

ChemComm

Accepted Manuscript



This is an *Accepted Manuscript*, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this *Accepted Manuscript* with the edited and formatted *Advance Article* as soon as it is available.

You can find more information about *Accepted Manuscripts* in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this *Accepted Manuscript* or any consequences arising from the use of any information it contains.

COMMUNICATION

Synthesis and Vibrational Spectroscopy of ^{57}Fe -Labeled Models of [NiFe] Hydrogenase: First Direct Observation of a Nickel-Iron Interaction

Cite this: DOI: 10.1039/x0xx00000x

Received 00th June 2014,
Accepted xxth xxx 2014

DOI: 10.1039/x0xx00000x

www.rsc.org/chemcomm

David Schilter,^a Vladimir Pelmeshnikov,^b Hongxin Wang,^{c,d} Florian Meier,^b Leland B. Gee,^c Yoshitaka Yoda,^e Martin Kaupp,^b Thomas B. Rauchfuss,^a and Stephen P. Cramer^{c,d}

A new route to iron carbonyls has enabled synthesis of ^{57}Fe -labeled [NiFe] hydrogenase mimic $(\text{OC})_3^{57}\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})$. Its study by nuclear resonance vibrational spectroscopy revealed Ni- ^{57}Fe vibrations, as confirmed by calculations. The modes are absent for $[(\text{OC})_3^{57}\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})]^+$, which lacks Ni- ^{57}Fe bonding, underscoring the utility of the analyses in identifying metal-metal interactions.

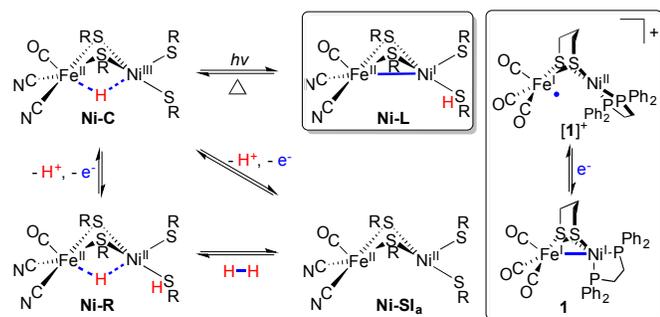
Despite our extremely low atmospheric concentration of dihydrogen (~ 1 ppm), this substrate is a key metabolite of many anaerobic bacteria.¹ In such living systems can be found the most prevalent enzymes for hydrogen processing, the nickel-iron hydrogenases ([NiFe]-H₂ases).^{1,2} These electrocatalysts specifically mediate the redox reaction $\text{H}_2 \rightleftharpoons 2\text{H}^+ + 2\text{e}^-$ at several hundred turnovers per second.³ Their heterobimetallic active sites exist in several states, some of which are summarized below (Fig. 1, left and centre).

The active sites feature Ni bound to four cysteinato residues, two of which bridge to an $\text{Fe}(\text{CO})(\text{CN})_2$ fragment. In the Ni-C state, Ni(III)Fe(II) centres bind a bridging hydride (H^-), reductive elimination of which affords Ni-L.⁴ Thus, H^+ is abstracted by a terminal cys ligand (Fig. 1, centre top), leaving a Ni(I)Fe(II) core with a 2e^- bond between the metals.⁴

The use of vibrational spectroscopy to study [NiFe]-H₂ase is convenient in that its active site features chromophores easily identifiable by such techniques. Spectral analyses are often aided by comparison to data from synthetic models⁵ whose structures are well understood. Specificity for ^{57}Fe -coupled modes is afforded by nuclear resonance vibrational spectroscopy (NRVS, *vide infra*). This has recently enabled observation of characteristic Fe-CN/Fe-CO bending and stretching modes in [NiFe]-H₂ase (Ni-A and Ni-R) and the Fe subsite model $[\text{Fe}(\text{benzenedithiolato})(\text{CN})_2\text{CO}]^{2-}$.⁶ However, no NRVS studies have reported on metal-metal bonding, which is expected for low-valent clusters like Ni-L and its models.^{4,7}

