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Investigating nalidixic acid adsorption onto ferrihydrite and maghemite surfaces: molecular-level insights *via* continuous-flow ATR-FTIR spectroscopy†

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Contaminants sequestration by iron oxides are of particular relevance in environmental chemistry and Fourier transform infrared (FTIR) spectroscopy in attenuated total reflectance (ATR) mode is used in this study to investigate the molecular-scale adsorption mechanism of the antibiotic nalidixic acid (NAL) onto ferrihydrite and maghemite surfaces. High-surface area ferrihydrite and nano-sized maghemite were prepared as model adsorbents. This work provides a complete band assignment for protonated and deprotonated NAL to better describe *in situ* adsorption processes. *In situ* ATR-FTIR spectroscopy suggests that NAL mainly adsorbs on both iron oxides by a bidentate mode. Results also suggest that two sorption sites coexist on Maghemite surfaces leading to two sorption modes through the carboxylate group: bidentate chelate and bidentate bridging. The obtained results can contribute significantly towards molecular level understanding of antibiotics binding modes on iron oxides, a valuable piece of information to develop environmental remediation strategies.

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Introduction

Organic pollution has raised significant environmental concern due to the uncertain long-term fate of pollutants within natural media. Prolonged exposure to toxic substances, such as pesticides, pharmaceuticals, and industrial additives, can modify microbial ecology, thereby posing risks to aquatic life and human health.^{1,2} One critical process controlling the spreading of contaminants is chemical sorption, particularly at the surface of reactive and high surface area minerals that are widespread in natural environment.^{1,3–7} Hence, the use of nanotechnology in remediation processes has expanded with iron substrates being at the forefront due to their remarkable versatility, sorption capacity and organics affinity. As a matter of fact, chemical adsorption to natural or geo-inspired substrates emerge as a potential alleviation mechanism for mitigating pollution.⁸

Extensive studies have elucidated the binding mechanism of various adsorbate and adsorbent (solid substrate) couples.^{1,3,6,9–14}

In the 90s, more abundant and reactive anionic solutes easily found in groundwater were often targeted in sorption studies, such as sulfates, phosphates and carbonates,^{3,4,9,11,15} while in the last decade, pharmaceuticals gained relevance due to their higher toxicological relevance and appearance even in drinking water.^{2,16}

Quinolones, *i.e.* a class of antibiotics, have garnered attention in recent decades due to their toxicological concerns associated with antimicrobial resistance enhancement and public health issues stemming from incomplete metabolism.^{2,17} Among these, nalidixic acid (NAL), namely 1-ethyl-7-methyl-4-oxo-1,8-naphthyridine-3-carboxylic acid (IUPAC nomenclature), is a synthetic antibiotic that has been the object of several sorption studies.^{1,8,18–22} NAL consists of a bicyclic ring system that contains some groups attached to a naphthyridine, *i.e.* two pyridine rings fused, which provides unique pharmacological properties and antibacterial activity; more precisely, an ethyl group (–CH₂CH₃) at position 1, a carboxylic acid group (–COOH) at position 3, a keto-carbonyl (–C=O) at position 4 and a methyl group (–CH₃) at position 8 (Fig. 1). Its pK_a value is reported to be in the 5.8–6.2 range^{1,18,19,23–26} and due to the presence of the carboxylic acid group, NAL can also exist in a deprotonated form, the nalidixate anion (NAL[–]) above *ca.* pH 6. Its solubility in water is thus strongly dependent on pH, varying from 33 (pH 5) to 27 600 mg L^{–1} (pH 9).²⁶

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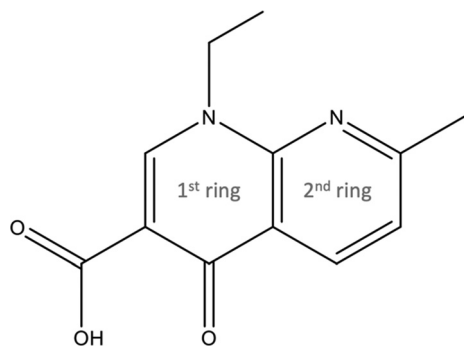


Fig. 1 The nalidixic acid structure.

The escalating presence of NAL in natural environments has elicited growing concerns due to its deleterious effects since it is recognized as toxic.²⁷ NAL has been detected in different environments all around the world: in river sediments downstream of Paris²⁸ (with a maximum concentration of $22 \mu\text{g kg}^{-1}$), in municipal waste water treatment plants effluents and rivers samples collected in Queensland, Australia¹⁶ (with maximum concentrations of 0.45 and $0.75 \mu\text{g L}^{-1}$) and in hospital effluents samples in Taiwan²⁹ (maximum concentration of $67 \mu\text{g L}^{-1}$).

In situ spectroscopy is vigorously favored to reach a molecular-scale understanding of sorption phenomena in order to fully integrate the effect of surface hydration on the geometry of sorption complexes,¹¹ mimicking colloidal interfaces in natural settings that typically retain water. Consequently, *in situ* studies mitigate the problems usually associated with *ex situ* investigations, such as structural modifications caused by drying.^{9,11} IR spectroscopy in attenuated total reflectance (ATR) mode arises as a key technique to perform *in situ* studies of chemical sorption by exploiting the principles of total reflection.³⁰ Its short path lengths reduce solvent interference, enabling the study of sorption modes even in presence of strongly IR-light absorbing media, such as water.

This work explores quinolone, specifically NAL, binding modes onto maghemite, an iron oxide $\gamma\text{-Fe}_2\text{O}_3$, and ferrihydrite, a hydrous ferric oxyhydroxide (used as model of iron oxides ubiquitous in natural environments) under dry and aqueous condition with IR spectroscopy in ATR mode to conduct *in situ* studies to discuss the molecular interactions at solid-liquid and solid-air interfaces. Maghemite was chosen since it is a direct oxidation product of magnetite and zero-valent iron³¹ and has a well-known surface reactivity and affinity for inorganic and organic pollutant sorption.³² Maghemite can exhibit high surface area when prepared in nano-sized form, with potential use as sorbent for different purposes, such as water

decontamination. Moreover, high-surface area ‘two-lines’ ferrihydrite was chosen as adsorbent because it plays a substantial role in the sequestration of contaminants from groundwater due to its abundance in natural media.³³

Materials and methods

Adsorbents synthesis and characterization

Maghemite, $\gamma\text{-Fe}_2\text{O}_3$. Nano-maghemite (Mh) was prepared by oxidizing synthetic nano-magnetite obtained by co-precipitating Fe(II) and Fe(III) in aqueous solutions at alkaline pH.^{8,34} For this purpose, a starting solution containing 2:1 parts of ferric and ferrous 1 M chloride solutions was prepared. By slowly adding 1 M NaOH solution, the pH was raised to 12 and the solution was then stirred for 24 hours. The synthesis was performed in a Jacomex glovebox under N_2 atmosphere (<10 ppm of O_2) with O_2 -free deionized water ($18.2 \text{ M}\Omega \text{ cm}$; Milli-pore Milli-Q system and degassed by bubbling N_2 for 45 min at 85°C) to avoid Fe(II) oxidation, improving the yield and quality of the synthesized nano-particles. Final coprecipitate was centrifuged, washed three times and vacuum dried. A black fine powder was recovered and then heated at 200°C for 18 h in open-air.⁸ A dark-brown powder was finally recovered.

Two-lines ferrihydrite. Two-lines ferrihydrite (Fh), with a typical formula $\text{Fe}_{10}\text{O}_{14}(\text{OH})_2 \cdot n\text{H}_2\text{O}$,³⁵ was synthesized according to the classical alkali precipitation method.³⁶ A 0.2 M nitrate ferric solution was adjusted to pH 7–8 by adding KOH solution. The solution was vigorously stirred and the precipitate was centrifuged, rinsed-centrifuged three times with Milli-Q water and air-dried. A red-brown powder was recovered after grinding in an agate mortar.

Mineralogical characterization. Selected properties for synthetic adsorbent samples are summarized in Table 1. Mineralogical purity of our synthetic samples was checked by powder X-ray diffraction (XRD) using a Panalytical X'pert Pro diffractometer (see Fig. S1 in ESI[†]). Mean-coherent-domain size (MCD) of Mh was calculated using the Scherrer formula from line-width values determined by Rietveld refinement performed with the XND code³⁷ assuming isotropic crystallite shape (Fig. S1, ESI[†]). Surface area (SA) was measured by nitrogen physisorption at 77 K and determined with the Brunauer–Emmett–Teller (BET) theory on samples pre-treated at 120°C under vacuum overnight.

ATR-FTIR measurements

A Bruker Vertex 80 spectrometer constantly purged with N_2 gas and equipped with mercury cadmium telluride (MCT, liquid

Table 1 Properties of the synthetic adsorbents used in this study. Mean-coherent domain size (MCD) of Mh was determined using Rietveld refinement of the XRD pattern. Surface area (SA) was determined using the BET method (see text). Refraction index (n) and point of zero charge (PZC) are reported based on literature data

Adsorbent	Formula	MCD	SA [$\text{m}^2 \text{g}^{-1}$]	Refraction index [n]	Point of zero charge
Maghemite, Mh	$\gamma\text{-Fe}_2\text{O}_3$	$12.7 \pm 0.2 \text{ nm}$	49	$2.54\text{--}2.73^{38}$	7.0^{39}
Ferrihydrite, Fh	$\text{Fe}_{10}\text{O}_{14}(\text{OH})_2 \cdot n\text{H}_2\text{O}$	Typically 2 nm^{40}	275	$2.24\text{--}2.32^{41,42}$	7.9^{43}

nitrogen) detector was used. It has to be noted that the detector is operational (and experiments are possible) up to 24 h after cooled down before being recharged with liquid nitrogen. Spectra were recorded in the rapid-scan mode with a 160 kHz speed and using the double-sided forward-backward interferogram recording. A flow-through horizontal attenuated total reflection (HATR) accessory from PIKE Technologies was used. The accessory is composed of a rectangular stainless-steel base, supplied with 10 reflections flat-long ($70 \times 10 \times 4$ mm) zinc selenide crystal (45° ZnSe), with a screwable metallic lid and two threaded-to-female Luer ports. Depending on the interface involved – varying between solid-liquid or solid-air-, background was differently collected by averaging 2000 scans. For static experiments, *i.e. ex situ*, the background spectrum was collected with the mineral substrate over the ZnSe crystal only. For dynamic experiments, *i.e. in situ*, the background was collected with the mineral substrate over the ZnSe crystal and with a constant aqueous flow circulating over it. For static measurements, samples were recorded with 2000 scans, while for dynamic experiments, kinetics was monitored by collecting 500 scans per minute and averaged by groups of 4 (*i.e.* 2000 scans total). Resolution was set at 4 cm^{-1} and data analysis was performed with Origin 2018 (OriginLab Corporation) and Python 3.10.11 open-source.

Nalidixic acid (NAL). NAL in analytical grade ($\text{C}_{12}\text{H}_{12}\text{N}_2\text{O}_3$, 98.5%) was obtained from Sigma-Aldrich and analyzed both in solid-state and in solution varying the pH. Clean ZnSe crystal was used as background for NAL solid-state characterization. NAL in polycrystalline powder form exhibiting white to slightly yellow color was directly deposited onto the ZnSe crystal. For $\text{NAL}_{(\text{s})}$ analysis, 5 drops of water were randomly spread onto the powder before drying overnight to assure homogeneous spreading of the powder onto the crystal before collecting the IR spectrum. Clean ZnSe crystal covered with carbonate-free water at the same pH was used as background for aqueous state characterization. For $\text{NAL}_{(\text{aq})}$ measurements, the quinolone powder was dissolved in water with a 0.1 mM concentration and pH adjusted to 5 (*i.e.* $\text{pH} < \text{pK}_a$) by adding 4 drops of HCl 0.1 M. Solutions were then loaded into the HATR cell and FTIR measurements were quickly performed to avoid pH fluctuations since no buffer was used.

