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COMMUNICATION

Bypassing a Highly Unstable Frustrated Lewis Pair: Dihydrogen Cleavage by a Thermally Robust Silylium-Phosphine Adduct

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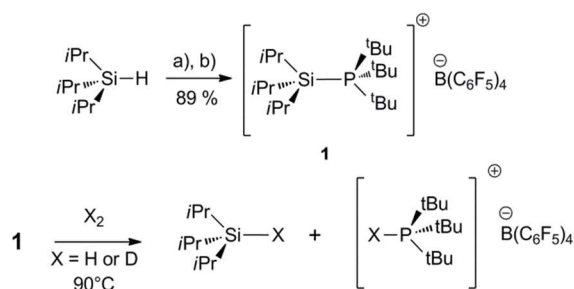
The thermally robust silylium complex $[iPr_3Si-PtBu_3]^+[B(C_6F_5)_4]^-$ (1**) activates H_2/D_2 at $90^\circ C$ (PhCl); no evidence for dissociation into the separated Lewis pair is found. DFT calculations show H_2 cleavage proceeds *via* Si–P bond elongation to form an encounter complex directly from the adduct, thus avoiding the non-isolable $iPr_3Si^+/PtBu_3$ FLP.**

Since their discovery in 2006, ‘frustrated Lewis pairs’ (FLPs) have continued to provide a fresh and novel approach to the field of bond activation chemistry.¹ FLPs can be defined as combinations of a Lewis acid and base which, by dint of steric hindrance, are unable to datively bind in the classical manner, leading to unquenched reactivity capable of activating small molecules. One of the most exciting properties of FLPs is their ability to induce heterolytic cleavage of H_2 into protic and hydridic components (H^+/H^-),² which can subsequently be transferred to reducible substrates either stoichiometrically (e.g. carbonyls, CO_2)³ or catalytically (e.g. imines, silyl enol ethers, alkenes).⁴ By far the most commonly used Lewis acids are organometallics of the Group 13 elements (e.g. R_3E ; $E = B$ or Al , $R =$ electron-deficient organyl), among which $B(C_6F_5)_3$ remains prevalent. Silylium ions (R_3Si^+) are both isoelectronic and isolobal with this class of compounds and, because of their potent electrophilicity,⁵ can demonstrate similar FLP reactivity when employed as the Lewis acid partner. However, there are few reports of stable R_3Si^+ cations, which is attributed to the difficulties in stabilising the diffuse $3p$ Si valence orbital either through hyper- or π -conjugation.⁶ This property explains, in part, their voracious affinity for nucleophiles; even traditionally ‘inert’ arene solvents⁷ and weakly coordinating anions (WCAs)⁸ can demonstrate interactions with the silicon centre. The work of Ozerov *et al.* reflects this extreme appetite for σ and π donors, in which decomposition of the robust $[B(C_6F_5)_4]^-$ WCA was observed when heated in the presence of Et_3Si^+ ,⁹ presumably *via* $[C_6F_5]^-$ abstraction or C–F activation.¹⁰ Such

indiscriminate reactivity may be suppressed *via* incorporation of intramolecular donors which moderate the potent electrophilicity while preserving sufficient reactivity for catalysis. In this respect, Oestreich and co-workers used the ferrocene-stabilised silylium ion $[(C_5H_5)Fe(C_5H_4Si/BuMe)]^+$ to selectively and catalytically hydrosilylate various ketones to the corresponding alkyl/silyl ethers ($RR'CH-O-SiR_3$; $R/R' =$ alkyl);¹¹ by contrast the fully reduced alkanes $R-CH_2-R'$ ¹² can be isolated when the ‘untamed’ $[R_3Si]^+$ is used as catalyst. Steric protection of the Si centre has also been adopted to prevent undesired silylation of aromatic solvents, as documented by Müller *et al.* Here, bulky triarylsilylium compounds (Ar_3Si^+ ; $Ar = C_6H_{6-x}Me_x$, $x = 3-5$), in conjunction with phosphine (R_3P ; $R =$ alkyl, aryl) or silylene bases,¹³ are shown to generate FLPs that engage in H_2 cleavage. Nonetheless, eventual solution-phase decomposition of these FLPs under ambient conditions was attributed to the long-term instability of the Ar_3Si^+ ions, which leads to protonated arene products.¹⁴

Herein we report the synthesis and characterisation of a *classical* donor-acceptor complex between tBu_3P and the less sterically encumbered, highly reactive, silylium ion iPr_3Si^+ . We show that the use of a Lewis adduct considerably stabilises the R_3Si^+ moiety, in comparison with the previously studied Ar_3Si^+ species. Furthermore, this species is shown to heterolytically cleave H_2 , the mechanism of which avoids the formation of a presumably highly unstable $iPr_3Si^+/PtBu_3$ separated Lewis pair.

