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Hydrogen peroxide assisted synthesis of fluorescent carbon nanoparticles from teak leaves for dye-sensitized solar cells†

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Fluorescent carbon nanoparticles (FCNs) have emerged as promising sensitizers for dye-sensitized solar cells (DSSCs) owing to their unique optical properties and low-cost fabrication. In this study, we synthesized three different types of FCNs, such as hydrogen peroxide (H₂O₂)-assisted FCNs (FCN_H), hydrothermal-assisted FCNs (FCN_A), and pyrolysis-assisted (FCN_P). Among them, FCN_H-based DSSCs exhibited superior fluorescence properties, indicating effective surface oxidation by H₂O₂. Then we fabricated DSSCs using FCN_P, FCN_A, and FCN_H as sensitizers and evaluated their photovoltaic activity. The results revealed that FCN_H-based DSSCs exhibited superior performance compared to FCN_P-based and FCN_A-based DSSCs with an open circuit voltage (*V*_{OC}) of 0.43 V, a short circuit current (*I*_{SC}) of 0.52 mA, and a fill factor (FF) of 0.45. The superior performance of FCN_H-based DSSCs can be attributed to the improved charge transfer and fluorescence properties of the FCNs after surface oxidation with H₂O₂. These findings emphasize the significance of surface oxidation in optimizing the performance of FCNs in DSSCs and suggest the potentiality of FCN_H as an efficient sensitizer in renewable energy as well as energy harvesting applications for wearable platforms.

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Sustainability spotlight

The unique optical properties and low fabrication cost of fluorescent carbon nanoparticles (FCNs) have made them a promising sensitizer for dye-sensitized solar cells (DSSCs). Typically, concentrated oxidizing acidic solutions, including HNO₃, H₂SO₄, H₂SO₄/HNO₃ mixture and H₃PO₄, have been used to induce the oxidation of carbon sources, resulting in the formation of carbon nanodots with surface functional moieties, such as carboxylic acids, hydroxyl groups, or carbonyl groups. Teak leaves (*Tectona grandis* L. f.) were utilized in this study as a natural resource to create affordable, environmentally beneficial H₂O₂-assisted functionalized carbon nanomaterials (FCN_H). The fluorescent carbon nanoparticles (FCNs) were applied in dye sensitized solar cells. The study's findings highlight the importance of surface oxidation in maximizing the performance of functionalized carbon nanomaterials (FCNs) in dye-sensitized solar cells (DSSCs) and point to the possibility of using H₂O₂-assisted FCNs as an effective sensitizer in energy harvesting and renewable energy applications for wearable platforms. The Sustainable Development Goals (SDGs) of the UN are met by our work: responsible consumption and production (SDG 12), sustainable cities and communities (SDG 11), and industry, innovation, and infrastructure (SDG 9).

1 Introduction

In recent years, fluorescent carbon nanomaterials (FCNs) have gained considerable attention from researchers due to their unique properties, such as tunability in size, shape, excellent

electron conductivity, fluorescence properties, surface chemistry, and optical properties.^{1–9} These properties make them attractive for a variety of applications, including optoelectronic devices,^{9,10} wearable sensors,¹¹ drug delivery,¹² supercapacitors,¹³ photocatalysis,¹⁴ solar cells,^{2,5,15,16} bioimaging and biosensing.^{17–19}

Various synthesis strategies have been developed for preparing FCNs, including top-down and bottom-up approaches.^{20,21} Top-down approaches involve the fragmentation of a larger carbon structure using physical or chemical methods, followed by size selection and surface modification. Bottom-up approaches involve the formation of carbon nanoparticles from smaller molecular precursors through chemical reactions. Other methods, such as electrochemical and photochemical synthesis, have also been reported.²² Typically, concentrated oxidizing acidic solutions, including HNO₃,²³

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H_2SO_4 ,²⁴ $\text{H}_2\text{SO}_4/\text{HNO}_3$ mixture,^{25,26} and H_3PO_4 ,²⁷ have been used to induce the oxidation of carbon sources, resulting in the formation of carbon nanodots with surface functional moieties, such as carboxylic acids, hydroxyl groups, or carbonyl groups.

In the last few years, H_2O_2 -assisted quantum dots from different sources, such as graphene,²⁸ sulfur,²⁹ and nitrogen-doped carbon,³⁰ have been synthesized by hydrothermal treatment. In this study, we used teak leaves (*Tectona grandis* L. f.) as a natural resource to synthesize eco-friendly and cost-effective H_2O_2 -assisted functionalized carbon nanomaterials (FCN_H). Teak leaves are a promising source of carbon due to their abundance of various organic compounds.³¹

Hydrothermal treatment is a cost-effective and convenient one-step approach for producing water-soluble FCNs with a narrow size distribution on a large scale.³² The use of H_2O_2 as an oxidizer during the synthesis process can incorporate oxygen-containing functional groups onto the surface of the carbon skeleton and improve the aqueous solubility and biocompatibility of the FCNs.³³ However, in this study, we also fabricated hydrothermal-assisted FCNs using only deionized (DI) water (referred to as FCN_A), and additional FCNs were synthesized by pyrolyzing teak leaves (referred to as FCN_P) by using H_2O_2 . It was observed that FCN_A exhibited a blue emission color and showed good stability in an aqueous solution under 365 nm UV irradiation, despite the absence of H_2O_2 .

