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1 Introduction

The acceptorless dehydrogenation and hydrogenation of Nheterocycles are key atom-economical and fundamental methods to afford various imine and amine derivatives with potential bioactivities.¹⁻¹¹ Additionally, the two kinds of redox reactions, especially the reversible acceptorless dehydrogenation and hydrogenation of N-heterocycles, involve the release and acceptance of H₂, which indicate that N-heterocycles are significant chemical hydrogen-storage materials. In fact, Nheterocycles have already been verified as an important type of chemical hydrogen-storage materials.^{12–19} The acceptorless dehydrogenation and hydrogenation of N-heterocycles as organic hydrogen carriers have been broadly studied, mainly focusing on the catalyst synthesis and design, investigation of the redox mechanisms, and extension of substrate scope.^{3–19}

Since the chemical processes for the acceptorless dehydrogenation and hydrogenation of N-heterocycles involve H_2 release and acceptance, therefore, the thermodynamics of the

Thermodynamic evaluations of the acceptorless dehydrogenation and hydrogenation of prearomatic and aromatic N-heterocycles in acetonitrile

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N-heterocycles are important chemical hydrogen-storage materials, and the acceptorless dehydrogenation and hydrogenation of N-heterocycles as organic hydrogen carriers have been widely studied, with the main focus on the catalyst synthesis and design, investigation of the redox mechanisms, and extension of substrate scope. In this work, the Gibbs free energies of the dehydrogenation of pre-aromatic N-heterocycles (YH₂) and the hydrogenation of aromatic N-heterocycles (Y), *i.e.*, $\Delta G_{H_2R}(YH_2)$ and $\Delta G_{H_2A}(Y)$, were derived by constructing thermodynamic cycles using Hess' law. The thermodynamic abilities for the acceptorless dehydrogenation and hydrogenation of 78 pre-aromatic N-heterocycles (YH₂) and related 78 aromatic N-heterocycles (Y) were well evaluated and discussed in acetonitrile. Moreover, the applications of the two thermodynamic parameters in identifying pre-aromatic N-heterocycles possessing reversible dehydrogen reductants in catalytic hydrogenation were considered and are discussed in detail. Undoubtedly, this work focuses on two new thermodynamic parameters of pre-aromatic and aromatic N-heterocycles, namely $\Delta G_{H_2R}(YH_2)$ and $\Delta G_{H_2A}(Y)$, which are important supplements to our previous work to offer precise insights into the chemical hydrogen storage and hydrogenation reactions of pre-aromatic and aromatic N-heterocycles.

dehydrogenation and hydrogenation of N-heterocycles are important thermodynamic parameters to evaluate the hydrogen-storage abilities and hydrogenation abilities of Nheterocycles, especially for pre-aromatic N-heterocycles possessing reversible dehydrogenation and hydrogenation abilities.³⁻⁸ This work follows on from our previous study into the thermodynamics of hydrogen transfer for amines.²⁰ This present work focuses on thermodynamic evaluations of the acceptorless dehydrogenation and hydrogenation of a special category of N-heterocycles, namely, pre-aromatic and aromatic N-heterocycles, in acetonitrile.

In our previous research work, we computed the Gibbs free energies of 84 amines (**YH**₂), including 78 pre-aromatic Nheterocycles and 6 general amines, releasing hydrides and their activation free energies for hydride self-exchange reactions, *i.e.*, $\Delta G_{\text{H}^-\text{R}}(\text{YH}_2)$ and $\Delta G^{\ddagger}(\text{YH}_2/\text{YH}^{+})$, using density functional theory (DFT) calculations in acetonitrile.^{20a} Earlier in 2023, thermodynamic evaluations were conducted of 84 amines releasing two hydrogen ions (H⁻ + H⁺) and the corresponding 84 imines accepting two hydrogen ions (H⁻ + H⁺), including $\Delta G_{\text{H}^-\text{PR}}(\text{YH}_2)$ and $\Delta G_{\text{H}^-\text{PA}}(\text{Y})$.^{20b} Combined with the above thermodynamic data, in this work, the thermodynamic values upon H₂ release and acceptance for the dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles

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(Scheme 1) were naturally derived *via* constructing thermodynamic cycles using Hess' law,²¹ and the thermodynamic abilities of the pre-aromatic N-heterocycles as chemical hydrogenstorage materials were well evaluated and compared. Moreover, application of the thermodynamic data for identifying prearomatic N-heterocycles possessing reversible dehydrogenation and hydrogenation abilities³⁻⁸ were investigated and discussed in detail. Undoubtedly, this work focuses on new thermodynamic parameters, which is an important supplement to our previous work to offer precise insights into the chemical hydrogen storage and hydrogenation reactions of pre-aromatic N-heterocycles.¹⁻¹⁹

2 Results and discussion

In this work, 78 various pre-aromatic N-heterocycles $(1H_2-78H_2)$ were designed to expand the scope of the investigated N-heterocyclic substrates, and 6 general amines $(79H_2-84H_2)$ were designed for comparison with pre-aromatic N-heterocycles. The chemical structures of 84 amines (YH_2) , including the 78 pre-aromatic N-heterocycles $(1H_2-78H_2)$ and 6 general amines $(79H_2-84H_2)$, are shown in Scheme 2.

To obtain the Gibbs free energy of YH₂ releasing H₂, the constructed thermodynamic cycle based on the processes of YH_2 releasing two hydrogen ions or H_2 in acetonitrile is shown in Scheme 3. As can be seen from Scheme 3, Step 1 is the chemical process of YH_2 releasing H_2 , $YH_2 \rightarrow Y + H_2$, and the thermodynamic driving force of YH2 dehydrogenation could be described by the Gibbs free energy of YH_2 releasing H_2 , ΔG_{H_2} - $_{R}$ (YH₂). Step 2 is the chemical process of YH₂ releasing two hydrogen ions, $YH_2 \, \rightarrow \, Y \, + \, H^- \, + \, H^+\!,$ and the corresponding thermodynamic driving force was defined as the Gibbs free energy of YH₂ releasing two hydrogen ions, ΔG_{H^-PR} (YH₂).²⁰ Step 3 is the chemical process of H^- reacting with H^+ to form H_2 , H^- + H⁺ \rightarrow H₂, and the related thermodynamic driving force was defined as the Gibbs free energy of one molar H⁻ reacting with one molar H^+ to generate one molar H_2 , $\Delta G_{PA}(H^-)$. The $\Delta G_{\rm PA}({\rm H}^{-})$ value was reported as -76.0 kcal mol⁻¹ in acetonitrile.²²⁻²⁴ Therefore, the value of $\Delta G_{H,R}(YH_2)$ could be calculated using eqn (1) in Table 1 by establishing a thermodynamic cycle according to Hess' law (Scheme 3),²¹ where $\Delta G_{\rm H,-}$ $_{\rm R}({\rm YH}_2) = \Delta G_{{\rm H}^-{\rm PR}}({\rm YH}_2) + \Delta G_{{\rm PA}}({\rm H}^-)$ (eqn (1)). For eqn (1), $\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathrm{YH}_2)$ values are available from our previous work and are displayed in the fourth column of Table 2. Since the $\Delta G_{\mathrm{H}^{-R}}(\mathrm{YH}_2)$ values were computed using the DFT method with