Treatment of iPr_3SiH with $[Ph_3C][B(C_6F_5)_4]$ (Bartlett-Condon-Schneider hydride transfer)¹⁵ in chlorobenzene afforded solutions of $[iPr_3Si-CIPh]^+$, as previously described.¹⁶ Subsequent *in situ* reaction with tBu_3P furnished $[iPr_3Si-PtBu_3]^+[B(C_6F_5)_4]^-$ (**1**), upon precipitation with hexanes and recrystallisation from PhCl, in excellent yield (Scheme 1). **1** has been characterised by 1H , ^{13}C , ^{29}Si and ^{31}P NMR spectroscopy, high resolution MS (ES+), and elemental analysis (see ESI).



Scheme 1. Synthesis of **1** and subsequent reactivity with H₂/D₂ (C₆D₅Cl or C₆H₅Cl solution, 0.096 M; H₂ and D₂ experiments respectively, 4 bar). Reagents and conditions: a) i) [Ph₃C][B(C₆F₅)₄], PhCl, RT, -Ph₃CH; b) tBu₃P added in situ.

Slow cooling of a PhF solution of **1** to -25°C produced large colourless plates which were suitable for single crystal X-ray diffraction,[†] and the solid state structure is shown in Figure 1.

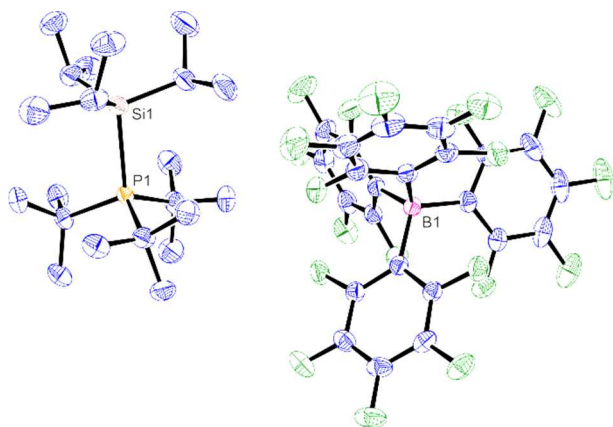


Figure 1. ORTEP plot of the X-ray structure of **1**. C atoms blue, P atom orange, B atom pink, F atoms green and Si atom light brown. H atoms have been removed for clarity, and thermal ellipsoids are shown at 50% probability; Si1–P1 = 2.4843(5) Å.

The [*i*Pr₃Si–PtBu₃]⁺ fragment in **1** is derived from a donor-acceptor interaction between the *i*Pr₃Si⁺ moiety and *t*Bu₃P, as exemplified by the pyramidalisation about the Si atom (0.5765(11) Å deviation from the plane of the three C atoms in the C₃Si unit). The C–Si–C bond angles, which range between 110.51(8) and 111.86(9)°, are much closer to the idealised tetrahedral angle (109.5°) than those found in *i*Pr₃Si(CHB₁₁H₅Cl₆) (117.3°),⁸ which possesses significant silylium character and hence approaches a trigonal geometry. The [B(C₆F₅)₄]⁻ anions are well separated from the cations, with no close Si to F contacts, and hence are non-coordinating. However, the Si–P bond distance is rather long (2.4843(5) Å; within the top 2% of those reported in the CSD), and may be compared to [PhMe₂Si–PtBu₃]⁺[HB(C₆F₅)₃]⁻ (Si–P = 2.376(2) Å), which has been prepared from reaction of PhMe₂SiH and the FLP system *t*Bu₃P/B(C₆F₅)₃.¹⁷ This increased distance may be attributed to the higher degree of steric strain due to crowding between the organic groups along the Si–P axis.

PhCl solutions of **1** proved to be stable for at least several months at room temperature, and these can be heated at 90°C for 24

