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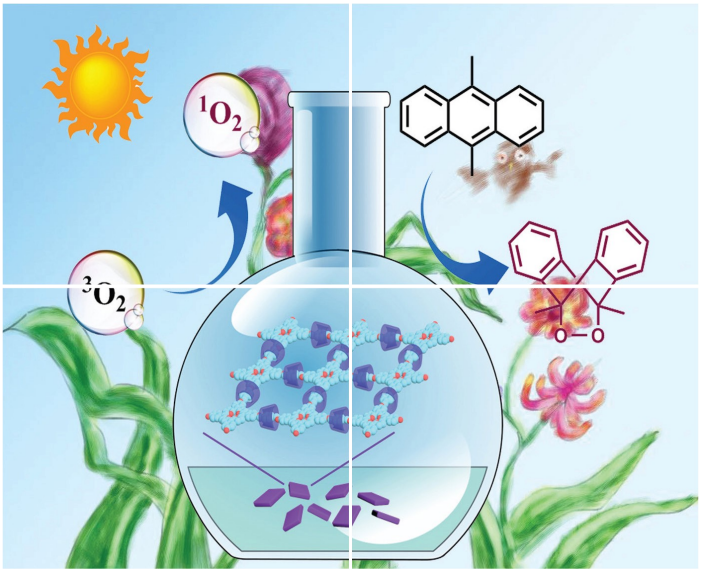

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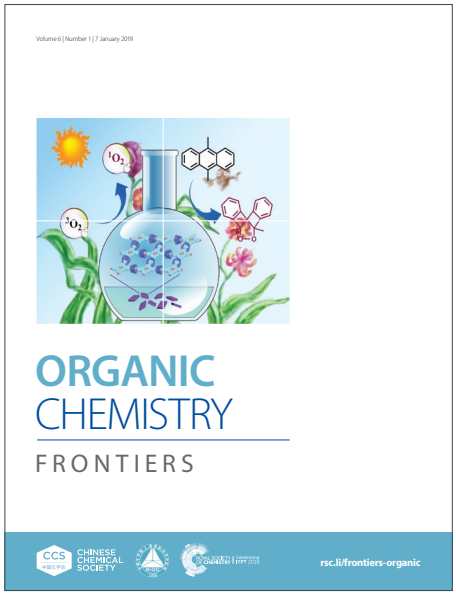

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Mechanism of Formation of Chiral Allyl SCF_3 Compounds via Selenium-Catalyzed Sulfenofunctionalization of Allylboronic Acids

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Abstract

The detailed reaction mechanism of diphenyl selenide-catalyzed sulfenofunctionalization of chiral α-CF_3 allylboronic acids is investigated by means of density functional theory calculations. It is demonstrated that the reaction starts with the transfer of the SCF_3 group from the $(PhSO_2)_2NSCF_3$ reagent to the Ph_2Se catalyst, a process that is shown to be assisted by the presence of Tf_2NH acid. After a proton transfer step, the SCF_3 group is transferred to the C=C double bond of the substrate to generate a thiiranium ion. Concerted deborylative opening of the thiiranium ion yields then the final product. Several representative substrates are considered by the calculations, and the origins of the stereoselectivity of the reactions are analyzed by comparing the energies and geometries of the transition states leading to the different products.

Keywords: Selenide, reaction mechanism, stereoselectivity, DFT calculations, chiral allylboronic acid

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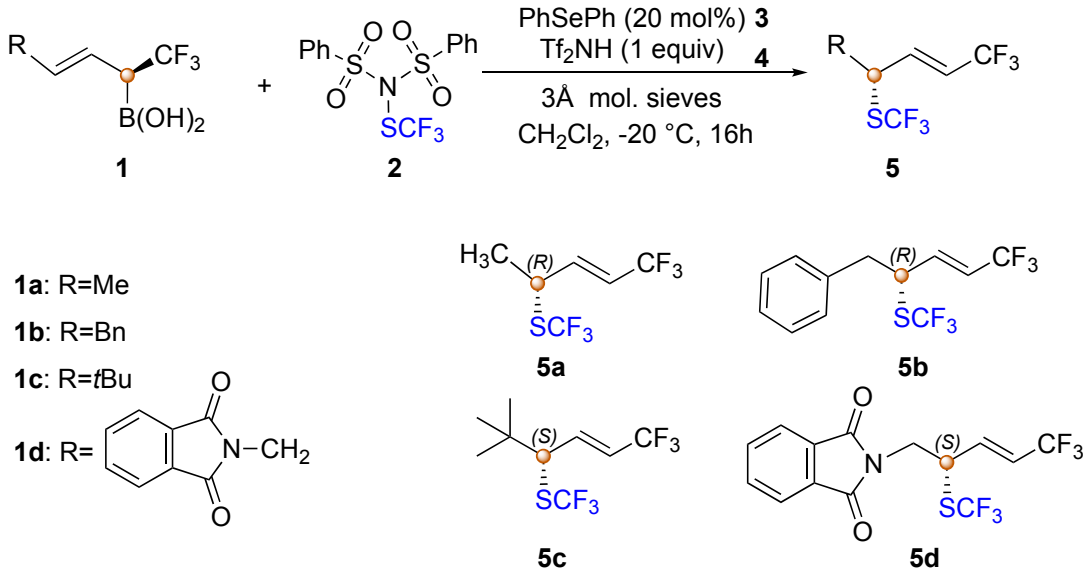
1. Introduction

Fluorine-containing pharmacophores are often applied in modern drug substances.[1] One of five commercial drugs contains at least one C-F bond.[2] The main reason for the widespread application of fluorine-containing groups in bioactive small molecules is their beneficial metabolic and pharmacokinetical properties.[3] A particularly important pharmacophore is the trifluoromethylthio (SCF_3) group, which substantially increases the lipophilicity of the drug substances and also has a pronounced electron-withdrawing character.[4-6] As a consequence, several excellent methods appeared for the introduction of the trifluoromethylthio group to organic small molecules.[7-21] In particular, the synthesis of chirally enriched SCF_3 compounds became an important but challenging area in modern organic chemistry.[22-35]

Very recently, Szabó and co-workers presented an efficient method to introduce the SCF_3 group by using $(PhSO_2)_2NSCF_3$ reagent[36] **2** with α-CF_3 allylboronic acid derivatives **1** to form a chiral alkenyl SCF_3 compound **5** (Scheme 1).[37] The reaction relies on the application of selenium-based Lewis base Ph_2Se **3** as a catalyst in the presence of Tf_2NH **4** as the activator. These reactions have a high degree of functional group tolerance and proceed with excellent stereo-, diastereo- and site-selectivity.



