

Cite this: *Dalton Trans.*, 2024, **53**, 2512Received 22nd November 2023,
Accepted 31st December 2023

DOI: 10.1039/d3dt03915c

rsc.li/dalton

Photo-induced valence tautomerism and polarization switching in mononuclear cobalt complexes with an enantiopure chiral ligand†

Wen-Huang Xu, Yu-Bo Huang, Wen-Wei Zheng, Sheng-Qun Su, Shinji Kanegawa,^{id} Shu-Qi Wu^{id}* and Osamu Sato^{id}*

Light-induced polarization switchable molecular materials have attracted attention for decades owing to their potential remote manipulation and ultrafast responsiveness. Here we report a valence tautomeric (VT) complex with an enantiopure chiral ligand. By a suitable choice of counter anions, a significant improvement in photoconversion has been demonstrated, leading to novel photo-responsive polarization switching materials.

Owing to the versatility of their switching properties, switchable molecular materials have become ideal candidates for energy storage and light-recording devices.^{1–4} Valence tautomeric (VT) compounds, which possess two nearly degenerate electronic states resulting from a charge transfer process between transition metal ions and redox-active ligands, exhibit desirable properties such as switchable valence, spin multiplicity, and polarization.^{4–17} The rich chemical tailorability in the molecular design and properties of VT compounds renders them highly attractive for the development of a wide variety of functional materials.^{5,13} Among a variety of external stimuli, light has received extensive research in recent years due to its fast response and remote control.^{18–25}

In 2020, our group reported a cobalt mononuclear compound, [Co(phendiox)(*rac*-cth)]ClO₄·0.5EtOH⁷ (H₂phendiox = 9,10-dihydroxyphenanthrene, *rac*-cth = racemic 5,5,7,12,12,14-hexamethyl-1,4,8,11-tetraazacyclotetradecane), with a thermally induced polarization change during the VT process, owing to the change of its molecular dipole moment in a polar crystallographic space group. Upon temperature variations, the pyroelectric effect was observed in the same temperature range as with the VT process. However, the low light conversion efficiency (8%) and relatively fast relaxation dynamics at low temperatures precluded the direct observation of the light-induced polarization switching properties.

Notably, various factors, such as counterions, solvents and crystal packing, play pivotal roles in modulating intermolecular interactions.^{26–30} These interactions, in turn, influence the electron transfer processes within VT compounds (Fig. 1).

Extensive efforts involving the substitution of co-crystallized solvents and counterions were undertaken to enhance the photoconversion of [Co(phendiox)(*rac*-cth)]⁺. However, most of such compounds tend to crystallize in nonpolar space groups,³¹ and the molecular dipole moments cancel out, leading to a cancellation of molecular dipole moments and consequently, a lack of polarization change in the crystalline state.

Herein, we introduced an enantiopure chiral ligand as the strategy to break the crystallographic centrosymmetry, yielding VT compounds [Co(SS-cth)(phendiox)]PF₆·0.5EtOH (1·0.5EtOH) and [Co(SS-cth)(phendiox)]AsF₆·0.5EtOH (2·0.5EtOH). Notably, complex 2·0.5EtOH crystallizes in the polar *P*₂₁ space group and exhibits the lowest VT transition temperature in this family of compounds. Correspondingly, it shows an improvement in the photoconversion properties; at 5 K, complex 2·0.5EtOH exhibits a conversion ratio of *ca.* 27% and an enhanced lifetime of electron-transferred metastable

Institute for Materials Chemistry and Engineering & IRCCS, Kyushu University, 744 Motoooka, Nishi-ku, Fukuoka 819-0395, Japan. E-mail: sato@cm.kyushu-u.ac.jp
jpwu.shuqi.152@m.kyushu-u.ac.jp

† Electronic supplementary information (ESI) available. CCDC 2309543 and 2309544. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d3dt03915c>

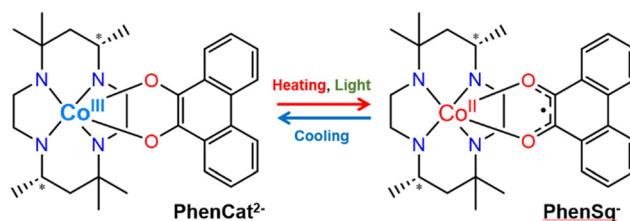


Fig. 1 Electron transfer behaviour in [Co(phendiox)(SS-cth)]⁺ motifs.



states of *ca.* 4000 seconds, which suggests that a significant polarization switching could be realized in these materials.

