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Effects of sponge-derived polybrominated diphenyl ethers on human cancer cell α -*N*-acetylgalactosaminidase and bacterial α -*D*-galactosidase and their antioxidant activity†

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During the search for glycosidase inhibitors among marine natural products that are environmental contaminants, we evaluated for the first time the potency of polybrominated diphenyl ethers 1–7 (5: OH-BDE47, 6: OH-BDE85, 7: OH-BDE137) isolated from *Dysidea* sponges on the hydrolyzing activity of cancer α -*N*-acetylgalactosaminidase (α -NaGalase) and the potency of dihydroxylated PBDEs 1 and 2 on the hydrolyzing activity of the recombinant α -*D*-galactosidase (α -PsGal) from the marine bacterium *Pseudoalteromonas* sp. KMM 701. The antioxidant activity of the spongean metabolites was assessed to test whether the antioxidant activity affects the activity of the enzymes under study. All OH-PBDEs/MeO-PBDEs tested had no effects on human cancer cell α -*N*-acetylgalactosaminidase. Dihydroxylated PBDEs 1 and 2 at micromolar levels quickly and irreversibly inhibited the bacterial α -*D*-galactosidase. Compounds 1–7 were shown to have low antioxidant activity at millimolar levels. It is likely that antioxidant activity is not responsible for the effectiveness of α -*D*-galactosidase inhibition due to the large difference in the concentrations required to display antioxidant activity (millimolar) and to inhibit an enzyme (micromolar).

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Environmental significance

PBDEs are under continuous research as naturally produced metabolites and synthetically produced brominated flame retardants. Both groups of PBDEs are environmental contaminants. One of the most abundant natural sources of PBDEs with various degrees of hydroxylation, methoxylation, and bromination is marine *Dysidea* sponges. Spongean PBDEs can be accumulated in the marine food web and transferred to humans. Enzymes glycoside hydrolases, such as α -*D*-galactosidase and α -*N*-acetylgalactosaminidase, may be one of the targets of OH-PBDEs effects on humans. Our results demonstrate that OH-PBDEs at micromolar levels irreversibly inhibit bacterial α -*D*-galactosidase of the GH36 family but have no effects on human cancer cell α -*N*-acetylgalactosaminidase.

Introduction

PBDEs are under continuous research as naturally produced metabolites and synthetically produced polybrominated flame retardants.¹ Synthetically produced PBDEs are widespread environmental contaminants, which are transformed by biological and nonbiological ways, and are distributed in soil, groundwater, and atmosphere.^{2–5} PBDEs, together with their hydroxylated and methoxylated analogues, have been identified ubiquitously in humans, wildlife, and marine biota.^{6–8}

Toxicology reports indicate that PBDEs cause thyroid dysfunction,^{9,10} affect ovarian granulosa cell function and follicular development,¹¹ and interfere with the male reproductive system.¹² Moreover, they have neurotoxic,¹³ carcinogenic¹⁴ and other toxic effects on reproductive health,¹⁵ including the effects on reproductive outcomes among women undergoing assisted reproductive technologies¹⁶ and declining semen quality.¹²

Marine organisms, in particular marine sponges, are an inexhaustible source of metabolites with various biological activities.¹⁷ Naturally produced PBDEs with various degrees of hydroxylation, methoxylation, and bromination have been mainly isolated with high yields from tropical marine sponges of the *Dysideidae* family.^{18–22} Spongean PBDEs possess antimicrobial,^{19,23} antifungal,²⁴ and antiviral²⁵ activities, inhibit human aldose reductase²⁶ and a range of enzymes that are implicated in tumor development, such as inosine monophosphate dehydrogenase, guanosine monophosphate synthetase,²⁷ 15-lipoxygenase,²⁸ and Tie2 kinase.²⁹ They inhibit the assembly of

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† Electronic supplementary information (ESI) available: Physical properties for compounds 3–7; dose–response curves of α -PsGal inhibition with compounds 1–3; the α -PsGal activity after PBDEs 1 and 2 treatment; effects of compounds 1–7 on cancer cell α -NaGalase. See DOI: <https://doi.org/10.1039/d2va00269h>



microtubuline proteins,³⁰ regulate interleukin-8 production,³¹ and exhibit anti-proliferative activity against cancer cells adapted to glucose starvation.³² Spongian PBDEs activate KCNQ potassium channels, which predominantly express in the heart and cardiomyocytes.³³ Screening of 300 naturally occurring compounds showed 4,6-dibromo-2-(2',4'-dibromophenoxy)phenol and 4,5,6-tribromo-2-(2',4'-dibromophenoxy)phenol as substances with the potential for anticancer therapy.³⁴ Spongian PBDEs are bioaccumulated in the marine food web with the potential transfer to humans through seafood consumption; dietary ingestion is understood to be an important uptake pathway for many OH/MeO-PBDEs.^{35–38} Due to the consumption of seafood containing spongian OH-PBDEs,³⁵ these compounds can get into humans and affect a range of enzymes, among which are glycoside hydrolases, such as α -D-galactosidase (α -Galase) and α -N-acetylgalactosaminidase (α -NaGalase).

