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CORRECTION

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Correction: A computational study of the thortveitite structure of zinc pyrovanadate, $Zn_2V_2O_7$, under pressure

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Correction for 'A computational study of the thortveitite structure of zinc pyrovanadate, $Zn_2V_2O_7$, under pressure' by S. Reza *et al.*, *RSC Adv.*, 2023, **13**, 17212–17221, https://doi.org/10.1039/D3RA02426A.

The authors regret that there was an error in the first line in the introduction of the original article.

The text originally read, "The search for chemical systems suitable for lithium-ion batteries with good charge-exchange ability and high energy density[1–3] has attracted the attention of researchers toward electrode materials to meet the energy demand in portable electronic devices such as cell phones, laptops, computers, digital cameras, and motor vehicles due to their storage capacity which commercial batteries made of graphite lack." This sentence should read, "The search for chemical systems suitable for lithium-ion batteries with good charge-exchange ability and high energy density[1–3] has attracted the attention of researchers toward electrode materials to meet the energy demand in portable electronic devices such as cell phones, laptops, computers, digital cameras, and motor vehicles due to their storage capacity instead of commercial batteries made of graphite."

The authors regret that eqn (3) was shown incorrectly in the original article. The correct version of eqn (3) is as shown below. The corresponding section of the text in the manuscript should be adjusted according to this change, as detailed below.

The value of Debye temperature Θ_D was estimated using the following expression:[47]

$$\Theta_D = \frac{h}{k_B}\left[\frac{3n}{4\pi V_0}\right]^{\frac{1}{3}} v_a \tag{3}$$

where h is Planck's constant, k_B is Boltzmann's constant, V_0 is the volume of the unit cell, n is the number of atoms and v_a is the average wave velocity.

The authors regret that eqn (4) was shown incorrectly in the original article. The correct version of eqn (4) is as shown below.

$$\frac{1}{m^*} = \frac{1}{\hbar^2}\frac{\partial^2 E(k)}{\partial k^2} \tag{4}$$

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

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