purple; C, gray; O, red; H, white.



formation of numerous strong intermolecular C–H $\cdots\pi$ interactions between adjacent clusters. Fig. S6–S8† demonstrate how nearby 1-NaphOPO₃²⁻ and ^tBuC \equiv C⁻ ligands in cluster **4** form intermolecular C–H $\cdots\pi$ interactions at distances ranging from 2.862 to 3.626 Å. In cluster **3-PhOPO₃**, intermolecular C–H $\cdots\pi$ interactions (at distances of 3.479 and 3.754 Å) are formed by PhOPO₃²⁻ ligands, while in cluster **3-NaphOPO₃**, intermolecular C–H $\cdots\pi$ interactions (at distances of 2.821, 3.342, and 3.598 Å) are formed by 1-NaphOPO₃²⁻ ligands. These strong interactions between adjacent clusters are crucial to their luminous properties.

The coordination modes of alkynyl and phosphonate ligands with Cu⁺ result in a diverse range of crystal structures, as demonstrated in Fig. S9.† Seven coordination modes, including μ_2 - η_{σ}^1 , η_{π}^1 ; μ_3 - η_{σ}^1 , η_{σ}^1 , η_{π}^1 ; μ_3 - η_{σ}^1 , η_{π}^1 , η_{π}^1 ; μ_3 - η_{σ}^1 , η_{σ}^1 , η_{σ}^1 ; μ_4 - η_{σ}^1 , η_{σ}^1 , η_{σ}^1 ; μ_4 - η_{σ}^1 , η_{σ}^1 , η_{π}^1 ; μ_4 - η_{σ}^1 , η_{σ}^1 , η_{π}^1 , are observed. Overall, the introduction of multiple VO₄ and PO₄ tetrahedrons with abundant terminal oxygen atoms during the synthesis of copper(i) alkynyl clusters not only balances the positive charges on the Cu(i) shells, but also effectively prevents or interrupts insoluble polymers formed by the reaction of alkynyl with Cu⁺.

Photophysical properties

The optical properties of seven Cu₃₆ and Cu₅₆ clusters in their crystalline state were investigated through a series of spectroscopic analyses (Fig. 6a). The absorption features of clusters constructed by VO₄ tetrahedrons were found to be similar. Similarly, the absorption features of clusters constructed by PO₄ tetrahedrons had peaks located at 200–400 and 720 nm (except **4**), indicating that the electronic structures of these clusters are similar and in agreement with the single crystal data. The analysis of the absorption spectra revealed that clusters constructed by PO₄ tetrahedrons have an optical band gap in the range of 1.93–2.64 eV (Fig. S19†).

Photoluminescence (PL) properties correlate with both the optical absorption behaviors and crystal structures.^{58,59} The PL spectra of cluster **2** in its crystalline state showed emission peaks at 444, 468, 507, 572, 654, and 670 nm, indicating a mechanism of multiple-exciton emission (Fig. 6b). As the temperature decreased, the luminescence intensity increased slightly, while the peak position remained unchanged. No luminescence thermochromism was observed, indicating that temperature has little effect on the tight packing and bond lengths in the crystal structure (argentophilic Cu \cdots Cu bond average distances of 2.723 Å and 2.710 Å were observed in **2** and **2-100 K**, respectively). Clusters **1** and **4** both exhibited yellow emission centered at 574 nm. Meanwhile, **2-Na**, **2-K**, **3-PhOPO₃**, and **3-NaphOPO₃** exhibited a main emission peak at 430 nm and a shoulder peak at 710 nm.