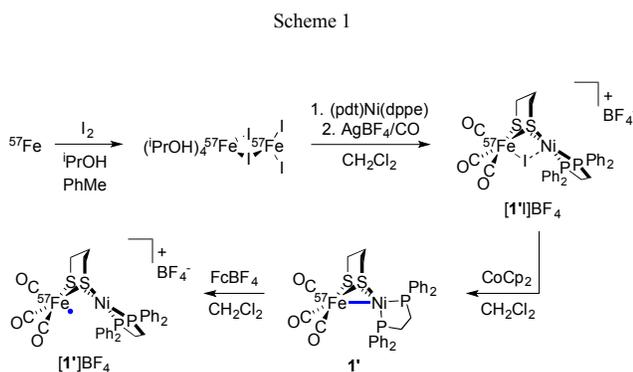
A near-complete [NiFe]-H₂ase mimic is the Ni(II)Fe(I) species $[(\text{OC})_3\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})]^+$ ($[1]^+$, $\text{pdt}^{2-} = ^-\text{S}(\text{CH}_2)_3\text{S}^-$; $\text{dppe} = 1,2$ -bis(diphenylphosphino)ethane), a model for Ni-L, albeit with metal oxidation states reversed (Fig. 1, right).⁸ This $S = 1/2$ model is prepared from $(\text{OC})_3\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})$ (**1**),⁹ itself the subject of density functional theory (DFT) and resonance Raman investigations.¹⁰ Disclosed here is methodology for ^{57}Fe -labeled prototypes $[(\text{OC})_3^{57}\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})]^{0/+}$ ($[1]^{0/+}$), enabling the study of metal-metal bonding with NRVS.

The Ni(I)Fe(I) complex **1** is usually accessed by interaction of $(\text{pdt})\text{Ni}(\text{dppe})$ with an Fe carbonyl such as $\text{Fe}_2(\text{CO})_9$ or $\text{Fe}(\text{CO})_4\text{I}_2$.¹¹⁻¹⁴ The precursor to these, $\text{Fe}(\text{CO})_5$, is not conveniently prepared from elemental Fe, a factor that

Fig. 1 Key [NiFe]-H₂ase states (left and centre) and two model complexes (right).

necessitated a new route adaptable to ^{57}Fe incorporation. Thus, metallic Fe was converted to the organo-soluble FeI_2 source $\text{Fe}_2\text{I}_4(\text{iPrOH})_4$,¹⁵ which, upon combination with $(\text{pdt})\text{Ni}(\text{dppe})$, gave the known diiodide $\text{I}_2\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})$.¹⁴ While the diiodide does not bind CO in CH_2Cl_2 , when treated with AgBF_4 it converts to the putative electrophile $[\text{IFe}(\text{pdt})\text{Ni}(\text{dppe})]^+$ (or perhaps its dimer), which undergoes carbonylation to afford $[(\text{OC})_3\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})]^+$ ($[\mathbf{1}]^+$).¹⁴ Reduction with CoCp_2 gave $\mathbf{1}$ in yields comparable to the $(\text{pdt})\text{Ni}(\text{dppe})/\text{Fe}_2(\text{CO})_9$ route.¹¹

In adapting the synthesis to the Ni^{57}Fe target, elemental ^{57}Fe was oxidized to $^{57}\text{Fe}_2\text{I}_4(\text{iPrOH})_4$,¹² which was converted into violet $[(\text{OC})_3^{57}\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})]^+$ (ESI-MS: m/z 829.5 $[\mathbf{1}]^+$ vs 828.9 $[\mathbf{1}]^+$), with reduction affording green $(\text{OC})_3^{57}\text{Fe}(\text{pdt})\text{Ni}(\text{dppe})$ ($\mathbf{1}'$) in 35% yield (Scheme 1).



While ν_{CO} energies and analytical and $^{31}\text{P}\{^1\text{H}\}$ NMR data for $\mathbf{1}'$ are virtually identical to those of natural abundance $\mathbf{1}$, LI-FDI-MS (liquid introduction field desorption ionization mass spectrometry) analyses are telling. Soft ionization of non-polar $\mathbf{1}'$ and $\mathbf{1}$ allowed for detection of parent cations (m/z 702.9 and 701.9, respectively). Oxidation of $\mathbf{1}'$ with FcBF_4 afforded mixed-valent salt $[\mathbf{1}']\text{BF}_4$, whose EPR signal is broadened relative to that of $[\mathbf{1}]\text{BF}_4$ due to hyperfine interactions.