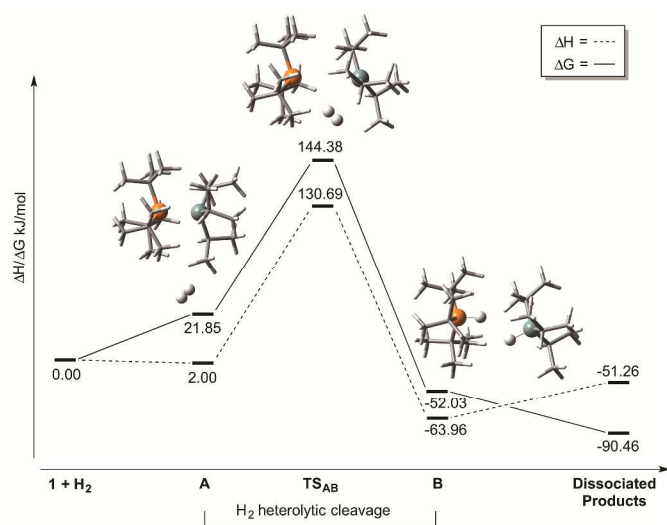
hr without evidence of decomposition. ¹H and ¹³C NMR spectra (C₆D₅Cl; 298 K) of **1** are commensurate with the solid state structure, and the upfield ²⁹Si resonance (δ = 43.1 ppm, ¹J_{SiP} = 23 Hz) reveals the Si–P bond to be intact in solution. However, the coupling constant is noticeably smaller than those reported for [(C₆Me₅)₃Si–PEt₃]⁺[B(C₆F₅)₄]⁻ or zwitterionic [Mes₂(SiPh₂H)P⁺CH₂CH₂B⁻(C₆F₅)₂] (¹J_{SiP} = 41.2 and 48.5 Hz, respectively),^{14,18} and the ³¹P NMR resonance (δ = 57.3 ppm, 298 K) is close to that observed for free *t*Bu₃P (δ = 62.0 ppm, 298 K); collectively these data would suggest a weak Si–P bond.

In contrast to the results obtained using the silylium FLPs, (C₆Me₅)₃Si⁺/PR₃,¹⁴ admission of H₂ to **1** (4 bar, C₆D₅Cl solvent) at room temperature led to no discernible reaction. However, heating these solutions to 90°C led to complete consumption of the adduct (8 hours; Scheme 1), concomitant with formation of *i*Pr₃Si–H (¹H NMR, δ = 3.43 ppm) and phosphonium borate [*t*Bu₃P–H]⁺[B(C₆F₅)₄]⁻ (¹H NMR, δ = 4.17 ppm, ¹J_{HP} = 430 Hz; ³¹P {¹H} NMR, δ = 60.3 ppm) in high conversion (90–94 %, four runs).¹⁹ Conducting these experiments under D₂ (C₆H₅Cl solvent) gave the deuterated products *i*Pr₃Si–D and [*t*Bu₃P–D]⁺, as shown by ²H, ³¹P and ²⁹Si NMR spectroscopy (see ESI for details),²⁰ this conclusively shows that H₂/D₂ is the source of H/D atoms in the formally hydridic silane, and protic phosphonium ion.²¹ Upon reaction completion, ¹⁹F and ¹¹B NMR spectra showed only resonances corresponding to the [B(C₆F₅)₄]⁻ counterion, demonstrating that silylium-mediated decomposition of the anion had not occurred. Since neither [H–B(C₆F₅)₃]⁻ nor B(C₆F₅)₃ could be observed in solution, the possibility that H₂ cleavage involves the known *t*Bu₃P/B(C₆F₅)₃ FLP pathway¹⁷ may be discounted.

In order to investigate the possibility that **1** may dissociate in PhCl solution to generate *t*Bu₃P and solvated silylium ion, [*i*Pr₃Si·CIPh]⁺, a variable temperature (VT) ³¹P NMR experiment was conducted (Ph₃P internal capillary reference; see Figure 7 in ESI). At low temperature the spectrum shows a very broad resonance (δ = 54.5 ppm, -40°C) which moves progressively downfield (δ = 60.6 ppm, 100°C), and markedly sharpens. However, it should be noted that the ³¹P NMR resonance for *t*Bu₃P also moves by ca. Δδ = 6 ppm over the same temperature range and the behaviour is likely due to a temperature-induced shift for both species. Furthermore, addition of *t*Bu₃P (1–10 eq., PhCl, 100°C) to **1** produced ³¹P NMR spectra consisting only of their separate respective resonances; neither a discernible change in the lineshape nor chemical shift position of the adduct was observed. If rapid exchange between **1** and appreciable concentrations of dissociated products were indeed occurring, introduction of extraneous *t*Bu₃P would be expected to lead to a significant perturbation of the ³¹P NMR resonance of **1**. Finally, we synthesised [*i*Pr₃Si·CIPh]⁺[B(C₆F₅)₄]⁻ in order to investigate its reactivity under the conditions of H₂ activation, in the absence of added phosphine: at 90°C this species degraded (40 min) *via* decomposition of the anion, producing B(C₆F₅)₃, *i*Pr₃SiF (¹⁹F NMR δ = -185.0 ppm; ¹J_{FSi} = 298 Hz) and other unidentified products. This implies that, if dissociation of *t*Bu₃P from **1** were to occur and generate *i*Pr₃Si⁺ or (more likely) [*i*Pr₃Si·CIPh]⁺, the rate at which this cation reacts with [B(C₆F₅)₄]⁻ would greatly exceed the observed rate of H₂ activation.