The photoluminescence quantum yield (PLQY) of these clusters ranges from 1.05% to 3.65% (ESI Table S3†). Clusters protected by PhOPO₃²⁻ and 1-NaphOPO₃²⁻ ligands exhibited three times higher luminescence intensity than those protected by ^tBuPO₃²⁻ ligands. There was a slightly enhanced quantum yield for clusters protected by PhOPO₃²⁻ and 1-NaphOPO₃²⁻ ligands. The strong intramolecular C–H $\cdots\pi$ interactions

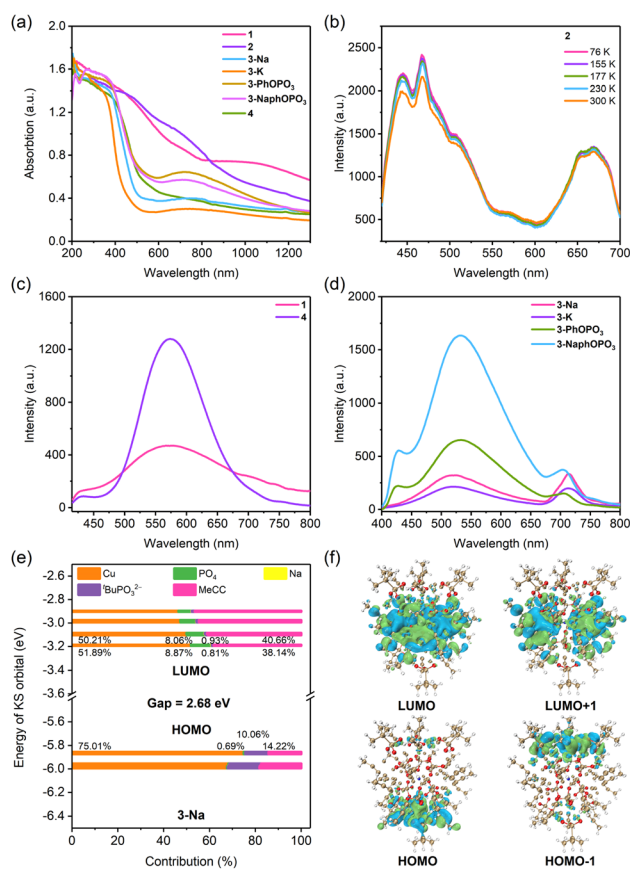


Fig. 6 Optical properties and electronic structures of clusters Cu₃₆ and Cu₅₆ in the crystalline state ($\lambda_{\text{ex}} = 380$ nm). (a) UV-Vis absorption spectra. (b) Temperature-dependent emission spectra of **2**. (c) Photoluminescence emission spectra of **1** and **4**. (d) Photoluminescence emission spectra of **3-Na**, **3-K**, **3-PhOPO₃**, and **3-NaphOPO₃**. (e) Composition analysis of frontier Kohn–Sham orbitals of **3-Na**. (f) Electronic density diagrams of the HOMO–1, HOMO, LUMO, and LUMO+1 of **3-Na** (isovalue = 0.010 a.u.).

connecting adjacent clusters restrict the rotation and vibration of some ^tBuC \equiv C⁻, PhOPO₃²⁻, and 1-NaphOPO₃²⁻ ligands and effectively decrease nonradiative decay, accounting for the enhanced emission. The PL lifetime decay curves of each metal cluster were examined to determine the key characteristics of their excited state dynamics. For **2**, decay curves of the 468 and 670 nm peaks were well fitted using a double exponential function, while other clusters required a triple exponential function for adequate fitting (Fig. S20–S22 and Table S3†). The exciton lifetimes τ_1 and τ_2 for the 670 nm peak of **2** at 300 K were determined to be 4.36 (97.72%) and 59.23 ns (2.28%), respectively, and were observed to remain relatively stable at different temperatures. In contrast, other clusters were found to have a contribution from microsecond lifetime processes; for example, in **3-Na**, the exciton lifetimes τ_1 , τ_2 , and τ_3 for the 670 nm peak were determined to be 38.20 (62.79%), 361.17 (21.31%), and 2315.16 ns (15.90%), respectively.

