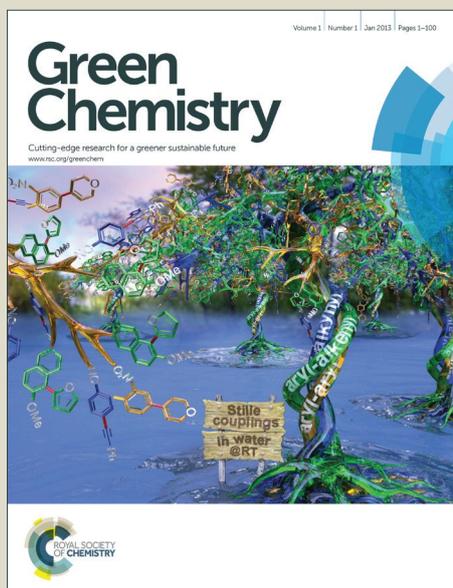


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COMMUNICATION

Iron-Catalyzed Olefin Hydrogenation at 1 bar H₂ with a FeCl₃/LiAlH₄ Catalyst

Cite this: DOI: 10.1039/x0xx00000x

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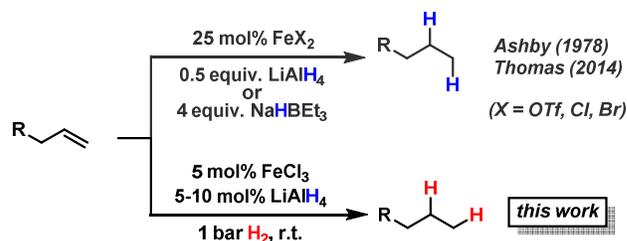
Received 00th January 2014,
Accepted 00th January 2014

DOI: 10.1039/x0xx00000x

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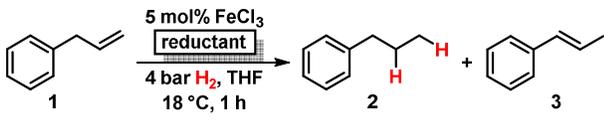
The scope and mechanism of a practical protocol for the iron-catalyzed hydrogenation of alkenes and alkynes at 1 bar H₂ pressure were studied. The catalyst is formed from cheap chemicals (5 mol% FeCl₃/LiAlH₄, THF). A homogeneous mechanism operates at early stages of the reaction while active nanoparticles form upon ageing of the catalyst solution.

Catalytic hydrogenations of olefins constitute one of the strongholds of transition metal catalysis within organic synthesis and technical processes.¹ The majority of methods involve noble metal catalysts based on Pd, Pt, Rh, Ir or toxic metals such as Ni or Co. Iron-catalyzed hydrogenations of olefins have only recently attracted great interest due to their expedient economic and environmental qualities.² Homogeneous iron catalysts were mostly reported with phosphine and pyridyl-2,6-diimine ligands, sometimes requiring high pressures of H₂.^{3,4} Nanoparticle Fe catalysts could be prepared by reduction of iron salts with Grignard reagents in the absence of a suitable ligand or by decomposition of iron carbonyls.⁵ Fe-catalyzed reductions of olefins were recently reported with cheap ferrous salt pre-catalysts FeX₂ in the presence of an excess of lithium *N,N*-dimethylaminoborohydride (10 equiv.) or sodium triethylborate (4 equiv.) and required a high catalyst loading or the addition of tetradentate ligands.⁶ Reductions of alkenes and alkynes with LiAlH₄ in the presence of various transition metal halides (NiCl₂, TiCl₂, CoCl₂, FeCl₃) were already reported in the 1960s and postulated to involve metal hydride species that engage in formal hydrometalations of the olefin.⁷ Here, we wish to present a synthetic and mechanistic study on a hydrogenation protocol using catalytic amounts of a cheap Fe salt and catalytic amounts of lithium aluminiumhydride (LiAlH₄) as catalyst activator under an atmosphere of 1 bar H₂ as stoichiometric hydrogen source (Scheme 1).^{7e} This method allows the use of standard (ambient pressure) equipment. H₂ is an abundant raw material; LiAlH₄ is an easy-to-handle reductant with numerous applications.⁸



Scheme 1 Iron-catalyzed reductions of olefins: Hydride vs. hydrogen methods.

