

RSC Sustainability

Accepted Manuscript

This article can be cited before page numbers have been issued, to do this please use: J. Blanco, I. Agirre, I. Gandarias, J. Moreno and J. Iglesias, *RSC Sustainability*, 2025, DOI: 10.1039/D5SU00106D.



This is an Accepted Manuscript, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this Accepted Manuscript with the edited and formatted Advance Article as soon as it is available.

You can find more information about Accepted Manuscripts in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this Accepted Manuscript or any consequences arising from the use of any information it contains.

Sustainable Statement

View Article Online
DOI: 10.1039/D5SU00106D

The relocation of potentially polluting industrial activities to countries with lenient environmental legislation has been constant in the chemical industry since the late 20th century. However, demand for products from such processes remains, merely shifting pollution without global-scale changes. Furfural is one of the chemical products obtained through processes with significant environmental impact and serves as a case study in this contribution, which evaluates the technical, economic, and environmental feasibility of its production in Europe. This evaluation guides the development of a more sustainable furfural production process using heterogeneous catalysis, reducing environmental impact by eliminating sulphuric acid and improving energy efficiency. It also valorises by-products like lignin and cellulose, aligning with UN SDGs 9 and 13.



Towards greener furfural: Evaluating the technical, economic and environmental feasibility of heterogeneous catalysis in Biomass Conversion

Jorge Blanco-Cejas^a, Ion Agirre^b, Inaki Gandarias^b, Jovita Moreno^{a,c}, José Iglesias^{a,c}

^a Chemical & Environmental Engineering Group. Universidad Rey Juan Carlos. C/ Tulipan s/n., 28933, Mostoles (Spain)

^b Department of Chemical and Environmental Engineering. Bilbao School of Engineering. University of the Basque Country (UPV/EHU). Plaza Ingeniero Torres Quevedo 1. 48013 Bilbao (Spain).

^c Instituto de Tecnologías para la Sostenibilidad. Universidad Rey Juan Carlos. C/ Tulipan s/n., 28933, Mostoles (Spain)

***Corresponding Authors: Jorge Blanco-Cejas. Email: jorge.blanco@urjc.es. Ion Agirre. Email: ion.agirre@ehu.eus.**

ABSTRACT

Furfural is a key biomass-derived platform chemical with a large market volume, yet its production has largely been outsourced from Europe due to the high energy demand for reactor heating and the significant environmental impact of acidic waste generation. Current industrial processes, predominantly the Chinese Batch Process (CBP), rely on sulphuric acid as a catalyst and require extensive steam stripping, contributing significant environmental constraints. This study explores the feasibility of a more sustainable furfural production by evaluating an alternative process based on a heterogeneous acid catalyst. The proposed process integrates scale-up considerations to improve reactor performance, replacing steam stripping with nitrogen stripping and sulphuric acid with Amberlyst-70® as heterogeneous catalyst. A detailed process simulation, techno-economic analysis (TEA), and life cycle assessment (LCA) were conducted to compare the material and energy balances of both processes and to assess the viability of the heterogeneous catalytic process (HCP). Results indicate that the current selectivity of Amberlyst-70® is insufficient for technical feasibility, as a 5.5-fold improvement in furfural-to-tar selectivity is required to match the steam consumption of CBP. If this target is met, both processes exhibit similar minimum selling prices (> €3,000/t), although significantly above current market levels. However, the HCP presents a potential cost reduction pathway (< €1,000/t) through the valorisation of lignin and cellulose by-products, offering a competitive advantage. Environmental analysis highlights key benefits of the HCP, including a significant reduction in freshwater ecotoxicity by eliminating sulphuric acid and improved energy efficiency through enhanced process integration. Nonetheless, energy consumption and maize cob usage remain critical environmental hotspots. Overall, the study identifies catalyst selectivity as the key bottleneck preventing the implementation of the HCP. Further development of a more selective and stable heterogeneous catalyst, alongside integrated biorefinery strategies, could enable the competitive and sustainable production of furfural.

KEYWORDS: acid-free furfural production, heterogeneous catalysis, techno-economic analysis, life cycle assessment, biomass valorisation.



1 INTRODUCTION

Climate change and resource depletion are critical environmental threats largely driven by fossil fuel dependence.^{1,2} The industrial sector heavily relies on fossil combustion and is the second economic sector on greenhouse gas emissions with 14 Gt CO₂-eq.³ The current trend involves exploring alternative carbon sources, with biomass emerging as a key renewable option due to its abundance, low cost, and independence from international supply chains.^{4,5} By 2022, the number of biorefineries in Europe had reached 1,250, including 500 facilities focused on chemical production.^{6,7}

Among the biomass-derived chemicals with the strongest market support, platform molecules such as lactic and succinic acids stand out as the most successful examples.⁸ Nevertheless, furfural is increasingly relevant, as its market volume exceeds USD 550 million and is expected to reach USD 900 million by the end of the decade,⁹ primarily driven by the demand for furfuryl alcohol (FOL).¹⁰ Furfural consists of a furanic ring with an aldehyde group, and it is obtained from pentoses naturally present in biomass. It finds uses as a precursor for a wide range of compounds including maleic acid, tetrahydrofurfuryl alcohol or the aforementioned succinic acid and FOL. Almost all furfural consumed in Europe is procured by Belgium and the Netherlands, and it is mainly imported from non-EU countries such as China, the Dominican Republic or Israel, with only a small percentage produced in Slovenia and Austria.^{11,12} The reason behind this outsourcing of production relates to the potential environmental damage associated with the operation of the production plants, as these consume large amounts of sulphuric acid.¹³

The industrial production of furfural dates back almost a century, with the technology patented by the Quaker Oats company.¹⁴ Despite the maturity of this process, and the associated technological developments, most of today's commercial processes follow a similar scheme based on a semi-continuous process (see Figure S1). The starting feedstock consists of agricultural or forestry residues, which are hydrolysed in contact with sulphuric acid. Upon hydrolysis, hemicellulosic pentose sugars –mostly xylose and arabinose– are released and subsequently dehydrated to furfural. Typically, both reactions occur in a single batch reactor, where the slurry of biomass and dilute sulphuric acid is subjected to moderate pressure and temperature conditions (about 160°C and 7 atm) by contact with medium-pressure steam.^{14,15} Steam is injected into the reactor as a heating and stripping agent for the rapid separation of the evolving furfural, avoiding condensation and resinification side reactions. At the end of the reaction, the unreacted solid, bearing the sulphuric acid, is filtered and landfilled, while furfural is purified via serial double distillation. In this operation, the first stage removes the carboxylic acids (acetic and formic) present in the mixture, while the second stage separates furfural from water through an azeotropic distillation, producing 98.5 - 99.5% grade furfural.

Some of most recognised technologies for furfural production include the Huaxia, SupraYield, Vedernikovs or Biofine processes.^{11,12} However, almost all furfural currently on the market is produced in China (Chinese batch process), the Dominican Republic (Quaker Oats) and South Africa (Rosenlew process), due to the reasons previously outlined.⁹ Among these technologies, the Chinese Batch Process (CBP) has the largest market share due to its production volume and the low price afforded.^{16–18} CBP feedstock typically comprises corn cobs from maize processing. This is a cheap and abundant material in China,¹⁹ with a high percentage of pentosans in its structure, which makes it ideal for furfural production. The process operates similarly to a Quaker-Oats-type process, although incorporating some energy improvements such as partial condensation of the reactor outlet to provide the heat at the reboiler of the first distillation column.^{14,15}



Beyond commercial ventures, furfural production has recently received considerable attention from research community. Numerous proposals have attempted to advance homogeneous catalysis using different solvents and substrates.^{20–26} However, a growing trend is the use of heterogeneous catalysts that allow the minimisation of corrosion effects and hazardous waste generation, thus contributing to minimize capital expenditures through the use of simpler reaction equipment, and to minimize environmental impacts through the prevention of waste production. Many catalytic routes are based on the use commercially available of supported metal oxides^{27–29} and zeolites. However, some authors have reported low yields in the production of furfural through heterogeneous catalyst driven processes,³⁰ and a significant deactivation after several reuse cycles.^{31,32} Gupta et al. achieved improved reusability by grafting sulfonic acid groups to H- β zeolite,³³ promoting a high Brønsted acidity similar to some ion exchange resins. Notably, the use of these resins in furfural production has gained momentum in recent years, because of the wide commercial availability and low cost of these materials. Some authors have explored the use of strong cation exchange resins such as NafionTM NR 50³⁴ or Amberlyst® 15³⁵ reporting moderate to high yields to furfural, but with the deactivation of the catalytic centres in consecutive reuses. Similarly, Sato et al.³⁶ tested the use of Amberlyst® 70 (a non-commercial cation exchange resin) with a continuous flow of CO₂ at 150°C, obtaining yields similar to those provided by commercial resins. Additionally, Hu et al.³⁷ tested the use of Amberlyst® 70 in water, although reaching very low yields due to the instability of the reaction intermediates. Despite this drawback, other authors have tested water as a solvent, mostly in homogeneous catalytic systems.^{38–40} Even though the use of water can be advantageous, it is mostly used in biphasic mixtures with organic compounds such as toluene^{41–43} or gamma-valerolactone^{30,44–47} to promote dehydration reactions. A comprehensive list of solvent and catalytic systems can be consulted at Edumujeze et al.⁴⁸ With a few exceptions,^{49–51} most of the investigations focus on the laboratory scale disregarding specific scale-up effects (e.g. reactor configuration). This low technological maturity, along with the scarcity on benchmark data, hinders the analysis of constraints on an industrial level and restricts the comparison in terms of economic profitability and environmental performance.

This article aims to identify the process level keys for a sustainable furfural production, for which three critical issues must be addressed. First, the process must be technically feasible on a commercial scale, considering reactor heating limitations and purification system characteristics. Second, furfural production must be economically competitive to significantly impact the current market shares. Third, environmental impact must be notably reduced, as this is the primary reason preventing the production of furfural in countries with developed environmental protection laws. To this end, the scale-up of a furfural production process based on the use of a heterogeneous catalyst (HCP) has been rigorously simulated, in line with previous investigations.^{52,53} This process has been selected for its potential to reduce energy costs and to avoid acid waste due to its characteristic reaction system configuration. The process is conducted in two stages. First, pentosans are hydrolysed into sugars in the absence of external catalyst and under moderate conditions. In a second reactor, the sugars are dehydrated to furfural using Amberlyst® 70 as a heterogeneous acid catalyst instead of H₂SO₄. The evolving furfural is separated using a nitrogen stripping stream instead of conventional steam, potentially reducing energy consumption. Finally, the product is subjected to a double distillation as in commercial processes, yielding 99.5 wt.% pure furfural. In parallel, comprehensive data on the industrial production of furfural has been generated through the rigorous simulation and optimisation of literature data.^{14,15,19} For this purpose, the Chinese Batch Process (CBP) has been selected as benchmark due to its current market share in the production of furfural. This inventory serves to support the comparison against the HCP and provide a reliable reference in the assessment of the economic and environmental performances. For this purpose, technoeconomic analysis (TEA) and lifecycle analysis (LCA) have been applied to both processes aiming to identify critical aspects that



influence the economic competitiveness and environmental impact of the HCP compared to the current furfural production. This analysis can guide future research efforts to address the identified bottlenecks, ultimately leading to the sustainable production of furfural.