To further elucidate the PL mechanism, density functional theory (DFT) calculations were performed using the ORCA 5.0.3 program.^{60,61} Based on the single-crystal data, ^tBuC \equiv C⁻ was



replaced with $\text{MeC}\equiv\text{C}^-$ to reduce the computational cost. In Fig. 6 and S23–S26,[†] the highest occupied molecular orbitals (HOMO and HOMO–1) of spherical Cu_{56} clusters were mainly located on the upper and lower cap-shaped structural units, while the lowest unoccupied molecular orbitals (LUMO and LUMO+1) were distributed on the wheel-shaped structural units at the waist. Comparative analysis of the contribution of different components to the LUMO and HOMO suggests that the PL mechanism arises from alkynyl \rightarrow $\text{Cu}(\text{i})$ $^3\text{LMCT}$ (ligand to metal charge transfer) excited states that are modified by encapsulated templates and outer phosphonate ligands within the crystal structure.⁶²

Electrochemical properties

ESI-MS reveals that $^t\text{BuC}\equiv\text{C}^-$ ligands can be isolated from the as-prepared clusters in solution. Fig. S27[†] illustrates the space-filling model structures of the spherical Cu_{56} clusters with the removal of a $^t\text{BuC}\equiv\text{C}^-$ ligand exposing several active sites formed by the Cu atoms. These spherical Cu_{56} clusters have similar configurations while differing in terms of their encapsulated templates and external phosphonate ligands, which could result in distinct catalytic behaviors for the CO_2RR .⁵⁵

We conducted electrochemical experiments on **2**, **3-Na**, and **3-NaphOPO₃** clusters for the CO_2RR at different constant current densities (-50 to -500 mA cm^{-2}) in a flow cell, analyzing the gas and liquid products by gas chromatography (GC) and ^1H NMR, respectively. At low current densities HCOOH, CO, and H_2 are the primary products, while at higher current densities, C_2H_4 and $\text{C}_2\text{H}_5\text{OH}$ emerge in greater quantities (Fig. 7 and S28–S32[†]). Among these clusters, **3-Na** and **3-NaphOPO₃** display similar CO_2RR properties. As the applied current density increases (-50 to -500 mA cm^{-2}), the faradaic efficiency for C_1 products initially rises and subsequently declines gradually, whereas the faradaic efficiency for C_2 products ($\text{FE}_{\text{C}_2\text{H}_4} + \text{FE}_{\text{C}_2\text{H}_5\text{OH}}$) generally increases over the same range

of current densities. Notably, **2** exhibits higher faradaic efficiency for H_2 than **3-Na** and **3-NaphOPO₃**. At a current density of -200 mA cm^{-2} , **3-Na** demonstrates the highest faradaic efficiency for C_1 products ($\text{FE}_{\text{CO}} \sim 30\% + \text{FE}_{\text{HCOOH}} \sim 30\%$) and the lowest faradaic efficiency for H_2 ($\sim 17\%$). Under current densities ranging from -50 to -500 mA cm^{-2} , **3-NaphOPO₃** displays higher FE_{C_2} products than the two other clusters, reaching up to $\sim 45\%$ at -500 mA cm^{-2} , with respective faradaic efficiencies for C_2H_4 and $\text{C}_2\text{H}_5\text{OH}$ of $\sim 30\%$ and $\sim 15\%$. Compared with the VO_4 template, the internal PO_4 template and external naphthyl groups can promote C–C coupling, resulting in an increase in C_2 products. This observation indicates that CO_2RR properties are not solely reliant on surface morphology but also influenced by the internal VO_4/PO_4 tetrahedral template.