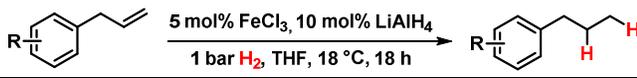
Reaction conditions and substrate scope. Initial experiments with the model substrate allylbenzene (**1**) aimed at the identification of a suitable catalytic reductant which assists the formation of a low-valent iron catalyst (with dark brown colour) from the commercial pre-catalyst FeCl₃ (Table 1).⁹ LiAlH₄ displayed excellent selectivity which exceeded that of earlier protocols with Grignard reagents.⁵ Isomerization of the terminal double bond into conjugation – which occurred in the related EtMgCl-mediated protocols (entries 2, 4) – was effectively suppressed.¹⁰ NaBH₄ was far less active even at elevated temperature and pressure (entries 6, 7). Interestingly, low ratios of LiAlH₄/FeCl₃ (1/1 to 2/1) fared optimal in the hydrogenation of **1** at 1 bar H₂. When employing a larger excess of LiAlH₄ (>2/1), the catalytic activity collapsed.^{7e} This stoichiometry differs from literature reports where large excess amounts of hydride reagents effected clean hydrogenations of olefins.^{6,7a-c} At 60°C, the FeCl₃/LiAlH₄ catalyst decomposed upon decolorization. The catalyst system comprises of cheap and easy-to-handle reagents (FeCl₃ or FeCl₂, LiAlH₄, THF); the reaction operates under ambient conditions (1 bar H₂, 20°C), which make the general protocol practical for every-day use in standard synthesis laboratories. The optimized conditions were applied to functionalized allylbenzenes and styrenes (tables 2 and 3).⁹

Table 1 Selected optimization experiments^a


Reductant (mol%)	Deviation from conditions	2 (%) ^b	3 (%) ^{ba}
1	-	0	2
2	EtMgCl (30)	42	56
3	EtMgCl (30)	16	0
4	EtMgCl (30)	60	36
5	Et ₂ Zn (20)	4	1
6	NaBH ₄ (100)	8	0
7	NaBH ₄ (100)	45	38
8	LiAlH ₄ (10)	97	3
9	LiAlH₄ (10)	98	1
10	LiAlH ₄ (30)	95	3
11	LiAlH ₄ (10)	96	1
12	LiAlH ₄ (10)	20	15

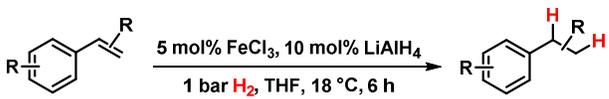
^a Conditions: 5 mol% FeCl₃ in THF (0.5 mL) under argon, addition of reductant at r.t., after 10 min addition of **1**, after 1 min exchange of Ar with 4 bar H₂; ^b quantitative GC-FID (internal reference); ^c during catalyst preparation.

Allylbenzenes underwent only minimal olefin isomerization.¹⁰ Styrenes exhibited low propensity to undergo polymerization (entry 13, table 3). The general protocol is compatible with several functional groups including F, Cl, Br, allyl and benzyl ethers, esters, carboxamides, pyridines and anilines. Clean hydrogenation was achieved with bulky, *ortho*-substituted, and electron-rich styrenes. For comparison, the FeCl₃/EtMgCl-derived catalyst effected undesired dehalogenation (Cl, Br)¹¹ and allylether cleavage¹², and showed no activity in the presence of carboxylates or cinnamates. Catalyst decomposition was effected by nitro groups, iodides, nitriles, ketones, and acidic protons (e.g. alkanols, pK_a~17), presumably by oxidation to catalytically incompetent Fe(II) species (decolorization). Tri-substituted styrenes gave low conversions. In general, bulky and functionalized substrates were more reactive at elevated pressures (10 bar H₂).¹³

Table 2 Hydrogenation of allylbenzenes at 1 bar H₂.