2 METHODOLOGY

2.1 Process modelling

The design of the furfural production plant was inspired by existing industrial facilities, which generally handle annual throughputs of several kilotons of furfural, with some reaching a maximum capacity of 35 kilotons per year.^{54,55} This study assumes an annual production scale of 15 kilotons of 99.5 wt.% furfural for both conventional and heterogeneous catalysis-based plants.

Process modelling was carried out using Aspen Plus® V14 software. The thermodynamic framework combined the NRTL model for liquid-phase activity coefficients with the Redlich-Kwong equation of state (EOS) for vapor-phase properties. Since Aspen Plus does not directly provide a pre-defined model for lignocellulosic biomass –e.g. corncobs–, its physical properties were sourced from NREL.⁵⁶

For the CBP, the kinetic data was sourced from the literature.^{57,58} According to these models (see Figure S2), xylans and arabinans are converted directly into their respective sugars (xylose and arabinose), while the acetyl groups are hydrolysed to acetic acid. Subsequently, free sugars are converted into intermediates, finally yielding furfural. Additionally, the hydrolysis of cellulose to produce glucose was also considered in the kinetic model of CBP, despite the low reaction extent at the operating conditions.

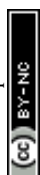
In the HCP, reaction kinetics were implemented using two CSTR reactor models. The first one, involves non-catalytic autohydrolysis reactions where xylans and arabinans are initially converted into their respective oligomers, which are subsequently broken down into sugars (see Figure S3). Similarly, acetyl groups are hydrolysed to acetic acid. Under the operating conditions of this stage, cellulose and lignin remain unaltered. The corresponding kinetic data were derived from the literature.⁵⁹ The second reactor focuses on the conversion of the free sugars into furfural through heterogeneous catalysis. In this case the kinetic data were obtained in the laboratory.⁵²

Both models account for the formation of tars via two mechanisms: the reaction between xylose and furfural, and the self-resinification of furfural. Moreover, the stripping effect was simulated considering the phase equilibria of the two outlet streams (waste liquid stream and the vapor stream derived from the stripping).

Distillation columns were modelled using the rigorous RadFrac model. The column configurations, including the number of stages, were optimized by minimizing Total Annualized Costs (TAC).⁶⁰ TAC, as defined in Eq. 1, accounts for operational expenses by combining capital costs, amortization periods, and utility expenditures.

$$TAC = \frac{\text{Capital cost}}{\text{Payback period}} + \text{Utility costs} \quad \text{Eq. 1}$$

A sensitivity analysis was conducted to identify the optimal feed stage, aiming to minimize reboiler energy consumption.



2.2 Economic Analysis

View Article Online
DOI: 10.1039/D5SU00106D

The economic evaluation encompassed both capital expenditures (CapEx) and overall operational expenditures (OpEx), with the minimum selling price (MSP) of furfural serving as the benchmark for assessing economic viability.⁶¹ The MSP was determined to ensure that the annual revenue aligns with the Equivalent Annual Operating Cost (EAOC) (Eq. 2). EAOC includes both annualized capital costs and operational expenses, calculated using a discounted cash flow approach. This analysis assumed 10-year project lifespan, 20% internal rate of return, and 40% corporate tax rate.⁶² The annualized capital cost (Eq. 3) was distributed over the project's duration and combined with OpEx to calculate the EAOC (Eq. 4). All costs were calculated in euros, assuming a currency exchange rate of 1 USD = 0.95 €.

$$MSP = \frac{\text{Annualized capital cost} + \text{operating cost}}{\text{Plant FOL Production Capacity}} \quad \text{Eq. 2}$$

$$EAOC = \text{Annualized capital cost} + \text{operating cost} \quad \text{Eq. 3}$$

$$\text{Annualized capital cost} = \frac{\text{Capital cost} \cdot i \cdot (1 + i)^n}{(1 + i)^n - 1} \quad \text{Eq. 4}$$

Where:

- n: Economic life of the plant (10 years).
- i: Internal rate of return (20%).

CapEx estimations adhered to standard methodologies, utilizing equipment cost data from Aspen Process Economic Analyzer (APEA-V14), based on first-quarter 2022 figures. This tool, integrated within Aspen Plus®, accounts for equipment costs, installation, bulk plant systems (e.g., power distribution, control systems), and indirect costs. Steam ejector costs were considered negligible.⁶³

In both the conventional CBP and the alternative HCP, carbon steel was selected as the construction material, as it is commonly employed in current furfural production plants. To withstand the corrosive effects of sulphuric acid, the reactor walls in the CBP are designed with an extraordinary thickness of 50 mm,¹⁴ and therefore this measure was considered in the CapEx estimation for the reactors in the conventional process. In contrast, for the HCP the wall thickness was adjusted according to the specific operating conditions due to the absence of acid.

Total OpEx included raw materials, utilities, catalysts, and supplementary expenses such as labour, maintenance, overheads, and administration.⁶² Raw material and utility costs were escalated annually by 3%.

Raw Materials

The raw materials used for both furfural production processes include corncobs and water. The cost of corncobs has been set at €81.42 per metric ton,⁶⁴ while water is priced at €1.61 per metric ton, as reported by a Spanish oil refining company. Corncobs are assumed to have a moisture content of 13%, with their dry composition consisting of 45.2% cellulose, 35% hemicellulose, 16.7% lignin, and 3.1% ash and other components. These values were calculated as a mean value of the different references published in Phyllis2 database.⁶⁵

The H₂SO₄ is considered an additional raw material in the CBP, with a cost of €57 per metric ton.⁶⁴ In contrast, the HCP is based on heterogeneous catalysis, utilizing Amberlyst® 70 resin, priced at €47.5 per kilogram.⁵⁴ The catalyst replacement time was estimated using a productivity ratio,



assuming that 1 kg of catalyst produces 1,000 kg of furfural before disposal.⁵⁴ After optimizing the required catalyst amount (as detailed in section 3.1) and applying this production ratio, the catalyst replacement interval was calculated to be 5.05 days. While this might seem relatively short, it aligns with experimental findings reported by Agirrezabal et al.,^{52,53} who observed that catalyst particles become coated with tars formed during course of the reaction.

These prices, along with estimated process consumption rates and costs for minor reagents such as sodium carbonate used in the CBP, are summarized in Table 1. Additionally, the table provides the market prices of furfural, cellulose, and lignin.

Table 1 Summary of the main costs of raw materials, catalysts and utilities.

Raw Materials	Price (€/Mt)	Consumption CBP (Mt/year)	Consumption HCP (Mt/year)
Corncoobs ⁶⁴	81.40	142,800.00	142,800.00
Process Water	1.61	164,529.00	21,797.00
H ₂ SO ₄ ⁶⁴	57.00	2,789.20	-
Heterogeneous Catalyst ⁵⁴	47,500	-	8.40
Sodium carbonate ⁶⁶	339.10	243.60	-

Utilities	Price
Steam@250 psi (€/Mt)	34.88
Steam@50 psi (€/Mt)	34.03
Electricity (€/MWh)	229.23
Cooling Water (€/GJ)	1.32

Product	Market price (€/Mt)
Furfural ^{18,47}	1000-1500
Cellulose ⁶⁷	776.10
Lignin ⁶⁸	190.00

^a Figures provided by a Spanish oil refining company (2023).

^b Updated using CEPCI value from 2023 (797.1).

Utilities

Utility prices were derived from data provided by a Spanish oil refining company for Q1 2023 (Table 1). Cooling water was supplied from a central facility with a cooling tower, including fans, makeup water, chemical injections, and pumps. Calculations accounted for 0.3% windage loss from mechanical draft towers, maximum salt concentration factor of 5, pump efficiency of 75%, and chemical cost of €0.148 per 1,000 kg of makeup water.⁶²

2.3 Life Cycle Assessment (LCA)

The methodology used for the LCA study is described below following the ISO 14044 structure.

Goal and scope

The LCA was carried out using an attributional approach to evaluate the production of 1 kg of furfural at the factory gate with 99.5 wt.% purity. The scope is cradle-to-gate, although the treatment of effluents coming out directly from the process is also considered. The LCA follows a comparative approach, where both product systems have the same boundaries and produce the same functional unit. The system boundaries are schematically depicted in Figure 1.



