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ARTICLE

Covalent Functionalization Of Transition Metal Dichalcogenides With Perylene For Light Harvesting Devices

Received 00th January 20xx,
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

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This study investigates the optical and electronic properties of eight two-dimensional transition metal chalcogenides (TMDs)—MoS₂, WS₂, MoSe₂, WSe₂, MoTe₂, WTe₂, MoO₂, and WO₂—covalently functionalized with perylene, forming zero-dimensional/two-dimensional hybrid materials. Comprehensive characterization was conducted using techniques including XPS, Raman, EDX, TEM, and AFM. Optical properties were assessed using UV-Vis-NIR absorption and photoluminescence spectroscopy, while electronic properties were examined through cyclic voltammetry and field-effect transistor devices. Notably, the spectroscopic signatures of isolated perylene predominate in the hybrid materials, while WSe₂ and MoSe₂ displayed a novel band in the near-IR region, and MoTe₂ exhibited enhanced conductivity. Perylene significantly boosted absorption between 400–600 nm, leading to remarkable improvements in the photo-response and responsivities showing values exceeding 2x10⁵% and 2x10⁴ mA/W, respectively. The presented hybrid materials rival the best examples of non-covalent functionalization, underscore the potential of covalent functionalization as a powerful technique for further tailoring the optical and electronic properties of 2D materials.

Introduction

The properties of single-layer transition metal dichalcogenides (TMDs), including MoS₂, WS₂, MoSe₂, WSe₂, and MoTe₂, are well-documented and widely recognized.^{1–3} TMDs exhibit either metallic or semiconducting properties, which can be precisely controlled during synthesis and further modified through temperature and chemical modulation.³ The purest semiconducting phases are stable under ambient conditions for several months and can persist for years in the absence of oxygen.^{4,5} WS₂ and WSe₂ have direct band gaps of 2.11 eV and 1.65 eV, with carrier mobilities of approximately 50 cm²/V-s and 180 cm²/V-s, respectively.^{6,7} Likewise, MoS₂ and

MoSe₂ are generally n-type, with band gaps of 1.89 eV and 1.55 eV, and carrier mobilities exceeding 150 cm²/V-s and 50 cm²/V-s, respectively.^{8,9} For MoTe₂, the metallic phase is predominant; however, recent studies have successfully achieved the synthesis of the semiconducting phase through chemical vapor deposition (CVD). The semiconducting polytype of MoTe₂ exhibits a band gap ranging from 1.00 to 1.2 eV and a carrier mobility of 16.5 cm²/V-s.¹⁰ Conversely, WTe₂ is purely metallic, rendering its pristine phase less attractive for optical applications.¹¹ On the other hand, oxygen-based metal oxides, such as MoO₂ and WO₂, lack layered structures. Instead, ultrathin materials ranging from 1 to 100 nm in thickness have been synthesized using CVD, typically resulting in metallic properties.^{12–15} Additionally, pristine TMDs devices exhibit promising performance as field-effect transistors (FETs), with on/off ratios exceeding 10⁸,¹⁶ displaying n-type, p-type, or ambipolar behavior. Additionally, TMDs-based transistors are activated by light from the UV-Vis to near-IR regions, making them highly attractive for fabricating ultrasmall and ultraefficient flexible phototransistors.¹⁷ Exploring the properties of 2D-TMDs is a daunting task, given their complexity and extreme sensitivity. Due to their large surface area, 2D materials are highly susceptible to environmental interactions, leading to significant changes in properties that are influenced by factors such as the number of layers,¹⁸ size,^{19–21} strain,^{22,23} defects,²⁴ temperature,²⁵ interlayer angle,²⁶ substrate,²⁷ and polytype.²⁸ While minor variations can lead to substantial differences, complicating

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†Supplementary Information available: See DOI: 10.1039/x0xx00000x



direct comparisons between studies, this inherent variability also offers distinct advantages. It enables the fine-tuning of TMDs properties, thereby broadening their potential applications. One of the most common strategies for tailoring the properties of TMDs involves doping, which can be either intrinsic,^{29,30} altering the internal structure, or extrinsic,² modifying the surface environment. Focusing on the latter, modifications in the optoelectronic properties—such as absorption, photoemission, or photo-response—can also be achieved by anchoring organic molecules (0-dimensional structures) onto TMDs.³¹ Examples include porphyrins,^{32,33} phthalocyanines,^{34–36} polyamide carbon dots,³⁷ fullerenes^{38,39} pyrenes⁴⁰ or perylenes.⁴¹ These modifications, achieved through simple interactions with other species, hold great promise in energy harvesting and nanotechnology. Unfortunately, the absence of π - π interactions in TMDs, compared to other nanomaterials like graphene, presents a challenge.⁴² Aromatic molecules tend to self-assemble, causing substantial aggregation on the surface of TMDs and hindering their complete isolation. Chemical functionalization ensures a strong bond, confining the organic molecules on the TMDs and enabling their complete isolation. This strategy has primarily been applied to MoS₂ and WS₂ rather than other TMDs.^{43,44} Reactions involving sulfur derivatives, such as thiols,^{45–47} dithiolanes,^{48–50} sulfanes,⁵¹ or dithiols,⁵² are the most commonly used because they do not damage the material during the bonding process, unlike other method such as diazonium salt,⁴⁸ or decatungstate acylation.⁵³ Specifically, sulfur derivatives selectively target chalcogen vacancies without altering the crystallinity of the material. The level of functionalization achieved through sulfur derivative techniques tends to be low, but sufficient to isolate organic molecules on the TMDs surface, effectively forming a 0D-2D mixed material^{54,55} and allowing for accurate examination using conventional methodologies. This approach has been used to study electron and energy transfer processes in MoS₂ and WS₂ with various organic molecules such as pyrene,^{56–58} perylenes,⁵⁹ porphyrins,^{55–61} or phthalocyanines,⁶² as well as with 0D nanomaterials like fullerene derivatives⁵⁴ or polyamide carbon dots.⁶³ Moreover, covalently functionalized hybrid materials have demonstrated significant potential in areas such as catalysis,^{50,61} and energy conversion.^{56–62} Notably, MoS₂ and WS₂ exhibit distinctive and unique characteristics after functionalization with each chromophore, suggesting broader comparisons with other TMDs.⁶³ Notably, the covalent functionalization of TMDs with organic chromophores for applications in optical nanodevices remains limited. Previous studies demonstrated that the photo-response of WS₂ doubles when functionalized with pyrene⁵⁶ under white light irradiation. This effect is further amplified when MoS₂ or WS₂ are functionalized with zinc porphyrin (ZnP)⁵⁵ or when MoS₂ or MoSe₂ are combined with PCBM,⁵⁴ improving the photo-response by more than tenfold and yielding final responsivities of 4.9×10^{-5} , 2.3×10^{-5} , 1.4 , and 2.1×10^{-2} A/W, respectively. Moreover, the only electron-acceptor molecule tested so far is PCBM, a spherical molecule that hinders optimal contact with the 2D material. A planar molecule would enable better contact with 2D materials, as its entire π - π skeleton can be deposited on the TMDs surface, maximizing

electronic interactions. In this context, perylene derivatives are electron-acceptor molecules with rich chemistry, making them promising candidates for further exploration.

To establish a standardized framework and advance the frontiers of covalent functionalization, we conducted an in-depth study of perylene modified with eight different 2D-TMDs materials. Specifically, we functionalized perylene derivatives with single layers of semiconducting MoS₂, MoSe₂, MoTe₂, WS₂, and WSe₂, as well as with metallic WTe₂, and ultrathin MoO₂ and WO₂. Our analysis focused on the electronic and optical properties, drawing attention to both the differences and similarities among the various functionalized TMDs. While perylene generally reduces the intrinsic emission of TMDs, we noted the emergence of new perylene-TMDs excitations, particularly in the selenide-based hybrid materials. Additionally, under light excitation at 450–550 nm, targeting perylene, the TMDs hybrids significantly alter their electronic states, largely enhancing the photo-response in MoS₂, MoSe₂, WSe₂, MoTe₂, and WO₂, whereas metallic WTe₂ and MoO₂ remained unaltered. Each hybrid material presents distinct advantages, targeting specific applications and proving attractive for various fields in nanotechnology.

Results and discussion

1 Synthesis

All TMDs (MoS₂, WS₂, MoSe₂, WSe₂, MoTe₂, WTe₂, MoO₂ and WO₂) were synthesized by CVD on SiO₂/Si substrates using our established CVD methodology (see Table S1).⁵⁴ Next, 3,4,9,10-Perylenetetracarboxylic dianhydride was coupled with cysteine via amidation, resulting in the formation of a perylene diamide bi-substituted.^{64,65} Subsequently, the TMDs on the substrate were immersed in perylene and stirred slowly during 4 days at temperatures between 40–50 °C, resulting in the formation of Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂, Per-MoTe₂, Per-WTe₂, Per-MoO₂, and Per-WO₂ hybrid materials. After this period the non-covalent species were extensively cleaned with dimethylacetamide (DMAC). During the reaction, sulfur reacts with the chalcogen vacancies on the surface of the TMDs, resulting in the presence of perylene across the entire surface of the TMDs, as shown in Figure 1. Due to the considerable length of perylene diimide between the two sulfur atoms, measuring 16.5 Å, it may cover a circular area of 854 Å². This increases the the probability that both sulfur atoms will find

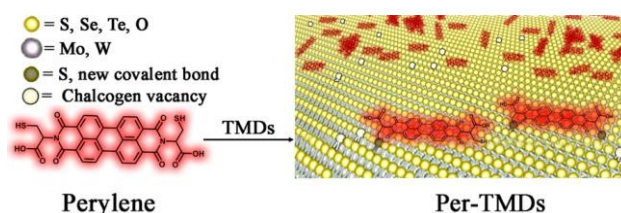


Figure 1. Representative reaction route for obtaining Per-TMDs hybrids.



insertion, which supports the covalent functionalization. It should also be noted that perylene signatures were observed at 1296 cm^{-1} , 1374 cm^{-1} , and 1573 cm^{-1} across all Per-TMDs hybrid materials.