α -D-Galactosidases catalyze the hydrolysis of the non-reducing terminal α -D-galactose from α -D-galactosides, galactooligosaccharides, and polysaccharides, such as galactomannans, galactolipids, and glycoproteins. It is known that α -galactosidase activity does not exist in the human intestine mucosa, and non-digestible α -galactosides are exclusively hydrolyzed by GH36 family α -galactosidases of the human intestinal symbiotic bacteria.^{39,40} So, α -Galase of the human intestinal symbiotic bacteria may be one of the targets of harmful effects of OH-PBDEs. Human α -N-acetylgalactosaminidase (α -NaGalase) is a lysosomal glycohydrolase, which hydrolyzes the O-glycosidic linkage between α -N-acetylgalactosamine and glycoproteins. α -NaGalase is produced by all

cancer cells and accumulates in the blood plasma of cancer patients, leading to immunosuppression in patients with advanced cancer.⁴¹ The level of enzyme activity increases in the blood serum at the initial stage of the disease and the stage of metastasis.⁴² Thus, α -NaGalase, an immunosuppressive agent in cancer patients, is considered as a potential therapeutic agent in cancer treatment. One of the causes of some diseases, including cancer, viral infections, and metabolic and genetic disorders, is an imbalance in glycosidase functions;^{43–46} glycosidase inhibitors can regulate or block the enzyme imbalance. So, much attention is paid to identifying potential inhibitors of α -NaGalase^{47,48} and α -Galase.^{49,50}

In the last years, marine organisms have received increased attention as a source of enzyme inhibitors due to their potential use in pharmacology.^{51,52} To date, few attempts have been made to find inhibitors of glycoside hydrolases, such as α -Galase and α -NaGalase, among low molecular marine natural products. The previously characterized low molecular marine inhibitors of α -galactosidase from the marine bacterium *Pseudoalteromonas* sp. (α -PsGal) are polyhydroxynaphthoquinone pigments from sea urchins and their synthetic analogues,⁵³ polybrominated diphenyl ether and dibenzo-*p*-dioxins from *Dysidea* sponges,⁵⁴ pentacyclic guanidine alkaloids from the sponge *Monanchora pulchra*,⁵⁵ and marine alkaloids of aaptamine and makaluvamine series.⁵⁶ The effect of spongian polybrominated diphenyl ethers on the activity of cancer α -NaGalase has not been previously studied.

The main goal of this study was to evaluate *in vitro* spongian metabolites 1–7 (Fig. 1) isolated from *Dysidea* sponges, some of

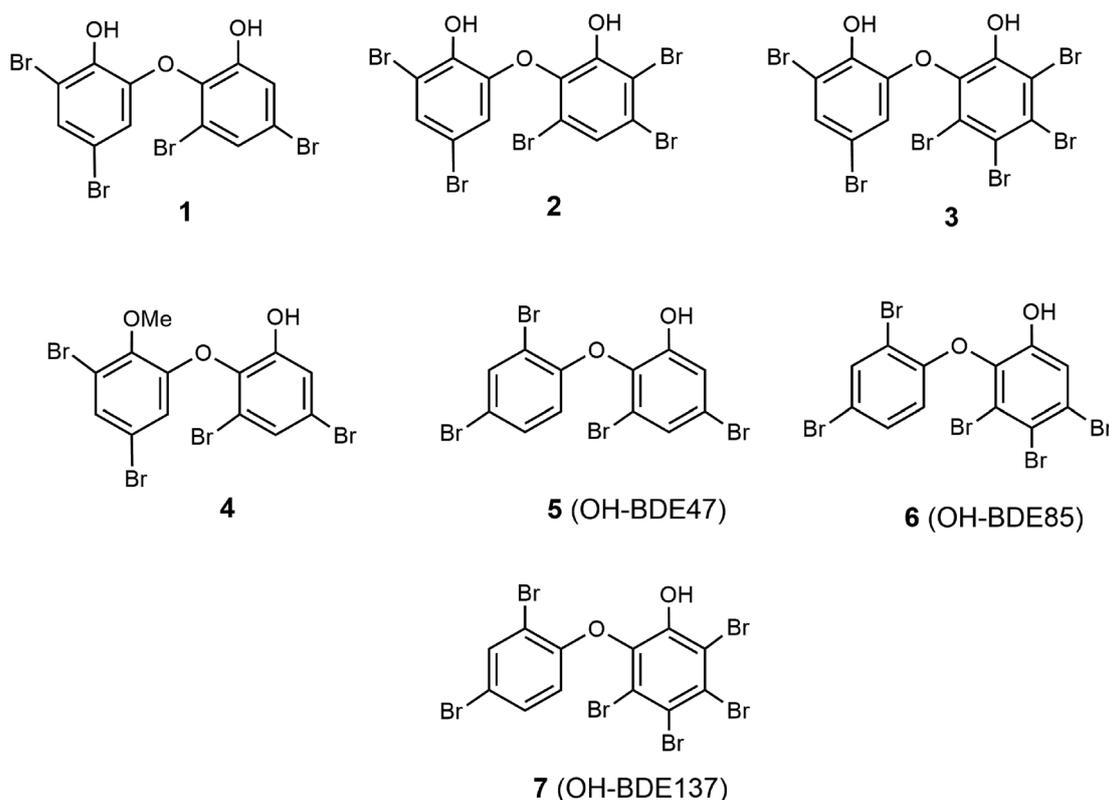


Fig. 1 Structures of spongian PBDEs 1–7.



