

Cite this: *Mater. Adv.*, 2024,  
5, 5896

# Highly-robust nanoplate-shaped V<sub>2</sub>O<sub>5</sub> as an efficient cathode material for aqueous zinc ion batteries

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The efficiency and performances of energy storage and conversion systems are highly dependent on the electrode performances, which have a direct impact on their functioning. The fabrication of electrode materials with novel nanostructures has a substantial positive impact on their electrochemical performances. Recently, two-dimensional nanomaterials have drawn a lot of attention due to their significant features, such as distinctive 2D-layered structure and infinite planar lengths as well as providing short routes for ion and electron transportation including large surface areas for additional adsorption sites. Herein, we have proposed a simple solvothermal synthesis for the fabrication of 2D nanoplates of a V<sub>2</sub>O<sub>5</sub> cathode for rechargeable aqueous zinc-ion batteries. The obtained high electrochemical results confirmed the potency of the V<sub>2</sub>O<sub>5</sub> nanoplate cathode for zinc ion batteries. Furthermore, the Zn<sup>2+</sup> ion storage mechanism within the V<sub>2</sub>O<sub>5</sub> crystal lattice is also discussed, which is based on the phase transition from pristine V<sub>2</sub>O<sub>5</sub> to zinc pyrovanadate (Zn<sub>x</sub>V<sub>2</sub>O<sub>5</sub>·nH<sub>2</sub>O) during reversible Zn<sup>2+</sup> (de)-intercalation in the open-structured hosts.

Received 18th April 2024,  
Accepted 30th May 2024

DOI: 10.1039/d4ma00403e

rsc.li/materials-advances

## Introduction

In today's modern civilization, the concerns related to energy storage, conversion, and distribution are crucial. An energy-storage device is considered to be ideal if it possesses the vital features of high safety, environmental friendliness, low cost, high specific capacity, fast charge/discharge capability, and high gravimetric and volumetric energy densities.<sup>1–4</sup> Lithium-ion batteries (LIBs) have garnered a lot of attention in recent years due to their numerous applications from small mobile gadgets to hybrid and electric cars. Despite the high energy and power density of LIBs, their wide-scale implementation is constrained by the safety risks caused by flammable organic electrolytes. In addition, the limited Li resources are also a major concern, which may restrict the further use of LIBs.<sup>5–8</sup> Recently, the rechargeable aqueous zinc-ion batteries have shown potential to be used as an alternative due to the low cost and high abundance of metallic zinc in the earth's crust (~2 USD per kg and ~75 ppm), excellent safety characteristics and low equilibrium redox potential (−0.76 V vs. SCE) of the Zn

electrode, and easy assembling procedures.<sup>9–14</sup> Similar to LIBs, the chemistry of aqueous zinc-ion batteries involves the movement of Zn<sup>2+</sup> ions between the cathode and anode. The theoretical capacity of metallic zinc is ~820 mA h g<sup>−1</sup> and it has a high kinetic overpotential for hydrogen evolution, which makes it stable with an aqueous electrolyte.<sup>9–12</sup> Despite many advantages of the zinc metal as an anode, a suitable cathode material is highly required, which is currently restricting the advancement of rechargeable Zn-ion batteries. Various cathode materials for aqueous Zn-ion batteries are currently under investigation, such as Prussian blue analogues,<sup>15</sup> polyanion metal phosphates,<sup>16</sup> quinone analogues,<sup>17</sup> manganese dioxide,<sup>18</sup> vanadium-based compounds, *etc.*<sup>19</sup> However, the undesirable phase transitions, co-insertion of H<sup>+</sup> along with Zn<sup>2+</sup>, structural failure and material disintegration were found to be major drawbacks associated with cathode materials.<sup>15–21</sup> Thus, to select a novel cathode material with high reversible capacity and a long cycle life is still a challenge for aqueous zinc ion batteries.

Recently, the V<sub>2</sub>O<sub>5</sub> electrode material was identified as a promising cathode due to its high specific capacity (~589 mA h g<sup>−1</sup>), rich resources, layered structure with a stable open-framework, steady cycling performance and outstanding rate capability. However, the structural instability during long-term cycling and the low electrical conductivity are still unresolved issues of the V<sub>2</sub>O<sub>5</sub> cathode.<sup>19–23</sup> Among all the approaches to optimize these issues, the priority is given to the fabrication of novel nanostructures, which in fact offers numerous intriguing properties such as more active sites for redox reactions, short diffusion paths for Zn<sup>2+</sup> ions, high surface area and

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large electrode/electrolyte contact area, as well as suppresses the structural breakdown to a certain extent induced by volume changes, resulting in excellent electrochemical performances.