Figure 2b shows the spectroscopic Raman mappings, highlighting the most intense modes of each TMDs and perylene. It is evident that the signatures of perylenes are exclusively on top of the nanosheets, suggesting that all non-covalently bonded species have been removed during the cleaning process, and only the strongly attached, covalently bound species remain.

2.2 XPS and IR

IR spectroscopy reveals rich signatures of perylene diimide. However, this technique was not applicable for evaluating Per-TMDs hybrid materials. It should be noted that the single layers of TMDs are isolated over a large area on a robust SiO_2 substrate, which reduces detectivity. Consequently, the absence of the already weak S-H signature in perylene is likely due to it falling below the detection limit.

In contrast, X-ray photoelectron spectroscopy (XPS) offers valuable insights into the bonding nature within hybrid materials and is commonly used to chemically analyze the interactions between perylene and TMDs. Given the large number of materials studied, this manuscript focuses on a representative material, Per- MoS_2 (see Figure 3), while the other hybrids and pristine materials are presented in Figure S1.

Firstly, the S 2p spectrum of pristine MoS_2 can be deconvoluted into five components: two peaks at 163.4 and 162.2 eV, which correspond to crystalline MoS_2 , and two peaks at 163.1 and 161.9 eV, associated with sulfur vacancies (see Figure 3a). All these bands exhibit similar areas, providing evidence that the basal plane of pristine MoS_2 contains significant defects. The fifth peak, observed at 168.0 eV, corresponds to oxidized sulfur species. After functionalization, the bands related to defects disappear, and the oxidized sulfur peak is significantly reduced (see Figure 3a, inset). This result supports the occurrence of covalent functionalization between the cysteine groups of perylene and the defective regions of the material.⁵⁷ A similar description applies to WS_2 .⁵⁶

The C-S-H bond from thiols, like free perylene, typically appears at 164 eV, merging with the sulfur $2p_{1/2}$ peak in Per- MoS_2 or Per- WS_2 , challenging its obtention.^{56,57} In TMDs where the chalcogen is not sulfur, the thiol signature was only marginally observed, even though the presence of carbon and nitrogen confirms the presence of perylene, suggesting that has been employed for covering a chalcogen vacancy.⁷⁵ This phenomenon has been previously observed in other sulfur-free TMDs functionalized with sulfur derivatives.^{54,76}

The C 1s signatures in all pristine TMDs appear broad and weak, indicating the presence of few amorphous carbon impurities. In contrast, the full width at half maximum (FWHM) of the carbon 1s peak at 284.73 eV is reduced in the hybrid materials at the same time that the intensity increases. Specifically, the FWHM of the carbon 1s peak ranges from 2.55 to 3.13 eV in pristine TMDs, whereas these values range from 1.33 to 1.88 eV in the hybrid materials. More notably, a new band at 288.55 eV and 286.55 eV confirms the presence of an amide bond (N-C=O) characteristic of perylene⁷⁷ and the C-S

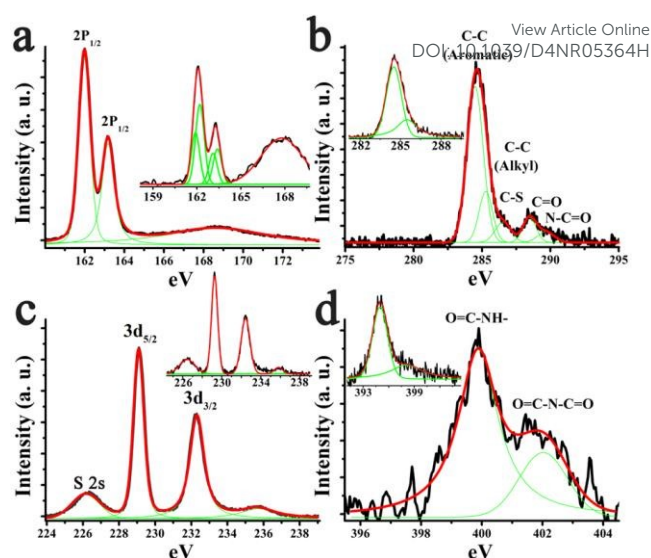


Figure 3. XPS spectra for (a) S 2p, (b) C 1s, (c) Mo 3d, and (d) N 1s of Per- MoS_2 , showing the acquired spectrum in black, deconvolution lines in green, and the sum of the spectra in red. The insets in Figure 3 correspond to the XPS analysis of (a) S 2p, (b) C 1s, (c) Mo 3d, and (d) N 1s of pristine MoS_2 .

bound (see Figure 3b). Similarly, the nitrogen 1s peak also exhibits a N-C=O band at 399.35 eV and an imine band at 402.00 eV (see Figure 3d).^{78,79} In conclusion, the covalent functionalization methodology has been unequivocally demonstrated.

For MoS_2 ,⁸⁰ MoSe_2 ,^{81,79} MoTe_2 ,⁸² and MoO_2 ,⁸³ strong signatures of molybdenum were recorded (see Figure 3c and S1). Specifically, the $3d_{5/2}$ peaks were observed at 229.25 eV, 228.61 eV, 228.23 eV, and 230.45 eV, while the $3d_{3/2}$ peaks appeared at 232.42 eV, 231.70 eV, 231.36 eV, and 233.00 eV, respectively. Additionally, an oxidized band related to MoO_3 species was recorded at 232.9 eV in MoTe_2 .⁵⁴ This band is not associated with oxidized species within the structure of MoTe_2 , but rather with the synthesis procedure, as MoTe_2 crystals grow hundreds of microns away from the bulk areas of MoO_3 . Given that our incident laser has a diameter of 0.5 mm^2 , we were unable to measure isolated crystals. In contrast, Raman spectroscopy confirms the absence of MoO_3 signatures once the surface of the layers is analyzed. This explanation also accounts for other oxide signatures, such as those recorded in WTe_2 . The $4f_{7/2}$ peaks of tungsten (W) in WS_2 ,^{5,84} WSe_2 , WTe_2 , and WO_2 ⁸⁵ were recorded at 32.83 eV, 32.79 eV, 31.58 eV, and 36.12 eV, respectively, while the $4f_{5/2}$ peaks appeared at 34.80 eV, 34.94 eV, 33.79 eV, and 38.19 eV.

The selenium $3d_{5/2}$ peaks in MoSe_2 and WSe_2 were observed at 54.30 eV and 54.60 eV, while the $3d_{3/2}$ peaks appeared at 54.98 eV and 55.35 eV, in the same order.^{79,81} The tellurium $3d_{5/2}$ peaks were observed at 574.15 eV for MoTe_2 and 572.54 eV for WTe_2 ,⁵⁴ while the $3d_{3/2}$ peaks were recorded at 584.61 eV for MoTe_2 and 583.06 eV for WTe_2 . Oxygen 1s signatures in MoO_2 ⁸³ and WO_2 ⁸⁵ could be deconvoluted into two peaks, indicating interactions between oxygen and Mo or W. Specifically, the peak at 532.45 eV corresponds to the substrate, while the peaks at 530.78 eV and 531.01 eV are attributed to MoO_2 and WO_2 , respectively.

After functionalization, the signatures of Mo, W, Se, Te, and O atoms show negligible spectral changes.⁵⁴ Thermogravimetric analysis



(TGA) is widely employed to understand the level of functionalization in organic materials. However, this technique is not applicable to the current system. This is because the weight of the single-layer nanosheets (~ 1 nm) on the SiO_2 surface (1×10^6 nm thick) is negligible, making it impossible to obtain reliable data. On the other hand, the XPS intensities of each component enable the calculation of the level of perylenes onto TMDs. Per-MoTe₂ had the highest perylene loading, followed by Per-MoSe₂, Per-WS₂, Per-WSe₂, and Per-MoS₂, with ratios of 1 perylene per 8, 10, 20, 16, and 20 TMDs units, respectively. According to the unit cell dimensions of TMDs, the areas of MoS₂, WS₂, MoSe₂, WSe₂, and MoTe₂ are 8.66, 8.60, 9.43, 9.31, and 10.73 Å², respectively, while the area of perylene is approximately 40 Å². Therefore, a single perylene molecule can cover 4.6, 4.7, 4.2, 4.3 and 3.7 unit cells of MoS₂, WS₂, MoSe₂, WSe₂, and MoTe₂, respectively. 1 functional group of perylene per every 20 MoS₂, 20 WS₂, 10 MoSe₂, 16 WSe₂, and 8 MoTe₂ unit cells correspond to coverages of 23%, 23%, 42%, 26%, and 46%, respectively. This coverage is sufficient to detect perylene while ensuring that the molecules remain isolated. The results for other materials, such as MoO₂ and WO₂, are uncertain because they are not layered materials.