which are environmental contaminants (5: OH-BDE47, 6: OH-BDE85, 7: OH-BDE137), on the hydrolyzing activity of α -NaGalase from cancer cells; additionally, to evaluate the effect of spongean metabolites 1 and 2 that had not been previously studied on the hydrolyzing activity of α -PsGal of the GH36 family from the marine bacterium *Pseudoalteromonas* sp. KMM 701.⁵⁷ This enzyme can be used as a suitable model for studying the potentially harmful effects of spongean OH-PBDEs. As polybrominated diphenyl ethers can potentially be reductive inhibitors of α -PsGal and α -NaGalase due to their phenolic moieties, we evaluated their antioxidant activity to check whether the antioxidant potency influences the activity of enzymes under research.

Materials and methods

Chemicals

Roswell Park Memorial Institute Medium (RPMI 1640), phosphate buffered saline (PBS), L-glutamine, penicillin-streptomycin solution, trypsin and fetal bovine serum (FBS), sodium hydrocarbonate (NaHCO₃), and agar were purchased from BioloT (Bolshoy Sampsonievsky Ave, St. Petersburg, Russia), *p*-nitrophenyl-*N*-acetyl- α -D-galactosaminide (*p*-NPNA- α -Gal), *p*-nitrophenyl- α -D-galactopyranoside (*p*NP- α -Gal), bovine serum albumin (BSA), and Bradford reagent were purchased from Sigma-Aldrich (St. Louis, MO, USA); IMAC Ni2+Sepharose, Q-Sepharose, Mono-Q, and Superdex-200 PG were purchased from GE Healthcare (Uppsala, Sweden); pET 40 b(+) plasmid was purchased from Invitrogen, Carlsbad, CA, USA; recombinant protein markers SDS-PAGE-electrophoresis – from BioRad (1000 Alfred Nobel Drive, Hercules, CA, USA). Silica gel (Sorb-polymer, Krasnodar, Russia), Sephadex LH-20 (Pharmacia Fine Chemicals) were used for column chromatography, Sorbfil plates coated with silica gel F254 (Sorb-polymer, Krasnodar, Russia) were used for TLC; linoleic acid, 99% (Sigma-Aldrich) and 2,2'-azobis(2-amidinopropane) dihydrochloride (AAPH) were used for assessing antioxidant activity.

Equipment

Microplate spectrophotometer (BioTek Instruments, Highland Park, Winooski, VT, USA) was used for measuring the optical density at 400 nm (D400). Ultrasonic homogenizer Bandelin Sonopuls (Bandelin electronic GmbH & Co., KG Heinrichstraße 3-4 12, 207, Berlin, Germany) was used for homogenization. GenBAflex-tubes 6–8 kDa (Scienova GmbH, Wildenbruchstabe 15 07745 Jena, Germany) were used for dialysis. ¹H NMR and ¹³C NMR spectra (δ in ppm, referenced to the solvent used) were recorded on a Bruker AVANCE DRX-500 NMR spectrometer at 500 and 125 MHz, respectively. High-resolution electron ionization mass spectrometry (HREI MS) was performed on an AMD-604 S mass spectrometer (Intectra, Germany). Melting points were determined on a Boetius apparatus.

Animal material

We used a freeze-dried sample of a sponge *Dysidea* sp. collected during scientific cruise of R/V "Academik Oparin", which is kept in the collection of Pacific Institute of Bioorganic Chemistry.

Extraction, isolation, and identification

Reisolation of compounds 3–7 from repository samples of *Dysidea* sponge was described earlier.⁵⁴ Data for compounds 3–7 are given in ESI (Tables S1 and S2†). A freeze-dried sponge sample *Dysidea* sp. (6 g) was extracted with CHCl₃. The CHCl₃ extract was concentrated under reduced pressure to yield a green gum. This material was partitioned into hexane solubles and hexane insolubles. The hexane-insoluble fraction contained a mixture of phenolic compounds. This mixture was chromatographed on a silica gel column in CHCl₃–hexane (5 : 1) to get compound 1 and then in CHCl₃ to obtain a mixture (1 : 2) of compounds 2 and 3. Compound 1 was purified by repeated column chromatography on a Sephadex LH-20 column in CHCl₃–MeOH (1 : 1). Crystallization from CHCl₃ gave pure 1 (1.8 mg, 0.03% on dry weight of the sponge) identified as 2-(3',5'-dibromo-2'-hydroxyphenoxy)-3,5-dibromophenol.⁵⁸ A mixture of compounds 2 and 3 was separated by repeated column chromatography on Sephadex LH-20 in CHCl₃–MeOH (1 : 1) to get 2 and 3. Pure 2 (9.0 mg, 0.15% on dry weight of the sponge) was obtained by several recrystallizations from acetone and identified as 2-(3',5'-dibromo-2'-hydroxyphenoxy)-3,4,6-tribromophenol.^{18–21}