Nalidixate anion (NAL^-). The NAL^- solutions were prepared by dissolving quinolone powder in water with concentrations ranging from 0.1 to 2 mM and adding 4 to 10 drops of NaOH 0.1 and/or 1 M to adjust the pH to 9 (*i.e.* $\text{pH} > \text{pK}_a$). The NAL^- solutions were analyzed by ATR-FTIR in the same manner that the $\text{NAL}_{(\text{aq})}$ one. The solution containing 2 mM was used as stock solution for following adsorption experiments, and thus was protected from light and stored at 4°C .

Layer deposition. To conduct *in situ* ATR-IR adsorption experiments effectively, it is crucial to uniformly deposit thin layers of mineral substrates with sufficient adherence to the internal reflection element (*i.e.* IRE), which in this study is ZnSe. Attachment is important since the aim is to submit this deposit to a constant aqueous flow, while IR spectra are recorded over time. Mineral layers were prepared by spreading

a sonicated mineral suspension (containing adsorbent material, either Fh or Mh) over the flat-long ATR crystal and letting it dry overnight as described elsewhere.⁴⁴ Suspensions of 2 mg mL^{-1} (in equal parts of ethanol and pure water) were used targeting a final amount of 2.4 mg iron(oxy)(hydr)oxide coatings. The crystal surface was previously gently rubbed with Q-tips soaked with isopropanol and dried under N_2 flow to assure cleanliness.

Static sorption experiments. For static measurements, quinolone is impregnated onto the substrate spread over the ATR crystal to reach a surface loading of 2 molecules per nm^2 . On Mh a total volume of 15.2 mL solution with a concentration of 0.026 mM was added in 10 steps due to the limited volume of the ATR cell (*i.e.* 1.5 mL). In between each step, the sample was vacuum-dried for about 20 minutes, removing most of the solvent. The final 1.5 mL added was left to dry at room temperature overnight. On Fh, a total volume of 16.1 mL solution with a concentration of 0.137 mM was used in similar manner. The total volumes used were chosen to mimic the volumes used in the dynamic adsorption (see below); both prepared diluting the nalidixate $\text{NAL}^-_{(\text{aq})}$ stock solution (2 mM with pH 9) at room temperature in pure water, aiming to reach neutral pH after equilibrium (later detailed for each experiment). Note that the stock solution is quickly filtered (Filter HVLP type Millipore $0.45 \mu\text{m}$) to remove potential traces of solid before used.

Dynamic sorption experiment. For these measurements, a set-up is required to properly investigate the solid-liquid interface. We implemented an attachable re-circulating system to the FTIR-ATR cell, which is composed of four parts as illustrated in Fig. S2 in ESI.† In this approach, the nalidixate $\text{NAL}^-_{(\text{aq})}$ stock solution (2 mM with pH 9) at room temperature is quickly filtered (in the same manner as for the static approach) and directly injected into a stable re-circulating water flow *via* Flask A. No background electrolyte was added in order to avoid any competition with putative outer-sphere surface complexes of $\text{NAL}^-_{(\text{aq})}$. The total water volume circulating is $15.0 \pm 0.1 \text{ mL}$, on which a variable volume of $\text{NAL}^-_{(\text{aq})}$ is added (in a 0.2–1.1 mL range) depending on substrate surface area – to reach a target surface coverage of 2 molecules nm^{-2} considering full adsorption. For Mh sorption, a final 0.026 mM solution containing quinolone was used within the re-circulating system, while for Fh, a 0.137 mM solution was used. A magnetic stirrer was added in the mixing chamber to ensure flux homogeneity, and a pH electrode in the laminar chamber to continuously monitor pH. The circulation was supported by a Gilson mini-plus peristaltic pump under a 0.35–0.45 mL min^{-1} flux with unidirectional counterclockwise motor.

Density functional theory calculation of IR frequencies. Density functional theory (DFT) calculations were performed using the ORCA software in the gas-phase,⁴⁵ thus not considering the dielectric constant of the surrounding medium. We employed the approximation B3LYP⁴⁶ to the exchange-correlation term and the Gaussian basis sets def2-TZVPP.⁴⁷ IR frequencies obtained were scaled with a factor 0.965 to counter the inaccuracy within the approximation B3LYP.

Results

NAL and NAL^- vibrational band assignments

ATR-FTIR spectra of solid-state NAL, aqueous state $\text{NAL}_{(\text{aq})}$ (pH = 5 *i.e.* pH < $\text{p}K_{\text{a}}$) and $\text{NAL}^-_{(\text{aq})}$ (pH = 9 *i.e.* pH > $\text{p}K_{\text{a}}$) are shown in Fig. 2. These spectra provide valuable information about the molecular structure of protonated and deprotonated quinolone molecules and thus serve as a basis for further assigning vibrational bands of surface-bounded NAL species.

Solid-state NAL. Experimental FTIR bands for $\text{NAL}_{(\text{s})}$ spectrum displayed in Fig. 2a are summarized in Table 2. The experimental values have been compared with literature data and with DFT calculations for a full band assignment. The most intense calculated bands (DFT) have been highlighted in bold in Table 2 and the experimental spectra are compared to calculated band positions and intensities in Fig. S3 in ESI.† Note that among the references, studies using other fluoroquinolones, such as ciprofloxacin (CIP) and ofloxacin, were also used due to their spectral similarities linked to quinolone nucleus and the presence of a carboxylic acid group.^{48,49} Additionally, when it comes to literature using NAL, some inconsistencies have been found concerning the carboxyl (COOH) and carbonyl (C=O) band assignments. For instance, bands at 1715–1712 cm^{-1} have been sometimes attributed to keto-carbonyl vibrational mode instead of carboxyl.^{1,22,50} Also, bands at 1620–1614 cm^{-1} have been occasionally wrongly attributed to $\nu_{\text{C}=\text{C}}$ vibration of the aromatic rings instead of keto-carbonyl.^{1,22,24,50} DFT calculations show that each vibrational mode involves simultaneous movement of several atoms/bonds, which might explain literature discrepancies in their

assignments. However, a better description of each vibrational mode is fundamental for further discussion.

The strong peak at 1710 cm^{-1} in $\text{NAL}_{(\text{s})}$ is attributed to C=O carboxylic group stretching (*i.e.* $\nu_{\text{C}=\text{O}}$ carboxyl). The equivalent DFT calculated value is 1736 cm^{-1} and a higher frequency is consistent since intermolecular interactions are not taken into account with DFT. The following experimental strong peak is centered at 1616 cm^{-1} in $\text{NAL}_{(\text{s})}$, which is largely identified in literature as the stretching of the C=O attached to the naphthyridine (*i.e.* $\nu_{\text{C}=\text{O}}$ carbonyl). Our DFT calculations support this assignment, despite a downshift in the calculated wavenumber. This downshift can be due to the fact that in gas-phase, the OH from the carboxyl and the keto-carbonyl of NAL (see Fig. 1) have a strong intramolecular hydrogen bond OH–O=C interaction inducing the formation of a particularly stable 6-membered ring. This interaction is probably somehow modified in solid-state NAL, due to the presence of intermolecular forces. Other bands and their proposed assignments are reported in Table 2.

Aqueous NAL and NAL^- species. To help further spectral interpretation of NAL sorption onto iron (oxyhydr)oxides, we also address ATR-FTIR spectra of the aqueous phase containing the NAL species as a function of pH. Both $\text{NAL}_{(\text{aq})}$ and $\text{NAL}^-_{(\text{aq})}$ spectra, see Fig. 2b and c, showed reduced absorbance when compared to solid-state NAL, which is expected due to their low concentration, and consequently have been magnified for better visualization.

The 0.1 mM $\text{NAL}_{(\text{aq})}$ spectrum at pH 5 is displayed in Fig. 2b. It is important to remind that NAL solubility is limited in water at this pH (only 0.142 mM^{26}), so that it was not possible to

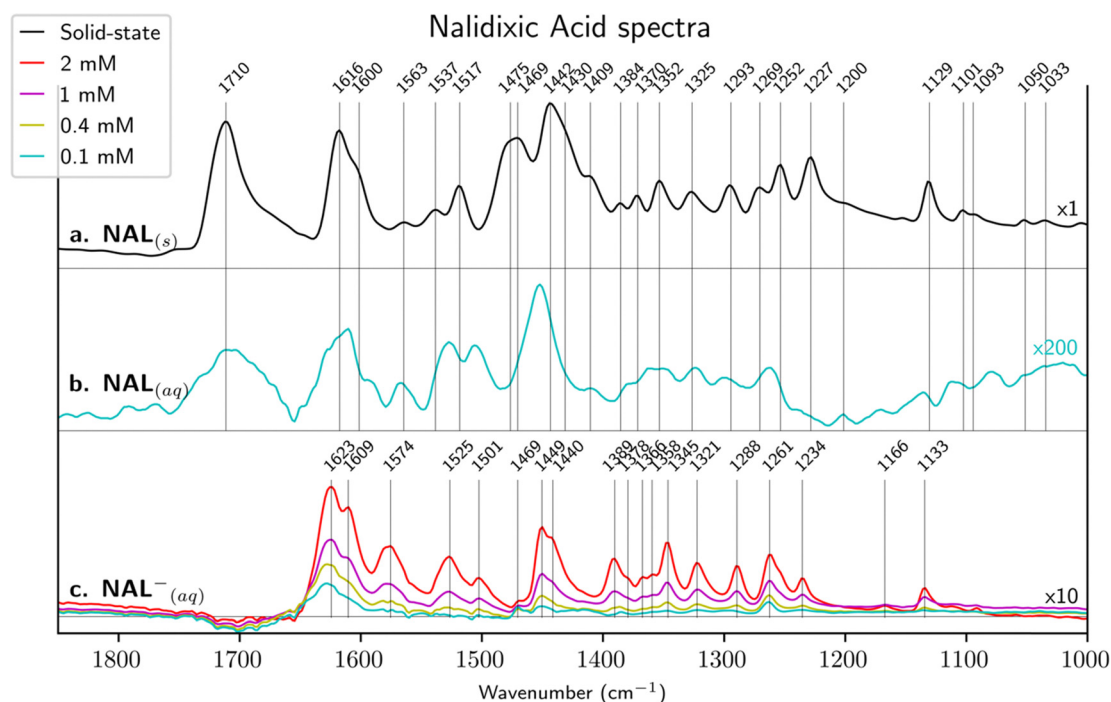


Fig. 2 ATR-FTIR spectra of (a) solid-state nalidixic acid ($\text{NAL}_{(\text{s})}$), (b) aqueous nalidixic acid ($\text{NAL}_{(\text{aq})}$, pH 5) and (c) aqueous nalidixate ($\text{NAL}^-_{(\text{aq})}$, pH 9). Aqueous species spectra have been magnified and vertically shifted for clarity.