The direct current magnetic susceptibilities (χ_m) of 1-0.5EtOH and 2-0.5EtOH were measured in a 2000 Oe external magnetic field (Fig. 2). To prevent loss of co-crystallized ethanol molecules, the measurements of the pristine samples were performed from 300 to 5 K, followed by heating up to 400 K. The samples were then maintained at 400 K for an adequate duration of time, and another complete measurement cycle between 400 and 5 K was conducted to determine the magnetic properties of the desolvated samples.

These complexes exhibit similar magnetic behavior. The room-temperature $\chi_m T$ values of both complexes were slightly larger than the spin-only values of an uncoupled $hs-Co^{II}$ -semi-quinonato (PhenSq⁻) pair (1: 2.75 cm³ K mol⁻¹, 2: 3.0 cm³ K mol⁻¹), indicating a predominant presence of this electronic configuration.⁷ Below 200 K, the $\chi_m T$ values dropped drastically to 0.5–0.7 cm³ K mol⁻¹, corresponding to an increase in the population in the diamagnetic $ls-Co^{III}$ -catecholato (PhenCat²⁻) ground state.³² Notably, the $T_{1/2}$ (defined as the peak temperature of the first order derivative of $\chi_m T$ - T plots) of 2 (*ca.* 190 K) is slightly lower than that of 1 (*ca.* 200 K), suggesting the enhanced stability of high-spin species in the lattice of the hexafluoroarsenate salt, potentially attributable to the negative chemical pressure exerted by the AsF₆⁻ anion owing to its larger size compared to PF₆⁻. During heating to 400 K, neither 1 nor 2 exhibit significant thermal hysteresis in $\chi_m T$ values. Upon solvent loss (Fig. S1[†]),^{33,34} the $T_{1/2}$ for 1 and 2 decrease by 15 and 5 K, respectively. The solvent effect observed here closely resembles that of the previously reported complex, [Co(*rac*-cth)(phenidiox)]ClO₄·0.5EtOH. However, given

that 1-0.5EtOH and 2-0.5EtOH exhibit significantly lower transition temperatures, an enhancement in the stability of the light-induced metastable high-spin state is anticipated, possibly leading to a better photoconversion performance.³⁵

The X-ray single-crystal structures of 2-0.5EtOH at 78 and 300 K were obtained at the Spring-8 synchrotron facility. Crystallographic parameters and refinement parameters are summarized in ESI tables.[†] Different from its racemic counterpart in the *C2/c* space group, complex 2-0.5EtOH crystallized in the polar *P2₁* space group.^{31,36} The use of chiral ligands plays a key role in the loss of the inversion symmetry in the crystal. In the molecular structure (Fig. 3), the Co center is octahedrally coordinated by four N atoms from the cth ligand and two O atoms from the phenidiox ligand. At low temperature, two crystallographically independent molecular motifs could be found in the unit cell, distinguishable by the Co–O bond lengths; one exhibits lengths of 1.898 and 1.907 Å (α site), characteristic of the Co^{III} ion, whereas the other shows lengths of 2.017 and 1.999 Å (β site).

The relatively longer Co–O bond lengths revealed an incomplete VT transition at β sites, which is consistent with the non-vanishing $\chi_m T$ value at this temperature. The difference in the