Compound 1. Colourless needles (CHCl₃); mp 181–183 °C; ¹H NMR (acetone-*d*₆), δ _H: 7.38 (1H, d, *J* = 2.1 Hz), 7.24 (1H, d, *J* = 2.1 Hz), 7.37 (1H, d, *J* = 2.1 Hz), 6.64 (1H, d, *J* = 2.1 Hz); ¹³C NMR, δ _C: 152.6 (C), 146.7 (C), 144.5 (C), 139.3 (C), 129.1 (CH), 126.8 (CH), 120.7 (CH), 119.6 (C), 118.6 (C), 116.3 (CH), 111.2 (C), 110.7 (C); HREI MS, *m/z*: 517.7698 (calcd for C₁₂H₆O₃⁷⁹Br₂⁸¹Br₂ 517.7539).

Compound 2. Colourless needles (acetone); mp 183–185 °C; ¹H NMR (acetone-*d*₆), δ _H: 7.47 (1H, s), 7.39 (1H, d, *J* = 2.1 Hz), 6.80 (1H, d, *J* = 2.1 Hz); ¹³C NMR, δ _C: 151.7 (C), 147.1 (C), 145.4 (C), 140.7 (C), 130.1 (CH), 128.1 (CH), 123.6 (C), 117.4 (C), 117.1 (CH), 115.4 (C), 112.2 (C), 111.2 (C); HREI MS, *m/z*: 595.6091, 597.6098 (calcd for C₁₂H₅O₃⁷⁹Br₃⁸¹Br₂ 595.6110; calcd for C₁₂H₅O₃⁷⁹Br₂⁸¹Br₃ 597.6091).

Production and purification of the recombinant α -D-galactosidase

The recombinant wild-type α -PsGal was produced as described earlier.⁵⁹ The plasmid DNA pET-40b(+) containing an insertion of the gene from the marine bacterium *Pseudoalteromonas* sp. KMM 701 encoding α -PsGal was transformed in the *Escherichia coli* strain Rosetta (DE3). Heterologous expression was carried out at optimal conditions as described previously.⁶⁰ Purification of the recombinant α -PsGal was performed according to the procedures described previously.^{59,60} The purification quality was assessed by 12% Laemmli-SDS-PAGE-electrophoresis.⁶¹

Cell cultured

Cancer cells RPMI-7951 (ATCC® no. HTB-66™), MDA-MB-231 (ATCC® no. HTB-26™), DLD-1 (ATCC® no. CCL-221), HT-29 (ATCC® no. HTB-38™), HCT-116 (ATCC® no. CCL-247), SK-MEL-28 (ATCC® no. HTB-72™), and mouse healthy epidermal cells JB6 Cl 41 (ATCC® no. CRL-2010™) were obtained from the



American Type Culture Collection (Manassas, VA, USA). Cells were cultured in an appropriate culture medium supplemented with 10% FBS and 1% penicillin–streptomycin solution. The cell cultures were maintained at 37 °C in a humidified atmosphere containing 5% CO₂.

Preparation of cell lysate

Cell lysates were prepared according to the procedure described previously.⁵⁶

Isolation and purification of α -NaGalase from cell lysates

α -NaGalase was isolated from cell lysates of melanoma RPMI-7951 and SK-Mel-28, breast cancer MDA-MB-231, intestinal carcinoma DLD-1, HT-29 and HCT-116, and healthy JB6 Cl 41 mouse epidermal cells according to the procedure described previously.⁶² The α -NaGalase activity was found in 75% ammonium sulphate precipitate. The purification quality was assessed by 12% Laemmli-SDS-PAGE.⁶¹

Enzyme and protein assays

Enzyme and protein assays were carried out on 96-well plates. The mixture, placed in one cell, contained 50 μ L of the enzyme solution and 100 μ L solution of *p*NP- α -Gal (1 mg mL⁻¹ (3.3 mM) 0.05 M sodium phosphate buffer (pH 7.3)) or *p*-NPNA- α -Gal (3 mg mL⁻¹ (8.8 mM) 0.05 M sodium citrate buffer (pH 5.0)) as the substrates for α -PsGal or α -NaGalase, respectively. The mixture was incubated at 20 °C or 37 °C for 10 or 60 min for α -PsGal or α -NaGalase, respectively. The reaction was stopped by adding 150 μ L of 1 M Na₂CO₃. The amount of the released *p*NP was determined spectrophotometrically at 400 nm ($\epsilon_{400} = 18\ 300\ \text{M}^{-1}\ \text{cm}^{-1}$). One unit of the standard activity was determined as the amount of the enzyme that releases 1 μ mol of *p*NP per 1 min at 20 °C or 1 nmol of *p*NP per 1 h at 37 °C for α -PsGal or α -NaGalase, respectively. The specific activity was calculated as units per mg protein. The protein concentrations were estimated by the Bradford method with BSA as a standard.⁶³ Buffer solutions of the α -PsGal (0.01 units per mL stock solution) and α -NaGalase (0.001 units per mL stock solution) were used in further experiments.