Moreover, we selected **3-Na** and **3-NaphOPO₃** clusters with a higher FE_{C_1} as examples to evaluate their stability using chronoamperometric measurements. The current densities of gaseous products remained stable for at least 720 minutes, as shown in Fig. 7d and S29.[†] Despite minor fluctuations in the faradaic efficiencies of CO, C_2H_4 , and H_2 at the start of the measurement, the overall current density remained relatively constant over the 720-minute period, indicating favorable catalytic sustainability. Finally, we compared these three clusters with atomically precise metal clusters exhibiting CO_2RR properties previously reported. As summarized in ESI Table 4,[†] it is worth noting that the CO selectivity of these three clusters is lower than that of gold clusters,⁴⁵ However, **3-NaphOPO₃** exhibits better faradaic efficiency for C_2 products, and copper, with its lower price, presents a more competitive alternative to gold.

Conclusions

In conclusion, we present template-assisted strategies for the customization of seven high-nuclearity copper(i) alkynyl clusters at the atomic level. Single-crystal structural analysis reveals the multishelled core architecture of spherical Cu_{36} and Cu_{56} clusters assembled by a $[\text{M}@\text{(V/PO}_4)_6]$ (M: Cu^{2+} , Na^+ , K^+) skeleton, while the peanut-shaped Cu_{56} cluster is constructed using four separate PO_4 templates. The crystalline-state emissions of these clusters are dependent on the inner templates and outer phosphonate ligands. The enhanced luminescence observed can be attributed to numerous well-arranged intermolecular C–H $\cdots\pi$ interactions between the ligands of phenyl and 1-naphthyl phosphate protected clusters. The electrocatalytic properties of CO_2 reduction reaction are affected by both the surface morphology and the internal VO_4/PO_4 template. Specifically, internal PO_4 templates and external naphthyl groups could promote an increase in C_2 products. This study provides ingenious strategies for tailoring the structures of nanoclusters at an atomically precise level and highlights the significance of atomic-level comparative studies of structure–property relationships.

Data availability

All experimental and computational data are included in the ESI.[†] Crystallographic data for the structures reported in this

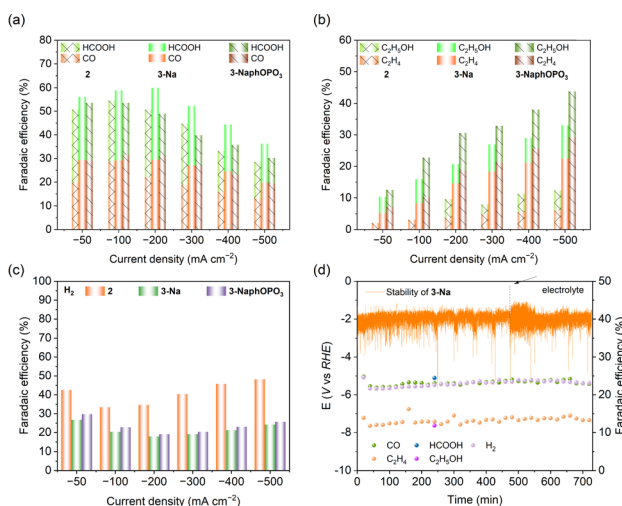


Fig. 7 Electrochemical performances. (a)–(c) Faradaic efficiencies of **2**, **3-Na**, and **3-NaphOPO₃** at different applied current densities (0.5 M KHCO_3). (d) Stability tests of the electrocatalysts for CO_2 reduction (**3-Na** at -200 mA cm^{-2}).



article have been deposited at the Cambridge Crystallographic Data Centre under deposition numbers CCDC 2219441, 2219447, 2116536, 2219449, 2219459, 2219463, 2219464, and 2285409.†

Author contributions

Y.-P. X. and X. L. conceived the original idea. Y.-P. X. designed and directed the synthesis experiments. J.-J. F. performed the experiments and analyzed the data. J.-J. F., Z. L., and Y.-L. S. performed the structural characterization. J.-J. F. collected and analyzed the ESI-MS, PL, and CO₂RR data. Y.-P. X. directed the DFT calculations. J.-J. F. conducted the DFT calculations. All authors have given approval to the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We gratefully acknowledge financial support from the National Natural Science Foundation of China (No. 21771071, 22171094, 21925104, and 92261204) and the Hubei Provincial Natural Science Foundation of China (No. 2021CFA020). We thank the staff in the Analytical and Testing Center at the Huazhong University of Science and Technology for all related measurements. We thank the staff at the BL17B beamline of the National Center for Protein Sciences Shanghai (NCPSS) at Shanghai Synchrotron Radiation Facility for the assistance with data collection.

