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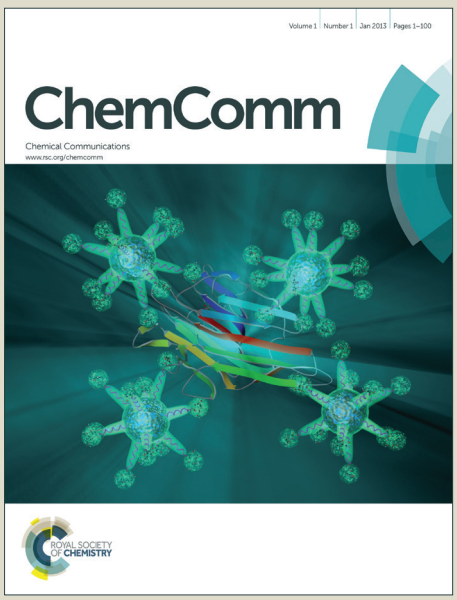

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ARTICLE

Discovery of a Novel Family of Polycyclic Aromatic Molecules with Unique Reactivity and Its Members Valuable for Fluorescent Sensing and Medicinal Chemistry

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Novel ploycyclic aromatic molecule, *i.e.* 1-oxo-1*H*-phenalene-2,3-dicarbonitrile (compound **1**, initially misidentified as 8-oxo-8*H*-acenaphtho[1,2-b]pyrrolecarbonitrile) was discovered by our group in 2005. This parent compound is highlighted for its unique oxidative S_NAr^H (nucleophilic substitution of aromatic hydrogen) reactivity that provides easy approaches to diverse derivatives with different long wavelength fluorescence and important biological activities. To date, a large number of derivatives have been synthesized and evaluated by several international research groups, indicating the formation of a new and valuable family of functional chemicals. Some members have been functionalized for molecular or nanoparticle-based probes applicable in chemical and environmental sensing, biomolecule imaging and tumor diagnosis. And the others have been qualified as high potency anticancer agents specifically targeting different functional proteins in tumor cells. Taking into the more and more attention paid to this new chemical family, it is good time to review the major achievements in order to promote the further and deeper investigations.

1. Introduction

Novel polycyclic aromatic compounds easily obtained from the petroleum or coal tar chemicals are important precursors for further modifications to develop variety of functional materials, *e.g.* dyes, sensors and medicines. The reasons are easily understandable: On the one hand, they have extended π-conjugation skeletons that are critical structural basis for chromophores and fluorophores; on the other hand, polycyclic materials are of rigidity, planarity and hydrophobicity, which render them the capability to insert into the biomacromolecules' cavities adapted with these small molecules in terms of size, hydrophobicity and other supramolecuar interactions. If these polycyclic precursors can undergo some chemical modifications under mild conditions, this will strengthen their importance as universal scaffolds, or leads toward practically applicable products.

In 2005, a new polycyclic compound **1** (initially misidentified to be 8-oxo-8*H*-acenaphtho[1,2-b]pyrrolecarbonitrile) was discovered.[1] From then on, compound **1** exhibited great attractive for researchers in different fields, not only because its planar and rigid chemical structure qualified it as desirable precursor to develop novel derivatives with diversified spectral properties or biological activities, but also because of its unique chemical reactivity favorable for versatile and simplified

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Fengyu Liu was born in Liaoning, in 1973. She received her B.S.and M.S. degrees from East China University of Science and Technology in 1995 and 2002, and Ph. D. Degree from Dalian Univesity of Technology in 2006. She became associate professor (2013) in Dalian Univesity of Technology. Her research interest focuses on electrogenerated chemiluminescence (ECL) and its applications.

derivations. To this date, a large number of compound **1**'s derivatives as well as their promising applications as chemosensors or antitumor therapy agents have been reported by several different research groups. Taking into the growing interest in this new chemical family, it will be necessary to summarize the important progresses from their discovery till now.

2. Discovery of the parent compound 1

During our research project to develop efficient fluorescent dyes for biological imaging, we noticed the unsatisfactory quantum yields of some fluorophores. It had been known that introducing electron-withdrawing cyano to the conjugated systems was favorable for fluorescence emitting.[2] Previously, quite a few fluorophores had been obtained by the Knoevenagel condensations between various aromatic aldehydes/ketones and malononitrile, as shown in scheme 1.[3] While these fluorophores demonstrated some advantages, such as long and tunable emission wavelength and environmental sensitivities, their fluorescence were not very strong in common solvents. This could be ascribed to the flexible conjugation structures; Obviously, the free rotation and vibration of the carbon-carbon double bonds connecting aromatic planes to cyanos dissipated excitation energy and thus quenched the luorescence. Additionally, according to the experiences on another type of fluorophores, 4-amino 1,8-naphthalimides that had fluorescence quantum yields up to 100%,[4] it was easy to conclude that the rigid polycyclic conjugation structures were essential for strong

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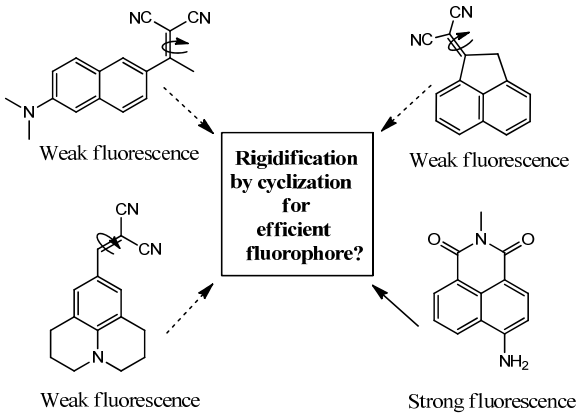
Weiping Zhu received his Bachelor's degree from the National University of Defense and Technology in 1989, and Master's degree from East China University of Science and Technology (ECUST) in 1992. After working at Yangzi Petrochemical Institute of SINOPEC for seven years, he returned and received his PhD degree at ECUST (2004). He spent one year as a visiting scholar at UCLA (2009), and now is a professor at ECUST. His research interests focus on the design and synthesis of fluorescent sensors and smart drug delivery systems.