Optical density was measured on a Microplate spectrophotometer at 400 nm. Results were read with a computer program Gen5 and treated with Excel computer program. All calculations were based on reactions with the consumption of 10% of the chromogenic substrate.

Determination of inhibitory potency of PBDEs 1–7 for α -NaGalase and PBDEs 1, 2, and 3 for α -PsGal

To preliminary test the inhibitory potency of the compounds, 25 μ L of the enzyme solutions of α -PsGal in 0.05 M sodium phosphate buffer (pH 7.0) were preincubated for 25 min at 20 °C with 2 μ L of ethanol solution of a test compound at various concentrations (from 10⁻³ M to 10⁻⁷ M in a probe) to allow enzyme and inhibitor interaction. In the control sample, 2 μ L of EtOH was added to 25 μ L of enzyme solution. The enzyme reaction was initiated by adding 120 μ L of the substrate *p*NP- α -

Gal (3.3 mM) for α -PsGal in the respective buffer solutions. After 10 min, for α -PsGal, the reaction was stopped by adding 150 μ L of 1 M Na₂CO₃. The amount of *p*NP was quantified by spectrophotometric detection at 400 nm. Results were evaluated as the percent of inhibition as described above.

The inhibitory activity was assessed according to the described method, and the enzyme α -PsGal was kept for 25 min in the presence of different concentrations of compounds, and after that, its standard activity was determined.^{54,55} Inhibition (%) was calculated with the formula: inhibition (%) = (1 - A/A₀) × 100, where A is the standard activity of enzymes under the action of an inhibitor or other factors, A₀ is the standard activity of enzymes in the absence of the influence of any factors. The concentration of inhibitors causing 50% inactivation of the enzyme (IC₅₀) was determined by fitting the dependences of the inhibition (%) on the concentration of the inhibitors to the Hill equation nonlinear regression with the computer software Origin 8.1 (OriginLab, Northampton, Massachusetts, USA).

To study the effect of compounds 1–7 on α -NaGalase, 5 μ L of the compound in ethanol was added to 25 μ L of the enzyme in 0.05 M sodium citrate buffer (pH 5.0) and kept for 30 min. The reaction was initiated by adding 120 μ L of the 8.8 mM substrate (*p*-NPNA- α -Gal) and incubating at 37 °C for 90 min. The reaction was stopped by adding 150 μ L of 1 M Na₂CO₃. The concentration of compounds 1, 2, 3, 4, 5, 6, and 7 in the sample (mM) was 1.3; 1.1; 1.0; 1.3; 1.0; 1.2, and 1.5, respectively. The following mixtures served as controls:

- (1) Excluding the effect of ethanol on the enzyme, 25 μ L of the enzyme and 5 μ L of ethanol were kept for 30 min, and then 120 μ L of the substrate was added and incubated for 90 min;
- (2) Mixtures controlling the absorption of the compound at 400 nm containing 25 μ L of working buffer and 5 μ L of the compound were kept for 30 min; then, 120 μ L of buffer was added and incubated for 90 min;
- (3) Mixtures controlling the absorbance of the substrate at 400 nm containing 25 μ L of working buffer and 5 μ L of ethanol were incubated for 30 min; then, 120 μ L of the substrate was added and incubated for 90 min.

The irreversibility of enzyme inhibition by compounds 1 and 2

To determine the irreversibility of the α -PsGal activity inhibition by compounds 1 and 2, 15 μ L of 0.625 mM EtOH solution of compounds 1 and 2 were added to 75 μ L of α -PsGal solution in 0.05 M sodium phosphate buffer (pH 7.0). After 30 min of the mixture incubation, the aliquot of 30 μ L was taken and 120 μ L of 3.3 mM *p*NPG solution was added to initiate the reaction; the reaction was stopped by the addition of 150 μ L of 1 M Na₂CO₃. The α -D-galactosidase activity was determined as described above. The remaining reaction mixture was dialyzed against 1 L of 0.02 M sodium phosphate buffer (pH 7.0) for 60 h at 4 °C. GenBAflex-tubes 6–8 kDa (Scienova 108 GmbH, Wildenbruchstabe 15 07745 Jena, Germany) were used for dialysis. The buffer was changed 3 times during the dialysis. The enzyme activity was determined as described above. A sample of α -PsGal treated with H₂O or EtOH in the absence of the inhibitor (75 μ L of the enzyme solution and 15 μ L of H₂O or 6 μ L of EtOH,



respectively) was dialyzed and used as a control for determination of the initial activity. The experiment was carried out in two replicates. The residual activity was calculated as described above.