In order to determine whether thermally-induced dissociation of **1** (and hence a typical FLP-mediated mechanism) was responsible for H₂ heterolysis, we examined

this system further using DFT calculations. The results of our computational calculations (M06-2X/6-311+G(d,p) level of theory; see ESI) which took into account secondary (conductor-like polarizable continuum model; C-PCM) solvent interaction, are presented in Scheme 2 (and Table S1, ESI). Various conformations of **1** were identified (Table S2, ESI), which are separated in free energy by only *ca.* 13–30 kJ mol⁻¹; this lends support to the solution-phase VT ³¹P NMR data whereby the dynamic lineshape observed for **1** at low temperatures can be explained through interconversion between conformers. In total, two intermediates (**A** and **B**), and a single transition state for their interconversion (**TS_{AB}**) were located along the reaction coordinate for H₂ activation, starting from the lowest energy conformer of **1**.



Scheme 2. Solvent-phase enthalpy (ΔH) and Gibbs free energy (ΔG) profile for H₂ activation mediated by **1** in PhCl solution, showing intermediates and transition states along the reaction coordinate ($T = 363$ K). P atom orange and Si atom pale green.

On progressing to the transition state for H₂ activation, **TS_{AB}**, the Si...P distance lengthens considerably (Table 1), to an extent that is greater than the sum of the van der Waals radii of the elements (3.90 Å).²² The incipient cavity accompanying this bond elongation permits entry of a molecule of H₂ while attractive electrostatic P...Si and secondary van der Waals interactions between *t*Pr₃Si and *t*Bu₃P fragments lead to binding beyond covalent distances, thus retarding *t*Bu₃P dissociation. This factor explains the experimental observation that the rate of H₂ activation outcompetes decomposition, (which would be anticipated to be the faster process if dissociation to a true FLP were to occur at these elevated temperatures). A range of van der Waals interactions give rise to QTAIM bond critical points (BCP) between C–H...H–C and C–H...P (ESI, Tables S3 and S4) which appear to hold the fragments in place, fulfilling a similar role to those of the C–H...F H-bonds reported for the frustrated encounter complex [*t*Bu₃P]...[B(C₆F₅)₃] (the precursor species to H₂ cleavage by that FLP system).²³ It is notable that the B...P distance in the latter (4.20 Å) compares well with the Si...P length in **TS_{AB}**, yet both are appreciably

shorter than that those calculated for the ‘encounter complexes’ in Müller’s (C₆Me₅)₃Si⁺/PR₃ systems (R = alkyl, aryl; range 4.45–5.73 Å); this likely reflects the much greater steric bulk of the Ar₃Si⁺ fragment in these ‘true’ FLPs.¹⁴ Heterolytic cleavage of H₂ subsequently proceeds to give [*t*Bu₃P–H]⁺ and H–Si*t*Pr₃ as a dihydrogen bonded intermediate (**B**), after which dissociation to form the free products is strongly entropically driven. The TS molecular orbitals exhibit H₂ contributions (ESI, Fig 19) while BCPs from H₂ to both P and Si are obtained (ESI, Table S5). Moreover, the early TS exhibits nascent NBO P→H₂(σ*) and H₂(σ)→Si electron transfer, approximately equal to 8 and 34 kJ mol⁻¹ respectively, which are expected to increase as the reaction proceeds. Heterolysis of H₂ in this manner is consistent with the electron transfer model,²⁴ with no observable deviation in H–H distance between that in **TS_{AB}** and free H₂, denoting an early transition state.

Table 1. Pertinent interatomic distances corresponding to selected H₂ activation intermediates and transition states (Å).