Table 2 FTIR band assignments for NAL_(s) and NAL_(aq), comparing experimental data with published literature data and with DFT calculations of NAL in gas phase. Stretching modes are noted as ν and deformation as δ . The most intense calculated bands have been highlighted in bold

Experimental position [cm ⁻¹]		Assignment based on literature		Assignment based on DFT	
NAL _(s)	NAL _(aq)	Vibrational mode	Range [cm ⁻¹]	Vibrational mode (the main bands involved for each mode are indicated)	DFT Position [cm ⁻¹]
1710	1709	$\nu_{\text{C=O}}$ carboxyl	1720–1706 ^{14,23,24,48,49,51–53}	$\nu_{\text{C=O}}$ carboxyl	1736
1616	1610	$\nu_{\text{C=O}}$ carbonyl	1640–1620 ^{48,51–53}	$\nu_{\text{C=O}}$ carbonyl , $\delta_{\text{C-H}}$ 1st ring	1596
1600	—	—	—	$\nu_{\text{C-C}}$ and $\nu_{\text{C=N}}$ rings	1585
1563	1567	$\nu_{\text{C=C}}$	1560–1544 ^{50,54}	$\nu_{\text{C=O}}$ carbonyl , $\delta_{\text{C-H}}$ 1st ring	1560
1537	1526	—	—	$\nu_{\text{C-C}}$ and $\nu_{\text{C=N}}$ rings	1522
1517	1505	$\nu_{\text{C=C}}$ OR $\nu_{\text{C=N}}$	1518–1519 ^{50,54}	$\delta_{\text{O-H}}$, ethyl, methyl, $\delta_{\text{C-H}}$ 2nd ring	1478
1475	1483	—	—	$\delta_{\text{C-H}}$ ethyl	1453
1469	—	$\nu_{\text{C=N}}$	1470–1473 ^{50,54}	Mainly $\delta_{\text{O-H}}$	1443
1442	1451	$\nu_{\text{C=N}}$	1444 ⁵⁴	$\delta_{\text{C-H}}$ ethyl	1440
—	—	—	—	$\delta_{\text{C-H}}$ methyl, ethyl	1434
1430	—	—	—	$\delta_{\text{C-H}}$ methyl, ethyl	1433
—	—	—	—	$\delta_{\text{C-H}}$ methyl	1425
1409	—	$\nu_{\text{C=N}}$	1409 ⁵⁰	Delocalized	1416
1384	—	$\delta_{\text{C-N-H}}$ OR $\nu_{\text{C=N}}$	1384–1385 ^{50,54}	$\delta_{\text{C-H}}$ ethyl	1368
1370	—	—	—	$\delta_{\text{C-H}}$ mainly methyl	1364
1352	—	$\delta_{\text{C-N-H}}$ OR $\nu_{\text{C-C}}$	1353 ^{50,54}	$\delta_{\text{C-H}}$ ethyl, methyl	1358
—	—	—	—	$\delta_{\text{C-H}}$ ethyl, $\delta_{\text{C-H}}$ 1st ring	1340
1325	—	$\nu_{\text{C-OH}}$	1327 ⁵⁰	70% $\delta_{\text{C-H}}$ ethyl	1335
1293	—	$\nu_{\text{C-OH}}$ and $\delta_{\text{C-OH}}$	1300–1260 ^{48,50}	70% $\delta_{\text{C-H}}$ ethyl	1309
1269	1264	Coupled mode $\nu_{\text{COOH}}/\delta_{\text{C-OH}}$	1258 ¹⁸	$\delta_{\text{C-H}}$ 1st ring, 15% $\delta_{\text{O-H}}$	1267
1252	—	δ_{CH}	1253–1228 ⁵⁰	70% ethyl	1253
1227	—	—	—	$\delta_{\text{C-H}}$ 1st ring	1230
1200	—	—	—	$\delta_{\text{C-H}}$ 2nd ring	1212
—	—	—	—	δ_{OH} , $\nu_{\text{C-COOH}}$, $\delta_{\text{C-H}}$ 2nd ring	1183
1129	—	$\delta_{\text{C-CH}}$ asy and sy	1130 ⁵⁴	$\delta_{\text{C-H}}$ rings, ethyl	1132
—	—	—	—	Torsion of the ethyl group and bending of C–H in rings	1112
1101	—	$\delta_{\text{C-CH}}$ OR $\nu_{\text{C-C}}$	1102 ^{50,54}	$\delta_{\text{C-H}}$ rings, Ethyl	1079
1093	—	—	—	70% $\delta_{\text{C-H}}$ ethyl	1071
1050	—	$\delta_{\text{C-CH}}$ asy and sy	1051 ^{50,54}	Delocalized	1027
1033	—	$\delta_{\text{C-C-H}}$ OR $\nu_{\text{C-C}}$	1035–1020 ^{50,54}	$\delta_{\text{C-H}}$ methyl	1022

increase further the concentration in order to improve the signal-to-noise ratio. By comparing NAL_(s) and NAL_(aq), many similarities are observable, and band assignments are consistent with those summarized in Table 2. Among modest differences, we may mention that the keto-carbonyl stretching slightly downshifts from 1616 to 1610 cm⁻¹. This and other shifts are likely to be ascribed to intermolecular interactions with surrounding water molecules.

The NAL_(aq)⁻ spectra at pH 9 for various concentrations are displayed in Fig. 2c and peak positions and assignments are summarized in Table 3. The alkaline conditions (pH > pK_a) lead to deprotonation of the carboxylic acid, inducing structural changes and favoring a resonance-stabilized structure as well as a much higher solubility (about 0.1 M²⁶). As previously, band assignments have also been compared with literature data and with results obtained from DFT calculations. The most intense calculated bands (DFT) have been highlighted in bold in Table 3 and the experimental spectra are compared to calculated band positions and intensities in Fig. S3 in ESI.† It is important to note that, to the best of our knowledge, only partial assignment^{18,23} has been provided for nalidixate in the literature up to now.

Deprotonation of NAL leads to the formation of a carboxylate moiety (COO⁻) resulting in the absence of the carboxyl (COOH) centered at 1710 cm⁻¹. In some previous studies

(see Table 3), bands assigned to the carboxylate group are located at about 1580 and 1390 cm⁻¹ for the carboxylate asymmetric and symmetric stretching modes respectively. However, DFT calculations performed in the present work disagree in part with those assignments. Asymmetric stretching of the carboxylate is found to contribute to both 1650 and 1614 cm⁻¹ bands, while the symmetric stretching is found to contribute to the 1276 cm⁻¹ band. In this scenario, both bands reflecting the asymmetric stretching overlaps with the position classically identified in the literature^{18,23,49,55,56} to the stretching mode of the keto-carbonyl (at 1623 cm⁻¹). This discrepancy can be explained by the conjugation of the carboxyl and carbonyl moieties in the quinolone molecule, so that the 1623 and 1609 cm⁻¹ bands (1650 and 1614 cm⁻¹ in DFT calculations) both reflect contributions from the carbonyl C=O and the COO⁻ asymmetric stretching. In contrast, the symmetric COO⁻ stretching mode found at 1276 cm⁻¹ with DFT can be correlated to the experimental bands at 1288 or 1261 cm⁻¹ (see Fig. S3, ESI†). However, Fig. 3 (see below) shows that the relative intensity of the band at 1262 cm⁻¹ is strongly modified after NAL⁻ adsorption, which suggests an implication of this vibrational mode upon adsorption. Consequently, the band at 1261 cm⁻¹ for NAL_(aq)⁻ in Table 3 is assigned to the symmetric COO⁻ stretching mode. Another important band to mention is the one centered at 1449 cm⁻¹, which is known to remain stable under pH modification and

Table 3 FTIR band assignments for aqueous nalidixate ($\text{NAL}^-_{(\text{aq})}$), comparing experimental data with data available in published literature and with DFT calculations of the anion in gas phase. Stretching modes are noted as ν and deformation as δ . The most intense calculated bands have been highlighted in bold

Experimental position [cm^{-1}]	Assignment based on literature		Assignment based on DFT	
	Vibrational mode	Range [cm^{-1}]	Vibrational mode	DFT Position [cm^{-1}]
$\text{NAL}^-_{(\text{aq})}$				
1623	$\nu_{\text{C=O}}$ carbonyl	1640–1620 ^{18,23,49,55,56}	$\nu_{\text{C=O}}$ carbonyl , ν_{COO^-} as	1650
1609	—	—	$\nu_{\text{C=O}}$ carbonyl , ν_{COO^-} as	1614
1574	ν_{COO^-} as	1586–1578 ^{14,18,23,48,49,53,55,56}	$\nu_{\text{C-C}}$ 1st ring , $\nu_{\text{C-H}}$	1574
1525	$\nu_{\text{C-C}}$	~1540–1528 ^{18,23,49,56}	$\nu_{\text{C-C}}$ and $\nu_{\text{C=N}}$ 2nd ring , $\delta_{\text{C-H}}$	1554
1501	$\nu_{\text{C-C}}$ or $\delta_{\text{C-H}}$	~1500–1498 ^{18,23,49}	$\nu_{\text{C-C}}$ and $\nu_{\text{C=N}}$ rings	1523
1469	$\nu_{\text{C-C}}$	~1470 ⁴⁹	$\delta_{\text{C-H}}$ ethyl	1454
1449	$\nu_{\text{C-C}}$ or $\delta_{\text{C-H}}$	1448 ²³	$\delta_{\text{C-H}}$ ethyl	1450
—	—	—	$\delta_{\text{C-H}}$ ethyl	1438
1440	—	—	$\delta_{\text{C-H}}$ ethyl, methyl	1436
—	—	—	$\delta_{\text{C-H}}$ ethyl, methyl	1429
—	—	—	$\delta_{\text{C-H}}$ methyl	1426
1389	ν_{COO^-} s	1395–1380 ^{14,18,23,48,49,53}	$\delta_{\text{C-H}}$ ethyl , $\delta_{\text{C-H}}$ 2nd ring	1412
1378	—	—	$\delta_{\text{C-H}}$ methyl and ethyl	1360
1366	—	—	$\delta_{\text{C-H}}$ methyl	1357
1358	—	—	$\delta_{\text{C-H}}$ ethyl	1351
1345	—	—	$\delta_{\text{C-H}}$ 1st ring	1339
1321	—	—	$\delta_{\text{C-H}}$ ethyl	1320
1288	δ_{NH^+} , δ_{COO^-}	1300–1290 ⁴⁸	$\delta_{\text{C-H}}$ ethyl	1289
1261	—	—	ν_{COO^-} s, $\nu_{\text{C-C}}$	1276
1234	—	—	$\delta_{\text{C-H}}$ ethyl, methyl	1268
—	—	—	$\delta_{\text{C-H}}$ 1st ring	1225
—	—	—	$\delta_{\text{C-H}}$ 2nd ring	1211
1166	—	—	$\delta_{\text{C-H}}$ both rings	1194
1133	—	—	$\delta_{\text{C-H}}$ 2nd ring	1116

consequently has been commonly used for spectra normalization in the literature.^{23,24} Previous works^{18,23} have assigned this band to ring stretching or C–H bending. Present DFT calculations suggest that the band at 1449 cm^{-1} is due to a bending mode of the ethyl group.

Sorbed NAL vibrational band assignments

IR spectra of sorbed NAL onto Mh and Fh surfaces are displayed in Fig. 3. In order to normalize the spectra of the sorbed species, we have used the $1450 \pm 3 \text{ cm}^{-1}$ band, which is mainly due to $\delta_{\text{C-H}}$ of the ethyl group (as suggested by DFT calculations). In this comparison, different interfaces are probed, solid–air (air-dried samples) for the static approach (Fig. 3a) and solid–liquid (aqueous conditions) for the dynamic one (Fig. 3b). In both cases, sorbed spectra are compared to the spectrum of the $\text{NAL}^-_{(\text{aq})}$ stock solution, which has been only multiplied 3-fold for clarity (not normalized). The intensity of the different spectra illustrates the surface-sensitivity of the ATR-FTIR technique. The intensity of IR bands assigned to the NAL^- species present considerably higher intensities when adsorbed than when dissolved in aqueous media (*i.e.* $\text{NAL}^-_{(\text{aq})}$) despite a lower number of molecules present in the HATR cell for the solid samples. For instance, by comparing Fh- NAL^- with $\text{NAL}^-_{(\text{aq})}$ the respective amount of quinolone varies from $2.2 \mu\text{mol}$ (16.1 mL of 0.136 mM solution) to $3 \mu\text{mol}$ (1.5 mL of a 2 mM solution). The difference in intensity between the sorption results and the $\text{NAL}^-_{(\text{aq})}$ arises mainly from the concentration of the adsorbed species on the absorbent layer deposited on the ATR crystal.