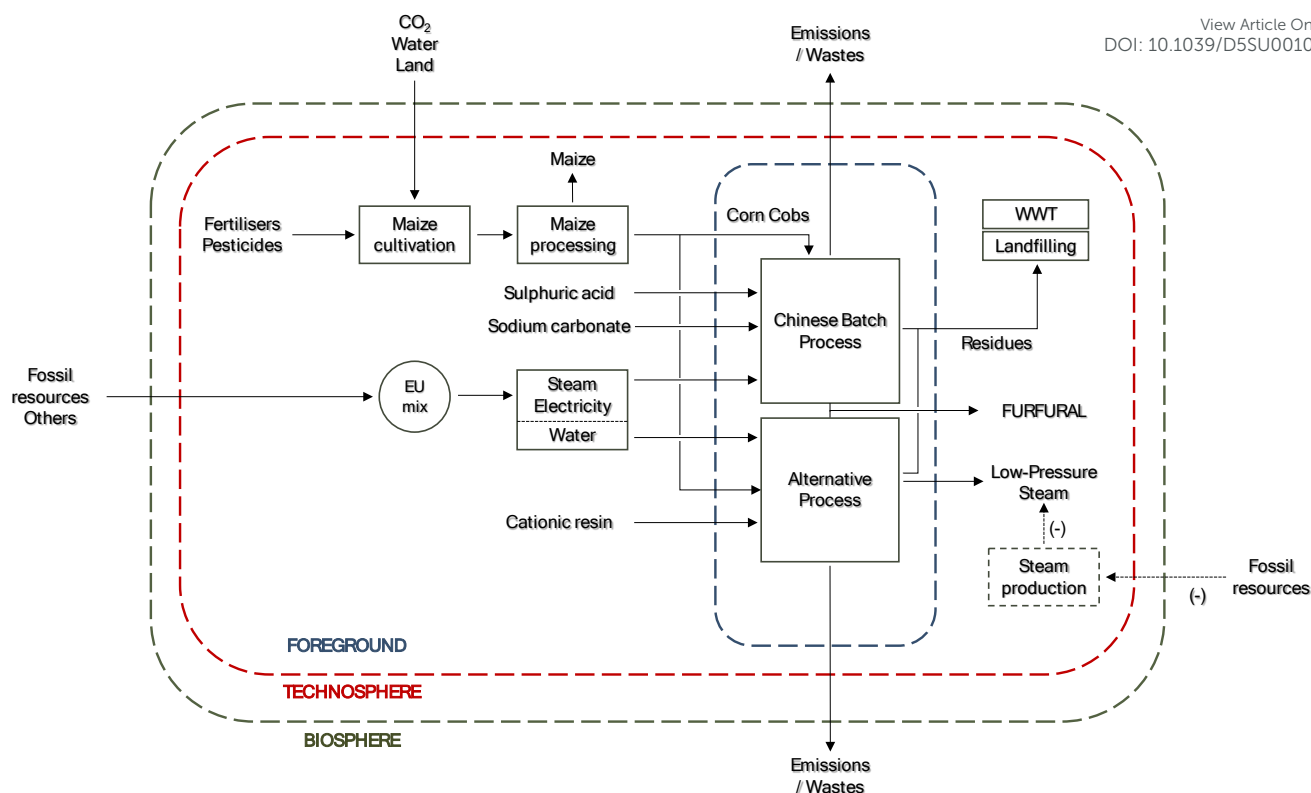


Figure 1 Boundaries defined for the systems under study. Only the most relevant activities at the background level are shown.

Life cycle inventory (LCI)

Data sources for the inventory modelling of both HCP and CBP systems are described in sections 2.1 and 2.2. The detailed inputs and outputs can be consulted in Tables S4.1 to S4.4. For the construction of the background system, ecoinvent version 3.10.1 was used. The APOS (at the point of substitution) model was chosen to capture the entire upstream supply chain, including the impact of products from waste recovery activities. All data were managed using the open-source tool Brightway v2.5 through the Activity Browser v2.11.⁶⁹

Life cycle impact assessment (LCIA)

The Environmental Footprint v3.1 was used as the calculation method due to its broad consensus for European data modelling.⁷⁰ All indicators included in this methodology have been assessed, although disaggregated data on biogenic carbon and the contribution of organic and inorganic substances to certain indicators are only provided in the supporting information (Table S9) to avoid cluttering the discussion. Given the low technology readiness (TRL) of the HCP and the scarcity of transparent CBP information, the uncertainty of the input data has been considered using a lognormal deviation of the inputs based on the Pedigree matrix.⁷¹ This uncertainty was propagated through the LCA calculations by Monte Carlo sampling (1000 runs). The distribution of impact data for each evaluated category is shown in Figures S5 to S19. Global sensitivity analyses (GSA) were also carried out in the case of critical deviations to identify and explain the parameters responsible for most of this uncertainty.

Interpretation

The interpretation of the results focused on the identification of the main hotspots of both systems, examining their contribution to the different environmental impact categories being assessed. The analysis has a process-oriented character, trying to map the impacts to the different sections of the



plant according to the design described in section 2.1 and Figure 2 and S1. This analysis perspective aims to derive the main differences between the two systems, the potential improvements, the challenges and the limitations of the study.

2.4 Assumptions and limitations

Several key assumptions were considered along the assessment. These assumptions and constraints are critical to interpreting the results and provide the context necessary for evaluating the economic and technical feasibility of the process:

- Although CBP is carried out in China, to ensure a fair comparison with the proposed technology, the techno-economic and environmental analyses have been carried out using European data. This avoids the price and impact differences associated with the regionalization of data. It is important to note that this may lead to deviations, such as an increase in the reference price of furfural due to the higher cost of critical raw materials, or a decrease in the impact associated with the energy mix. In any case, it ensures a fair comparison framework, which better aligns with the objective of this work.
- In CBP, two waste streams are generated: spent liquor from the reactor emptying and the bottom stream from the first distillation column. The first stream contains approximately 27 wt.% of solids consisting of cellulose, lignin, other oligomers, ash and tar. These solids are filtered and sent to a controlled landfill. The liquid fraction, containing a high acid load, undergoes water treatment prior to discharge. However, some acetic and sulphuric acids are retained in the solid cake, together with other compounds such as furfural and other by-products. In the HCP, two waste streams are also produced in the autohydrolysis and stripping reactors. The first stream consists of a solid fraction with a similar composition to that of the conventional reactor, although free of mineral acids and other harmful compounds. The second reactor produces a stream containing a minor tar fraction, which is filtered and sent to a controlled landfill, while the aqueous fraction is sent to conventional water treatment.
- The CBP may produce by-products like methanol or acetic acid, as mentioned by some authors.^{14,19} Information on methanol production is only found in the description by Win,¹⁹ albeit it is not supported by transparent data and refers to the continuous process (Huaxia as modified by Westpro). Similarly, to the best of the authors' knowledge, there are no references to methanol production from methylated biomass structural polymer sections. Therefore, it was decided not to include this by-product in the simulation and LCA calculations. On the other hand, the bottom stream of the first distillation column contains 1.2 wt% acetic acid, which falls within the range reported in the literature,¹⁴ although is too low to justify its recovery according to consulted references.¹⁵ Therefore, the purification train simulated does not comprise the recovery of this compound, and it was not considered as a by-product. Finally, in the HCP, although the potential to obtain high-quality cellulose and lignin through purification — comprising filtration, washing, and precipitation — has been considered, this design was not implemented. Consequently, these by-products were assumed to undergo landfilling as a part of the waste solid cake.
- The HCP process exhibits a higher integration than the CBP due to the flexibility in its design approach. While the proprietary process was modelled to maximize performance, the CBP simulation aimed to provide a representative benchmark. Thus, only design parameters explicitly reported in the literature were used, with optimization limited to the ranges provided (e.g., stripping steam flow rate, distillation column design). However, commercial implementations may be more optimized than described in the literature, introducing a potential limitation to this comparison.



3 RESULTS AND DISCUSSION

View Article Online
DOI: 10.1039/D5SU00106D

3.1 Basic Process Engineering: HCP vs CBP

As previously disclosed, HCP was initially featured with distinct characteristics as compared to CBP which were expected to strongly improve the environmental footprint of furfural production while preserving its techno-economic feasibility. Two main aspects have been investigated:

- the use of nitrogen (N_2) as a stripping agent instead of steam,
- and the utilization of a heterogeneous catalyst to eliminate the need for sulphuric acid (H_2SO_4)

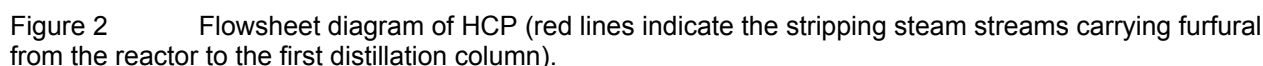
The use of N_2 as a stripping agent has been published as a promising alternative due to its lower cost.⁵² However, process simulations about this option suggest that N_2 would remove a significant portion of the water from the reaction medium since it would leave the reactor saturated in steam. This phenomenon increases the formation of tars due to the increase in the concentration of tar-precursors in the reaction medium. Moreover, it must be considered that besides acting as entrainer, steam also acts as a heating media, so that if using N_2 , an external heating source must be implemented. For this reason, N_2 was discarded as a stripping agent in the second reactor, and steam continued to be used as in the conventional CBP.

The second potential improvement focuses on the use of commercial Amberlyst® 70 as a catalyst. Unlike H_2SO_4 , a heterogeneous catalyst hydrolyse biomass as a solid within a liquid reaction media. Therefore, the use of this catalyst implies that the hydrolysis must be divided into two stages to favour the contact between the heterogeneous catalyst and the carbohydrate raw material (see Figure 2). The first stage involves autothermal hydrolysis (R101), where biomass is partially hydrolysed at 190 °C (16 bar),⁵⁹ aided by the presence of acetic acid evolving from biomass. This is conducted in a continuously stirred tank reactor (CSTR) with an optimized residence time of 7.5 minutes under the specified conditions. These conditions enable the depolymerization of the biomass, achieving the hydrolysis of 92% of the biomass, with the majority of xylans and arabinans (71%) converted to oligomers and 15% to free monosaccharides solubilized in the liquid water. Various temperatures (160–190 °C) and residence times were simulated in Aspen before determining the optimal conditions. Figure S4 shows the evolution of the different species that are present in the reaction media in function of the residence time and the temperature. Higher residence times led to increased biomass hydrolysis and furfural production, but also promoted tar formation, making 7.5 minutes at 190 °C the optimal setting, keeping the solids (the sum of unhydrolyzed xylans and arabinans and tars) below 15%. Before moving to the second stage, a filter is required to separate the solid fraction, consisting of cellulose, lignin, and unhydrolyzed hemicellulose (7-8%). This stream has potential value in an integrated biorefinery concept.

The second hydrolysis step (R-102) of the HCP requires higher temperature conditions as compared to the CBP (200 °C instead of 160 °C). The main reason is the increased time required to conduct the hydrolysis of the oligomers into monosaccharides according to kinetic data.^{52,53} Lower operating temperatures, such as 175 °C, could be used as suggested by Agirrezabal-Telleria et al.⁵² However, it is important to consider that in R-102, oligomers continue depolymerizing into sugars, and at 175 °C, this process occurs significantly slower than at 200 °C (see Figure S4). As a result, insufficient xylose and arabinose would be available to achieve the furfural production target set in this study, given a residence time similar to that of the CBP process, which is 4 hours. Besides, since the previous hydrolysis step is conducted at 190 °C, maintaining a similar temperature seems appropriate. This is further supported by the fact that medium-pressure steam has a temperature of 205 °C, allowing the reactor to operate adiabatically at 200 °C. Bearing this in mind, multiple parallel stripping



View Article Online
DOI: 10.1039/C5SU00010G



Initial simulations of the HCP under the mentioned conditions evidence that 325,000 kg/h of steam was required to achieve the same furfural production rate, i.e. 2,550 kg/h, (127.5 kg of vapor/kg of furfural distilled, see Figure 3) at the stripping outlet as in the CBP, which needed only 63,671 kg/h (25.4 kg of vapor/kg of furfural distilled). These values refer to the overall values in R-102 system, i.e., the sum of the vapor required and furfural produced in each parallel reactor. This seemed counterintuitive, as the relative volatility of furfural with respect to water is slightly higher at 200 °C than at 160 °C (0.172 vs. 0.158). This discrepancy suggests that the rate of tar formation in the heterogeneous catalysts is excessively high, and therefore very large stripping vapor rates are required to limit furfural degradation.

