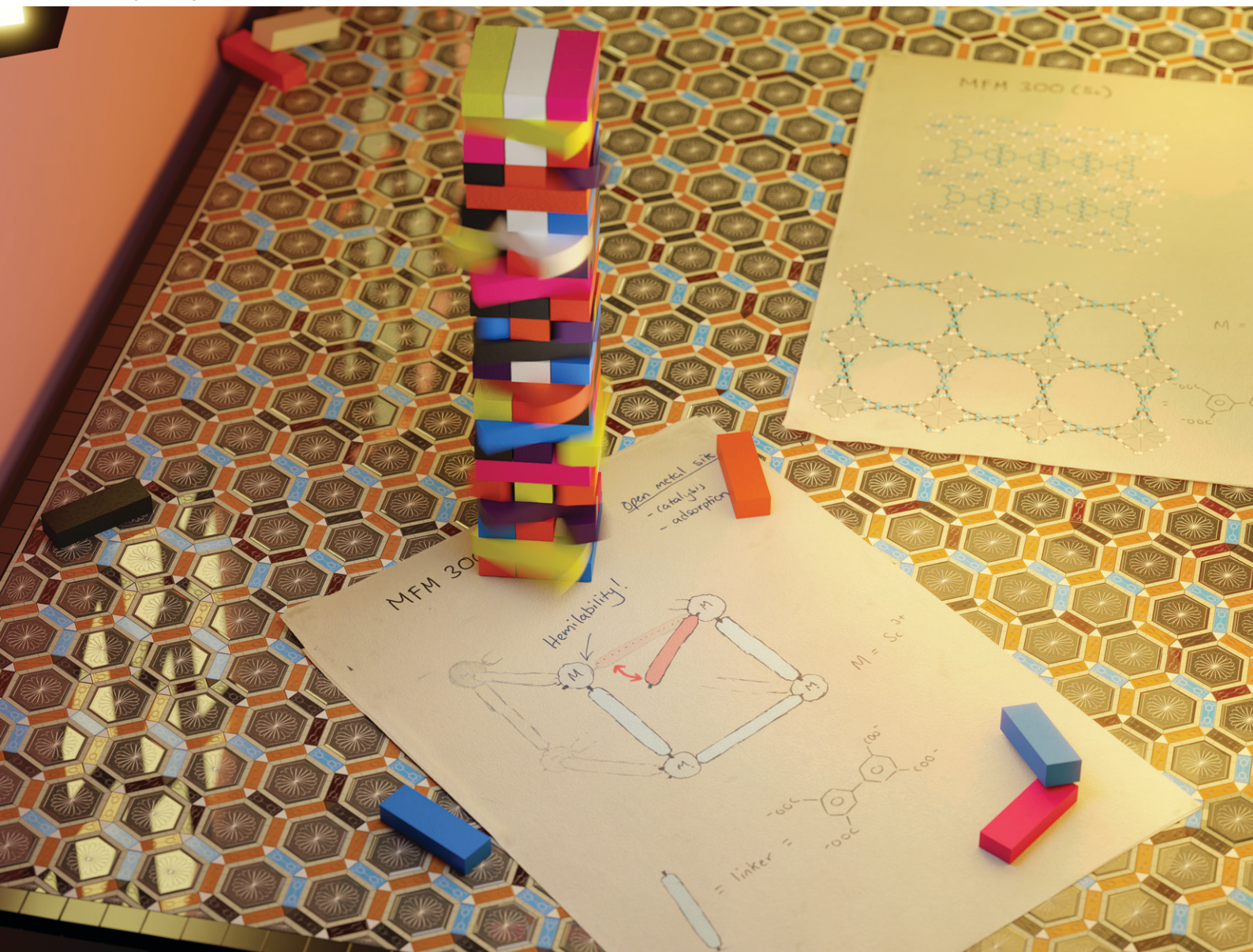


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HIGHLIGHT

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Dynamic metal-ligand bonds in MFM-300: the use of semi-open metal sites to “teach an old dog new tricks”


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Dynamic metal–ligand bonds in MFM-300: the use of semi-open metal sites to “teach an old dog new tricks”

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Developing a deep understanding of the fascinating properties of MFM-300 (MFM = Manchester framework material) opens up a world of new and thrilling applications. The unique dynamic metal–ligand bonding for MFM-300(Sc) and MFM-300(In) is chronologically presented here, highlighting the exacting applications that were associated with such special behaviour (*i.e.*, description of complex NH₃ adsorption mechanisms, conversion of toxic H₂S to polysulfides for promising fabrication of lithium–sulphur batteries and heterogeneous catalysis). This highlight aims to showcase the current advances in semi-open metal sites inherent to MFM-300(Sc) and MFM-300(In) to inspire new investigations to be carried out to develop new applications of this promising family of MOF materials.

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1. Introduction

Metal–organic frameworks (MOFs), or porous coordination polymers (PCPs), are a promising family of hybrid materials constructed by metal ions coordinated to organic ligands.¹ MOFs were first reported back in the 1980s, and since then they have drawn significant attention due to their chemical variability, high crystallinity, extraordinary apparent surface area, and, in some specific cases, high chemical stability.² Also, their integrated design and ordered lattice permit chemical, spatial and dynamic properties to be carefully adjusted in ways not available in other porous materials (*e.g.*, carbon molecular sieves, zeolites and porous metal oxides). Thus, not surprisingly, these appealing properties have postulated MOFs as attractive contenders for applications such as gas storage/separation,³ catalysis,⁴ drug delivery,⁵ environmental air and water remediation,^{6,7} and molecular sensing.⁸ Interestingly, more than 100 000 reported MOF structures constitute a rapidly expanding collection of chemical functionalities originating from their organic ligands, inorganic metal centres and different post-synthetic modifications.^{9,10} Such an appealing set of characteristics enhances the crystal engineering of MOFs as one of the most researched fields in the past decades.