Allylbenzene	R	Yield (%) ^a
1	H	93
2	Me	79 (92) ^b
3	OMe	84 (89) ^c
4	Me	95
5	OAc	93 (95) ^{c,d}
6		99
7	OMe	100
8	F	86

^a quantitative GC-FID (internal reference); conversion (%) in parentheses if <95%; ^b 7% 1-propen-1-ylbenzene (*E/Z* 9/1); ^c 24 h; ^d by NMR (vs. CH₂Br₂).

Table 3 Hydrogenation of styrene derivatives.


Styrene	R	Yield (%) ^a
1	H	100
2	Me	98
3	OMe	98
4	Cl	93 ^b
5	Br	94 ^b
6	F	77 ^c
7	OBn	100 ^d
8	NH ₂	97 ^c
9	CO ₂ Me	97
10	H	100 ^e
11	OMe	84
12		100
13		86 ^d
14	H	100 ^d
15	Cl	85 (89) ^f
16	Br	44 (56) ^f
17	Br	92 ^{c,f}
18	OMe	100 ^d
19	Cl	74 (86) ^b
20	OBn	100 ^d
21	Ph	100 ^d
22	Bn	100 ^d
23	CO ₂ Et	58 (68) ^c
24		48 (54)
25		88 ^c
26		33 (58) ^e
27		75 ^{d,g}
28		18 (18) ^c

^a quantitative GC-FID (internal reference); conversion (%) in parentheses if <95%; ^b <12% ethylbenzene; ^c 20 h, 10 bar H₂; ^d 5 mol% LiAlH₄, 3 h; ^e 20 h; ^f <5% cumene; ^g mixture of partial and full hydrogenation products (~6/1)

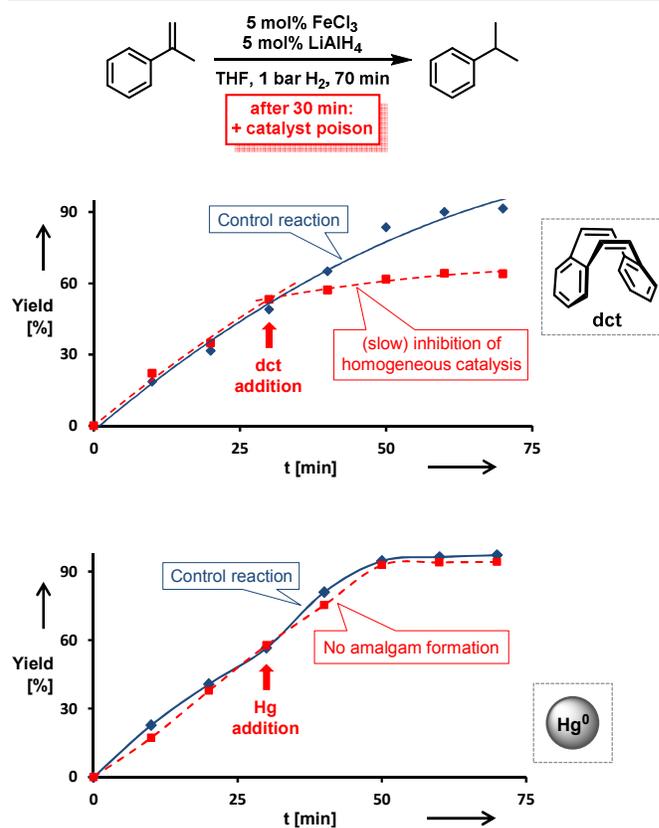
Hydrogenations of aliphatic alkenes were also catalyzed by FeCl₃/LiAlH₄ under similar conditions (Table 4).⁹ Terminal olefins were only slowly isomerized (~10%).¹⁰ Surprisingly, substrates containing moderately acidic protons (pK_a ~25)¹⁴ underwent hydrogenation with high selectivity (entries 10-13).¹⁵ Alkynes underwent *Z*-selective semi-hydrogenation,¹⁶ whereas complete hydrogenation to the alkanes was observed at longer reaction times or elevated pressures.