Thus, we have fabricated here a two-dimensional (2D) nanoplate-shaped  $V_2O_5$  cathode material using a facile solvothermal synthesis method and tested its suitability for aqueous zinc ion batteries. Recently, 2D nanomaterials have been significantly used as electrode materials for energy storage applications due to their high mechanical integrity as they are stacked by van der Waals forces and made up by the substantial overlapping of the layers. In addition, the major advantages regarding atomic arrangement, coordination number, bond lengths, atomic or molecular thickness and infinite planar lengths of 2D nanomaterials make them superior to their bulk counterparts.<sup>24,25</sup> The obtained results demonstrate that the electrochemical performances of a nanoplate-shaped  $V_2O_5$  cathode at high current rate are still stable. This research is expected to provide a feasible strategy for designing and creating promising vanadium-based cathode materials for rechargeable aqueous zinc-ion batteries.

## Experimental

### Material synthesis

Initially, 6 mmol of vanadium acetylacetonate (Sigma Aldrich, India, 97%) was dissolved in 30 mL each of anhydrous carbinol and ethylene glycol under continuous stirring. Thereafter, liquid ammonia ( $NH_4OH$ ) was added dropwise to maintain pH  $\sim 9$  in the above precursor solution under vigorous stirring. Then, after continuous stirring for 45 min, the obtained solution was further transferred to a Teflon-lined stainless-steel autoclave and heated up to 180 °C for 12 h. The as-prepared product was collected after natural cooling followed by centrifugation and washing with distilled water and ethanol several times to remove the organic impurities. After overnight

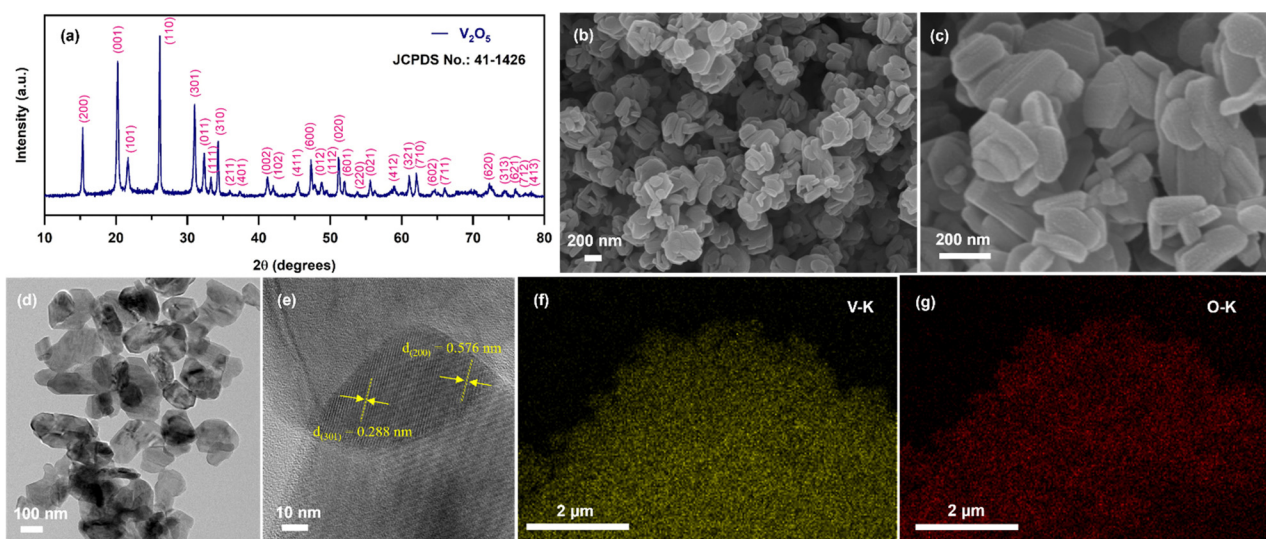
drying of the as-prepared sample at 80 °C, the calcination was performed in an air atmosphere at 450 °C for 1 h with the heating rate of 1° min<sup>-1</sup>.