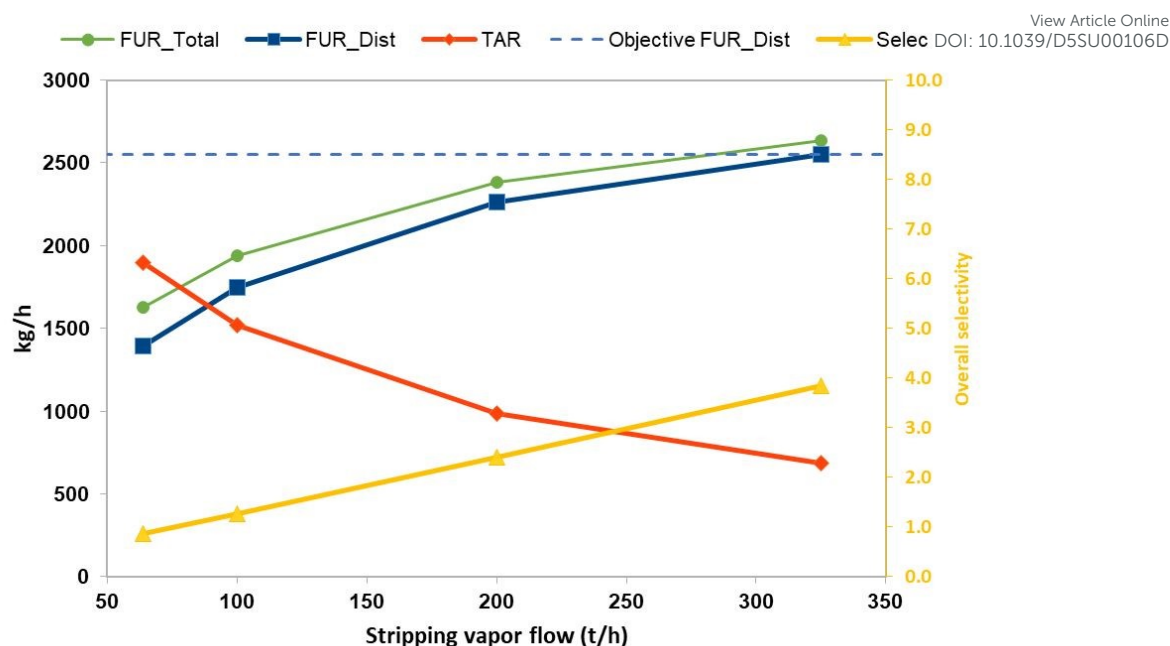


Figure 3 Furfural and tars production and the overall selectivity among them (kg of furfural distilled/kg of tars) in function of the stripping vapor flow. (200 °C). “FUR_total” accounts for the total furfural output (distillated + waste stream) in the reactor. “FUR_Dist” is the flow of furfural that leaves the reactor with the stripping steam. “Objective FUR_Dist” accounts for the furfural flow that leaves the reactor with the stripping steam in CBP.

As it is described in the techno-economic assessment results (Section 3.2, *vide infra*), the cost of stripping steam is very significant and accounts for almost 32.4% of the total EAOC in the CBP. Moreover, larger steam flows also increase distillation costs as more diluted furfural is obtained. Therefore, the use of Amberlyst® 70 makes the HCP economically unfeasible due to its low selectivity, which demands impractically high steam stripping flows. To enable a viable process, catalyst improvements are necessary.

Under this background, it is critically important to determine the global selectivity value required for a future heterogeneous catalytic process to enable the economically viable production of furfural on an industrial scale. This parameter can serve as a key benchmark to steer the scientific community's efforts toward the development of optimized solid acid catalysts that not only surpass the performance of Amberlyst® 70 but also compete effectively with homogeneous catalysts. As illustrated in Figure 4, an increase in the rate of furfural formation relative to the rate of tar formation (instantaneous selectivity), while keeping the stripping vapor flow fixed at the level used in the CBP process (63,671 kg/h), gradually improves the furfural yield, bringing it closer to the furfural production achieved in CBP (2550 kg/h). Based on these findings, the development of a catalyst exhibiting a global selectivity of 4.8 kg of furfural per kg of tar is essential. This represents a 5.5-fold improvement over Amberlyst® 70 and is necessary to match the CBP steam stripping consumption while maintaining the same furfural production rate (25 kg of vapor/kg of furfural distilled).- Interestingly, the overall selectivity in the CBP is 1.5 kg of furfural per kg of tar when operating under the same conditions. Initially, it was hypothesized that the excessive steam requirement in HCP could be attributed to a high rate of tar formation on the heterogeneous catalyst. However, a closer examination of the overall selectivity (kg of furfural produced/kg of tars produced) values suggests a different explanation. In the CBP process, the overall selectivity is 1.5, whereas in HCP, when operating with 325,000 kg/h of stripping steam, the selectivity increases to 3.84. This difference arises because, unlike in CBP, only approximately



80% of the hemicellulose undergoes complete hydrolysis to monomeric sugars in HCP in 4 h. Therefore, the heterogeneous catalyst must achieve a higher selectivity than sulphuric acid in CBP to ensure that the available sugars, which are fewer in HCP, are directed proportionally more towards furfural rather than tars, compared to the CBP process. These variations can be observed in Figure 3 and Figure 4.

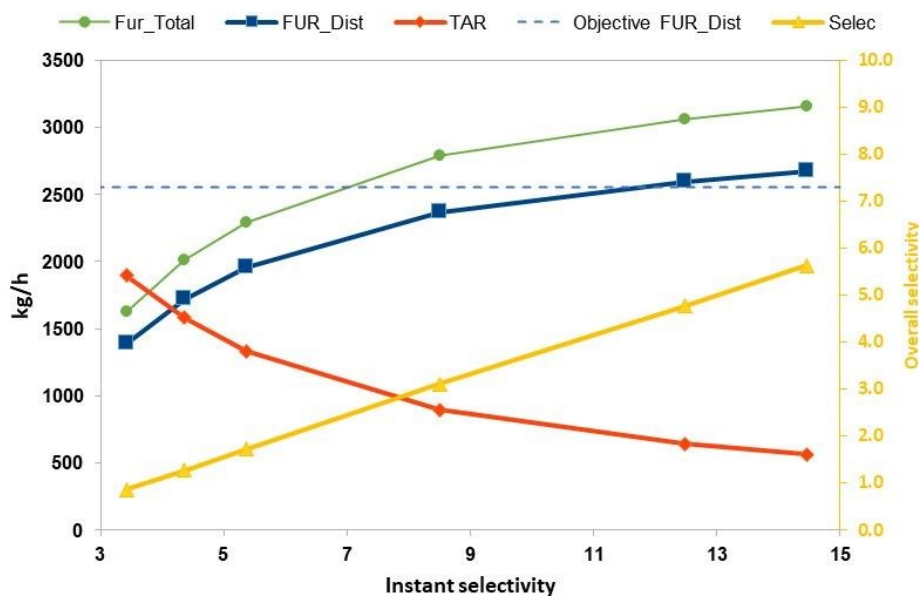


Figure 4 Furfural and tars production and the overall selectivity (kg of furfural distilled/kg of tars) among them in function of the instantaneous selectivity (rate of furfural production / rate of tars production). (200 °C & 63,671 kg/h of stripping vapor). "FUR_total" accounts for the total furfural output (distilled + waste stream) in the reactor. "FUR_Dist" is the flow of furfural that leaves the reactor with the stripping steam. "Objective FUR_Dist" accounts for the furfural flow that leaves the reactor with the stripping steam in CBP.

A sensitivity analysis was conducted to identify the optimum amount of catalyst in each reactor (see Figure 5). By varying the total catalyst amount (sum of the 17 parallel reactors) between 50 and 1000 kg, furfural production was found to reach its maximum at 250 kg, equivalent to 14.6 kg per reactor. Moreover, when fixing the stripping steam flow at 63,671 kg/h (25 kg steam/kg furfural), only a certain amount of furfural can be effectively removed from the reaction medium. Beyond a certain catalyst loading, while the catalyst promotes both furfural and tar formation, the additional furfural generated is not fully carried away by the stripping steam, remaining in the reactor. This leads to its resinification and therefore, increasing tar formation.

In addition, at least 1000 kg of furfural per kg of catalyst must be produced before disposal of the catalyst, based on the stability criteria for industrial catalyst viability established by Lange.⁵⁴ Meeting this target is particularly challenging, as heterogeneous catalysts like Amberlyst® 70 are prone to deactivation due to fouling caused by tar and other heavy products deposited on the active sites. Consequently, in addition to developing a more selective catalytic material, it is essential to ensure that it can be easily regenerated. For Amberlyst® 70, regeneration through simple calcination is not feasible due to the organic nature of the material, which limits its thermal stability, and other regeneration strategies with dilute acids are required for restoring its activity.⁷² In this regard, inorganic materials present the advantage of allowing thermal regeneration, and therefore, they represent a more promising alternative. If a catalytic material meeting the lifetime criterion of 1000 kg of furfural per kg of catalyst is developed, the catalyst consumption is established in 15 Mt per operating year.



