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# A fast and simple procedure for the synthesis of a zinc and 1,4-benzene dicarboxylic acid metal–organic framework and its evaluation as a sorbent for dispersive micro solid phase extraction of pesticide residues†

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In this work, Zn-1,4-benzenedicarboxylate metal–organic framework was synthesized by a simple hydrothermal process and used in dispersive micro solid phase extraction of some pesticide from various fruit juice and water samples. Seven widely consumed pesticides in agriculture (chlorpyrifos, haloxyfop-*R*-methyl, oxadiazon, diniconazole, clodinafop-propargyl, fenpropathrin, and fenaxoprop-*p*-ethyl) were selected as target analytes. In this work, dispersive micro solid phase extraction was followed by a liquid phase microextraction method to achieve more enrichment of the analytes, and the enriched analytes were quantified using a gas chromatography–flame ionization detector. The sorbent was authenticated by Fourier transform infrared spectrophotometry, X-ray diffraction, energy dispersive X-ray analysis, and scanning electron microscope imaging. The factors affecting the extraction efficiency of the developed method were investigated, and the validation of the method under the optimized extraction conditions presented satisfactory results for precision and trueness, with limits of detection and quantification in the ranges of 0.50–0.90 and 1.5–2.7  $\mu\text{g L}^{-1}$ , respectively. Enrichment factors and extraction recoveries were in the ranges of 239–392% and 47–78%, respectively. One river water and some fruit and vegetable juice samples were analyzed by the recommended method, and the obtained recoveries were between 90% and 102%.

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## 1. Introduction

The use of pesticides is a common and essential practice in different stages of cultivation to control pests, preserve crop quality, and enhance their yield. Although they have many advantages, their overuse can lead to detrimental effects on untargeted organisms, including humans, through contaminated water and leaving their residues in agricultural products.<sup>1</sup> Hence, pesticide residue detection and quantification in water and agricultural products have become the subject of many analytical studies.<sup>2–4</sup> Common analytical methods for the

determination of pesticide residues utilize chromatography techniques such as gas chromatography (GC)<sup>5,6</sup> and liquid chromatography.<sup>7,8</sup> Generally, due to the complex matrices of food samples and the low concentration of pesticide residues in them, utilizing an efficient sample preparation process prior to instrumental analysis is needed. Solid phase extraction (SPE) is an efficient and commonly used method for the purification and enrichment of pesticide residues in food samples.<sup>9–11</sup> Dispersive micro solid phase extraction (D- $\mu$ SPE) is a variant of SPE and is used in food analysis frequently.<sup>12,13</sup> In D- $\mu$ SPE, sorbent particles are dispersed into a sample solution and a close contact is established between the solute molecules in the sample solution and the sorbent particles.<sup>14</sup> Therefore, analytes are adsorbed onto the sorbent particles and most of the other components of the sample remain in the liquid phase.<sup>15–18</sup> The prominent advantage of dispersive modes of SPE is the increasing contact area between the analytes and sorbent. Also, labors of performing SPE in the classic mode such as high and controlled vacuum or pressure for passing the sample solution through the cartridge and clogging the cartridge (when the samples having suspended particles are processed) are eliminated.<sup>19</sup> Up to now, various materials have been utilized as the

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sorbent for the extraction of pesticide residues from fruit and vegetable juices such as molecularly imprinted polymers,<sup>20</sup> carbon nanoparticles,<sup>21</sup> graphene,<sup>22</sup> functionalized silica,<sup>23</sup> and metal–organic frameworks (MOFs).<sup>24</sup> MOFs are a class of inorganic–organic hybrid crystals with porous structures. High adsorption ability, pore size and surface function tunability, synthetic versatility, and large specific surface area make them ideal for use in various fields such as gas storage,<sup>25</sup> catalysis,<sup>26</sup> drug delivery,<sup>27</sup> chemical sensors,<sup>28</sup> and luminescent materials.<sup>29</sup> They also are capable of being applied as the chromatographic stationary phase<sup>30</sup>, and adsorbents for SPE<sup>31</sup> and solid phase microextraction.<sup>32</sup>

In this study, MOF-5 consisting of zinc (as the metal ion) and 1,4-benzene dicarboxylic acid (terephthalic acid) (as the organic linker) was synthesized by a hydrothermal procedure and its efficiency as a sorbent for D- $\mu$ SPE of some pesticide residues from water and some fruit and vegetable juices was evaluated. The successful synthesis of the MOF-5 was verified by using techniques such as Fourier transform infrared (FTIR), X-ray diffraction (XRD), energy dispersive X-ray (EDX), and scanning electron microscopy (SEM) imaging. A dispersive liquid–liquid phase extraction (DLLME) step was performed after the D- $\mu$ SPE process and the extracted pesticides were quantified using a GC-flame ionization detector (FID).