Scheme 1. *Se-catalyzed sulfenofunctionalization of allylboronic acid to form chiral allyl SCF_3 compound investigated in the present study.*

The application of electrophilic sulfenofunctionalization in the presence of selenium catalysis has been documented by the pioneering studies of the groups of Denmark,[38] Zhao,[26] and others.[39] However, the application of allylboronate substrates for trifluoromethylthiolation revealed a couple of new, interesting mechanistic aspects.[37] An important feature is the excellent stereochemistry of the reaction, including chirality transfer and high *E*-selectivity for the formation of the new double bond. The studies also pointed out the benefits of using allylboronic acid type of substrates. Replacement of the $B(OH)_2$ group of **1** with Bpin leads to a significant decrease in the yield of the product, indicating that allyl-Bpin species have a substantially lower reactivity than allylboronic acids.

In the present work, we have performed density functional theory (DFT) calculations to elucidate the mechanism for the formation of allyl SCF_3 compounds via the selenium-catalyzed sulfenofunctionalization of allylboronic acids. We consider the reactions of several representative substrates (**1a-1d** in Scheme 1) and discuss the origins of selectivity for each of them.

2. Computational Details

The B3LYP-D3(BJ) functional,[40, 41] which includes the D3 dispersion correction with the Becke-Johnson damping function,[42, 43] was used for all calculations presented in this work, and the Gaussian 16 program[44] was employed. Geometry optimizations were carried out with the 6-31G(d,p) basis set. To obtain more accurate energies, single-point calculations were performed on the optimized structures using the larger basis set 6-311+G(2d,2p). Analytical frequency calculations were performed at the level of theory of the geometry optimization to calculate the Gibbs free energy corrections. Solvation effects were considered by performing single-point calculations on the optimized structures using the SMD method[45] with the parameters of dichloromethane.

To evaluate the effect of performing the geometry optimizations with the smaller basis set and in the gas phase, the geometries of the first step of the reaction (**2**+**3**+**4** → **Int1**, see below) were re-optimized twice, in implicit DCM solvent and with the larger basis set. The

calculations showed that the geometries and energies were not affected significantly (see SI).

Intrinsic reaction coordinate (IRC) calculations were performed on all transition states to confirm the nature of the connecting intermediates. A thorough manual conformation search was carried out on all stationary points to ensure that the structures with the lowest energy were located.

3. Results and Discussion

We start the mechanistic investigation by considering the reaction of model substrate **1a**, in which the R group is a simple methyl substituent. Although this compound was not included in the experimental studies,[37] it can serve as a good reference for γ-alkyl allylboronates. As representatives for different classes of substrates, we subsequently investigate the mechanisms of three other substrates that have been examined explicitly by the experiments,[37] namely those in which the R is a relatively bulky benzyl (**1b**) and tert-butyl (**1c**) groups, as well as a nitrogen-containing phthalimide derivative **1d** (Scheme 1).

3.1 Reaction of model substrate 1a

The first step of the reaction mechanism is the transfer of the SCF_3 group from the $(PhSO_2)_2NSCF_3$ transfer reagent **2** to the Ph_2Se catalyst **3**, generating the $SePh_2SCF_3$ cation **7** (Scheme 2a). We found that this step is assisted by the participation of Tf_2NH acid **4** that forms a hydrogen bond to the nitrogen of the transfer reagent (**TS1**, Figure 1). The calculated barrier is 24.2 kcal/mol, and the energy of the resulting intermediate **Int1**, in which acid **4**, the $(PhSO_2)_2N^-$ anion **6** and the $SePh_2SCF_3$ cation **7** are in complex with each other, lies at 21.3 kcal/mol (Figure 2). The hydrogen bond provided by acid **4** stabilizes the negative charge developing on the nitrogen atom, as seen from the H···N distance, which is 1.98 Å at **TS1** and 1.86 Å at **Int1**. Transfer of the SCF_3 group without the participation of the acid was also considered and was found to have a much higher energy barrier of 50.4 kcal/mol (see **SI**).

Next step is a proton transfer from **4** to $(PhSO_2)_2N^-$ anion **6** via **TS2**, generating $(PhSO_2)_2NH$ **8** and Tf_2N^- anion **9** (Scheme 2b). The energy barrier of this step is only 0.6