### Material characterisation

The crystallinity and purity of the calcined sample was determined by X-ray diffraction analysis (XRD, Shimadzu X-ray diffractometer with Cu K $\alpha$  radiation,  $\lambda = 1.5406 \text{ \AA}$ ). The surface morphology, particle size, and shape of the calcined sample were analyzed by field-emission scanning electron microscopy (FE-SEM, S-4700 Hitachi) and field emission transmission electron microscopy (FE-TEM, Philips TecnaiF20 at 200 kV). Furthermore, the elemental mapping and surface chemical compositions of the calcined sample were characterized by energy dispersive spectroscopy (EDS) and X-ray photoelectron spectroscopy (XPS; M/s Thermo Fischer Scientific Instrument UK with a monochromatic Al K $\alpha$  X-ray source), respectively. Raman analysis was carried out using a Renishaw Raman Microscope, with a 514 nm laser to analyse the chemical bonding states of the product. The surface area including pore size distribution was investigated by the Brunauer–Emmett–Teller method using  $N_2$  adsorption/desorption isotherms (BET, Micromeritics Instrument Corp. Gemini-V).

### Electrode fabrication

A CR-2032 coin-type half-cell of the  $V_2O_5$  cathode was assembled using zinc metal foil as a reference electrode, glass fibre as a separator and 2 M  $ZnSO_4$  as an electrolyte. Initially, to fabricate a working electrode, the active materials (70 wt%), acetylene black (15 wt%) and polyacrylic acid (PAA) (15 wt%) were pressed onto a stainless-steel mesh, which was then dried under vacuum at 120 °C for 12 h. The cells were assembled in an open air environment and they were ideally kept for 12 h for aging before the electrochemical measurements. The electrochemical measurements were performed using a WBCS 3000 (WonATech) battery tester between the potential



**Fig. 1** (a) XRD pattern; (b, c) FE-SEM images at different magnifications; (d) FE-TEM image; (e) HR-TEM image; and (f, g) representative EDX mapping images of the  $V_2O_5$  sample.



range of 0.2–1.6 V. Cyclic voltammetry (CV) was also carried out using an AUTOLAB PGSTAT302N potentiostat for the initial five cycles at the constant scan rate of  $0.2 \text{ mV s}^{-1}$  under the same potential range.

## Results and discussion

### Physiochemical properties

Fig. 1(a) shows the XRD pattern of the calcined sample, which clearly demonstrates sharp peaks with high intensity, indicating high crystallinity of the sample. All the diffraction peaks of the sample are well consistent with the standard diffraction pattern of the orthorhombic phase of  $\text{V}_2\text{O}_5$  with  $Pmmn$  space group (JCPDS no: 00-041-1426). In addition, there were no other impurity peaks visible in the diffraction pattern, which implies that the calcined sample is highly pure. The morphological analysis of the calcined sample was then further investigated and the obtained results are shown in Fig. 1(b)–(d). Fig. 1(b) and (c) represent the FE-SEM images taken at different

magnifications, which clearly shows 2D nanoplate-type morphology. As can be seen from the figure, the nanoplates are highly aggregated with smooth surfaces. As shown in Fig. 1(d), TEM analysis was also performed for detailed microstructural analysis. The obtained TEM image is in accordance with the FE-SEM results. The nanoplate sizes are in the range of  $\sim 50$ – $200 \text{ nm}$ . The high-resolution TEM (HR-TEM) image displayed in Fig. 1(e) indicates the high-crystallinity of the calcined sample. The HR-TEM image demonstrates clear lattice fringes with interplanar spacing of  $\sim 0.571 \text{ nm}$  and  $0.288 \text{ nm}$ , corresponding to the (200) and (301) planes respectively, which is matched with the XRD pattern of the  $\text{V}_2\text{O}_5$  sample. The elemental-mapping images displayed in Fig. 1(f) and (g) confirmed the uniform dispersion of vanadium (V-K; yellow) and oxygen (O-K; red) elements without the existence of any other impurities throughout the sample, which indicates that the nanoplates are mainly composed of  $\text{V}_2\text{O}_5$ .

In order to investigate the oxidation states of the elements, XPS analysis was carried out and the obtained results are displayed in Fig. 2(a) and (b). Fig. 2(a) shows the survey