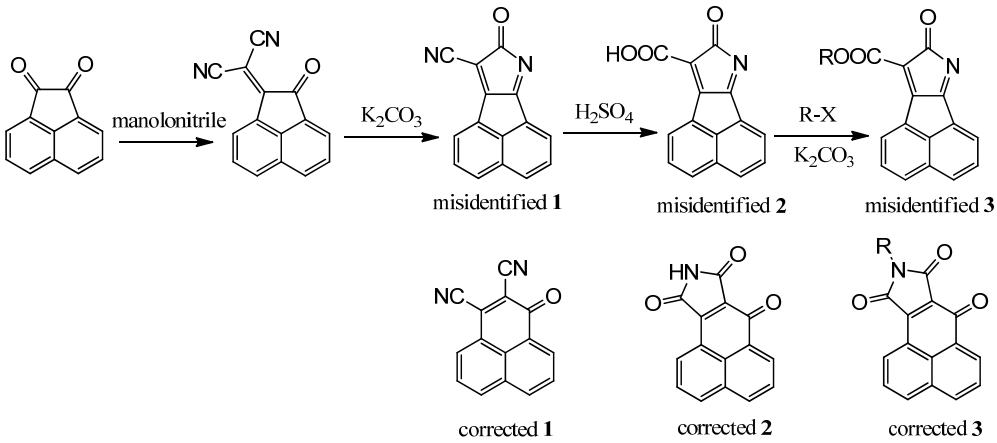


Scheme 1: comparison of the structures of three Knoevegal adducts with weak fluorescence and that of strong fluorescent 4-amino 1,8 naphthalimide indicating the strategy to get efficient fluorophores

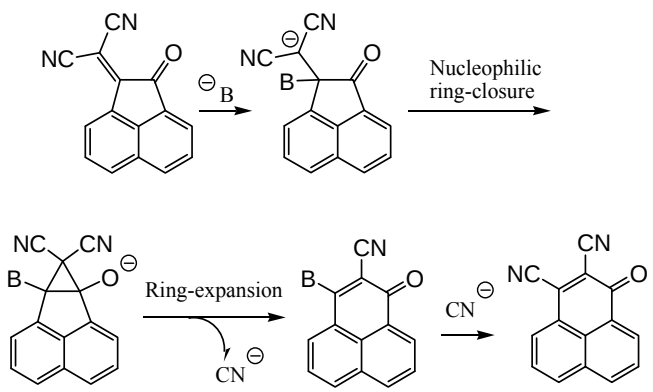
Yufang Xu was born in Jiangsu, in 1965. She received her Ph.D. degree in Chemistry from East China University of Science and Technology (ECUST) in 1994. Then she joined in ECUST. From 1999 to 2001 she worked as postdoctoral fellow in Tokyo Medical and Dental University. Her research interests focus on small organic molecules to regulate the bioprocess, includes the design and synthesis of fluorescent probes to response the hypoxia and related enzymes to diagnose and treat related diseases especially in solid tumour.

Xuhong Qian was born in Jiangsu, in 1962. He received his B.S., M.S. and Ph. D. degrees from East China University of Science and Technology in 1982, 1985 and 1988, respectively. Then he was associate researcher in Lamar University, US and Humboldt postdoctoral fellow in Wuerzburg University, West Germany (1989-1991), he was back and became associate professor (1992) and professor (1995-2000, 2004-present) in his Alma Mate. During the period (2000-2004), he was Chongkong professor in Dalian University of Technology.
He was FRSC (2009), academician of Chinese Academy of Engineering (2011), Honorary Doctor of Science, Queen's Univerisity, Belfast (2012). His research interest covers some aspects of bioorganic chemistry & engineering related with dyes and pesticides, e.g. fluorescent sensors and antitumor agents derived from dyes, as well as of green insecticides and plant-activators.

fluorescence. Thus, in order to construct novel and strong fluorophores, it became a noteworthy idea to search for a ring-formation strategy to fix the double bonds formed by Knoevenagel condensation.



Scheme 2: The syntheses and structure identifications of novel polycyclic parent molecule **1** and its hydrolysis derivatives **2** and **3**



Scheme 3: Mechanism of the base-catalyzed ring expansion proposed by Wang *et al*, [6b] and Lebreton *et al*. [7]

Then, acenaphthenequinone was chosen as starting material because it was one of the common and cheap coal chemical products. Knoevenagel condensation of acenaphthenequinone with malononitrile was carried out, as shown in Scheme 2. When there was no basic catalyst, the monoadduct with two cyano groups were the only product. We predicted that, under basic conditions, the remaining ketone group should be activated to attack the adjacent electron-deficient double bond, which might promote a further ring-close transformation. So, this strategy was attempted by adding anhydrous K_2CO_3 to the solution of the monoadduct intermediate, which immediately generated a yellow precipitate *i.e.* compound **1** with high yield up to 88% and high purity. Through characterization with HRMS, ^{1}H NMR, ^{13}C NMR, IR and so on, compound **1** was identified as 8-Oxo-8*H*-acenaphtho[1,2-b]pyrrolecarbonitrile, which had been proved an error very recently. The correct structure of compound **1** should be 1-oxo-1*H*-phenalene-2,3-dicarbonitrile (phenalene dicarbonitrile in the rest of this review). The hydrolysis of cyano groups in compound **1** were carried out to facilely get another important compound **2** (naphtho[1,8-*ef*]isoindole-7,8,10(9H)-trione, abbreviated as phenalene imide). [5] This molecule was also misidentified, as a result of the incorrect structure of the parent. Thankfully, by two international teams, *i.e.* Wang and Qian *et al*,[6] and Lebreton *et al*, [7] these misidentifications had been corrected lately, through the analyses of X-ray crystallographic structures and 2D NMR spectra.