Investigations into Ni–Fe bonding in the new ^{57}Fe -labeled variants of the reduced and oxidized complexes (respectively $\mathbf{1}'$ and $[\mathbf{1}']^+$) were undertaken using NRVS. This technique, enabled by the development of third generation synchrotron sources, insertion devices and advanced X-ray optics,^{16–18} involves scanning an extremely monochromatic (*e.g.* 0.8 meV) X-ray beam through a nuclear resonance of a Mössbauer-active isotope (*e.g.* 14.4 keV for ^{57}Fe). Subsequent relaxation causes the generation/annihilation of phonons, the detection of which reveals all modes in which the ^{57}Fe nucleus moves along the direction of the incident X-ray. NRVS has several advantages over traditional IR and resonance Raman spectroscopies,¹⁹ not least in terms of element and isotope specificity and absence of the optical selection rules, which have allowed for the resolution of $\nu_{\text{Fe-X}}$ ($X = \text{S}, \text{P}, \text{Cl}, \text{CO}, \text{CN}, \text{NO}$) vibrations in complicated systems.^{6,13,19,20} This relatively new but powerful technique in inorganic and biological iron chemistry is applied here to $[\mathbf{1}']^{0/+}$. The analysed spectra in terms of ^{57}Fe partial vibrational density of states (PVDOS) are given in Fig. 2.

Intense bands assigned to $\nu_{\text{Fe-CO}}$ and $\delta_{\text{Fe-CO}}$ modes were observed at 440–630 cm^{-1} , with full-range NRVS spectra and Fe–C(O) kinetic energy distribution (KED) diagrams presented in Fig. S9. Similar signals were found for $[\text{NiFe}]\text{-H}_2\text{ase}$.⁶ It was expected that $\nu_{\text{Fe-S}}$, $\delta_{\text{Fe-S}}$ and $\nu_{\text{Fe-Ni}}$ bands, if observable, would lie at low energies ($\leq 400 \text{ cm}^{-1}$). Upon comparing data for $\mathbf{1}'$ and $[\mathbf{1}']^+$ (Fig. 2a), a sharp and prominent NRVS peak at 158 cm^{-1} for $\mathbf{1}'$ was noticed. This band, absent from the spectra of $[\mathbf{1}']^+$, was tentatively ascribed to vibration of the Ni–Fe bond, such interactions not being significant in $[\mathbf{1}']^+$. Differences in NRVS data of $\mathbf{1}'$ and $[\mathbf{1}']^+$ were less marked in other spectral regions, although peaks for $[\mathbf{1}']^+$ were broader.

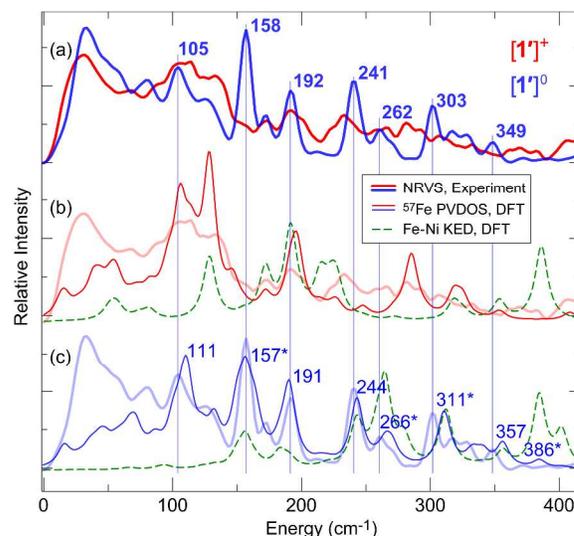


Fig. 2 Observed NRVS spectra for $[\mathbf{1}']^+$ (thick red lines, (a) and (b)) and $\mathbf{1}'$ (thick blue lines, (a) and (c)) vs. DFT calculated ^{57}Fe PVDOS spectra for $[\mathbf{1}']^+$ (thin red line, (b)) and $\mathbf{1}'$ (thin blue line, (c)). Calculated Fe–Ni KED (green) is given in (b) for $[\mathbf{1}']^+$, and in (c) for $\mathbf{1}'$. Key bands observed for $\mathbf{1}'$ are labelled in (a), with DFT counterparts in (c) indicated by vertical lines. Modes giving rise to bands with significant Fe–Ni character are marked (*) and shown in Fig. 3 and Fig. S10. For 0–650 cm^{-1} spectra, see Fig. S9.