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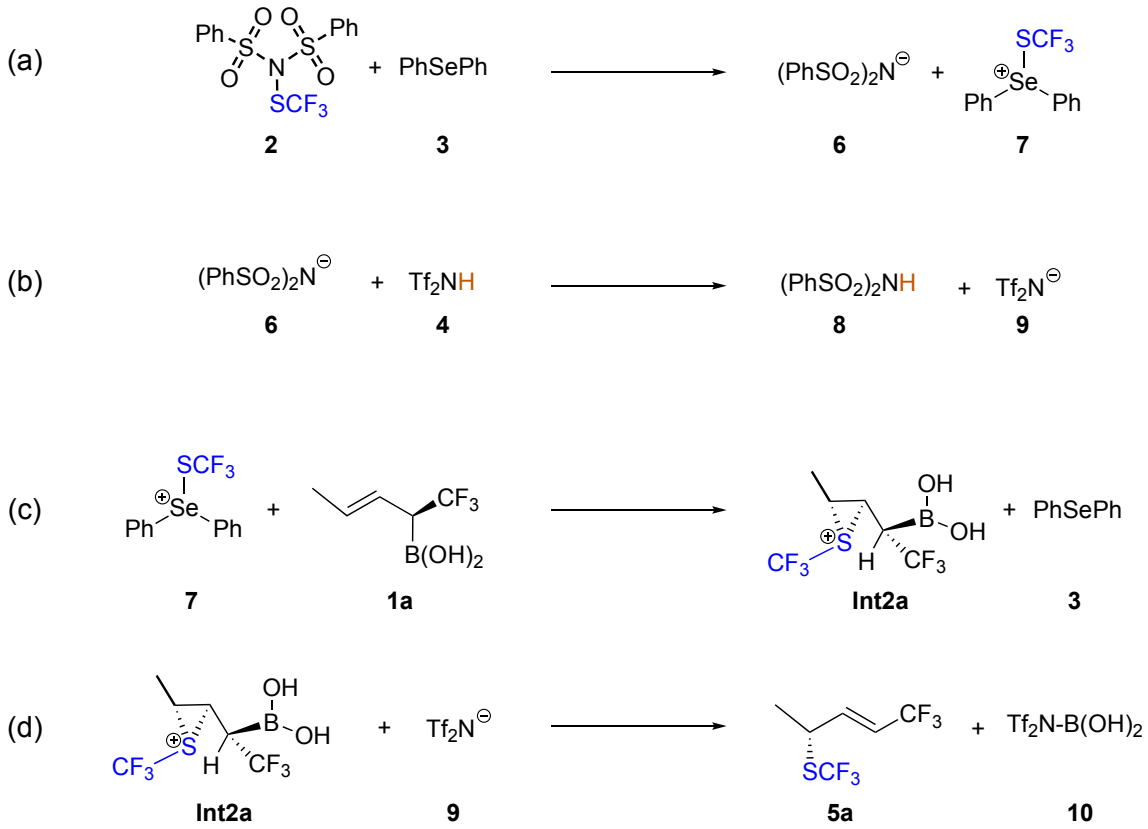
kcal/mol relative to **Int1**.

We also calculated the alternative pathway with the reversed step order, i.e. in which a proton is first transferred from acid **4** to $(PhSO_2)_2NSCF_3$, followed by the transfer of the SCF_3 group from the generated $(PhSO_2)_2NHSCF_3$ cation to the Ph_2Se catalyst **3**. However, the energy barrier of this scenario was found to be very high, 52.9 kcal/mol (see **SI**).



Scheme 2. *Reaction steps investigated in the present study.*

In the following step of the mechanism, the SCF_3 group of the $SePh_2SCF_3$ cation **7** is transferred to the C=C double bond of the substrate to generate a thiiranium ion (Scheme 2c). Two competing stereoselective pathways are possible and were investigated, in which the SCF_3 group is transferred either to the *Re*-face of the substrate via **TS3a(3*R*)** to form the (*R*)-configuration **Int2a(3*R*)**, or to the *Si*-face via **TS3a(3*S*)** to generate the (*S*)-configuration **Int2a(3*S*)**. The optimized structures of the transition states and intermediates are shown in Figure 3. The energies of **TS3a(3*R*)** and **TS3a(3*S*)** are calculated to be 18.4 and 18.6 kcal/mol, respectively, relative to the previous intermediate, and the resulting **Int2a(3*R*)** and **Int2a(3*S*)** are 10.6 and 11.2 kcal/mol higher, respectively.

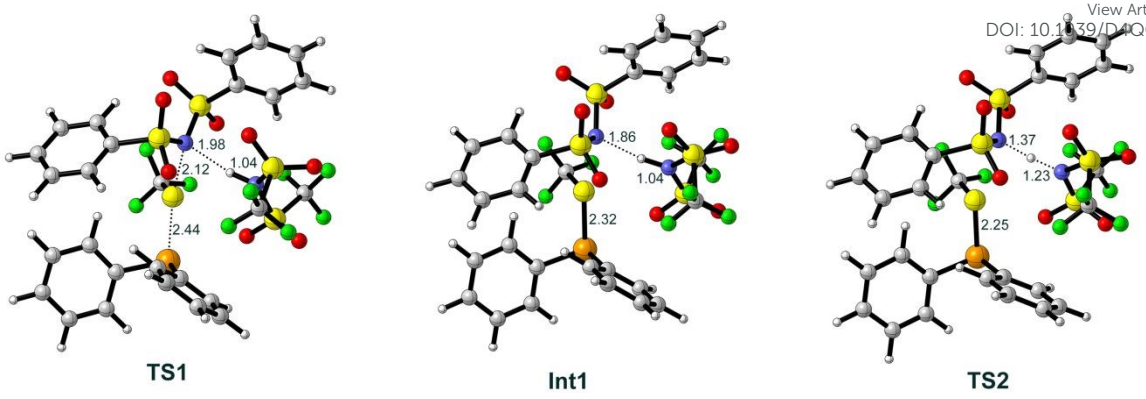


***Figure 1**. Optimized structures of the transition states and intermediates of **TS1**, **Int1**, and **TS2**. Selected bond distances are indicated in Å.*

For comparison, we also investigated the uncatalyzed reaction, in which the SCF_3 group is transferred directly from reagent **2** to substrate **1a,** generating **Int2a** without the participation of catalyst **3** (see **SI**). The energy barrier was calculated to be 29.9 kcal/mol for both the 3*R*- and 3*S*- pathways, which is significantly higher than the case with catalyst **3**. Accordingly, using PhSePh catalyst (**3**) increases the reactivity of **2**. In the positively charged **7** the electrophilicity of SCF_3 is substantially increased compared to **2**. In addition, the cleavage of the weak Se-S bond in **7** is also easier than the cleavage of the N-S bond of **2**. The stability of **7** is poor under ambient conditions, and therefore **2** is converted to **7** *in situ* (in the presence of **4**) under the applied reaction conditions.