References

- R. Hamze, J. L. Peltier, D. Sylvinson, M. Jung, J. Cardenas, R. Haiges, M. Soleilhavoup, R. Jazzar, P. I. Djurovich, G. Bertrand and M. E. Thompson, *Science*, 2019, **363**, 601–606.
- M. Cao, R. Pang, Q. Wang, Z. Han, Z. Wang, X. Dong, S. Li, S. Zang and T. C. W. Mak, *J. Am. Chem. Soc.*, 2019, **141**, 14505–14509.
- S. Li, X. Du, B. Li, J. Wang, G. Li, G. Gao and S. Zang, *J. Am. Chem. Soc.*, 2018, **140**, 594–597.
- R. Huang, Y. Wei, X. Dong, X. Wu, C. Du, S. Zang and T. C. W. Mak, *Nat. Chem.*, 2017, **9**, 689–697.
- O. Fuhr, S. Dehnen and D. Fenske, *Chem. Soc. Rev.*, 2013, **42**, 1871–1906.
- H. Schmidbaur and A. Schier, *Chem. Soc. Rev.*, 2012, **41**, 370–412.
- Y. Sun, E. Wang, Y. Ren, K. Xiao, X. Liu, D. Yang, Y. Gao, W. Ding and Y. Zhu, *Adv. Funct. Mater.*, 2019, **29**, 1904242.
- Z. Wang, M. Wang, Y. Li, P. Luo, T. Jia, R. Huang, S. Zang and T. C. W. Mak, *J. Am. Chem. Soc.*, 2018, **140**, 1069–1076.
- P. Yuan, R. Chen, X. Zhang, F. Chen, J. Yan, C. Sun, D. Ou, J. Peng, S. Lin, Z. Tang, B. K. Teo, L. S. Zheng and N. Zheng, *Angew. Chem., Int. Ed.*, 2019, **58**, 835–839.
- R. W. Huang, X. Y. Dong, B. J. Yan, X. S. Du, D. H. Wei, S. Q. Zang and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2018, **57**, 8560–8566.
- L. L. Zhang and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2017, **56**, 16228–16232.
- X. Chang, K. Low, J. Wang, J. Huang and C. Che, *Angew. Chem., Int. Ed.*, 2016, **55**, 10312–10316.
- S. S. Y. Chui, M. F. Y. Ng and C. Che, *Chem.–Eur. J.*, 2005, **11**, 1739–1749.
- R. P. Brocha Silalahi, G. Huang, J. Liao, T. Chiu, K. K. Chakrahari, X. Wang, J. Cartron, S. Kahlal, J. Saillard and C. W. Liu, *Inorg. Chem.*, 2020, **59**, 2536–2547.
- Y. Jin, S. Li, Z. Han, B. J. Yan, H. Y. Li, X. Y. Dong and S. Q. Zang, *Angew. Chem., Int. Ed.*, 2019, **58**, 12143–12148.
- K. K. Chakrahari, R. P. B. Silalahi, T. H. Chiu, X. Wang, N. Azrou, S. Kahlal, Y. C. Liu, M. H. Chiang, J. Y. Saillard and C. W. Liu, *Angew. Chem., Int. Ed.*, 2019, **58**, 4943–4947.