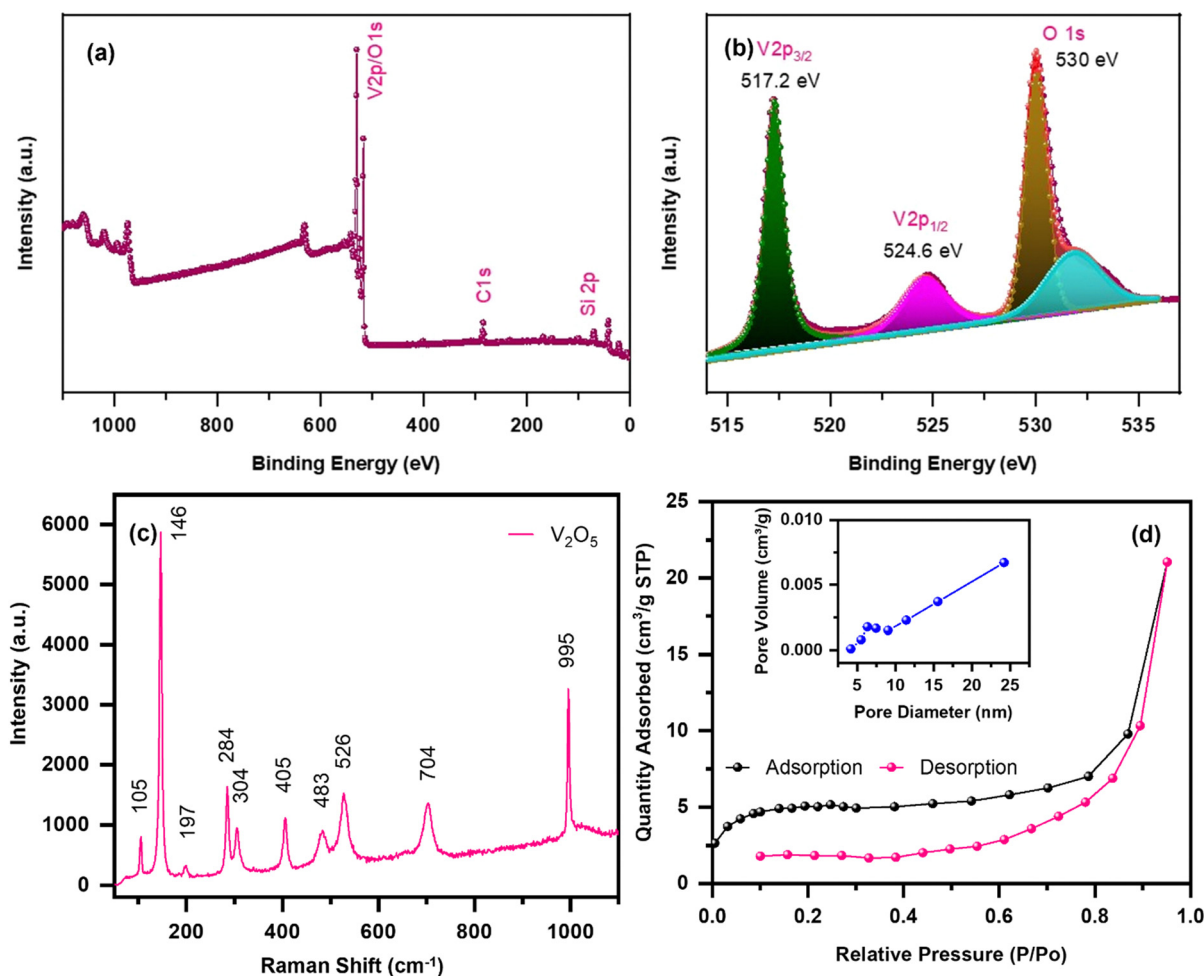


Fig. 2 XPS: (a) survey spectrum; (b) V-2p and O-1s spectrum; (c) Raman spectrum; (d) BET analysis with pore size distribution (inset) of the nanoplate-shaped  $\text{V}_2\text{O}_5$  sample.