3 Morphological characterization

3.1 AFM

AFM reveals surface roughness on the nanosheets across all analyzed Per-TMDs, while the substrate remains unaffected, suggesting the presence of perylene exclusively on the basal plane of the nanosheets.^{54,55,86} The variation in surface roughness is less than 3 nm, with these features uniformly distributed across the basal plane of the nanosheets. This suggests that perylenes form pseudo-monolayer films when covalently attached to Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂, Per-MoTe₂, or Per-WTe₂ (see Figure 4 a-f). The variation in the perylene nanofilm thickness by 1 to 2 nm can be attributed to the different ratio of chalcogen vacancies around the surface of the TMDs crystals. In materials with a higher number of chalcogen vacancies, the level of functionalization increases, causing perylene to adopt a vertical orientation rather than a horizontal one. This observation aligns with the dimensions of perylene (0.3 x 0.5 x 1.6 nm³). Additionally, contact mode can sweep away any non-covalent functionalization species, ensuring that all the roughness observed is related to strong interactions between the organics and nanosheets.

Conversely, Per-MoO₂ and Per-WO₂ (Figure 4 g and h) lack a purely two-dimensional character, leading to the formation of ultrasmall 2D structures with thicknesses of approximately 15 nm and 50 nm, respectively.

When Per-MoS₂, Per-WS₂, Per-MoSe₂, and Per-WSe₂ were annealed at 600 °C in the presence of the corresponding chalcogen, all the organic species associated with perylene were removed, recovering the original smooth surface characteristic of pristine TMDs (see Figure S2a-d). Additionally, pristine MoTe₂, MoO₂, and WO₂ can be observed in Figure S2e-g. In both cases, the thickness of the nanosheets is clearly reduced.

3.2 TEM

Transmission Electron Microscope (TEM) was used for atomic resolution imaging, clearly demonstrating that the hexagonal lattice

characteristic of semiconducting TMDs is preserved after functionalization with perylene.^{55,54} Atoms of molybdenum, tungsten, selenium, and tellurium were visualized at 1.5 million magnification, but lighter atoms such as sulfur, oxygen, carbon, and nitrogen show low contrast, making them invisible to our instrument. Consequently, the hybrid materials exhibited structures similar to those of the pristine materials (see Figure 4i-m for Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂, and Per-MoTe₂, respectively). However, energy-dispersive X-ray spectroscopy (EDX) integrated into TEM confirms the presence of perylene in all TMDs. The studied transition metals (Mo or W) and chalcogens (O, S, Se, or Te) were shown with similar intensities. Additionally, carbon, nitrogen, oxygen, and sulfur were unambiguously detected on the basal plane of suspended TMDs, far from the amorphous carbon edges, confirming the incorporation of perylene into all Per-TMDs (see Figure 4n).

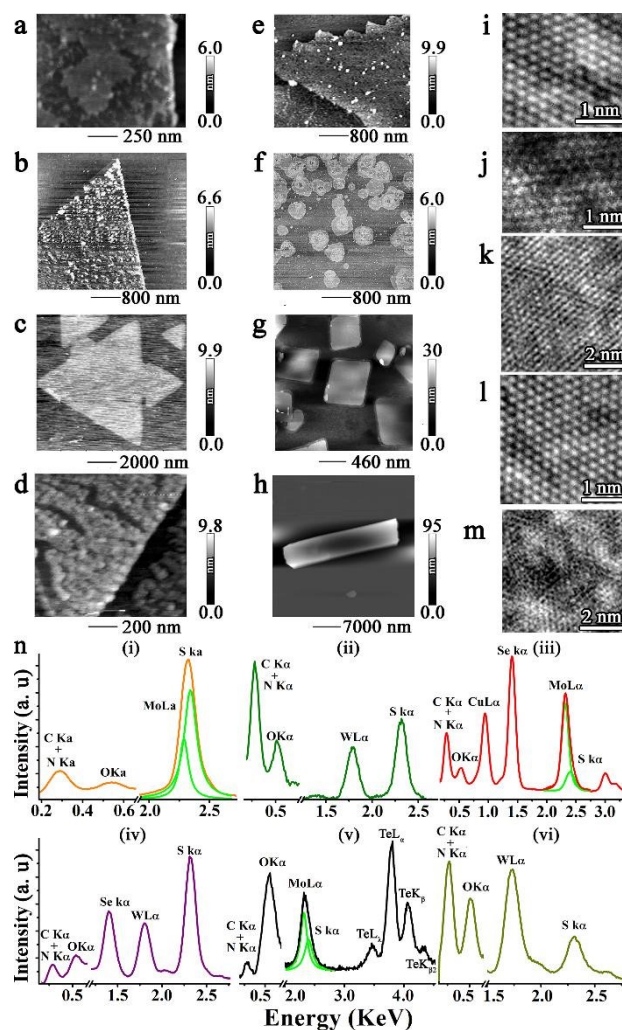


Figure 4. AFM images at different magnifications for (a) Per-MoS₂, (b) Per-WS₂, (c) Per-MoSe₂, (d) Per-WSe₂, (e) Per-MoTe₂, (f) Per-WTe₂, (g) Per-MoO₂, and (h) Per-WO₂. High-magnification TEM images of (i) Per-MoS₂, (j) Per-WS₂, (k) Per-MoSe₂, (l) Per-WSe₂, and (m) Per-MoTe₂. (n) EDX mapping in the same regions for (i) Per-MoS₂, (ii) Per-WS₂, (iii) Per-MoSe₂, (iv) Per-WSe₂, (v) Per-MoTe₂ and (vi) Per-WO₂.



4 Optical properties

4.1 ABSORPTION

UV-Vis-Near-IR spectroscopy confirms the semiconducting nature of MoS₂, WS₂, MoSe₂, WSe₂, and MoTe₂, as these materials exhibit pronounced absorption bands between 400 nm and 1205 nm (Figure 5a). Semiconducting TMDs typically exhibit two excitonic transition bands at lower energies, associated with spin-orbit splitting, and additional bands at higher energies corresponding to electronic transitions between the highest energy states in the valence band and the lowest energy states in the conduction band.⁸⁷ For MoS₂,¹⁹ the bands were recorded at 397, 445, 611, and 673 nm while for WS₂²⁰ are found at 422, 459, 529, and 640 nm. Next, MoSe₂⁵⁴ exhibited bands at 453, 710, and 816 nm, and WSe₂⁵⁴ at 430, 631 and 690 nm. In addition, MoTe₂ also shows the semiconducting polytype with bands at 684, 935, and 1205 nm.⁵⁴ The absence of any bands in MoO₂, WO₂, or WTe₂ can be attributed to their metallic character.⁵⁴ The bands of perylene were observed at 484, 514, and 559 nm in liquid media, corresponding to the electronic S₀₋₁ transitions at 0-0, 0-1, and 0-2 vibrations, respectively.^{59,88,89} In contrast, in solid states, the bands appeared very broad, merging and becoming almost indistinguishable (see Figure 5a). These spectroscopic modifications suggest that in liquid states, perylenes are molecularly insolated, whereas in solid states, there is significant perylene-perylene aggregation. On the other hand, the perylene signatures within the hybrid materials were perfectly distinguishable, like free perylene in liquid media. For example, perylene bands appear at 457, 487, and 519 nm in Per-MoS₂, with a shift of ±3 nm in other hybrid materials. Conversely, the intensity of the TMDs signatures is slightly modified, whereas a global red shift has been observed. Specifically, all bands shift by 2-7 nm in MoS₂, 1-4 nm in WS₂, 2-8 nm in MoSe₂, and 1-7 nm in WSe₂. Only MoTe₂ displays a blue shift of its bands at 1205 nm, achieving a value of 1188 nm in Per-MoTe₂. These changes in the spectroscopic morphology indicate strong interactions between the orbitals of perylene and TMDs.^{56,59,60,63}

4.2 EMISSION

The PL emission of solid Perylene aggregates is broad, extending from 650 to 850 nm, with a maximum at 732 nm (1.70 eV). This contrasts sharply with isolated perylenes, which exhibit emission at 550 nm (2.26 eV) in liquid conditions, where π-π interactions are suppressed.⁵⁹ The solid-state PL emission of the hybrid materials exhibits a perylene signature similar to that of isolated perylenes, with emission bands appearing between 530 and 615 nm (1.96 to 2.35 eV). More specifically, Per-MoS₂ exhibits three perylene emission bands at 529, 571, and 611 nm (2.34, 2.17, and 2.03 eV), forming a mirror image of its UV-Vis absorption spectrum. Per-WS₂, Per-MoO₂, and Per-WO₂ show similar signatures, with negligible shifts. Conversely, the signatures of Per-MoSe₂ and Per-WSe₂ are slightly shifted and broadened, with peaks at 585, 616, and 666 nm (2.12, 2.01, and 1.86 eV), indicating partial perylene-perylene interactions. These interactions are moderate in Per-MoTe₂ and pronounced in Per-WTe₂. Figure 5b-f displays a representative spectrum of each material, extracted between 10-100 Kelvin, whereas Figure S3 provides the full spectra across the entire temperature range. By employing the Varshni equation (see Eq. S1),

the TMDs band gap of Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂ was calculated to be 1.91, 2.04, 1.60 and 1.71 eV, respectively.