Table 4 Hydrogenation of other alkenes and alkynes.

Substrate	Product	Yield (%) ^a
1		82 ^b
2		69 ^b
3		82 ^c
4		89 (89) ^c
5		65 (65) ^c
6		21 (41) ^b
7		100 ^d
8		2 (10)
9		64 (65) ^d
10		12 (16)
11		96 ^{d,e}
12		38 (38) ^{d,e}
13		69 (69) ^{d,e}
14	R = H	100 ^c
15	CO ₂ Me	92 ^c
16		75 (80) ^g
17		90 ^g

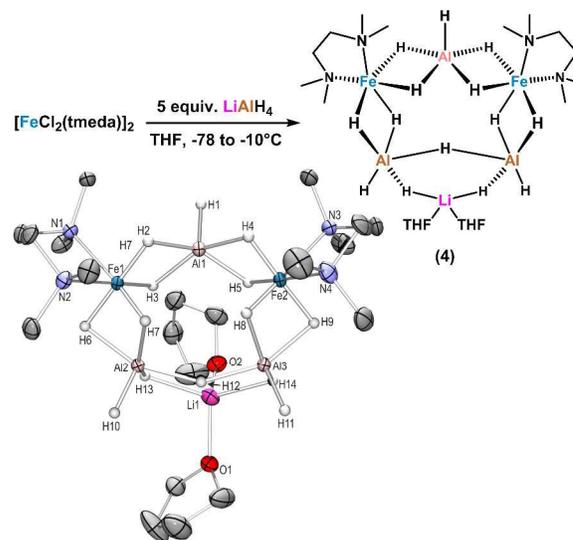
^a quantitative GC-FID (internal reference); conversion (%) in parentheses if <95 %; ^b alkene isomers; ^c 20 h; ^d 10 bar, 20 h; ^e 10 mol% LiAlH₄; ^f 60 °C; ^g by NMR (vs. CH₂Br₂).

Mechanistic studies. The distinction between homogeneous and heterogeneous catalysts is a challenging task.¹⁷ However, kinetic experiments with selective poisons can provide valuable information on the topicity of the catalyst species. We have performed two sets of poisoning experiments which appear to support a homogeneous mechanism. Dibenzo[*a,e*]cyclooctatetraene (dct) is a selective ligand for homogeneous metal species due to its rigid tub-like structure and π -acceptor properties.¹⁸ Upon addition of 30 mol% dct (6 equiv. per [Fe]) to the hydrogenation of isopropenylbenzene at 1 bar H₂ after 30 min, the catalyst activity was significantly inhibited (Scheme 2, top).^{9,19} A similar conclusion can be derived from a poisoning experiment with 3 equiv. Hg (60 equiv. Hg per [Fe]). A potential amalgam formation²⁰ was not observed and no significant change of the catalyst activity was observed in comparison with the control reaction (Scheme 2, bottom).⁹ These results suggest the operation of a homogeneous catalyst species during the early stage of the catalytic hydrogenation. Previous studies showed that the reaction of FeCl₃ with an excess of LiAlH₄ ultimately leads to the formation of iron metal and AlH₃ via the intermediate formation of a thermally unstable iron(II) compound with the composition Fe(AlH₄)₂.^{21,22} In an attempt to gain deeper insight into the catalyst species operating



Scheme 2 Top: Poisoning experiment with 30 mol% dibenzo[*a,e*]cyclooctatetraene (dct, dashed curve) vs. control reaction (solid line). Bottom: Poisoning with 3 equiv. Hg (dashed) vs. control reaction (solid line).