Static adsorption. The absence of a band at 1710 cm^{-1} confirms that static sorption occurs with the deprotonated

NAL molecule in agreement with the pH of the solutions used –6.3 and 6.9 for Mh and Fh, respectively. Spectra for the sorbed species are similar to the deprotonated molecule (Fig. 2c, $\text{NAL}^-_{(\text{aq})}$ at pH 9), presenting only slight shifts once adsorption takes place. The strongest peak upshifts from 1623 cm^{-1} to 1628 and 1632 cm^{-1} – respectively at Mh and Fh surfaces – and shows an extra shoulder at about 1650 cm^{-1} . The band upshift has been extensively reported, even with other quinolones,^{14,48,49} while the broadening has been less discussed, despite also being observed.¹⁴ These spectral modifications might be a result of a H-bonding involvement or an interaction induced by the deprotonated carboxyl group.^{14,55} This particular band has been largely assigned as the keto–carbonyl stretching (*i.e.* $\nu_{\text{C=O}}$ carbonyl) but according to our DFT calculations, this vibration is coupled with asymmetric stretching of the carboxylate as already discussed. Therefore, this band arises both from the keto–carbonyl stretching and from the asymmetric stretching of the carboxylate (*i.e.* $\nu_{\text{C=O}}$ carbonyl and ν_{COO^-} as). The shoulder at 1609 cm^{-1} observed in the $\text{NAL}^-_{(\text{aq})}$ is present after static sorption but slightly upshifted to 1612 cm^{-1} on both substrates. At higher wavenumbers, a shoulder arises around 1650 cm^{-1} during adsorption – more evidenced at Mh than in Fh-, which can also be observed in other sorption study¹⁸ but is not discussed. As far as the 1574 and 1389 cm^{-1} band are concerned – identified in literature as carboxylate vibrational modes (see Table 3) –, they shift to 1580 and 1396 cm^{-1} for Mh- NAL^- and to 1567 and 1391 cm^{-1} for Fh- NAL^- . Modifications of the asymmetric carboxylate band is explained by a distortion of the carbon angle (*i.e.* C–O bond lengths), which depends on the absorbent involved and might upshift and downshift^{48,57} or even

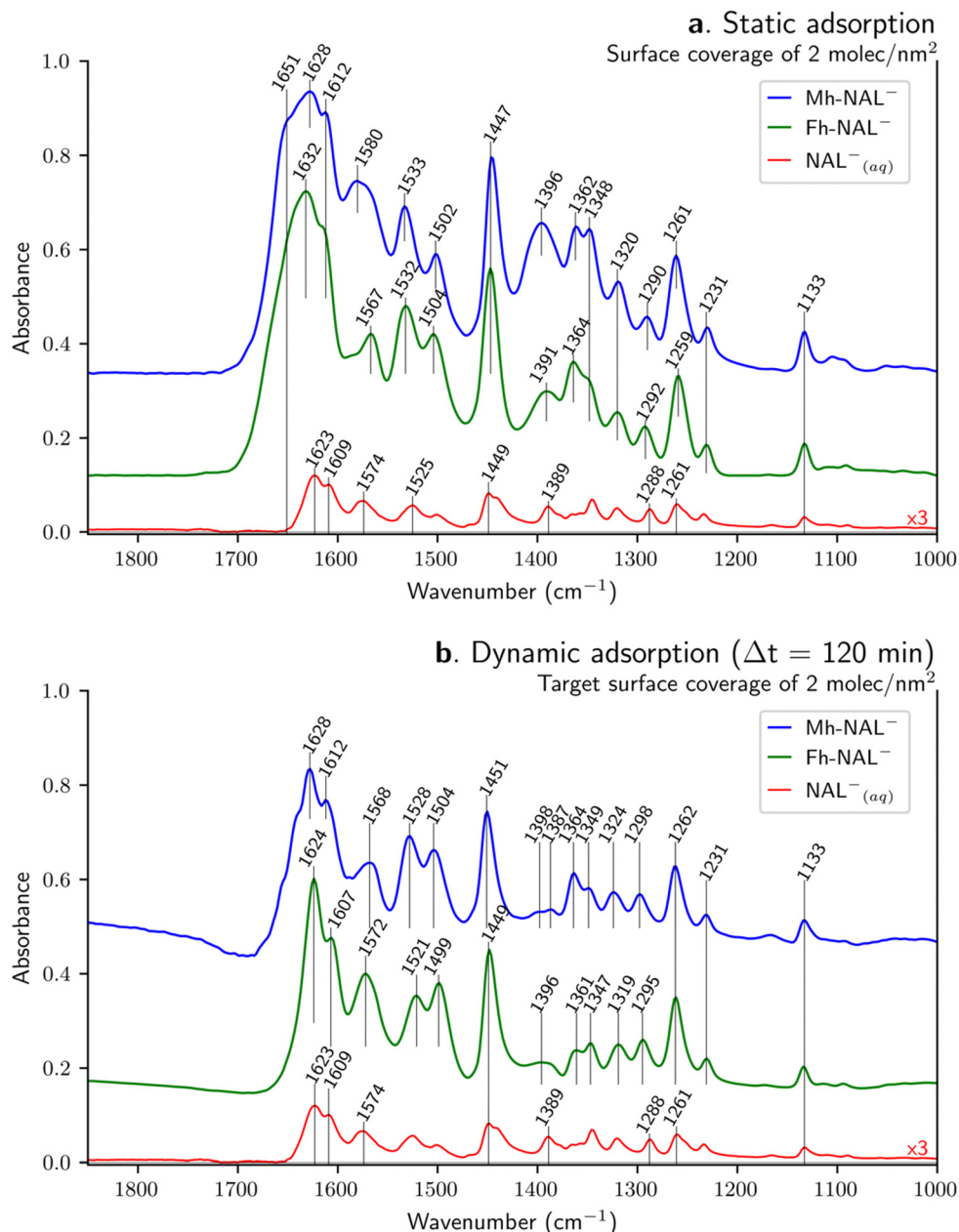


Fig. 3 ATR-FTIR spectra for adsorbed NAL onto Mh and Fh (at neutral pH) compared to $\text{NAL}^-_{(aq)}$ stock solution (2 mM with pH 9). The target surface loading of NAL is 2 molecules nm^{-2} performed via (a) static and (b) dynamic (after 120 minutes circulating) approaches. NAL spectra sorbed on Mh are normalized with respect to Fh with the intensity of the peak at about 1450 cm^{-1} . Spectra have been vertically shifted for clarity.

not be affected by adsorption.²³ Differences concerning the symmetric carboxylate band suggest molecular interaction of this group with surface hydroxo-groups with various hydrogen-bonds strengths, leading to an inhomogeneous band broadening,²³ while the upshifts can be attributed to the formation of a surface bonded complex.^{49,58} However, it was detailed above that our DFT calculations on isolated NAL^- disagree with these assignments (1574 cm^{-1} for ν_{COO^-} as and 1389 cm^{-1} for ν_{COO^-} s). DFT calculations suggest that the asymmetric carboxylate stretching for NAL^- is at 1623 and 1609 cm^{-1} while the symmetric COO^- stretching is at 1261 cm^{-1} . Thus, once adsorption takes place,

the asymmetric COO^- bands upshifts to $1628\text{--}1612\text{ cm}^{-1}$ for Mh-NAL^- and to $1632\text{--}1612\text{ cm}^{-1}$ for Fh-NAL^- , while the symmetric band is weakly affected. In this scenario, the asymmetric COO^- bands might upshift upon carbon angle distortion, as already pointed out, while being simultaneously affected by possible H-bonding involvement due to the conjugated keto-carbonyl stretching mode. Following DFT calculations, the bands previously assigned to ν_{COO^-} (as) and ν_{COO^-} (s) at around 1580 and 1390 cm^{-1} are reassigned to C–C and C–H stretchings as well as C–H bending, coming mainly from the rings and might upshift due to complexation. Lastly, concerning the C=C

ring stretching mode, the shift from 1525 cm^{-1} to 1533 or 1532 cm^{-1} band for NAL^- species when adsorbed, can be attributed to variations of the electronic distribution in the NAL molecule upon surface complexation.^{23,24}

Dynamic adsorption. Similar to what has been disclosed for dried samples, NAL^- adsorption under the dynamic approach leads to a deprotonated sorbed molecule as demonstrated by the absence of a carboxyl band (at 1710 cm^{-1}), which is also in agreement with the pH measured after equilibrium, respectively 6.3 for Mh- NAL^- and 6.9 for Fh- NAL^- . It is noteworthy that the experimental conditions (NAL concentration and pH) were set up to be similar for both approaches (static and dynamic) and the main impact on sorption spectra comes from the different interfaces. For instance, the shape of the combined keto-carbonyl/carboxylate band changes noticeably when compared to the static approach, mainly narrower in the dynamic mode (water/solid interface). The shoulder contribution at around 1650 cm^{-1} in the static approach is much smaller in intensity (and even absent for Fh) in the dynamic mode suggesting that this feature is highly affected by solvation. The most intense band for Mh- NAL^- remains centered at 1628 cm^{-1} (as with the air-dried sample) while for Fh- NAL^- it downshifts from 1632 to 1624 cm^{-1} . In this case, upshifts with respect to $\text{NAL}^-_{(\text{aq})}$ due to H-bonding involvement and/or to the carbon angle distortion seem to happen but to be somehow attenuated due to solvation of the sorbed species. Actually, shifts are mostly smaller in the dynamic (solvated) approach when comparing to static (dried) one. For instance, the shoulder at 1609 cm^{-1} slightly shifts to 1612 (Mh) or 1607 cm^{-1} (Fh), while the COO^- symmetric stretching mode stays at the same position ($1261 \pm 1\text{ cm}^{-1}$ for both Mh and Fh). It can also be noted that the band assigned to the combined keto-carbonyl/carboxylate band at about 1625 cm^{-1} is larger for Mh than for Fh suggesting the existence of more than one sorption site for the former.

Time-course evolution of the sorbed NAL^- species in dynamic mode

The adsorption of NAL^- on Mh and Fh were followed *in situ* by time-resolved FTIR spectroscopy. The results shown previously (in Fig. 3b) are the final spectra collected after 120 minutes of circulating flow. However, more information might be acquired by following their sorption over time as shown in Fig. 4.

Fig. 5 illustrates the progressive NAL^- sequestration at the sorbent surface by plotting the relative intensity of target bands as a function of time (details on the calculation of peak intensities are provided in ESI[†] as well as detailed numerical values concerning the fitting – shown as dashed lines in Fig. 5).

Bands at 1525 ± 4 , 1500 ± 4 and $1450 \pm 1\text{ cm}^{-1}$ have been chosen because they are related to the quinolone ring (see Table 3) and thus are considered as related to vibrations involving chemical moieties far from adsorption sites.