The assignment of vibrational bands was elaborated using DFT calculations on $[\mathbf{1}']^{0/+}$ as detailed in the SI; simulated NRVS (^{57}Fe PVDOS) and Fe–Ni KED diagrams for $[\mathbf{1}']^+$ and $\mathbf{1}'$ are also presented in Fig. 2b and Fig. 2c, respectively. Band positions and intensities in the calculated and observed spectra are largely in agreement (particularly in Fig. 2c; see also Fig. S9). In the case of $[\mathbf{1}']^+$, some differences are assigned to impurities ($\mathbf{1}'$ and/or $[\mathbf{1}'\text{H}]^+$). The band for $\mathbf{1}'$ at 158 cm^{-1} , calculated by DFT at 157 cm^{-1} , indeed involves stretching of the Fe–Ni bond symmetrical to a Ni–P1 stretch (see Fig. 3, and the SI for animations). Such a vibrational coupling in the Fe–Ni–P1 triad implicates a strong Fe–Ni interaction, and it has no complement in the normal modes pool calculated for $[\mathbf{1}']^+$.

While very prominent in NRVS (owing to the large ^{57}Fe displacement), the 157 cm^{-1} band is weaker (~5%) in the Ni–Fe KED diagram, which reflects *relative* motion of Fe and Ni. Other Ni–Fe stretches, such as those calculated at 266, 311, and 386 cm^{-1} (see in Fig. 2c), are considerably stronger (representing

13%, 9%, and 11% total Ni–Fe KED, respectively). Yet only the first two modes can be associated with bands observed at 262 and 303 cm^{-1} , while the last one has vanishing NRVS intensity (Fig. 2a). The lower ^{57}Fe PVDOS intensity for the latter modes at 266, 311, and 386 cm^{-1} is accounted for by the fact that they involve displacement mostly of Ni (rather than Fe), this movement being evident from the DFT results (Fig. S10). Analysis of Fe–Ni vibrations is further complicated by vibrational coupling to C/P/S atoms, in particular the bridging S donors. Notably, mixed $\nu_{\text{Fe-Ni}}$ modes in the 220–360 cm^{-1} region involving up to 14% contribution from the Fe–Ni stretch were also reported for **1**.¹⁰

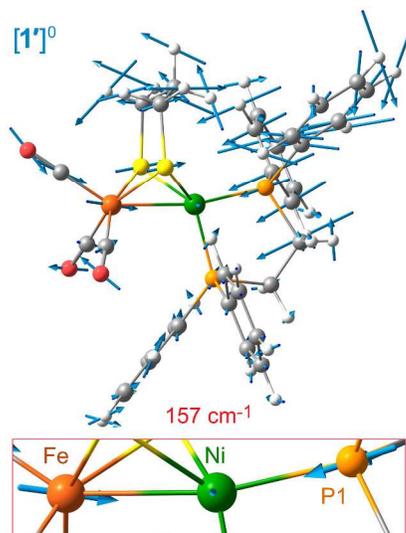


Fig. 3 Scaled arrow depiction of nuclear displacements for the normal mode calculated for **1*** at 157 cm^{-1} (a symmetric Fe–Ni–P1 stretch, see corresponding ^{57}Fe PVDOS band in Fig. 2c). Key $[\mathbf{1}]^{0/+}$ modes are animated in the SI.

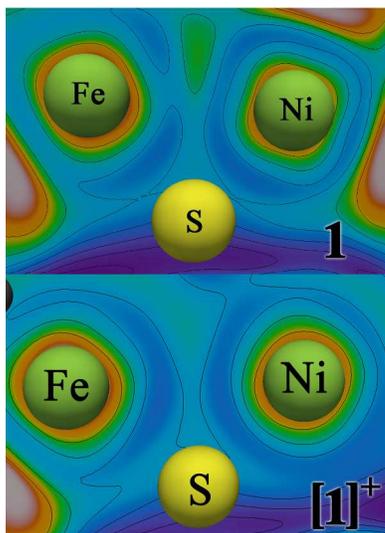


Fig. 4 Electron localization function (ELF) analysis of the Ni–Fe bonding in $[\mathbf{1}]^{0/+}$ (top/bottom). Ni–Fe bond attractor position for $[\mathbf{1}]^0$ is indicated by the localized area in green (center-top), absent for $[\mathbf{1}]^+$. See Figs. S11 and S12 for alternative bonding representations.