## 2. Experimental

### 2.1. Solutions and reagents

Analytical-grade zinc nitrate hexahydrate, terephthalic acid, acetone, acetonitrile, isopropanol, methanol, potassium chloride, hydrochloric acid (37%, w/w), sodium sulfate, sodium hydroxide, ammonia solution (25%, w/v), and sodium chloride were obtained from Merck (Darmstadt, Germany). 1,1,2,2-Tetrachloroethane (1,1,2,2-TCE), 1,1,2-trichloroethane (1,1,2-TCE), carbon tetrachloride, and 1,2-dibromoethane (1,2-DBE) were from Janssen (Beerse, Belgium). All the studied pesticides including chlorpyrifos, fenpropathrin, oxadiazon, fenoxaprop-*p*-ethyl, haloxyfop-*R*-methyl, diniconazole, and clodinafop-propargyl (purity > 98%) were from Dr Ehrenstorfer company (Augsburg, Germany). A stock solution of the mentioned pesticides (100 mg L<sup>-1</sup> of each pesticide) was prepared in methanol and diluted with deionized water daily for use in the optimization and validation steps.

### 2.2. Samples

Fresh apples, watermelon, and onions were bought from local stores (Tabriz, Iran). They were squeezed using a juice extractor. The river water sample was collected from the Meydan Chay River (Tabriz, Iran). Prior to the extraction process, all the juices were diluted at a ratio of 1 : 1 with deionized water and used in the experiments. River water was subjected to the proposed method without any dilution or treatment.

### 2.3. Apparatus

Determination of the extracted pesticides was carried out on a Shimadzu 2014 gas chromatograph with a Zerbon™-1

capillary column (5% phenyl and 95% dimethylpolysiloxane) (Phenomenex, Torrance, CA, USA). Column dimensions and film thickness were as follows: 30 m × 0.25 mm and 0.25  $\mu$ m, respectively. The temperature programming of the column oven was: 60 °C for 1 min and ramped to 300 °C at 18 °C min<sup>-1</sup> (kept for 6 min). Each run time was nearly 21 min. The gas chromatograph was equipped with a split/splitless injection port and an FID. The injection port was operated at 300 °C in a splitless/split mode (split ratio and sampling time were 1 : 10 and 1 min, respectively). Helium gas as the carrier gas (99.999%) was flown at a constant linear velocity of 30 cm s<sup>-1</sup>. The make-up gas was helium at 30 mL min<sup>-1</sup>. The temperature of FID was set at 300 °C and flow rates of hydrogen and air were 30 and 300 mL min<sup>-1</sup>, respectively. pH adjustment was performed using a Metrohm-654 pH meter (Herisau, Switzerland). A D-7200 Hettich centrifuge (Kirchlengern, Germany) was used for centrifugation. Vortex agitation was performed using an L46 vortex (Labinco, Breda, The Netherlands). EDX analysis and SEM (Tescan, Czech) imaging were conducted to study the surface composition and morphology of the prepared sorbent. The FTIR spectrum of the prepared sorbent was recorded using a Bruker FTIR spectrophotometer (Billerica, USA). The XRD pattern of the synthesized sorbent was recorded on a Siemens AG (Karlsruhe, Germany) XRD apparatus operated at 35 kV.

### 2.4. MOF-5 synthesis

For the synthesis of MOF-5, 500 mg of terephthalic acid was dissolved in 10 mL of concentrated ammonia solution (25%, v/v) and dropwise added to 100 mL 1 mol L<sup>-1</sup> solution of Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O prepared in degassed deionized water. The reaction was continued at 80 °C for 2 h under stirring. The reaction product was collected by vacuum filtration, washed with deionized water, and dried at room temperature.

### 2.5. Extraction procedure

**2.5.1. D- $\mu$ SPE.** Five milligrams of MOF-5 were added to a 10 mL test tube poured with 5 mL of the sample solution (see Section 2.2) and/or model solution at a concentration of 100  $\mu$ g L<sup>-1</sup> of each pesticide. Then, sodium chloride was dissolved in it at a concentration of 5% (w/v). The resultant mixture was vortexed for 5 min to disperse the sorbent particles into the solution. Thereafter, it was centrifuged for 7 min at 7000 rpm and the supernatant was discarded. Afterward, 2 mL of isopropanol was added to the sorbent particles and the analytes were desorbed by vortexing for 5 min. Finally, centrifugation was performed (for 7 min at 7000 rpm) and the supernatant was utilized as a disperser in the subsequent DLLME step.

**2.5.2. DLLME.** To further enrich the extracted analytes from the previous step, the obtained eluate was mixed with 35  $\mu$ L 1,2-DBE as an extractant. The obtained solution was speedily injected into 5 mL 5% (w/v) NaCl solution and placed into a 10 mL conical bottom glass test tube. A cloudy solution was formed. It was centrifuged at 7000 rpm for 7 min and a 10  $\pm$  1  $\mu$ L organic phase was sedimented. One microliter of the sedimented phase was injected into GC-FID to quantify the analytes.



The overall extraction time was almost 30 min. The proposed D- $\mu$ SPE-DLLME method is schematically shown in Fig. 1.