The final step of the mechanism (Scheme 2d) involves the Tf_2N^- anion **9** performing a nucleophilic attack on the boronate group of **Int2a(3*R*)** or **Int2a(3*S*)**, triggering the concerted deborylative opening of the thiiranium ion via **TS4** to yield the four possible forms of the final product. The (*3R*)-configured products are formed from **Int2a(3*R*)** and can result in either the *E*-configuration through the *anti*-elimination pathway via **TS4$_{anti}$(3*R*)**, or alternatively, rotation of the C_α-C_β bond leads to the *syn*-elimination pathway via **TS4$_{syn}$(3*R*)**, resulting in the *Z*-configuration. Similarly, the *E-(3S)*- or *Z*-(*3S*)-configured products can be achieved via **TS4$_{syn}$(3*S*)** or **TS4$_{anti}$(3*S*)**, respectively. The optimized structures of these transition states are shown in Figure 4, while the optimized structures of the products are given in the **SI**.

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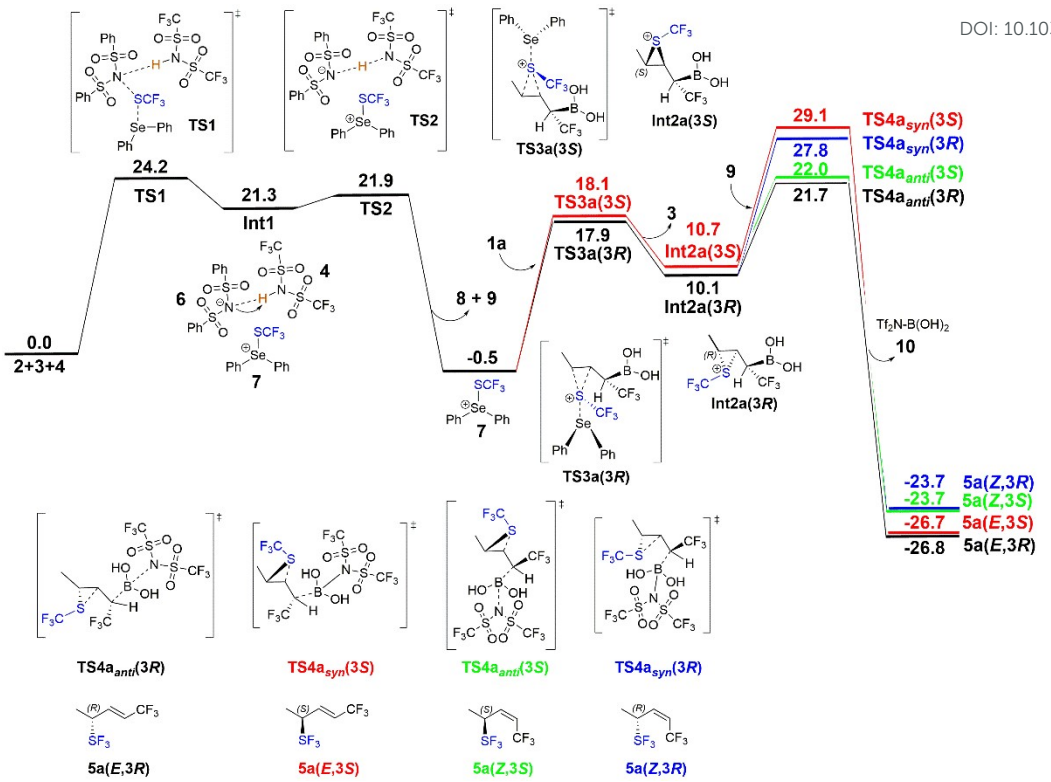

Figure 2. *Calculated free energy profile (kcal/mol) of model substrate* ***1a****.*

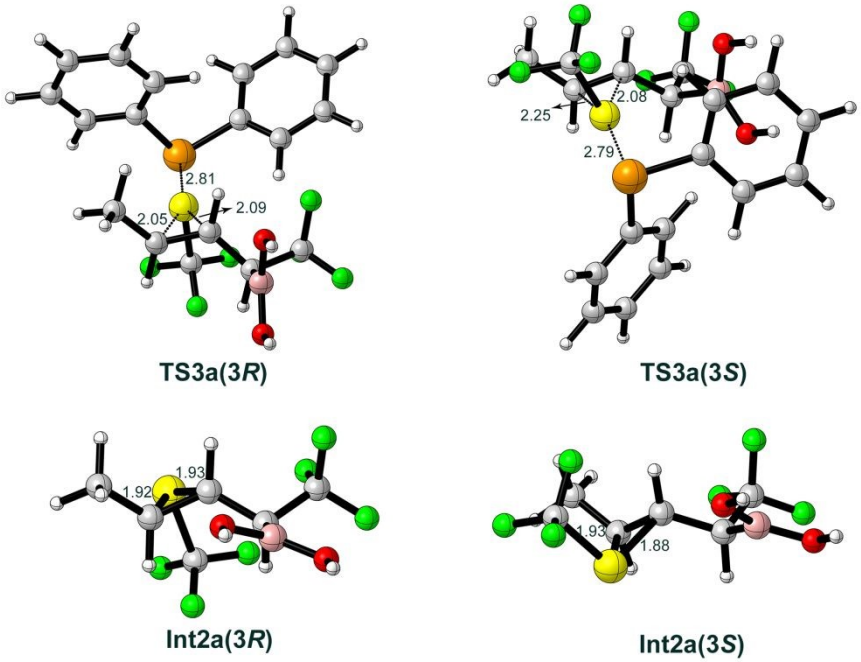

Figure 3*. Optimized structures of the transition states and intermediates for the generation of the thiiranium ion* ***Int2a****. Selected bond distances are indicated in Å.*

Although the chiral center is formed at **TS3**, the calculations show that the energy barriers for the eliminations of the boronate group via **TS4** are higher and irreversible, which means that the latter step is the stereoselectivity-determining step of the reaction.