Scheme 1 Chemical equations for the acceptorless dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles.

a precision of 1.1 kcal mol⁻¹ in our previous work,^{20a} and the pK_a values of YH⁺ were predicted using XGBoost with a 0.87 pK_a error,^{20a} the $\Delta G_{H_2R}(\mathbf{YH}_2)$ values were derived from $\Delta G_{H^-R}(\mathbf{YH}_2)$ and $pK_a(YH^+)$, whereby $\Delta G_{H_2R}(\mathbf{YH}_2) = \Delta G_{H^-R}(\mathbf{YH}_2) + 1.37 pK_a(YH^+)$, and therefore, the precision of the $\Delta G_{H_2R}(\mathbf{YH}_2)$ values could be estimated within 2.3 kcal mol⁻¹, which was a suitable precision to give practical guidance on the chemical hydrogen storage and hydrogenation reactions, *etc.*

Step 4 in Scheme 2 is the chemical process of **Y** accepting H_2 to offer \mathbf{YH}_2 , $\mathbf{Y} + H_2 \rightarrow \mathbf{YH}_2$, and the corresponding thermodynamic driving force could be described as the Gibbs free energy of **Y** accepting H_2 to generate \mathbf{YH}_2 , $\Delta G_{H_2A}(\mathbf{Y})$. Since the **Y** hydrogenation and \mathbf{YH}_2 dehydrogenation are reverse chemical reactions, the $\Delta G_{H_2A}(\mathbf{Y})$ value was the opposite of the $\Delta G_{H_2R}(\mathbf{YH}_2)$ value,²¹ *i.e.*, $\Delta G_{H_2A}(\mathbf{Y}) = -\Delta G_{H_2R}(\mathbf{YH}_2)$ (eqn (2) in Table 1). Herein, the chemical processes, thermodynamic parameters, and data sources or computed equations of step 1–step 4 for \mathbf{YH}_2 dehydrogenation and **Y** hydrogenation in acetonitrile are presented in Table 1, meanwhile, the $\Delta G_{H^-PR}(\mathbf{YH}_2)$, $\Delta G_{H^-PA}(\mathbf{Y})$, $\Delta G_{H_2R}(\mathbf{YH}_2)$, and $\Delta G_{H_2A}(\mathbf{Y})$ values of the 84 amines (\mathbf{YH}_2) dehydrogenation and their relevant imines (**Y**) hydrogenation in acetonitrile are listed in Table 2.

2.1. Thermodynamic abilities of YH₂ dehydrogenation as organic hydrogen-storage materials

For better comparing the relationship between the thermodynamic abilities and chemical structures, all 84 amines (YH_2) were divided into 4 groups,²⁰ comprising 1,4-dihydropyridine compounds $(Y_IH_2, 1H_2-62H_2)$, 1,2-dihydropyridines $(Y_{II}H_2,$ $63H_2-73H_2)$, 1,2-dihydro-3-substituted-indoline analogs $(Y_{III}H,$ $74H_2-78H_2)$, and general amines $(Y_{IV}H_2, 79H_2-84H_2)$ (Scheme 4). Among the 4 groups, Y_IH_2 , $Y_{II}H_2$, and $Y_{III}H_2$ were 1,4-dihydro nitric six-member heterocycles, 1,2-dihydro nitric six-member heterocycles, and 1,2-dihydro nitric five-member heterocycles, respectively, which all belong to pre-aromatic N-heterocycles. As pre-aromatic N-heterocycles, the resulting dehydrogenation products Y (Y_I , Y_{II} , and Y_{III}) are aromatic N-heterocycles (Scheme 4) after YH_2 releases H_2 , $YH_2 \rightarrow Y + H_2$.

According to the definition of $\Delta G_{H_2R}(\mathbf{YH}_2)$, if the $\Delta G_{H_2R}(\mathbf{YH}_2)$ value is more negative than 0, $\Delta G_{H_2R}(\mathbf{YH}_2) < 0$, the chemical process of \mathbf{YH}_2 releasing H_2 is thermodynamically favorable, and \mathbf{YH}_2 is recognized as a thermodynamically excellent H_2 donor.²¹ In contrast, if the $\Delta G_{H_2R}(\mathbf{YH}_2)$ value is greater than 0, $\Delta G_{H_2R}(\mathbf{YH}_2) > 0$, the chemical process of \mathbf{YH}_2 releasing H_2 is thermodynamically unfavorable, and \mathbf{YH}_2 is not a thermodynamically feasible H_2 donor.

As can be seen from Table 2, the $\Delta G_{H_2R}(YH_2)$ scale of the 78 pre-aromatic N-heterocycles $(1H_2-78H_2)$ investigated in this work ranged from -15.8 kcal mol⁻¹ to 22.0 kcal mol⁻¹, which spanned a very wide thermodynamic range of 37.8 kcal mol⁻¹. Moreover, among the 78 pre-aromatic N-heterocycles $(1H_2-78H_2)$, $21H_2$ from Y_1H_2 was thermodynamically the best hydrogen donor or carrier $(-15.8 \text{ kcal mol}^{-1})$, even better than HCO₂H $(-5.9 \text{ kcal mol}^{-1})$,²⁵ while $44H_2$ from Y_1H_2 was thermodynamically the worst hydrogen donor or carrier (22.0 kcal mol^{-1}). For more refined thermodynamic analysis,



Scheme 2 Chemical structures of 84 amines (YH_2), including the 78 pre-aromatic N-heterocycles ($1H_2-78H_2$) and 6 general amines ($79H_2-84H_2$) investigated in this work.



Scheme 3 Constructed thermodynamic cycle based on the processes of YH_2 releasing two hydrogen ions or H_2 in acetonitrile.

the distribution of $\Delta G_{\mathrm{H_2R}}(\mathbf{YH_2})$ values for the 78 pre-aromatic Nheterocycles ($\mathbf{1H_2}$ -78 $\mathbf{H_2}$) in every 5 kcal mol⁻¹ is clearly shown in Fig. 1 with $\mathbf{YH_2}$ amounts as the ordinate and $\Delta G_{\mathrm{H_2R}}(\mathbf{YH_2})$ ranges as the abscissa.