spectrum of the calcined  $V_2O_5$  sample, which clearly shows the presence of vanadium (V) and oxygen (O) elements. More precisely, Fig. 2(b) represents the combined de-convoluted XPS spectra of V-2p and O-1s. As can be seen, the characteristic peaks obtained at binding energies of 517.2 and 524.6 eV correspond to  $V^{5+} 2p_{3/2}$  and  $2p_{1/2}$ , respectively. Since the binding energy difference between the core levels is found to be 7.4 eV, it is in good agreement with the electronic states of  $V_2O_5$ . The peaks located at binding energies of 530 eV and 531.9 eV of the O-1s spectrum belong to the V–O and V–OH bonds of the  $V_2O_5$  sample.<sup>26,27</sup> Raman analysis is also employed to investigate the bonding configurations using a variety of molecular vibrational modes, as shown in Fig. 2(c). The Raman spectrum was measured in the Raman shift range of 50–1100  $cm^{-1}$ , which exhibits various high and intense peaks located at  $\sim 105$ ,  $\sim 146$ ,  $\sim 197$ ,  $\sim 284$ ,  $\sim 304$ ,  $\sim 405$ ,  $\sim 483$ ,  $\sim 526$ ,  $\sim 704$  and  $\sim 995$   $cm^{-1}$ . The peaks present at 146 and 197  $cm^{-1}$  are related to the  $B_{3g}$  mode of V–O–V chains, which indicates the relative motion of  $V_2O_5$  layers with regard to one another, indicating high crystallinity of the sample. However, the other two peaks present at 284 and 405  $cm^{-1}$  show the bending vibration of V=O type bonding. On the other hand, the shoulder peaks obtained at 483 and 304  $cm^{-1}$  belong to bending vibrations of V–O–V (doubly coordinated oxygen) and  $V_3$ –O (triply coordinated oxygen) type, respectively.<sup>28,29</sup> The peaks positioned at 526 and 704  $cm^{-1}$  correspond to the stretching vibrations of triply co-ordinated oxygen ( $V_3$ –O) and doubly co-ordinated oxygen ( $V_2$ –O), respectively. Finally, the Raman shift present at 995  $cm^{-1}$  is due to the terminal oxygen ( $V^{5+}=O$ ) stretching mode.<sup>29,30</sup> In addition, BET analysis was also conducted to evaluate the specific surface area and porosity of the material. As depicted in Fig. 2(d), the  $N_2$  adsorption-desorption isotherms show a type IV isotherm, indicating mesoporosity in the sample. The BET surface area was found to be 15.94  $m^2 g^{-1}$ . Moreover, the Barrett–Joyner–Halenda (BJH) pore size distribution plot, as shown in the inset of the figure, demonstrates an average pore diameter of  $\sim 16.8$  nm and total pore volume of  $\sim 0.032$   $cm^3 g^{-1}$ , which also suggests that the material possesses mesopores. It is well known that a porous nanostructure and large surface area offers a short diffusion length, better electrolyte impregnation, numerous active sites for Zn-ion adsorption, *etc.*, and thereby excellent electrochemical performances.<sup>30</sup>

### Electrochemical investigation

Fig. 3 displays CV curves for the initial 5 cycles at a constant scan rate of 0.2  $mV s^{-1}$ . As can be seen, the first cycle shows two reduction peaks located at  $\sim 0.57$  V and  $\sim 0.9$  V, which is due to the insertion processes of Zn-ions in the  $V_2O_5$  nanoplate electrode. During the anodic scan, the oxidation peaks observed at  $\sim 1.08$  and  $\sim 1.35$  V can be attributed to the de-insertion of  $Zn^{2+}$  ions.<sup>31</sup> However, compared to the first cycle, the peak positions and peak current from the second cycle onwards are different, which may be due to the incomplete extraction of inserted  $Zn^{2+}$  ions from the  $V_2O_5$  nanostructure.<sup>19</sup> In addition, the cathodic peak observed at  $\sim 0.9$  V in the first

cycle splits into two peaks during the second cycle and appeared to be  $\sim 0.91$  V and  $\sim 1.1$  V, which could be described by the multiple insertion reactions of  $Zn^{2+}$  ions because of the chemical defects of vanadium oxide.<sup>32</sup> The peak located at  $\sim 0.91$  V and  $\sim 0.57$  V corresponds to the redox reaction of  $V^{5+}/V^{4+}$  and  $V^{4+}/V^{3+}$ , respectively.<sup>33</sup>