By comparing the energy profiles of the above four pathways (Figure 2), the calculations of the model substrate **1a** show that the energy barriers of *anti*-eliminations are considerably lower than the *syn*-eliminations, 21.7 kcal/mol in **TS4**$_{anti}$**(3*R*)** *vs.* 27.8 kcal/mol in **TS4**$_{syn}$**(3*R*)**, and 22.0 kcal/mol in **TS4**$_{anti}$**(3*S*)** *vs.* 29.1 kcal/mol in **TS4**$_{syn}$**(3*S*)**. Inspection of the optimized structures in Figure 4 shows that the reason for this energy difference is mainly the steric repulsion between the SCF_3 group and the leaving boronate group, as these two moieties point toward each other in the *syn*-elimination, **TS4**$_{syn}$**(3*R*)** and **TS4**$_{syn}$**(3*S*)**, while in the *anti*-elimination, **TS4**$_{anti}$**(3*R*)** and **TS4**$_{anti}$**(3*S*)**, they point away from each other.

The difference in energy between **TS4**$_{anti}$**(3*R*)** and **TS4**$_{anti}$**(3*S*)**, which lead to the ***E*-(3*R*)-5a** and ***Z*-(3*S*)-5a** products, respectively, is calculated to be only 0.3 kcal/mol in favor of the former. The calculations show thus that already the model substrate **1a**, with the small methyl substituent, qualitatively reproduces the experimental selectivity trend, in that the formation of the *E*-(3*R*)-configured product is associated with the lowest-energy pathway, albeit with a small energy. However, as will be shown below, the calculations using the experimentally employed substrates, with bulkier substituents, yield a more quantitative agreement with the experiments.

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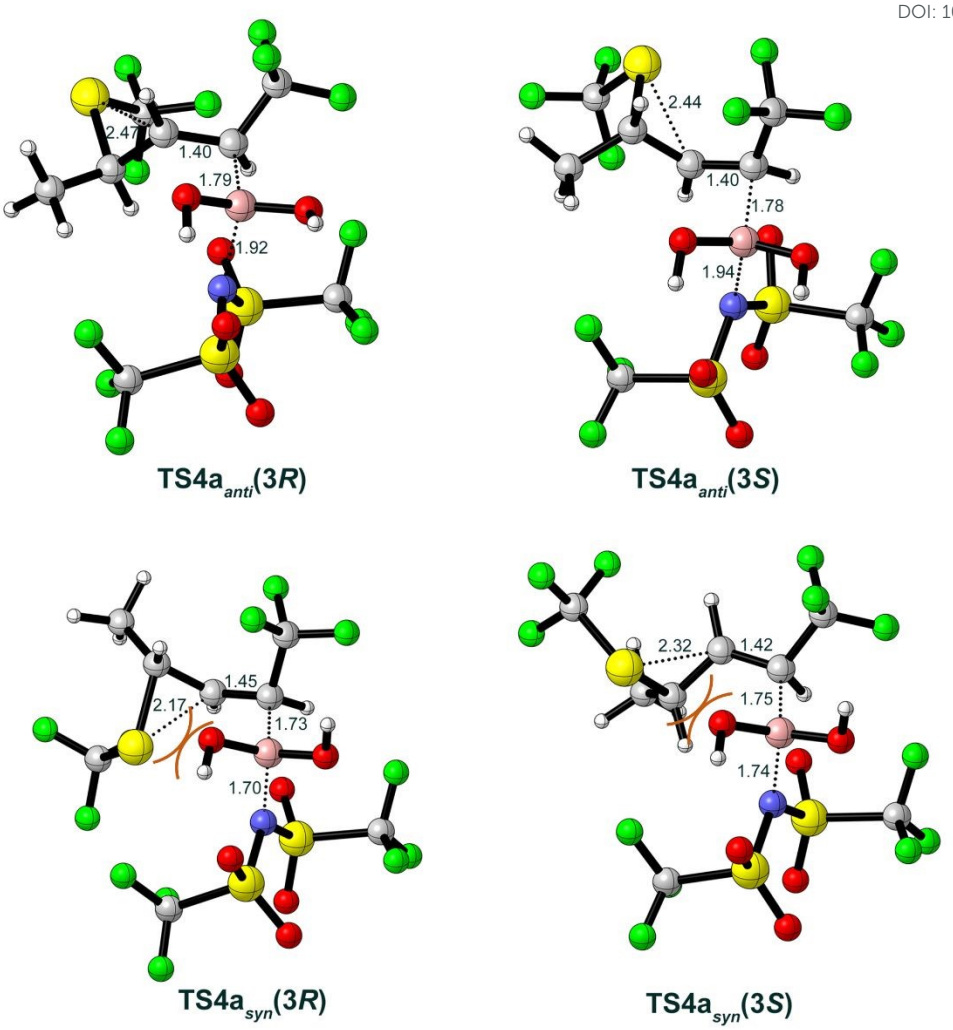

Figure 4. *Optimized structures of the transition states for the step of deborylative elimination in model substrate* ***1a****. Selected bond distances are indicated in Å.*

To summarize, the mechanism proposed on the basis of the current calculations is given in Scheme 3. The obtained overall energy profile (Figure 2) indicates that the first step, i.e. the acid-assisted transfer of the SCF_3 group via **TS1**, is the rate-determining step (RDS) for the model substrate, with a barrier of 24.2 kcal/mol. However, the final step, i.e. deborylative opening of the thiiranium ring via **TS4**, has an overall barrier of 22.2 kcal/mol, which is quite close in energy, and it is therefore not possible to determine confidently the nature of the RDS based only on the calculations. In particular, various substituents on the substrate may lead to significant changes in the energy of the final step (see below).