The classic perception of MOFs as static crystalline structures has transited to flexible (*i.e.*, dynamic metal–ligand interactions or hemilabile coordinate bonds), which has been recognised for contributing to unusual physiochemical properties.^{11–13} Although MOFs can typically generate functionality from their building blocks (*i.e.*, functional groups incorporated into the organic ligands or with access to open metal sites), functionality is progressively derived from such dynamic processes (flexibility around the coordination bonds).^{14–16}

One of the most relevant concepts in coordination chemistry is ligand lability, and, for example, labile ligands provide access to the active metal sites in catalysis.^{17,18} Hemilabile ligands are multidentate ligands (*i.e.*, “hard” and “soft” donors) in which one donor is labile and can be reversibly disrupted to generate an open metal site to which substrates bind.^{19,20} In MOF chemistry, labile metal–ligand coordination can be dislocated by different stimuli without compromising the crystal structure of the material (*e.g.*, framework connectivity), in other words, the generation of transient open metal sites.²¹

Similarly, the assembly of crystallinity in extended metal–organic lattices (such as MOFs, metal–organic cages and HOFs) is based on the reversible exchange of coordinate bonds.^{22,23} Thus, such reversible metal–ligand bonding, which can be exemplified by stimuli-responsive, adaptable, dynamic, and self-healing bonding, has been lately presented to elucidate key phenomena in MOFs, from their crystal growth to their phase transitions.²⁴ Typically, the perception of metal–ligand dynamics has been exclusively attributed to temperature increase as the external stimulus. However, recently, it has been

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demonstrated that guest molecules can trigger such reversible bond reorganisation.^{20,25} Such studies on metal-ligand dynamics in MOFs have demonstrated that it is fundamental to recognise phase transitions to discover uncommon adsorption mechanisms.²⁶

This highlight aims to establish the role of dynamic bonding in a particular family of MOF materials (MFM-300, MFM = Manchester framework material), which has been observed so far only for MFM-300(Sc) and MFM-300(In), by presenting the principal contributions related to such special behaviour and its associated exciting applications while providing relevant perspectives on the future of MOF chemistry in this field, particularly with the aim to encourage the scientific community to investigate the other members of the MFM-300 family (*i.e.*, MFM-300(Al), MFM-300(Cr), MFM-300(V), MFM-300(Ga) and MFM-300(Fe)).

2. Structural description

The MFM-300 family is synthesised using a tetracarboxylic ligand H₄BPTC (biphenyl-3,3',5,5'-tetracarboxylic acid) and different trivalent metal centres (Sc(III),²⁷ Al(III),²⁸ In(III),²⁹ Cr(III),³⁰ V(III),³¹ Ga(III)³² and Fe(III)³³) under different reaction conditions; Schröder and Yang have elegantly reported this family. The constituents (BPTC⁴⁻ and metal ions) of the MFM-300 family are arranged such that every ligand (BPTC⁴⁻) is coordinated to eight metal centres *via* four carboxylate groups (Fig. 1a), affording an infinite chain of [MO₄(OH)₂] which constituted the secondary building unit (SBU) (Fig. 1b). The metal centres display an octahedral geometry associated by two mutually *cis*-μ₂-OH (hydroxo groups) anions

that comprise the helical SBUs which are connected by the BPTC ligands to form a one-dimensional cylindrical channel (Fig. 1c).

3. The first reversible guest-induced evidence: NH₃ adsorption

The first reported material of the MFM-300 family was described by Champness and Schröder in 2011 (MFM-300(Sc), also known as NOTT-400), and H₂ was the first molecule (not considering N₂ at 77 K to estimate the surface area of this material) to be investigated inside the pores of MFM-300(Sc).²⁷ From that point, adsorption of many molecules was investigated (CO₂, CH₄, H₂O, I₂, C₆H₆ and SO₂), showing interesting properties such as the high chemical stability of this Sc(III)-based material, high CO₂ adsorption and CO₂/CH₄ selectivity, facile adsorption-desorption, H₂O adsorption for promising heat-pump applications, very high I₂ uptake for relevant radioactive decontamination, benzene stabilization and effective SO₂ capture with remarkable CO₂/SO₂ selectivity.³³⁻³⁷ However, none of these molecules provided any evidence of dynamic bonding behaviour.

Then, in 2021, Dincă and Maurin investigated an unusual ammonia adsorption behaviour for MFM-300(Sc).²⁶ First, upon the adsorption of NH₃ (at 298 K and 1 bar) on MFM-300(Sc) a very unusual isotherm was shown (see Fig. 2), which was not correlated to any of the isotherms previously collected for other gases under the same conditions (*e.g.*, H₂, CH₄ and SO₂).^{27,35,37} Moreover, this particular adsorption/desorption behaviour was preserved over multiple cycles (corroborating the chemical stability of the material and

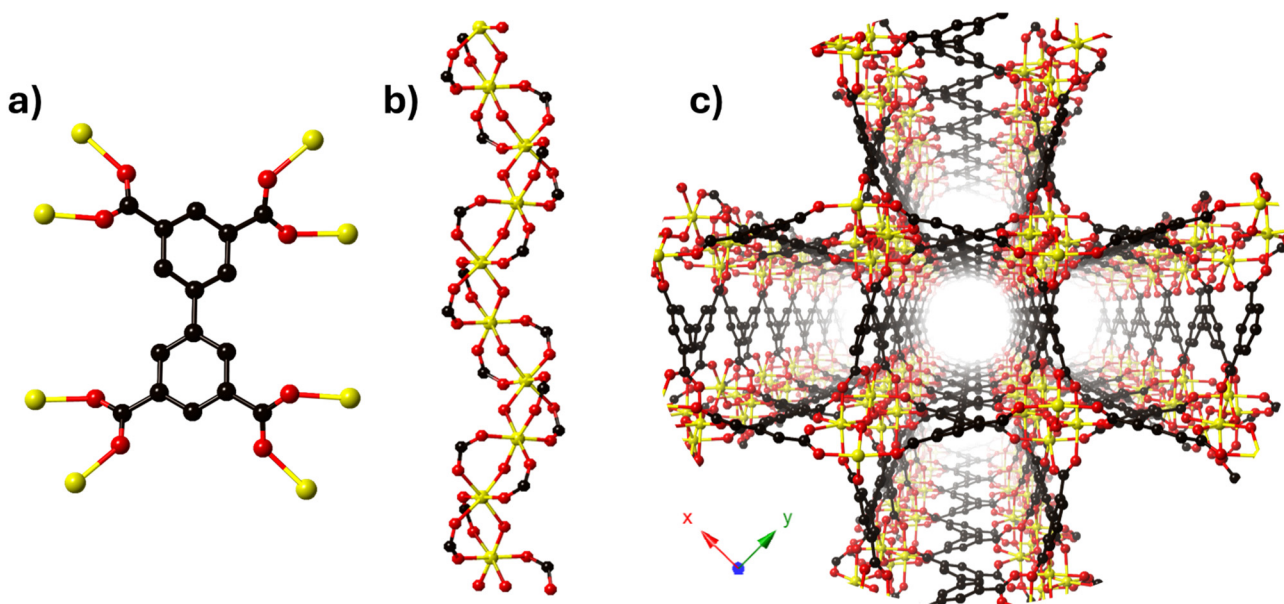


Fig. 1 MFM-300(M) (M = Sc(III), Al(III), In(III), Cr(III), V(III), Ga(III) and Fe(III)) ligand-metal framework structure constitution. (a) Coordination arrangement of BPTC⁴⁻ ligand (C atoms in grey, O atoms in red, hexacoordinated metal ions in yellow (Sc, In)); (b) M(OH)(CO₂)₂, 1D chiral chain; (c) MFM-300(Sc, In) cylindrical channels viewed along the c axis.