Furthermore, the charge/discharge data of the  $V_2O_5$  electrode was recorded at a constant current density of 0.5  $A g^{-1}$  for the initial 15 cycles. And out of 15 cycles, Fig. 3(b) shows the charge/discharge curves of the 1st, 2nd, 3rd, 4th, 5th and 10th cycles. Though the voltage profiles are similar, the high reversibility and stability can be easily seen in all the cycles except for the first discharge cycle of the electrode material. The significant irreversible capacity loss observed in the initial discharge cycle can be attributed to the formation of the SEI layer, which causes the kinetic limitations by consuming more number of  $Zn^{2+}$  ions during discharge, but does not allow extraction of the same number of  $Zn^{2+}$  ions during charge.<sup>19,34</sup> It is clear from the figure that from the second cycle onwards, the voltage profile shows high reversibility, which indicates a decrease in charge transfer resistance at the electrode/electrolyte contact, resulting in low polarization. The small polarization demonstrates fast  $Zn^{2+}$  transfer kinetics within the electrode. The electrode is able to deliver the charge/discharge capacities of  $\sim 141.2/170.06$ ,  $\sim 132.2/137.38$ ,  $\sim 123.7/124.26$ ,  $\sim 127.8/126.26$ ,  $\sim 132.2/132.03$  and  $\sim 141.06/143.46$   $mA h g^{-1}$  for the 1st, 2nd, 3rd, 4th, 5th and 10th cycles, respectively, which clearly indicates  $\sim 100\%$  coulombic efficiency of the electrode. On the other hand, Fig. 3(c) represents the galvanostatic charge/discharge cycles of the  $V_2O_5$  electrode for the 30th, 60th, 100th, 500th and 1000th cycles at a constant current density of 5  $A g^{-1}$ . It can be observed that the capacity decreases with increase of the cycle numbers. However, the coulombic efficiency was still  $\sim 100\%$  throughout the cycling and the charge/discharge capacity was found to be  $\sim 69/69.8$   $mA h g^{-1}$ ,  $\sim 67.8/68.8$   $mA h g^{-1}$ ,  $\sim 63.40/61.34$   $mA h g^{-1}$ ,  $56.34/56.34$   $mA h g^{-1}$  and  $\sim 43.81/43.81$   $mA h g^{-1}$ , respectively for the above-mentioned cycle numbers. In addition, it is also believed that the high surface area along with the electrode may also tolerate the absolute volume changes of the electrode during the insertion/de-insertion of  $Zn^{2+}$  ions. In order to investigate the stability of the nanoplate-shaped  $V_2O_5$  electrode, C-rate was performed at various current densities of 0.5  $A g^{-1}$  to 20  $A g^{-1}$  and the obtained results are displayed in Fig. 3(d). The  $V_2O_5$  cathode delivers the specific capacities of  $\sim 183$ ,  $\sim 168$ ,  $\sim 156$ ,  $\sim 133$ ,  $\sim 104$  and  $\sim 59$   $mA h g^{-1}$  at current densities of 0.5, 1, 2, 5, 10 and 20  $A g^{-1}$ , respectively. Impressively, when the current rate returns back to the initial current rate of 0.5  $A g^{-1}$  after 60 cycles, the electrode is still able to regain the original capacity of  $\sim 180$   $mA h g^{-1}$  with  $\sim 100\%$  coulombic efficiency, which clearly indicates the high stability of the electrode material. In addition, the continuous capacity loss with increase of the current density during cycling is a common phenomenon as  $Zn^{2+}$  ions do not get sufficient time to (de)-intercalate at high current rates.<sup>35,36</sup> More importantly, it can be seen that the electrode exhibits 32% of the initial capacity at such a high current density of 20 000  $mA g^{-1}$  (20  $A g^{-1}$ ), which also suggests that the nanoplate morphology maintains the integrity of





the electrode during volume expansion/contraction resulting from  $\text{Zn}^{2+}$  (de)-insertion. Fig. 3(e) shows the specific capacity *versus* cycle number graph including the corresponding coulombic efficiency data of the 2D nanoplate-shaped  $\text{V}_2\text{O}_5$  electrode at the current density of  $0.5 \text{ A g}^{-1}$  for the first 15 cycles followed by  $5 \text{ A g}^{-1}$  for a further 1000 cycles. Since the electrode

delivers the initial discharge/charge capacity of  $\sim 180/141 \text{ mA h g}^{-1}$ , the capacity was maintained at  $\sim 44 \text{ mA h g}^{-1}$  after 1000 cycles. Significant capacity loss can be noticed after 15 cycles and the capacity reaches  $\sim 75 \text{ mA h g}^{-1}$ , which can be ascribed to the fact that the inserted  $\text{Zn}^{2+}$  ions could not be completely extracted from the  $\text{V}_2\text{O}_5$  lattice. These unextracted  $\text{Zn}^{2+}$  ions locate at the “dead  $\text{Zn}^{2+}$ ”