From Fig. 1, several interesting conclusions could be drawn as follows. (1) The distribution of YH₂ amounts exhibited an excellent normal distribution, and $\Delta G_{H,R}(YH_2)$ values of prearomatic N-heterocycles ranging from 0 to 5 kcal mol^{-1} were the most common (29.5%). (2) It was found that the $\Delta G_{H,R}(\mathbf{YH}_2)$ values of 30 YH2 were more negative than 0, meaning that 30 YH₂ were thermodynamically feasible H₂ donors, and belonged to potential chemical hydrogen-storage materials. While the $\Delta G_{H_{2}R}(\mathbf{YH}_{2})$ values of 48 \mathbf{YH}_{2} were greater than 0, indicating that the 48 YH_2 were thermodynamically unfeasible H_2 donors. (3) It was also discovered that 75 $\Delta G_{H,R}(YH_2)$ values ranged from -15 kcal mol⁻¹ to 15 kcal mol⁻¹, indicating that \sim 96% of the thermodynamic driving forces for the pre-aromatic Nheterocycles dehydrogenations were between -15 kcal mol⁻¹ to 15 kcal mol⁻¹, while less than 4% of the thermodynamic driving forces for the pre-aromatic N-heterocycles dehydrogenation were more negative than -15 kcal mol⁻¹ or greater than 15 kcal mol⁻¹. Accordingly, for an unknown pre-aromatic Nheterocycle, the thermodynamic driving force for YH2 dehydrogenation is generally between -15 kcal mol⁻¹ to 15 kcal mol⁻¹ (~96% possibilities).

Table 1 Chemical processes, thermodynamic parameters, and data sources or computed equations of step 1-step 4 for YH₂ dehydrogenation and Y hydrogenation in acetonitrile

Chemical processes		Thermodynamic parameters	Sources or computed equations	Equation X	
Step 1	$\mathbf{YH}_2 \rightarrow \mathbf{Y} + \mathbf{H}_2$	$\Delta G_{\rm H, P}({\bf YH}_2)$	$\Delta G_{\rm HLP}(\mathbf{Y}\mathbf{H}_{2}) = \Delta G_{\rm HLPP}(\mathbf{Y}\mathbf{H}_{2}) + \Delta G_{\rm PA}(\mathbf{H}^{-})$	1	
Step 2	$\mathbf{YH}_2 \rightarrow \mathbf{Y} + \mathbf{H}^- + \mathbf{H}^+$	$\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathrm{YH}_{2})$	Ref. 20	_	
Step 3	$H^{-} + H^{+} \rightarrow H_{2}$	$\Delta G_{\rm PA}({\rm H}^{-})$	$-76.0 \text{ kcal mol}^{-1}$ (ref. 22–24)	_	
Step 4	$\mathbf{Y} + \mathbf{H}_2 \rightarrow \mathbf{Y}\mathbf{H}_2$	$\Delta G_{\mathrm{H}_{2}\mathrm{A}}(\mathbf{Y})$	$\Delta G_{\mathbf{H}_{2}\mathbf{A}}(\mathbf{Y}) = -\Delta G_{\mathbf{H}_{2}\mathbf{R}}(\mathbf{Y}\mathbf{H}_{2})$	2	

 $\label{eq:control} \mbox{Table 2} \quad \Delta G_{H^-PR}(YH_2), \ \Delta G_{H^-PA}(Y), \ \Delta G_{H_2R}(YH_2), \ \mbox{and} \ \ \Delta G_{H_2A}(Y) \ \ \mbox{values of} \qquad \mbox{Table 2} \quad (Contd.)$ the 84 considered amines (YH_2) and their corresponding imines (Y) in ace

acetonitrile (I	unit: kcal mol ⁻¹)			-				$\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathbf{YH_2})^a$	$\Delta G_{\mathrm{H}_{2}\mathrm{R}}(\mathrm{YH}_{2})^{b}$
			$\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathrm{YH}_2)^a$	$\Delta G_{\mathrm{H_2R}}(\mathbf{YH_2})^b$	Compounds	Structures	G	$-\Delta G_{\mathrm{H}^{-}\mathrm{PA}}(\mathbf{Y})^{a}$	$-\Delta G_{\mathrm{H_2A}}(\mathbf{Y})^c$
Compounds	Structures	G	$-\Delta G_{\mathrm{H}^-\mathrm{PA}}(\mathbf{Y})^a$	$-\Delta G_{\mathrm{H}_{2}\mathrm{A}}(\mathbf{Y})^{c}$	44H ₂		NO	98.0	22.0
					45H2		NO_2	93.2	17.2
	нун				46H ₂	HG	CH ₃	73.5	-2.5
11.		_	68.0	-8.0	47H ₂		^t Bu	84.9	8.9
1112	N N		08.0	-0.0	48H ₂		Ph	74.9	-1.1
2H ₂		CH_3	68.8	-7.2		Ĥ.			
3H2	нн	^t Bu	70.6	-5.4	49H ₂		CH_3	76.2	0.2
4H ₂	X	Ph	66.5	-9.5	50H ₂	нн	^t Bu	80.1	4.1
5H2		NH_2	63.5	-12.5	$51H_2$	\sim	Ph	75.9	-0.1
6H ₂	N°G	СНО	75.6	-0.4	$52H_2$		NH_2	74.7	-1.3
7H ₂		CN	75.2	-0.8	53H ₂	Gʻ~N	СНО	79.0	3.0
8H2		NO	78.9	2.9	$54H_2$	п	CN	80.6	4.6
9H ₂		NO_2	71.7	-4.3	55H ₂		NO	80.4	4.4
10H ₂		CH ₂	69.7	-6.3	56H ₂		NO_2	81.2	5.2
11H ₂	нн	^t Bu	71.8	-4.2	2	нн	2		
12H ₂	G C	Ph	70.2	-5.8					
13H ₂	Į J	NHa	67.8	-8.2	57H ₂		_	67.6	-8.4
14H ₂	N [×]	CHO	80.6	4.6		✓ N ¹			
15H ₂	н	CN	78 7	2.7		п			
16H ₂		NO	87.3	11.3					
17H ₂		NOa	85.7	9.7	58Ha		н	87 5	11.5
18H ₂		CH ₂	67.4	-8.6	30112	Ľ ∕ <mark>Ņ</mark> ∕Ľ ∕∕	11	07.5	11.5
19H ₂			72.0	-4.0		Ĥ			
20H	H G	Ph	64.8	-11.2	59H ₂	G H	CH_3	87.5	11.5
21H		NH	60.2	-15.8	60H ₂	\sim	Ph	87.3	11.3
22H ₂	N [×]	CHO	72.3	-3.7					
23H	п	CN	74.1	-19		✓ N ✓			
24H		NO	76.6	0.6					
25H2		NO	84.6	8.6		H H			
20112	нн	1102	0110	0.0	61H.		_	76.3	0.3
	MeO ₂ CCO ₂ Me				01112	^N		70.0	0.0
26H ₂		—	83.1	7.1		Ĥ			
	H ₃ C N CH ₃					ннО			
						NH			
	EtO ₂ C、X,CO ₂ Et				62H ₂		_	87.6	11.6
27H ₂	ΥΎ	_	83.3	7.3		- N N O CH₃ H			
	н₃С́Ņ́СН₃								
	н					Хн			
					63H ₂	N N	_	64.5	-11.5
28H.			83.4	74					
20112	н₃С <mark>́М</mark> ́СН₃		03.4	7.4		нн			
	Ĥ Î				64H.	NXNH	_	73.7	-2.3
	н <mark>н</mark>				01112	i j			210
						~			
29H ₂		—	77.9	1.9					
	Ĥ				65H ₂	N TH	_	74.6	-1.4
30H.		CH	75.8	-0.2		н ^и			
31H	u u		79.5	3 5		нн			
32H ₂		Ph	75.9	-0.1	66H2	M ^H		77.6	1.6
33H-		NH ₂	69.6	-6.4	2	L J			
34H ₂	`` <mark>Ņ</mark> ´``G	CHO	86.0	10.0		~ ~			
35H ₂	н	CN	86.4	10.4		, × н			
36H_		NO	91.0	15.0	67H-	ſ ヽヽヽ N´`'	_	79.6	3.6
37H2		NO ₂	81.8	5.8	0711 <u>2</u>				5.0
38H2	нн	CH ₂	76.7	0.7					
39H ₂	G	^t Bu	81.1	5.1		H Ph			
40H ₂	L I	Ph	78.8	2.8		M H			
41H ₂	N N	NH_2	76.8	0.8	68H ₂		_	78.4	2.4
42H ₂	п	Сно	89.4	13.4		Ĩ [
43H ₂		CN	87.3	11.3		\sim			