The PL emission associated with the semiconducting properties of TMDs sharply decreases after functionalization with perylene. Starting with pristine materials, the signatures of MoS₂, WS₂, MoSe₂, WSe₂ and MoTe₂ are shown at 664, 620, 779, 752 and 1125 nm (1.87, 2.00, 1.59, 1.65 and 1.1 eV), respectively. Following functionalization, emission peaks were recorded at 661, 622, 775, 742, and 1097 nm (1.88, 1.99, 1.59, 1.67, and 1.13 eV) for Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂, and Per-MoTe₂, in the same order.⁵⁴ At low temperatures, several excitonic species, including neutral excitons, trions, and biexcitons, can typically be distinguished, influenced not only by temperature but also by incident power. Although these signatures are recorded in pristine materials, the spectral morphology remains invariant in the hybrid materials, probably due to the interaction with perylene. More interesting, Per-WS₂, Per-MoSe₂, and Per-WSe₂ (Figure 5c-e) exhibit an additional band at lower energies after functionalization, at 957, 1189, and 1006 nm (1.30, 1.04, and 1.23 eV), respectively, which is likely related to electron-hole recombination between the perylenes and TMDs.⁵⁴ This band saturates at high power densities, showing α<1, as expected due to the relatively small loading of perylenes. A representative example of the power dependence of Per-WS₂ is provided in the Supplementary Information (Figure S4a-b).

Next, measuring the lifetimes of the hybrid materials proved challenging due to significant quenching effects from perylene.⁹⁰ Nonetheless, it is important to note that all pristine materials typically exhibited lifetimes shorter than 0.1 ns.⁵⁴ Perylene exhibited a lifetime of 3 ns, while the overall lifetime of Per-MoSe₂ decreased to 0.5 ns, resulting in an ultrafast decay in other materials that could not be accurately measured with our instrument (see Figure S4c). This reduction in lifetime might be related to electron-hole or energy transfer between perylenes and TMDs.

Further, photoluminescence excitation (PLE) assays performed on Per-MoS₂, Per-WS₂, Per-MoSe₂ and Per-WSe₂ exhibit maximum intensity between 500–600 nm, following the same pattern as the absorption spectrum of perylene (see Figure S5). Conclusively, in the particular configuration presented in this contribution, signatures of perylene were recorded as isolated species, which tend to block the intrinsic fluorescence properties of all TMDs.

5 Electrochemical properties

Electrochemical assays are of great interest as they offer insights into the reductive and oxidative potentials of each species, providing a fundamental basis for understanding the electronic interactions between perylenes and TMDs. Firstly, the square wave voltammetry (SWV) in Figure 6a shows the first oxidative and reductive potentials of perylene at 1.27 and -0.94 V corresponding[‡] to LUMO and HOMO values of -3.86 and -6.07 eV, respectively, in the energy vacuum diagram, assuming a Fermi level of -4.8 eV for ferrocene (Fc/Fc⁺). The interpretation of TMDs is more complex. For example, MoS₂ and MoSe₂ exhibit first oxidation potentials at 0.3 V and 0.08 V, respectively, which do not correspond to the valence band maximum (VBM) but rather to defective regions of the material.⁵³ Similar



behavior has been observed in WS_2 and WSe_2 , with oxidation potentials recorded at 0.04 V and 0.09 V, respectively.^{91,92} These short peaks are not related to the VBM but are instead associated with trap states, which may also be relevant for explaining the electronic processes involved in the interaction with perylene. The VBM for MoS_2 , MoSe_2 , WS_2 , and WSe_2 are associated with oxidation potentials of 0.46, 0.25, 0.86 and 0.27 V, showing values of -5.26, -5.05, -5.66 and -5.07 eV. Similarly, the conduction band minimum (CBM) for MoS_2 , MoSe_2 , WS_2 , and WSe_2 corresponds to the reductive potentials at -1.39, -1.30, -1.40, and -1.41 V, yielding energy values of -3.41, -3.50, -3.51, and -3.39 eV, respectively.^{54, 91, 92} Table 1 shows the most relevant oxidation and reduction peaks for perylene, MoS_2 , MoSe_2 , WS_2 and WSe_2 . Regarding Per- MoS_2 , Per- MoSe_2 , Per- WS_2 and Per- MoSe_2 the reduction and oxidation peaks related with the CBM and VBM of the TMDs or perylene exhibit minimal variation, less than ± 0.2 V. Additionally, electrochemical bands associated with trap states were significantly reduced in the hybrid materials, suggesting

that the inclusion of perylene effectively passivates these energetic states.⁹³ Notably, the new emissive bands observed in WS_2 , MoSe_2 , and WSe_2 at 1.30, 1.04, and 1.23 eV, respectively, align well with the differences between the redox reduction potential of perylene at -0.94 V and the trap states of WS_2 , MoSe_2 , and WSe_2 at 0.36, 0.08, and 0.27 V, resulting in electrochemical gaps of 1.30, 1.02, and 1.21 eV. This suggests that the presence of these new bands corresponds to electron-hole pairs formed between perylene and the TMDs (see Figure S6a).

Through the integration of optical and electrochemical data, an orbital diagram, depicted in Figure 6b, may be constructed to elucidate the behavior of the photodevices. Before proceeding, it is important to highlight for readers with a background in physics or related fields that our nomenclature might be puzzling, as perylene behaves as a molecule with distinct molecular orbitals, while the

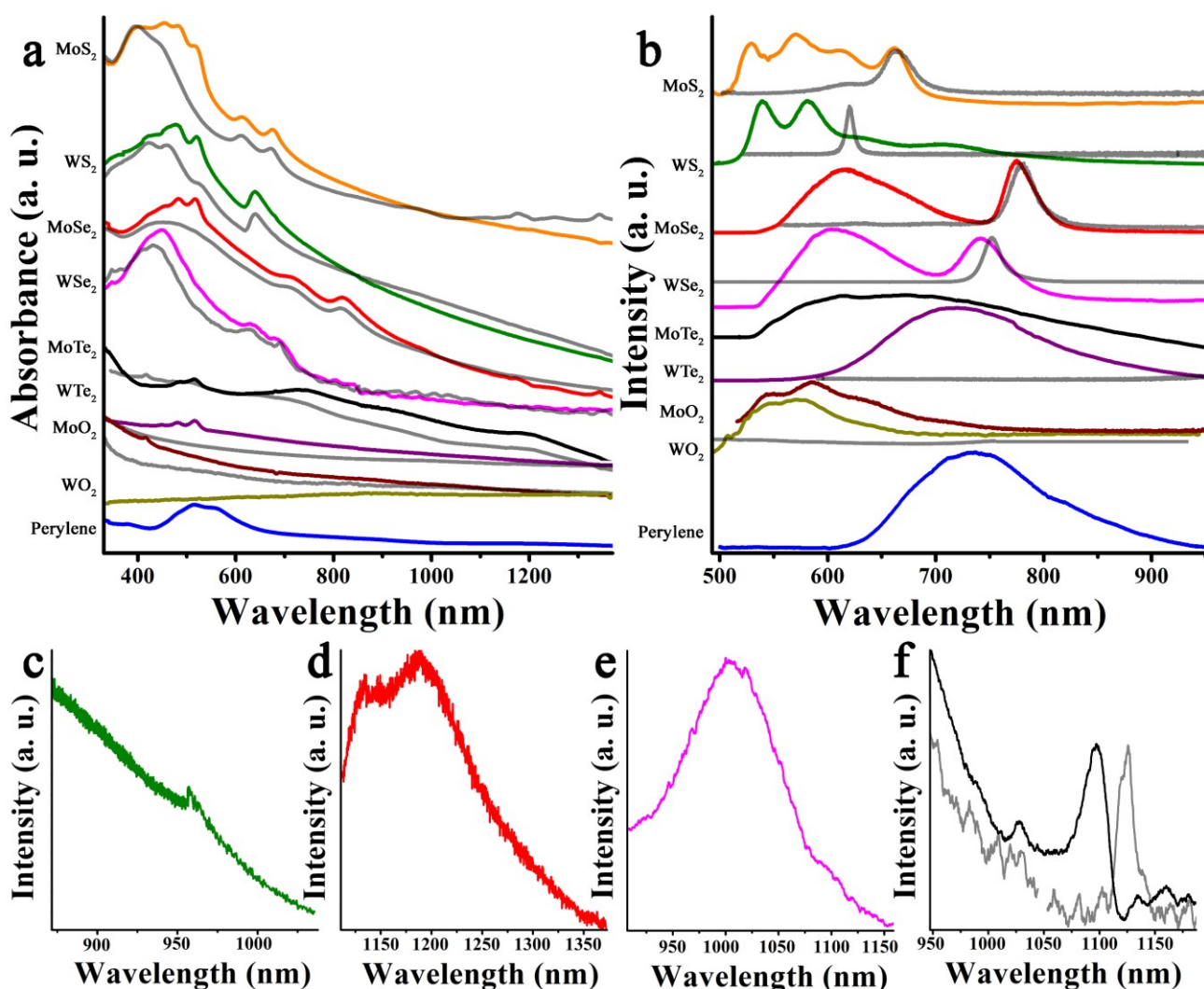


Figure 5. (a) Absorption and (b) emission spectra ($\lambda=488$ nm; 80 K) of TMDs: Per- MoS_2 (orange), Per- WS_2 (olive), Per- MoSe_2 (red), Per- WSe_2 (pink), Per- MoTe_2 (black), Per- WTe_2 (purple), Per- MoO_2 (wine), Per- WO_2 (dark yellow), and perylene (blue). The intensity of all materials has been normalized with pristine TMDs (gray) under the same conditions. Magnification of Near-IR emission bands of (c) Per- WS_2 , (d) Per- MoSe_2 , (e) Per- WSe_2 and (f) pristine MoTe_2 (gray) and Per- MoTe_2 (black).