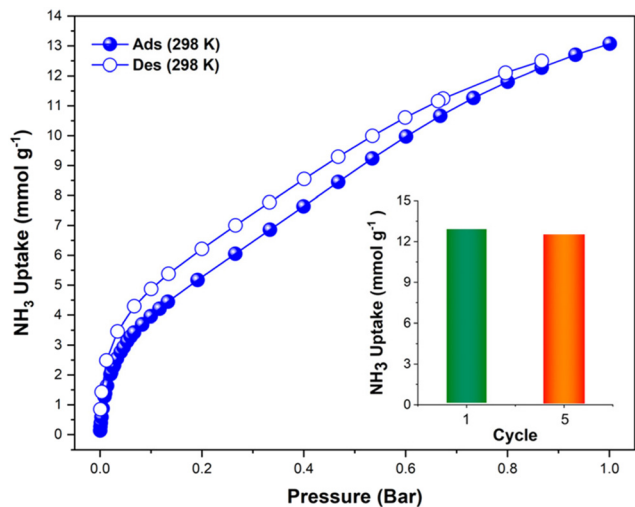


Fig. 2 Experimental NH_3 adsorption–desorption isotherms of MFM-300(Sc) (filled blue circles = adsorption; open blue circles = desorption) at 298 K and up to 1 bar. Inset: comparison of NH_3 uptake of MFM-300(Sc) from the first and fifth adsorption cycles (13.1 and 12.6 mmol g^{-1} , respectively, 3.8% decrease) at 298 K and up to 1 bar. Reproduced with permission from ref. 26. Copyright 2021, American Chemical Society.

confirmed by PXRD), which strongly suggested a fully reversible capture/release mechanism. In particular, the presence of an abnormal low-pressure step change (see Fig. 2) with an uptake of approximately 1 NH_3 molecule per $\mu_2\text{-OH}$ group provided the clue to a possible “partial” coordination of the NH_3 molecule to the Sc(III) metal centre. Thus, force field-based Monte Carlo simulations, periodic DFT calculations and *ab initio* molecular dynamics (AIMD) simulations demonstrated a Sc- NH_3 coordinative mode that remained stable over time, corresponding to an energetically favourable configuration (Fig. 3). Interestingly, AIMD calculations also showed that once NH_3 is disconnected from the metal centre (Sc(III)), the MOF structure self-heals, and the initial crystal structure configuration is restored. Thus, such NH_3 -triggered dynamic metal–ligand bonding is demonstrated to be fully reversible, explaining the characteristic experimentally observed adsorption/desorption behaviour for MFM-300(Sc).

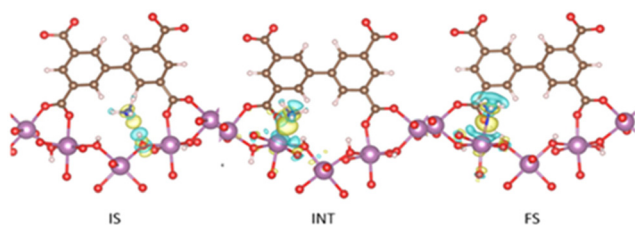


Fig. 3 DFT-derived charge differences for the adsorption of NH_3 in IS, INT, and FS with an isovalue of $\pm 0.002 e \text{ \AA}^{-3}$; yellow and cyan colours represent the charge accumulation and depletion regions, respectively. Colour code: carbon (brown), oxygen (red), nitrogen (blue), hydrogen (white), and scandium (purple). Reproduced with permission from ref. 26. Copyright 2021, American Chemical Society.

Finally, *in situ* DRIFT spectroscopy experiments were performed at 298 K, (by increasing the amount of pressure of NH_3 , see Fig. 4), which corroborated the computational predicted NH_3 adsorption mechanism. We evidenced that when starting with a sample loaded with 0.24 bar NH_3 , upon applying vacuum (1.7×10^{-3} Torr) it is possible to fully recover the IR spectrum collected for the pristine sample (see Fig. 4), confirming the reversibility of the NH_3 adsorption process. Although the NH_3 adsorption on MFM-300(Sc) was the first concrete evidence of metal–ligand bonding (or semi-open metal sites), this interesting phenomenon was almost caught some time ago by the adsorption of H_2S (*vide infra*).

4. H_2S adsorption: the unexpected formation of polysulfides

The H_2S capture on MFM-300(Sc) was investigated years before the capture of NH_3 , and finally, in 2020, these unusual results were published by Gutierrez-Alejandre and Maurin.³⁸ Thus, by a set of H_2S breakthrough experiments, it was consistently found that the total H_2S uptake in MFM-300(Sc) was approximately 16.55 mmol g^{-1} (Fig. 5).³⁸ Interestingly, when trying to reuse MFM-300(Sc) after the first adsorption of H_2S by reactivating the material (at high temperatures and under vacuum), it was not possible to get back to the original uptake; instead, such total uptake went down to approximately 10 mmol g^{-1} , even for five consecutive cycles (see Fig. 5). Although PXRD experiments exhibited the retention of the crystalline structure of MFM-300(Sc), after the exposure to H_2S , N_2 adsorption experiments at 77 K showed a reduction of the pore volume of approximately

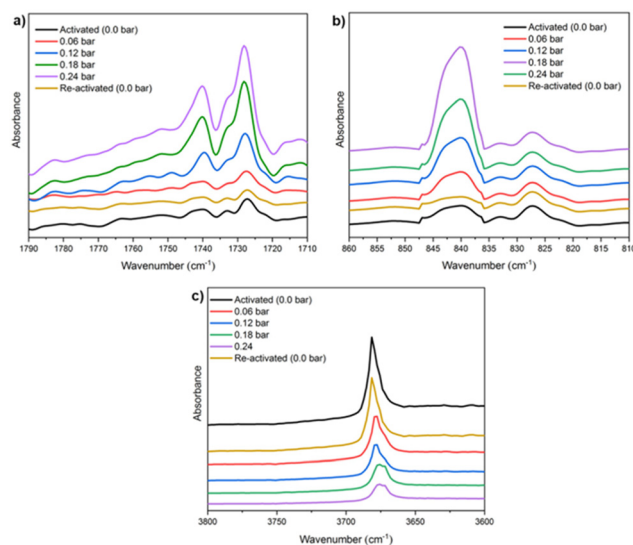


Fig. 4 DRIFT spectra of NH_3 adsorbed at different pressures (from 0 to 0.24 bar) over activated MFM-300(Sc) at 298 K, split into three wavenumber regions: (a) 1790–1710 cm^{-1} ; (b) 860–810 cm^{-1} ; and (c) 3800–3600 cm^{-1} . Reproduced with permission from ref. 26. Copyright 2021, American Chemical Society.