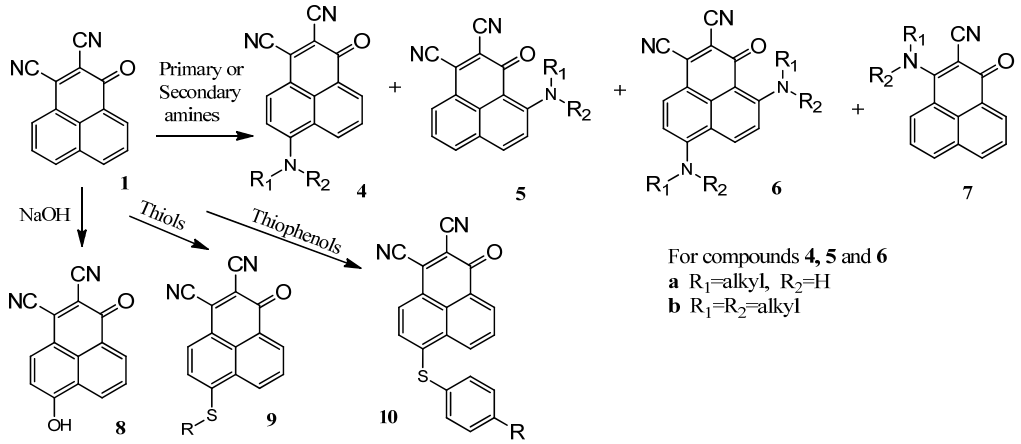
In 2014, Wang *et al* tried to optimize the reaction via using different basic catalyst, and found that the best one was an organic base, DMAP, and it could improve the yield of compound **1** to 92% that was a little higher than K_2CO_3 catalysis. They also proposed a reasonable mechanism of the base-promoted ring expansion, as shown Scheme 3.[6b] Several months Later, Lebreton *et al*,[7] also proposed the very similar mechanism and confirmed it through using K^{13}CN to exchange cyano on position 3. Understanding this mechanism proved to be very helpful to develop new phenalene-containing molecules analogue to compound **1**, as Lebreton *et al* had used 4-nitrophenylacetonitrile, instead of malononitrile to react with acenaphthenequinone in the presence of K_2CO_3 and obtained a novel nitrophenyl substituted phenalene derivative.[7]

3. Unique oxidative nucleophilic substitution of aromatic hydrogen (S_NAr^H): versatile and easy derivations

Looking at the structure of compound **1** and thinking over about its feature, it could be classified into an electron-deficient arene, for the strong electron-withdrawing carbonyl and cyano groups directly on the conjugated skeleton. According to previous studies, nucleophilic substitutions of a very few electron-deficient arenes took placed on the aromatic hydrogen atoms, which could be classified a kind of green chemistry because it was the hydrogen but not a conventional leaving group that was substituted.[8] Such a rare reaction was called oxidative nucleophilic substitution of aromatic hydrogen (S_NAr^H). With these considerations in mind, we decided to find out if it was feasible for compound **1** to undergo oxidative S_NAr^H.

Fortunately, compound **1** proved to be highly reactive to common N, S and O type nucleophiles, as shown in Scheme 4.[1,]

[6,7] Consequently, the oxidative S_NAr^H reactions can produce variety of derivatives. Generally, the reaction conditions were



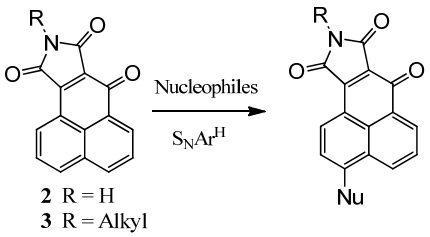
Scheme 4: Nucleophilic substitution reactions of compound **1** with different types of nucleophiles

very mild. For example, in acetonitrile, S_NAr^H reactions with the amines and hydroxide processed rapidly at room temperature. And it was found that both H-6 and H-9 on the phenalene ring were readily substituted by amines, which not only produced the regioisomeric derivatives **4** and **5**, but also the bissubstituted compound **6**. The yield distributions of the amine-substitution products were complicated and highly dependent on the types of amines, the reaction conditions *e.g.* mass ratios and reaction times. If reaction times were short, 6-amino derivatives would be the dominant products (yields 20-50%) when equivalent amines were used; this seemed to indicate that H-6 was most activated aromatic hydrogen. However, if excess amounts of amines were added and the reaction time was extended, the yields of **6** would increase and it would then become the major product. For instance, when the 4-6 equivalence of piperidine reacted with compound **1** in acetonitrile solution at room temperature for 2 hours, 6-piperidino product was up to 49%, while 9-piperidino isomer was 30% but no bissubstituted product was observed; when the reaction time was extended to 12h, the yield of 6,9-bispiperidino derivative increased remarkably to 48%. In acetonitrile, the S_NAr^H reactivity of thiols was not as high as amines: after 72h heating, the yield of 6-dodecylsulfanyl derivative was 42%. As for hydroxide, its reaction with compound 1 could smoothly processed in DMSO-H_2O at room temperature overnight to give compound **6** with a moderate yield (38%).