Determination of antioxidant activity of compounds 1–7

Linoleic acid autoxidation inhibition assay. Antioxidant activity of compounds was determined using described methods⁶⁴ with a slight modification. A straight vial (15 mm diameter; 40 mm height) containing 3.58 mL of 0.1 M phosphate buffer (pH 7.4), 20 μ L of either a 20 mM EtOH solution of a test compound (0.1 mM final concentration) or 20 μ L of EtOH as a control, and 200 μ L of 0.1 M micelle solution of linoleic acid (5.0 mM final concentration) in 0.1 M solution of Tween-40 in 0.1 M phosphate buffer was stirred vigorously and then sonicated in a bath sonicator at 50 $^{\circ}$ C for 3 min to give a clear micelle solution. Oxidation was initiated by the addition of 200 μ L of 20 mM AAPH solution in 0.1 M phosphate buffer (1.0 mM final concentration). The reaction mixture (4 mL total volume) was maintained at a constant temperature (50 $^{\circ}$ C) while continuously shaking. A 400 μ L aliquot was taken at 10 min intervals from the reaction mixture and the absorbance of conjugated diene hydroperoxides at 234 nm was measured for 1 h in a 1 mm pathlength cell. The rate of conjugated diene hydroperoxide formation was calculated as follows: oxidation rate (V , μ M min^{-1}) = $[(A_f - A_0)/(T_f - T_0) \times \epsilon^{-1}b^{-1}] \times 10^6$, where ϵ is the molar extinction coefficient of the conjugated dienes at 234 nm ($\epsilon = 27\,000\text{ M}^{-1}\text{ cm}^{-1}$);⁶⁴ $b = 0.1$ cm, the light path for the cell; A_0 and A_f , the absorbance at zero (T_0) and final time (T_f) of oxidation. Inhibition (%) = $100 - (V_{\text{sample}} \times 100/V_{\text{control}})$.

Ferric reducing antioxidant power (FRAP) assay. The FRAP assay was adapted from previous studies.^{65,66} The reaction mixture (2 mL total volume) consisted of EtOH solution of tested compounds (final concentration 25, 50, 100, 200, and 300 μ M), 100 μ L of 0.1% solution of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ in H_2O , 200 μ L of 0.25% EtOH solution of α, α' -dipyridyl, and 50% aqueous EtOH to 2 mL of the total value. BHT was used as a reference compound. After 60 min, the absorbance at 520 nm was measured using a Shimadzu UV 1240 spectrophotometer. The FRAP values for compounds were calculated according to the following equation: $\text{FRAP} = C_{\text{Fe}}/C_x$, where C_{Fe} is the concentration of Fe^{2+} (μ M) formed in the reaction; C_x is the concentration of tested compounds in the reacting mixture. The concentration of Fe^{2+} (μ M) formed in the reaction was calculated using the calibration curve obtained for different concentrations of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$.

Statistical analysis

Data were analyzed using Origin 8.1 software. All of the experiments were conducted in triplicate, and the data were expressed as the mean values \pm standard deviation.

Results and discussion

Characterization of the spongean OH-PBDEs

For the present study, polybrominated diphenyl ethers 2-(2',4'-dibromophenoxy)-3,5-dibromophenol (**1**, yield 0.03% on the dry

weight of the sponge) and 2-(2',4'-dibromophenoxy)-3,4,5-tribromophenol (**2**, yield 0.15%) were reisolated from repository sample of a *Dysidea* sponge and identified by comparison of their spectroscopic data (^1H -, ^{13}C -NMR, HREI MS) with published values.^{18–21,58} Reisolation of compounds **3–7** was described previously.⁵⁴

α -PsGal inhibitory potency of compounds 1 and 2

Previously, we showed that polybrominated diphenyl ethers **3–7** (Fig. 1) isolated from marine sponges of the genus *Dysidea* are irreversible inhibitors of recombinant α -PsGal from the marine bacterium *Pseudoalteromonas* sp. KMM 701.⁵⁴ During that investigation, the following tendency was established. It was found that in the series of polybrominated diphenyl ethers **3–7**, the presence of an additional hydroxyl group in ether **3** significantly increased its activity. An increase in the number of bromine atoms in monohydroxylated diphenyl ethers from **5** to **7** also increased their activity, which is associated with the increase of the hydrophobicity of the compounds. The presence of the methoxyl group reduced the activity of **4** (Table 1).⁵⁴

The effect of compounds **1** and **2** on the hydrolytic activity of bacterial α -PsGal has not been investigated previously. The standard activity of α -PsGal in the presence of compounds **1** and **2** was determined according to the procedure described earlier.⁵⁴ Plots of dependence of inhibition on the concentration of compounds had a sigmoid form and were approximated by the Hill equation. The theoretical dependences were fitted to the experimental values using the OriginLab 8.1 software. Values of IC_{50} were determined from the coefficients of non-linear regression.⁵⁴ We have repeated the determination of IC_{50} for compound **3** in this work for comparison. It was shown that the IC_{50} value of compound **3** determined in this study was 4.48 μ M, which is similar to the reported IC_{50} value of 4.26 μ M (ref. 54) (ESI, Fig. S1[†]).