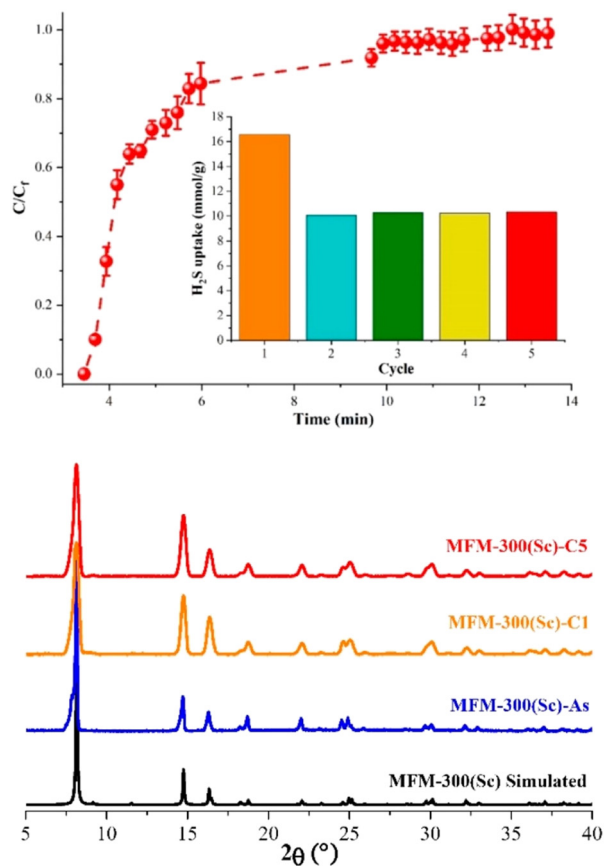


Fig. 5 (Top) Breakthrough curves of H₂S capture on MFM-300(Sc) at 25 °C and 1 bar with a H₂S feed concentration of 10 vol%. The total H₂S/N₂ flow rate was 25 cm³ min⁻¹. The inset shows the proportional H₂S adsorption capacities for each cycle. (Bottom) PXRD patterns for MFM-300(Sc): simulated (black), as-synthesized (blue), after the first H₂S adsorption-desorption cycle (cycle 1, MFM-300(Sc)-C1, orange), and after 5 H₂S adsorption-desorption experiments (cycle 5, MFM-300(Sc)-C5, red). Reproduced with permission from ref. 38. Copyright 2020, American Chemical Society.

34%, which suggested a partial blockage of the pores by the remaining molecules, possibly associated with a chemical reaction inside the pores of the material (*i.e.*, chemisorption). Interestingly, PXRD experiments showed the retention of the crystalline structure (Fig. 5) after every H₂S adsorption cycle, discarding a chemical reaction with MFM-300(Sc), which can partially or totally decompose the structure of the material.

Thus, puzzled by the unusual performance of MFM-300(Sc) upon H₂S adsorption, cyclic H₂S adsorption experiments were carried out on the isostructural material MFM-300(In), showing a total H₂S uptake of 9.10 mmol g⁻¹ for the first cycle. When performing a second cycle (after reactivating this material after the first capture of H₂S), the total uptake dramatically decreased to almost 0.2 mmol g⁻¹, and upon estimating the pore volume, a drastic reduction to 0.02 cm³ g⁻¹ was observed. However, PXRD demonstrated the retention of the crystallinity for MFM-300(In), similarly to MFM-300(Sc). Thus, for both materials, it was possible to suggest an unexpected chemisorption process, which

presumably produced unknown sulphur chemical species inside the pores without decomposing the crystal structure of MFM-300(Sc) and MFM(In). Raman spectroscopy on both “sulphur” samples was employed to uncover the nature of these chemisorbed sulphur species. Thus, the characteristic signals observed in the range 350 to 520 cm⁻¹ are associated with the S-S stretching vibrational modes of different chain length ($n = 8, 6, 4, 2, 1$) polysulfides. Finally, the key approach to distinguish such polysulfides was based on the implementation of electrochemical experiments, which demonstrated that low-order polysulfides ($n = 2$) were produced in MFM-300(Sc), while high-order polysulfides ($n = 6$ and 8) are formed in MFM-300(In). Open-circuit potential (OCP) experiments exhibited a relatively small difference in the potential of the materials (0.36 V): MFM-300(Sc) (2.29 V) and MFM-300(In) (2.65 V) (see Fig. 6), which is responsible for the formation of different polysulfides.

At that time, in 2020, the proposed catalytic mechanism for the formation of polysulfides was not fully understood, and it was postulated based on the effect of the pore

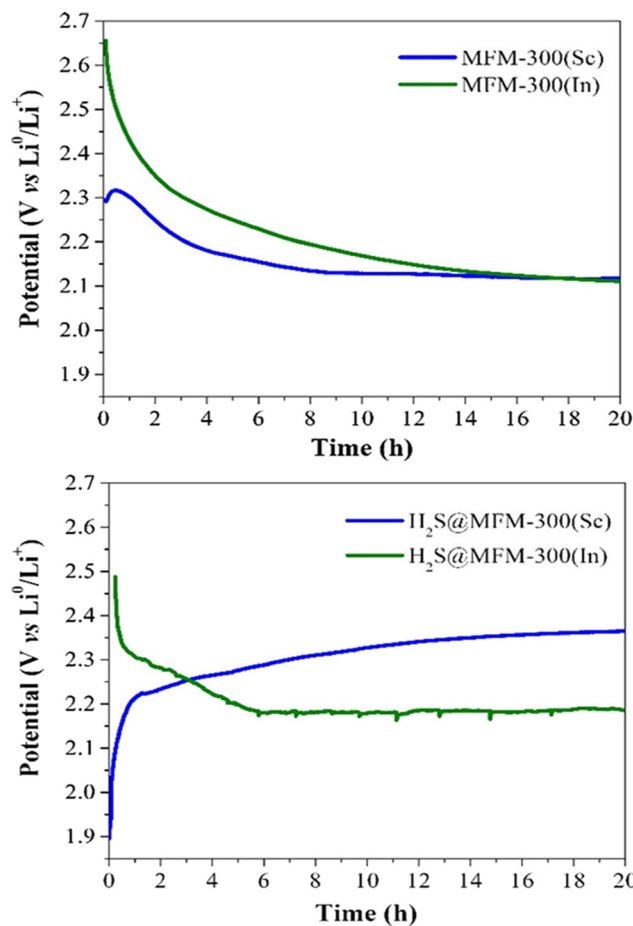


Fig. 6 Open-circuit potential (OCP) monitoring of MFM-300(M) as active materials in cathodes for lithium-ion batteries to describe the cell stabilisation processes. (Top) Pristine materials MFM-300(Sc) and MFM-300(In); (bottom) H₂S saturated materials H₂S@MFM300(Sc) and H₂S@MFM-300(In). Reproduced with permission from ref. 38. Copyright 2020, American Chemical Society.