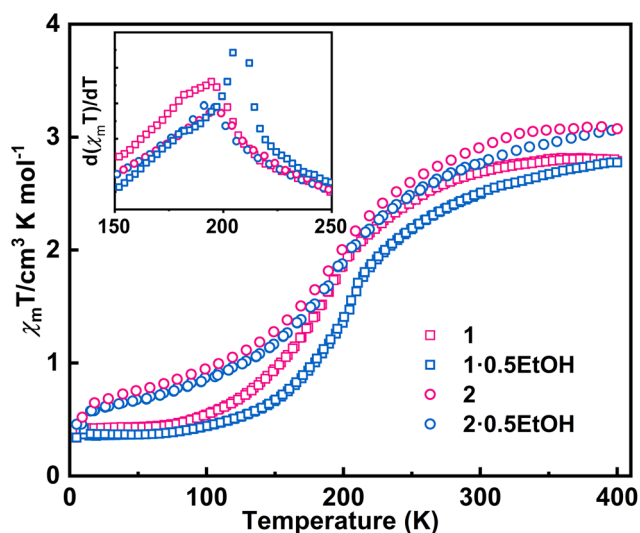


Fig. 2 Magnetic properties of complexes 1 and 2. The temperature-dependent magnetic susceptibility was performed at a sweeping rate of 5 K min⁻¹. The magnetic susceptibilities of compounds 1 and 2 are shown by open squares and circles, respectively. Magnetic properties before and after desolvation are marked in red and blue, respectively. Inset: First-order derivatives of the $\chi_m T$ product in the cooling process.

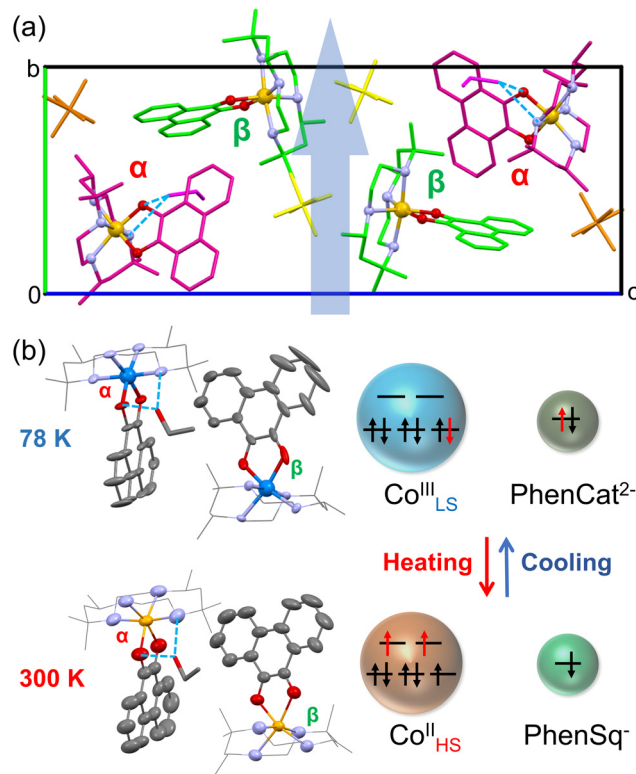


Fig. 3 Crystal structure and packing of 2-0.5EtOH. (a) Molecular packing viewed along the *a*-axis. The Co is coordinated by four nitrogen atoms (in lilac) and two oxygen atoms (in red). Inequivalent motifs are labelled in red (α site) and green (β site), respectively. The polarization direction is indicated by a blue arrow. (b) Structure of the asymmetric unit in the low temperature phase (upper) and the high temperature phase (lower). Counter ions and hydrogen atoms are omitted for clarity.



completeness of the VT transition might stem from the distinct hydrogen-bonded structures at each site. Molecules at α sites show stronger hydrogen bond interactions, with an N-H...O distance of 3.019 Å and an O-H...O distance of 2.885 Å. In contrast, no significant hydrogen bonds form between ethanol and molecules at the β site. The different chemical environment leads to a different spin state at a low temperature. At 300 K, the Co-O bond lengths at both sites increased to the typical values of the Co^{II} ion, ranging from 2.050 Å to 2.092 Å. Such changes in bond length are consistent with the usual VT behavior in the [Co(cth)(phendiox)]⁺ family.

Temperature-dependent infrared (IR) and UV-visible absorption spectra also confirmed the presence of electron transfer between metal and ligand sites of complex 2 (Fig. 4); the absorption peaks at approximately 1354 and 1378 cm⁻¹, which decreased upon heating, correspond to the in-phase C-O vibration and C=C stretching in phenCat²⁻. The absorption feature at around 1600 cm⁻¹ is identified as the C-C-C stretching mode, typical of the [Co^{III}(cth)(phencAT)]⁺ family.⁷ Meanwhile, the absorption peak at approximately 1458 cm⁻¹ is

ascribed to the C-O and C=C stretching modes in phenSq⁻. Similarly, the characteristic π - π^* transition band of phenCat²⁻ centered at 520 nm increases with heating, confirming the electron transfer nature of the temperature-dependent electronic structures.