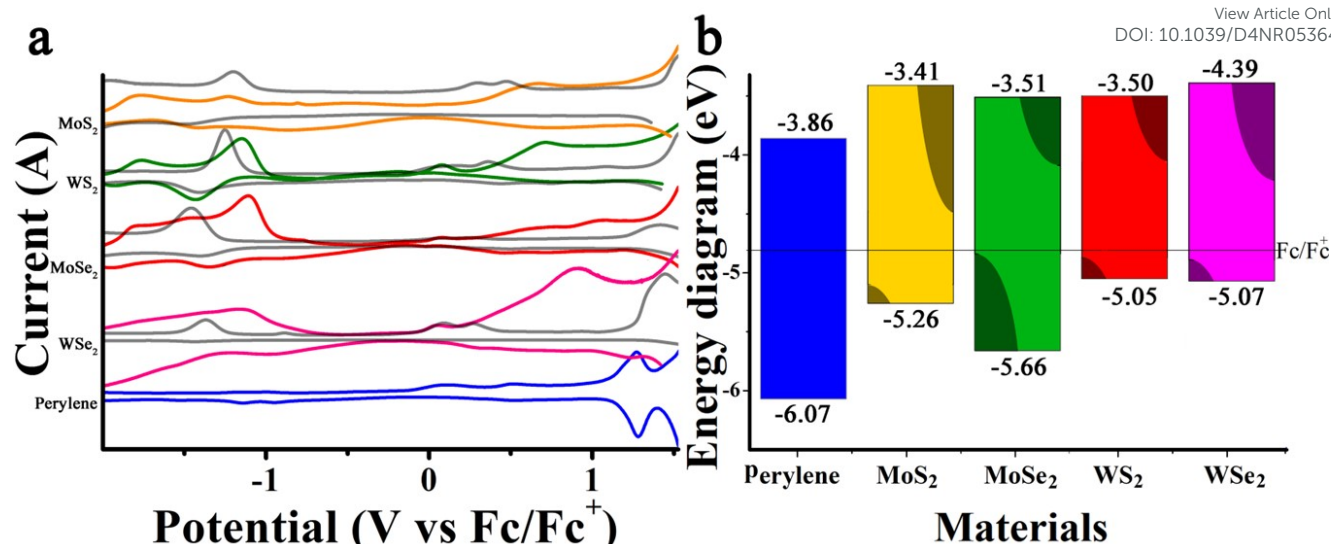


Figure 6. (a) SWV of Per-MoS₂ (orange), Per-WS₂ (olive), Per-MoSe₂ (red), Per-WSe₂ (pink), Per-MoTe₂ (black), Per-WTe₂ (purple), Per-MoO₂ (wine), Per-WO₂ (dark yellow), and perylene at 514 nm (blue). All materials have been normalized with pristine TMDs (gray) under the same conditions in dry and deoxygenated acetonitrile, with 0.1 M of tetrabutylammonium hexafluorophosphate. (b) Electrochemical energy vacuum diagram.

TMDs possess energy bands, resulting in a hybrid system of mixed orbitals and bands. Having established this premise, the explanation begins by noting that the HOMO and LUMO levels of perylene are deeper than the VBM and CBM of the TMDs, leading to a type II alignment for MoS₂, WS₂, MoSe₂, and WSe₂. Given that the electrochemical assays reveal a significant number of hole-trap states, the TMDs materials contain energy states below the LUMO of perylene, causing the system to behave like a pseudo-type I alignment. Specifically, under dark conditions, spontaneously generated electrons in the CBM in the TMDs can transfer to the LUMO of perylene, depleting the number of carriers, while spontaneously generated electrons in the LUMO and holes in the HOMO of perylene may transfer to trap states near the conduction or valence bands of TMDs (see an example in see Figure S6b and c). Under light illumination, however, the number of holes transferred

from the HOMO of perylene might saturate the surrounding trap states, thereby increasing the population of holes in the valence band and consequently enhancing conductivity or carrier mobility (see Figure S6d). The electrons excited to the LUMO of perylene by the light do not have enough energy to reach the conduction band of MoS₂, WS₂, MoSe₂, or WSe₂, but they might contribute to saturating a significant portion of the TMDs hole trap states. As a result, the photogenerated electrons in the conduction band of the TMDs will not be easily trapped, leading to increased carrier mobility. Finally, due to the large band gap of perylene, energy transfer is also plausible. Under these circumstances, a photon absorbed by perylene may excite an electron from the valence band or trap states of the TMDs to their conduction band or trap states, resulting in conditions similar to those expected in a type I alignment (see Figure S6e).

Table 1. Oxidative and reductive potentials via sweep pulse voltammetry.

Material	Oxidation (V)			Reduction (V)			HOMO (eV)	LUMO (eV)	GAP (eV)		
Perylene	1.27			-0.94	-1.139	-2.3	0.94	-6.07	-3.86	2.21	
MoS ₂	0.3	0.46	1.51	-1.39	-2.1			-5.26	-3.41	1.85	
MoSe ₂	0.08	0.25	1.41	-1.3				-5.05	-3.5	1.55	
WS ₂	0.04	0.16	0.36	0.86	-1.4	-1.29	-2.11	0.04	-5.66	-3.51	2.15
WSe ₂	0.09	0.27	1.45		-1.41				-5.07	-3.39	1.68
Per-MoS ₂	0.47	0.66	1.31	-0.86	-1.41				-5.46	-3.94	1.52
Per-MoSe ₂	0.08	0.73	1.06	1.21	-0.917	-1.339			-4.88	-3.883	1.0
Per-WS ₂	0.07	0.71	1.001		-0.82	-1.426	-1.91		-5.51	-3.374	2.14
Per-WSe ₂	0.055	0.913	1.22		-0.96	-1.52			-4.855	-3.84	1.02



In summary, under dark conditions, perylene has a minimal effect on the valence band of TMDs due to the presence of trap states (see Figure S6c). However, it negatively affects the conduction band in the TMDs hybrids, leading to a reduction in electron mobility (see Figure S6b). Under light illumination, perylene may inject hole carriers into the TMDs through ground-state interactions, while the negative effects observed in the excited states are significantly diminished or even reversed (see Figure S6d).

6 Devices properties

6.1 PHOTOTRANSISTOR

3-terminal devices with a FET configuration were fabricated by contacting a single layer of semiconducting TMDs, grown via CVD, with Bi/Au electrodes (20 nm/50 nm) through metal deposition on a SiO₂ substrate with a thickness of 270 nm. The devices featured a channel length of 5 μm targeting monocrystalline single layers of MoS₂, WS₂, MoSe₂, and WSe₂, while MoTe₂ was polycrystalline. For WTe₂, a few-layered nanofilm was used. Regarding MoO₂ and WO₂, the channel contains nanomaterial flakes with thicknesses of 15 nm and 50 nm, respectively (see Figure S7). The same TMDs-based devices were measured under identical conditions, both before and after functionalization with perylene (see Figure 7a). While replicating the exact conditions for the last three materials proved challenging, all parameters were meticulously standardized to ensure comparability with our prior research based on the functionalization of TMDs with PCBM.⁵⁴ A brief comparison is provided later.

First, the conductivity of the materials significantly decreases after functionalization, with Per-MoS₂, Per-WS₂, Per-MoSe₂, Per-WSe₂, Per-MoO₂, and Per-WO₂ showing 1000-, 100-, 50-, 40-, 10-, and 10,000-fold reductions, respectively, compared to the pristine materials under dark conditions (see Figure 7b-i). This phenomenon is consistent with the discussion in Section 5 and Figure S6b-c, which explains how perylene reduces the number of available carriers in the TMDs. Although low conductivity can present challenges in various systems, the signal-to-noise ratio remains unaffected, making these materials still of great interest for the fabrication of photodetectors and other nanotechnology applications. Interestingly, the dark current of Per-MoTe₂ increases by approximately 10-fold (see Figure 7f). Despite being unable to evaluate the energy levels via electrochemistry, the literature indicates that the valence band of MoTe₂ is at 4.9 eV (via UPS)⁹⁴ with a bandgap of 1.1 eV. Therefore, the conduction band is estimated to be at approximately -3.8 eV,⁹⁵ which is close to the LUMO of perylene, suggesting a type I alignment. Consequently, the previously described negative impact of perylene on the conduction band is mitigated, leading to enhanced conductivity (see Figure S6f). Under light illumination at 532 nm, the photo-response of Per-MoS₂ increases more than 1000-fold compared to pristine MoS₂ (see Figure 7b). This indicates that hole transfer from the HOMO of perylene substantially covers the hole trap states of MoS₂, significantly increasing the charge carrier density (see Figure S6d). Outside the excitation range, perylene does not significantly inject carriers to MoS₂ (see Figure S6b and c), reducing the enhancement of the photo-response (e.g., only a 0.2-fold increase at 254 nm). A

similar effect is observed in Per-MoSe₂ and Per-WSe₂, with improvements in photo-response by 33 and 150-fold at 532 nm, respectively, compared to pristine MoSe₂ and WSe₂ (see Figure 7d and e).

The photo-response of Per-WS₂ (see Figure 7c) under perylene excitation (at 405 nm or 532 nm) decreases by 20% compared to pristine WS₂, suggesting that the LUMO of perylene can efficiently recombines with holes in the VBM or nearby trap states (see Figure S6a). The recombination between the trap states, below CBM of WS₂ and HOMO of perylene likely results in the emission of a new band recorded at 1.30 eV (see Figure 5c). However, the photo-response of Per-WS₂ increases outside the perylene excitation range. Specifically, at 650 nm, the photo-response of Per-WS₂ enhances 28-fold compared to pristine WS₂. This complex effect likely involves multiple factors, one significant possibility being that perylene is directly excited to the trap states in the conduction band of WS₂, allowing holes in perylene to transfer to the valence band of WS₂, thereby enhancing the system's conductivity (see Figure S6g).