## 2.6. Calculation of the enrichment factor (EF) and extraction recovery (ER)

Two important factors of EF and ER indicating the efficiency of the proposed method were calculated from eqn (1) and (2), respectively:

$$EF = \frac{C_{\text{sed}}}{C_0} \quad (1)$$

$$ER = \left( \frac{n_{\text{sed}}}{n_0} \right) \times 100 = \frac{C_{\text{sed}} \times V_{\text{sed}}}{C_0 \times V_{\text{aq}}} \times 100 = \frac{V_{\text{sed}}}{V_{\text{aq}}} \times EF \times 100 \quad (2)$$

where  $C_0$  and  $C_{\text{sed}}$  are the initial concentrations of analytes in the sample solution and the sedimented phase, respectively.  $n_0$  and  $n_{\text{sed}}$  are the total and extracted amounts of the analyte, respectively.  $V_{\text{aq}}$  and  $V_{\text{sed}}$  are the volumes of the sample solution and sedimented extractant, respectively.

## 3. Results and discussion

### 3.1. Characterization of the synthesized MOF-5

Characterization of the crystalline structure of the prepared MOF-5 was performed by XRD analysis. The obtained XRD pattern (Fig. 2a) showed that the diffractions at  $2\theta$  of  $6.1^\circ$ ,  $8.1^\circ$ ,  $12.2^\circ$ ,  $15.9^\circ$ ,  $16.1^\circ$ ,  $18.4^\circ$ ,  $24.5^\circ$ ,  $25.9^\circ$ ,  $29.8^\circ$ , and  $33.8^\circ$  are related to the crystal planes of (200), (220), (400), (420), (422), (440) (711), (731), (751), and (911), respectively, and they are matched with those from the previous reports in the

literature.<sup>33,34</sup> FTIR spectroscopy was used to confirm the functional groups in MOF-5 (Fig. 2b). Two peaks at  $1392$  and  $1605 \text{ cm}^{-1}$  are assigned to stretching vibrations of C–O and C=O bonds of terephthalic acid, respectively. The morphology of MOF-5 particles was investigated by SEM. As shown in Fig. 2c, MOF-5 particles exhibit spherical and irregular shapes. The EDX analysis (Fig. 2d and e) indicates Zn, C, and O peaks with weight percentages of 33.26%, 31.65%, and 35.09%, respectively. The BET analysis was performed to investigate the surface properties of the prepared sorbent. The following values were obtained: average pores diameter;  $24.365 \text{ nm}$ ; total pores volume:  $0.111 \text{ cm}^3 \text{ g}^{-1}$ , and surface area;  $18.227 \text{ m}^2 \text{ g}^{-1}$ .

### 3.2. Optimization of sorbent composition and weight

The amount of terephthalic acid was changed in the synthesis process in the range of 100–750 mg to reach an optimal composition for the prepared sorbent. The amount of  $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  was kept constant at 29.75 g. The reaction product in each case was collected, washed, and dried. Twenty milligrams of each product were subjected to the proposed procedure. Considering the results in Fig. S1,† the difference in ERs of the analytes by the use of each product as the sorbent is not significant ( $p$  values  $> 0.05$ ). Therefore, MOF-5 was synthesized in each case. Given that the yield of the synthesis reaction was high using 500 mg of terephthalic acid, it was selected for the preparation of the sorbent.

The weight of the sorbent was investigated in the range of 2–25 mg (2.0, 3.0, 4.0, 5.0, 10, 20, and 25 mg). The results from Fig. 3 show that after increasing the sorbent weight from 2 to 5 mg, the ERs of the analytes increased. Further increasing the sorbent amount has no significant effect on the extraction