Structural (X-ray diffraction) and DFT studies on unlabeled $[\mathbf{1}]^{0/+}$ have suggested that the reduced Ni(I)Fe(I) species may feature Ni–Fe bonding,⁹ while the oxidized Ni(II)Fe(I) does not.⁸ This was confirmed by re-analysing bonding in $[\mathbf{1}]^{0/+}$ using ELF,²¹ ELI-D,²² and QTAIM²³ electron density-based methods (**Error! Reference source not found.**, S11 and S12), all of which indicated a Ni–Fe bond in **1**. Notably, the ELF/ELI-D bond attractor and the bond critical point found by QTAIM are shifted from the Ni–Fe vector away from the bridging S atoms, such that Fe(I) is (pseudo)octahedral. Bonding involves overlap of two singly-occupied $d(z^2)$ orbitals and is absent when Ni is oxidized, with $[\mathbf{1}]^+$ featuring a Fe-localized $d(z^2)$ SOMO. In contrast, Ni–L has a SOMO with Ni $d(z^2)$ and $d(x^2 - y^2)$ character,⁴ and a $2e^-$ Ni→Fe dative bond (optimized Fe–Ni distance: 2.47 Å).⁴ While $[\mathbf{1}]^+$ and Ni–L have the same electron count for the $[\text{NiFe}]^{3+}$ site, only one has Ni–Fe bonding. A key distinction is Ni geometry, which is planar in $[\mathbf{1}]^+$ but SF_4 -like in Ni–L.

Since the Fe(I) centre in $[\mathbf{1}]^+$ is electron poor (average $\nu_{\text{CO}} = 2010 \text{ cm}^{-1}$)⁸ it cannot supply $2e^-$ for a Fe→Ni coordinate bond. Another possibility, a covalent Ni–Fe bond, is unlikely due to the low donicity of Ni(II). The metal centres in $[\mathbf{1}]^+$ are distant (our DFT result: 2.80 Å), while those in electron-rich **1** (average $\nu_{\text{CO}} = 1977 \text{ cm}^{-1}$) are sufficiently proximal (experimental: 2.47 Å,⁹ DFT: 2.46 Å) for covalent Fe–Ni bonding (see Table S2 and Fig. S3 for structural and IR details).

Obtaining direct evidence of metal–metal bonding in molecular systems is nontrivial, and distances do not guarantee presence/absence of bonding. For example, EXAFS studies reveal similar Fe–Ni distances for Ni–L and Ni–C,²⁴ despite Ni not being bonded to Fe in the latter. This highlights that (i) $[\text{NiFe}]$ -H₂ase active site ligation is inflexible (relative to $[\mathbf{1}]^{0/+}$) and that (ii) EXAFS/XRD studies typically only afford nuclear positions via core electron density, with bonding electron density between metals being poorly resolved.²⁵ In rare cases when very high quality single crystal data are obtained, multipole analysis, followed by topological analysis of static electron density (versus inspection of difference maps) can give insights into metal–metal bonding, as exemplified by (re)investigation of the archetypal $\text{Mn}_2(\text{CO})_{10}$.²⁶

Metal–metal bonding is common in organoiron chemistry, but it also plays a role in the reduced states of some metalloenzymes, stabilizing low-valent metal centres poised for substrate activation.⁷ In addition to $[\text{NiFe}]$ -H₂ase, Ni–Fe interactions are also proposed for carbon monoxide dehydrogenase (CODH),^{7,27} with Ni–Ni and Fe–Fe bonds being present in acetyl-CoA synthetase (ACS)⁷ and $[\text{FeFe}]$ -H₂ase,²⁸ respectively. NRVS is demonstrably effective in unambiguous identification of low energy ^{57}Fe -coupled modes.^{13,20,29} Compared to IR and resonance Raman, it avoids interference from solvent and ‘fingerprint’ bands, enabling identification of Fe-coupled vibrations, such as the 158 cm^{-1} $\nu_{\text{Fe-Ni}}$ mode here. The use of NRVS to probe iron-metal modes in everything from small molecules to iron enzymes is anticipated to provide a wealth of information on these catalysts.