Regarding Per-MoTe₂, the material exhibits a clear enhancement of its photo-response compared to pristine MoTe₂ (see Figure 7f). Per-MoTe₂ shows small but consistent photo-responses outside the perylene regions, with slight magnification at 532 nm and 407 nm when perylene is irradiated. This behavior is consistent with a type I alignment, where perylene acts as an antenna, enhancing the photo-response throughout the entire spectral region analyzed in this study (see Figure S6f). Lastly, some intriguing effects have been observed with intrinsically metallic materials, that typically exhibit negligible or no photo-response in their pristine forms. Per-WTe₂ (see Figure 7g) shows a negative photo-response, particularly under perylene excitation, with decreases of 90% at 532 nm and 25% at 650 nm. This effect is likely due to the trapping of holes and/or electrons by perylene, which reduces the conductivity in the hybrid material. In contrast, the photo-response of Per-WO₂ increases 8-fold at 407 nm compared to the pristine material (refer to Figure 7i), suggesting the injection of electrons or holes into the system. Finally, metallic MoO₂ remains insensitive to light after functionalization with perylene (see Figure 7h).

The recovery time after a light pulse is ultrafast in MoSe₂, WSe₂, MoTe₂, and WTe₂, making it challenging to accurately measure with our instruments. After functionalization, this ultrafast response is preserved, ensuring the good performance of the system. Conversely, the recovery time for pristine MoS₂ and WS₂ was critically slow, fitting well with a biexponential function. We observed fast recovery components with values of 1.0 and 21 seconds, and slow components of 31 and 165 seconds for MoS₂ and WS₂, respectively. The slow recovery components are attributed to trap states and defects that hinder rapid relaxation. After functionalization, the recovery times are partially shortened due to the trap state passivation by perylene, resulting in recovery times of 1 and 16 seconds for Per-MoS₂, and 1 and 13 seconds for Per-WS₂, respectively.^{34, 55, 56} A representative fitting for Per-MoS₂ is included in the Supporting Information as Figure S8. Although this is not our primary focus, such long-time recovery times are of interest for applications in neurological memory photodevices.



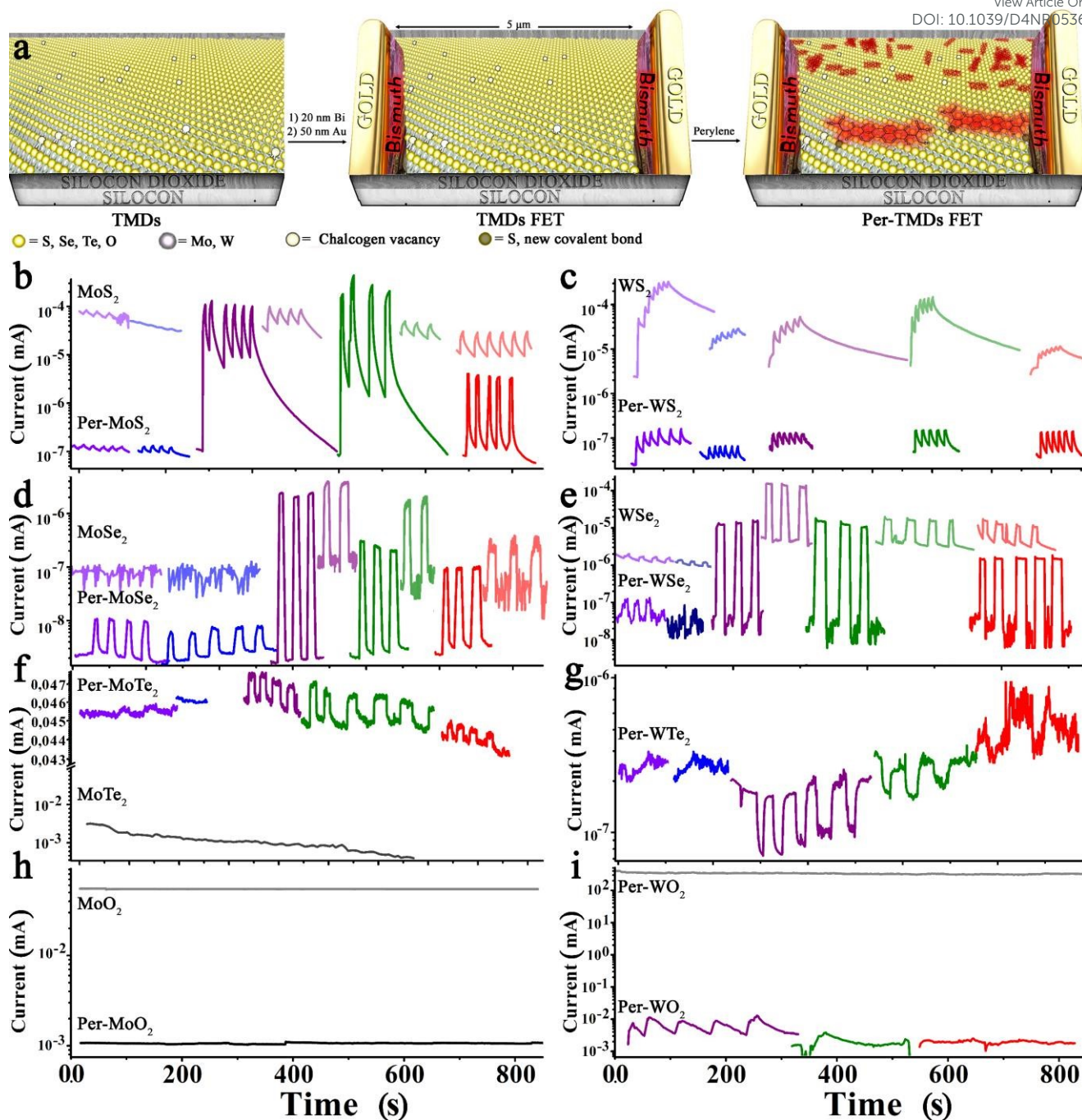


Figure 7. (a) Fabrication of Per-TMDs-based phototransistors. Photo-response wavelength dependence for (b) Per-MoS₂ (opaque line) and pristine MoS₂ (semitransparent line), (c) Per-WS₂ (opaque line) and pristine WS₂ (semitransparent line), (d) Per-MoSe₂ (opaque line) and pristine MoSe₂ (semitransparent line), (e) Per-WSe₂ (opaque line) and pristine WSe₂ (semitransparent line), (f) Per-MoTe₂ (opaque line) and pristine MoTe₂ (semitransparent line), (g) Per-WTe₂ (opaque line) and pristine WTe₂ (semitransparent line), (h) Per-MoO₂ (opaque line) and pristine MoO₂ (semitransparent line), and (i) Per-WO₂ (opaque line) and pristine WO₂ (semitransparent line) at 254 (violet), 365 (blue), 407 (purple), 532 (green), and 650 nm (red) light excitation pulses of 5s and V_{D,S} of 1 V. The power intensity for 254nm, 365 nm, 407 nm, 532 nm and 650 nm were 10 mW/cm², 10 mW/cm², 30 mW/cm², 17 mW/cm² and 7 mW/cm².

Based on factors such as incident light wavelength, laser power, channel dimensions, and applied voltage, several key parameters were calculated, including photo-response, responsivity, external quantum efficiency (EQE), specific detectivity (D*), superficial conductivity (σ_{sq}), transconductance (G_m), and carrier mobilities

(see Eq. S2–S8). This manuscript summarizes the best performance of the hybrid materials in Figure 9, while comprehensive results are detailed in the Supplementary Information (Tables S2–7). Optimal results were consistently achieved under 532 nm laser irradiation, with the highest photo-responses observed for Per-WSe₂ and Per-



MoSe₂, with values of $3.9 \times 10^{5\%}$ and $2.6 \times 10^{5\%}$ each. The best responsivity and EQE were recorded for Per-MoSe₂ and Per-MoTe₂, with values of 4.6 and 26 A/W, and 1.1×10^3 and $6.1 \times 10^{3\%}$, respectively. The greatest specific detectivity was obtained for Per-MoS₂, followed by Per-WSe₂ and Per-MoSe₂, with values of 1.1×10^{12} , 1.0×10^{11} , and 2.8×10^{10} Jones, respectively. Additionally, the highest transconductance and carrier mobilities were also achieved in Per-MoS₂, with values of $9.0 \mu\text{A/V}$ and $92.3 \text{ cm}^2/\text{V}\cdot\text{s}$ respectively. The best superficial conductivity was observed in the metallic material Per-MoO₂ with a value of $3.4 \times 10^{-3} \text{ S/sq}$, while among the semiconducting materials, Per-MoTe₂ showed the highest conductivity at $5.8 \times 10^{-6} \text{ S/sq}$, followed by Per-MoS₂ at $3.9 \times 10^{-8} \text{ S/sq}$. Volumetric conductivity, commonly expressed in S/m, is often incorrectly applied to 2D materials. In our data, this would suggest an improvement of more than 10^9 -fold, which renders direct comparisons invalid and should be avoided. Therefore, the Per-MoS₂ device appears to be the most effective across the majority of analyses.

6.2 TRANSFER CURVES AND OUTPUT CHARACTERISTICS

Other relevant parameters commonly evaluated in FET transistors include transfer curves and output characteristics. Specifically, transfer curves were fundamental in determining the carrier mobility discussed in Section 6.1. A representative example illustrating the performance of MoS₂ is provided in Figure 8, while data for the other materials are included in Supplementary Information, Figure S9.

