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# N-Heterocyclic carbene-initiated epoxide/anhydride ring-opening copolymerization: effective and selective organoinitiators for the production of various polyesters†

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The present report describes the first use of unprotected and “free” N-heterocyclic carbenes (NHCs), such as IMes, IPr and air-stable Cl-IPr carbenes, as single component initiators for the alternating ring-opening co-polymerization (ROCOP) of cyclic anhydrides (phthalic anhydride, PA; succinic anhydride, SA) and various epoxides to efficiently and selectively access the corresponding polyesters in a well-defined manner and as metal-free materials. In the case of ROCOP runs with PA as the anhydride source, control experiments are in line with an initiation *via* ring-opening of PA by the NHC moiety, as established with the synthesis and structural characterization of the IMes-PA ring-opened product **1**. The present NHC systems were further exploited as ROCOP initiators of PA and bio-sourced epoxides such as eugenyl glycidyl ether (EGE) and saffrole glycidylether (SO) yielding the production of regioregular p(PA-*alt*-EGE) polyester, as well as p(PA-*alt*-SO), a novel polyester material.

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## Introduction

The production of sustainable materials is currently of key importance for society for replacement of petroleum-based materials.<sup>1</sup> In this regard, aliphatic polyesters are promising polymers as they may incorporate sustainable and/or bio-sourced monomers and be readily degraded to non-toxic and re-usable components.<sup>2</sup> Industrially, polyesters are typically produced by polycondensation of diacids and diols and ring-opening polymerization (ROP) of cyclic esters. For instance, the production of a biodegradable polyester of current importance like polylactide (PLA) is industrially performed by ring-opening polymerization (ROP) of lactide using a Sn(II)-based initiator.<sup>3</sup>

Over the past fifteen years, the alternating ring-opening copolymerization (ROCOP) of epoxides and anhydrides has

become a powerful approach to access structurally diverse polyesters in a controlled manner using a variety of well-defined metal complexes as ROCOP initiators.<sup>4</sup> Notably, Co-, Cr-, Zn- and Al-based complexes when operating in the presence of a co-catalyst (typically a halide salt) were shown to selectively produce polyesters *via* ROCOP.<sup>4a,5</sup> More recently, due to the potential toxicity of some metal salts,<sup>6</sup> organocatalyzed and/or initiated epoxide/anhydride ROCOP has emerged as an effective and attractive tool to access well-defined metal-free polyesters materials.<sup>7</sup> Though the field is only emerging, a number of two-component catalytic systems, most notably those involving boron-based Lewis acidic organocatalysts in combination with a halide source acting as a nucleophile/co-catalyst, were demonstrated to display high activity and ester selectivity in epoxide/anhydride ROCOP.<sup>8</sup> Other two-component organocatalysts, including several ingeniously designed bifunctional borane-ammonium catalysts, organic superbases combined with alcohol/Brønsted acid sources, were also shown to mediate epoxide/anhydride ROCOP.<sup>9</sup> In contrast, ROCOPs initiated by a single component/nucleophile, most frequently halide salts and various amines,<sup>7,10</sup> remain less studied since typically thought to exhibit lower catalytic performance. In the area, a recent report showed that carbon(0) species of the type carbodicarbene may efficiently initiate epoxide/anhydride ROCOP to produce polyesters.<sup>11</sup>

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Due to their exceptional steric/electronic properties and ready synthesis,<sup>12</sup> N-heterocyclic carbenes (NHCs) are ubiquitous organic initiators able to mediate various ROP processes. Thus, acting as strong nucleophiles, they are well-established ROP initiators of various cyclic esters, and efficiently mediate the homo-polymerization of epoxides.<sup>13</sup> The use of NHCs as single component initiators of the alternating epoxides and anhydrides ROCOP remain little explored. In that regard, Buchmeiser showed that latent NHC-metal/NHC-CO<sub>2</sub> adducts, which may be seen as “protected” NHC sources, are excellent epoxide/anhydride ROCOP initiators for a broad scope of monomers.<sup>14</sup> Herein, we report that unprotected, “free” NHCs such as classical IMes, IPr and air-stable Cl-IPr carbenes (Fig. 1) may act as single component initiators for the effective and selective ROCOP of various epoxides (including bio-sourced derivatives) and anhydrides.