Compounds **1–3** are dihydroxylated PBDEs, compounds **5–7** are monohydroxylated PBDEs, and PBDE **4** is a methoxylated analogue of compound **1** (Fig. 1). As can be seen in Table 1, values of IC_{50} of the compounds for inhibition of α -PsGal decrease from compound **1** to **3**, which indicates an increase in their inhibitory activity. Compounds **1**, **2**, and **3** have four, five, and six bromine atoms in the structure, and their hydrophobicity, calculated using the ChemOffice 2008/ChemBio3D Ultra 11.0 program, increases from **1** to **3** (cLogP 6.15, 6.77, 7.67, respectively). These data are consistent with the previously established tendency for monohydroxylated polybrominated diphenyl ethers, and an increase in hydrophobicity leads to an increase in the inhibitory activity of hydroxylated

Table 1 Values of IC_{50} for compounds **1–7** in inhibition of α -D-galactosidase (α -PsGal)

| Compound | 1 ^a | 2 ^a | 3 ^a | 3 ^b | 4 ^b | 5 ^b | 6 ^b | 7 ^b |
|---|-----------------------|-----------------------|-----------------------|-----------------------|-----------------------|-----------------------|-----------------------|-----------------------|
| IC_{50} , μM ^c | 12.88 | 8.78 | 4.48 | 4.26 | 60.5 | 41.8 | 36.0 | 16.0 |

^a The value of IC_{50} was determined in this experiment. ^b Published data.⁵⁴ ^c Average values of three determinations.



polybrominated diphenyl ethers.⁵⁴ Dependence of the activity of OH-PBDEs and the number of bromine atoms in their structures was reported earlier regarding the activity of OH-PBDEs in the inhibition of lipoxygenases²⁶ and binding to thyroid hormone receptors.¹⁰ So, data obtained for dihydroxylated diphenyl ethers **1** and **2** confirm this tendency.

It was shown that polybrominated diphenyl ethers **1** and **2** irreversibly inactivate the α -D-galactosidase. The activity of the enzyme did not recover after dialysis against the buffer solution for 60 h, but the enzyme, in the absence of the inhibitors, kept 100% activity during this dialysis process (ESI, Table S3†).

The effect of PBDEs 1–7 on α -NaGalase

According to the modern classification of carbohydrate active enzymes (CAZy), the α -NaGalase from human cancer cells belongs to the GH27 family, and α -PsGal belongs to the GH36 family. These enzymes of the GH27 and GH36 families are related, and they share a common ancestor gene and form clan D, and their active site structure is highly homologous;⁴⁶ therefore, we decided to test whether spongean PBDEs 1–7 will inhibit cancerous α -NaGalase, which is closely related. We hoped that effective α -PsGal inhibitors would inhibit the highly malignant enzyme α -NaGalase in cancer cells. All PBDEs 1–7 were tested for their influence on the hydrolyzing activity of the α -NaGalase isolated from cells of human intestinal carcinoma DLD-1, HT-29, and HCT-116, breast cancer MDA-MB-231, melanoma SK-MEL-28 and RPMI-7951, and mouse healthy epidermal cells JB6 Cl 41. Our experiments showed that none of the tested compounds 1–7 had any direct inhibitory effects on the activity not only against human cancer cell α -NaGalase as well against α -NaGalase from mouse healthy epidermal cells (ESI, Fig. S2†).

The elevated level of the enzyme α -NaGalase is a well-recognized feature of cancer cells; moreover, the ability to metastasize, impaired programmed cell death, and drug resistance are the most apparent features of cancer cells. Jafari *et al.*⁶⁷ found that after α -NaGalase silencing in cells of human gastric adenocarcinoma, no significant difference was observed for daunorubicin sensitivity, but α -NaGalase downregulation

had augmentative and regressive effects on cell death and migration. Clark *et al.*⁴⁸ showed that the iminosugar DGJNac can inhibit, stabilize, and chaperone human α -NaGalase both *in vitro* and *in vivo*. Ayers *et al.*⁴⁷ showed that galacto-iteamine, unnatural analogue of alkaloid iteamine (*o*-aminobenzyl α -D-glucopyranoside) isolated from *Itea virginica* L. inflorescence, was the weak direct inhibitor of chicken liver α -NaGalase. Probably, compounds 1–7 had no inhibitory effects on cancer α -NaGalase since 1–7 are not *O*-glycosylated.

Antioxidant activity of compounds 1–7

Polybrominated diphenyl ethers 1–7 can potentially be reductive inhibitors of α -PsGal due to their phenolic moieties. We evaluated their antioxidant activity to check if the antioxidant activity influences the activity of α -PsGal. It has been established that under the conditions of oxidation, dihydroxylated diphenyl ethers 1–3 at a concentration of 0.1 mM exhibit weak antioxidant properties, reducing the rate of oxidation of linoleic acid by a maximum of 18.8%. An increase in the concentration of compounds 1–3 to 0.2 mM reduced the oxidation rate by 34.0–34.2% (Table 2). The standard antioxidant BHT (2,6-di-*tert*-butyl-4-methylphenol) at a concentration of 0.1 mM and 0.2 mM reduced the oxidation rate by 88.1% and 91.4%, respectively.