- T. C. Higgs, P. J. Bailey, S. Parsons and P. A. Tasker, *Angew. Chem., Int. Ed.*, 2002, **41**, 3038–3041.
- P. Putaj and F. Lefebvre, *Coord. Chem. Rev.*, 2011, **255**, 1642–1685.
- H. Lv, Y. V. Geletii, C. Zhao, J. W. Vickers, G. Zhu, Z. Luo, J. Song, T. Lian, D. G. Musaev and C. L. Hill, *Chem. Soc. Rev.*, 2012, **41**, 7572–7589.
- J. Zhang, Y. Huang, G. Li and Y. Wei, *Coord. Chem. Rev.*, 2019, **378**, 395–414.
- J. Liu, L. Feng, H. Su, Z. Wang, Q. Zhao, X. Wang, C. Tung, D. Sun and L. Zheng, *J. Am. Chem. Soc.*, 2018, **140**, 1600–1603.
- Z. Wang, H. Su, M. Kurmoo, C. Tung, D. Sun and L. Zheng, *Nat. Commun.*, 2018, **9**, 2094.
- Y. Xie, J. Jin, X. Lu and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2015, **54**, 15176–15180.
- Y. Xie and T. C. W. Mak, *J. Am. Chem. Soc.*, 2011, **133**, 3760–3763.
- G. Gao, P. Cheng and T. C. W. Mak, *J. Am. Chem. Soc.*, 2009, **131**, 18257–18259.
- L. Chen, F. Jiang, Z. Lin, Y. Zhou, C. Yue and M. Hong, *J. Am. Chem. Soc.*, 2005, **127**, 8588–8589.
- S. Wang and G. Yang, *Chem. Rev.*, 2015, **115**, 4893–4962.
- F. A. Paz, J. Klinowski, S. M. Vilela, J. P. Tome, J. A. Cavaleiro and J. Rocha, *Chem. Soc. Rev.*, 2012, **41**, 1088–1110.
- K. J. Gagnon, H. P. Perry and A. Clearfield, *Chem. Rev.*, 2012, **112**, 1034–1054.
- J. Goura and V. Chandrasekhar, *Chem. Rev.*, 2015, **115**, 6854–6965.
- L. Zhang, R. Clérac, P. Heijboer and W. Schmitt, *Angew. Chem., Int. Ed.*, 2012, **51**, 3007–3011.
- M. Haga, K. Kobayashi and K. Terada, *Coord. Chem. Rev.*, 2007, **251**, 2688–2701.
- S. Bao and L. Zheng, *Coord. Chem. Rev.*, 2016, **319**, 63–85.
- X. W. Lv, C. C. Weng, Y. P. Zhu and Z. Y. Yuan, *Small*, 2021, **17**, 2005304.
- Y. Shen, J. Jin, J. Fang, Z. Liu, J. Shi, Y. Xie and X. Lu, *Inorg. Chem.*, 2021, **60**, 6276–6282.
- J. Jin, Y. Xie, H. Cui, G. Duan, X. Lu and T. C. W. Mak, *Inorg. Chem.*, 2017, **56**, 10412–10417.