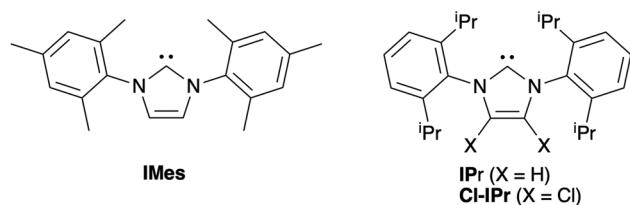


Fig. 1 NHC initiators used for epoxide/anhydride ROCOP.

## Results and discussion

ROCOP tests were first performed with CHO and PA as monomers using IPr and IMes carbenes as initiators, and all results are compiled in Table 1. For initial tests, a low monomer feed was selected (60/70/1 PA/CHO/NHC ratio) to first probe/validate the possibility of NHC-mediated ROCOP. Thus, carbene IPr was satisfyingly found to effectively mediate PA/CHO ROCOP to afford quantitatively and selectively well-defined and monodisperse p(PA-CHO) polyester after 16 h at 90 °C ( $M_n = 10\,200\text{ g mol}^{-1}$ ,  $D = 1.05$ , >98% ester linkage; entry 1, Table 1).

Optimization of the reaction showed that the ROCOP is complete within 6 h at 90 °C (entry 2, Table 1). Polymerization also proceeds efficiently when the less sterically hindered IMes carbene was used as initiator leading to the selective formation of p(PA-CHO) polyester material (entry 3, Table 1). It may also be noted that *in situ* generated IPr (from deprotonation of salt [IPr-H][BF<sub>4</sub>] by K[N(SiMe<sub>3</sub>)<sub>2</sub>], with a subsequent pentane wash) also mediates PA/CHO ROCOP with the selective production of polyester, though with lower activity relative to pure IPr (entry 2 vs. 4, Table 1). For both IPr and IMes, lowering the reaction temperature to 70 °C with an identical monomer/initiator feed ratio led to no reaction after 6 h (entries 5 and 6, Table 1). Overall, these initial results show that carbenes IMes and IPr may behave as suitable and effective nucleophiles for epoxide/anhydride ROCOP for polyester production. As a comparison,

Table 1 ROCOP of PA and CHO/PO initiated by NHCs IMes, IPr and Cl-IPr<sup>a</sup>

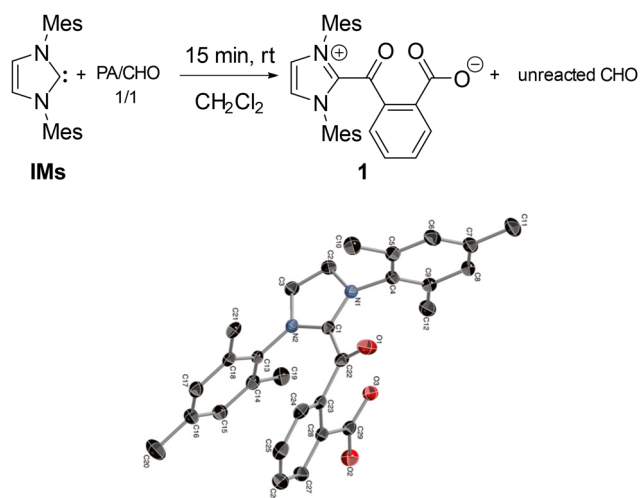
Entry	NHC	epoxide	PA/epoxide/NHC ratio	Time (h)	Conv PA <sup>b</sup> (%)	Ester linkage <sup>c</sup> (%)	$M_n^d$ (kg mol <sup>-1</sup> )	$D^d$
1	IPr	CHO	60 : 70 : 1	16	95	>98	10.2	1.05
2	IPr	CHO	60 : 70 : 1	6	95	>98	13.4	1.13
3	IMes	CHO	60 : 70 : 1	6	80	>98	13.0	1.19
4	IPr (crude) <sup>e</sup>	CHO	60 : 70 : 1	6	73	>98	10.1	1.23
5	IPr <sup>f</sup>	CHO	60 : 70 : 1	6	0	—	—	—
6	IMes <sup>f</sup>	CHO	60 : 70 : 1	6	0	—	—	—
7	[PPN]Cl	CHO	60 : 70 : 1	6	<20	—	—	—
8	IPr	CHO	250 : 300 : 1	4	72	>98	26.7	1.11
9	IPr	CHO	250 : 300 : 1	6	100	>98	32.0	1.12
10	Cl-IPr	CHO	250 : 300 : 1	6	100	>98	25.2	1.20
11	Cl-IPr <sup>h</sup>	CHO	250 : 300 : 1	16	53	>98	13.4	1.37
12	IMes-PA (1)	CHO	250 : 300 : 1	16	97	>98	29.7 <sup>g</sup>	1.17 <sup>g</sup>
13	IPr	PO	250 : 300 : 1	16	80	>98	35.0	1.21
14	IMes	PO	250 : 300 : 1	16	83	>98	33.0	1.21
15	Cl-IPr	PO	250 : 300 : 1	16	89	>98	29.2	1.22