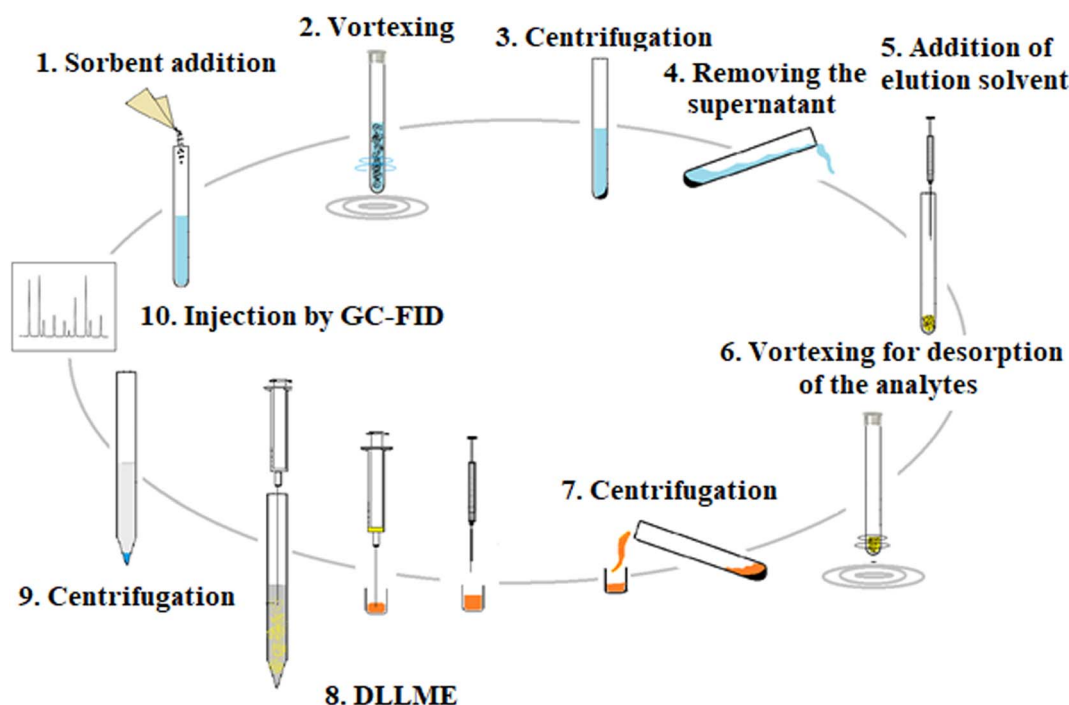


Fig. 1 Extraction procedure steps.



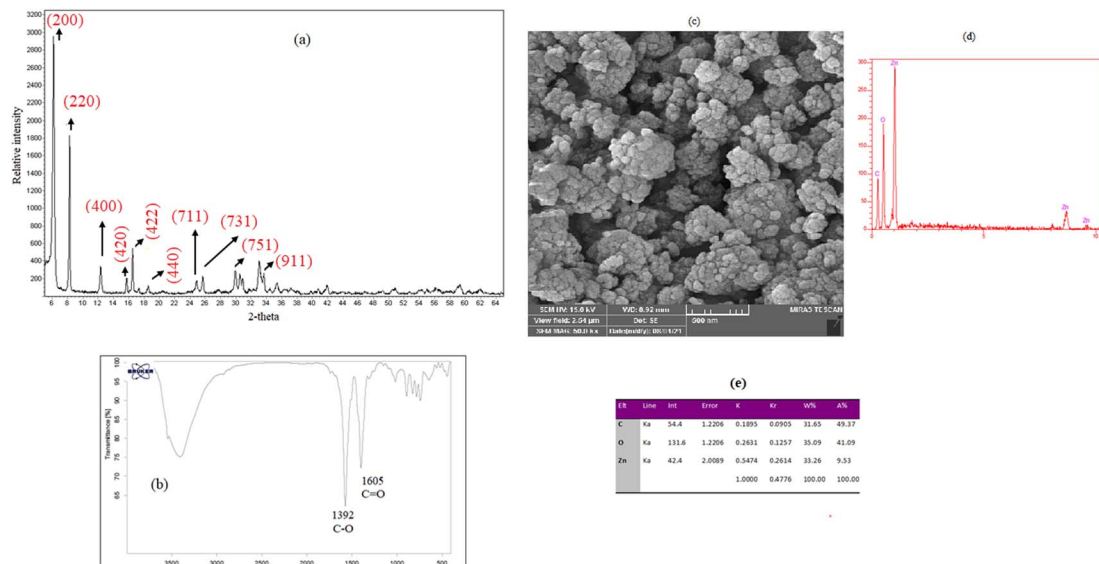


Fig. 2 (a) XRD pattern, (b) FTIR spectrum, (c) SEM images, (d) EDX spectrum (e), and EDX data of the synthesized sorbent.

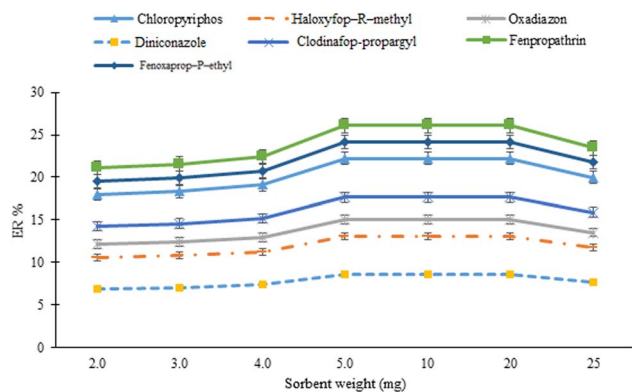


Fig. 3 Optimization of sorbent amount. Extraction conditions: sorbent amount, 20 mg; sample volume, 5 mL deionized water spiked with  $100 \mu\text{g L}^{-1}$  of each analyte; terephthalic acid amount used for MOF synthesis, 500 mg; vortexing time in adsorption and desorption steps, 9 min; elution solvent (volume), acetonitrile (1 mL); DLLME aqueous phase, 5 mL deionized water; extraction solvent (volume), 1,2-DBE (45  $\mu\text{L}$ ); and centrifugation time and rate; 7 min and 7000 rpm, respectively. Error bars indicate minimum and maximum of three repeated determinations.

efficiency of the method. Increasing the method efficiency by increasing the sorbent amount is attributed to increasing the active sites for the adsorption of the analytes. However, by further increasing the sorbent amount, aggregation of the sorbent particles and incomplete elution of the analytes limit any additional improvement in the efficiency of the method. Therefore, 5 mg of the sorbent was used in the rest of the study.