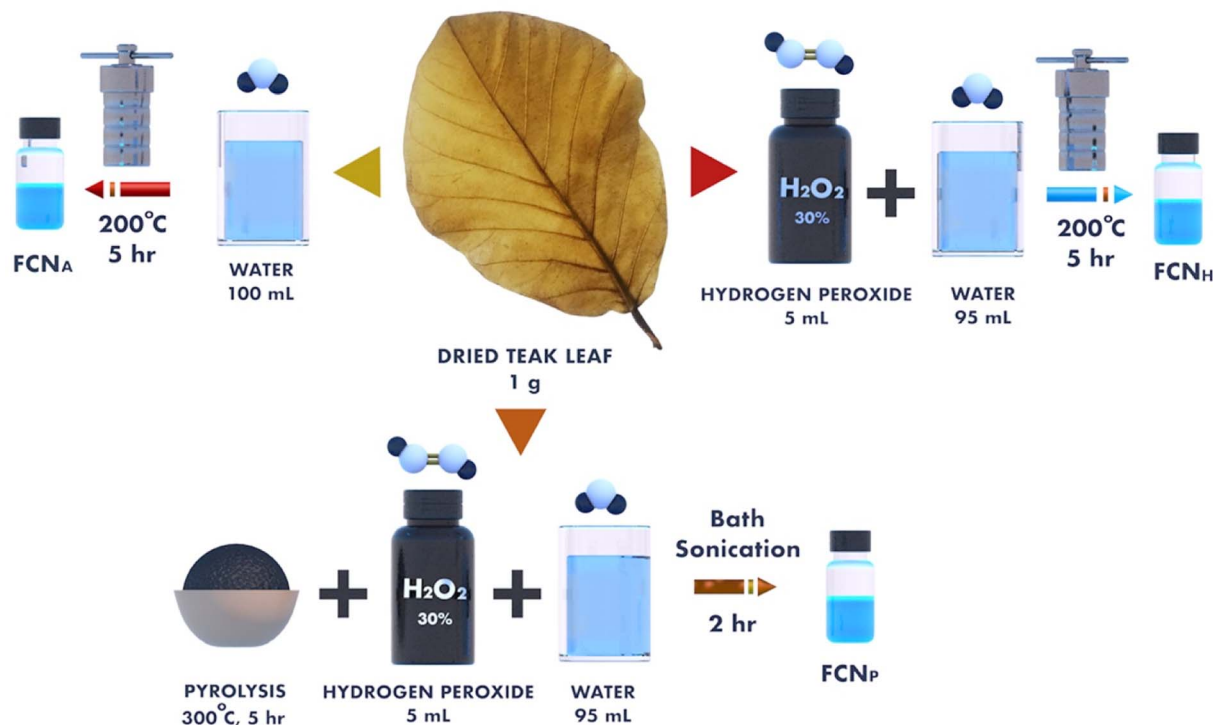
To date, in the field of dye-sensitized solar cells (DSSCs), a limited number of applications of FCNs have been reported.³⁴ In DSSCs, FCNs act as sensitizers, enhancing light absorption in the visible region and thereby improving overall solar cell efficiency.³⁴ Nevertheless, organic dyes, particularly ruthenium (Ru) dyes,^{35,36}

and D- π -A structured porphyrin dyes,^{34,37-39} have also been used as photosensitizers in DSSCs. Generally, organic-based sensitizers have demonstrated superior performance compared to FCNs. However, due to certain drawbacks associated with organic dyes, such as low stability, time-consuming synthesis process and toxicity, the scientific community has emphasized the need for the development of simple, cost-effective and eco-friendly DSSCs to overcome these limitations.³⁴ From this perspective, the use of non-toxic and environmentally friendly FCNs can potentially contribute towards fabricating DSSCs. In this study, blue emitting FCNs are synthesized from teak leaves to construct DSSCs. To the best of our knowledge, this study represents the first attempt to employ H_2O_2 -assisted FCNs in DSSCs, with results indicating excellent activity compared to traditional FCNs.

2 Experimental section

2.1 Chemicals and apparatus

Analytical grade hydrogen peroxide (30%) was purchased from Sigma-Aldrich. Dried teak tree (*Tectona grandis* L. f.) leaves were collected near the Chittagong University of Engineering & Technology area in the Chattogram region of Bangladesh. The titanium dioxide (TiO_2) powder (Degussa P-25) for the photoanode preparation was acquired from the Institute of Chemical Education, University of Wisconsin–Madison, Department of Chemistry, Madison, WI, USA. Fluorine tin oxide (FTO) conducting glass slides were purchased from Hartford Glass Company, Hartford City, Indiana, USA. Acetone, ethanol, and acetic acid were obtained from Sigma-Aldrich (St. Louis, MS, USA) and used without additional purification. DI water was used throughout all segments in this experiment.



Scheme 1 Different synthesis routes of FCNs from dried teak leaves.



2.2 Experimental

2.2.1 Preparation of FCN_H. FCNs were synthesized using dried teak leaves as a carbon source through hydrothermal and pyrolysis methods (Scheme 1). For the preparation of FCN_H, the teak leaves were washed with deionized water, dried, and sieved (0.3 mm) to remove the veins and midrib. Next, 1 g fine teak leaf powder was mixed with 5 mL H₂O₂ and 95 mL DI water, sonicated and transferred into a PPL-lined autoclave for heating at 200 °C for 5 hours. After cooling down at room temperature, the resulting light-yellow solution was purified and dialyzed against DI water to remove the excess amount of H₂O₂. The synthesized FCNs are referred to as FCN_H. Moreover, FCN_A and FCN_P were also synthesized for the purpose of comparison in this study. All samples were collected in powder form through a freeze-drying process.

2.2.2 Synthesis of FCN_A. To prepare FCN_A, 1 g of fine dried-teak leaf powder was sonicated in 100 mL of DI water for 1 hour, without the addition of H₂O₂. The resulting solution was then transferred to a PPL-lined hydrothermal autoclave and heated at 200 °C for 5 hours. After allowing it to cool at room temperature, the light-yellow solution was filtered through a 0.22 μm membrane filter to remove larger particles and the resulting powder was obtained by freeze-drying.