<sup>a</sup> Conditions (unless indicated otherwise): neat conditions (no solvent), 90 °C. <sup>b</sup> Determined by <sup>1</sup>H NMR of the crude polymer mixture (integration of the polyester signals relative to those of unreacted PA). <sup>c</sup> Determined from <sup>1</sup>H NMR by integration of CH-ester signals (typically in the 4.0–5.5 ppm region) vs. CH-O polyether resonances (typically in the 3.0–3.5 ppm region). <sup>d</sup> Determined by GPC analysis (THF, room temperature) using polystyrene standards. <sup>e</sup> Crude IPr was generated by reaction of [IPr-H][BF<sub>4</sub>] with 0.9 equiv. K[N(SiMe<sub>3</sub>)<sub>2</sub>] and the crude mixture was only washed with pentane prior to ROCOP. <sup>f</sup> ROCOP performed at 70 °C. <sup>g</sup> The GPC trace is shouldered to the lower masses, possibly due to the presence of impurities leading to transesterification side reactions. <sup>h</sup> Performed under air (closed vial).



the use of a halide salt such as [PPN]Cl as an initiator under identical ROCOP conditions (60/70/1 PA/CHO/[PPN]Cl ratio, 90 °C, 16 h) only allowed low monomer conversion (<20%) and no well-defined p(PA-CHO) material could be isolated (entry 7, Table 1).

The selective formation of polyesters in the NHC-mediated ROCOP was confirmed by NMR data and mass spectrometric MALDI-TOF analysis.  $^1\text{H}$  and  $^{13}\text{C}$  NMR data of the crude and isolated materials agree in all cases with the exclusive presence of ester linkages, and thus the absence of any polyether material that would arise from CHO homopolymerization by the NHC initiator. In addition,  $^1\text{H}$  DOSY NMR data of the material produced from a 60/70/1 PA/CHO/IPr run only display signals correlating with one diffusion coefficient value, in line with the presence of one monodisperse copolymer (see ESI†). MALDI-TOF data for the ROCOP runs with IMes and IPr both display peaks alternatively spaced by CHO (98 u.a.) and PA (148 u.a.), consistent with an alternating p(PA-alt-CHO) polyester, as shown in Fig. 2 in the case of a PA/CHO ROCOP initiated by IPr. These data also agree with NHC and OH end groups at the  $\alpha$  and  $\omega$  chain ends, respectively, confirming that the NHC moiety acts as the actual ROCOP initiator and in line with a zwitterionic polymer chain growth. Taking the example of a PA/CHO ROCOP initiated by IPr (Fig. 2), a closer inspection of the MALDI-TOF peaks sequence (and associated mass values) is consistent with a polymer chain of the type [IPr-PA-p(CHO-alt-PA)-H] $^+$ , in line with a ring-opening of PA by IPr as the initiating step. ROCOP initiation through ring-opening of PA (instead of the epoxide source) contrasts with documented and well-established ROCOP processes initiated by various nucleophiles, most notably halide salts, for which the halide source typically ring-opens the epoxide as the initiating step.<sup>7,10,15</sup> To gain additional insight, stoichiometric control experiments were performed. Thus, the NMR scale reaction of one equiv. of IMes with a 1/1 PA/CHO mixture ( $\text{CD}_2\text{Cl}_2$ , RT) led to the immediate (<15 min) and quantitative formation of zwitterion IMes-PA (**1**, Scheme 1), arising from a selective nucleophilic