### 3.3. Optimization of D- $\mu$ SPE conditions

**3.3.1. Ionic strength of sample solution.** Ionic strength is an important factor in microextraction techniques. Increasing ionic strength can represent two contrary effects: (I) a decrease

in the solubility of the analyte in aqueous solution and favor their transferring into the extractive medium and (II) an increase in viscosity of the aqueous solution and a negative influence on the kinetic of the extraction. To investigate the effect of this factor, 5 mL of 2.5%, w/v, solutions of NaCl, KCl, and  $\text{Na}_2\text{SO}_4$  spiked with  $100 \mu\text{g L}^{-1}$  of each analyte were used as a sample solution, separately. The results in Fig. S2(a)† indicate that NaCl is a bit better than other salts and it can be related to better solvation of NaCl ions with water and release of the organic compounds.

To optimize the concentration of NaCl, it was varied in the range of 1–10%, w/v. The results shown in Fig. S2(b)† indicate that the extraction efficiency is maximum at 5%, w/v, NaCl and decreases thereafter. Increasing the method efficiency by adding NaCl shows a salting-out effect. However, at higher concentrations, increasing the sample solution viscosity reduces the analytes migration rate and decreases the method efficiency. Attending to the results, 5%, w/v, NaCl was used for further studies.

**3.3.2. Sample solution pH.** The effect of pH on the extraction efficiency of the method was tested by adjusting the pH of the aqueous solution over the range of 3–9. The obtained results in Fig. S3† show that the maximum ER of each analyte was obtained in the pH range of 4–8, which matches the pH of the used samples in this study. Hydrolysis or dissociation of the analytes is the major reason for decreasing the method's efficiency.<sup>35</sup> Therefore, the following experiments were performed without pH adjusting.

**3.3.3. Optimization of adsorption and desorption time.** In the dispersive mode of SPE, vortex mixing is usually used to form better interfacial contact between the analytes and sorbent particles in the adsorption step and effective elution of the analytes from the sorbent particle's surface in the desorption step. Therefore, it is an important parameter and affects the performance of the method. Additionally, the D- $\mu$ SPE process



needs optimum time durations to complete the adsorption and desorption of the analytes. Hence, vortex mixing times of adsorption and desorption steps were investigated in the range of 1–9 min. According to the results (Fig. S4†), the ERs increase with increasing vortex mixing time from 1–5 min and then remain almost constant in both adsorption and desorption steps. Increasing the analytical data by increasing adsorption and desorption time is related to the increasing contact area of the sorbent with the elution solvent. It must be noted that 5 min is sufficient time for adsorption or desorption of the analytes. Hence, 5 min time was opted for vortex mixing time for both of adsorption and elution steps in the subsequent experiments.

**3.3.4. Selection of elution solvent type and volume.** In the proposed D- $\mu$ SPE-DLLME method, the selection of a viable elution solvent for quantitative desorption of the analytes from the surface of the sorbent particles is of great importance. Additionally, the dispersion efficiency of the extraction solvent used in the DLLME step depends on the type of the elution solvent. For this purpose, 1 mL of methanol, isopropanol, acetonitrile, and acetone were tested as the possible eluents and their desorption/dispersion capabilities are compared in Fig. 4a. Regarding the results, the highest ERs for the studied analytes are achieved using isopropanol as the elution solvent. It can be related to the higher solubility of the analytes in isopropanol or its better role as a disperser solvent in DLLME. Therefore, it was selected for use in the following experiments.

The effect of the elution solvent volume on the extraction efficiency of the method was evaluated using different volumes of isopropanol ranging from 0.5 to 2.5 mL and the outcomes are given in Fig. 4b. As can be seen, increasing the volume of isopropanol up to 2.0 mL results in increasing the extraction efficiency of the method. A decrease in the extraction efficiency in the volume of 2.5 mL can be assigned to the decreased polarity of the aqueous phase used in DLLME. This leads to increasing the analyte's solubility in the aqueous phase. Therefore, 2.0 mL was used in subsequent steps.

### 3.4. Optimization of DLLME conditions

**3.4.1. Salt addition to the DLLME aqueous phase.** Five milliliters of  $\text{Na}_2\text{SO}_4$ ,  $\text{NaCl}$ , and  $\text{KCl}$  aqueous solutions at a concentration of 2.5% (w/v) were used as the DLLME aqueous phases to investigate the influence of salt addition in the DLLME step. The results shown in Fig. S5a† indicate that the addition of  $\text{NaCl}$  improves extraction efficiency for all of the analytes.

Different concentrations of  $\text{NaCl}$  in the concentration range of 1–10%, w/v, were tested to investigate the influence of salt concentration. By increasing the salt concentration from 1 to 10%, w/v, the volume of the extractant was decreased from 45 to 25  $\mu\text{L}$  and in all cases  $10 \pm 1 \mu\text{L}$  of the sedimented extractant was obtained. The results shown in Fig. S5b† indicate that the salt addition up to 5.0%, w/v, leads to an increase in ERs of the analytes and that increasing the  $\text{NaCl}$  concentration provides no further improvement in the extraction efficiency of the method. Increasing the method efficiency by the addition of salt can be related to decreasing the solubility of the analyte in the aqueous phase used in the DLLME step. Therefore, 5 mL  $\text{NaCl}$  solution at a concentration of 5.0%, w/v, was selected as the aqueous phase in the DLLME step.