A major difference among Fig. 5a and b is the order of magnitude of absorbance, being about 10 times higher for Fh than for Mh. Since the target surface coverage is the same ($2\text{ molecules nm}^{-2}$) the main variable is the iron substrate

involved. ATR-FTIR spectroscopy relies on optical properties of the material under study and Fig. 5 show that Fh is more transparent than Mh. This can be explained either by a porosity effect or a refractive index difference – n_{Fh} is about 2.3 while n_{Mh} is about 2.6 (see Table 1). Total reflection, the basis of ATR spectroscopy, occurs when light traveling through a medium with a higher refractive index hits the boundary of a medium with a lower refractive index at an angle greater than the critical angle, resulting in all the light being reflected into the original medium. In other words, the n_{IRE} must be superior to n_{sample} to guarantee the formation of an evanescent wave and probe the interface of interest. Since n_{znse} is a constant in our set-up with a value of 2.4, when n_{sample} increases the measurement might be globally attenuated – less light is reflected – and consequently less interface with the substrate is probed, which explains a lower absorbance for Mh.

In the case of NAL^- adsorption onto Mh surface, bands at 1528 , 1504 and 1451 cm^{-1} (Fig. 5a) can be smoothly fitted with a biexponential kinetic function, meaning that data can be model with two different time constants, which suggests the presence of two different adsorption sites. Changes in band intensities imply rapid adsorption on one site and slower adsorption on another one. The variations in relative intensities for the bands at 1528 and 1504 cm^{-1} for the sample on Fig. 4a are in agreement with this hypothesis. The intensity of the two bands is almost identical for the 12-minute spectrum (green line) while the band at 1528 cm^{-1} is more intense for the 120-minute spectrum (black line). This change also suggest that these two bands follow two slightly different kinetics, which would be in agreement with the presence of two binding sites in competition. In the same spirit, Fig. 4a shows that the band at about 1390 cm^{-1} for pure $\text{NAL}^-_{(\text{aq})}$ (Fig. 2c) splits into two bands at 1398 and 1387 cm^{-1} once NAL is sorbed onto Mh under the dynamic approach suggesting also the existence of two sorption sites.

Conversely, for NAL^- adsorbed on Fh, the evolution of bands at 1521 , 1498 and 1449 cm^{-1} (Fig. 5b) can be properly fitted both by a monoexponential or by a biexponential function since the adjusted R^2 values are very close for both functions (see ESI[†]). However, no clear changes in relative intensity of bands can be seen from the spectra shown in Fig. 4b. Therefore, it may be concluded that progressive NAL adsorption onto Fh likely occurs through a single adsorption site, or, if more than one adsorption site is present, they all show similar adsorption kinetics.

Discussion

NAL^- binding modes as derived from ATR-FTIR band assignment

Information on the binding mode of sorbed molecules can be extracted with IR spectroscopy since a modification of peak positions is observed with respect to the molecules in solution.¹⁰ Carboxylate stretching modes are very informative in this respect. Table 4 summarizes their peak positions with different assignments already discussed above.

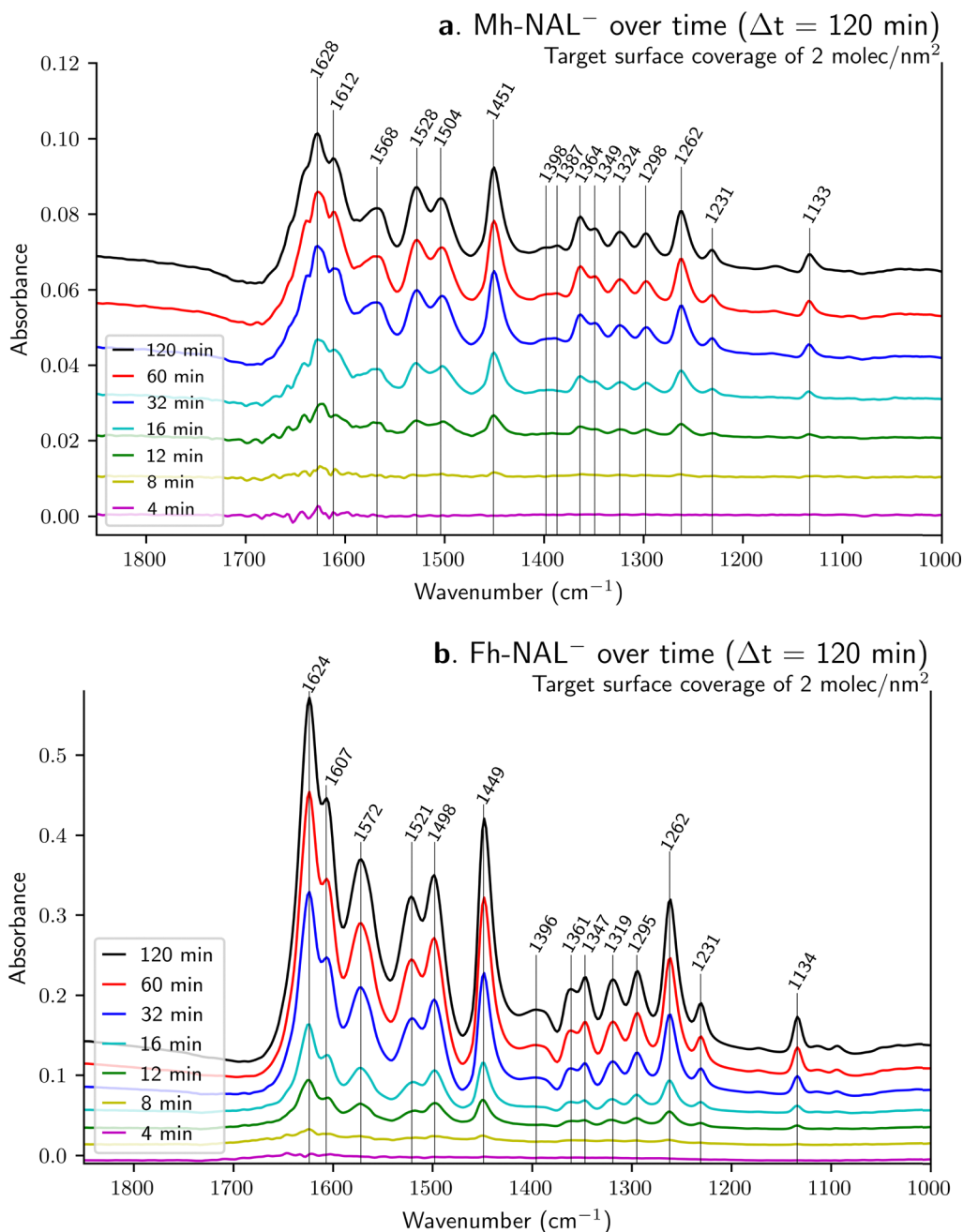


Fig. 4 Temporal evolution of the dynamic adsorption using ATR-FTIR spectra of NAL onto (a) Mh and (b) Fh with initial circulating concentrations of 0.026 and 0.137 mM and a final pH of 6.3 and 6.9 after equilibrium respectively. Spectra have been vertically shifted for clarity and no normalization has been performed.

Surface complexes from carboxylate-containing molecules are usually forming either monodentate or bidentate structures. A way to interpret binding mode of acetate groups in pure metal acetate was first proposed by Nakamoto.⁵⁹ This method is based on assessing the difference between the asymmetric and symmetric carboxylate stretching modes for surface complexes, noted as $\Delta\nu$, and now largely used in literature to determine molecular binding mode.^{5,59,60} By following this methodology, if $\Delta\nu$ carboxylates for surface complexes is substantially greater than for the free anion

($\Delta\nu_{\text{Me-COO}^-} > \Delta\nu_{\text{COO}^-}$), it is suggested that one oxygen of the carboxylate binds to one surface site as a monodentate complex.^{48,59,60} If the contrary happens ($\Delta\nu_{\text{Me-COO}^-} < \Delta\nu_{\text{COO}^-}$), it is suggested that two oxygens bind to one surface site as a bidentate chelate.^{48,59} Lastly, if $\Delta\nu$ remains similar, it is suggested that two oxygens bind to two surface sites as a bridging bidentate complex. Fig. 6 illustrates the possible binding modes for NAL onto iron(oxy)(hydr)oxides *via* carboxylate group attachment involving also potential additional hydrogen bonding discussed in the text.

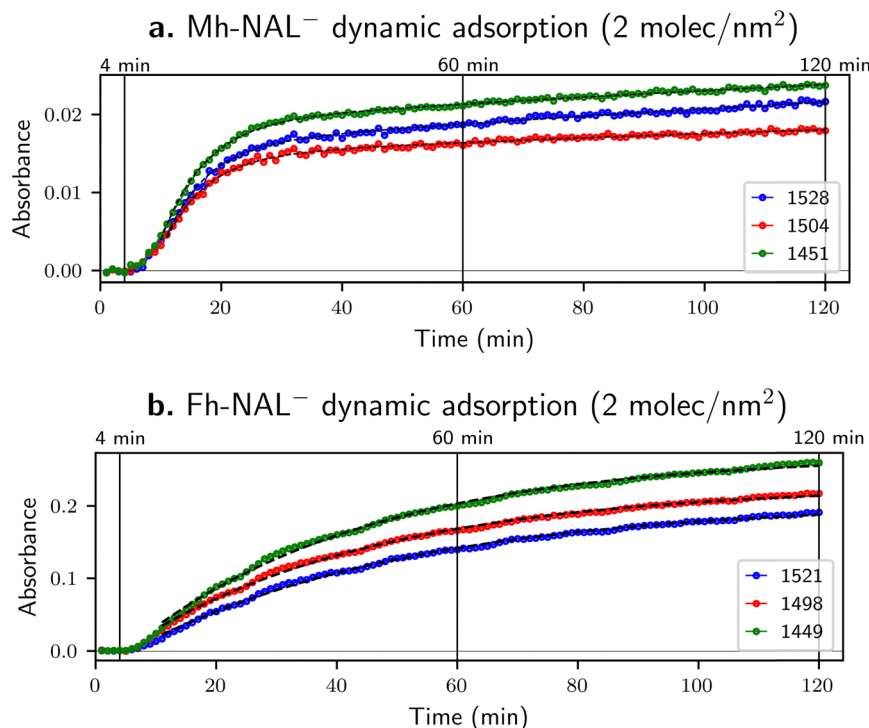


Fig. 5 Kinetics extracted from ATR-FTIR spectra for dynamic sorption of $\text{NAL}^-_{(\text{aq})}$ onto (a) Mh and (b) Fh at a target surface coverage of 2 molecules nm^{-2} for three individual bands. Absorbance data were extracted and fitted with exponential functions (dashed lines) in a 10–120-minute window (more details available in ESI†).

Table 4 Carboxylate peak positions [cm^{-1}] for NAL species, whereas asymmetric stretching COO^- mode ($\nu_{\text{COO}^- \text{ as}}$), symmetric stretching COO^- mode ($\nu_{\text{COO}^- \text{ s}}$) and the difference between these two bands ($\Delta\nu_{\text{COO}^-}$) are given. Peak positions are based on Fig. 2c ($\text{NAL}^-_{(\text{aq})}$) and Fig. 3 (NAL^- sorbed on Mh and Fh)

		Carboxylate assignments via literature (Table 3)			Carboxylate assignments with DFT calculations (this work, Table 3)		
		$\nu_{\text{COO}^- \text{ as}}$, cm^{-1}	$\nu_{\text{COO}^- \text{ s}}$, cm^{-1}	$\Delta\nu_{\text{COO}^-}$, cm^{-1}	$\nu_{\text{COO}^- \text{ as}}$, cm^{-1}	$\nu_{\text{COO}^- \text{ s}}$, cm^{-1}	$\Delta\nu_{\text{COO}^-}$, cm^{-1}
DFT calculated	NAL^-	—	—	—	1650 1614	1276	374 338
Solution	$\text{NAL}^-_{(\text{aq})}$	1574	1389	185	1623 1609	1261	362 348
NAL sorption via Static approach	Mh- NAL^-	1580	1396	184	1628 1612	1261	367 351
	Fh- NAL^-	1567	1391	176	1632 1612	1259	373 353
NAL sorption via dynamic approach	Mh- NAL^-	1568	1398	170	1628 1612	1262	366 350
	Fh- NAL^-	1572	1396	176	1624 1607	1262	362 345

It is noteworthy that there is no consensus on the definitive value from which $\Delta\nu$ can be considered similar for the surface complex and the free anion. Trivedi and Vasudevan noted as $\Delta\nu_{\text{Me-COO}} \approx \Delta\nu_{\text{COO}^-}$ when the difference was equal or below 5 cm^{-1} , under the argument that 5 wavenumber units is not statistically significant.⁴⁸ Results using this methodology from published literature (obtained by ATR-FTIR spectroscopy on wet-pastes recovered from batch experiments) of NAL and

ciprofloxacin (CIP, another quinolone) sorbed on Goethite (Gt) – the most abundant iron form in nature – and Magnetite (Mt) are summarized in Table 5.