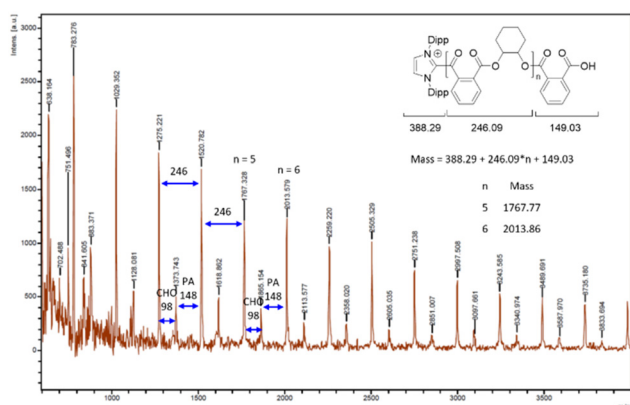
**Scheme 1** Reaction of carbene IMes with a 1/1 PA/CHO mixture and molecular structure (ORTEP view) of ring-opened product **1**.

philic ring-opening of PA by IMes, along with unreacted CHO, according to  $^1\text{H}$  and  $^{13}\text{C}$  NMR data. Species **1** was also prepared through an independent synthesis by a 1/1 PA/IMes reaction leading to its isolation in 79% yield.

In addition, the molecular structure of **1** was further established through X-ray crystallographic analysis (Scheme 1), confirming the ring-opening reaction and the zwitterionic nature of **1**. Though stable enough to be isolated, species **1** exhibits a limited stability in  $\text{CD}_2\text{Cl}_2$  (room temperature) overtime and slowly decomposes to unidentified products over the course of several days ( $t_{1/2} \approx 4$  days) as monitored by  $^1\text{H}$  NMR. An identical ring opening reactivity was observed upon mixing IPr or IPr-Cl with 1 equiv. of PA, with the fast and complete formation of the corresponding ring opening products ( $\text{CD}_2\text{Cl}_2$ , 15 min, RT) according to  $^1\text{H}$  NMR data, thus indicating that all three studied NHCs allow fast and complete PA ring opening. However, in the case of the more sterically bulky IPr and IPr-Cl, the ring opened products are highly unstable ( $\text{CD}_2\text{Cl}_2$ , room temperature) and decompose within a few hours to unknown species, precluding their isolation in a pure form.

DFT calculations were also performed to rationalize the observed higher reactivity of IMes with PA (vs. CHO) leading to **1** (with A, B, C, D and E as models of IMes, PA, CHO, **1**, and IMes-CHO ring-opened product, respectively; Fig. 3). As depicted in Fig. 3, the nucleophilic attack by A of a PA  $\text{C}=\text{O}$  moiety proceeds *via* a concerted low energy transition state (TS) and thus a low energy barrier ( $\Delta\Delta G = 5.1 \text{ kcal mol}^{-1}$ ) to afford the ring-opened product D, which is in line with the experimentally observed fast formation of **1** at room temperature. In contrast, the energy barrier for a nucleophilic attack of model carbene A on CHO model C to form E was computed to occur with a much higher energy transition state (TS',  $\Delta G = 26.1 \text{ kcal mol}^{-1}$ ), rationalizing the exclusive ring-opening of PA by IMes at room temperature.

Next, we further studied the present NHC-initiated ROCOP to probe its scope/performances (entries 8–15, Table 1).



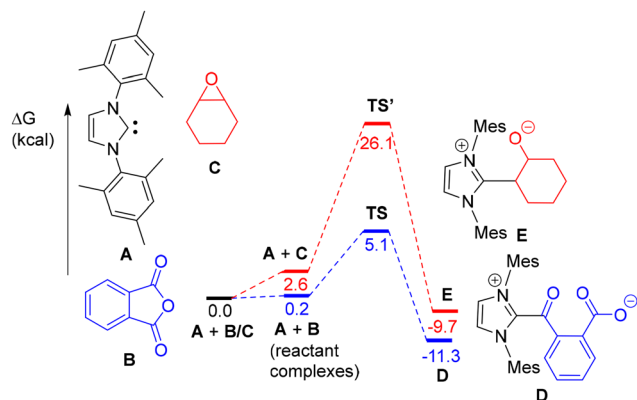


Fig. 3 DFT-estimated reactions (GAUSSIAN 16/ $\omega$ B97XD) of IMes model A with PA (B) and CHO (C) to form the ring-opened products D and E, respectively.