Based on the density functional theory (DFT) calculations reported previously,⁷ the [Co(cth)(phendiox)]⁺ motif possesses a permanent electric dipole moment of 13.26 Debye for a ls-Co^{III}-catecholato state and 7.47 Debye for a hs-Co^{II}-semiquinonato state corresponding to the LT and HT phases, respectively. However, the change in net molecular dipole moments in a unit cell is largely cancelled owing to the approximately perpendicular alignment of the [Co(cth)(phendiox)]⁺ motif with respect to the polar b -axis. The dipole moment direction could be estimated from the O-Co-O bisection line, which forms angles of 100.15° and 89.75° with the b -axis for α and β molecules, respectively, at 300 K. At 78 K, these angles change slightly to 103.09° and 90.00°. Projecting the dipole moments to the polar axis, the change induced by the VT process is approximately 3.64 Debye per unit cell, corresponding to a polarization change of 353 nC cm⁻².

Photomagnetic susceptibility measurements were performed using a 532 nm wavelength laser. An increase in $\chi_m T$ from 0.4 to 1.2 cm³ mol⁻¹ K at 5 K indicated a photoinduced charge transfer process from ls-Co(III) to hs-Co(II). A gradual decrease in the $\chi_m T$ value upon increasing the temperature was then observed, followed by a collapse back to the normal (in the absence of photoirradiation) value at 53 K (Fig. 5). As the temperature increased, the curve coincided with the $\chi_m T$ vs. T plot of the bulk sample. The percentage of photoconversion was estimated to be nearly 27% using the reported high-temperature $\chi_m T$ value (approximately 3.0 cm³ mol⁻¹ K) of the same family of compounds as the upper limit.

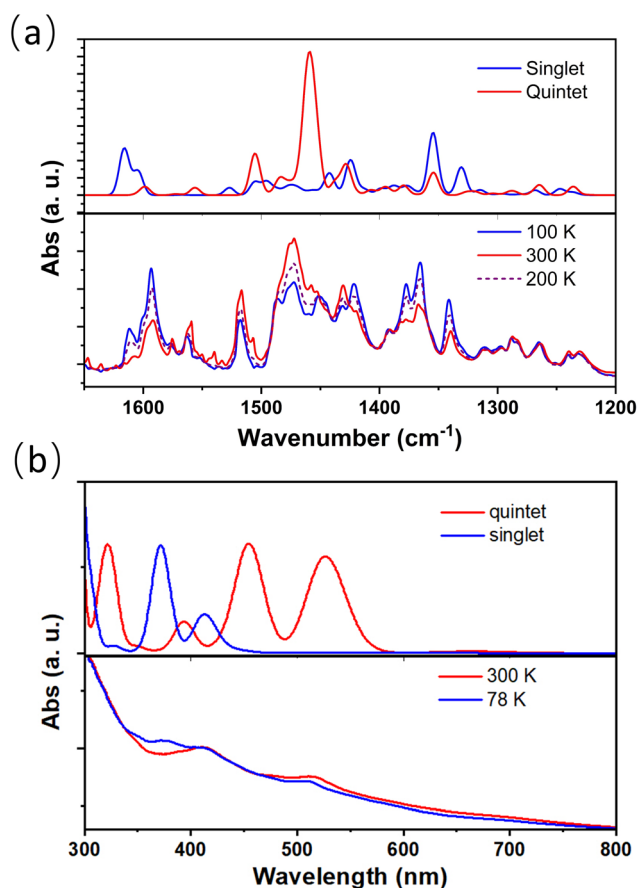


Fig. 4 Experimental and calculated temperature-dependent IR and UV spectra of complex 2·0.5EtOH. (a) Calculated IR absorption for singlet and quintet states (upper), and temperature-dependent IR spectra recorded between 100 and 300 K (lower). (b) Calculated UV absorption spectra for singlet and quintet states (upper), and temperature-dependent UV spectra recorded between 78 and 300 K (lower).