Table 2 (Contd.)

			$\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathrm{YH}_{2})^{a}$	$\Delta G_{\mathrm{H}_{2}\mathrm{R}}(\mathbf{Y}\mathbf{H}_{2})^{b}$
Compounds	Structures	G	$-\Delta G_{\mathrm{H}^-\mathrm{PA}}(\mathbf{Y})^a$	$-\Delta G_{\mathrm{H_2A}}(\mathbf{Y})^c$
69H ₂	H H N-H	_	84.1	8.1
70H ₂	H H H H H	_	76.6	0.6
71H ₂	HN H	_	80.8	4.8
72 H ₂	HZ H H N-H		77.0	1.0
73H ₂	H H N-H	_	84.5	8.5
74H ₂	N Ph	_	68.6	-7.4
75H ₂	N H	_	85.5	9.5
76H ₂	N H O Ph	_	78.2	2.2
77H ₂	K K K K K K K K K K K K K K K K K K K	_	82.7	6.7
78H ₂	N H N H	_	76.9	0.9
79H ₂	H _N ^H H ₃ C H ₃ CH ₃	_	86.3	10.3
80H ₂	$H_{N'}^{H}$ Ph H_{H}^{H} CH ₃	_	87.6	11.6
81H ₂	H _N , Ph H ₃ C H CH ₃	_	91.2	15.2
82H ₂	H _N , Ph Ph H	_	93.8	17.8
83H ₂	H _N ,Ph Ph CH₃	_	93.7	17.7

Table 2	(Contd.)
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			$\Delta G_{\mathrm{H}^{-}\mathrm{PR}}(\mathrm{YH}_{2})^{a}$	$\Delta G_{\mathrm{H}_{2}\mathrm{R}}(\mathbf{Y}\mathbf{H}_{2})^{b}$
Compounds	Structures	G	$-\Delta G_{\mathrm{H}^-\mathrm{PA}}(\mathbf{Y})^a$	$-\Delta G_{\mathrm{H}_{2}\mathrm{A}}(\mathbf{Y})^{c}$
84H ₂		_	94.3	18.3

^{*a*} $\Delta G_{\mathrm{H}^-\mathrm{PR}}(\mathbf{YH}_2)$ and $\Delta G_{\mathrm{H}^-\mathrm{PA}}(\mathbf{Y})$ values are from ref. 20. ^{*b*} $\Delta G_{\mathrm{H}_2\mathrm{R}}(\mathbf{YH}_2) = \Delta G_{\mathrm{H}^-\mathrm{PR}}(\mathbf{YH}_2) + \Delta G_{\mathrm{PA}}(\mathrm{H}^-)$. ^{*c*} $\Delta G_{\mathrm{H}_2\mathrm{A}}(\mathbf{Y}) = -\Delta G_{\mathrm{H}_2\mathrm{R}}(\mathbf{YH}_2)$.

To clearly reveal the relationship between the structural features and thermodynamic abilities, the $\Delta G_{\text{H}_2\text{R}}(\text{YH}_2)$ scales of 4 groups of YH₂, along with the $\Delta G_{\text{H}_2\text{R}}(\text{XH}_2)$ values of HCO₂H, H₂, **67H₂**, **78H₂**, and HEH₂ (Hantzsch ester, **27H₂**) releasing H₂ in acetonitrile are shown in Scheme 5,^{25–27} because HCO₂H,²⁸ H₂,^{29–33} **67H₂**,^{34,35} **78H₂**,^{36,37} and HEH₂ (**27H₂**)^{38–40} are the most frequently-used hydrogen reductants (denoted as XH₂) used in the hydrogenation reactions.