Conclusions

While primordial routes to iron carbonyls have been reported (e.g. the one-pot preparation of $(\text{CO})_3\text{Fe}(\text{pdt})\text{Fe}(\text{CO})_3$ from FeCl_2),³⁰ they are typically limited in scope and reproducibility. In contrast, the $\text{Fe}_2\text{I}_4(\text{PrOH})_4/\text{CO}/\text{CoCp}_2$ strategy will likely be generalizable and afford iron carbonyls of relevance to H_2 ases and organoiron chemistry in general. With the isolation of ^{57}Fe -labeled $[\text{NiFe}]\text{-H}_2\text{ase}$ mimics $[\mathbf{1}']^{0/+}$, the element and isotope selectivity of NRVS was exploited to obtain the first vibrational spectroscopic evidence of Fe–Ni interactions. These results serve as an important benchmark, opening the door to work probing metal-metal bonding in redox-active H_2ase enzymes as well as related enzymatic and model systems.

Thanks are given to Drs Mark J. Nilges and Haijun Yao for assistance with EPR and LI-FDI-MS, respectively. Financial support was provided by the National Institutes of Health (GM061153-10 to T.B.R. and GM-65440 to S.P.C.), U.S. Department of Energy Office of Biological and Environmental Research (DOE OBER) (S.P.C.), and the ‘Unifying Concepts in Catalysis’ initiative of the German Research Council (V.P.). NRVS experiments performed at SPring-8 BL09XU were funded by JASRI (beamtime proposal 2013A0032).

Notes and references

^aDepartment of Chemistry, University of Illinois at Urbana-Champaign, Urbana, IL 61801, USA. E-mail: schilter@illinois.edu

^bInstitut für Chemie, Technische Universität Berlin, 10623 Berlin, Germany

^cDepartment of Chemistry, University of California, Davis, CA 95616, USA

^dPhysical Biosciences Division, Lawrence Berkeley National Laboratory, Berkeley, CA 94720, USA

^eJASRI, SPring-8, Sayo-gun, Hyogo 679-5198, Japan

†Electronic Supplementary Information (ESI) available: experimental procedures, spectral data, computational chemistry details, animated vibrational modes as GIFs. See DOI: 10.1039/c000000x/