	Si...P	H...H	Si...H	P...H
1	2.50	0.74 ^[a]		
A	2.50	0.74	5.36	6.09
TS_{AB}	4.10	0.74	2.87	2.98
B	4.63	2.34	1.50	1.41
Products	-	-	1.50	1.40

[a] Bond length corresponding to free H₂

Formation of **TS_{AB}** is both enthalpically unfavourable owing to the weakening of the Si–P interaction, and entropically unfavourable due to the increased ordering as a result of H₂ coordination (Table S1). The substantial energy barrier (122.53 kJ mol⁻¹; **A**→**TS_{AB}**) associated with H₂ activation is testament to the strong Si–P dative bond in **1** and explains the elevated temperatures required experimentally to achieve bond elongation and access the encounter complex. Comparable rate-determining steps have been observed elsewhere in the literature; Lammertsma *et al.* report that the experimentally observed insertion of CO₂ into dimeric P/Al-based Lewis pairs proceeds at room temperature, despite having computed a significant energy barrier (*ca.* 140 kJ mol⁻¹).²⁵

Whilst a number of Lewis pairs have been documented that exhibit classical/frustrated borderline reactivity with H₂ or alkynes,²⁶ spectroscopic evidence for the existence of the dissociated constituents has always been demonstrated, due to the stability of the Lewis acid as an independent entity. In our particular example, however, we have shown that the critical development of an encounter complex prior to H₂ activation can be achieved directly from the Lewis pair adduct *via* simple bond elongation and weakening, obviating the need for the formation of a true FLP; this is especially important when the Lewis acid partner (in our case *i*Pr₃Si⁺) is too reactive to isolate in the free form.

In conclusion, we have shown that a classical donor-acceptor adduct incorporating a highly electrophilic trialkylsilylium ion can activate H₂; this reaction is thus not rigidly confined to truly separated R₃Si⁺/base FLP combinations. Indeed, the formation of a stable adduct incorporating such species may provide a general strategy towards the protection of highly reactive Lewis acids/bases in Lewis pair systems, without excluding such systems from participating in characteristic FLP-type chemistry. We are currently exploring the small molecule reactivity of other [R₃Si–(base)]⁺ adducts, in addition to investigating their potential use in the catalytic hydrogenation of unsaturated substrates.

Notes and references

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† Crystallographic data for 1: C₄₅H₄₈BF₂₀PSi, *M* = 1038.70, triclinic, *a* = 11.6772(5), *b* = 12.5736(4), *c* = 17.0286(6) Å, *α* = 79.442(3), *β* = 75.102(3), *γ* = 77.292(3)°, *U* = 2335.61(15) Å³, *T* = 173 K, space group *P*-1 (no. 2), *Z* = 2, *ρ*_{calcd} = 1.477 g cm⁻³, *μ*(Cu_{Kα}) = 1.792 mm⁻¹, colourless blocks, Agilent Xcalibur PX Ultra A diffractometer; 8923 independent measured reflections (*R*_{int} = 0.0180), *F*² refinement,²⁷ *R*₁(obs) = 0.0354, *wR*₂(all) = 0.0948, 7708 independent observed absorption-corrected reflections [*F*_o > 4σ(*F*_o)], 2θ_{max} = 147°, 613 parameters. CCDC 970728. Electronic Supplementary Information (ESI) available: Experimental section including computational details, characterisation data and copies of the NMR spectra compounds and reactions. See DOI: 10.1039/c000000x/

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- 1** proved to be highly reactive to even trace amounts of H₂O. Accordingly, the synthesis and solution-phase studies required meticulous drying of solvents, conducting reactions in Teflon[®] vials, and performing NMR experiments in Teflon[®] inserts. Reactions in dilute solution, such as those under H₂, invariably become exposed to trace amounts of H₂O, leading to [tBu₃PH]⁺[B(C₆F₅)₄]⁻ and (iPr₃Si)₂O (2:1). The yield for H₂ conversion is most reliably calculated *via* relative integration of ¹H NMR signals for the iPr₃Si-H resonance against that of tBu₃P-H (1:1 from H₂). Accordingly, hydrolysis produces twice the amount of phosphonium salt than the reaction with H₂, and accounts for the slightly sub-quantitative yields (90-94%). In support of this conclusion, ²H NMR spectroscopy of the D₂ cleavage reaction shows a 1:1 tBu₃P-D: D-SiPr₃ ratio, since adventitious moisture is introduced as H₂O, and not D₂O (see Figure 8 in ESI).
- A full kinetic study by ¹H NMR was hampered by iPr₃SiH interaction with **1**, which masks the actual amount of silane produced until reaction completion (i.e. complete conversion of **1**). This phenomenon has been noted for R₃Si(μ-H)SiR₃⁺ species,²⁸ wherein the silane proton is rendered invisible to ¹H NMR spectroscopy. It is plausible that a similar bridging silane interaction is also operational here.
- Heating solutions of independently prepared [tBu₃P-H]⁺[B(C₆F₅)₄]⁻ with iPr₃SiH (PhCl, 90°C, 60 hr) led to no reaction, demonstrating that the H₂ cleavage reaction is irreversible. Furthermore, conducting the same experiments under D₂ did not produce any HD, which would be expected to form due to H/D scrambling.
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