2.2.3 Synthesis of FCN_P. To synthesize FCN_P through the pyrolysis process, 1 g of teak leaf powder was heated at 300 °C for 5 hours in a muffle furnace. After cooling to room temperature, the pyrolyzed sample was dispersed in 5 mL H₂O₂ and 95 mL of DI water with the assistance of a bath-sonicator (Powersonic) for 2 hours. FCN_P was filtered through a 0.22 μm membrane filter and then collected in dried form through freeze-drying.

2.3 Fabrication of dye sensitized solar cells

The titanium dioxide (TiO₂) films were prepared utilizing a previously documented spin-coating technique.^{40–42} To prepare the working electrode, the fluorine-doped transparent tin oxide (FTO) substrate (Hartford Glass Co., Inc. TEC 7, 1" × 1" × 2.2 mm thickness, Solar) was cleaned with detergent, deionized (DI) water and ethanol using an ultrasonic bath for 30 min. To prepare the titanium dioxide (TiO₂) paste, TiO₂ powder (Degussa P-25) was initially mixed with glacial acetic acid. Then, the prepared TiO₂ paste was applied to the conductive side of FTO using a spin coater and annealed at 450 °C for 30 min.^{43–45} The substrate was then immersed in FCN solutions (1 mg of FCNs in 1 mL pure water) overnight for sensitization. Loosely bound particles were removed by rinsing with water and ethanol and the photoanode was dried at room temperature. The counter electrode (cathode) was prepared by painting colloidal graphite on the cleaned FTO-coated glass slide and allowed it to dry at room temperature. Finally, to fabricate the DSSCs, the FCN-sensitized slide and the carbon electrodes were put on top of each other and the redox (I⁻/I₃⁻) electrolyte solution was placed between them.

2.4 Characterization and measurements

The fluorescent nanoparticles and the sensitized film's morphology were assessed using field emission scanning

electron microscopy (JSM-7100FA JEOL USA, Inc.). Absorption spectroscopy was conducted using a UV-3600 Plus from Shimadzu, MD, USA, while emission spectroscopy measurements were performed with an RF-5301PC from Shimadzu, MD, USA. The zeta potentials of the samples were determined using an SZ-100 series Dynamic Light Scattering Particle Size Distribution Analyzer, HORIBA Instruments Inc., Irvine, CA, USA. Transmission Electron Microscopy (TEM) images were obtained using a JEM-1400 PLUS (JEOL USA, Peabody, Massachusetts, USA). The TiO₂ paste was deposited onto FTO glass using a WS-650 Series Spin Processor from Laurell Technologies Corporation, PA, USA. Carbon paint obtained from TED PELLA, INC, USA was used for making cathode slides.

2.5 Photovoltaic measurements of DSSCs

The cell performance was evaluated using a 150 W fully reflective solar simulator with standard illumination of air-mass 1.5 global (AM 1.5 G) and an irradiance of 100 mW cm⁻² (Sciencetech Inc., London, Ontario, Canada). A GAMRY Instruments Reference 600 Potentiostat/Galvanostat/ZRA (Warminster, PA) was employed for the measurements. Cells were externally biased and the photocurrent generated was measured to determine the current–voltage characteristics of the cells under these conditions. The electrocatalytic performance of the samples was further evaluated by electrochemical impedance spectroscopy (EIS).

3 Results and discussion

Surface oxidation is a recognized method for enhancing fluorescence in carbon-based materials by introducing oxygen-containing functional groups *via* the generation of surface defects.⁴⁶ H₂O₂ is known to be effective in promoting the generation of these defects, resulting in an abundance of carbonyl and carboxylic groups on the carbon structure, thereby improving the solubility in aqueous solutions.⁴⁷ The prepared H₂O₂-supported aqueous solution of FCN_H exhibited blue emission under 365 nm UV light and formed stable-colored suspensions in DI water, which is likely due to the larger number of functional groups, such as hydroxyl and carboxylic groups.

The UV-vis spectra of all FCNs have been investigated and are presented in Fig. 1a. The spectra showed a weak absorption in the UV region at around 220 nm attributed to the π–π* electronic transition of the sp² conjugated carbon (C=C) bond⁴⁸ and a weak absorption peak at around 285 nm due to the functional groups (C=O) present on the surface of all FCNs.⁴⁹ Notably, the absorbance peak position of FCN_H shifted towards the blue region compared to the other two FCNs, signifying a larger number of hydroxyl and carboxylic groups incorporated into the carbon structure through oxidation with H₂O₂.

The surface functional groups of the three different types of the synthesized FCNs were investigated using Fourier transform infrared (FTIR) spectroscopy, as presented in Fig. 1b. The FCN_H sample exhibited a higher overall peak intensity compared to FCN_A and FCN_P. All FCNs exhibited identical C=C stretching