Cheng and co-workers¹⁸ presented results for NAL adsorbed on Gt where $\Delta\nu$ increases upon binding which suggests a monodentate coordination following literature ($\Delta\nu_{\text{NAL-COO}^-} > \Delta\nu_{\text{COO}^-}$). The authors suggested that the keto-group (*i.e.* the ketone group from the quinoline ring) is involved in the

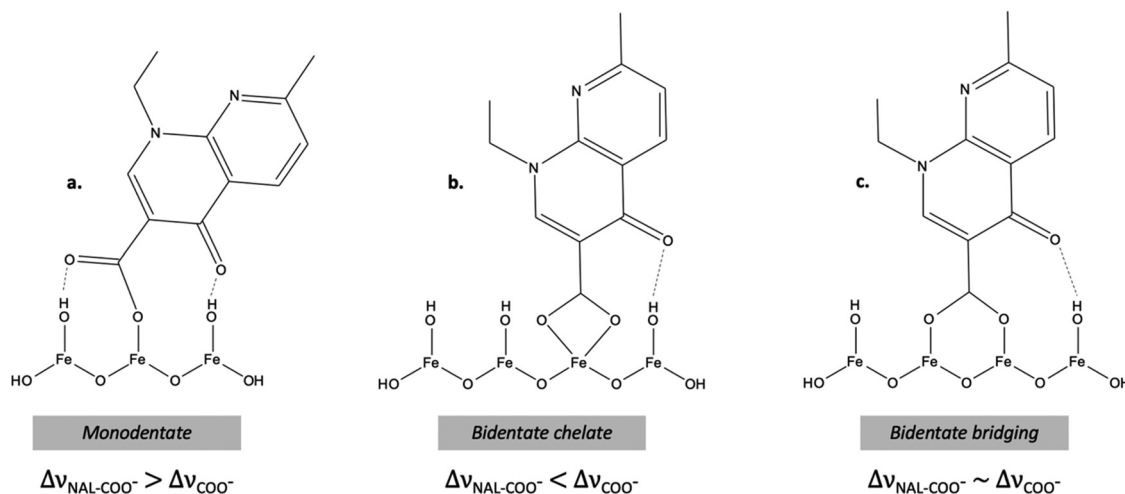


Fig. 6 Possible binding modes for NAL sorbed onto iron oxides via carboxylate attachment. Using only one iron site as a (a) monodentate or a (b) bidentate chelate or using two surface sites as a (c) bidentate bridging mode. Potential additional hydrogen bonding is also shown. These models are suggested based on literature data for quinolones adsorption.^{1,23,49}

Table 5 Overview on peak positions from the literature for different quinolones (nalidixic acid (NAL) and ciprofloxacin (CIP)) adsorbed on goethite (Gt) or on magnetite (Mt)

Absorbent-absorbate	pH	Free carboxylate			Substrate-quinolone complexed			Result
		$\nu_{\text{COO}^- \text{ as}}$	$\nu_{\text{COO}^- \text{ s}}$	$\Delta\nu_{\text{COO}^-}$	$\nu_{\text{NAL-COO}^- \text{ as}}$	$\nu_{\text{NAL-COO}^- \text{ s}}$	$\Delta\nu_{\text{NAL-COO}^-}$	
Gt-NAL ¹⁸	7	1583	1394	189	1583	1387	196	$\Delta\nu_{\text{NAL-COO}^-} > \Delta\nu_{\text{COO}^-}$
Gt-NAL ²³	4–6	1578	1392	186	Not affected by adsorption	Not marked	Not marked	$\Delta\nu_{\text{NAL-COO}^-} > \Delta\nu_{\text{COO}^-}$
Gt-CIP ⁴⁸	6	1577	1381	196	1534	1384	150	$\Delta\nu_{\text{CIP-COO}^-} < \Delta\nu_{\text{COO}^-}$
Gt-CIP ¹⁴	6	1577	1390	187	1572	1391	181	$\Delta\nu_{\text{CIP-COO}^-} < \Delta\nu_{\text{COO}^-}$
Mt-CIP ⁵³	4–8	1580	1385	195	1577	1390	187	$\Delta\nu_{\text{CIP-COO}^-} < \Delta\nu_{\text{COO}^-}$

bonding mode either through a direct chemical bond with surface iron atoms or through hydrogen bonding such as that shown in Fig. 6a in agreement with other works.^{49,58} Overall, the involvement of the keto-group is usually suggested in literature when the band attributed to C=O of the carbonyl blueshifts upon adsorption, which is the case for Cheng and co-workers.¹⁸ The same group²³ had previously reached the same conclusion for NAL on Gt but the exact band position was not reported for the symmetric carboxylate mode as well as for $\Delta\nu_{\text{NAL-COO}^-}$.

Table 5 also reports results obtained with another quinolone-type antibiotic, the Ciprofloxacin (CIP). Trivedi and Vasudevan,⁴⁸ Gu and co-workers¹⁴ and Rakshit *et al.*⁵³ show that $\Delta\nu$ decreases upon CIP adsorption which suggests a bidentate chelate adsorption mode ($\Delta\nu_{\text{CIP-COO}^-} < \Delta\nu_{\text{COO}^-}$). The only difference in their conclusions is coming from the potential involvement of the keto-group. Trivedi and Vasudevan⁴⁸ suggest that the band shift of the carbonyl band is driven by the electron withdrawing nature of the *ortho* carboxylate-iron bond, while Gu and co-workers¹⁴ on Goethite and Rakshit *et al.*⁵³ on Magnetite suggest that the band widening and shifting is the signature of the involvement of the keto-group in the surface complex (see Fig. 6c for NAL). It is noteworthy that the involvement of two functional groups of quinolones onto the surface coordination might explain its high affinity.⁵³

Experimental data from the present work are summarized in Table 4. Two approaches are possible regarding the carboxylate binding mode, whether one considers the carboxylate peak assignment often reported in the literature (*i.e.* $\nu_{\text{COO}^- \text{ as}} = 1574 \text{ cm}^{-1}$ and $\nu_{\text{COO}^- \text{ s}} = 1389 \text{ cm}^{-1}$ for $\text{NAL}^-_{(\text{aq})}$ species free in solution) or the DFT calculations discussed above (*i.e.* $\nu_{\text{COO}^- \text{ as}} = 1650$ and 1614 cm^{-1} and $\nu_{\text{COO}^- \text{ s}} = 1276 \text{ cm}^{-1}$ for gas-phase NAL^-).

In the first case, based on the carboxylate peak assignment often reported in the literature, $\Delta\nu_{\text{COO}^-}$ is 185 cm^{-1} for the reference $\text{NAL}^-_{(\text{aq})}$. Experimental data for Mh under the static (dried) approach provide a $\Delta\nu$ very close to the reference (184 cm^{-1} instead of 185 cm^{-1}) while, when hydrated (dynamic approach), some ambiguity arises from the fact that peak assigned as $\nu_{\text{COO}^- \text{ symmetric}}$ is very low in intensity and splits in two bands as previously discussed. It is proposed to follow the $\Delta\nu$ with both positions leading to $\Delta\nu$ of 170 and 181 cm^{-1} . If this is the case, the former suggests a bidentate chelate ($\Delta\nu_{\text{NAL-COO}^-} < \Delta\nu_{\text{COO}^-}$) while the second a bridging bidentate ($\Delta\nu_{\text{NAL-COO}^-} \simeq \Delta\nu_{\text{COO}^-}$) – which agrees with the kinetic results suggesting that two sorption sites exist for Mh. Experimental data for Fh provided the same $\Delta\nu$ of 176 cm^{-1} for both approaches (static or dynamic), suggesting a bidentate chelate surface complex ($\Delta\nu_{\text{NAL-COO}^-} < \Delta\nu_{\text{COO}^-}$).

The alternative way to interpret values in Table 4 relies on peak assignment based on DFT calculations, where the COO^-

asymmetric stretching mode is conjugated with the carbonyl, appearing at higher frequencies, simultaneously at 1650 and 1614 cm^{-1} for the gas-phase NAL while the symmetric mode is centered at 1276 cm^{-1} , giving a theoretical $\Delta\nu$ of 374 cm^{-1} or 338 cm^{-1} depending on the peak chosen for the asymmetric mode. When it comes to solvated $\text{NAL}^-_{(\text{aq})}$, values are slightly downshifted – which is understandable since DFT is not explicitly considering the solvent – leading to a $\Delta\nu$ of 362 cm^{-1} or 348 cm^{-1} . Under the static (dried) approach, sorption onto Mh leads to a $\Delta\nu$ of 367 cm^{-1} and 351 cm^{-1} , which is very similar to $\text{NAL}^-_{(\text{aq})}$ and consistent with a bridging bidentate sorption complex previously discussed. In contrast, sorption onto Fh leads to a $\Delta\nu$ of 373 cm^{-1} and 353 cm^{-1} . From this, it is more difficult to draw a definite conclusion since the former suggests a monodentate ($\Delta\nu_{\text{NAL-COO}^-} > \Delta\nu_{\text{COO}^-}$) while the second a bridging bidentate ($\Delta\nu_{\text{NAL-COO}^-} \simeq \Delta\nu_{\text{COO}^-}$). Under the dynamic approach, $\Delta\nu$ of 366 cm^{-1} and 350 cm^{-1} for Mh and 362 cm^{-1} and 345 cm^{-1} for Fh are found, which are similar to $\text{NAL}^-_{(\text{aq})}$ and indicate a bidentate bridging mode ($\Delta\nu_{\text{NAL-COO}^-} \simeq \Delta\nu_{\text{COO}^-}$) for both substrates. Hence, all these results show that the Nakamoto⁵⁹ methodology does not allow one to discriminate binding modes without any ambiguity based on the DFT assignments presented in this work.

Overall, the spectroscopic approach followed in this work points, predominantly, to the possible existence of two sorption modes for the aqueous nalidixate ion: a bidentate chelate and a bidentate bridging mode depending on the substrate (Maghemite and Ferrihydrite) and the interface considered (solid/air or solid/liquid). Two approaches have been considered by using the Nakamoto⁵⁹ methodology based on the asymmetric and symmetric carboxylate stretching modes. These approaches differ in the assignment of the carboxylate modes whether literature data or DFT calculations performed in this work are considered. However, a bidentate mode appears to be predominant and a monodentate adsorption can be mostly excluded which is at variance with the results of Cheng and co-workers¹⁸ and Xu *et al.*²³ on goethite. The existence of two adsorption modes for Maghemite (Mh), bidentate chelate and bridging bidentate, is also in line with the hypothesis that two sorption sites are present for this oxide.