**3.4.2. Selection of extraction solvent type and volume.** Different density than water, low aqueous solubility, high extraction ability for the analytes, and good chromatographic behavior are the required conditions of a proper extraction solvent for the proposed D- $\mu$ SPE-DLLME method. Considering the mentioned requirements, 1,1,2,2-TCE (50  $\mu\text{L}$ ), 1,2-DBE (35  $\mu\text{L}$ ), 1,1,1-TCE (40  $\mu\text{L}$ ), and carbon tetrachloride (30  $\mu\text{L}$ ) were examined and  $10 \pm 1 \mu\text{L}$  of the sedimented extractant was obtained in each case. The results (Fig. 5) illustrate that 1,2-DBE presents the highest efficiency among the tested solvents due to more solubility of the analytes in this solvent or its better dispersion by the disperser solvent and providing high contact area. Therefore, it was used in the further steps.

To study the effect of extraction solvent volume on the method efficiency, the volume of 1,2-DBE was altered in the range of 35–55  $\mu\text{L}$ . According to the results (data not shown here), the analytical signals of the analytes decrease with increasing volume of the extractant due to the dilution effect. On the other hand, using volumes lower than 35  $\mu\text{L}$  led to obtaining an insufficient sedimented organic phase, which reduced the repeatability of results. Considering a proper balance between the analytes enrichment and repeatability, 35  $\mu\text{L}$  1,2-DBE was selected for the subsequent experiments.

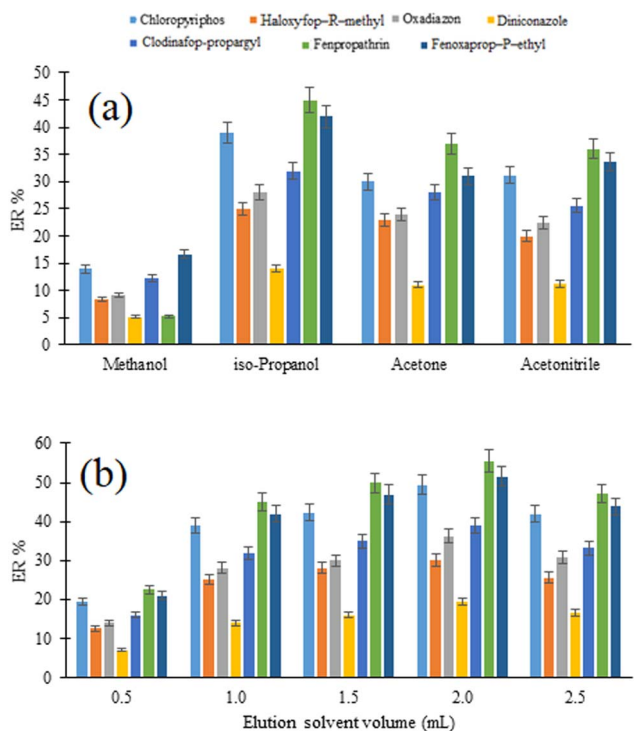


Fig. 4 Optimization of elution solvent type (a) and (b) volume. (a) Extraction conditions: are the same as those used in Fig. 3, except adsorption and desorption times were 5 min. (b) Extraction conditions: are the same as those used in (a), except iso-propanol was chosen as the elution solvent.



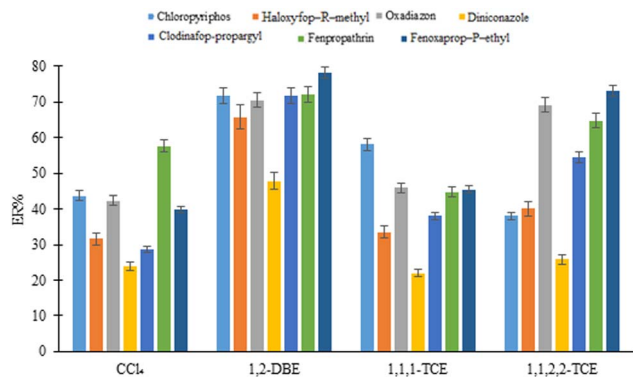


Fig. 5 Selection of extraction solvent in DLLME. Extraction conditions: are the same as those used in Fig. 4(b), except 5.0%, w/v, NaCl was dissolved in the DLLME aqueous phase.

### 3.5. Method validation

The calibration curves plotting was performed by preparing various solutions at concentrations of 3, 10, 25, 50, 100, 250, 500, and 1000  $\mu\text{g L}^{-1}$  and subjecting them to the method. It is noted that the maximum concentration of the calibration curve for haloxypop-*p*-ethyl and diniconazole was 500  $\mu\text{g L}^{-1}$ . Satisfactory coefficients of determination in the range of 0.994–0.997 were obtained for all analytes. The precision of the method was expressed using relative standard deviations (RSDs). Intra- and inter-day precisions were in the ranges of 5–6 and 6–8% ( $n = 5$ ), respectively. The limit of detection (LOD) and quantification (LOQ) values, determined at the signal-to-noise ratios of 3 and 10, were in the ranges of 0.50–0.90 and 1.5–2.7  $\mu\text{g L}^{-1}$ , respectively. The ER and EF values were between 47 and 78% and 239 and 392, respectively. Table 1 summarizes the values.