Above results confirmed that the compound **1** was a new and highly efficient platform molecule for oxidative S_NAr^H reactions and it was feasible to get different substituted derivatives by choosing nucleophiles or tuning the reaction conditions.

Other groups were inspired to try S_NAr^H reactions on compound **1** by using other kinds of neucleophiles not initially attempted by us, and also got the positive results. For example, Wang *et al* [6b] had successfully achieved the reactions of compound **1** with thiophenols and obtained the corresponding 6-substituted derivatives **10**. Besides S_NAr^H reactions, Lebreton *et al* also found that the cyano-3 of compound **1** could also be substituted by amines to produce small amount of compounds **7** (*e.g.* butylamine substitution yield 3%).[7] Although this conventional nucleophilic substitution needing leaving group was not as efficient as oxidative S_NAr^H reactions, it helped to expand the diversity of derivatives from parent **1**.

We reasoned that compound **2** and **3** should inherit high S_NAr^H reactivity from its parent **1** because their conjugation structures remained highly electron-deficient. As expected, compound **2** and **3** could also be readily substituted by various neucleophiles to produce corresponding 6-substituted derivatives, as shown in scheme 5.[5] As their unique S_NAr^H reactions with variety of nucleophiles provided versatile and easy approaches to a large number of derivatives, compound **1**, **2** and **3** became valuable scaffolds for the development of new functional chemicals.



Scheme 5: Oxidative S_NAr^H reactions of compounds **2** and **3**

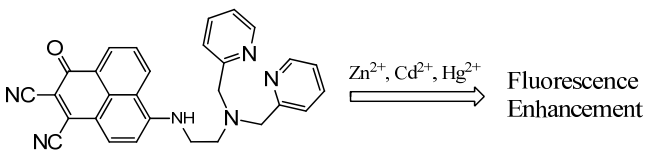
4. Favourable Fluorescence characteristics for the development of fluorescent sensors

While the parent compound **1** is nonfluorescent, most of the 6-substituted derivatives presented strong long wavelength fluorescence.[1,6b] For instances, the fluorescence quantum yields of compound **4a-1** (butyl amine substituted), **8** and **9** (dodecanethiol substituted) were 0.81 (in dichloromethane), 0.87 (in chloroform) and 0.43 (in chloroform), respectively. Firstly, this fluorogenic phenomenon should be ascribed to these derivatives' ICT (intramolecular charge transfer) nature: with strong electron-withdrawing cyano and carbonyl groups on one side of the conjugation skeleton, the introduction of electron-donating groups (nucleophiles) to the other opposite side (position-6) constructed the donor-π-acceptor structure. That was the origin of their ICT processes. The 9-substituted derivatives had been reported to emit much weaker fluorescence, due to improper location of the donating group unfavorable for ICT. The ICT fluorophores are usually

sensitive to the environmental polarity. That is especially true for 6-piperidino compound **4b-1** (on behalf all the 6-secondary amine substituted derivatives) whose fluorescence quantum yield is up to 0.38 in toluene but is almost nonfluorescent in acetonitrile. Another advantage of these new ICT fluorophores was their high photostability, because the electron-deficiency inherited from their parent made them more resistant against the photooxidiation (one of the major reasons for light induced decomposition). More interestingly, different 6-substituted derivatives exhibited tunable fluorescence in the long wavelength range of visible light, because nucleophiles with different electron donating capability strengthened the ICT to different extents. As a result, the emission maxima of compound **4a-1**, **9** and **8** were at around 595, 567and 548nm, respectively, and so, their fluorescence were in different colors (orange, yellow and green). Owing to the advantageous fluorescence properties *e.g.* longer and tunable wavelength, high photostability, environmental sensitivity and the convenient synthesis *etc.*, the fluorophores from this phenalene dicarbonitrile family have been adopted for the constructions of various small-molecule probes and nanoparticle-based multifunctional chemosensors.

4.1 Molecular probes

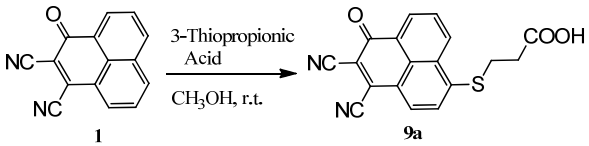
As 6-amino phenalene dicarbonitrile fluorophores for the long wavelength emission and convenient synthesis through S_NAr^H, Akkaya *et al* [9] employed this fluorophore for the first time to develop a fluorescent chemosensor for metal ions (compound **4a-2** in Scheme 6). In their design, they adopted aminoethyl dipicolylamine as the receptor of the cations, and facilely introduced this unit to the 6-position of the phenalene dicarbonitrile through S_NAr^H under the same reaction conditions previously provided by us. Compound **4a-2** selectively responded to group IIB cations by remarkable fluorescence enhancement through inhibiting the fluorescence quenching process of PET (photoinduced electron transfer) from the dipicolylamine to the excited fluorophore. For example, the emission intensity at the peak emission wavelength of 588 nm increased 7-fold upon Zn(II) binding. Benesi–Hildebrand analysis of the binding data revealed strong association constants: 2.3×10^6 for Zn(II), 5.4×10^5 for Cd(II) and 2.9×10^6 for Hg(II), all in M^{-1}.