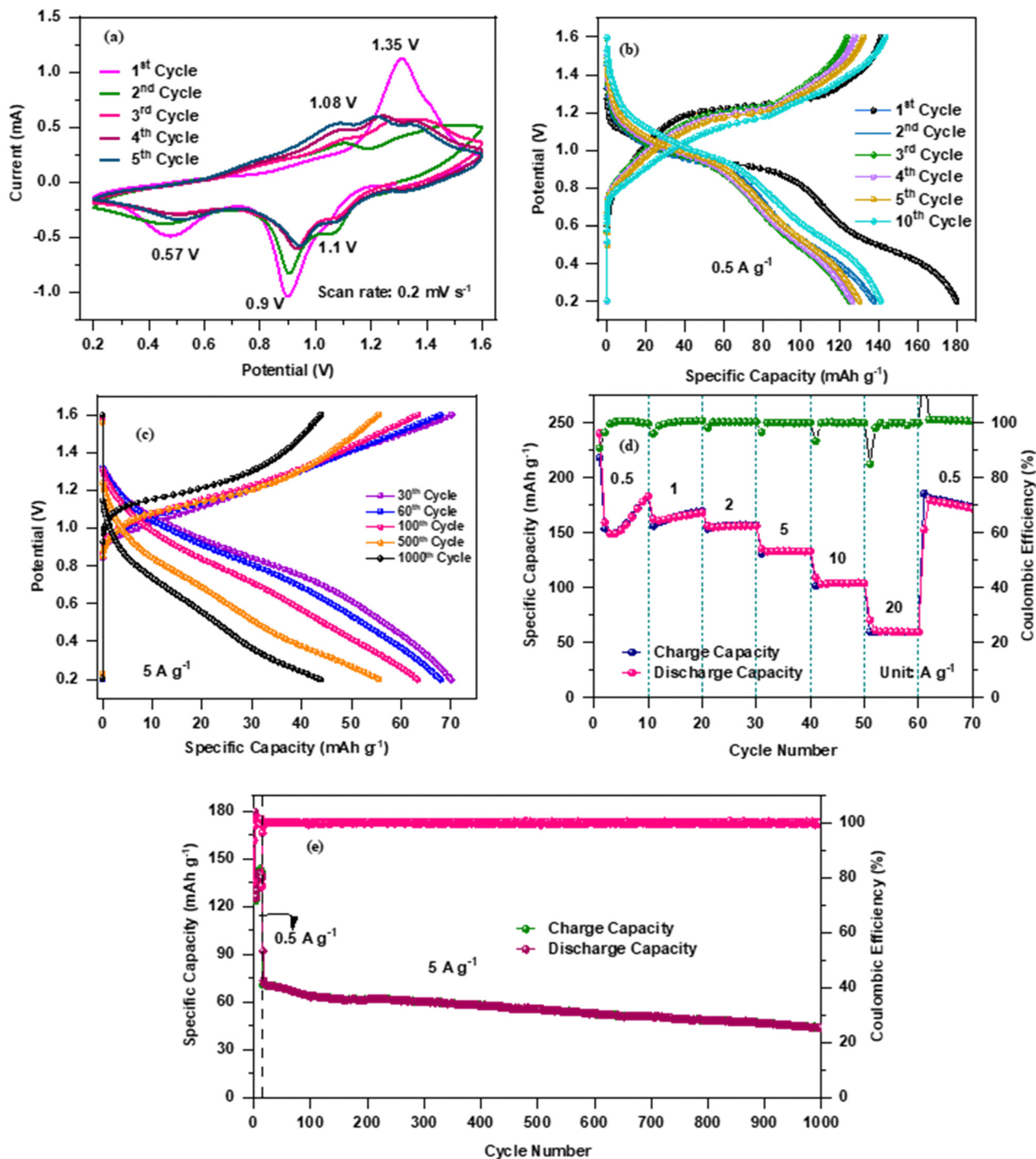


Fig. 3 (a) Cyclic voltammograms at a constant scan rate of  $0.2 \text{ mV s}^{-1}$ ; (b) discharge/charge curves at a current density of  $0.5 \text{ A g}^{-1}$  for the initial five cycles; (c) discharge/charge curves at a current density of  $5 \text{ A g}^{-1}$  for 1000 cycles; (d) C-rate performance at different current densities and (e) cycling performance at a constant current density of  $5 \text{ A g}^{-1}$  until 1000 cycles of the nanoplate-shaped  $\text{V}_2\text{O}_5$  electrode.