- 37 Y. Xie and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2012, **51**, 8783–8786.
- 38 L. Zhang and T. C. W. Mak, *J. Am. Chem. Soc.*, 2016, **138**, 2909–2912.
- 39 L. L. Zhang, G. Zhou, G. Zhou, H. Lee, N. Zhao, O. V. Prezhdo and T. C. W. Mak, *Chem. Sci.*, 2019, **10**, 10122–10128.
- 40 J. Fang, Y. Shen, Z. Liu, C. Liu, Y. Xie and X. Lu, *Inorg. Chem.*, 2021, **60**, 13493–13499.
- 41 V. W. W. Yam and K. K. W. Lo, *Chem. Soc. Rev.*, 1999, **28**, 323–334.
- 42 Z. N. Chen, N. Zhao, Y. Fan and J. Na, *Coord. Chem. Rev.*, 2009, **253**, 1–20.
- 43 H. Xu, D. Rebollar, H. He, L. Chong, Y. Liu, C. Liu, C. J. Sun, T. Li, J. V. Muntean, R. E. Winans, D. J. Liu and T. Xu, *Nat. Energy*, 2020, **5**, 623–632.
- 44 S. Nitopi, E. Bertheussen, S. B. Scott, X. Liu, A. K. Engstfeld, S. Horch, B. Seger, I. Stephens, K. Chan, C. Hahn, J. K. Nørskov, T. F. Jaramillo and I. Chorkendorff, *Chem. Rev.*, 2019, **119**, 7610–7672.
- 45 D. Raciti and C. Wang, *ACS Energy Lett.*, 2018, **3**, 1545–1556.
- 46 Y. Song, Y. Li, M. Zhou, X. Liu, H. Li, H. Wang, Y. Shen, M. Zhu and R. Jin, *Sci. Adv.*, 2021, **7**, eabd2091.
- 47 X. Ma, J. Jia, P. Luo, Z. Wang, S. Zang and T. C. W. Mak, *Nano Res.*, 2022, **15**, 5569–5574.
- 48 S. Nematulloev, R. W. Huang, J. Yin, A. Shkurenko, C. Dong, A. Ghosh, B. Alamer, R. Naphade, M. N. Hedhili, P. Maity, M. Eddaoudi, O. F. Mohammed and O. M. Bakr, *Small*, 2021, **17**, 2006839.
- 49 J. Wang, F. Xu, Z. Y. Wang, S. Q. Zang and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2022, **61**, e202207492.
- 50 Z. J. Guan, J. J. Li, F. Hu and Q. M. Wang, *Angew. Chem., Int. Ed.*, 2022, **61**, e202209725.
- 51 H. Seong, V. Efremov, G. Park, H. Kim, J. S. Yoo and D. Lee, *Angew. Chem., Int. Ed.*, 2021, **133**, 14684–14691.
- 52 S. Verma, Y. Hamasaki, C. Kim, W. Huang, S. Lu, H. M. Jhong, A. A. Gewirth, T. Fujigaya, N. Nakashima and P. J. A. Kenis, *ACS Energy Lett.*, 2018, **3**, 193–198.
- 53 X. Wan, J. Wang, Z. Nan and Q. Wang, *Sci. Adv.*, 2017, **3**, e1701823.
- 54 Q. Tang, Y. Lee, D. Li, W. Choi, C. W. Liu, D. Lee and D. Jiang, *J. Am. Chem. Soc.*, 2017, **139**, 9728–9736.
- 55 L. J. Liu, Z. Y. Wang, Z. Y. Wang, R. Wang, S. Q. Zang and T. C. W. Mak, *Angew. Chem., Int. Ed.*, 2022, **61**, e202205626.
- 56 X. Ma, F. Sun, L. Qin, Y. Liu, X. Kang, L. Wang, D. Jiang, Q. Tang and Z. Tang, *Chem. Sci.*, 2022, **13**, 10149–10158.
- 57 P. Yang and U. Kortz, *Acc. Chem. Res.*, 2018, **51**, 1599–1608.
- 58 S. Peng, H. Yang, D. Luo, M. Xie, W. Tang, G. Ning and D. Li, *Inorg. Chem. Front.*, 2022, **9**, 5327–5334.
- 59 Y. Xie, Y. Shen, G. Duan, J. Han, L. Zhang and X. Lu, *Mater. Chem. Front.*, 2020, **4**, 2205–2222.
- 60 T. Lu and F. Chen, *J. Comput. Chem.*, 2012, **33**, 580–592.
- 61 F. Neese, F. Wennmohs, U. Becker and C. Riplinger, *J. Chem. Phys.*, 2020, **152**, 224108.
- 62 G. F. Manbeck, W. W. Brennessel, R. A. Stockland and R. Eisenberg, *J. Am. Chem. Soc.*, 2010, **132**, 12307–12318.