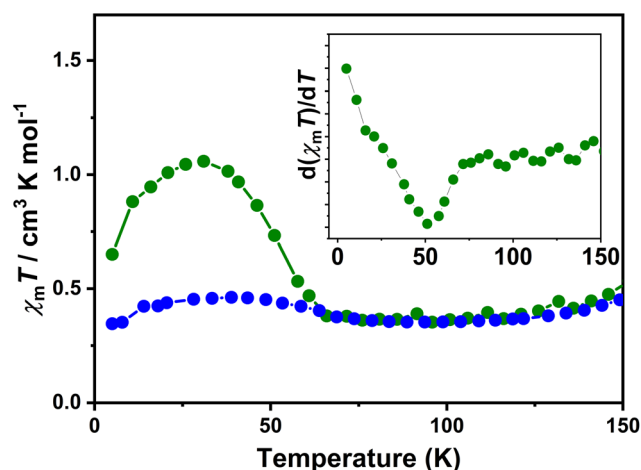


Fig. 5 Temperature-dependent magnetic susceptibilities of complex 2·0.5EtOH after irradiation (green). Light-induced magnetization is performed by 532 nm irradiation. Magnetization of the bulk sample is shown in blue circles. Inset: First-order derivatives of the light induced $\chi_m T$ product.



To further investigate the relaxation dynamics from the photoinduced metastable state, the time dependence of the magnetic susceptibility after irradiation was measured at different temperatures. The fraction of the metastable state as a function of time can be defined as

$$\gamma(t) = \frac{M(t) - M_{\text{OFF}}}{M(0) - M_{\text{OFF}}} \quad (1)$$

where $M(t)$ is the magnetization value at time t after irradiation, $M(0)$ is the magnetization at $t = 0$ when the system was saturated by a laser, and M_{OFF} is the magnetization in the absence of irradiation. The normalized HS ratio vs. time (t) curve was fitted using the stretched exponential function (2) as follows:

$$\gamma(t) = (1 - \gamma)e^{-\left(\frac{t}{\tau}\right)^\beta} + \gamma \quad (2)$$

The parameter γ indicates the phenomenological HS fraction that persists in the infinite time limit, τ denotes the characteristic relaxation time and β indicates its distribution. Remarkably, the relaxation time for complex **2** reached *ca.* 4000 s at 5 K, which is over 3 times longer than that observed for [Co(*rac*-cth)(phendiox)]ClO₄ and [Co(*SS*-cth)(phendiox)]PF₆. This suggests the enhanced stability of the metastable state, which correlates with the substantial improvement of the photoconversion performance in this system (Fig. 6). Such a trend aligns with the established understanding that lower thermally-induced transition temperatures are indicative of slower relaxation times for the light-induced excited spin states within spin crossover systems.³⁷ Furthermore, it is important to recognize that complex **2**, with its polar crystal structure, allows for achieving a persistent polarization change through light irradiation by trapping the electron-transferred state. By assuming the Gaussian broadening of the pyroelectric coefficient during the relaxation process, we estimate its maximal value as 3.19 nC cm⁻² K⁻¹, which lies within the

detectable range of a standard Keithley 6517B electrometer. However, the practical realization of such measurements poses significant challenges due to the extremely thin thickness of the crystals of complex **2**, making experimental observations difficult to carry out.

Conclusions

In this study, we introduced an enantiopure cth ligand to obtain two isostructural VT compounds crystallized in the polar $P2_1$ space group. By using the hexafluoroarsenate anion in complex **2**, a lowering of VT transition temperatures by *ca.* 10 K has been observed, suggesting the improved stability of the hs-Co(II)-semiquinonato species in this lattice. Correspondingly, the light conversion efficiency and the lifetime of the metastable state were improved to approximately 27% and 4000 s at 5 K, respectively. As the electric dipole moments change substantially during the electron transfer process, light-induced polarization switching should be expected for this family of compounds. These results suggest the feasibility of achieving photoinduced persistent polarization by these strategies, which would lead to a promising outcome for applications in molecular electronics and optically rewritable information storage.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Wen-Huang Xu was supported by JST SPRING, Grant Number JPMJSP2136. Yu-Bo Huang thanks the China Scholarship Council for support. This work was supported by JSPS KAKENHI (21K05085, 22K14694, 20H00385 and 21K05086). This work was performed under the Cooperative Research Program of "Network Joint Research Center for Materials and Devices". This work was supported by the MEXT Project of "Integrated Research Consortium on Chemical Sciences". Shu-Qi Wu and Sheng-Qun Su are grateful for support from the Iketani Science and Technology Foundation (0341056A and 0351039A).