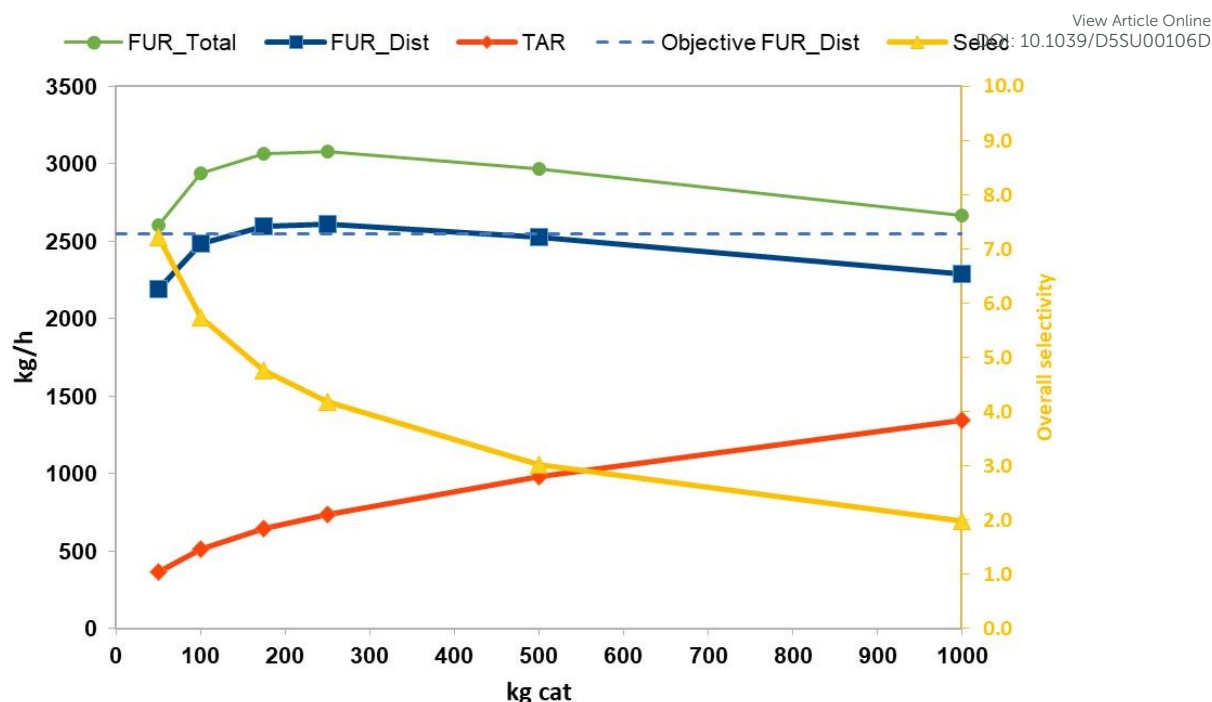


Figure 5 Effect of the catalyst amount on the furfural and tars production and their selectivity at 200 °C and using 25 kg of stripping vapor/kg furfural. “FUR_total” accounts for the total furfural output (distilled + waste stream) in the reactor. “FUR_Dist” is the flow of furfural that leaves the reactor with the stripping steam. “Objective FUR_Dist” accounts for the furfural flow that leaves the reactor with the stripping steam in CBP.

In summary, the intended technology substitution could be achieved by developing a more selective (with an overall selectivity of at least 4.17 kg of furfural per kg of tar, after optimizing the catalyst amount) and stable (with a lifetime of at least 1000 kg of furfural per kg) heterogeneous catalyst. It is expected that under these conditions economic parity would be accomplished while avoiding the use of sulphuric acid, which would contribute to a significant environmental benefit. In this context, and to point toward promising directions for future research, advances should focus on both functionality (selectivity increase) and catalytic support stability. Considering the first point, efficient furfural production from xylose typically benefits from catalysts combining Brønsted and Lewis acid sites (or Brønsted base sites), which promote isomerization to xylulose and subsequent dehydration.^{28,32} The current system uses a sulfonated ion-exchange resin providing Brønsted acidity leading to overall good results, but which may be enhanced by adding this double functionality according to literature.⁷³ On the other hand, regarding the low thermal stability of ionic resins, alternatives like Nafion NR-50 are more thermally stable -up to 240°C- but less selective.³⁴ Zeolites, though tuneable and suitable for functionalization, often present diffusion limitations due to small pore sizes in the case of beta structures and are unstable in aqueous systems.^{31,33} Similarly, silica-based catalysts degrade in water, and while metal oxides are more robust and water-tolerant, their catalytic activity tends to be lower.^{27,29} A more promising direction may lie in acid-functionalized carbon-based catalysts, particularly those doped with nitrogen, which has demonstrated high furfural yields (>88%) from lignocellulosic feedstocks, provided gamma-valerolactone (GVL) is used as a solvent.⁷⁴ However, their performance in water drops significantly due to deactivation, so this should remain a key focus.



3.2 Techno-economic assessment

Figure 6 presents the EAOOC calculated for both processes. In the HCP, the CapEx accounts for 10.2% of the EAOOC, whereas in the CBP process, it represents 11.2%. The higher CapEx in the CBP process can be attributed to two primary factors:

1. Number of stripping reactors. As stated in section 3.1, to achieve the targeted furfural production, the conventional process requires 20 stripping reactors (R-101, Figure S1) arranged in parallel, while the HCP requires only 17.
2. Material and wall thickness of reactors. The conventional process operates with sulphuric acid, necessitating reactor walls with a thickness of 50 mm,¹⁴ which exceeds the thickness required by the operational pressure and temperature conditions. In HCP, despite the higher operating pressure would typically require thicker reactor walls, using a heterogeneous catalyst the wall thickness is not as thick as in CBP, thereby construction costs are cheaper.

It should also be noted that the HCP includes an additional reactor for the autothermal hydrolysis stage (R-101). However, due to the short residence time in this stage (7.5 min), the reactor volume is only 6.3 m³, approximately half the size of the stripping reactors. Consequently, the cost savings from eliminating three stripping reactors more than offsets the investment required for this initial hydrolysis reactor.

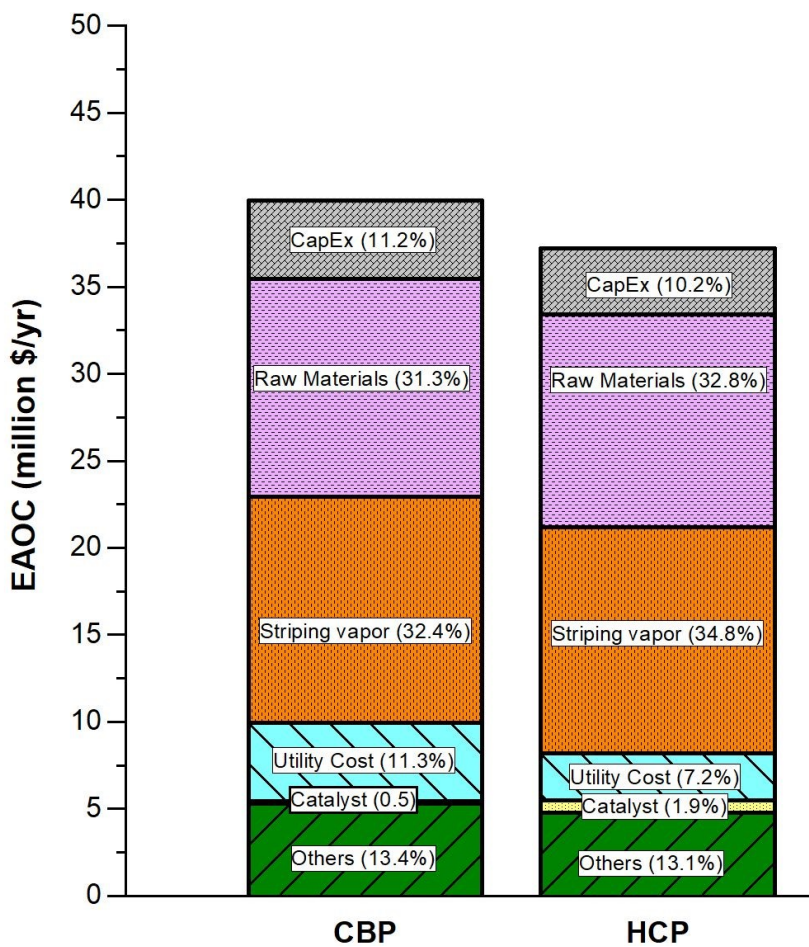


Figure 6 Distribution of EAOOC among the capital cost (plain) and operating costs (pattern blocks).

The breakdown of the operating costs for both processes is summarized in Table 2. The cost of raw materials is similar for both alternatives, amounting to €12.2M in HCP and €12.5M in the CBP, corresponding to 36.5% and 35.2% of their respective total OPEX. This slight difference is attributed to the water recirculation implemented in HCP. The cost of the biomass feedstock, however, is identical in both cases.

Regarding utility costs (excluding the stripping vapor), the CBP incurs slightly higher expenses (€4.5M vs €2.7M). This difference is primarily attributed to the enhanced energy integration achieved in the HCP. It is important to note that the selectivity of the catalyst in HCP was adjusted to ensure that the stripping steam consumption is equivalent in both processes.

The costs associated with the treatment of the reactor effluent, containing organic residues and spent catalysts (sulphuric acid in the CB process and a heterogeneous catalyst in the HCP), were considered negligible. Although, in CBP, neutralization of the sulphuric acid would be required, these treatment costs account for less than 1% of the total operating costs in both cases and were not considered.

As expected, the cost of the heterogeneous catalyst in the HCP is higher than the cost of sulphuric acid in the CBP (€709,000/year vs €159,000/year). Nevertheless, this expense represents a minor fraction of the total OpEx, contributing 2.1% and 0.4%, respectively.

Once CapEx and OpEx were accounted for, the MSP calculated for the two processes was €2,685 per metric ton for the CBP and €2,499 per metric ton for HCP. These results indicate that the production cost of furfural would be quite similar in both cases. Given the $\pm 30\%$ margin of error typically associated with such calculations,⁶² it cannot be conclusively stated that one process is superior to the other. However, under current assumptions the proposed HCP demonstrates economic viability.

When compared to the current market price of furfural, which ranges between €1,000 and €2,000 per metric ton –and even lower–, the estimated MSP is higher.¹⁸ This discrepancy can be attributed to several factors that were not considered in this study but could significantly reduce production costs:

1. Energy price assumptions. This study uses the current energy prices in Europe, which are significantly higher than those in China, where the CBP is primarily implemented.
2. Utilization of reactor residue. In the existing processes, the biomass residue exiting the stripping reactor is dried and burned to produce heat for generating the stripping steam, which reduces the associated vapor cost.¹⁴

In the HCP, the unhydrolyzed cellulose and lignin fraction forms a separate stream. While this stream could also be dried and burned as in the conventional process, the absence of mineral acids and its unhydrolyzed state open possibilities for valorisation within a bio-refinery scenario. Although the complete design of this integrated bio-refinery is beyond the scope of this study, a preliminary economic estimate was conducted to assess how the valorisation of this stream could impact the MSP of furfural. A selling price of €776.1/Mt and €190.0/Mt (see Table 1) was assigned to the cellulose and lignin streams, respectively. For cellulose, this estimate takes a conservative approach, reflecting the current market price of cellulose pulp. However, further valorisation of this fraction into higher-value products, such as 2,5-furandicarboxylic acid, could substantially increase its market price. In order to estimate the cost of processing this stream –including annualized capital and operating expenditures– into purified lignin and cellulose pulp, the previously described furfural production cost structure was applied, considering an annualized capital and operating costs (excluding raw



materials) of 67.2% of the MSP value. Assuming this proportion remains constant, the net income from valorising lignin and cellulose streams represents the 32.8% of the selling price: €255/Mt for cellulose and €62.3/Mt for lignin. Using this potential net income, the recalculated MSP for furfural was determined to be €1455/Mt. This value can be deemed competitive considering the current market price of furfural, demonstrating that a fully integrated biorefinery concept, which efficiently utilizes all biomass fractions, can lead to significant cost reductions.