4a-2

Scheme 6: a fluorescent molecular probe of group IIB cations

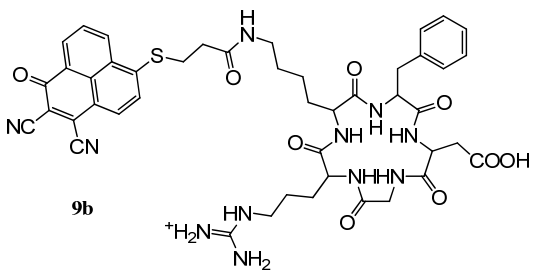
While our group reported that the S_NAr^H reaction of the parent **1** with n-dodecyl thiol required heating in acetonitrile,[1b] Li *et al* optimized the conditions by using protic solvents. As shown in scheme 7, they found that in methanol, the reaction of **1** with 3-thiopropionic acid can efficiently occur at room temperature to produce corresponding 6-thiol substituted compound **9a**. Because Cys and Hcy, are similar in chemical structure to 3-



1 **9a**

Scheme 7: S_NAr^H recation under optimized conditions

thiopropionic acid, Li *et al* then reasoned that the reaction of **1** with Cys or Hcy would readily promote a fluorescent recognition of these thiol-containing amino acids critical for cell functions. As they had expected, the parent **1** proved to be an ideal 'nake-eye' visible and fluorescent probe for Cys or Hcy. In a mixture of methanol and HEPES solution at pH 7, upon addition of Cys or Hcy, the absorption band of **1** centered at 430 nm gradually decreased and a new absorption band centered at 580 nm appeared, and an emission peak at 588 nm increased rapidly. The observed spectral changes indicated that the oxidative S_NAr^H reaction took place. The final fluorescence enhancement was up to 75 fold in the presence of 40 equa Hcy, which revealed a high sensitivity. Subsequently, compound **1** was successfully applied for the imaging of Cys/Hcy under confocal laser scanning microscopy and two-photon laser scanning microscopy. Since this new sensor's excitation and emission were in longer wavelength ranges than the previous reported Cys/Hcy probes, compound **1** was of great benefit for studying the effects of Cys/Hcy in biological systems.[10]

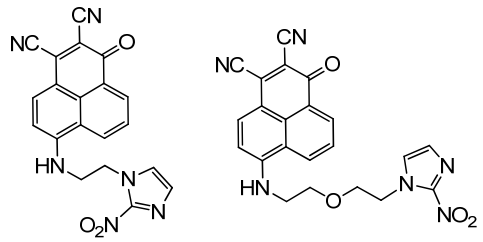


9b

Scheme 8: a fluorescent probe targeted imaging of tumor cells

In order to achieve the specific tumor imaging, Li *et al* [11] developed probe (compound **9b** in Scheme 8) composed of a 6-thiol phenalene dicarbonitrile as fluorophore and a cyclic peptide, c(RGDFK) as the targeting unit, based on the knowledge that an arginine–glycine–aspartic acid tripeptide sequence could specifically bind integrin $\alpha_v\beta_3$ of tumor cells. Although there had been some other probes constructed by the same strategy, the major limitation was their intolerance of photobleach. Li *et al* carried out the photodecompose test in cells under the confocal microscope and found that probe **9b** had much better photostability than Acridine Orange, one of the common used cellular fluorescent labels. Moreover, probe **9b** possessed high solubility and exhibited very low cytotoxicity.

There features were favorable for cell biology applications. Two different cell lines were chosen: one was human cervical carcinoma HeLa cells over expressing integrin $\alpha_v\beta_3$, and the other was human hepatocellular carcinoma SMMC-7721 cells as a control. A strong fluorescence signal was observed in the HeLa cells, whereas the SMMC-7721 cells (expressing lowly $\alpha_v\beta_3$) showed much less fluorescence. This comparison confirmed probe **9b** was specifically targeting $\alpha_v\beta_3$, and thus, qualified it as a practical diagnostic tool for detecting certain tumor cells expressing $\alpha_v\beta_3$.



4a-3 **4a-4**

Scheme 9: two fluorescent probes for the hypoxic environment of solid tumors

The solid tumor cells are hypoxic and these hypoxic cells are more resistant to radiation than cells under well-oxygenated conditions. Thus, to better predict treatment outcome and select appropriate therapies, it is very meaningful to accurately differentiate the hypoxic cells from cells under normal oxygenation status. In 2006, our group attached the novel 6-amino phenalene dicarbonitrile fluorophore to the classical hypoxic environment targeting molecule nitroimidazole, and developed long wavelength (maximum emission at 592nm) hypoxic fluorescent indicators, *e.g.* **4a-3** and **4a-4**, as shown in Scheme 9.[12] This work were carried out because a number of previously developed intracellular hypoxic fluorescent markers had unsatisfactory detection accuracy or sensitivity, as their emissions in the relatively shorter wavelength ranges might be interfered by biological background fluorescence. The molecular design utilized two properties of nitroimidazole: it had strong fluorescence quenching effect, and it could undergo bioreduction in hypoxic cells to generate the product with less quenching ability. That was why the markers could release 'turn-on' fluorescent signals in hypoxic cells while in normal cells no fluorescence enhancement was observable. After 3h incubation of V79 379A Chinese hamster cells, **4a-4** and **4a-3** showed remarkable 15 and 11 times fluorescence enhancements, respectively in the hypoxic cells compared to the cells at normal oxygenation status. The reason that **4a-4** outperformed **4a-3** could be partly ascribed to the better solubility and higher cell uptake brought by the ester-type spacer between the fluorophore and the nitroimidazole.