Highlight

confinement within MFM-300, where a H_2S molecule can spontaneously react with another H_2S molecule to form polysulfides.³⁸ However, thanks to the understanding related to the NH_3 -triggered dynamic metal–ligand bonding, in 2021,²⁶ a different catalytic mechanism was proposed by Gutiérrez-Alejandre.³⁹ Thus, the generation of polysulfides is feasible due to the direct interaction of Sc(III) metal centres (semi-open metal sites) with H_2S to establish a “classic catalytic reaction” (see Fig. 7), concluding that the inclusion of H_2S inside the pores of MFM-300(Sc) and MFM-300(In) triggers the dynamic metal–ligand bonding. This fascinating and unexpected catalytic discovery provided the incommensurable curiosity to explore a “classic catalytic reaction” (Strecker hydrocyanation reaction) by MFM-300 (*vide infra*).

5. Heterogeneous catalysis: a highly desired application for MOFs

MOFs have been envisioned as excellent candidates for heterogeneous catalysis due to their structural tunability, vast diversity in the incorporation of functional groups, high internal surface areas, access to active metal sites, and highly crystalline structures.^{40–42} Typically, MOFs show access to active metals by the removal of labile terminal ligands or molecules (*e.g.*, H_2O , DMF and EtOH), affording what is known as uncoordinated metal centres or open metal sites (OMS).^{43,44} By this approach, countless examples of MOFs with OMS for remarkable catalytic applications have been reported in the literature.^{45–48}

However, the introduction of the new concept in MOFs related to dynamic metal–ligand bonding or semi-open metal sites was only recently reported in 2022 by Cheon-Jeong and

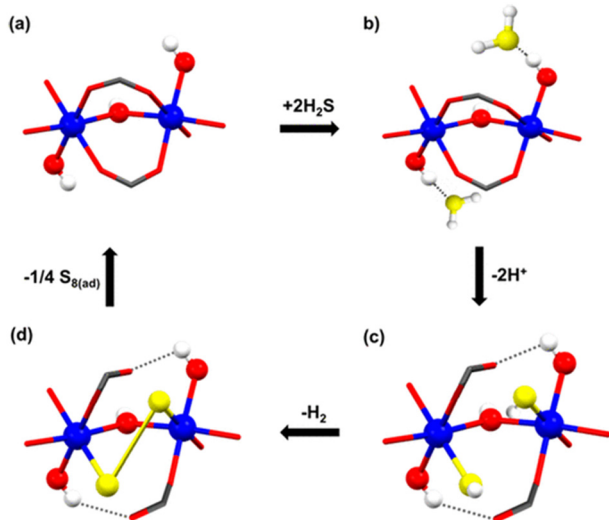


Fig. 7 Schematic representation of the proposed catalytic steps (sequence) for the transformation of H_2S to polysulfides in MFM-300(Sc). Scandium in blue, oxygen in red, hydrogen in white, and sulphur in yellow. Reproduced with permission from ref. 39. Copyright 2021, American Chemical Society.

Maurin, where MFM-300(Sc) was investigated as a catalyst for the well-known Strecker hydrocyanation reaction.²⁰ Typically, the Strecker reaction is performed under homogeneous conditions, proceeding by the nucleophilic addition of cyanide ions to imines using different Lewis acid catalysts, such as Sc(III) metal centres (as dissolved cations).⁴⁹ Thus, MFM-300(Sc) was employed to heterogeneously catalyse the formation of 2-phenyl-2-(phenylamino)acetonitrile from the reaction between *N*-benzylideneaniline and TMSCN ,²⁰ by taking advantage of the metal–hemilabile ligand bond dynamics in MFM-300(Sc), as was demonstrated for the conversion of H_2S to polysulfides.³⁹ Continuous ^1H NMR monitoring showed that MFM-300(Sc) was an effective catalyst for this Strecker reaction with a conversion of approximately 45% (see Fig. 8). Five catalytic cycles showed a consistent catalytic performance (see Fig. 8), and PXRD experiments demonstrated the retention of the crystalline structure after the catalytic cycling (Fig. 8).

The molecule responsible for the dynamic metal–ligand bonding to have access to the Sc(III) metal centres (semi-open metal sites) is HCN , which was corroborated by DFT calculations, finding a similar mechanism observed for the NH_3 adsorption.²⁰ The proposed Strecker mechanism (see Fig. 8) begins with the formation of a bond between the $\text{C}(\text{HCN})$ and the Sc(III) semi-open metal site, followed by a proton transfer from the coordinated HCN to the $\text{N}(\text{imine})$, and the CN^- group remained coordinated to the unsaturated Sc(III) . Then, the second step continues with a change of the CN^- to NC^- , which is perfectly fitting for a nucleophilic attack on the intermediate imine species, affording the formation of the cyanylated product.

Later, in 2023, Cheon-Jeong and Maurin evaluated this same reaction with an isostructural material to MFM-300(Sc) based on In(III) : MFM-300(In).⁵⁰ This material also showed access to semi-open metal sites, which was examined by variable temperature resonance Raman spectroscopy (VT-

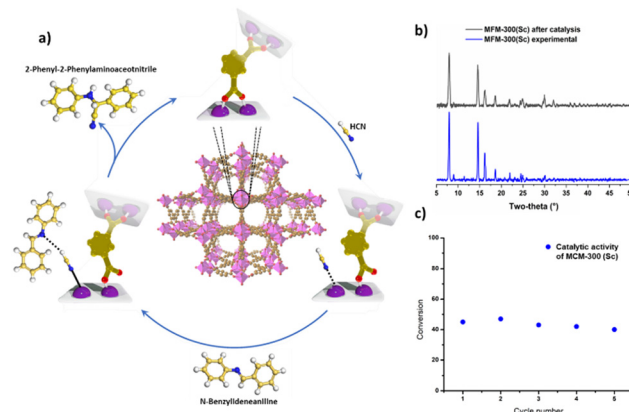


Fig. 8 (a) Illustration of the catalytic mechanism for the production of 2-phenyl-2-(phenylamino)acetonitrile driven by the metal–hemilabile linker dynamics in MFM-300(Sc) (C, gold; H, white; N, blue; O, red; Sc, pink). (b) PXRD data for pre- and post-catalysis of MFM-300(Sc). (c) Catalytic activity of MFM-300(Sc) for 5 cycles, in which significant loss of activity was not detected. Reproduced with permission from ref. 20.