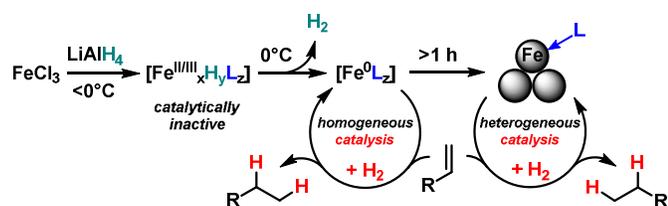
in homogeneous solution, we treated [FeCl₂(tmeda)]₂ (tmeda = *N,N,N',N'*-tetramethylethylenediamine) with LiAlH₄ at -70 °C and obtained dark red crystals of the oligohydride compound [Li(thf)₂{Fe(tmeda)}₂(AlH₅)(Al₂H₉)] (4, Scheme 3).⁹ The hexa-metallic macrocyclic cage contains 14 bridging hydrido ligands and two Fe atoms with distorted octahedral coordination geometries. Unfortunately, the thermal instability prevented further spectroscopic characterization.



Scheme 3 Synthesis of the soluble LiAlFe-oligohydride complex 4.

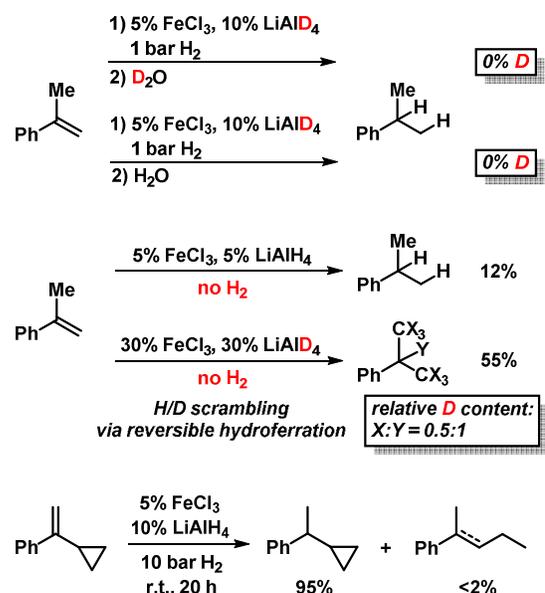
However, complex **4** showed no activity in hydrogenations of styrenes (1-10 bar H₂, -10°C) and maintained its red colour throughout the reaction. Above -10°C, the complex rapidly decomposed upon H₂ evolution to give a brown paramagnetic species which afforded good yields in hydrogenations at 20°C and 4 bar H₂. The crystallographic characterization of **4** documents that this or similar oligonuclear Fe(II) aluminohydride complexes may be intermediates *en route* to the formation of catalytically active low-valent iron species.²³

The initially homogeneous dark-brown catalyst species (possibly in the oxidation states 0 and/or +1)²³ experience rapid ageing and particle formation after appr. 1 h under the reductive conditions. Several methods of synthesis and characterization techniques of naked Fe(0) nanoparticles (prepared by reduction of ferric and ferrous halides) have been reported.^{5,7,23,24} DLS measurements (dynamic light scattering) of freshly prepared catalyst solutions (5 mol% FeCl₃/LiAlH₄, THF, r.t., 10 min, then 100 nm nanofiltration) documented the presence of poly-disperse particles of 250-1500 nm size after 30 min of ageing under anaerobic conditions in the absence of substrates. The aged species are much less catalytically active than their homogeneous counterparts. Catalyst solutions (FeCl₃/LiAlH₄ (1/1) in THF) stored at 0°C under argon for 6 h, 24 h, and 48 h afforded 42%, 12%, and 5% conversion of α -methylstyrene under standard conditions (see entry 16 in table 2), respectively. We postulate a homogeneous mechanism of soluble, low-valent iron catalyst in the initial stage of the hydrogenation reactions (Scheme 4). Such species form by reduction of FeCl₃ (or L_nFeCl₂) with LiAlH₄ at above 0°C and are typically characterized by the dark brown colour. The absence of suitable ligands leads to the formation of Fe(0) nanoclusters^{5,22,24} which require higher H₂ pressures than the homogeneous species to maintain catalytic activity.