References

- O. Sato, *Nat. Chem.*, 2016, **8**, 644–656.
- S. Horiuchi and Y. Tokura, *Nat. Mater.*, 2008, **7**, 357–366.
- M. D. Manrique-Juárez, S. Rat, L. Salmon, G. Molnár, C. M. Quintero, L. Nicu, H. J. Shepherd and A. Bousseksou, *Coord. Chem. Rev.*, 2016, **308**, 395–408.
- Y. Sekine, R. Akiyoshi and S. Hayami, *Coord. Chem. Rev.*, 2022, **469**, 214663.

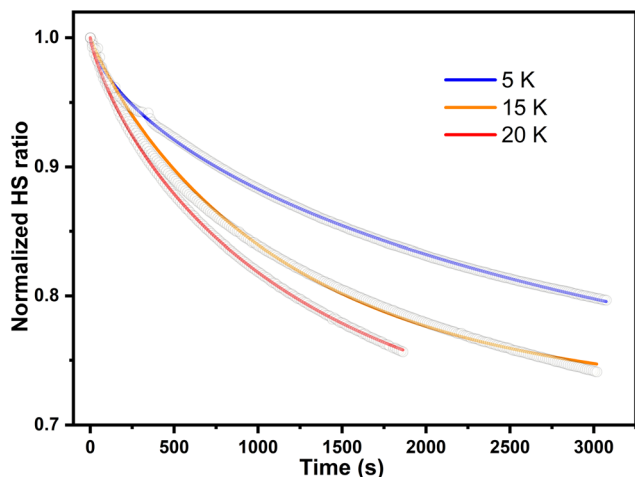


Fig. 6 The relaxation plots of photoinduced magnetization of complex **2**·0.5EtOH at 5, 15 and 20 K, which are shown in blue, orange, and red respectively. The wavenumber of excitation light is 532 nm.