Conclusions

Continuous-flow FTIR-ATR spectroscopy provides a real-time, *in situ*, monitoring of nalidixate (NAL^-) adsorption on iron oxy-hydroxides, Maghemite (Mh) and Ferrihydrite (Fh), at the oxide water interface highlighting the critical role of surface hydration. A complete band assignment for nalidixic acid and nalidixate is discussed based on literature results and DFT calculations. The experimental results suggest that the nalidixate ion (NAL^-) adsorbs onto maghemite and ferrihydrite surfaces *via* a bidentate mechanism mostly excluding monodentate adsorption. However, these result does not permit to fully conclude on the prevalence of the two bidentate sorption modes: bidentate chelate *versus* bidentate bridging. Nevertheless, the temporal

evolution of IR spectra on Mh suggests the existence of two sorption sites: (i) band broadening (1628 cm^{-1}), (ii) band splitting (1390 cm^{-1}) and (iii) kinetic fitting (biexponential function). Hence, the two sorption modes (bidentate chelate and bidentate bridging) on two different sorption sites could coexist for Mh while only one sorption site is identified for Fh.

This study advances our molecular-level comprehension of NAL adsorption, demonstrating the potential of iron oxides in environmental remediation applications. The information gained could help in the development of more effective strategies for mitigating pharmaceutical contamination in natural water systems, emphasizing the environmental relevance of iron oxides in contaminant sequestration.

Data availability

The data supporting this article (XRD characterization, full set of IR spectra of the time-resolved adsorption experiments – dynamic approach as well as kinetic fitting data), have been included as part of the ESI.†

Conflicts of interest

There are no conflicts of interest to declare.

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References

- 1 Q. Wu, Z. Li and H. Hong, Adsorption of the Quinolone Antibiotic Nalidixic Acid onto Montmorillonite and Kaolinite, *Appl. Clay Sci.*, 2013, **74**, 66–73, DOI: [10.1016/j.clay.2012.09.026](https://doi.org/10.1016/j.clay.2012.09.026).
- 2 B. Halling-Sørensen, S. Nors Nielsen, P. F. Lanzky, F. Ingerslev, H. C. Holten Lützhøft and S. E. Jørgensen, Occurrence, Fate and Effects of Pharmaceutical Substances in the Environment – A Review, *Chemosphere*, 1998, **36**(2), 357–393, DOI: [10.1016/S0045-6535\(97\)00354-8](https://doi.org/10.1016/S0045-6535(97)00354-8).
- 3 G. Lefèvre and M. Fédoroff, Sorption of Sulfate Ions onto Hematite Studied by Attenuated Total Reflection-Infrared Spectroscopy: Kinetics and Competition with Other Ions, *Phys. Chem. Earth Parts ABC*, 2006, **31**(10–14), 499–504, DOI: [10.1016/j.pce.2006.04.001](https://doi.org/10.1016/j.pce.2006.04.001).
- 4 D. Peak, R. G. Ford and D. L. Sparks, An *in situ* ATR-FTIR Investigation of Sulfate Bonding Mechanisms on Goethite,

- J. Colloid Interface Sci.*, 1999, **218**(1), 289–299, DOI: [10.1006/jcis.1999.6405](https://doi.org/10.1006/jcis.1999.6405).
- 5 G. Lefèvre, T. Preočanin and J. Lützenkirchen, Attenuated Total Reflection – Infrared Spectroscopy Applied to the Study of Mineral – Aqueous Electrolyte Solution Interfaces: A General Overview and a Case Study, in *Infrared Spectroscopy – Materials Science, Engineering and Technology*, ed. T. Theophanides, InTech, 2012, DOI: [10.5772/36459](https://doi.org/10.5772/36459).
 - 6 T. Paul, J. Liu, M. L. Machesky and T. J. Strathmann, Adsorption of Zwitterionic Fluoroquinolone Antibacterials to Goethite: A Charge Distribution-Multisite Complexation Model, *J. Colloid Interface Sci.*, 2014, **428**, 63–72, DOI: [10.1016/j.jcis.2014.04.034](https://doi.org/10.1016/j.jcis.2014.04.034).
 - 7 Q. Liu, K. Xu, G. Hu, F. Zeng, X. Li, C. Li and Y. Zhang, Underwater Superelastic MOF/Polyacrylamide/Chitosan Composite Aerogel for Efficient 2,4-Dichlorophenoxyacetic Acid Adsorption, *Colloids Surf., A*, 2022, **635**, 127970, DOI: [10.1016/j.colsurfa.2021.127970](https://doi.org/10.1016/j.colsurfa.2021.127970).
 - 8 S. G. Ardo, S. Nélieu, G. Ona-Nguema, G. Delarue, J. Brest, E. Pironin and G. Morin, Oxidative Degradation of Nalidixic Acid by Nano-Magnetite via Fe²⁺/O₂-Mediated Reactions, *Environ. Sci. Technol.*, 2015, **49**(7), 4506–4514, DOI: [10.1021/es505649d](https://doi.org/10.1021/es505649d).
 - 9 S. J. Hug, In Situ Fourier Transform Infrared Measurements of Sulfate Adsorption on Hematite in Aqueous Solutions, *J. Colloid Interface Sci.*, 1997, **188**(2), 415–422, DOI: [10.1006/jcis.1996.4755](https://doi.org/10.1006/jcis.1996.4755).
 - 10 P. Roonasi and A. Holmgren, An ATR-FTIR Study of Sulphate Sorption on Magnetite; Rate of Adsorption, Surface Speciation, and Effect of Calcium Ions, *J. Colloid Interface Sci.*, 2009, **333**(1), 27–32, DOI: [10.1016/j.jcis.2008.12.080](https://doi.org/10.1016/j.jcis.2008.12.080).
 - 11 Y. Arai and D. L. Sparks, ATR-FTIR Spectroscopic Investigation on Phosphate Adsorption Mechanisms at the Ferrihydrite–Water Interface, *J. Colloid Interface Sci.*, 2001, **241**(2), 317–326, DOI: [10.1006/jcis.2001.7773](https://doi.org/10.1006/jcis.2001.7773).
 - 12 A. M. Awad, R. Jalab, A. Benamor, M. S. Nasser, M. M. Ba-Abbad, M. El-Naas and A. W. Mohammad, Adsorption of Organic Pollutants by Nanomaterial-Based Adsorbents: An Overview, *J. Mol. Liq.*, 2020, **301**, 112335, DOI: [10.1016/j.molliq.2019.112335](https://doi.org/10.1016/j.molliq.2019.112335).
 - 13 S. B. Johnson, T. H. Yoon, B. D. Kocar and G. E. Brown, Adsorption of Organic Matter at Mineral/Water Interfaces. 2. Outer-Sphere Adsorption of Maleate and Implications for Dissolution Processes, *Langmuir*, 2004, **20**(12), 4996–5006, DOI: [10.1021/la036288y](https://doi.org/10.1021/la036288y).
 - 14 X. Gu, Y. Tan, F. Tong and C. Gu, Surface Complexation Modeling of Coadsorption of Antibiotic Ciprofloxacin and Cu(II) and onto Goethite Surfaces, *Chem. Eng. J.*, 2015, **269**, 113–120, DOI: [10.1016/j.cej.2014.12.114](https://doi.org/10.1016/j.cej.2014.12.114).
 - 15 J. R. Bargar, J. D. Kubicki, R. Reitmeyer and J. A. Davis, ATR-FTIR Spectroscopic Characterization of Coexisting Carbonate Surface Complexes on Hematite, *Geochim. Cosmochim. Acta*, 2005, **69**(6), 1527–1542, DOI: [10.1016/j.gca.2004.08.002](https://doi.org/10.1016/j.gca.2004.08.002).
 - 16 A. J. Watkinson, E. J. Murby, D. W. Kolpin and S. D. Costanzo, The Occurrence of Antibiotics in an Urban Watershed: From Wastewater to Drinking Water, *Sci. Total Environ.*, 2009, **407**(8), 2711–2723, DOI: [10.1016/j.scitotenv.2008.11.059](https://doi.org/10.1016/j.scitotenv.2008.11.059).
 - 17 R. Sugden, R. Kelly and S. Davies, Combatting Antimicrobial Resistance Globally, *Nat. Microbiol.*, 2016, **1**(10), 16187, DOI: [10.1038/nmicrobiol.2016.187](https://doi.org/10.1038/nmicrobiol.2016.187).
 - 18 W. Cheng, J. Li, J. Sun, T. Luo, R. Marsac, J.-F. Boily and K. Hanna, Nalidixic Acid and Fe(II)/Cu(II) Coadsorption at Goethite and Akaganéite Surfaces, *Environ. Sci. Technol.*, 2023, **57**(41), 15680–15692, DOI: [10.1021/acs.est.3c05727](https://doi.org/10.1021/acs.est.3c05727).
 - 19 S. Monti, I. Manet, F. Manoli, M. L. Capobianco and G. Marconi, Gaining an Insight into the Photoreactivity of a Drug in a Protein Environment: A Case Study on Nalidixic Acid and Serum Albumin, *J. Phys. Chem. B*, 2008, **112**(18), 5742–5754, DOI: [10.1021/jp711261n](https://doi.org/10.1021/jp711261n).
 - 20 E. Clervil, M. Usman, E. Emmanuel, V. Chatain and K. Hanna, Sorption of Nalidixic Acid onto Sediments under Batch and Dynamic Flow Conditions, *Chem. Geol.*, 2013, **335**, 63–74, DOI: [10.1016/j.chemgeo.2012.10.041](https://doi.org/10.1016/j.chemgeo.2012.10.041).
 - 21 X. Fan, H. Hao, Y. Wang, F. Chen and J. Zhang, Fenton-like Degradation of Nalidixic Acid with Fe³⁺/H₂O₂, *Environ. Sci. Pollut. Res.*, 2013, **20**(6), 3649–3656, DOI: [10.1007/s11356-012-1279-0](https://doi.org/10.1007/s11356-012-1279-0).
 - 22 M. Usman, S. Martin, N. Cimetière, S. Giraudet, V. Chatain and K. Hanna, Sorption of Nalidixic Acid onto Micrometric and Nanometric Magnetites: Experimental Study and Modeling, *Appl. Surf. Sci.*, 2014, **299**, 136–145, DOI: [10.1016/j.apsusc.2014.01.197](https://doi.org/10.1016/j.apsusc.2014.01.197).
 - 23 J. Xu, R. Marsac, D. Costa, W. Cheng, F. Wu, J.-F. Boily and K. Hanna, Co-Binding of Pharmaceutical Compounds at Mineral Surfaces: Molecular Investigations of Dimer Formation at Goethite/Water Interfaces, *Environ. Sci. Technol.*, 2017, **51**(15), 8343–8349, DOI: [10.1021/acs.est.7b02835](https://doi.org/10.1021/acs.est.7b02835).
 - 24 U. Neugebauer, A. Szeghalmi, M. Schmitt, W. Kiefer, J. Popp and U. Holzgrabe, Vibrational Spectroscopic Characterization of Fluoroquinolones, *Spectrochim. Acta, Part A*, 2005, **61**(7), 1505–1517, DOI: [10.1016/j.saa.2004.11.014](https://doi.org/10.1016/j.saa.2004.11.014).
 - 25 A. C. Hari, R. A. Paruchuri, D. A. Sabatini and T. C. G. Kibbey, Effects of pH and Cationic and Nonionic Surfactants on the Adsorption of Pharmaceuticals to a Natural Aquifer Material, *Environ. Sci. Technol.*, 2005, **39**(8), 2592–2598, DOI: [10.1021/es048992m](https://doi.org/10.1021/es048992m).
 - 26 D. Ross and C. Riley, Aqueous Solubilities of Some Various Substituted Quinolone Antimicrobials, *Int. J. Pharm.*, 1990, **63**(3), 237–250, DOI: [10.1016/0378-5173\(90\)90130-V](https://doi.org/10.1016/0378-5173(90)90130-V).
 - 27 California Office of Environmental Health Hazard Assessment, Expedited Risk Assessment of Three Priority Carcinogens: 1,3-Butadiene, Ethylene Oxide, and 2-Amino-3-Methylimidazo[4,5-f]Quinoline (IQ), 2002, <https://oehha.ca.gov/media/downloads/proposition-65/chemicals/expedit1aug2002.pdf>.
 - 28 F. Tamtam, F. Van Oort, B. Le Bot, T. Dinh, S. Mompelat, M. Chevreuil, I. Lamy and M. Thiry, Assessing the Fate of Antibiotic Contaminants in Metal Contaminated Soils Four Years after Cessation of Long-Term Waste Water Irrigation, *Sci. Total Environ.*, 2011, **409**(3), 540–547, DOI: [10.1016/j.scitotenv.2010.10.033](https://doi.org/10.1016/j.scitotenv.2010.10.033).