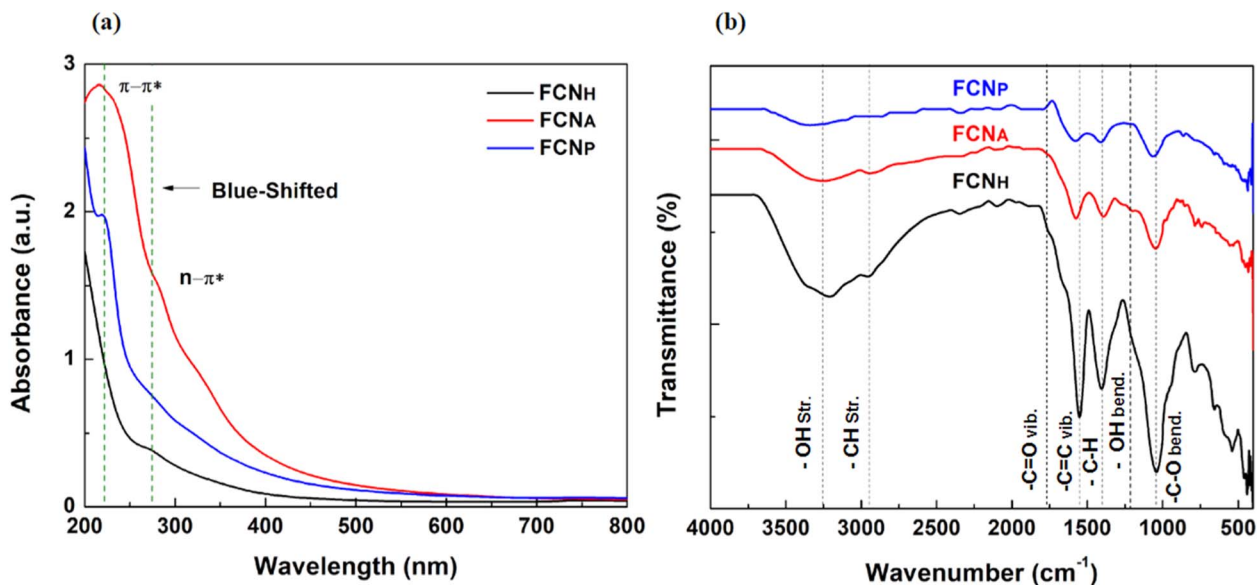


Fig. 1 (a) UV-vis spectra and (b) FT-IR of the as-synthesized FCNs.

and C-H bending peaks at around 1600 and 1027 cm^{-1} , respectively, attributed to the alkyl hydrocarbon from the teak leaf. The broad signature band peak of -OH stretching at 3200 cm^{-1} to 3400 cm^{-1} was detected in all FCNs, confirming the adsorption of water molecules onto sample surfaces, along with an -OH bending peak at around 1390 cm^{-1} . Additionally, a distinguished C=O stretching peak was detected at around 1740 cm^{-1} with high intensity in FCN_H, confirming the successful incorporation of a considerable number of oxygen functional groups into the carbon moieties after treatment with H₂O₂. A superimposed C-H stretching peak of each sample was also observed in the region 2400–2950 cm^{-1} .⁵⁰

The photoluminescence (PL) spectra were recorded with an excitation wavelength of 340 nm. All synthesized carbon nanoparticles showed fluorescence emission peaks in the blue region, ranging from 430 nm to 435 nm. These results indicated the maximum wavelength at which the emitted fluorescence was observed for all types of FCNs (Fig. 2).

The zeta potential, which quantifies the surface charge of colloidal particles in a solution, was measured and is presented

in Fig. 3. The zeta potential values of the three types of FCNs were assessed using electrophoretic light scattering in an aqueous solution. All samples exhibited negative zeta potential values, confirming negatively charged surfaces. This negative charge is attributed to the presence of oxygenated functional groups on the particle surfaces, which dissociate in the solution and generate negatively charged ions. Among the three samples, FCN_H displayed the most negative zeta potential value of -27.7 mV, demonstrating a higher surface negative charge density. The formation of a larger number of functional groups on the particle surface can be attributed to H₂O₂ treatment. The H₂O₂ treatment likely facilitated the introduction of more oxygen functionalities, leading to an increased negative charge on the surface. On the other hand, FCN_A exhibited a zeta potential value of -11.2 mV, while FCN_P showed a more negative value of -23.6 mV, suggesting a notable increase in the negative surface charge compared to FCN_A. Notably, FCN_H exposed an even stronger negative charge than FCN_P, demonstrating a certain level of passivation or resistance toward chemical reactions.^{51,52} This passivation effect could limit or

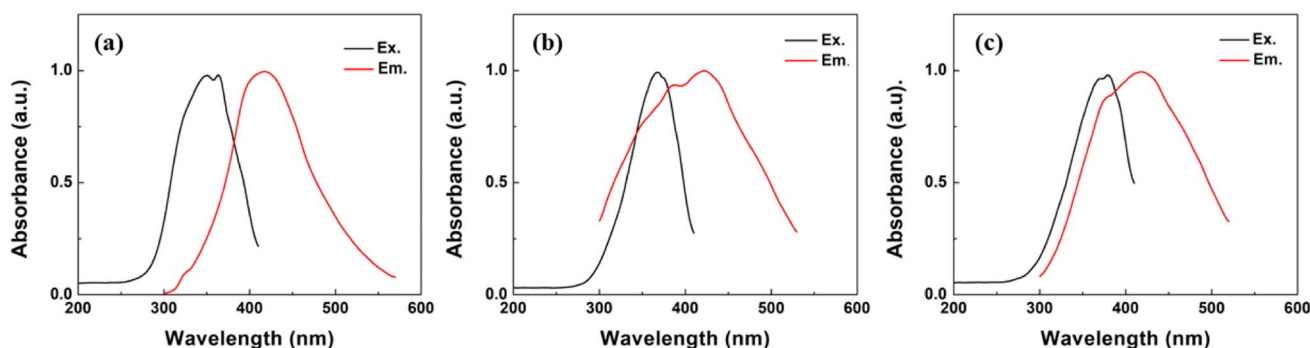


Fig. 2 PL spectra of (a) FCN_H, (b) FCN_A and (c) FCN_P.