4.2 Nanoparticle/polymer-based fluorescent chemosensors

Small molecular probes have some limitations: they have relatively low sensitivity, due to the relatively low fluorescence intensity of a single dye molecule; they cannot be reused; and usually a molecular probe has a single function. Nanoparticle-based fluorescent sensors can overcome these major problems. In order to develop silca nanoparticles as multifunctional sensing materials, we introduced siloxane unit to 6-amino phenalene dicarbonitrile fluorophore.[13] This fluorescent siloxane derivative underwent hydrolysis to form a core for the core-shell nanoparticles; and this core emitted stable long wavelength fluorescence (peaked at 595nm) from 6-amino phenalene dicarbonitrile derivative, which was not interfered by environmental factors, *e.g.* pH; and thus this fluorescence signal could act as the inner references. Outside the core, a silica shell loaded with fluorescent probe was formed to selectively sense Zn^{2+} with the 'turn-on' signals. As shown in

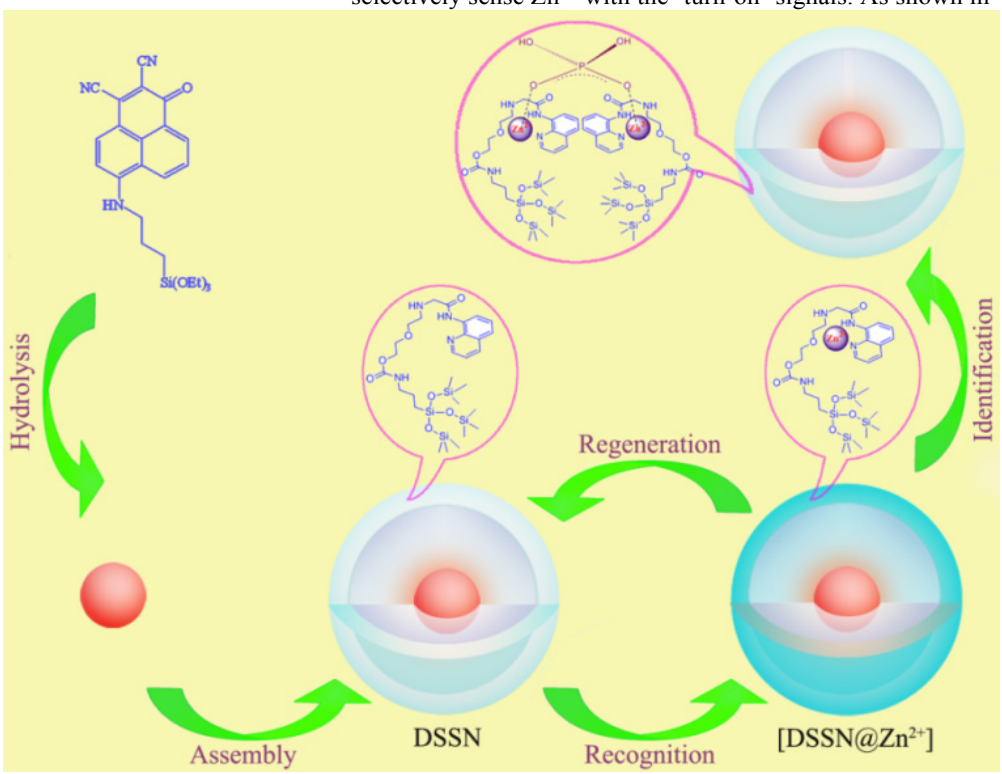
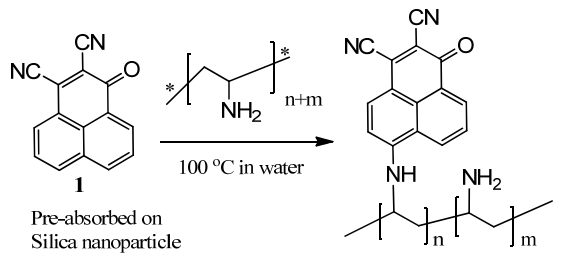


Figure 1: the structure of a siloxane-modified fluorophore and the design concept of nanoparticle-based ratiometric chemosensors for Zn^{2+} and $H_2PO_4^-$

Figure 1, since the shell's short wavelength fluorescence (centered at 480nm) from the Zinc probe could be clearly differentiated from that of the core's, this created an ideal ratiometric sensing platform (**DSSN**) for accurately quantifying Zinc ions not only in aqueous solution but also in live cells. After **DSSN** binding Zn^{2+} to form the new nanoparticle named **DSSN@Zn^{2+}**, its regeneration was simple and the recovered **DSSN** could be reused again. It was proved that **DSSN** could be recycled at least four times without losing the Zn^{2+} sensitivity. It was more interesting that **DSSN@Zn^{2+}** could be used to selectively recognize $H_2PO_4^-$ against other anions. This was because $H_2PO_4^-$ was an efficient Zn^{2+} binder, and such binding decreased the probe's fluorescence intensity and changed the intensity ratio of the Zinc probe to the inner reference fluorophore. At a neutral aqueous solution, **DSSN@Zn^{2+}** exhibited a detection limit lower than 6×10^{-6} M, indicating a high sensitivity to $H_2PO_4^-$.