- 5 G. K. Gransbury and C. Boskovic, *Encyclopedia of Inorganic and Bioinorganic Chemistry*, pp. 1–24. DOI: [10.1002/9781119951438.eibc2785](https://doi.org/10.1002/9781119951438.eibc2785).
- 6 S.-Q. Su, S.-Q. Wu, S. Kanegawa, K. Yamamoto and O. Sato, *Chem. Sci.*, 2023, **14**, 10631–10643.
- 7 S.-Q. Wu, M. Liu, K. Gao, S. Kanegawa, Y. Horie, G. Aoyama, H. Okajima, A. Sakamoto, M. L. Baker, M. S. Huzan, P. Bencok, T. Abe, Y. Shiota, K. Yoshizawa, W. Xu, H.-Z. Kou and O. Sato, *Nat. Commun.*, 2020, **11**, 1992.
- 8 O. Sato, T. Iyoda, A. Fujishima and K. Hashimoto, *Science*, 1996, **272**, 704–705.
- 9 H. W. Liu, K. Matsuda, Z. Z. Gu, K. Takahashi, A. L. Cui, R. Nakajima, A. Fujishima and O. Sato, *Phys. Rev. Lett.*, 2003, **90**, 167403.
- 10 S. Kanegawa, Y. Shiota, S. Kang, K. Takahashi, H. Okajima, A. Sakamoto, T. Iwata, H. Kandori, K. Yoshizawa and O. Sato, *J. Am. Chem. Soc.*, 2016, **138**, 14170–14173.
- 11 F. Cheng, S. Wu, W. Zheng, S. Su, T. Nakanishi, W. Xu, P. Sadhukhan, H. Sejima, S. Ikenaga, K. Yamamoto, K. Gao, S. Kanegawa and O. Sato, *Chem. – Eur. J.*, 2022, **28**, e202202161.
- 12 J. Tao, H. Maruyama and O. Sato, *J. Am. Chem. Soc.*, 2006, **128**, 1790–1791.
- 13 A. Dei, D. Gatteschi, C. Sangregorio and L. Sorace, *Acc. Chem. Res.*, 2004, **37**, 827–835.
- 14 O. Drath and C. Boskovic, *Coord. Chem. Rev.*, 2018, **375**, 256–266.
- 15 J. Ferrando-Soria, J. Vallejo, M. Castellano, J. Martínez-Lillo, E. Pardo, J. Cano, I. Castro, F. Lloret, R. Ruiz-Garcia and M. Julve, *Coord. Chem. Rev.*, 2017, **339**, 17–103.
- 16 G. K. Gransbury, B. N. Livesay, J. T. Janetzki, M. A. Hay, R. W. Gable, M. P. Shores, A. Starikova and C. Boskovic, *J. Am. Chem. Soc.*, 2020, **142**, 10692–10704.
- 17 T. Tezgerevska, K. G. Alley and C. Boskovic, *Coord. Chem. Rev.*, 2014, **268**, 23–40.
- 18 P. Sadhukhan, S.-Q. Wu, S. Kanegawa, S.-Q. Su, X. Zhang, T. Nakanishi, J. I. Long, K. Gao, R. Shimada, H. Okajima, A. Sakamoto, J. G. Chiappella, M. S. Huzan, T. Kroll, D. Sokaras, M. L. Baker and O. Sato, *Nat. Commun.*, 2023, **14**, 3394.
- 19 S.-Q. Su, S.-Q. Wu, Y.-B. Huang, W.-H. Xu, K.-G. Gao, A. Okazawa, H. Okajima, A. Sakamoto, S. Kanegawa and O. Sato, *Angew. Chem., Int. Ed.*, 2022, **61**, e202208771.
- 20 A. Tashiro, S. Kanegawa, O. Sato and Y. Teki, *Polyhedron*, 2013, **66**, 167–170.
- 21 T. J. Penfold, J. O. Johansson and J. Eng, *Coord. Chem. Rev.*, 2023, **494**, 215346.
- 22 A. Witt, F. W. Heinemann and M. M. Khusniyarov, *Chem. Sci.*, 2015, **6**, 4599–4609.
- 23 H. Kuramochi, G. Aoyama, H. Okajima, A. Sakamoto, S. Kanegawa, O. Sato, S. Takeuchi and T. Tahara, *Angew. Chem., Int. Ed.*, 2020, **59**, 15865–15869.
- 24 R. Ash, K. Zhang and J. Vura-Weis, *J. Chem. Phys.*, 2019, **151**, 104201.
- 25 O. Sato, A. Cui, R. Matsuda, J. Tao and S. Hayami, *Acc. Chem. Res.*, 2007, **40**, 361–369.
- 26 E. Evangelio, C. Rodriguez-Blanco, Y. Coppel, D. N. Hendrickson, J. P. Sutter, J. Campo and D. Ruiz-Molina, *Solid State Sci.*, 2009, **11**, 793–800.
- 27 Y. Mulyana, G. Poneti, B. Moubaraki, K. S. Murray, B. F. Abrahams, L. Sorace and C. Boskovic, *Dalton Trans.*, 2010, **39**, 4757–4767.
- 28 C. Reichardt, *Chem. Rev.*, 1994, **94**, 2319–2358.
- 29 A. Dei, A. Feis, G. Poneti and L. Sorace, *Inorg. Chim. Acta*, 2008, **361**, 3842–3846.
- 30 G.-L. Li, S. Kanegawa, Z.-S. Yao, S.-Q. Su, S.-Q. Wu, Y.-G. Huang, S. Kang and O. Sato, *Chem. – Eur. J.*, 2016, **22**, 17130–17135.
- 31 A. Bencini, A. Caneschi, C. Carbonera, A. Dei, D. Gatteschi, R. Righini, C. Sangregorio and J. V. Slageren, *J. Mol. Struct.*, 2003, **656**, 141–154.
- 32 G. Poneti, M. Mannini, B. Cortigiani, L. Poggini, L. Sorace, E. Otero, P. Sainctavit, R. Sessoli and A. Dei, *Inorg. Chem.*, 2013, **52**, 11798–11805.
- 33 V. L. Nadurata and C. Boskovic, *Inorg. Chem. Front.*, 2021, **8**, 1840–1864.
- 34 V. L. Nadurata, M. A. Hay, J. T. Janetzki, G. K. Gransbury and C. Boskovic, *Dalton Trans.*, 2021, **50**, 16631–16646.
- 35 R. D. Schmidt, D. A. Shultz and J. D. Martin, *Inorg. Chem.*, 2010, **49**, 3162–3168.
- 36 A. Caneschi, A. Dei, F. Fabrizi de Biani, P. Gütllich, V. Ksenofontov, G. Levchenko, A. Hofer and F. Renz, *Chem. – Eur. J.*, 2001, **7**, 3926–3930.
- 37 V. A. Money, C. Carbonera, J. Elhaik, M. A. Halcrow, J. A. K. Howard and J.-F. Létard, *Chem. – Eur. J.*, 2007, **13**, 5503–5514.

