



Cite this: *Chem. Soc. Rev.*, 2025, 54, 5735

DOI: 10.1039/d5cs90043c

rsc.li/chem-soc-rev

Correction: Electrochemical nitrogen fixation and utilization: theories, advanced catalyst materials and system design

Wenhan Guo,^a Kexin Zhang,^a Zibin Liang,^a Ruqiang Zou^{*a} and Qiang Xu^{*bc}

Correction for 'Electrochemical nitrogen fixation and utilization: theories, advanced catalyst materials and system design' by Wenhan Guo *et al.*, *Chem. Soc. Rev.*, 2019, 48, 5658–5716, <https://doi.org/10.1039/C9CS00159J>.

The authors regret that the pK_a of hydrazine was incorrectly given in the original article. The correct value should be 7.93. This means that eqn (8) and (9) and Fig. 3 were incorrect. The correct versions are shown below. This also applies to the sentence beginning "For hydrazine, the case is similar...", where the boundary should be given as 7.93.

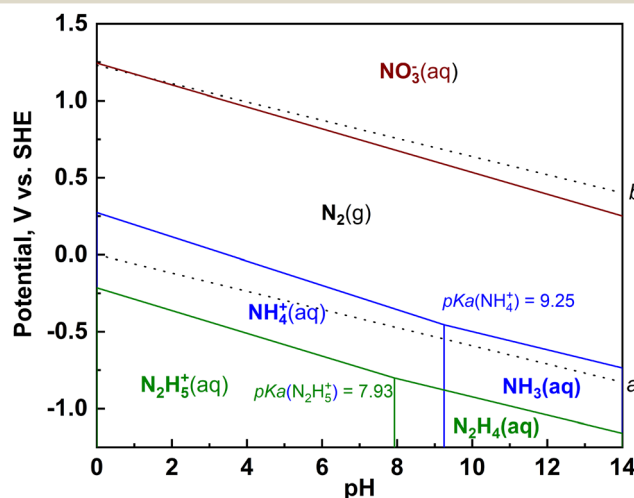
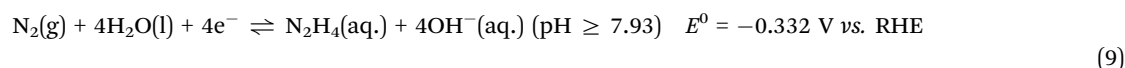
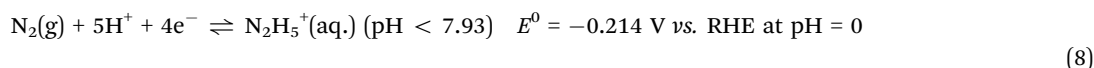


Fig. 3 Partial Pourbaix diagram of the N_2 - H_2O system including N_2 , NH_3 , N_2H_4 and NO_3^- . Region between dotted lines a (HOR/HER) and b (OER/ORR) correspond to the condition of water stability. Reproduced with standard electrode potentials in water at 298.15 K from ref. 48.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

^a Beijing Key Laboratory for Theory and Technology of Advanced Battery Materials, Department of Materials Science and Engineering, College of Engineering, Peking University, Beijing 100871, P. R. China. E-mail: rzou@pku.edu.cn

^b AIST-Kyoto University Chemical Energy Materials Open Innovation Laboratory (ChEM-OIL), National Institute of Advanced Industrial Science and Technology (AIST), Sakyo-ku, Kyoto 606-8501, Japan. E-mail: q.xu@aist.go.jp

^c School of Chemistry & Chemical Engineering, Yangzhou University, Yangzhou 225009, P. R. China. E-mail: qxuchem@yzu.edu.cn