Monohydroxylated diphenyl ethers 4–7 at a concentration of 0.1 mM did not show a noticeable inhibitory effect on the rate of linoleic acid oxidation. At a concentration of 0.2 mM, diphenyl ethers 4–7 showed a very weak antioxidant activity, reducing the oxidation rate by a maximum of 3.4% (Table 2).

In addition to the ability of compounds 1–7 to scavenge peroxy radicals during the oxidation of linoleic acid, the iron-reducing activity of diphenyl ethers 1–7 was tested. The reduction process of Fe(III) ions was followed by an increase in the concentration of Fe(II) ions, which form a colored complex with α , α' -dipyridyl.⁶⁶ The iron-reducing activity of diphenyl ethers 1–3 containing two hydroxyl groups was lower than the activity of BHT by more than 2 times. The iron-reducing activity of the monohydroxylated diphenyl ethers 4–7 was an order of magnitude lower than that of the dihydroxylated diphenyl ethers 1–3 (Table 2).

Table 2 Antioxidant activity of compounds 1–7

| Compound | AAPH-initiated oxidation of linoleic acid | | | | Reduction of Fe(III), $\mu\text{M Fe(II)}$ per μM antioxidant ^a |
|----------------------|---|--------|---------------|--------|---|
| | Rate, $\mu\text{M min}^{-1}$ | | Inhibition, % | | |
| | 0.1 mM | 0.2 mM | 0.1 mM | 0.2 mM | |
| 1 | 7.95 | 6.45 | 18.8 | 34.2 | 0.19 |
| 2 | 8.01 | 6.47 | 18.2 | 34.0 | 0.18 |
| 3 | 7.97 | 6.46 | 18.6 | 34.1 | 0.19 |
| 4 | 9.78 | 9.49 | n/a | 3.2 | 0.03 |
| 5 | 9.80 | 9.48 | n/a | 3.3 | 0.03 |
| 6 | 9.80 | 9.47 | n/a | 3.4 | 0.03 |
| 7 | 9.79 | 9.50 | n/a | 3.1 | 0.03 |
| BHT | 1.17 | 0.84 | 88.1 | 91.4 | 0.45 |
| Control ^b | 9.80 | 9.80 | — | — | — |

^a Average value of three determinations. ^b Oxidation of linoleic acid without antioxidants.



As can be seen from the data shown in Table 2, the number of bromine atoms in the molecules of brominated diphenyl ethers did not affect the rate of oxidation of linoleic acid and the iron-reducing activity. A comparison of compounds 1 and 4 shows that the replacement of one hydroxyl group in ether 1 with a methoxyl group leads to a decrease in the antioxidant activity of ether 4 (Table 2). Thus, the observed antioxidant activity of dihydroxylated diphenyl ethers 1–3 is associated with the presence of an additional hydroxyl group in their structure. As can be seen in Table 1, dihydroxylated PBDEs are more active and their α -PsGal inhibitory activity increases in the following order: 3 > 2 > 1 and the activity of less active monohydroxylated PBDEs increases in the order: 7 > 6 > 5. Although the antioxidant activity of PBDEs is low, the influence of hydroxyl groups is observed; the activity of compounds 3, 2, 1 is higher than that of 7, 6, 5. So, there is a direct correlation between the α -PsGal inhibitory and antioxidant activity of tested PBDEs.

Conclusion

Here, we characterized the properties of OH-PBDEs 1 and 2 as irreversible inhibitors of the bacterial α -D-galactosidase of GH36 family (α -PsGal) from the marine bacterium used as a suitable model. Bacterial α -D-galactosidase of GH36 family of the human intestine may be one of the targets of the harmful effects of OH-PBDEs. The study showed that OH-PBDEs from marine sponges of the genus *Dysidea* inhibit α -PsGal but do not act on cancer α -NaGalase.

Spongean OH-PBDEs at high concentrations exhibit weak antioxidant activity. Comparison of the results of inhibition of α -PsGal activity by compounds 1–7 and antioxidant activity allows us to draw the following conclusions: the presence of a hydroxyl group is a necessary condition for the inhibition of α -PsGal and antioxidant activity; dihydroxylated compounds are more active than monohydroxylated ones; the introduction of a methoxyl group into the molecule of PBDE reduces its inhibitory action on α -PsGal and reduces the antioxidant activity of PBDE; an increase in the degree of bromination increases the inhibitory activity of PBDEs on the enzyme, while the antioxidant activity does not depend on the number of bromine atoms in the molecule. Antioxidant activity is not responsible for the effectiveness of α -PsGal inhibition due to the large difference in the concentrations required for enzyme inhibition (μ M) and antioxidant activity (mM).

Our findings add to the knowledge of OH-PBDEs as environmental toxins and their potential action on human health.

Conflicts of interest

There are no conflicts of interest to declare.

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