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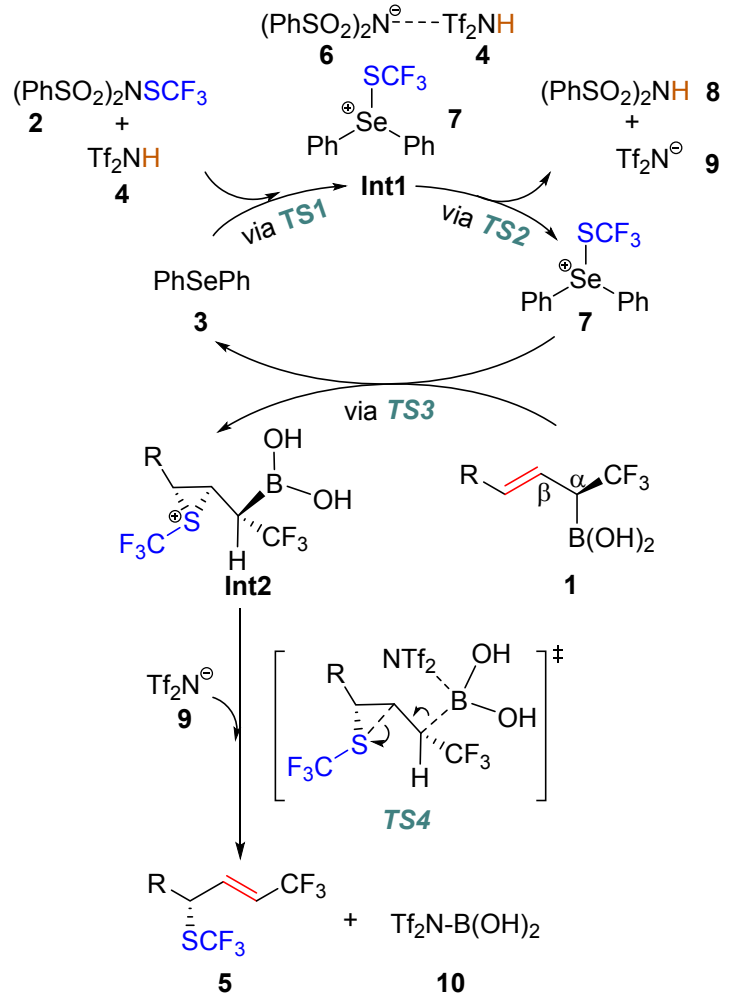

Scheme 3. *Reaction mechanism suggested on the basis of the current calculations.*

Here, it is interesting to mention two previous mechanistic studies on sulfenofunctionalizations of alkenes catalyzed by selenides, where DFT calculations were employed to investigate various aspects of the reactions. However, none of these studies involved a deborylation step, which is a novel aspect of the present study. Denmark and co-workers analyzed the geometries and energies of the transition states for the thiiranium ion formation step, which was assumed to be the enantio-determining step of the reaction,[46] while Zhao and co-workers investigated the mechanism of selenide-catalyzed trifluoromethylthiolation of *gem*-diaryl tethered alkenes to synthesize trifluoromethylthiolated tetrahydronaphthalenes.[27]

In addition to the results discussed above, we have also considered some other mechanistic alternatives that turned out to have higher energy barriers. As seen from Figures 2 and 4, catalyst **3** does not participate in the deborylative elimination step in **TS4**. We have considered whether it can assist this step, but the energy barriers for this scenario were found to be higher (see **SI**). We also considered whether the $(PhSO_2)_2NH$ species **8** could act as the nucleophile to attack the boronate group of **Int2a**, but the calculated energy barriers for this pathway were calculated to be very high as compared to when the anionic Tf_2N^- **9** is the nucleophile (see **SI**). Finally, the experiments reveal that replacing the $B(OH)_2$ group of the substrate with Bpin significantly decreases the yield of the product.[37] Consistently with this result, the calculations show that the barrier for the case of Bpin is 3.7 kcal/mol higher than the case of the $B(OH)_2$ group (see **SI**).

3.2 Reactions with other substrates

Next, we calculated the mechanisms when the R group of the substrate is Bn (**1b**), *t*Bu (**1c**), and the phthalimide substituent (**1d**), all of which have been employed in the experimental study.[37] As seen from Figure 2, the reaction mechanism up to the formation of $SePh_2SCF_3$ cation **7** is independent of the substrate, and therefore we investigated the reactions of the other substrates starting from this point.

For substrate **1b**, with the benzyl substituent, the mechanism was calculated to be very similar to that of the model substrate **1a** shown in Scheme 3 (see calculated energy profile in Figure 5). One small difference is that the formation of the *Z*-(3*R*)-configured product through the *syn*-elimination was found to occur in a stepwise manner (see **SI** for detailed results). The calculations show that the overall barrier for substrate **1b** is ca 2 kcal/mol lower than for **1a**, and very importantly, the extent of the stereo-differentiation is well-reproduced.

Similarly to substrate **1a**, the barriers of *anti*-eliminations for **1b** are considerably lower than the *syn*-eliminations due to steric repulsion between the SCF_3 group and the leaving boronate group. In addition, the pathway leading to the ***E*-(3*R*)-5b** product is now 3.7 kcal/mol lower than that leading to the ***Z*-(3*S*)-5b** product (Figure 5), due to a steric

repulsion between the benzyl substituent of the substrate and the SCF_3 group (see **SI**). The calculated energy difference is in good agreement with the experimentally observed ee of 93% in favor of the *E*-(3*R*) product.

For substrate **1c** with the *t*Bu substituent (see **SI**), the situation is very similar to substrate **1b**, with both *syn*-eliminations found to take place in a stepwise manner. The overall barrier was calculated to be ca 3 kcal/mol higher than for **1b**, and the selectivity is determined by the same factors, with an energy difference of 4.2 kcal/mol, in good agreement with the experimental outcome of 98% ee.