### 3.6. Real samples analysis

It is a known fact that the efficiency of microextraction techniques can be influenced by the composition of the sample matrix. Therefore, river water, and fresh apple, watermelon, and onion juices were prepared as mentioned in Section 2.2, spiked with 50 and 100  $\mu\text{g L}^{-1}$  for each analyte, and analyzed under the optimized experimental conditions. The found relative recoveries (RR) (expressed as the ratio of the analyte peak areas in

real samples and deionized water spiked with the same concentration of analyte multiplied by 100) are listed in Table 2. The obtained RR values of the studied analytes are between 90 and 102%. The results indicate that the sample matrices have no significant effect on the extraction efficiency of the method. It should be mentioned that the obtained RR values in undiluted juice samples were relatively low. Therefore, they were diluted with deionized water at a ratio of 1:1 (v/v) before extraction. It is noteworthy that, none of the studied analytes was detected in the analyzed juice and water samples. Therefore, the studied pesticides were not present in the studied samples or their concentrations were below the method LODs.

### 3.7. Comparison of the developed method and some other reported methods

The efficiency of the proposed method was compared with that of the other techniques applied for pesticide residue

Table 2 Results of relative recovery tests performed on the studied samples

Analyte	Mean relative recovery (%) $\pm$ standard deviation ( $n = 3$ )			
	Apple	Watermelon	Orange	River water
<b>All samples were spiked with each analyte at a concentration of 50 <math>\mu\text{g L}^{-1}</math></b>				
Chlorpyrifos	90 $\pm$ 5	92 $\pm$ 5	95 $\pm$ 5	99 $\pm$ 6
Haloxypop- <i>p</i> -ethyl	90 $\pm$ 5	93 $\pm$ 5	95 $\pm$ 6	98 $\pm$ 5
Oxadiazon	91 $\pm$ 5	92 $\pm$ 6	94 $\pm$ 6	97 $\pm$ 5
Diniconazole	90 $\pm$ 5	93 $\pm$ 5	95 $\pm$ 5	98 $\pm$ 5
Clodinafop-propargyl	92 $\pm$ 6	93 $\pm$ 6	96 $\pm$ 5	97 $\pm$ 6
Fenpropathrin	91 $\pm$ 5	92 $\pm$ 5	96 $\pm$ 5	98 $\pm$ 5
Fenoxaprop- <i>p</i> -ethyl	90 $\pm$ 5	94 $\pm$ 5	95 $\pm$ 5	96 $\pm$ 5
<b>All samples were spiked with each analyte at a concentration of 100 <math>\mu\text{g L}^{-1}</math></b>				
Chlorpyrifos	95 $\pm$ 5	100 $\pm$ 5	96 $\pm$ 5	100 $\pm$ 6
Haloxypop- <i>p</i> -ethyl	96 $\pm$ 5	102 $\pm$ 6	96 $\pm$ 5	99 $\pm$ 6
Oxadiazon	96 $\pm$ 6	98 $\pm$ 6	98 $\pm$ 6	95 $\pm$ 5
Diniconazole	97 $\pm$ 5	96 $\pm$ 5	98 $\pm$ 5	101 $\pm$ 6
Clodinafop-propargyl	95 $\pm$ 6	96 $\pm$ 6	98 $\pm$ 6	99 $\pm$ 6
Fenpropathrin	98 $\pm$ 5	98 $\pm$ 6	100 $\pm$ 5	100 $\pm$ 5
Fenoxaprop- <i>p</i> -ethyl	95 $\pm$ 5	95 $\pm$ 5	96 $\pm$ 6	95 $\pm$ 6

Table 1 Analytical features of the developed D- $\mu$ SPE-DLLME-GC-FID method for pesticides analysis

Analytes	LOD <sup>a</sup>	LOQ <sup>b</sup>	LR <sup>c</sup>	$r^{2d}$	<sup>e</sup> RSD%		EF $\pm$ SD <sup>f</sup>	ER $\pm$ SD <sup>g</sup>
					Intra-day	Inter-day		
Chlorpyrifos	0.90	2.7	2.7–1000	0.995	6	7	359 $\pm$ 24	71 $\pm$ 5
Haloxypop- <i>p</i> -ethyl	0.50	1.5	1.5–500	0.994	5	7	329 $\pm$ 25	65 $\pm$ 5
Oxadiazon	0.50	1.5	1.5–1000	0.994	6	8	353 $\pm$ 16	70 $\pm$ 3
Diniconazole	0.90	2.7	2.7–500	0.994	5	6	239 $\pm$ 14	47 $\pm$ 3
Clodinafop-propargyl	0.90	2.7	2.7–1000	0.995	5	7	359 $\pm$ 35	71 $\pm$ 5
Fenpropathrin	0.62	1.8	1.8–1000	0.997	5	6	360 $\pm$ 15	72 $\pm$ 3
Fenoxaprop- <i>p</i> -ethyl	0.51	1.5	1.5–1000	0.996	6	8	392 $\pm$ 16	78 $\pm$ 3