- 29 A. Y.-C. Lin, T.-H. Yu and C.-F. Lin, Pharmaceutical Contamination in Residential, Industrial, and Agricultural Waste Streams: Risk to Aqueous Environments in Taiwan, *Chemosphere*, 2008, **74**(1), 131–141, DOI: [10.1016/j.chemosphere.2008.08.027](https://doi.org/10.1016/j.chemosphere.2008.08.027).
- 30 S. J. Hug and B. Sulzberger, In Situ Fourier Transform Infrared Spectroscopic Evidence for the Formation of Several Different Surface Complexes of Oxalate on TiO₂ in the Aqueous Phase, *Langmuir*, 1994, **10**(10), 3587–3597, DOI: [10.1021/la00022a036](https://doi.org/10.1021/la00022a036).
- 31 S. P. Schwaminger, D. Bauer, P. Fraga-García, F. E. Wagner and S. Berensmeier, Oxidation of Magnetite Nanoparticles: Impact on Surface and Crystal Properties, *CrystEngComm*, 2017, **19**(2), 246–255, DOI: [10.1039/C6CE02421A](https://doi.org/10.1039/C6CE02421A).
- 32 T. Tatarchuk, L. Soltys and W. Macyk, Magnetic Adsorbents for Removal of Pharmaceuticals: A Review of Adsorption Properties, *J. Mol. Liq.*, 2023, **384**, 122174, DOI: [10.1016/j.molliq.2023.122174](https://doi.org/10.1016/j.molliq.2023.122174).
- 33 F. M. Michel, L. Ehm, S. M. Antao, P. L. Lee, P. J. Chupas, G. Liu, D. R. Strongin, M. A. A. Schoonen, B. L. Phillips and J. B. Parise, The Structure of Ferrihydrite, a Nanocrystalline Material, *Science*, 2007, **316**(5832), 1726–1729, DOI: [10.1126/science.1142525](https://doi.org/10.1126/science.1142525).
- 34 G. Ona-Nguema, G. Morin, Y. Wang, A. L. Foster, F. Juillot, G. Calas and G. E. Brown, XANES Evidence for Rapid Arsenic(III) Oxidation at Magnetite and Ferrihydrite Surfaces by Dissolved O₂ via Fe²⁺-Mediated Reactions, *Environ. Sci. Technol.*, 2010, **44**(14), 5416–5422, DOI: [10.1021/es1000616](https://doi.org/10.1021/es1000616).
- 35 T. Hiemstra, Surface and Mineral Structure of Ferrihydrite, *Geochim. Cosmochim. Acta*, 2013, **105**, 316–325, DOI: [10.1016/j.gca.2012.12.002](https://doi.org/10.1016/j.gca.2012.12.002).
- 36 U. Schwertmann and R. M. Cornell, *Iron Oxides in the Laboratory: Preparation and Characterization*, John Wiley & Sons, 2008.
- 37 J. F. Béjar and G. Baldinozzi, XND Code: From X-Ray Laboratory Data to Incommensurately Modulated Phases, *Commun. Powder Diffract Newsl.*, 1998, **20**, 3–5.
- 38 Maghemite, Webmineral, <https://www.Webmineral.Com/Data/Maghemite.Shtml>, Accessed 9 June 2024.
- 39 V. O. Leone, M. C. Pereira, S. F. Aquino, L. C. A. Oliveira, S. Correa, T. C. Ramalho, L. V. A. Gurgel and A. C. Silva, Adsorption of Diclofenac on a Magnetic Adsorbent Based on Maghemite: Experimental and Theoretical Studies, *New J. Chem.*, 2018, **42**(1), 437–449, DOI: [10.1039/C7NJ03214E](https://doi.org/10.1039/C7NJ03214E).
- 40 F. Maillot, G. Morin, Y. Wang, D. Bonnin, P. Ildefonse, C. Chaneac and G. Calas, New Insight into the Structure of Nanocrystalline Ferrihydrite: EXAFS Evidence for Tetrahedrally Coordinated Iron(III), *Geochim. Cosmochim. Acta*, 2011, **75**(10), 2708–2720, DOI: [10.1016/j.gca.2011.03.011](https://doi.org/10.1016/j.gca.2011.03.011).
- 41 Ferrihydrite, Webmineral, <https://www.Webmineral.Com/Data/Ferrihydrite.Shtml>, Accessed 9 June 2024.
- 42 N. Kumar, J. Lezama Pacheco, V. Noël, G. Dublet and G. E. Brown, Sulfidation Mechanisms of Fe(III)-(Oxyhydr)-Oxide Nanoparticles: A Spectroscopic Study, *Environ. Sci.: Nano*, 2018, **5**(4), 1012–1026, DOI: [10.1039/C7EN01109A](https://doi.org/10.1039/C7EN01109A).
- 43 J. A. Davis and J. O. Leckie, Surface Ionization and Complexation at the Oxide/Water Interface II. Surface Properties of Amorphous Iron Oxyhydroxide and Adsorption of Metal Ions, *J. Colloid Interface Sci.*, 1978, **67**(1), 90–107, DOI: [10.1016/0021-9797\(78\)90217-5](https://doi.org/10.1016/0021-9797(78)90217-5).
- 44 T. Bürgi and A. Baiker, Attenuated Total Reflection Infrared Spectroscopy of Solid Catalysts Functioning in the Presence of Liquid-Phase Reactants, *Advances in Catalysis*, Elsevier, 2006, vol. 50, pp. 227–283, DOI: [10.1016/S0360-0564\(06\)50005-7](https://doi.org/10.1016/S0360-0564(06)50005-7).
- 45 F. Neese, F. Wennmohs, U. Becker and C. Riplinger, The ORCA Quantum Chemistry Program Package, *J. Chem. Phys.*, 2020, **152**(22), 224108, DOI: [10.1063/5.0004608](https://doi.org/10.1063/5.0004608).
- 46 A. D. Becke, Density-Functional Thermochemistry. I. The Effect of the Exchange-Only Gradient Correction, *J. Chem. Phys.*, 1992, **96**(3), 2155–2160, DOI: [10.1063/1.462066](https://doi.org/10.1063/1.462066).
- 47 F. Weigend and R. Ahlrichs, Balanced Basis Sets of Split Valence, Triple Zeta Valence and Quadruple Zeta Valence Quality for H to Rn: Design and Assessment of Accuracy, *Phys. Chem. Chem. Phys.*, 2005, **7**(18), 3297, DOI: [10.1039/b508541a](https://doi.org/10.1039/b508541a).
- 48 P. Trivedi and D. Vasudevan, Spectroscopic Investigation of Ciprofloxacin Speciation at the Goethite–Water Interface, *Environ. Sci. Technol.*, 2007, **41**(9), 3153–3158, DOI: [10.1021/es061921y](https://doi.org/10.1021/es061921y).
- 49 R. Marsac, S. Martin, J.-F. Boily and K. Hanna, Oxolinic Acid Binding at Goethite and Akaganéite Surfaces: Experimental Study and Modeling, *Environ. Sci. Technol.*, 2016, **50**(2), 660–668, DOI: [10.1021/acs.est.5b04940](https://doi.org/10.1021/acs.est.5b04940).
- 50 S. Gunasekaran, R. K. Natarajan, R. Rathikha and D. Syamala, Vibrational Spectra and Normal Coordinate Analysis of Nalidixic Acid, *Indian J. Pure Appl. Phys.*, 2005, **43**(07), 503–508.
- 51 S. Romero, P. Bustamante, B. Escalera, P. Mura and M. Cirri, Influence of Solvent Composition on the Solid Phase at Equilibrium with Saturated Solutions of Quinolones in Different Solvent Mixtures, *J. Pharm. Biomed. Anal.*, 2004, **35**(4), 715–726, DOI: [10.1016/j.jpba.2004.02.006](https://doi.org/10.1016/j.jpba.2004.02.006).
- 52 P. Grubb, *Analytical Profiles of Drug Substances*, Academic Press, Inc., 1979, vol. 8.
- 53 S. Rakshit, D. Sarkar, E. J. Elzinga, P. Punamiya and R. Datta, Mechanisms of Ciprofloxacin Removal by Nano-Sized Magnetite, *J. Hazard. Mater.*, 2013, **246–247**, 221–226, DOI: [10.1016/j.jhazmat.2012.12.032](https://doi.org/10.1016/j.jhazmat.2012.12.032).
- 54 S. Muthu, E. E. Porchelvi and Y. Erdogdu, Density Functional Theory and Abinitio Studies of Vibrational Spectra of Nalidixic Acid, *J. Int. Acad. Phys. Sci.*, 2011, **15**(2), 205–222.
- 55 T. Paul, M. L. Machesky and T. J. Strathmann, Surface Complexation of the Zwitterionic Fluoroquinolone Antibiotic Ofloxacin to Nano-Anatase TiO₂ Photocatalyst Surfaces, *Environ. Sci. Technol.*, 2012, **46**(21), 11896–11904, DOI: [10.1021/es302097k](https://doi.org/10.1021/es302097k).
- 56 K. W. Goynes, J. Chorover, J. D. Kubicki, A. R. Zimmerman and S. L. Brantley, Sorption of the Antibiotic Ofloxacin to Mesoporous and Nonporous Alumina and Silica, *J. Colloid Interface Sci.*, 2005, **283**(1), 160–170, DOI: [10.1016/j.jcis.2004.08.150](https://doi.org/10.1016/j.jcis.2004.08.150).
- 57 M. Nara, H. Torii and M. Tasumi, Correlation between the Vibrational Frequencies of the Carboxylate Group and the Types of Its Coordination to a Metal Ion: An *Ab Initio*

- Molecular Orbital Study, *J. Phys. Chem.*, 1996, **100**(51), 19812–19817, DOI: [10.1021/jp9615924](https://doi.org/10.1021/jp9615924).
- 58 J. Xu, R. Marsac, C. Wei, F. Wu, J.-F. Boily and K. Hanna, Cobinding of Pharmaceutical Compounds at Mineral Surfaces: Mechanistic Modeling of Binding and Cobinding of Nalidixic Acid and Niflumic Acid at Goethite Surfaces, *Environ. Sci. Technol.*, 2017, **51**(20), 11617–11624, DOI: [10.1021/acs.est.7b02900](https://doi.org/10.1021/acs.est.7b02900).
- 59 K. Nakamoto, *Infrared and Raman Spectra of Inorganic and Coordination Compounds, Part B: Applications in Coordination, Organometallic, and Bioinorganic Chemistry*, John Wiley & Sons, 4th edn, 1986.
- 60 J. E. Tackett, FT-IR Characterization of Metal Acetates in Aqueous Solution, *Appl. Spectrosc.*, 1989, **43**(3), 483–489, DOI: [10.1366/0003702894202931](https://doi.org/10.1366/0003702894202931).