Firstly, under dark conditions, the transfer curves of MoS₂ exhibit n-type semiconductor characteristics with on-off ratios spanning six orders of magnitude, a behavior that is retained in Per-MoS₂ (Figure 8a and b). Notably, this ratio becomes more pronounced following functionalization, sharply increasing from 40 V to 60 V, in contrast to the pristine material, where the current increases over the range of

-60 V to 60 V. Additionally, at negative voltages ranging from -40 to -80 V, Per-MoS₂ demonstrates enhanced conductivity compared to pristine material, likely due to hole injection into the valence band by perylene. Under light irradiation, carrier injection increases the current across all wavelengths, with the effect being most pronounced in the perylene region, at 407 and 532 nm.

WS₂ and Per-WS₂ display on-off current ratios of 4 and 3 orders of magnitude, respectively, similar to WSe₂ and Per-WSe₂, while maintaining an n-type character. Per-MoSe₂ exhibits an on-off ratio with one-order-of-magnitude improvement after functionalization, revealing a p-type semiconductor character. Upon light irradiation, Per-WS₂, Per-WSe₂, and Per-MoSe₂ exhibit a similar effect to Per-MoS₂, with carrier injection enhancing the current throughout the system. For Per-MoTe₂, Per-WTe₂ and Per-WO₂, the drain-to-source (V_{D-S}) current remains stable across a range of gate voltages (V_G), indicating that these devices are better characterized as robust photodetectors rather than phototransistors.

Finally, the output characteristics in Figure 8c and d reveal an asymmetry in the current flow for Per-MoS₂ compared to pristine MoS₂, indicating inhomogeneities caused by the presence of perylene.^{54,55} An asymmetry in the output characteristics should not be mistaken for a drawback. In fact, it is quite the opposite. The larger conductivity at positive voltages, which is also photoactive, is characteristic of photodiodes. Thus, Per-MoS₂ can function not only as a phototransistor but also as a photodiode, meaning that the current passing through the system can be modulated not only by the gate voltage but also by its direction and light exposure, further enhancing its potential as a nanodevice. The examined voltage range shows no current saturation, confirming that the devices can operate effectively across a broad window of drain-to-source voltages, both positive and negative, while demonstrating a strong response to gate voltage. A similar effect is observed in Per-WS₂, Per-MoSe₂, and Per-WSe₂. In contrast, MoTe₂, WTe₂, MoO₂, and WO₂ exhibit no significant changes with gate voltage variation, exhibiting symmetry in both positive and negative voltages, thereby confirming bidirectional drain-to-source current behavior.

6.3 COMPARATIVE ANALYSIS WITHIN THE FRAMEWORK OF ADVANCED TMDS PHOTODEVICES

Certainly, extensive research has focused on functionalizing TMDs for the fabrication of ultrasmall photodevices, especially phototransistors. The 0D-2D structures hold particular interest for this area of nanotechnology.³¹ For instance, PbS quantum dots on MoS₂ exhibits the best performance to our knowledge, with a remarkable photoresponsivity of $2 \times 10^6 \text{ A/W}$.⁹⁶ The highest reported value for 1D/2D structures is found in 1D Te/2D MoTe₂,⁹⁷ which is two orders of magnitude lower, similar to the 3D/2D hybridization between MoS₂ and GaN.⁹⁸

Close to this level of performance are organic-inorganic hybrids, MoS₂ functionalized non-covalently with perylene (MoS₂/Perylene) standing out due to its highest photoresponsivity of $1 \times 10^3 \text{ A/W}$ and an impressive photo-response of $1.5 \times 10^4 \%$,⁹⁹ followed closely by MoS₂ functionalized with zinc-phthalocyanine (MoS₂/ZnPc),³⁴ which shows values of $4.3 \times 10^2 \text{ A/W}$ and $3.0 \times 10^5 \%$. A brief reminder is that

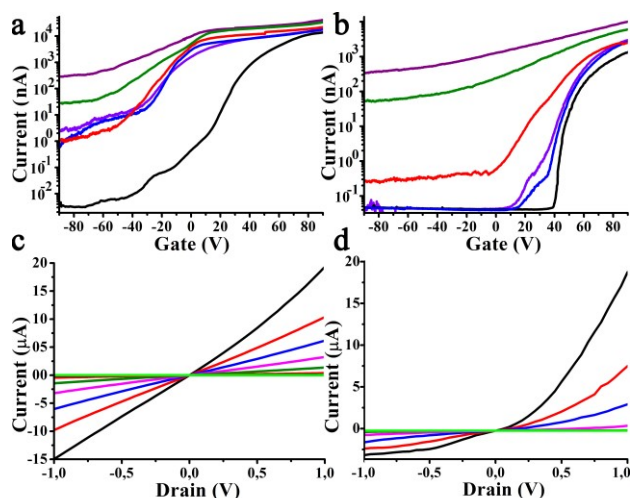


Figure 8. Transfer curve of (a) pristine MoS₂ and (b) Per-MoS₂ at dark and at 254 (violet), 365 (blue), 407 (purple), 532 (green), and 650 nm (red) light excitation. Output characteristics of (c) pristine MoS₂ and (d) Per-MoS₂ at different gate voltage, 80 V (black), 70 V (red), 60 V (blue), 50 V (pink), 40 V (olive), 30 V (red), 20 V or less (green). The V_{D-S} was 1 V. The power intensity for 254 nm, 365 nm, 407 nm, 532 nm and 650 nm were 10 mW/cm^2 , 10 mW/cm^2 , 30 mW/cm^2 , 17 mW/cm^2 and 7 mW/cm^2 .



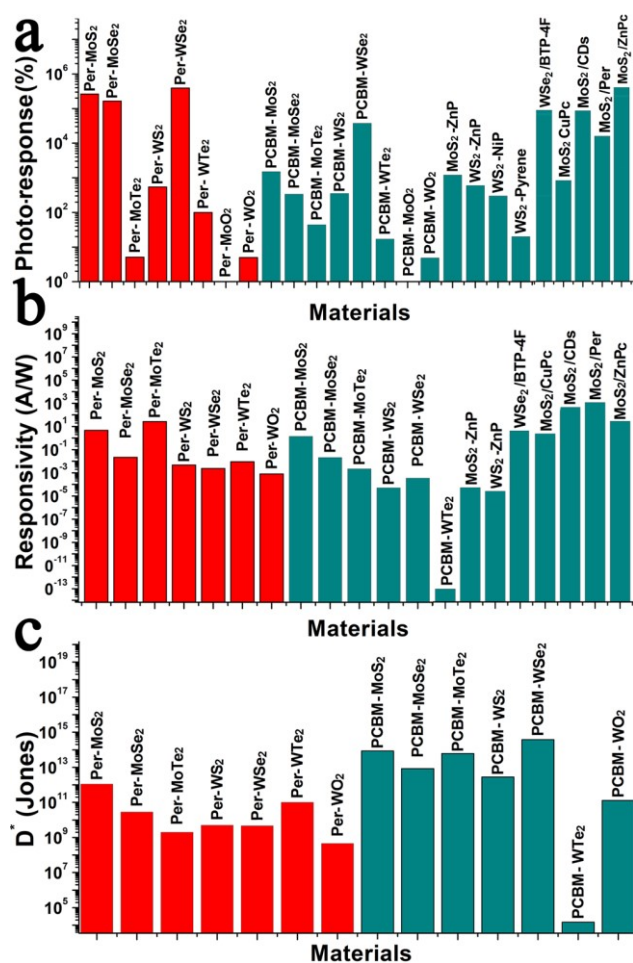


Figure 9. Comparison of (a) photo-response, (b) responsivity, and (c) detectivity of Per-TMDs with PCBM-TMDs and other relevant studies.^{34,36,54-56,61,99,100,101}

the highest responsivity in this contribution is achieved with Per-MoTe₂, while Per-WSe₂ exhibits the most significant photo-response, with values of 26 A/W and 3.9×10⁵%, respectively. However, the study of covalent functionalization with organic chromophores in this context remains largely unexplored. As mentioned previously, covalent functionalization facilitates thorough material cleaning and isolates individual molecules. Additionally, it prevents electron diffusion in organic-organic systems caused by aggregation, thereby enhancing performance. However, only a few contributions have targeted devices using this approach.

On the other hand, the performance of the devices depends on a wide range of factors, including the substrate,⁵ the synthesis of the TMDs and the functionalization techniques,¹⁸ challenging an accurate comparison. Said that, we previously functionalized TMDs with a PCBM derivative,⁵⁴ by employing not only the same conditions but even a piece of same CVD pristine TMDs substrate, which was divided into different fractions. Both PCBM and perylene were functionalized with cysteine, which enables covalent interaction with TMDs via thiol groups. The reactions with TMDs were conducted in parallel under identical conditions, resulting in sibling materials that allow for the best possible comparison. In fact, the best photo-response, responsivity, and detectivity performance of Per-TMDs

were compared with PCBM-TMDs in Figure 9. Although detectivity decreases when comparing each TMDs with its homologue, the responsivity remains comparable, and the photo-response is superior in Per-TMDs.

PCBM exhibited a similar alignment to perylene across most materials, except for MoS₂, where PCBM forms type I alignment in contrast with perylene, that forms a type II alignment. As a result, the conductivity of PCBM-MoS₂ increases, but its photo-response is two orders of magnitude smaller than that of Per-MoS₂. In contrast, Per-MoTe₂ exhibits a 50-fold increase in both responsivity and conductivity compared to PCBM-MoTe₂, at 532 nm. For other TMDs hybrid materials, perylene and PCBM show similar behavior, with notable enhancements of 10-fold for Per-MoSe₂ and 5-fold for Per-WSe₂ at 532 nm. However, the most significant difference lies in the photo-response outside this wavelength range. Perylene exhibits strong absorption between 450 and 600 nm, tailoring the photo-response primarily within this range, which is a desirable property for selective photo-devices. In contrast, PCBM displays a broader absorption spectrum, leading to an overall enhancement of photo-response across the visible spectrum, making it more suitable for non-selective photo-devices.