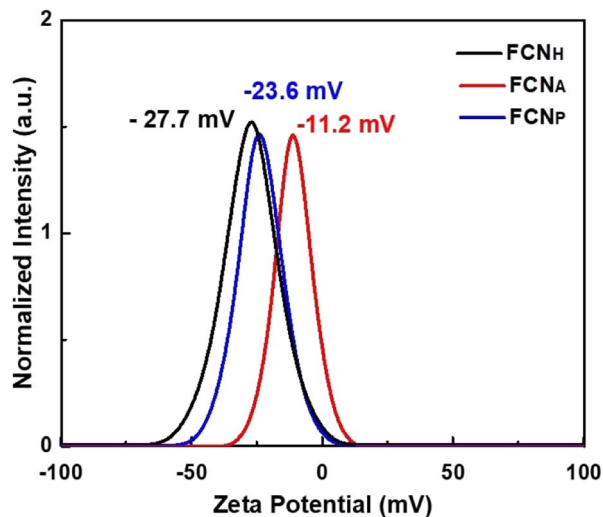


Fig. 3 Zeta Potential measurement of all FCNs.

slow down the functionalization reactions, even with H_2O_2 treatment. Even though H_2O_2 introduced oxygenated functional groups, the surface properties of pyrolytic carbon nanoparticles might impede the further incorporation of functional groups,

resulting in lower zeta potential values compared to FCN_H .⁵³ The observed percentages in EDX (Fig. S1 and Table S1†) reflect the relative abundance of carbon (C) and oxygen (O) atoms within each sample. FCN_H exhibited a composition of 79.53% C and 20.47% O atoms, indicating a relatively higher oxygen content. The treatment with H_2O_2 mainly introduced oxygen-containing functional groups onto the surface of the carbon dots through oxidation reactions. Similarly, FCN_A displayed a composition of 81.59% C and 16.52% O, which confirmed the slightly lower oxygen content compared to FCN_H . The reduction in oxygen content might be due to the breaking of weaker oxygen-carbon bonds under the high-temperature and pressure conditions of hydrothermal treatment.⁵⁴ The composition of 42.71% carbon and 14.35% oxygen in FCN_P suggests a significant reduction in both carbon and oxygen contents compared to the other treatments. The reduction in both carbon and oxygen contents implies that the pyrolysis process led to the removal of various oxygen-functional groups, as well as a partial decomposition of carbon components, resulting in a composition with relatively lower carbon and oxygen concentrations.

In this study, we also investigated the surface morphology of FCN_H using transmission electron microscopy (TEM) and field-emission scanning electron microscopy (FE-SEM). The TEM images displayed individual particles within a size range of 30–

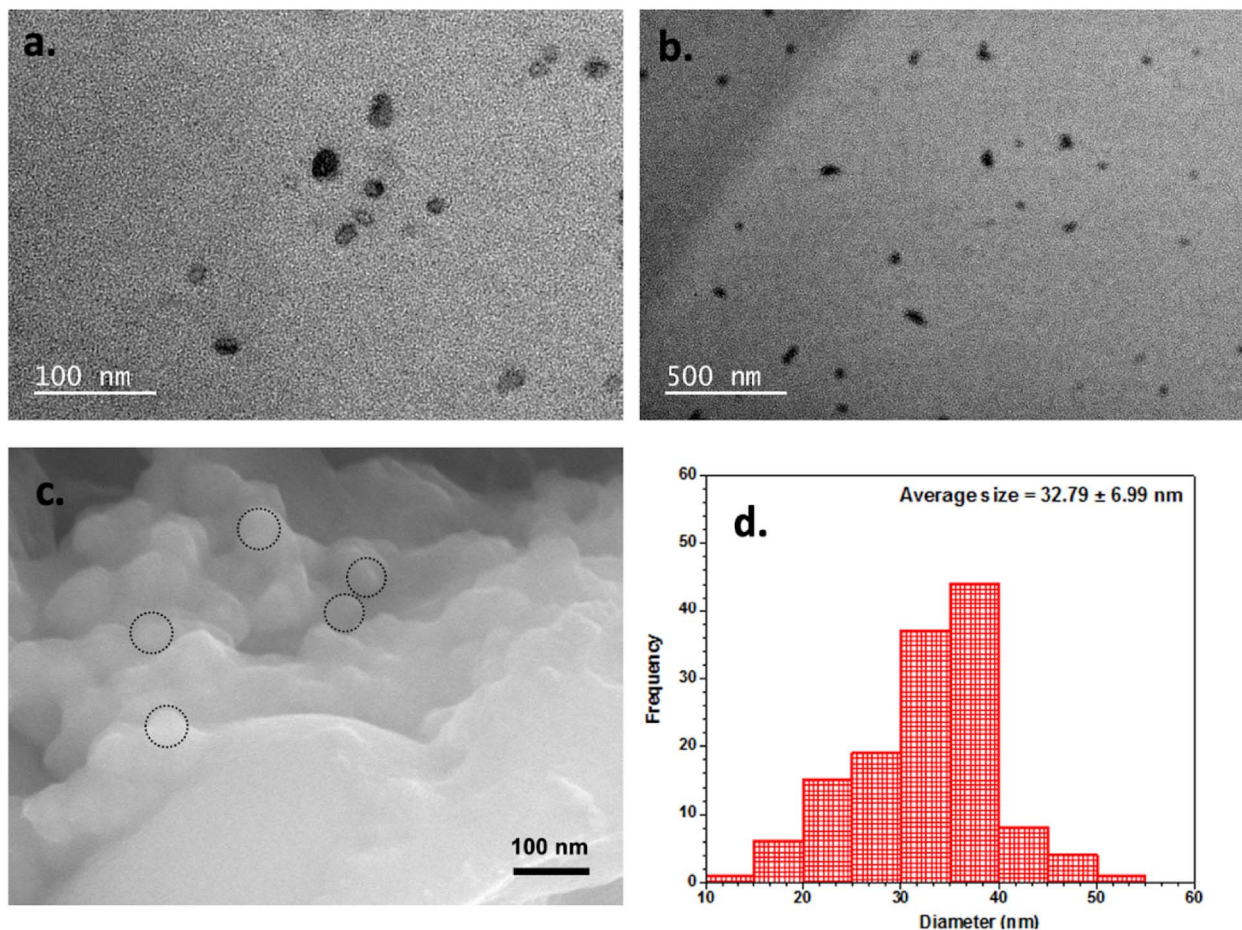


Fig. 4 (a) High-resolution TEM image, (b) TEM image, (c) FE-SEM image and (d) particle size distribution histogram of FCN_H .