Lastly, it is worth noting that in the alternative HCP, all available heat streams can be used to generate low-pressure steam. Although this was not included in the economic calculations, it could further support the valorisations of the cellulose and lignin fraction in an integrated biorefinery context.

Table 2 Costs analysis of furfural production process: CBP vs HCP.

Installed Cost (€)				
	CBP		HCP	
Reaction section	13,610,000	83.6 %	9,937,900	75.8%
Furfural purification	1,875,200	11.5 %	2,047,300	15.6%
Others	797,000	4.9%	1,132,700	8.6%
Total Installed Cost	16,282,200		13,117,900	
Added costs ^a	2,512,600		2,838,700	
Total Capital Cost	18,794,800		15,956,600	
Annualized Capital Cost (€/yr)	4,482,988		3,806,012	
^a Indirect costs, plus contingency and fee, plus auxiliary facilities.				
Operating costs (€/yr)				
	CBP		HCP	
Raw Materials Cost	12,543,765	35.2 %	12,239,965	36.5%
Utilities Cost	4,518,890	12.7 %	2,670,750	8.0%
Stripping vapor Cost	12,978,735	36.5 %	12,978,735	38.8%
Catalyst Cost	158,984	0.4%	709,012	2.1%
Others ^b	5,362,310	15.1 %	4,882,350	14.6%
Operating Cost	35,598,427		33,492,062	
^b Operating labour, maintenance, overheads...				
Final economic assessment				
	CBP		HCP	
EAOC (€/yr)	40,081,415		37,298,075	
Furfural production capacity (kt/yr)	14.9		14.9	
MSP (€/t)	2,685 (1,000-2,000 ^c)		2,499 (1,455 ^d)	
^c Market price				
^d Considering the valorisation of all biomass fractions				



3.3 Life Cycle Assessment

View Article Online
DOI: 10.1039/D5SU00106D

Life cycle assessment has been performed on both CBP and HCP following the same assumptions than for the TEA (section 3.2). Figure 7 summarises the impact of both processes on each environmental category. For each pair of cells, the colour scheme reflects the magnitude of the proportional difference between the values, with the highest value considered 100% (red) and the lower values graded according to their proportional difference. Values with a higher difference are shifted towards green for easier interpretation. At first glance, it is easy to see an improved environmental performance of the HCP for all measured environmental indicators. Overall, three critical inputs are identified that explain most of the impacts in both processes. These are the production of the heating steam by burning fossil fuels, the consumption of corncobs as raw material and the waste generation. Therefore, the differences in these flows between the two processes account for most of the differences in the calculated values. Thus, the lower energy requirements of the HCP and the cleaner discharge of the first reactor are the main responsible factors for the ostensible differences shown in Figure 7. This section provides a detailed discussion for each critical input, analysing their impact on the environmental indicators and linking these effects to the configuration of both processes.

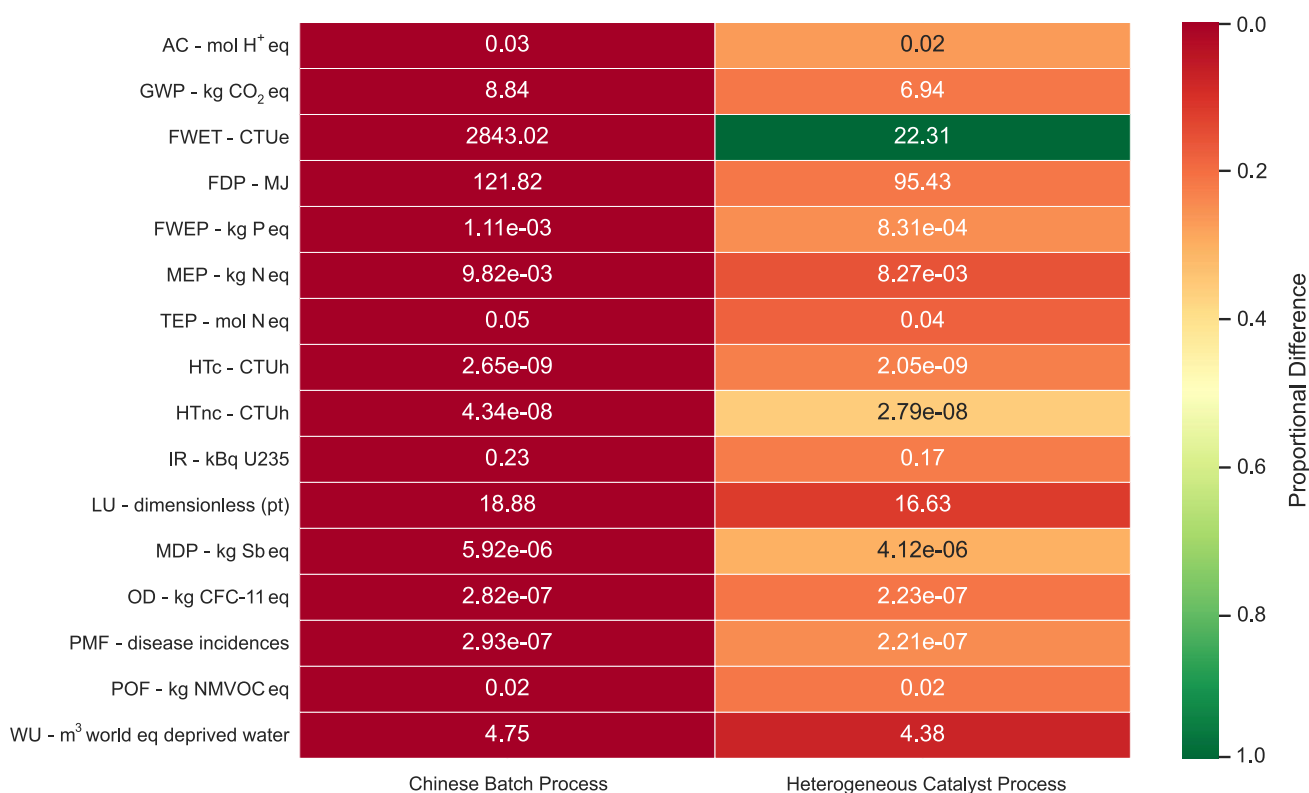


Figure 7 Overview of the impact values of both furfural production processes over the studied environmental indicators.

Energy requirements

Meeting the energy needs of both systems has the highest environmental cost, being in both cases the main driver of the environmental impacts. Fossil fuel combustion for steam generation has the highest percentage impact on acidification, global warming, fossil fuel consumption, eutrophication, human toxicity, ionising radiation, ozone depletion, particulate matter and oxidant formation indicators. At the process level, the highest energy consumption is allocated on the stripping reactor. This means that this stage bears the greatest impact on the overall production of furfural. The main



reason is the configuration of the reaction system itself, which must effectively separate the evolving furfural from the liquid phase to avoid side-reactions (condensation and resinification) which become extensive for prolonged contact times, for which stripping steam is used (see section 2.1). As both processes use the same steam flow, the effects associated with the reaction system are very similar. This might be regarded as a weakness of the HCP as it achieves equivalent performance but does not reduce the impact at the most critical point in the process (R-202), so future work should address this issue.

Despite this, the HCP reaches a better overall performance due to energy reductions in other sections of the plant. On the one hand, significant differences derive from the preheating system and the reboiler of the second distillation column, being the first one the most important. In the conventional process, 2.6 kg of steam is required per kilogram of furfural, whereas the HCP reduces this demand to 1.4 kg. This is attributed to enhanced energy integration within the process, as in the HCP, diluting water is fed at 187 °C, compared to only 25 °C in the baseline technology. This temperature increase is achieved by recirculating a portion of the water (45 wt.%) from the bottom of the first column (C-101), containing 99 wt.% water, at approximately 100 °C. This stream is subsequently heated by contact with the product stream exiting the stripping reactor (HX-101), reaching 187 °C prior to the introduction to the HX-102 exchanger.

On the other hand, the HCP avoids the consumption of low-pressure steam (LPS) from the second distillation column (C-102). This is achieved by forcing an identical configuration to the first tower by passing the product stream through the reboiler of this second column. In this way, the product partially condenses before entering the purification train, releasing its latent heat to allow this equipment to operate satisfactorily. This improvement saves 0.12 kg of steam per kg of furfural produced.

Finally, the HCP produces an LPS stream by contacting the product stream from the second reactor (R-102) with water at boiler B-101. This is done for total condensation before entering the C-101 column, preventing the production of steam using the conventional European mix, generating an environmental credit as shown in Figure 2. The LPS production is particularly relevant, as for each kilogram of furfural produced, the generation of 4.3 kg of heating steam is avoided. The positive impact of this by-product has a significant impact on the comparison between the two processes, although as shown in Figures S7.1-S7.16, this reduction is not decisive for their comparative performance.

Corn production

Although cobs from maize production are an agricultural residue, they still carry an impact because they are also an integral part of maize production and represent a fraction of the resources used to grow it –corn cobs might represent 10 wt% of the overall plant on a dry basis–. Therefore, a 0.1 allocation factor has been applied based on their proportional weight. Both processes have almost identical consumption of this raw material, therefore, the differences are negligible on an absolute scale and the deviation is only noticed in relative terms, as the energy impacts are less significant for the HCP.

Corn cob consumption significantly impacts indicators such as eutrophication, human toxicity, mineral resources consumption, particulate matter, and land and water use. These effects are primarily linked to the application of agrochemicals, particularly fertilizers and pesticides. Consequently, efforts to mitigate these impacts should focus on the supply chain level, as these are more extended issues.⁷⁵ At the process level (foreground), solutions are more limited in terms of feedstock, as materials with



a high pentosan content and at a competitive price are a current requisite. The effects of biogenic carbon are largely overshadowed by emissions associated with energy production.

Notably, the largest land use stems from both maize cultivation and steam production for heating. The latter becomes significant due to the high reactor consumption mentioned earlier, which drives intensive use of hardwood chips and coal despite their relatively low shares in the industrial mix (0.9% and 8.6%, respectively). However, around 50% of land use is due to the consumption of maize, which logically follows from cultivation activities. In addition, about 7 to 8% of the impact is related to the landfilling of the solid waste produced in both cases. An option to reduce this impact would be to avoid landfilling by recovering energy from the solid cake from the autohydrolysis reactor. This would be done at a lower energy cost than in the case of CBP, as the drying requirements are considerably lower. In addition, the combustion would be cleaner due to the absence of acids and furfural in this solid. However, this would mean a poor utilisation of this fraction as it contains cellulose and lignin with high purity. Purification and valorisation of both would lead to a much more favourable scenario. No disaggregated data are included as this purification system has not been designed —i.e. the impacts associated with its operation are not known— but a cursory analysis reveals that a net positive effect could be achieved for most indicators if the system is expanded to include these products within the envisaged functional unit.