Experiments performed under covalent functionalization of ZnP with MoS₂ (ZnP-MoS₂) and WS₂ (ZnP-WS₂) were conducted under similar conditions.⁵⁴ However, a key difference was that the MoS₂ and WS₂ nanosheets were directly transferred onto a gold electrode, forming a Schottky barrier that hindered optimal performance. Focusing solely on the photo-response, ZnP-MoS₂ and ZnP-WS₂ exhibited values of 1200% and 600% under 405 nm laser irradiation. In contrast, their homologues, Per-MoS₂ and Per-WS₂, reached values of 96300% and 442%, respectively, with orbital alignment being the primary factor responsible for the differences observed between MoS₂ and WS₂. Specifically, ZnP acts as an electron donor, whereas PCBM and perylenes function as electron acceptors. WS₂ covalently functionalized with pyrene (WS₂-pyrene) generally shows worse results than our materials.⁵⁶ However, a direct comparison is not applicable, as those experiments were conducted on nanofilms rather than on a single WS₂ layer.

For instance, considering non-covalent functionalization with perylene, not under its best performance but under conditions comparable to our experiments—namely, power densities of 12.5 or 25 mW/cm², 5s pulses of 532 nm laser illumination, V_{DS} = 3 V, and V_G = 0 V—the reported photo-response is approximately 10-fold.⁹⁹ In contrast, our results show an enhancement of 2630-fold, at 17 mW/cm² and V_{DS} = 1 V. A plausible explanation is that, in our system, perylene-perylenes interactions are minimized, suppressing competitive charge transfer pathways that would otherwise reduce carrier mobility. Additionally, the responsivity of non-covalent MoS₂/perylenes falls below 1 A/W. Similarly, the responsivity of MoS₂/ZnPc,³⁴ drops to 8 A/W under 532 nm illumination at 1.0 V and 3 mW/cm² when V_G = 0 V. These values are comparable to those obtained for Per-MoS₂ and Per-MoTe₂, which exhibit responsivities of 4.6 and 26.2 A/W, respectively, at 17 mW/cm².

Without any additional in-depth analysis, Per-MoS₂, Per-MoSe₂, and Per-WSe₂ exhibit a better photo-response than MoS₂ functionalized



with Cu-phthalocyanine (MoS_2/CuPc)³⁶ or with CDs (MoS_2/CDs),¹⁰⁰ or WSe_2 functionalized with BTP-4F ($\text{WSe}_2/\text{BTP-4F}$).¹⁰¹ On the other hand, the responsivity of MoS_2/CDs ¹⁰⁰ is higher but still close to Per- MoS_2 and Per- Te_2 . Moreover, Per- MoS_2 and Per- Te_2 show better performance than MoS_2/CuPc ³⁶ or $\text{WSe}_2/\text{BTP-4F}$ ¹⁰¹ (see Figure 9). It must be noted, however, that the photo-response and responsivity can be tuned by applying a gate voltage, potentially increasing the values. Nevertheless, this aspect was beyond the primary focus of this contribution, which aims to fully understand the electronic mechanisms involved between TMDs and isolated planar electron-acceptor molecules such as perylene.

Therefore, although a precise analysis is not feasible due to variations in experimental conditions, a comparison with literature reports under similar conditions reinforces the exceptional performance of our system.

Experimental

Synthesis:

CVD synthesis: This procedure has been previously published.⁵⁴ As example, MoS_2 (or WS_2) 15 mg of MoO_3 (or WS_2) containing 15% of NaCl was deposited at the centre of an alumina boat ($10 \times 1.5 \text{ cm}^2$) with a height of 1.1 cm. The boats were put in a quartz tube with 3 cm in diameter. Argon gas was flowed with a flow rate of 113 ml/min, and MoO_3 was gradually heated up to 725 °C (780 °C) in 13 minutes. And then, Sulphur was also heated up to 210 °C in 10 minutes. All CVD parameters were kept constant for 30 additional minutes before a slow cooling down. All conditions can be described on Table S1.

Synthesis of Perylene (N,N'-Bis-(L-Cysteine)-3,4,9,10-perylene tetracarboxylic diimide): The synthesis followed a procedure previously described in the literature.^{64,65} Briefly, 3,4,9,10-perylene tetracarboxylic acid dianhydride (PTCDA) (1 g, 1 equiv.), L-cysteine (0.7 g, 5 equiv.), and 100 mL of DMF were added to a Schlenk flask and heated at 120 °C overnight. After cooling, the reaction mixture was subjected to vacuum to remove the solvent. The resulting crude solid was powdered and washed multiple times with hot water, followed by methanol and acetone. To reduce the disulfide bonds formed during the reaction, the product was acidified with trifluoroacetic acid, and excess zinc powder was added over several hours. The excess zinc was removed by filtration through filter paper, and the pH was neutralized with sodium carbonate, causing perylene to precipitate. Finally, the material was washed with hot water to remove excess cysteine. Yields: 50% NMR: (8H, perylene) 8.5-7.5 ppm, multiplet; (4H, Cys α , thiol) 6.0-5.0 ppm, multiplet; (4H, Cys β) 3.25-2.75 ppm, multiplet, partially masked by solvent peaks. IR: COOH (Cys) stretch 1700-1800 cm^{-1} ; O=C-N-C=O (perylene) 1600-1700 cm^{-1} . MALDI: ($\text{C}_{30}\text{H}_{16}\text{N}_2\text{NaO}_8\text{S}_2$) calculated: 619.02, 620.03, 621.02, 621.03; found: 619.346, 620.360, 621.360.

TMDs-Per: MoS_2 (or WS_2) or other TMDs on SiO_2/Si was immersed in 5ml of DMAc with 1 mg of perylene-Cys. Pellets of Zn was added to prevent cysteine-cysteine dimerization of perylene, and the system was protected against light, oxygen or moisture. Reactions were kept under minimum stirring during 4 days in vacuo at 50 °C (40 °C) or room temperature and dark. Then, it was cleaned with pure DMAc

four times by immersion. Lately, non-covalent spices were completely removed employing DMAc under Soxhlet reflux during 12 h.

Please refer to Supplementary Information for instrumentation, methods and techniques.

Conclusions

In this contribution, perylene was successfully insolated on the surface of various single-layer TMDs through covalent functionalization. The materials underwent detailed characterization, followed by an in-depth analysis of their optical and electronic properties. Perylene forms a type II alignment with all semiconducting TMDs, except for MoTe_2 , where a type I alignment is observed. However, trap states play a crucial role, causing all the hybrid materials to behave as if they exhibit a pseudo-type I alignment. The intrinsic PL emission of all the TMDs undergoes a significant reduction. In contrast, the light absorption is greatly enhanced in the perylene region (400-600 nm) in hybrid materials. Interestingly, the photo-response emerges immediately after functionalization in tungsten-based metallic materials, such as Per- WO_2 , and improves in semiconducting materials, with a 1000-fold increase observed in Per- MoS_2 and Per- WSe_2 reaching values of $3.9 \times 10^5\%$ and $2.6 \times 10^5\%$, respectively. The highest carrier mobility was achieved by Per- MoS_2 , with a value of 92.3 $\text{cm}^2/\text{V}\cdot\text{s}$, while the best responsivity was observed in Per- MoTe_2 , reaching 26.2 A/W. The photo-response, responsivity, and specific detectivity values achieved through this study are comparable to the highest-performing systems involving non-covalent functionalization. This approach is both straightforward and effective, allowing for the precise integration of molecules onto TMDs surfaces while tailoring their properties. We demonstrate that perylene can uniquely modify each TMDs, and we discuss its potential target applications. Furthermore, we compare our results with those reported in the literature, confirming the outstanding performance of the Per-TMDs devices. As a result, this method holds significant potential for advancing applications in nanotechnology and energy conversion.

Conflicts of interest

"There are no conflicts to declare".

Data availability

The data supporting this article have been included as part of the Supplementary Information.

Acknowledgements

This work was supported by the Project on Design & Engineering by Joint Inverse Innovation for Materials Architecture (DEJ12MA) from the Ministry of Education, Culture, Sports,



Science and Technology (MEXT) and from Japan Society for the Promotion of Science (JSPS) postdoctoral fellowship grant agreement No P19368.

Notes and references

† Electrochemistry may be considered controversial for evaluating the positions of energy gaps in materials. The primary reason for this is that during oxidative or reductive processes, the material could undergo chemical changes that modify its electronic states. However, it is well known that perylenes and TMDs are chemically stable under oxidative or reductive potentials, suggesting that the current methodology is applicable. Cyclic voltammetry, therefore, might not be suitable for evaluating novel materials. However, since our system involves the interaction of two well-known and chemically stable materials (TMDs and perylenes), simple math calculations should suffice to verify whether the oxidative or reductive positions align with the values reported in the literature. This approach acknowledges the limitations of the system while providing reliable results. Additionally, the HOMO and LUMO gaps of perylene align well with its optical gap, while the conduction and valence bands of the TMDs correspond closely to the gaps calculated using the Varshni Eq. S1.

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Data availability statements

- The data supporting this article have been included as part of the Supplementary Information.