40 nm, as depicted in Fig. 4a and b. The FE-SEM results, depicted in Fig. 4c, along with the size distribution measurement (Fig. 4d), corroborated and strengthened the findings from TEM, confirming the presence of nano-sized particles in the sample. Additionally, the FE-SEM images in Fig. S2† revealed that FCN_P possessed smaller particle sizes compared to FCN_A. Notably, the introduction of H₂O₂ during the synthesis process resulted in a more significant reduction in particle size for FCN_P compared to FCN_A. The DLS data (Fig. S3†) obtained in this study were also consistent with the FE-SEM images, supporting the observation that the particle size of FCN_P is smaller than that of FCN_A.

The photoluminescence (PL) lifetime values of the carbon dots were measured for FCN_H, FCN_A, and FCN_P (Fig. 5). For FCN_H, the PL lifetime values were 2.20 ns and 8.69 ns, pointing to two distinct decay processes. Similarly, FCN_A exhibited two PL lifetime values of 1.28 ns and 7.07 ns, indicating the presence of two decay processes in its emission behavior. On the

Table 1 Fluorescence lifetime measurement of FCNs

| Sample name | Lifetime | | | |
|------------------|---------------|------------------------|---------------|------------------------|
| | τ_1 (ns) | Std dev. | τ_2 (ns) | Std dev. |
| FCN _H | 2.20 | 1.21×10^{-11} | 8.69 | 2.74×10^{-11} |
| FCN _A | 1.28 | 8.39×10^{-11} | 7.07 | 3.03×10^{-11} |
| FCN _P | 2.09 | 1.32×10^{-11} | 8.57 | 3.00×10^{-11} |

other hand, FCN_P displayed PL lifetime values of 2.09 ns and 8.57 ns when tested with the N719 dye. Therefore, it can be inferred from the PL lifetime values that FCN_H has a longer PL lifetime compared to FCN_A and FCN_P, proposing a lower rate of non-radiative recombination (Table 1). This characteristic potentially contributes to the enhanced solar-to-electric power conversion efficiency in the FCN_H-based DSSCs. Conversely, FCN_A and FCN_P, with their shorter PL lifetimes, are likely to exhibit higher rates of non-radiative recombination, thus resulting in a lower solar-to-electric power conversion efficiency in the DSSCs that utilized these dyes.

3.1 Use of FCNs as sensitizers in dye sensitized solar cells

When they are exposed to sunlight, dye molecules (S) absorbed on the TiO₂ film absorb photons and are excited from the highest occupied molecular orbital (HOMO) to the lowest unoccupied molecular orbital (LUMO) state as shown in Scheme 2. Upon photoexcitation, the dye species (S*) injects an electron into the conduction band of the TiO₂ electrode, leading to its oxidation (S⁺). The oxidized dye species accepts an electron from the electrolyte (I⁻), resulting in the restoration of the ground state of the dye (S). The injected electron is then transported through the nano-porous TiO₂ film to the FTO layer and further conducted through an external circuit to a load, where it is converted into electrical energy. The electron from the external load subsequently diffuses to the cathode and is transferred to the electrolyte (I₃⁻), thus regenerating the electrolyte system.

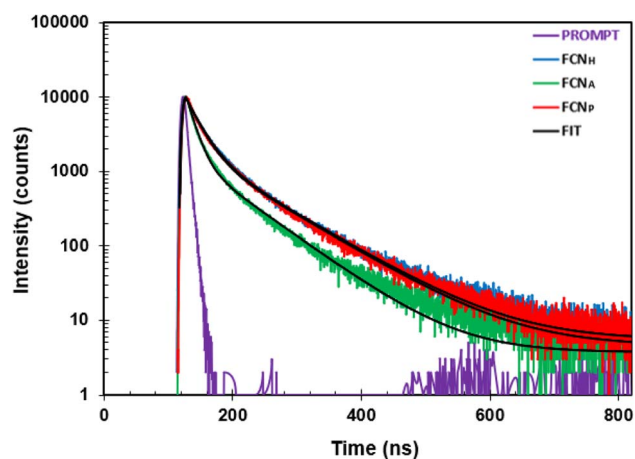
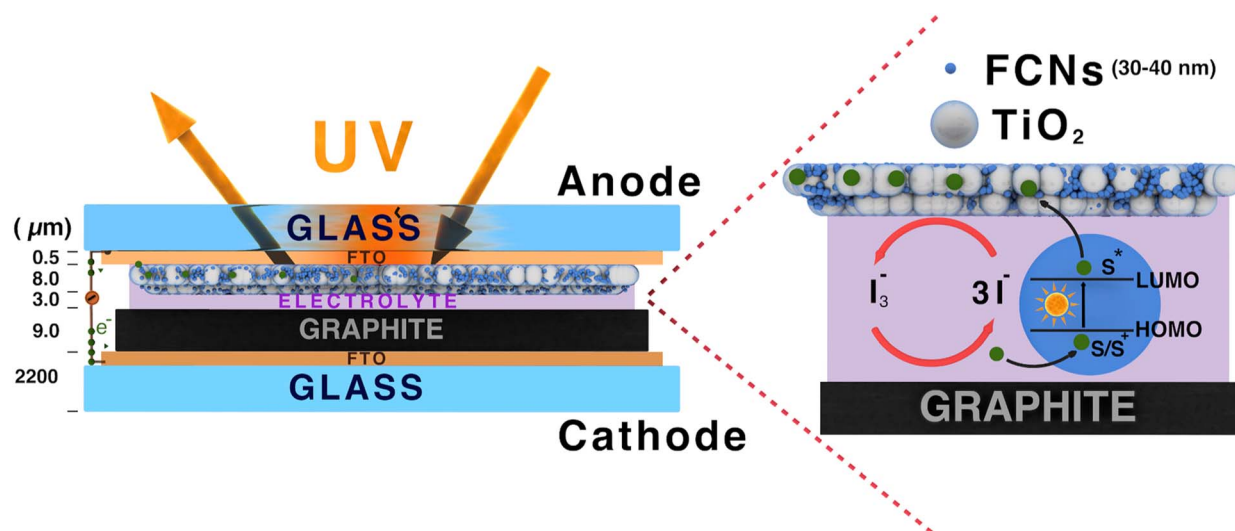