The indicator of water consumption is dominated by the maize cultivation phase, as it accounts for 93% (HCP) and 88% (CBP) of the impact, with the remainder mainly due to steam generation equipment and extraction systems. As expected, this impact is associated in both cases with irrigation activities.


Waste production

Waste generation is not particularly relevant as compared to the consumption of energy and corncobs in terms of environmental impacts. However, it is considered a critical emission due to its severe impact on freshwater ecotoxicity in the case of CBP technology. As mentioned above, the treatment of solid waste generates significant land use. This solid cake is acidic in the case of CBP (see 2.4), although the main effect is related to the fate of the liquid fraction. The design of the conventional process includes a neutralisation stage between the two distillation columns to treat the remaining sulphuric acid. However, much of this acid is dissolved in the liquid fraction of the reactor stream. In addition, the first distillation column is designed to maximise the bottom recovery of acetic acid, so that it does not reach the neutraliser. Therefore, both waste streams contain a high acid load. In total, 0.26 kg of acetic acid and 0.19 kg of sulphuric acid are emitted for every kilogram of furfural produced (see Table S8). These streams are mixed and sent to wastewater treatment. However, if only anaerobic secondary treatment is used,¹⁴ part of the acid is not degraded and is released into the environment. In the case of CBP, the loss of some of the sulphuric and acetic acids to surface water accounts for more than 98% of the impact. This means that the potential impact reduction of the proposed process using a heterogeneous catalyst is greater than 99%. This is a key point as the emission of these quantities of acid to surface water makes the process unfeasible. In addition, the production of sulphuric acid also has a noticeable impact on acidification, mineral resource consumption and particulate emissions indicators.

4 CONCLUSIONS

This study presents a comprehensive techno-economic and environmental comparison of two processes for furfural production: the conventional process based on homogeneous catalysis with



Open Access Article. Published on 28 May 2025. Downloaded on 6/2/2025 7:02:40 PM.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.


View Article Online
DOI: 10.1039/D5SS000106D

sulphuric acid (CBP), and an alternative process employing heterogeneous catalysis (HCP). The latter was evaluated as a promising route for a more sustainable furfural production by addressing key process limitations, including reactor heating and acid usage. The HCP incorporates a two-stage reaction system, beginning with biomass auto-hydrolysis induced by temperature, followed by the conversion of released sugars into furfural using a stripping reactor with a heterogeneous catalyst. Both processes were simulated under comparable assumptions to ensure direct comparison.

The study demonstrated that replacing steam with nitrogen (N₂) as a stripping agent in the HCP is not viable, as it removes excessive water from the reaction medium, increasing tar formation and destabilizing the process. Furthermore, steam acts not only as a stripping agent but also as a heating fluid, making its replacement with nitrogen impractical from both a technical and economic standpoint.

For the heterogeneous catalyst used in the HCP process, Amberlyst® 70 resin was found to lack sufficient selectivity, resulting in unfeasible steam requirements for stripping. The catalyst selectivity would need to be 4.9 times higher to match the steam consumption of the conventional process. Moreover, the results indicate that the required overall selectivity must be 2.8 times higher than the one showed by H₂SO₄ in CBP process due to the lower sugar availability in the reaction medium. However, with a hypothetical catalyst meeting this selectivity threshold, the HCP process could become viable. The economic analysis revealed a similar minimum selling price (MSP) for furfural in both processes: €2,499/t for HCP and €2,685/t for CBP. Although the calculated MSP for the conventional process is relatively high, it should be noted that European energy prices were used in the calculations, and an important factor was not considered: the utilization of reactor residue. In existing processes, the biomass residue exiting the stripping reactor is typically dried and burned to generate heat for producing the stripping steam, which significantly reduces vapor-related costs. Furthermore, by leveraging the additional valorisation of lignin and unhydrolyzed cellulose streams in an integrated HCP biorefinery, the MSP could be reduced to €1,455/t, making it competitive with current market prices.

The environmental assessment identified energy consumption, maize cob usage, and waste emissions as key impact factors. Steam generation, particularly for reactor heating, remains the most critical aspect in both processes, and should be the focus of future developments. However, the proposed process significantly reduces steam consumption through enhanced energy integration, such as preheating via recycled streams and efficient reboiler configurations, which also generate a net environmental credit by producing low-pressure steam. Maize cob consumption is nearly identical in both processes and remains a major environmental concern due to the agrochemical inputs in maize cultivation, land use, and irrigation requirements. This highlights the need for alternative upstream solutions in the supply chain, particularly to mitigate the impacts of pesticides and fertilizers. Additionally, land use could be reduced through the recovery of solid waste, which is rich in lignin and high-purity cellulose. Since direct energy recovery would be a suboptimal use of these resources, future work should focus on developing a purification train to recover both fractions, thereby reducing environmental impacts while improving economic viability. The study also found that waste emissions, particularly acid loss in the CBP, severely impact freshwater ecotoxicity. The HCP mitigates this issue by eliminating sulphuric acid, achieving over a 99% reduction in this category using a heterogeneous catalyst.

Overall, the findings indicate that HCP process is technically and economically promising pathway for a sustainable furfural production, provided that sufficiently selective and stable heterogeneous catalyst can be developed. However, a key challenge remains: the rapid deactivation of Amberlyst® 70 due to the deposition of tar-like materials, which compromises its long-term viability. To enable an

industrially feasible process, a heterogeneous catalyst must not only exhibit a selectivity at least 4.9 times higher than Amberlyst® 70 and 2.8 times higher than H₂SO₄ but also demonstrate long-term stability under continuous operation. Given that organic resins such as Amberlyst® 70 are thermally unstable and cannot be regenerated by calcination, alternative solid acid catalysts—particularly inorganic materials—offer a more promising route due to their higher thermal and chemical stability, allowing effective regeneration strategies. A reactor configuration that integrates the catalyst regeneration would be advisable. Addressing these limitations will be crucial for HCP to achieve technical and economic parity with conventional sulphuric acid-based processes. Additionally, integrating this process within a biorefinery concept could further enhance its environmental sustainability by maximizing resource utilization and minimizing environmental impacts, paving the way for a competitive and environmentally viable furfural industry.

Abbreviations

AC	Acidification
GWP	Global Warming Potential
FWET	Freshwater Ecotoxicity
FDP	Fossil Depletion Potential
FWEP	Freshwater Eutrophication
MEP	Marine Eutrophication
TEP	Terrestrial Eutrophication
HTc	Human Toxicity, cancer
HTnc	Human Toxicity, non-cancer
IR	Ionising Radiation
LU	Land Use
MDP	Metals / Minerals Depletion Potential
OD	Ozone Depletion
PMF	Particulate Matter Formation
POF	Photochemical Oxidant Formation
WU	Water Use

CONFLICTS OF INTEREST

'There are no conflicts to declare'.

ACKNOWLEDGEMENTS

This work received financial support from the Spanish Ministry of Science and Innovation through Cat4BioMon Project (PID2021-122736OB-C43 & PID2021-122736OB-C44), being funded through MCIN/AEI/10.13039/501100011033/FEDER, UE.

REFERENCES

- 1 Climate Change and Health (CCH), *Quantitative risk assessment of the effects of climate change on selected causes of death, 2030s and 2050s*, 2024.
- 2 M. Höök and X. Tang, *Energy Policy*, 2013, **52**, 797–809.
- 3 H. Lee and J. Romero, *IPCC, 2023: Climate Change 2023: Synthesis Report. Contribution of Working Groups I, II and III to the Sixth Assessment Report of the Intergovernmental Panel on Climate Change*, Geneva, 2023.



- 4 L. Lange, K. O. Connor, S. Arason, U. Bundgård-Jørgensen, A. Canalis, D. Carrez, J. Gallagher, N. Gøtke, C. Huyghe, B. Jarry, P. Llorente, M. Marinova, L. O. Martins, P. Mengal, P. Paiano, C. Panoutsou, L. Rodrigues, D. B. Stengel, Y. van der Meer and H. Vieira, *Front Bioeng Biotechnol*, 2021, **8**, 1–16. View Article Online
DOI: 10.1039/D5SU00106D
- 5 A. Nabera, I. R. Istrate, A. J. Martín, J. Pérez-Ramírez and G. Guillén-Gosálbez, *Green Chemistry*, 2023, **25**, 6603–6611.
- 6 E. Baldoni, P. Reumerman, C. Parisi, R. Platt, H. Gonzalez Hermoso, K. Vikla, J. Vos and R. M'Barek, *Chemical and material driven biorefineries in the EU and beyond : database and dashboard visualisation*, Publications Office of the European Union, 2021.
- 7 European Commission, Chemical and material biorefineries in the EU | Knowledge for policy, https://knowledge4policy.ec.europa.eu/visualisation/chemical-material-biorefineries-eu_en, (accessed 27 January 2025).
- 8 J. J. Bozell and G. R. Petersen, *Green Chemistry*, 2010, **12**, 539–554.
- 9 Grand View Research, *Furfural Market Size, Share, Growth & Trends Report, 2030*, San Francisco, 2023.
- 10 D. S. S. Jorqueira, L. F. de Lima, S. F. Moya, L. Vilcocq, D. Richard, M. A. Fraga and R. S. Suppino, *Appl Catal A Gen*, 2023, **665**, 119360.
- 11 K. J. Yong, T. Y. Wu, C. B. T. L. Lee, Z. J. Lee, Q. Liu, J. M. Jahim, Q. Zhou and L. Zhang, *Biomass Bioenergy*, 2022, **161**, 106458.
- 12 W. Adhami, A. Richel and C. Len, *Molecular Catalysis*, 2023, **545**, 113178.
- 13 Mehdi Dashtban, Allan Gilbert and Pedram Fatehi, in *Journal of Science & Technology for Forest Products and Processes*, 2012, vol. 2, pp. 44–53.
- 14 K. J. Zeitsch, in *Sugar Series*, Elsevier, 2000, vol. 13, pp. 36–74.
- 15 H. E. Hoydonckx, W. M. Van Rhijn, W. Van Rhijn, D. E. De Vos and P. A. Jacobs, in *Ullmann's Encyclopedia of Industrial Chemistry*, John Wiley & Sons, Ltd, 2007.
- 16 A. K. Mathew, A. Abraham, K. K. Mallapureddy and R. K. Sukumaran, in *Waste Biorefinery: Potential and Perspectives*, Elsevier, 2018, pp. 267–297.
- 17 W. Kubic, Jr., X. Yang, C. Moore and A. Sutton, *A Process for Converting Corn Bran to Furfural without Mineral Acids*, Los Alamos, NM (United States), 2020.
- 18 M. Z. R. Mohammed, Z. W. Ng, A. Putranto, Z. Y. Kong, J. Sunarso, M. Aziz, S. H. Zein, J. Giwangkara and I. Butar, *Clean Technol Environ Policy*, 2023, **25**, 1551–1567.
- 19 D. T. Win, 2005.
- 20 L. Zhang, G. Xi, J. Zhang, H. Yu and X. Wang, *Bioresour Technol*, 2017, **224**, 656–661.
- 21 R. J. H. Grisel, J. C. Van Der Waal, E. De Jong and W. J. J. Huijgen, *Catal Today*, 2014, **223**, 3–10.
- 22 T. L. K. Yong, N. Mohamad and N. N. M. Yusof, *Procedia Eng*, 2016, **148**, 392–400.
- 23 S. Peleteiro, V. Santos, G. Garrote and J. C. Parajó, *Carbohydr Polym*, 2016, **146**, 20–25.