RRS) and variable temperature diffuse reflectance infrared Fourier-transform spectroscopy (VT-DRIFTS), showing that at higher temperatures, more available active In(III) semi-open metal sites are formed (see Fig. 9). Thus, the Strecker reaction performed with MFM-300(In) afforded a 41% conversion to 2-phenyl-2-(phenylamino)acetonitrile, which was similar to that obtained with MFM-300(Sc). Remarkably, since active In(III) sites can be obtained by simply increasing the temperature, the same Strecker reaction was carried out at a higher temperature (323 K), achieving a higher conversion of 93%.

Finally, MFM-300(Sc) was investigated for the acetone hydrogenation reaction, finding that the semi-open metal sites were active for this gas-phase reaction, affording the formation of methyl isobutyl ketone.⁵¹ This reaction was successfully performed at 453 K for 4 hours, yielding a conversion of approximately 28%. After two additional catalytic cycles, the catalytic performance was maintained, and PXRD experiments corroborated the retention of the crystalline structure of MFM-300(Sc) with no significant changes in the BET surface. The proposed catalytic reaction mechanism involves all three steps that occur at the Sc(III) semi-open metal sites (Fig. 10).

These catalytic results for MFM-300(Sc) and MFM-300(In) demonstrate that due to the access to semi-open metal sites (*i.e.*, Sc(III) and In(III)), these chemically stable materials can effectively work as a heterogeneous catalyst in either aqueous or gaseous medium, highlighting the dynamic metal–ligand bonding as the key role in high performances for novel heterogeneous catalysts with remarkable regeneration properties.

6. Conclusions

This highlight provides a chronological account of the experimental and computational observations that revealed

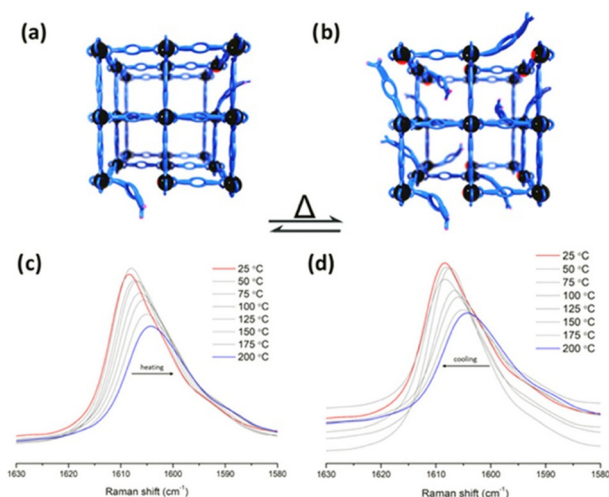


Fig. 9 Representation of transient open metal sites generated by hemilabile metal–ligand bonds and poised for catalysis at (a) lower temperatures and (b) higher temperatures. VT-RRS studies of InOF-1 (c) from 200 to 25 °C and (d) from 25 to 200 °C. Reproduced with permission from ref. 50. Copyright 2023, American Chemical Society.

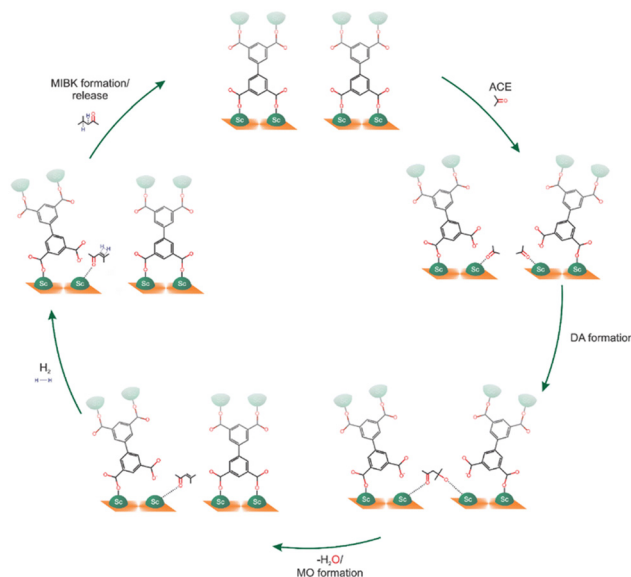


Fig. 10 Scheme showing the proposed acetone hydrogenation mechanism in MFM-300(Sc). Reproduced with permission from ref. 51. Copyright 2023, the Royal Society of Chemistry.

the unique dynamic metal–ligand bonding for MFM-300(Sc) and MFM-300(In). Remarkably, such dynamic binding, in the case of MFM-300(Sc), provided access to temporary open metal sites (semi-open metal sites), opening new portals for exciting and unexpected chemistry, such as the description of complex adsorption mechanisms (*i.e.*, NH_3 adsorption), paving the way toward advance developments of MOFs for the capture of highly corrosive molecules.

The unexpected *in situ* chemical transformation of H_2S inside the pores of MFM-300(Sc) and MFM-300(In) to generate polysulfides not only is an exciting discovery but also can signify the permanent sequestration and conversion of toxic H_2S for new environmental remediation technologies against air pollution. In addition, the homogeneous insertion of polysulfides (the homogeneous sulphur distribution affords higher electrochemical performance) within MOFs can provide new energy platforms for the development of novel MOF–sulphur battery electrodes, as recently demonstrated by Marquez and Mullins, who showed how SU-101 (a Bi(III)-based MOF with “classic” open metal sites) transformed H_2S into polysulfides to fabricate lithium–sulphur batteries.⁵² Thus, exploring MFM-300 for the investigation of how toxic H_2S emissions can be directly converted and applied in lithium–sulphur batteries, considering that the different metals that constitute the MFM-300 family can provide different polysulfides, opens exciting options for different capacities to power novel lithium–sulphur batteries.

MFM-300(Sc) and MFM-300(In) demonstrated promising results in developing novel “intelligent” catalysts, taking advantage of how molecule inclusion can trigger dynamic metal–ligand bonding (semi-open metal sites). Thus, understanding that selected examples of MOFs (a great and



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fascinating task to investigate) are dynamic rather than static environments can dramatically expand the reservoir of candidate MOF materials which can be contemplated for classic and new heterogeneous catalytic applications.

Finally, the fundamental understanding of MOFs remains developing. In particular, the discovery of dynamic metal–ligand bonding in a wide range of MOF materials (starting with the MFM-300 family and a remarkable example recently reported by Farha and co-workers where a Zr(IV)-based MOF demonstrated unexpected SO₂-induced modulator-node dynamics⁵³) turns out to be crucial for discovering new horizons with exciting applications.