Fig. 5 Photoluminescence decay curves at a 420 nm emission wavelength of the FCNs.



Scheme 2 A schematic diagram of the proposed DSSC (not to scale).



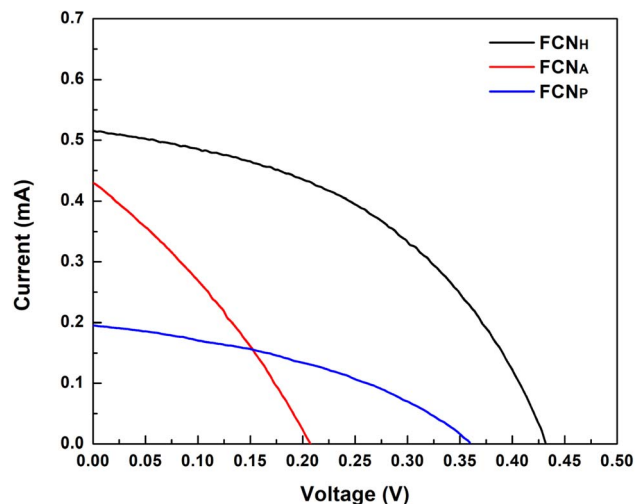


Fig. 6 Current–voltage characteristics of DSSCs fabricated from FCNs.

The current–voltage (I – V) measurement of the samples was carried out with a standard illumination of air-mass 1.5 global (AM 1.5 G) having an irradiance of 100 mW cm^{-2} . The I – V values of the samples are used to calculate the solar-to-electric energy efficiency of the solar cells. The I – V values of the measurements are displayed in Fig. 6. It was observed that FCN_H exhibits promising characteristics as it is evident from its open circuit voltage (V_{OC}) of 0.43 V and short-circuit current (I_{SC}) of 0.52 mA, thus demonstrating the successful functioning of the proposed DSSC (Table 2). Furthermore, the fill factor (FF) of 0.45 obtained from the fabricated DSSC ensures a good electrical contact and low series resistance. However, the power conversion efficiency (PCE) of the fabricated dye sensitized solar cell for FCN_H was 0.10%, offering further improvement in terms of light absorption and charge separation efficiency (Table 2).

3.2 Electrochemical impedance spectroscopy (EIS)

Electrochemical impedance spectroscopy (EIS) studies were carried out to understand the transport and recombination of the photoexcited electrons in the fabricated DSSC. The EIS studies expressed in Bode phase and Nyquist plots are displayed in Fig. 7 and 8, respectively.

The charge transfer process at the photoanode/electrolyte interface as well as the electron lifetime (τ) corresponds to the peak frequency in the Bode phase plot (Fig. 7). It is desirable to have a large electron lifetime as in the case of FCN_H . Additionally, the inclusion of FCN_H leads to a notable decrease in

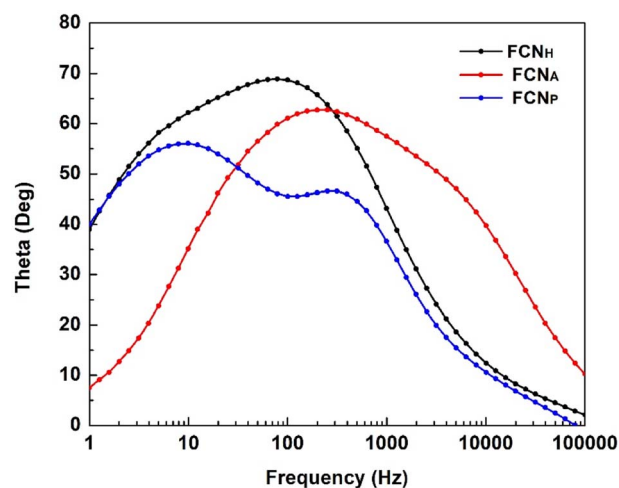


Fig. 7 Bode plot of the FCN DSSCs.