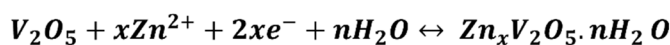
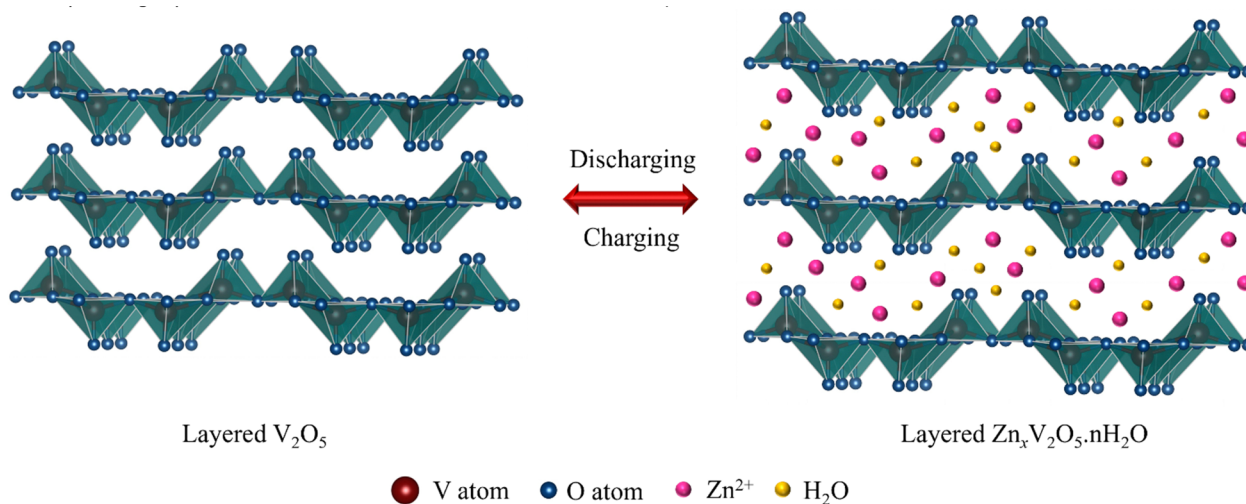


Fig. 4 Schematic illustration of the Zn-ion storage mechanism in the layered V<sub>2</sub>O<sub>5</sub> crystal system.

sites" making the layered crystalline structure stable with promising stable capacity until 1000 cycles.<sup>2,23</sup> The layered V<sub>2</sub>O<sub>5</sub> cathode permits a high extent of reversibility for hydrated Zn<sup>2+</sup> (de)-intercalation. It is believed that the co-insertion of H<sub>2</sub>O along with Zn<sup>2+</sup> ions causes structural evolution and changes the layered V<sub>2</sub>O<sub>5</sub> into a layered Zn<sub>x</sub>V<sub>2</sub>O<sub>5</sub>·nH<sub>2</sub>O. On the other hand, the subsequent charging process enables the reversal of the structural evolution followed by the extraction of hydrated Zn<sup>2+</sup> ions that have occurred during discharge. The co-intercalated H<sub>2</sub>O molecules act as a buffer against the high charge density of Zn<sup>2+</sup>, which facilitates quick cation transfer and high-rate capability. The reversible reaction of the Zn–V<sub>2</sub>O<sub>5</sub> battery is schematically illustrated in Fig. 4 with the corresponding equation.<sup>37</sup>

## Conclusions

Herein, we have fabricated a 2D nanoplate-shaped V<sub>2</sub>O<sub>5</sub> cathode material through a facile solvothermal method. The electrochemical investigation results clearly demonstrate that the V<sub>2</sub>O<sub>5</sub> electrode exhibits high specific capacity and better rate capability even at a high current density of 20 A g<sup>-1</sup>. The excellent performance of the V<sub>2</sub>O<sub>5</sub> electrode can be attributed to the unique features of the nanoplates *i.e.*, larger specific surface area and high mechanical stability, which increases the conductivity by shortening the diffusion paths for ions. The cathode delivers an initial charge capacity of ~141 mA h g<sup>-1</sup> at a current density of 0.5 A g<sup>-1</sup> and retains a capacity of ~44 mA h g<sup>-1</sup> after 1000 cycles at a current density of 5 A g<sup>-1</sup>. The obtained results conclude that the high performance of the V<sub>2</sub>O<sub>5</sub> nanoplates provides potential feasibility as a cathode material for ZIBs.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

One of the authors (Priya Yadav) would like to thank CSIR, New Delhi, for the award of a senior research fellowship (09/0045(11332)/2021-EMR-I). Dr Alok Kumar Rai is grateful to the University Grants Commission (UGC), India for the funding support under the UGC-FRP scheme (FRP ID: 57304). This work was mainly funded by the Council of Scientific and Industrial Research (CSIR) Research Grant (File No. 01(3015)/21/EMR-II) and partially supported by the National Research Foundation of Korea (NRF) grant funded by the Korea government (MSIT) (NRF-2018R1A5A1025224).

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