Scheme 4 Proposed formation and catalysis of low-valent iron species.

Deuterium incorporation was observed at higher catalyst concentrations (30 mol% FeCl₃/LiAlD₄) in the absence of H₂ which gave ~55% hydrogenation product (Scheme 5, center).⁹ Such H₂-free conditions can effect H/D scrambling in the starting material and product (via reversible hydroferration) and the formation of radical intermediates (with participation of THF as H donor).⁹ However, the radical mechanism is very unlikely to operate under hydrogenation conditions in the presence of H₂ gas (Scheme 5):⁹ Reaction work-up with deuterium oxide (D₂O) and employment of lithium aluminium-deuteride (LiAlD₄) showed no deuterium incorporation into the products, respectively (Scheme 5, top right). Further, the intermediacy of free C-radicals is unlikely: Employment of the radical probe 1-cyclopropyl-1-phenylethylene²⁵ resulted in less than 2% ring opening (Scheme 5, bottom).⁹ The hydrogenation of various styrenes (1 bar H₂) was unaffected by the presence of 1 equiv. 1,1-diphenylethene. On the other hand, the addition of TEMPO (2,2,6,6-tetramethylpiperidinyloxy, 1 equiv.) inhibited conversion of α -methylstyrene (no TEMPO adduct detected), possibly by irreversible catalyst oxidation as indicated by the decolorization of the solution.



Scheme 5 Mechanistic studies with deuterated reagents (top), in the absence of H₂ (center), and with radical probe (bottom).

Conclusions

In summary, we have studied the iron-catalyzed hydrogenation of various styrenes, alkenes, and alkynes under an atmosphere of 1 bar H₂. This method uses cheap and easy-to-handle reagents (FeCl₃, LiAlH₄, THF, H₂) which allow facile implementation in standard synthesis labs. Alkynes underwent Z-selective semi-hydrogenation. Sterically hindered and functionalized olefins showed higher conversions at elevated H₂ pressures. Mechanistic studies support the notion of a homogeneous catalyst species at the outset of the hydrogenation reactions (<1 h) while catalyst ageing results in the formation of particles which exhibited somewhat lower catalytic activity. The crystallographically characterized homogeneous Fe(II) oligohydride complex **4** can serve as starting point for further model catalyst preparations.

Acknowledgement

We thank the Evonik foundation (T.N.G.) and the Fonds der Chemischen Industrie (M.V.) for doctoral fellowships.

Notes and references

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† **General procedure:** A 10 mL vial was charged with a freshly prepared solution of FeCl₃ in THF (0.50 mL, 0.05 M) and an aliquot of a vigorously stirred suspension of LiAlH₄ in THF (0.50 mL, 0.1 M) under an argon atmosphere. After stirring the dark brown mixture for 10 min (during which H₂ evolution can be observed), the alkene (0.50 mmol) was added and the vial purged with dihydrogen gas (1 min). For reactions under higher H₂ pressures, the vial was transferred to a high pressure reactor (ParrTM), the reactor purged with H₂ (1 min), sealed, and the

internal pressure adjusted to 1 bar. After 3–20 h at room temperature, the reaction was quenched with saturated aqueous NaHCO₃ (1 mL) and extracted with ethyl acetate (2×2 mL). The organic phases were dried (Na₂SO₄) and subjected to flash chromatography (SiO₂, pentane/ethyl acetate) or analyzed by quantitative GC-FID analysis vs. pentadecane as internal reference.

Electronic Supplementary Information (ESI) available: See DOI: 10.1039/c000000x/

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