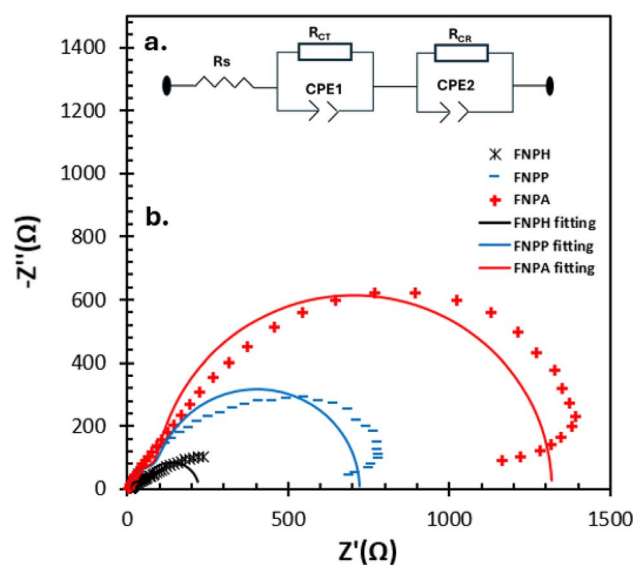


Fig. 8 (a) A full cell's equivalent circuit model and (b) matching Nyquist plots of FCN_H , FCN_A , and FCN_P -based DSSCs in the dark at a bias of V_{OC} between 0.1 Hz and 100 Hz with an amplitude of 10 mV.

charge transfer resistance, thus resulting in the enhancement of the overall performance of the DSSC.

The effective electron lifetime (τ) in a DSSC can be determined by analyzing the maximum frequency in the Bode plot using eqn (1) (ref. 55 and 56)

$$\tau = \frac{1}{2\pi f_{\text{max}}} \quad (1)$$

Table 2 Current–voltage characteristics of FCNs in DSSCs

| Sample | Electrolyte | V_{max} | I_{max} | V_{OC} [V] | I_{SC} [mA cm^{-2}] | FF [%] | PCE [%] |
|-----------------------|---------------------------|------------------|------------------|---------------------|---|--------|---------|
| FCN_H | I^-/I_3^- | 0.29 | 0.35 | 0.43 | 0.52 | 0.45 | 0.10 |
| FCN_P | | 0.11 | 0.25 | 0.21 | 0.43 | 0.31 | 0.03 |
| FCN_A | | 0.22 | 0.12 | 0.36 | 0.20 | 0.39 | 0.03 |



Table 3 Calculated charge transfer resistance and lifetime parameters

| Samples | R_S (Ω) | R_{CT} (Ω) | R_{CR} (Ω) | τ (ms) |
|------------------|--------------------|-----------------------|-----------------------|-------------|
| FCN _H | 47.79 | 11.84 | 161.7 | 2.01 |
| FCN _P | 132.7 | 6.398 | 582 | 0.504 |
| FCN _A | 140.2 | 6.141 | 1173 | 0.632 |

The Bode plot analysis of FCN_H reveals a peak at 79 Hz, corresponding to an estimated electron lifetime of 2 ms. These findings are depicted in Fig. 8 and summarized in Table 3. The longer electron lifetime observed in FCN_H can be associated with a reduced occurrence of recombination within the TiO₂ network, thus resulting in an enhanced electron diffusion through lower trapping states. This phenomenon accelerates electron transfer, increases electron density, and ultimately improves the overall performance of the device.

The Nyquist plots of FCN_H, FCH_A, and FCH_P-based DSSCs measured in the dark at forward bias (V_{OC}) are displayed in Fig. 8b. Each spectrum exhibits two distinct semicircles: the high-frequency region on the left side corresponds to the charge transfer at the electrolyte/counter electrode interface, while the intermediate-frequency region on the right side represents the back charge transfer from the photoanode to the electrolyte.^{57,58} Fig. 8a (inset) shows the equivalent circuit that was used to analyze the EIS spectra. The two interfaces were represented by a parallel combination of a capacitor and a resistance. The device's total series resistance (R_S), charge transfer resistance (R_{CT}) at the electrolyte/counter electrode interface, and charge recombination resistance (R_{CR}) at the photoanode/electrolyte interface are depicted in the figure. By fitting an equivalent circuit, the values of R_S , R_{CT} , and R_{CR} are obtained and shown in Table 3. In comparison to FCN_H (47.79 Ω), a large value of R_S (132.7 and 140.2 Ω) in FCH_A and FCH_P, respectively, suggests a slower charge transport in these two contexts. Although the R_{CT} value in FCN_H is marginally higher, all devices' R_{CT} values are consistent because the electrolyte and counter electrode were identical. In FCH_A, a noticeably high charge recombination resistance (1173 Ω) was found. This might be due to a gradual recombination in FCH_A. The study on the I - V curves of all the DSSCs indicates that FCN_H has higher power conversion efficiency compared to FCN_A and FCN_P. The obtained charge transfer resistance (R_{CT}) is much less in the case of FCN_H than FCN_A and FCN_P, and therefore is clearly associated with the reduced charge recombination and consequently the enhanced electron lifetime due to the abundance of oxygen containing functional groups.⁵⁹

4 Conclusion

In this study, we have demonstrated an easy H₂O₂-based FCN synthesis from an economical and available natural precursor teak leaf. H₂O₂ not only increases the degree of surface oxidation of the carbon framework, but also improves their stability in aqueous solution. The prepared all FCNs showed blue emission fluorescence with high aqueous stability. Then, DSSCs were fabricated using FCN_H, FCN_A and FCN_P as sensitizers and

their photovoltaic performance was evaluated. The FCN_H-based DSSCs showed superior performance over FCN_P-based and FCN_A-based DSSCs, achieving an open circuit voltage (V_{OC}) of 0.43 V, a short circuit current (I_{SC}) of 0.52 mA, and a Fill Factor (FF) of 0.45. This enhanced performance of FCN_H-based DSSCs was achieved through surface oxidation with H₂O₂. We anticipate that these bio-based FCNs could provide new insights into the fabrication process of DSSCs. Overall, the development of H₂O₂-assisted FCNs has opened new avenues for the application of carbon-based nanomaterials in various fields, and their potential uses are still being explored.

Conflicts of interest

There are no conflicts to declare.

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