From Scheme 5, the following valuable conclusions could be made: (1) the $\Delta G_{H_2R}(\mathbf{Y}_{\mathbf{I}}\mathbf{H}_2)$ scale covered from -15.8 kcal mol⁻¹ to 22.0 kcal mol⁻¹, which spanned the widest thermodynamic range by 37.8 kcal mol⁻¹ among the 4 groups of amines Y_IH_2 - $\mathbf{Y}_{IV}\mathbf{H}_2$; (2) according to the $\Delta G_{H,R}(\mathbf{YH}_2)$ scales of $\mathbf{Y}_{II}\mathbf{H}_2$ (-11.5 to 8.5 kcal mol⁻¹), $Y_{III}H_2$ (-7.4 to 9.5 kcal mol⁻¹), and $Y_{IV}H_2$ (10.3-18.3 kcal mol^{-1}), the dehydrogenation abilities decreased in the order of $Y_{II}H_2 \approx Y_{III}H_2 > Y_{IV}H_2$; (3) in view of the relations between sets, it was found that the $\Delta G_{H,R}(YH_2)$ values displayed the following regular pattern of $\{\Delta G_{H,R}(\mathbf{Y}_{II}\mathbf{H}_2) \cup \Delta G_{H,R}(\mathbf{Y}_{III}\mathbf{H}_2)\}$ $\bigcup \Delta G_{H_1R}(\mathbf{Y}_{IV}\mathbf{H}_2) \subseteq \Delta G_{H_1R}(\mathbf{Y}_{I}\mathbf{H}_2);$ (4) for the pre-aromatic Nheterocycles, the $\Delta G_{H_2R}(YH_2)$ scales ranged from -15.8 to 22.0 kcal mol⁻¹ for Y_IH₂, from -11.5 to 8.5 kcal mol⁻¹ for Y_{II}H₂, and from -7.4 to 9.5 kcal mol⁻¹ for Y_{III}H₂, respectively. Since the $\Delta G_{H_1R}(YH_2)$ scales of Y_1H_2 , $Y_{II}H_2$, and $Y_{III}H_2$ crossed negative and positive values $(-15.8 \text{ to } 22.0 \text{ kcal mol}^{-1})$, it was indicated that not all the acceptorless dehydrogenation of prearomatic N-heterocycles is a thermodynamically uphill or downhill process under ambient conditions,³⁻¹⁹ and not all prearomatic N-heterocycles are thermodynamically feasible to serve as chemical hydrogen-storage materials. $\Delta G_{H,R}(YH_2)$ is absolutely an important thermodynamic parameter to guide chemists to discover more potentially excellent chemical hydrogen-storage materials.

Initially, we tried to explain the effects of substituents on $\Delta G_{\text{H}_2\text{R}}(\mathbf{YH}_2)$, but we failed to draw a meaningful conclusion. For example, an electron-withdrawing group could decrease the hydride-donating ability of \mathbf{YH}_2 , $\Delta G_{\text{H}^-\text{R}}(\mathbf{YH}_2)$, while it increased the proton-donating ability of \mathbf{YH}^+ , $\Delta G_{\text{PR}}(\mathbf{YH}^+)$. Since $\Delta G_{\text{H}_2\text{R}}(\mathbf{YH}_2)$ is derived from $\Delta G_{\text{H}^-\text{R}}(\mathbf{YH}_2)$ and $\Delta G_{\text{PR}}(\mathbf{YH}^+)$ essentially, *i.e.*, $\Delta G_{\text{H}_2\text{R}}(\mathbf{YH}_2) = \Delta G_{\text{H}^-\text{R}}(\mathbf{YH}_2) + \Delta G_{\text{PR}}(\mathbf{YH}^+) + \Delta G_{\text{PA}}(\text{H}^-)$, for \mathbf{YH}_2 possessing an electron-withdrawing group, whether the $\Delta G_{\text{H}_2-\text{R}}(\mathbf{YH}_2)$ value increases or decreases depends on the *D*-values of $\Delta G_{\text{H}^-\text{R}}(\mathbf{YH}_2)$ decreasing and $\Delta G_{\text{PR}}(\text{YH}^+)$ increasing, *i.e.*,



Scheme 4 Classifications of the 84 amines (YH₂) and imines (Y) and the corresponding 4 groups of YH₂ and Y, respectively, considered in this work



Fig. 1 Distributions of $\Delta G_{H_2R}(YH_2)$ ranges for the 78 pre-aromatic N-heterocycles in every 5 kcal mol^{-1} from -20 kcal mol^{-1} to 25 kcal mol^{-1}.

 $\Delta\Delta G_{H^-R}(YH_2)$ and $\Delta\Delta G_{PR}(YH^+)$. In addition, the steric factors from the substituents will have a significant effect on reaction kinetics, and a slight effect on the reaction thermodynamics.

2.2. Thermodynamic abilities of Y hydrogenation to generate YH₂

According to the physical meaning of $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$, if the $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$ value is more negative than 0, $\Delta G_{\text{H}_3\text{A}}(\mathbf{Y}) < 0$, the **Y** hydrogenation

by H₂ is thermodynamically favorable. The more negative the $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$ value, the larger the thermodynamic driving force of \mathbf{Y} hydrogenation by H₂. While if the $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$ value is greater than 0, $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y}) > 0$, the \mathbf{Y} hydrogenation by H₂ is thermodynamically unfavorable, and the larger the $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$ value, the smaller the thermodynamic driving force of \mathbf{Y} hydrogenation by H₂. From Table 2, among the 84 imines (\mathbf{Y}), aromatic N-heterocycle **44** (-22.0 kcal mol⁻¹) had the largest thermodynamic driving force to accept H₂ producing **44H**₂, while the aromatic N-heterocycle **21** (15.8 kcal mol⁻¹) had the smallest thermodynamic driving force to obtain H₂ generating **21H**₂.

As for the $\Delta G_{H,A}(\mathbf{Y})$ values, we can safely draw some interesting conclusions to inform the hydrogenation reactions of imines, including aromatic N-heterocycles (1-78) and general imines (79-84), from Scheme 6. Specifically: (1) for aromatic Nheterocycles, the $\Delta G_{H,A}(\mathbf{Y})$ scales ranged from -22.0 kcal mol⁻¹ to 15.8 kcal mol⁻¹ for Y_I, from -8.5 to 11.5 kcal mol⁻¹ for Y_{II}, and from -9.5 kcal mol⁻¹ to 7.4 kcal mol⁻¹ for Y_{III}, respectively. The $\Delta G_{H,A}(\mathbf{Y})$ scale of aromatic N-heterocycles $(\mathbf{Y}_{I}-\mathbf{Y}_{II})$ ranged from -22.0 kcal mol⁻¹ to 15.8 kcal mol⁻¹, which covered a very large range of 37.8 kcal mol⁻¹. The $\Delta G_{H_2A}(\mathbf{Y})$ scales indicated that the H_2 (0 kcal mol⁻¹),²²⁻²⁴ even the great hydrogenreductant HCO₂H (-5.9 kcal mol⁻¹),²⁵ could not hydrogenate all the aromatic imines (YI, YII, and YIII) to offer pre-aromatic Nheterocycles $(Y_IH_2, Y_{II}H_2, and Y_{III}H_2)$. It also could be deduced that some pre-aromatic N-heterocycles from Y_IH₂, Y_{II}H₂, and Y_{III}H₂ were thermodynamically better hydrogen reductants than H_2 in hydrogenation reactions. In practice, $27H_2$ (HEH₂),³⁸⁻⁴⁰ 67H₂,^{34,35} 73H₂,⁴¹ 74H₂,⁴² and 78H₂ (ref. 36 and 37) have already



Scheme 5 $\Delta G_{H_2R}(YH_2)$ scales of 4 groups of YH₂ (Y₁H₂-Y_{1V}H₂), along with the $\Delta G_{H_2R}(XH_2)$ values of HCO₂H, H₂, 67H₂, 78H₂, and HEH₂ releasing H₂ in acetonitrile (kcal mol⁻¹).