Data availability

All data are available in the main text.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- B. F. Hoskins and R. Robson, *J. Am. Chem. Soc.*, 1989, **111**, 5962–5964.
- H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 1230444.
- S. Liu, R. Dupuis, D. Fan, S. Benzaria, M. Bonneau, P. Bhatt, M. Eddaoudi and G. Maurin, *Chem. Sci.*, 2024, **15**, 5294–5302.
- R. A. Peralta, M. T. Huxley, J. D. Evans, T. Fallon, H. Cao, M. He, X. S. Zhao, S. Agnoli, C. J. Sumby and C. J. Doonan, *J. Am. Chem. Soc.*, 2020, **142**, 13533–13543.
- M. Moharramnejad, A. Ehsani, M. Shahi, S. Gharanli, H. Saremi, R. E. Malekshah, Z. S. Basmenj, S. Salmani and M. Mohammadi, *J. Drug Delivery Sci. Technol.*, 2023, **81**, 104285.
- J. L. Obeso, D. R. Amaro, C. V. Flores, A. Gutiérrez-Alejandre, R. A. Peralta, C. Leyva and I. A. Ibarra, *Coord. Chem. Rev.*, 2023, **485**, 215135.
- E. Martínez-Ahumada, M. L. Díaz-Ramírez, M. d. J. Velásquez-Hernández, V. Jancik and I. A. Ibarra, *Chem. Sci.*, 2021, **12**, 6772–6799.
- T. K. Pal, *Mater. Chem. Front.*, 2023, **7**, 405–441.
- J. L. Obeso, M. T. Huxley, J. A. de los Reyes, S. M. Humphrey, I. A. Ibarra and R. A. Peralta, *Angew. Chem., Int. Ed.*, 2023, **62**, e202309025.
- J. D. Evans, C. J. Sumby and C. J. Doonan, *Chem. Soc. Rev.*, 2014, **43**, 5933–5951.
- A. B. Andreeva, K. N. Le, L. Chen, M. E. Kellman, C. H. Hendon and C. K. Brozek, *J. Am. Chem. Soc.*, 2020, **142**, 19291–19299.
- A. B. Andreeva, K. N. Le, K. Kadota, S. Horike, C. H. Hendon and C. K. Brozek, *Chem. Mater.*, 2021, **33**, 8534–8545.
- A. Halder and D. Ghoshal, *CrystEngComm*, 2018, **20**, 1322–1345.
- P. K. Allan, K. W. Chapman, P. J. Chupas, J. A. Hriljac, C. L. Renouf, T. C. A. Lucas and R. E. Morris, *Chem. Sci.*, 2012, **3**, 2559–2564.
- M. Wahiduzzaman, D. Lenzen, G. Maurin, N. Stock and M. T. Wharmby, *Eur. J. Inorg. Chem.*, 2018, **2018**, 3626–3632.
- D. Fairen-Jimenez, S. A. Moggach, M. T. Wharmby, P. A. Wright, S. Parsons and T. Düren, *J. Am. Chem. Soc.*, 2011, **133**, 8900–8902.
- A. J. Martínez-Martínez, C. G. Royle, S. K. Furfari, K. Suriye and A. S. Weller, *ACS Catal.*, 2020, **10**, 1984–1992.
- S. D. Pike, A. L. Thompson, A. G. Algarra, D. C. Apperley, S. A. Macgregor and A. S. Weller, *Science*, 2012, **337**, 1648–1651.
- K. Leus, I. Muylaert, M. Vandichel, G. B. Marin, M. Waroquier, V. Van Speybroeck and P. Van Der Voort, *Chem. Commun.*, 2010, **46**, 5085–5087.
- R. A. Peralta, P. Lyu, A. Lopez-Olvera, J. L. Obeso, C. Leyva, N. C. Jeong, I. A. Ibarra and G. Maurin, *Angew. Chem., Int. Ed.*, 2022, **61**, e202210857.
- C.-C. Cao, C.-X. Chen, Z.-W. Wei, Q.-F. Qiu, N.-X. Zhu, Y.-Y. Xiong, J.-J. Jiang, D. Wang and C.-Y. Su, *J. Am. Chem. Soc.*, 2019, **141**, 2589–2593.
- E. Sánchez-González, M. Y. Tsang, J. Troyano, G. A. Craig and S. Furukawa, *Chem. Soc. Rev.*, 2022, **51**, 4876–4889.
- B. Wang, R.-B. Lin, Z. Zhang, S. Xiang and B. Chen, *J. Am. Chem. Soc.*, 2020, **142**, 14399–14416.
- A. Pórolniczak and A. Katrusiak, *Mater. Adv.*, 2021, **2**, 4677–4684.
- J. L. Obeso, M. T. Huxley, C. Leyva, J. Gabriel Flores, N. Martín-Guaregua, M. Viniegra, J. Aguilar-Pliego, J. A. de los Reyes, I. A. Ibarra and R. A. Peralta, *Coord. Chem. Rev.*, 2023, **496**, 215403.
- P. Lyu, A. M. Wright, A. López-Olvera, P. G. M. Mileo, J. A. Zárate, E. Martínez-Ahumada, V. Martis, D. R. Williams, M. Dincă, I. A. Ibarra and G. Maurin, *Chem. Mater.*, 2021, **33**, 6186–6192.
- I. A. Ibarra, S. Yang, X. Lin, A. J. Blake, P. J. Rizkallah, H. Nowell, D. R. Allan, N. R. Champness, P. Hubberstey and M. Schröder, *Chem. Commun.*, 2011, **47**, 8304–8306.
- X. Han, H. G. W. Godfrey, L. Briggs, A. J. Davies, Y. Cheng, L. L. Daemen, A. M. Sheveleva, F. Tuna, E. J. L. McInnes, J. Sun, C. Drathen, M. W. George, A. J. Ramirez-Cuesta, K. M. Thomas, S. Yang and M. Schröder, *Nat. Mater.*, 2018, **17**, 691–696.
- J. Qian, F. Jiang, D. Yuan, M. Wu, S. Zhang, L. Zhang and M. Hong, *Chem. Commun.*, 2012, **48**, 9696–9698.
- L. Briggs, R. Newby, X. Han, C. G. Morris, M. Savage, C. P. Krap, T. L. Easun, M. D. Frogley, G. Cinque, C. A. Murray, C. C. Tang, J. Sun, S. Yang and M. Schröder, *J. Mater. Chem. A*, 2021, **9**, 7190–7197.
- Z. Lu, H. G. W. Godfrey, I. da Silva, Y. Cheng, M. Savage, F. Tuna, E. J. L. McInnes, S. J. Teat, K. J. Gagnon, M. D. Frogley, P. Manuel, S. Rudić, A. J. Ramirez-Cuesta, T. L.