Increasing monomer feed to a 250/300/1 PA/CHO/NHC ratio did not alter ROCOP performances, leading to the quantitative and selective formation of p(PA-*alt*-CHO) with a higher number average molecular mass ( $M_n = 32.0 \text{ kg mol}^{-1}$ , entry 9, Table 1). Air-stable Cl-IPr (less nucleophilic than IPr and IMes) may also effectively be used as ROCOP initiator with the quantitative production of p(PA-*alt*-CHO) after 16 h at 90 °C (entry 10, Table 1). When performed under air (16 h, 90 °C), ROCOP with IPr-Cl also selectively produced polyester, albeit with a significantly lower catalytic activity than under inert atmosphere (entry 10 *vs.* 11, Table 1). The ring-opened product 1 logically performs similarly to the NHC systems with the isolation of a polyester material ( $M_n = 29.1 \text{ kg mol}^{-1}$ ,  $D = 1.17$ ; entry 12, Table 1). In addition, using PO as an epoxide source, all three NHCs initiators also mediate PA/PO ROCOP (250/300/

1, 16 h, 90 °C), with high PA conversion (80 to 89%) to exclusively afford regioirregular p(PA-*alt*-PO) polyesters ( $M_n = 29.2\text{--}35 \text{ kg mol}^{-1}$ ,  $D = 1.21\text{--}1.22$ ; entries 13–15, Table 1).

The present NHC systems were also tested in SA/CHO (SA: succinic anhydride) and SA/PO ROCOP and all results are compiled in Table S2 (ESI†). As deduced from GPC and NMR data, IMes, IPr and IPrCl carbenes polymerize a 60/70 or 250/300 SA/CHO mixture to produce p(SA-*alt*-CHO) polyester ( $M_n = 1.5$  to  $5.8 \text{ kg mol}^{-1}$ ,  $D = 1.32$  to  $1.79$ ) with a good ester linkage selectivity (80 to 88%) and high SA conversion (82 to 86%) after 16 h at 90 °C under bulk conditions. Thus, going from PA to SA as anhydride source decreased ester linkage selectivity and only afforded lower molecular weight p(SA-*alt*-CHO) materials, as frequently observed.<sup>4b,7</sup> This can be ascribed to the lower reactivity of the less ring-strained SA (*vs.* PA), resulting in a decreased ROCOP activity as well as a possible competition with CHO homopolymerization (leading to polyether segments) and thus a decreased ester linkage selectivity. However, a SA/PO ROCOP test with IPr as initiator (250/300/1 SA/PO/IPr ratio) exclusively led the high yield production of well-defined p(SA-*alt*-PO) polyester (96% SA conversion, ester linkage >98%,  $M_n = 5.92 \text{ kg mol}^{-1}$ ,  $D = 1.82$ ). Since PO is a significantly less reactive monomer than CHO towards ring-opening due to lower ring strain, PO homopolymerization may not compete with the PO/SA ROCOP process, thus the observed higher ester linkage selectivity in PO/SA *vs.* CHO/SA ROCOP.

Last, to further expand the scope of the present NHC systems, carbene IPr was exploited as ROCOP initiator of PA and bio-sourced epoxides eugenyl glycidyl ether and safrole glycidylether (EGE, SO, Table 2). Epoxides EGE and SO, which are derived from natural products eugenol and safrole, respectively, were prepared according to literature procedures.<sup>16</sup> As an initial run, IPr was found to polymerize a 70/85 PA/EGE

Table 2 PA/EGE and PA/SO ROCOP initiated by IPr carbene<sup>a</sup>

Entry	Epoxide	PA/epoxide/IPr ratio	Time (h)	Conv PA <sup>b</sup> (%)	Ester linkage <sup>c</sup> (%)	$M_n^d$ (kg mol <sup>-1</sup> )	$D^d$
1	EGE	70 : 85 : 1	5	83	>98	12.1	1.20
2	EGE	250 : 300 : 1	16	>98	>98	43.5	1.60
3	SO	250 : 300 : 1	16	85	>98	45.3	1.30

<sup>a</sup> Conditions (unless indicated otherwise): neat conditions (no solvent), 90 °C. <sup>b</sup> Determined by <sup>1</sup>H NMR of the crude polymer mixture (integration of the polyester signals relative to those of unreacted PA). <sup>c</sup> Determined from <sup>1</sup>H NMR by integration of CH-ester signals (typically in the 4.0–5.5 ppm region) *vs.* CH–O polyether resonances (typically in the 3.0–3.5 ppm region). <sup>d</sup> Determined by GPC analysis (THF, room temperature) using polystyrene standards.







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