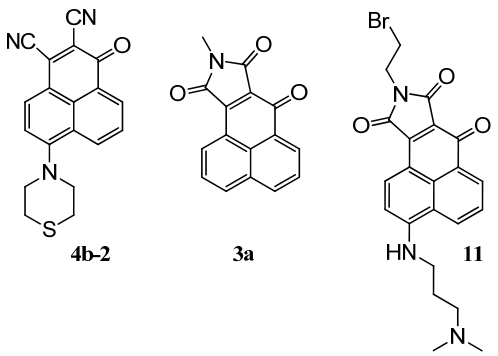
Scheme 10: the synthesis strategy of polymer-silica hybrid fluorescent particle by nucleophilic aromatic substitution of **1** adsorbed onto silica particles

Different from our strategy to synthesize siloxane substituted fluorophore for directly modifications of Silica nanoparticles, Spange *et al* [14] estabilished a novel and convenient two-step approach toward fluorophore-functionalized silica particles. As demonstrated in Scheme 10, they used silica as a 'solubilizer' which could absorb the water insoluble precursor **1** and water soluble poly(vinyl amine) (PVAm) and simultaneously provided them a large specific surface area to undergo S_NAr^H reaction. If without silica, such a reaction between the two incompatible reactants needed the aid of cyclodextrins. They also found that the fluorophore substituted PVAm exhibited pH-sensitive fluorescence spectral shifts: in the alkaline aqueous solution the fluorescent polymer possessed two emission bands peaked at 515 and 579 nm, respectively; With decreasing pH value, the intensity of the short wavelength emission band decreases, whereas for the emission band at the longer wavelength a slightly bathochromic shift was observed. These phenomena should be associated with the characteristics of the PVAm: under the acidic conditions, the amino groups were protonated and the mutual exclusions between the positive charged ammoniums resulted in polymer chains stretch, and thus the interactions between fluorophores were slight; but in basic solution, curl chains of PVAm maked fluorophores close to each other and thus strength their intramolecular interactions. Therefore, Spange *et al* successfully developed a multifunctional hybrid material possessing the advantages of silica nanoparticle,, the 'smart' pH-responses of PVAm and the environment–sensitive fluorescence of the new fluorophore.

5. Phenalene dicarbonitrile and imide derivatives as anticancer drug leads

5.1 Discovery of anticancer activities: DNA intercalators or protein inhibitors

Some electron-deficient polycyclic aromatic molecules, *e.g.* naphthalimides, anthracyclines exhibit antitumor activities and had become important drug leads. For long time, our group had been investigating anticancer agents based on naphthalimide as a type of DNA intercalators.[15] There was no surprise that we immediately began experimental evaluating the biological activities of the new phenalene dicarbonitrile and imide derivatives, taking into account their similarity in terms of chemical structure and electron deficiency to naphthalimides. As was expected, quite a few derivatives exhibited strong cytotoxicity to cancer cells, with submicromolar IC_{50} values, which qualified their status as drug leads. Scheme 11 listed three representative ones of higher activities: Compound **4b-2** showed IC_{50} of 0.17 μM against human cervical carcinoma (HeLa) cell line, compound **3a** showed IC_{50} of 0.45, 0.80 μM against A549 and P388 cell lines, and IC_{50} values of compound **10** against A549 and P388 cell line were 0.14 and 0.019 μM, respectively.



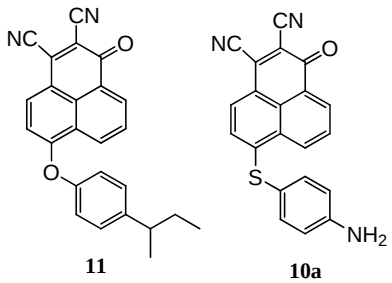
Scheme 11: three early discovered compounds with high antitumor activities

As naphthalimide-based drugs were the well-known DNA intercalators,[16] we initially considered to explain the antitumor activities of these new compounds by their DNA interacting abilities. Actually, they indeed demonstrated strong DNA binding in solution, as we had evaluated in DNA melt curves and circular dichroism experiments. [17] However, during the preliminary investigation on mechanism of the compound **4b-2**, it was discovered that this compound could inhibit the anti-apoptotic B-cell lymphoma 2 (Bcl-2) interacting protein and thus induce tumor cells apoptosis.[18] This result indicated that the some new phenalene dicarbonitrile and imide derivatives, at least compound **4b-2**, may specifically bind to and influence certain functional proteins, instead of DNA. Therefore, the

subsequent study should be focused on the identification of targeted proteins to lay the foundation to achieve higher specificity to and stronger inhibition of tumors.

5.2 Anticancer mechanism 1: Dual inhibitors of Mcl-1 and Bcl-2

Zhang, one of the former members at our group, independently carried out further research on the antitumor mechanisms of compound **4b-2**. Subsequently it was found that **4b-2** inactivated not only Bcl-2, but also Mcl-1, indicating that this compound was an inhibitor of *pan*-BCl-2 proteins. Zhang successfully clarified that compound **4b-2** could insert into and occupy the hydrophobic BH3 groove, the critical domain on the surface of some prosurvival Bcl-2-like proteins including Bcl-2 and Mcl-1. Nanomolar compound **4b-2** could inhibit them (K_i = 58 and 310 nM against Bcl-2 and Mcl-1 respectively) and thus induced tumor cells apoptosis efficiently. [19]