- 24 Y. Nie, Q. Hou, W. Li, C. Bai, X. Bai and M. Ju, *Molecules* 2019, Vol. 24, Page 594, 2019, **24**, 594. Open Access Article. Published on 28 May 2025. Downloaded on 6/2/2025 7:02:40 PM. This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence. DOI: 10.1039/D5SU000106D
- 25 C. B. T. L. Lee, T. Y. Wu, C. H. Ting, J. K. Tan, L. F. Siow, C. K. Cheng, J. Md. Jahim and A. W. Mohammad, *Bioresour Technol*, 2019, **278**, 486–489.
- 26 V. da Silva Lacerda, J. B. López-Sotelo, A. Correa-Guimarães, S. Hernández-Navarro, M. Sánchez-Bascones, L. M. Navas-Gracia, P. Martín-Ramos, E. Pérez-Lebeña and J. Martín-Gil, *Bioresour Technol*, 2015, **180**, 88–96.
- 27 C. Moreno-Marrodan, P. Barbaro, S. Caporali and F. Bossola, *ChemSusChem*, 2018, **11**, 3649–3660.
- 28 P. Bhaumik and P. L. Dhepe, *ChemCatChem*, 2017, **9**, 2709–2716.
- 29 R. K. Mishra, V. B. Kumar, A. Victor, I. N. Pulidindi and A. Gedanken, *Ultrason Sonochem*, 2019, **56**, 55–62.
- 30 S. M. Bruce, Z. Zong, A. Chatzidimitriou, L. E. Avci, J. Q. Bond, M. A. Carreon and S. G. Wettstein, *J Mol Catal A Chem*, 2016, **422**, 18–22.
- 31 Y. Wang, Y. Dai, T. Wang, M. Li, Y. Zhu and L. Zhang, *Fuel Processing Technology*, 2022, **237**, 107472.
- 32 W. Song, H. Liu, J. Zhang, Y. Sun and L. Peng, *ACS Catal*, 2022, **12**, 12833–12844.
- 33 A. Gupta, S. U. Nandanwar, P. Niphadkar, I. Simakova and V. Bokade, *Biomass Bioenergy*, 2020, **139**, 105646.
- 34 S. Le Guenic, D. Gergela, C. Ceballos, F. Delbecq and C. Len, *Molecules*, 2016, **21**, 1102.
- 35 A. Mittal, D. A. Ruddy, X. Chen and D. K. Johnson, *Energy and Fuels*, 2023, **37**, 13115–13125.
- 36 O. Sato, N. Mimura, Y. Masuda, M. Shirai and A. Yamaguchi, *J Supercrit Fluids*, 2019, **144**, 14–18.
- 37 X. Hu, R. J. M. Westerhof, D. Dong, L. Wu and C. Z. Li, *ACS Sustain Chem Eng*, 2014, **2**, 2562–2575.
- 38 T. Tongtummachat, A. Jaree and N. Akkarawatkhoosith, *RSC Adv*, 2022, **12**, 23366–23378.
- 39 J. K. Raman and E. Gnansounou, *Ind Crops Prod*, 2015, **69**, 371–377.
- 40 T. H. Kim, H. J. Ryu and K. K. Oh, *Bioresour Technol*, 2016, **218**, 367–372.
- 41 C. Li, D. Ding, Q. Xia, X. Liu and Y. Wang, *ChemSusChem*, 2016, **9**, 1712–1718.
- 42 X. Wang, C. Zhang, Q. Lin, B. Cheng, F. Kong, H. Li and J. Ren, *Ind Crops Prod*, 2018, **123**, 118–127.
- 43 S. Peleteiro, A. M. da Costa Lopes, G. Garrote, R. Bogel-Lukasik and J. C. Parajó, *Ind Crops Prod*, 2015, **77**, 163–166.
- 44 Z. Xu, W. Li, Z. Du, H. Wu, H. Jameel, H. min Chang and L. Ma, *Bioresour Technol*, 2015, **198**, 764–771.

- 45 P. Liu, S. Shi, L. Gao and G. Xiao, *Reaction Kinetics, Mechanisms and Catalysis*, 2022, **135**, 795–810. DOI:10.1039/D5SC000106D
- 46 H. Lin, J. Chen, Y. Zhao and S. Wang, *Energy and Fuels*, 2017, **31**, 3929–3934.
- 47 D. M. Alonso, S. H. Hakim, S. Zhou, W. Won, O. Hosseinaei, J. Tao, V. Garcia-Negron, A. H. Motagamwala, M. A. Mellmer, K. Huang, C. J. Houtman, N. Labbé, D. P. Harper, C. T. Maravelias, T. Runge and J. A. Dumesic, *Sci Adv*, 2017, **3**, 1–7.
- 48 D. Edumujeze, M. C. Fournier-Salaün and S. Leveneur, *Fuel*, 2025, **381**, 133423.
- 49 G. Zang, A. Shah and C. Wan, *J Clean Prod*, 2020, **260**, 120837.
- 50 M. S. Hossain, C. Theodoropoulos and A. Yousuf, *Biochem Eng J*, 2019, **144**, 89–103.
- 51 S. Farzad, M. A. Mandegari, M. Guo, K. F. Haigh, N. Shah and J. F. Görgens, *Biotechnology for Biofuels* 2017 10:1, 2017, **10**, 1–24.
- 52 I. Agirrezabal-Telleria, J. Requies, M. B. Güemez and P. L. Arias, *Green Chemistry*, 2012, **14**, 3132–3140.
- 53 I. Agirrezabal-Telleria, A. Larreategui, J. Requies, M. B. Güemez and P. L. Arias, *Bioresour Technol*, 2011, **102**, 7478–7485.
- 54 J.-P. Lange, E. van der Heide, J. van Buijtenen and R. Price, *ChemSusChem*, 2012, **5**, 150–166.
- 55 J. Slak, B. Pomeroy, A. Kostyniuk, M. Grilc and B. Likozar, *Chemical Engineering Journal*, 2022, **429**, 132325.
- 56 R. J. Wooley and V. Putsche, *Development of an ASPEN PLUS Physical Property Database for Biofuels Components*, Golden (Colorado, USA), 1996.
- 57 N. Bhandari, D. G. Macdonald and N. N. Bakhshi, *Biotechnol Bioeng*, 1984, **26**, 320–327.
- 58 V. Krzelj, J. Ferreira Liberal, M. Papaioannou, J. van der Schaaf and M. F. Neira d'Angelo, *Ind Eng Chem Res*, 2020, **59**, 11991–12003.
- 59 D. Nabarlantz, X. Farriol and D. Montané, *Ind Eng Chem Res*, 2004, **43**, 4124–4131.
- 60 W. L. Luyben, *Distillation design and control using aspen simulations*, John Wiley & Sons, Hoboken, New Jersey, 2nd Editio., 2013.
- 61 F. K. Kazi, A. D. Patel, J. C. Serrano-Ruiz, J. A. Dumesic and R. P. Anex, *Chemical Engineering Journal*, 2011, **169**, 329–338.
- 62 R. Turton, R. C. Bailie, W. B. Whiting, J. A. Shaeiwitz and D. Bhattacharyya, *Analysis, Synthesis, and Design of Chemical Processes*, Prentice Hall, Upper Saddle River, New Jersey, Fourth Edi., 2012.
- 63 M. S. Peters, K. D. Timmerhaus and R. E. West, *Plant Design and Economics for Chemical Engineers*, McGraw-Hill, New York, 5th Ed., 2003.
- 64 H. Zhang, L. Hou, Y. Lin, X. Liu, S. Zhao, C. Xu and C. Chang, *Bioresour Technol*, 2025, **418**, 131897.



- 65 TNO Biobased and Circular Technologies, Phyllis2 Database. View Article Online
DOI: 10.1039/D5SU00106D
- 66 ChemAnalyst, <https://www.chemanalyst.com/Pricing/Searchpricing>.
- 67 <https://straitresearch.com/report/cellulose-fiber-market>.
- 68 A. J. Robinson, A. Giuliano, O. Y. Abdelaziz, C. P. Hultberg, A. Koutinas, K. S. Triantafyllidis, D. Barletta and I. De Bari, *Bioresour Technol*, 2022, **364**, 128004.
- 69 B. Steubing, D. de Koning, A. Haas and C. L. Mutel, *Software Impacts*, 2020, **3**, 100012.
- 70 S. Andreasi Bassi, F. Biganzoli, N. Ferrara, A. Amadei, A. Valente, S. Sala and F. Ardente, *Updated characterisation and normalisation factors for the Environmental Footprint 3.1 method*, Publications Office of the European Union, Luxembourg (Luxembourg), 2023.
- 71 A. Ciroth, S. Muller, B. Weidema and P. Lesage, *Int J Life Cycle Assess*, 2016, **21**, 1338–1348.
- 72 G. Sampath and S. Kannan, *Catal Commun*, 2013, **37**, 41–44.
- 73 X. Teng, Z. Si, S. Li, Y. Yang, Z. Wang, G. Li, J. Zhao, D. Cai and P. Qin, *Ind Crops Prod*, 2020, **151**, 112481.
- 74 J. Gao, H. Wang, X. Cao, Z. Li, H. Guo, X. Yang, W. Wang, N. Guo and Y. Ma, *Molecular Catalysis*, 2023, **535**, 112890.
- 75 Y. Lyu, M. Raugei, X. Zhang, S. Mellino and S. Ulgiati, *Renewable and Sustainable Energy Reviews*, 2021, **151**, 111604.



Data availability statement

[View Article Online](#)
DOI: 10.1039/D5SU00106D

The data supporting this article have been included as part of the Supplementary Information.