- Easun, S. Yang and M. Schröder, *Nat. Commun.*, 2017, **8**, 14212.
- 32 C. P. Krap, R. Newby, A. Dhakshinamoorthy, H. García, I. Cebula, T. L. Easun, M. Savage, J. E. Eyley, S. Gao, A. J. Blake, W. Lewis, P. H. Beton, M. R. Warren, D. R. Allan, M. D. Frogley, C. C. Tang, G. Cinque, S. Yang and M. Schröder, *Inorg. Chem.*, 2016, **55**, 1076–1088.
- 33 X. Zhang, I. da Silva, H. G. W. Godfrey, S. K. Callear, S. A. Sapchenko, Y. Cheng, I. Vitórica-Yrezábal, M. D. Frogley, G. Cinque, C. C. Tang, C. Giacobbe, C. Dejoie, S. Rudić, A. J. Ramirez-Cuesta, M. A. Denecke, S. Yang and M. Schröder, *J. Am. Chem. Soc.*, 2017, **139**, 16289–16296.
- 34 M. Rivera-Almazo, E. Perez-Sanchez, E. Martínez-Ahumada, A. Martínez, J. Garza, I. A. Ibarra and R. Vargas, *J. Phys. Chem. C*, 2022, **126**, 6465–6471.
- 35 M. R. Gonzalez, J. H. González-Estefan, H. A. Lara-García, P. Sánchez-Camacho, E. I. Basaldella, H. Pfeiffer and I. A. Ibarra, *New J. Chem.*, 2015, **39**, 2400–2403.
- 36 I. A. Ibarra, A. Mace, S. Yang, J. Sun, S. Lee, J.-S. Chang, A. Laaksonen, M. Schröder and X. Zou, *Inorg. Chem.*, 2016, **55**, 7219–7228.
- 37 J. A. Zárate, E. Sánchez-González, D. R. Williams, E. González-Zamora, V. Martis, A. Martínez, J. Balmaseda, G. Maurin and I. A. Ibarra, *J. Mater. Chem. A*, 2019, **7**, 15580–15584.
- 38 J. G. Flores, J. A. Zarate-Colin, E. Sanchez-Gonzalez, J. R. Valenzuela, A. Gutierrez-Alejandre, J. Ramirez, V. Jancik, J. Aguilar-Pliego, M. C. Zorrilla, H. A. Lara-Garcia, E. Gonzalez-Zamora, G. Guzman-Gonzalez, I. Gonzalez, G. Maurin and I. A. Ibarra, *ACS Appl. Mater. Interfaces*, 2020, **12**, 18885–18892.
- 39 A. López-Olvera, J. G. Flores, J. Aguilar-Pliego, C. K. Brozek, A. Gutiérrez-Alejandre and I. A. Ibarra, *Chem. Mater.*, 2021, **33**, 6269–6276.
- 40 J. L. Obeso, J. G. Flores, C. V. Flores, M. T. Huxley, J. A. de los Reyes, R. A. Peralta, I. A. Ibarra and C. Leyva, *Chem. Commun.*, 2023, **59**, 10226–10242.
- 41 A. Dhakshinamoorthy, A. M. Asiri and H. García, *Trends Chem.*, 2020, **2**, 454–466.
- 42 A. Bavykina, N. Kolobov, I. S. Khan, J. A. Bau, A. Ramirez and J. Gascon, *Chem. Rev.*, 2020, **120**, 8468–8535.
- 43 Ü. Kökçam-Demir, A. Goldman, L. Esrafilı, M. Gharib, A. Morsali, O. Weingart and C. Janiak, *Chem. Soc. Rev.*, 2020, **49**, 2751–2798.
- 44 J. Bae, E. J. Lee and N. C. Jeong, *Chem. Commun.*, 2018, **54**, 6458–6471.
- 45 J. L. Obeso, J. G. Flores, C. V. Flores, V. B. López-Cervantes, V. Martínez-Jiménez, J. A. de los Reyes, E. Lima, D. Solís-Ibarra, I. A. Ibarra, C. Leyva and R. A. Peralta, *Dalton Trans.*, 2023, **52**, 12490–12495.
- 46 J. G. Flores, J. L. Obeso, V. Martínez-Jiménez, N. Martín-Guaregua, A. Islas-Jácome, E. González-Zamora, H. Serrano-Espejel, B. Mondragón-Rodríguez, C. Leyva, D. A. Solís-Casados, I. A. Ibarra, R. A. Peralta, J. Aguilar-Pliego and J. A. de los Reyes, *RSC Adv.*, 2023, **13**, 27174–27179.
- 47 S. H. Park, R. A. Peralta, D. Moon and N. C. Jeong, *J. Mater. Chem. A*, 2022, **10**, 23499–23508.
- 48 J. L. Obeso, J. Gabriel Flores, C. V. Flores, R. Rios-Escobedo, J. Aguilar-Pliego, A. Ken Inge, J. A. de los Reyes, R. A. Peralta, I. A. Ibarra and C. Leyva, *ChemCatChem*, 2023, **15**, e202300471.
- 49 S. Kobayashi and T. Busujima, *Chem. Commun.*, 1998, 981–982, DOI: [10.1039/A801464G](https://doi.org/10.1039/A801464G).
- 50 R. A. Peralta, M. T. Huxley, P. Lyu, M. L. Diaz-Ramirez, S. H. Park, J. L. Obeso, C. Leyva, C. Y. Heo, S. Jang, J. H. Kwak, G. Maurin, I. A. Ibarra and N. C. Jeong, *ACS Appl. Mater. Interfaces*, 2023, **15**, 1410–1417.
- 51 J. L. Obeso, A. Lopez-Olvera, C. V. Flores, R. A. Peralta, I. A. Ibarra and C. Leyva, *Chem. Commun.*, 2023, **59**, 3273–3276.
- 52 R. A. Marquez, J. L. Obeso, R. R. Vaidyula, V. B. López-Cervantes, R. A. Peralta, P. Marín Rosas, J. A. de los Reyes, C. B. Mullins and I. A. Ibarra, *J. Mater. Chem. A*, 2024, DOI: [10.1039/D4TA03620D](https://doi.org/10.1039/D4TA03620D).
- 53 W. Gong, Y. Xie, A. Yamano, S. Ito, X. Tang, E. W. Reinheimer, C. D. Malliakas, J. Dong, Y. Cui and O. K. Farha, *J. Am. Chem. Soc.*, 2023, **145**, 26890–26899.