<sup>a</sup> Limit of detection (S/N = 3) ( $\mu\text{g L}^{-1}$ ). <sup>b</sup> Limit of quantification (S/N = 10) ( $\mu\text{g L}^{-1}$ ). <sup>c</sup> Linear range ( $\mu\text{g L}^{-1}$ ). <sup>d</sup> Coefficient of determination. <sup>e</sup> Relative standard deviation for intra- ( $n = 5$ ) and inter-day ( $n = 5$ ) precisions at a concentration of 25  $\mu\text{g L}^{-1}$  of each analyte. <sup>f</sup> Enrichment factor  $\pm$  standard deviation ( $n = 3$ ). <sup>g</sup> Extraction recovery  $\pm$  standard deviation ( $n = 3$ ).



Table 3 Comparison of the developed method with other reported methods for pesticides determination

Analyte	Method	Sample	LR <sup>a</sup>	LOD <sup>b</sup>	EF <sup>c</sup>	<sup>d</sup> RSD%	Ref.
Chlorpyrifos	MDSPE-HPLC-UV <sup>e</sup>	Water and rice	10–100	1.0	401	4.8	36
Diniconazole	MDSPE-UHPLC-MS/MS <sup>f</sup>	Water	1.5–500	0.5	—	3.7	37
Oxadiazon	MDSPE-HPLC-UV <sup>e</sup>	Water	0.5–10	0.07	180	7.2	38
Fenprothrin	MWAE-DSPE-GC-MS <sup>g</sup>	Radix astragali	0.2	0.2–10	—	—	39
Chlorpyrifos	D-μSPE-DLLME-GC-FID <sup>h</sup>	Fruit and vegetable juices and river water	2.7–1000	0.90	359	6	This study
Haloxfob- <i>p</i> -ethyl			1.5–500	0.50	329	5	
Oxadiazon			1.5–1000	0.50	353	6	
Diniconazole			2.7–500	0.90	239	5	
Clodinafob-propargyl			2.7–1000	0.90	359	5	
Fenprothrin			1.8–1000	0.62	360	5	
Fenoxaprop- <i>p</i> -ethyl			1.5–1000	0.51	392	6	

<sup>a</sup> Linear range ( $\mu\text{g L}^{-1}$ ). <sup>b</sup> Limit of detection ( $\mu\text{g L}^{-1}$ ). <sup>c</sup> Enrichment factor. <sup>d</sup> Relative standard deviation. <sup>e</sup> Magnetic dispersive solid phase extraction-high performance liquid chromatography-ultraviolet detection. <sup>f</sup> Magnetic dispersive solid phase extraction-ultra high performance liquid chromatography-tandem mass spectrometry. <sup>g</sup> Microwave-assisted extraction-dispersive solid phase extraction-gas chromatography-mass spectrometry. <sup>h</sup> Dispersive micro solid phase extraction-dispersive liquid-liquid microextraction-gas chromatography-flame ionization detection.

determination by comparing their figures of merit (Table 3). As can be seen, the developed method possesses lower LODs compared to other methods except the used for the analysis of oxadiazon in which an inherently sensitive instrument (ultra-performance liquid chromatography-tandem mass spectrometry) was used as the determination system.<sup>36</sup> Wide linear ranges regarding the applied analysis system were obtained in comparison to other methods. Also, the EF and RSD values of the recommended method are higher or comparable with those of the mentioned methods. Generally, the developed D-μSPE-DLLME-GC-FID method can be considered a fast, low-cost, efficient, and reliable sample pretreatment method in the sample clean-up, extraction, and enrichment of the pesticide residues from various aqueous samples.

## 4. Concluding remarks

In the present study, MOF-5 was synthesized by a simple and fast hydrothermal procedure and successfully applied for the sample cleanup and extraction of some pesticides from water and fruit juice samples. Appropriate LODs, EFs, and precision were achieved and the obtained experimental results proved that the D-μSPE method using the MOF-5 combined with DLLME has a good potential for the extraction of the selected pesticides. The use of only 5 mg of the sorbent, consumption of relatively low volumes of organic solvents, and simplicity, qualifies the method as an economically and environmentally benign technique for the pesticide analysis.

## Abbreviations

D-μSPE	Dispersive liquid-liquid microextraction
ER	Extraction recovery
EF	Enrichment factor
GC	Gas chromatography
LOD	Limit of detection
LOQ	Limit of quantification
LR	Linear range

SPE	Solid phase extraction
FTIR	Fourier transform infrared
XRD	X-ray diffraction
EDX	Energy dispersive X-ray
SEM	Scanning electron microscopy
DLLME	Dispersive liquid-liquid phase extraction
MOF	Metal-organic framework
FID	Flame ionization detector
1,1,2,2-TCE	1,1,2,2-Tetrachloroethane
1,1,2-TCE	1,1,2-Trichloroethane
1,2-DBE	1,2-Dibromoethane
RSD	Relative standard deviation
RR	Relative recovery

## Data availability

The data are confidential.

## Conflicts of interest

The authors have declared no conflict of interest.

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