Scheme 6 $\Delta G_{H_2A}(\mathbf{Y})$ scales of 4 groups of \mathbf{Y} ($\mathbf{Y}_1 - \mathbf{Y}_{1V}$), along with the $\Delta G_{H_2R}(\mathbf{X}H_2)$ values of HCO₂H, H₂, **67H₂**, **78H₂**, and HEH₂ in acetonitrile (kcal mol⁻¹).

been extensively researched in hydrogenating various unsaturated compounds, including aldehydes ketones, alkenes, imines, and heterocycles;^{36–42} (2) in contrast, since \mathbf{Y}_{IV} were general imines without aromatic structures, all the $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y}_{IV})$ values (-18.3 to -10.3 kcal mol⁻¹) were all greater than the $\Delta G_{\text{H}_2\text{R}}(\text{H}_2)$ values (0 kcal mol⁻¹), meaning that H₂ could hydrogenate the related general imines (\mathbf{Y}_{IV}) in organic synthesis under suitable catalytic conditions, which has been proved by many published studies in the literature.^{43–47} Similarly, HCO₂H ($\Delta G_{\text{H}_2\text{R}}(\text{HCO}_2\text{H}) = -5.9$ kcal mol⁻¹),²⁸ 78H₂ ($\Delta G_{\text{H}_2\text{R}}(78H_2) = 0.9$ kcal mol⁻¹),^{36,37} 67H₂ ($\Delta G_{\text{H}_2\text{R}}(67H_2) = 3.6$ kcal mol⁻¹),^{34,35} and HEH₂ ($\Delta G_{\text{H}_2\text{R}}(\text{HEH}_2) = 7.3$ kcal mol⁻¹),^{38–40} could be applied to hydrogenate \mathbf{Y}_{IV} to prepare general amines $\mathbf{Y}_{IV}\mathbf{H}_2$ in organic synthesis from a thermodynamics viewpoint.

2.3. Thermodynamics for the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles

Many groups have focused on the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic Nheterocycles, mainly including discovering better pre-aromatic N-heterocycles carriers and developing novel metal-organic catalysts.¹⁻¹⁹ Since we have obtained so much valuable thermodynamic data of pre-aromatic N-heterocycles dehydrogenation and aromatic N-heterocycles hydrogenation, it seemed meaningful to clarify why the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles could happen, as well as why they possess the typical thermodynamic feature of the corresponding pre-aromatic Nheterocycles.

Theoretically, according to the reaction rule, for the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles, if the Gibbs free energy of a **YH**₂ dehydrogenation (**YH**₂ \rightarrow **Y** + **H**₂) is more negative than 0, $\Delta G_{\text{H},\text{R}}(\text{YH}_2) < 0$, then the **YH**₂ dehydrogenation reaction is judged as thermodynamically favorable; while if the Gibbs free energy of the corresponding **Y** hydrogenation reaction (**Y** + **H**₂ \rightarrow **YH**₂) is greater than 0, $\Delta G_{\text{H}_2\text{A}}(\text{Y}) > 0$, then the hydrogenation reaction of **Y** is thermodynamically unfavorable. In contrast, if the Gibbs free energy of a **YH**₂ dehydrogenation (**YH**₂ \rightarrow **Y**+**H**₂) is greater than 0, $\Delta G_{\text{H},\text{R}}(\text{YH}_2) > 0$, and the dehydrogenation reaction is considered as thermodynamically unfavorable; while if the Gibbs free energy of the corresponding **Y** hydrogenation $(\mathbf{Y} + \mathbf{H}_2 \rightarrow \mathbf{Y}\mathbf{H}_2)$ is more negative than 0, $\Delta G_{\mathbf{H}_2\mathbf{A}}(\mathbf{Y}) < 0$, and the related **Y** hydrogenation reaction is regarded as thermodynamically favorable. Therefore, it is very curious why the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles could happen, even using the same metal-organic catalyst.³⁻⁹

Examining the previous literature, it is believed that the H₂ plays a very important role in regulating the reversible dehydrogenation and hydrogenation thermodynamics.³⁻⁹ In YH₂ dehydrogenation, the H₂ is released from the reaction system, which greatly affects the dehydrogenation equilibrium.3-9 Therefore, the Gibbs free energy of YH₂ dehydrogenation could decrease to make the reaction happen if the $\Delta G_{H,R}(YH_2)$ value is not too much greater than 0. Furthermore, in Y hydrogenation, the high H₂ pressure would prompt more H₂ to dissolve in the reaction solvent, which would have a strong influence on the hydrogenation equilibrium.³⁻⁹ Consequently, the Gibbs free energy of Y hydrogenation could decrease to make the reaction occur if the $\Delta G_{H,A}(\mathbf{Y})$ is not too much greater than 0. Herein, this leads to the other key question: what are the restrictions on the $\Delta G_{H_{2}R}(\mathbf{YH}_{2})$ value for the reversible dehydrogenation and hydrogenation of a pre-aromatic and aromatic N-heterocycle?³

Investigating previous work, the Gibbs solvation energy of H₂ was estimated as 3.4 kcal mol⁻¹ in acetonitrile.⁴⁸ That is, the Gibbs free energy of H₂ release from acetonitrile solution was -3.4 kcal mol⁻¹, and the H₂ release from the dehydrogenation system could decrease the $\Delta G_{H_2R}(YH_2)$ by 3.4 kcal mol⁻¹. Therefore, it is reasonable to deduce that if the $\Delta G_{H,R}(YH_2)$ value of a pre-aromatic N-heterocycle is greater than -3.5 kcal mol^{-1} and more negative than 3.5 kcal mol^{-1} , -3.5 kcal mol^{-1} < $\Delta G_{\rm H_2R}(\rm YH_2) < 3.5 \ \rm kcal \ mol^{-1}$, the pre-aromatic N-heterocycle could be considered as a potentially reversible chemical organic hydrogen material, which could release H₂ and be regenerated by H₂ too. According to the above judgment criterion, the $\Delta G_{H_2R}(YH_2)$ values of general amines $(Y_{IV}H_2)$ (10.3– 18.3 kcal mol⁻¹) are much greater than 3.5 kcal mol⁻¹, and so Y_{IV}H₂ could not be designed as reversible dehydrogenation and hydrogenation materials, which is proved by published works.³⁻⁹ Moreover, the $\Delta G_{H,R}(YH_2)$ scale of 28 pre-aromatic Nheterocycles from YIH2, YIIH2, and YIIIH2 ranged from