Substrate **1d**, with the phthalimide substituent, represents an interesting case, because the carbonyl group of **1d** may perform a nucleophilic attack at the C_β atom of the thiiranium ion through an intramolecular mechanism,[25, 47, 48] leading to the opening of the thiiranium ion and yielding a six-membered ring intermediate **Int3d'** (see Figure 6). The calculations show that the energy barrier for this competing intramolecular nucleophilic attack via **TS4d'** is much lower than for the intermolecular reaction of the thiiranium ion with the Tf_2N^- anion **9** via **TS4d**, which was found for the other substrates. However, from **Int3d'**, the barriers for the following steps, which would be the nucleophilic attack of Tf_2N^- and the deborylative opening of the six-membered ring, were found to be higher in energy compared to the intermolecular reaction (see **SI**), indicating that the intramolecular nucleophilic attack is a reversible process. Thus, formation of **Int3d'** can be regarded as an unproductive dead-end for the deborylative trifluoromethylthiolation process.

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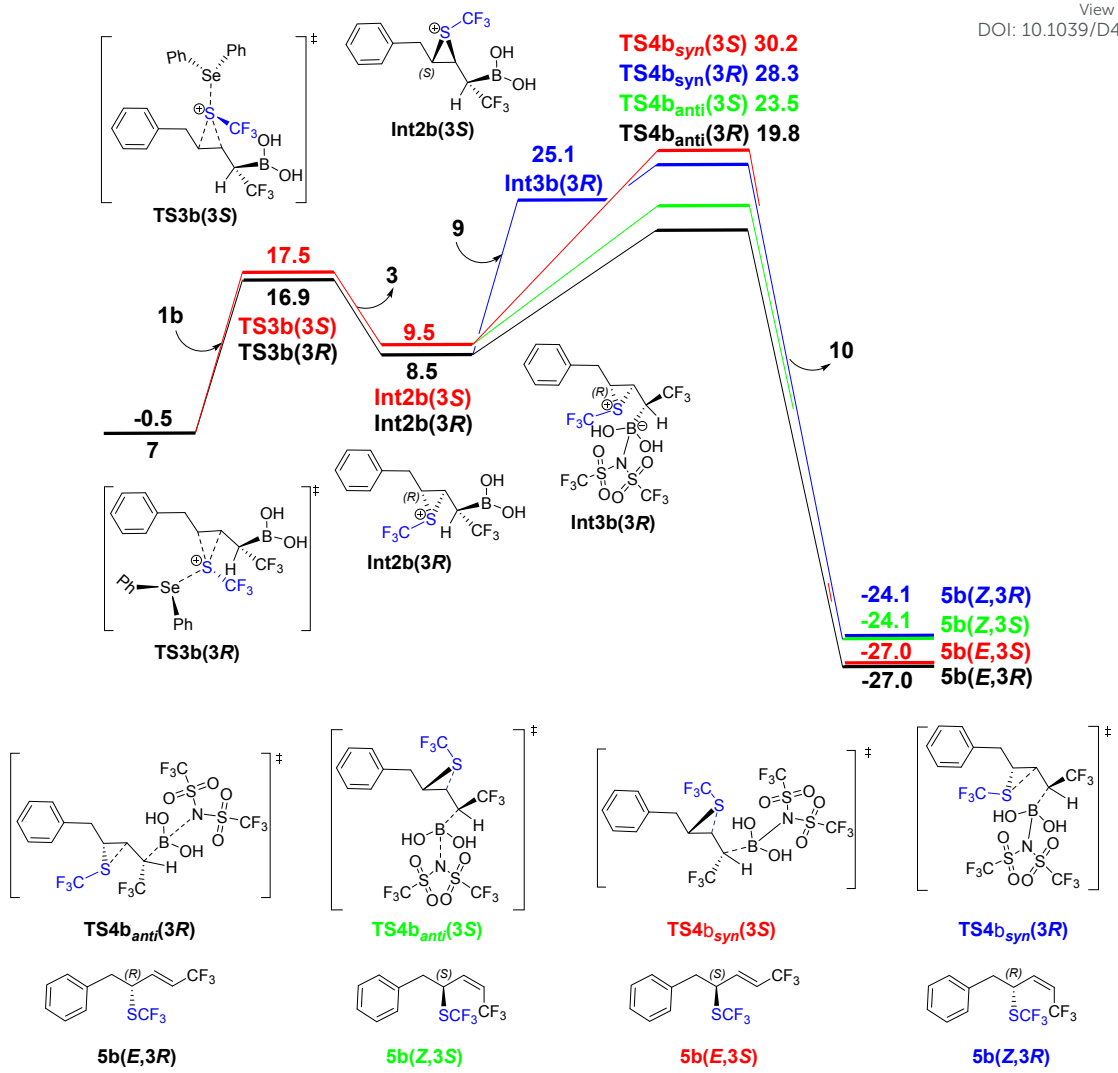

***Figure 5**. Calculated free energy profile (kcal/mol) of substrate **1b**.*

The calculations show thus that the reaction of substrate **1d** also follows the mechanism of the model substrate **1a**. However, as seen from Figure 6, the energy of **Int3d'** is the lowest point on the energy profile before **TS4d**, which means that the overall barrier should be calculated relative to **Int3d'**, resulting in a slightly higher barrier as compared to the other substrates (24.9 kcal/mol compared to 22.2, 20.3, and 23.2 kcal/mol, for **1a**, **1b** and **1c**, respectively). Importantly, the stereoselectivity is reproduced also for substrate **1d**, with a selectivity-determining energy difference of 2.5 kcal/mol, in good agreement with the 97% ee observed experimentally. The origins of the selectivity are found to be the same as for the other substrates.