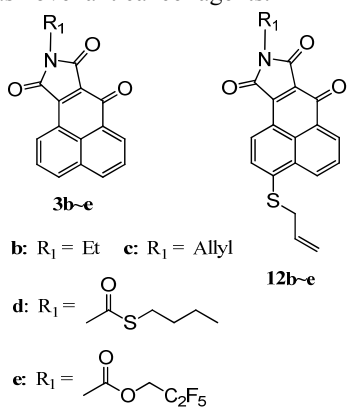


Scheme 12: two highly efficient dual inhibitors of Mcl-1 and Bcl-2

Based on their verifying compound **4b-2** as a novel molecular BH3 mimetic and a potent drug lead, Zhang *et al* focused on searching for more efficient *pan*-BCl-2 inhibitors derivative from the family of phenalene dicarbonitrile. [20,21] They designed and synthesized two series of 6- phenoxy and 6- phenylthio substituted phenalene dicarbonitrile derivatives. Among them, the *p*-sBu phenoxy compound **11**, and the *p*-amino phenylthio compound **10a** (Scheme 12) were screened to exhibit nanomalar affinities toward Mcl-1 and Bcl-2, as well as nanomalar cytotoxicity activity against multiple cancer cell lines. The affinity for **11** toward Mcl-1 and Bcl-2 (K_i = 24 nM and 158 nM, respectively) was about 2-3 times enhanced than the lead compound **4b-2**. The IC_{50} of compound **10a** to Mcl-1 (5nm) was significantly better than that of compound **4b-2**, while it maintained the similar affinity to Bcl-2. It was also exciting that compound **10a**'s cytotoxicity against tumor cells (IC_{50} = 12nM and 16 nM against MCF-7 and MMC-7721, respectively) was 10 folds stronger than that of the compound **4b-2**.

5.3 Anticancer mechanism 2: FGFR1 inhibitors

Inhibition of FGFRs (the fibroblast growth factor receptors 1-4) represents an attractive tumor therapeutic strategy. However, on account of structure homology of catalytic domains of various kinases, most existing FGFRs inhibitors inhibit different tyrosine kinases nonspecifically, which might cause side effects. In 2011, we reported for the first time potent and selective FGFR1 inhibitors based on phenalene imide derivatives.[22] After verifying the great in vitro antiproliferative effects of a simple phenalene imide (Compound **3b,** in scheme 13), we continued the mechanism investigation and successfully identify FGFR1 as the molecular target. Compound **3b** not only demonstrated potent inhibition (IC_{50} = 74 nM) on FGFR1 in vitro, and exhibited remarkable selectivity toward FGFR1 compared to other tested tyrosine kinases, including other FGFRs and highly homologous enzymes PDGFRs and VEGFRs. Subsequently, a number of phenalene imides and 6-thiol substituted phenalene imides were designed and synthesized for the systematic SAR investigation. In our design, we were trying to use the unsaturated hydrocarbon, fluorine or sulfur-containing side chains to modify the precursor, because, based on our research experience, these were beneficial factors to improve biological activities. It was found that for the phenalene imides, the introduction of different side chains to the imide nitrogen was critic for the FGFR1 inhibition. Especially, compound **3c** possessing an allyl goup on imide presented the strongest inhibition with a much lower IC_{50} (19 nM) than compound **3b**. While the 6-thiol substituted phenalene imides demonstrated weaker inhibition than corresponding non substituted phenalene imides, it was interesting that the groups on the thiol sulfur atom affected the activities remarkably. Again, the allyl thiol substituted molecule (Compound **12c**) showed the lowest IC_{50} of 216 nM. In Vitro antiproliferative assay was also carried out and the results confirmed that these new synthesized FGFR1 inhibitors inhibited the proliferation of human cancer cell lines much efficiently than normal cells. The above discovery of phenalene imides as novel FGFR1 inhibitors provided a promising starting point for further drug optimization and development as novel anticancer agents.



Scheme 13: highly efficient FGFR1 Inhibitors

6. Conclusion and Prospect

Starting from acenaphthenequinone, through the base-promoted ring expansion, we synthesized a novel rigid polycylic molecule **1**, 1-oxo-1*H*-phenalene-2,3-dicarbonitrile.We discovered that compound **1**'s aromatic hydrogen atoms on

postion 6 and 9 were highly activated to undergo oxidative S_NAr^H by N, O and S types of nucleophiles. By choosing different nucleophiles and by adjusting the reaction conditions, a large number of 6- or 9- substituted derivatives and 6,9-bissubstituted derivatives could be conveniently prepared. Additionally, hydrolyzing the two cyano groups of compound **1** would result in another 5-member dicarboxylic acid imide ring fused to the phenalene plane. More importantly, aromatic hydrogen-6 of the NH imde **2** and the alkylated imdes **3** inherited the oxidative S_NAr^H reactivity from the parent **1**.

Sincethe discovery of the polycyclic aromatic parent **1** as well as its unique oxidative S_NAr^H reactions, the family of phenalene dicarbonitrile have already attracted much attention for their rigid a planar structures, convenient syntheses and modifications, diverse fluorescence properties, biological activities and so on. As typical ICT fluorophores with long wavelength and high photostability, they have proved to be good candidates for the construction of molecular or nanoparticle probes/sensors. And as drug leads, at least two types of highly potent inhibitors specifically targeting different tumor regulating proteins, *i.e. pan*-BCl-2 and FGFR1, have been rationally developed.

It is believed that the interests in this promising chemical family will last for long time and the investigations to promote their application will go on. From the viewpoint of fluorescence imaging, currently, the important and forefront fields are single molecular imaging and supraresolution imaging which require the fluorophore to be highly photostable and intensive fluorescent; these are exactly the characteristics of this family's members; thus, they are highly recommendable for these sophisticated imaging applications. From the viewpoint of drug leads, identification of the targeted biological macromolecules, especially, the functional proteins, are extremely important; To this date, the work to find the potential targets of this new chemical family is still very limited; The combination of the modern bioinformatics and high-throughput drug screening technology may help to identify more targets other than *pan*-BCl-2 and FGFR1, and thus open new avenues to efficient antitumor agents; and combinatorial synthesis based on oxidative S_NAr^H of compound **1**, **2** or **3** should be utilized to provide large libraries of compounds for SAR study and structure optimization.

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