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-3.5 kcal mol⁻¹ to 3.5 kcal mol⁻¹, while the corresponding 28 pre-aromatic N-heterocycles were identified as reversible dehydrogenation and hydrogenation materials. The 28 pre-aromatic N-heterocycles included $6H_2$ (-0.4 kcal mol⁻¹), $7H_2$ $(-0.8 \text{ kcal mol}^{-1})$, 8H₂ (2.9 kcal mol⁻¹), 15H₂ (2.7 kcal mol⁻¹), $23H_2$ (-1.9 kcal mol⁻¹), $24H_2$ (0.6 kcal mol⁻¹), $29H_2$ $(1.9 \text{ kcal mol}^{-1})$, $30H_2 (-0.2 \text{ kcal mol}^{-1})$, $31H_2 (3.5 \text{ kcal mol}^{-1})$, $32H_2$ (-0.1 kcal mol⁻¹), $38H_2$ (0.7 kcal mol⁻¹), $40H_2$ $(2.8 \text{ kcal mol}^{-1}), 41H_2 (0.8 \text{ kcal mol}^{-1}), 46H_2 (-2.5 \text{ kcal mol}^{-1}),$ $48H_2$ (-1.1 kcal mol⁻¹), $49H_2$ (0.2 kcal mol⁻¹), $51H_2$ $(-0.1 \text{ kcal } \text{mol}^{-1}), 52H_2 (-1.3 \text{ kcal } \text{mol}^{-1}), 53H_2$ $(3.0 \text{ kcal mol}^{-1})$, **61H**₂ $(0.3 \text{ kcal mol}^{-1})$, **64H**₂ $(-2.3 \text{ kcal mol}^{-1})$, $65H_2$ (-1.4 kcal mol⁻¹), $66H_2$ (1.6 kcal mol⁻¹), $68H_2$ $(2.4 \text{ kcal mol}^{-1})$, 70H₂ (0.6 kcal mol⁻¹), 72H₂ (1.0 kcal mol⁻¹), $76H_2$ (2.2 kcal mol⁻¹), and $78H_2$ (0.9 kcal mol⁻¹), which need further validation and support in experimental work. Examining the chemical structures of the investigated N-heterocycles in previous literature,³⁻¹⁹ they were not exactly the same prearomatic N-heterocycle structure, and most cases involved two H₂ molecules release and acceptance from N-heterocycles, which could not provide direct experimental data to validate the thermodynamic model.

2.4. Application of thermodynamic parameters to the selection of pre-aromatic N-heterocyclic hydrogen reductants for catalytic hydrogenation reactions

It is well known that H₂ is the greenest hydrogenation reagent with 100% atomic economy.43-47 Besides H2, the organic hydrogen molecule reductants, most of which are pre-aromatic N-heterocyclics (such as 27H₂, 67H₂, 73H₂, 74H₂, and 78H₂),³⁶⁻⁴² are irreplaceable hydrogen carriers in hydrogenation reactions. Zhou et al. developed dual hydrogenation strategies and utilized catalytic amounts of Hantzsch ester (HEH2 or 27H2)49 and dihydrophenanthridines (67H₂)50 as hydrogen reductants to asymmetrically reduce imines by using H₂ as a real reductant for regeneration of the oxidated Hantzsch esters (HE or 27) under a catalyst of metal complexes (Scheme 7). During their work, the reduction of imines, the oxidation of organic hydrogen reductants (pre-aromatic N-heterocycles), as well as the regeneration of oxidated organic hydrogen reductants (aromatic N-heterocycles) all involved H₂ transfer. Therefore, the Gibbs free energies for amines dehydrogenation and imines hydrogenation are important thermodynamic parameters to



Scheme 7 Dual hydrogenation strategies involving utilizing catalytic amounts of pre-aromatic N-heterocycles as hydrogen reductants to asymmetrically reduce imines by using H_2 as a real reductant for the regeneration of aromatic N-heterocycles.

evaluate the reducing abilities of amines and the hydrogenation difficulties of imines, which could help us discover thermodynamically excellent organic hydrogen reductants or chemical hydrogen-storage compounds, choose appropriate hydrogen reductants, especially for catalytic amounts of hydrogen reductants during the hydrogenation of unsaturated substrates.

If an unsaturated substrate (Sub) is hydrogenated to a reduced unsaturated substrate (SubH₂), Sub + H₂ \rightarrow SubH₂, whose thermodynamic driving force is defined as the Gibbs free energy of an unsaturated substrate (Sub) accepting H₂ to afford SubH₂ (equal to the opposite of Gibbs free energy of SubH₂ releasing H₂ to give Sub, SubH₂ \rightarrow Sub + H₂), $\Delta G_{\text{H}_2\text{A}}(\text{Sub}) =$ $-\Delta G_{\text{H}_2\text{R}}(\text{SubH}_2)$, then the larger the thermodynamic driving force for the pre-aromatic N-heterocycle reductant dehydrogenation, the more favorable the Sub hydrogenation by **YH**₂, *i.e.*, Sub + **YH**₂ \rightarrow SubH₂ + **Y**.²¹

Herein, the applications of thermodynamic parameters on choosing suitable catalytic amounts of pre-aromatic Nheterocycle reductants in hydrogenation reactions are displayed in Scheme 8 to aid a clear discussion.

It should be noted that not all unsaturated substrates hydrogenation reactions could be designed as organic hydrogen-reductant catalytic reactions. For unsaturated substrates (Sub), if the $\Delta G_{H,R}(SubH_2)$ values are more negative than -3.5 kcal mol $^{-1}$, due to the regeneration of aromatic N-heterocycles being impossible $(\Delta G_{\mathrm{H},\mathrm{R}}(\mathrm{YH}_2))$ -3.5 kcal mol⁻¹), the hydrogenation reactions could not be designed as pre-aromatic N-heterocyclic reductant catalyzed reactions. In contrast, if the $\Delta G_{H_3R}(SubH_2)$ values of unsaturated substrates are greater than 3.5 kcal mol^{-1} , whose thermodynamic driving force scale of SubH₂ releasing H₂ is denoted as $\Delta G_{H_2R}(SubH_2)_{min} \sim \Delta G_{H_2R}(SubH_2)_{max}$, then prearomatic N-heterocycles with $\Delta G_{H_2R}(YH_2)$ values more negative than $\Delta G_{H,R}(SubH_2)_{min}$ and greater than -3.5 kcal mol⁻¹ can be recognized as potential catalytic pre-aromatic Nheterocycle reductants, that is, -3.5 kcal mol⁻¹ < $\Delta G_{\rm H,R}$ (- YH_2) < $\Delta G_{H_2R}(SubH_2)_{min}$. Specially, if 3.5 kcal mol⁻¹ < ΔG_{H_2} $_{R}(\mathbf{YH}_{2}) < \Delta G_{H_{2}R}(\mathrm{SubH}_{2})_{\mathrm{min}}$, then the corresponding prearomatic N-heterocycles could be designed as thermodynamically good catalytic hydrogen reductants. If $0 < \Delta G_{H,R}$ YH_2 < 3.5 kcal mol⁻¹, the corresponding pre-aromatic Nheterocycles could be designed as thermodynamically better catalytic hydrogen reductants. If -3.5 kcal mol⁻¹ < $\Delta G_{H,R}$ (- YH_2) < 0 kcal mol⁻¹, the corresponding pre-aromatic Nheterocycles could be designed as the thermodynamically best catalytic hydrogen reductants. If $\Delta G_{H_2R}(YH_2)$ < -3.5 kcal mol⁻¹, the corresponding pre-aromatic Nheterocycles could be considered thermodynamically very excellent reductants, but they could not be regenerated by H₂ and designed as catalytic hydrogen reductants.