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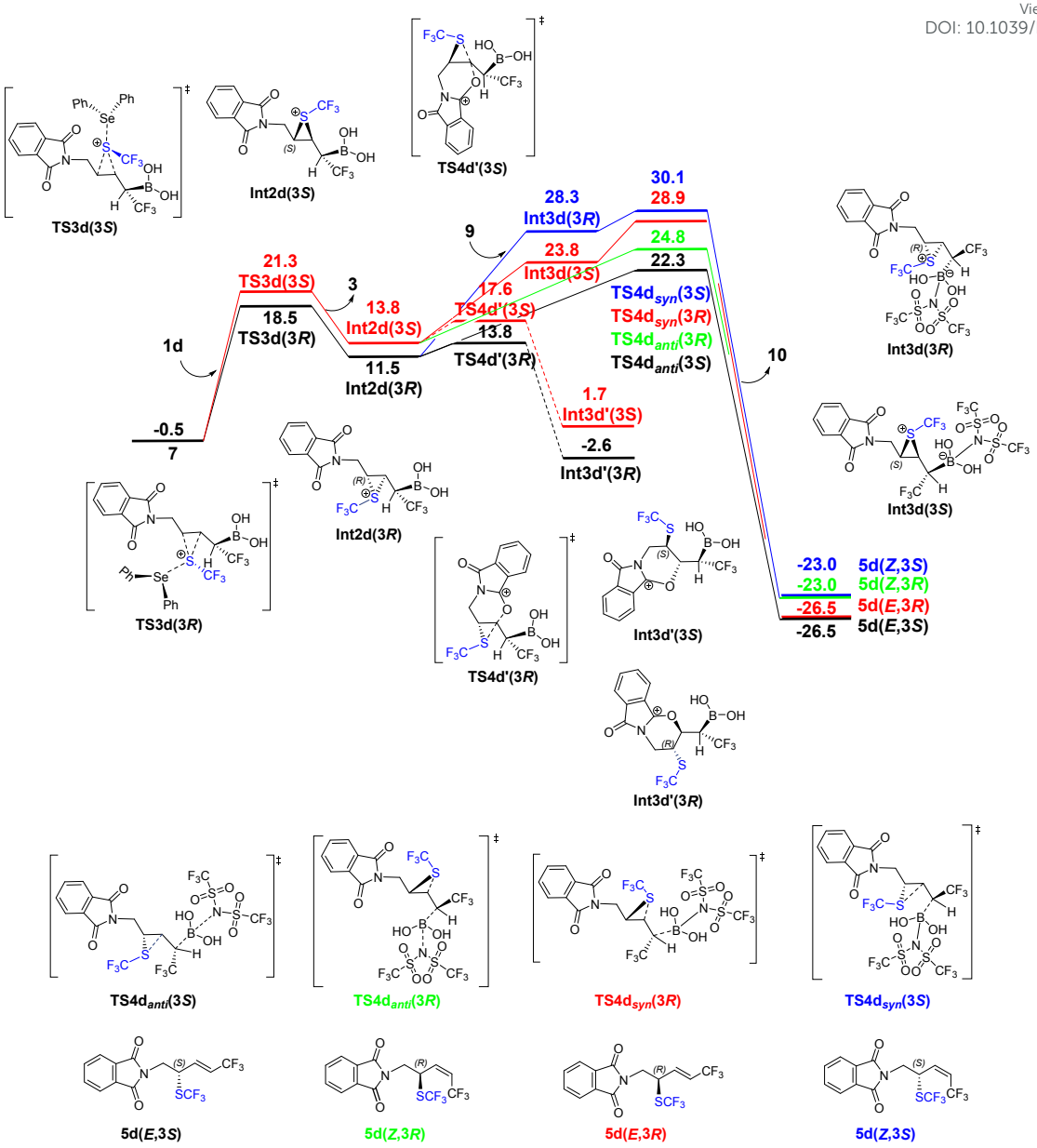

Figure 6. *Calculated free energy profile (kcal/mol) of substrate* ***1d****.*

4. Conclusions

In the present work, the reaction mechanism for the formation of chiral allyl SCF_3 compounds via diphenyl selenide-catalyzed sulfenofunctionalization of allylboronic acids has been investigated using DFT calculations. Several allylboronic acid substrates were considered, and the mechanism suggested on the basis of the calculations is shown in Scheme 2.

The reaction starts with the generation of the catalytically active $SePh_2SCF_3$ cation, a

process that takes place is in a stepwise manner, with a transfer of the SCF_3 group from the $(PhSO_2)_2NSCF_3$ reagent **2** to the Ph_2Se catalyst **3**, followed by a proton transfer from acid Tf_2NH **4** to the formed $(PhSO_2)_2N^-$ anion **6**. Interestingly, the Tf_2NH acid stabilizes the negative charge that develops on the nitrogen anion of the $(PhSO_2)_2N^-$ species **6** through a hydrogen bond interaction, lowering thus the barrier for the SCF_3 transfer.

Next, the SCF_3 group of the formed $SePh_2SCF_3$ cation is transferred to the C=C bond of the α-CF_3 allylboronic acid, generating the thiiranium ion species **Int2**. Two different pathways are possible, depending on whether the SCF_3 group is transferred to the *Si* or *Re* face of the C=C bond, which eventually lead to the *S*- or *R*-configurations of the product, respectively.

Finally, the Tf_2N^- anion **9** performs a nucleophilic attack at the boronate group of **Int2**, triggering the opening of the thiiranium ion and the leaving of the boronate group. Rotation around the C_α-C_β bond of **Int2** leads to either the *syn*- or *anti*-elimination, generating the *E*- or *Z*- configurations of the product. This step constitutes the selectivity-determining step of the reaction, and the calculations show a clear preference for formation of the *E*-form, in excellent agreement with the high *E*-selectivity reported in the experimental studies.[37] The enantioselectivity is also very well reproduced by the calculations, and analysis of the transition state structures shows that the selectivity is mainly controlled by steric repulsions between the SCF_3 group and both the leaving boronate group and the substituent of the α-CF_3 allylboronic acid.

The mechanistic insights provided by the current study will be valuable for the development of new regio-, stereo- and diastero-selective selenide-catalyzed, deborylative electrophilic sulfenofunctionalization reactions.

Conflicts of Interest

There are no conflicts to declare.

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Data Availability

The data supporting this article have been included as part of the Supplementary Information

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Supporting Information

Additional computational results discussed in the text, absolute energies and energy corrections, and Cartesian coordinates.

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Data Availability

The data supporting this article have been included as part of the Supplementary Information