- P. M. Vignais and B. Billoud, *Chem. Rev.*, 2007, **107**, 4206.
- H. Ogata, W. Lubitz, Y. Higuchi, *Dalton Trans.*, 2009, 7577.
- M. Frey, *ChemBioChem*, 2002, **3**, 153.
- M. Kampa, M.-E. Pandelia, W. Lubitz, M. van Gastel, F. Neese, *J. Am. Chem. Soc.*, 2013, **135**, 3915.
- S. Kaur-Ghumaan, M. Stein, *Dalton Trans.*, 2014, **43**, 9392.
- S. Kamali, H. Wang, D. Mitra, H. Ogata, W. Lubitz, B. C. Manor, T. B. Rauchfuss, D. Byrne, V. Bonnefoy, F. E. Jenney, M. W. W. Adams, Y. Yoda, E. Alp, J. Zhao, S. P. Cramer, *Angew. Chem. Int. Ed.*, 2013, **52**, 724.
- P. A. Lindahl, *J. Inorg. Biochem.*, 2012, **106**, 172.
- D. Schilter, M. J. Nilges, M. Chakrabarti, P. A. Lindahl, T. B. Rauchfuss, M. Stein, *Inorg. Chem.*, 2012, **51**, 2338.
- W. Zhu, W. A. C. Marr, Q. Wang, F. Neese, D. J. E. Spencer, A. J. Blake, P. A. Cooke, C. Wilson, M. Schröder, *Proc. Nat. Acad. Sci. U.S.A.*, 2005, **102**, 18280; M. T. Huynh, D. Schilter, S. Hammes-Schiffer, T. B. Rauchfuss, *J. Am. Chem. Soc.*, **2014**, DOI: 10.1021/ja505783z
- H. S. Shafaat, K. Weber, T. Petrenko, F. Neese, W. Lubitz, *Inorg. Chem.*, 2012, **51**, 11787.
- B. E. Barton, T. B. Rauchfuss, *J. Am. Chem. Soc.*, 2010, **132**, 14877.
- M. E. Carroll, J. Chen, D. E. Gray, J. C. Lansing, T. B. Rauchfuss, D. Schilter, P. I. Volkers, S. R. Wilson, *Organometallics*, 2014, **33**, 858.
- Y. Guo, H. Wang, Y. Xiao, S. Vogt, R. K. Thauer, S. Shima, P. I. Volkers, T. B. Rauchfuss, V. Pelmenschikov, D. A. Case, E. Alp, W. Sturhahn, Y. Yoda, S. P. Cramer, *Inorg. Chem.*, 2008, **47**, 3969.
- D. Schilter, T. B. Rauchfuss, *Dalton Trans.*, 2012, **41**, 13324.
- G. G. Nunes, R. C. R. Bottini, D. M. Reis, P. H. C. Camargo, D. J. Evans, P. B. Hitchcock, G. J. Leigh, E. L. Sá, J. F. Soares, *Inorg. Chim. Acta*, 2004, **357**, 1219.
- M. Seto, Y. Yoda, S. Kikuta, X. W. Zhang, M. Ando, *Phys. Rev. Lett.*, 1995, **74**, 3828.
- W. Sturhahn, T. S. Toellner, E. E. Alp, X. Zhang, M. Ando, Y. Yoda, S. Kikuta, M. Seto, C. W. Kimball, B. Dabrowski, *Phys. Rev. Lett.*, 1995, **74**, 3832.
- Y. Yoda, Y. Imai, H. Kobayashi, S. Goto, K. Takeshita, M. Seto, *Hyperfine Interact.*, 2012, **206**, 83.
- M. C. Smith, Y. Xiao, H. Wang, S. J. George, D. Coucouvanis, M. Koutmos, W. Sturhahn, E. E. Alp, J. Zhao, S. P. Cramer, *Inorg. Chem.*, 2005, **44**, 5562.
- Z. J. Tonzetich, H. Wang, D. Mitra, C. E. Tinberg, L. H. Do, F. E. Jenney, Jr., M. W. W. Adams, S. P. Cramer, S. J. Lippard, *J. Am. Chem. Soc.*, 2010, **132**, 6914.
- A. D. Becke, K. E. Edgecombe, *J. Chem. Phys.* **1990**, *92*, 5397; A. Savin, O. Jepsen, J. Flad, O. K. Andersen, H. Preuss, H. G. Vonschnering, *Angew. Chem. Int. Ed.* **1992**, *31*, 187; M. Kohout, A. Savin, *Int. J. Quantum. Chem.* **1996**, *60*, 875.
- M. Kohout, *Int. J. Quantum. Chem.* **2004**, *97*, 651; M. Kohout, K. Pernal, F. R. Wagner, Y. Grin, *Theor. Chem. Acc.* **2004**, *112*, 453.
- R. F. W. Bader, *Atoms in Molecules: A Quantum Theory*, Oxford University Press, Oxford, **1990**.
- J. P. Whitehead, R. J. Gurbel, C. Bagyinka, B. M. Hoffman, M. J. Maroney, *J. Am. Chem. Soc.*, 1993, **115**, 5629.
- T. S. Koritsanszky, P. Coppens, *Chem. Rev.*, 2001, **101**, 1583.
- R. Bianchi, G. Gervasio, D. Marabello, *Inorg. Chem.*, 2000, **39**, 2360.
- J.-H. Jeoung, H. Dobbek, *J. Am. Chem. Soc.*, 2009, **131**, 9922.
- Y. Nicolet, A. L. de Lacey, X. Vernède, V. M. Fernandez, E. C. Hatchikian, J. C. Fontecilla-Camps, *J. Am. Chem. Soc.*, 2001, **123**, 1596.
- S. P. Cramer, Y. Xiao, H. Wang, Y. Guo, M. C. Smith, *Hyperfine Interact.*, 2006, **170**, 47.
- P. I. Volkers, C. A. Boyke, J. Chen, T. B. Rauchfuss, C. M. Whaley, S. R. Wilson, H. Yao, *Inorg. Chem.*, 2008, **47**, 7002.