Thus, it can be seen that the two new thermodynamic parameters $\Delta G_{\text{H}_2\text{R}}(\mathbf{YH}_2)$ and $\Delta G_{\text{H}_2\text{A}}(\mathbf{Y})$ can not only help us evaluate and identify good chemical hydrogen-storage materials, but also choose appropriate hydrogen reductants, especially for catalytic amounts of N-heterocyclic hydrogen reductants during the hydrogenation of unsaturated substrates.



Scheme 8 Applications of the thermodynamic parameters on choosing suitable catalytic amounts of N-heterocyclic hydrogen reductants in hydrogenation reactions.

3 Conclusions

In this work, the Gibbs free energies of pre-aromatic Nheterocycles (YH₂) dehydrogenation and corresponding aromatic N-heterocycles (Y) hydrogenation, $\Delta G_{H,R}(YH_2)$ and $\Delta G_{\mathrm{H},\mathrm{A}}(\mathbf{Y})$, were derived by constructing thermodynamic cycles using Hess' law. The thermodynamic abilities on the acceptorless dehydrogenation and hydrogenation of 78 pre-aromatic N-heterocycles (YH_2) and related 78 aromatic N-heterocycles (Y) were well evaluated and discussed in acetonitrile. Several valuable conclusions could be drawn as follows:

(1) $\Delta G_{\text{H},\text{R}}(\text{YH}_2)$ is an important thermodynamic parameter to guide chemists to discover more potentially excellent hydrogen carriers. Not all pre-aromatic N-heterocycles are thermodynamically feasible chemical hydrogen-storage materials. The $\Delta G_{H_1R}(YH_2)$ scale of the considered 78 pre-aromatic Nranged from -15.8kcal mol^{-1} heterocycles to 22.0 kcal mol⁻¹, with 30 ΔG_{H_2R} (YH₂) values being more negative than 0, meaning that the related 30 YH₂ were identified as thermodynamically feasible H₂ donors, and belong to potential chemical hydrogen-storage materials. The $\Delta G_{H_2R}(Y_1H_2)$ scale $(-15.8 \text{ to } 22.0 \text{ kcal mol}^{-1})$ spanned the widest thermodynamic range by 37.8 kcal mol⁻¹ among $Y_IH_2-Y_{IV}H_2$, and the dehydrogenation abilities of $Y_{II}H_2 - Y_{IV}H_2$ decreased in the order of $Y_{II}H_2$ $\approx Y_{III}H_2 > Y_{IV}H_2.$

(2) As for the thermodynamic abilities of Y hydrogenation, the $\Delta G_{H,A}(\mathbf{Y})$ data indicated that for H_2 , even the great hydrogen-reductant HCO_2H (-5.9 kcal mol⁻¹), could not hydrogenate all the aromatic imines $(Y_{I}, Y_{II}, and Y_{III})$ to form pre-aromatic N-heterocycles. Y_{IV} are general imines without aromatic structures, and all their $\Delta {\it G}_{\rm H_2A}(Y_{IV})$ values (–18.3 to $-10.3 \text{ kcal mol}^{-1}$) were greater than $\Delta G_{\text{H},\text{R}}(\text{H}_2)$ (0 kcal mol}^{-1}), meaning that H₂ could hydrogenate the related general imines $(\mathbf{Y}_{\mathbf{IV}})$ in organic synthesis under suitable catalytic conditions.

(3) H₂ plays an important role in regulating reversible dehydrogenation and hydrogenation thermodynamics. The thermodynamic features in the reversible dehydrogenation and hydrogenation of pre-aromatic and aromatic N-heterocycles were clarified such that if $-3.5 \text{ kcal mol}^{-1} < \Delta G_{\text{H,R}}(\text{YH}_2) <$ 3.5 kcal mol⁻¹, the pre-aromatic N-heterocycle is considered as a thermodynamically potential reversible chemical organic hydrogen material, which could release H₂ and be regenerated by H₂ too.

(4) The application of thermodynamic parameters to the selection of pre-aromatic N-heterocyclic hydrogen reductants in catalytic hydrogenation was exhibited in this work. If the $\Delta G_{H_2R}(SubH_2)$ values of unsaturated substrates (SubH₂) are greater than 3.5 kcal mol⁻¹, pre-aromatic N-heterocycles with $-3.5 \text{ kcal mol}^{-1} < \Delta G_{H,R}(YH_2) < \Delta G_{H,R}(SubH_2)_{min}$ could be recognized as thermodynamically potential catalytic reductants. Specially, if 3.5 kcal mol⁻¹ < $\Delta G_{H,R}(YH_2)$ < $\Delta G_{H,R}(SubH_2)_{min}$, the related pre-aromatic N-heterocycles could be designed as thermodynamically good catalytic hydrogen reductants. If 0 < $\Delta G_{\mathrm{H}_{2}\mathrm{R}}(\mathrm{YH}_{2})$ < 3.5 kcal mol⁻¹, the related pre-aromatic Nheterocycles could be designed as thermodynamically better catalytic hydrogen reductants. If $-3.5 \text{ kcal mol}^{-1} < \Delta G_{\text{H},\text{R}}(\text{YH}_2)$ < 0 kcal mol⁻¹, the related pre-aromatic N-heterocycles could be designed as the thermodynamically best catalytic hydrogen reductants.

In summary, this work focused on two new thermodynamic parameters, namely $\Delta G_{H,R}(YH_2)$ and $\Delta G_{H,A}(Y)$, which is an important supplement to our previous work to offer precise insights into the chemical hydrogen storage of pre-aromatic Nheterocycles, and hydrogenation reactions.

Conflicts of interest

There are no conflicts of interest to declare.

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