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ARTICLE

MoS₂ Nanosheet Production by the Direct Exfoliation of Molybdenite Minerals from Several Type-Localities

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Samples of the mineral molybdenite from three classic molybdenum mining localities were examined as a potential source of molybdenum disulfide (MoS₂) nanosheets. In all cases, ultrasonication-promoted exfoliation of these samples in *N*-methylpyrrolidone (NMP) was found to produce MoS₂ as dispersed nanosheets with lateral sizes in the range of 200-600 nm and thicknesses between 1 and 10 atomic trilayers. The MoS₂ nanosheets obtained were found to be highly crystalline and largely defect-free, but tend to contain small amounts of aggregates on their surfaces. The exfoliated MoS₂ dispersions were characterised by UV-Vis spectroscopy, Raman spectroscopy, scanning electron microscopy (SEM), (scanning) transmission electron microscopy ((S)TEM) and energy dispersive X-ray (EDX) spectroscopy. This work raises the possibility that mined, unrefined minerals could be a source of low-dimensional MoS₂.

Introduction

Atomically thin graphene exhibits exceptional electronic, optical and mechanical properties.¹ Since graphenes' initial discovery² there has been an explosion of interest³ in other 2D materials, leading to the exploration of graphene analogues such as the family of transition metal dichalcogenides, or TMDCs (MX₂; *e.g.* M = Mo, W and X = S, Se, Te). TMDCs all display similar layered-graphitic superstructures.⁴ These structures are comprised of a layer of transition metal atoms coordinated by two layers of close-packed chalcogen atoms, held together by strong covalent bonds. In the bulk material, these sheets are bound together by weak van der Waals forces, just as in graphite.⁵ Of these 2D-TMDC materials, MoS₂ nanosheets are potentially a useful semiconducting component in a series of devices.^{6,7} In addition, molybdenum disulfide exhibits considerable anisotropy, giving rise to novel electronic, optical, structural and mechanical properties in its 2D form (2D-MoS₂). Due to these properties 2D-MoS₂ has been shown to be a suitable material for inclusion in that in turn gives rise to a myriad of potential applications in photochemical sensors,^{8,9} optoelectronics,^{10,11} bioimaging,^{12,13} catalysis,^{14,15} solid-state lubricants,¹⁶ and energy storage.^{17,18}

In many cases, laboratory-scale 2D-MoS₂ production is based on liquid exfoliation of molybdenum(IV) sulfide powder.¹⁹ The powder is refined from natural MoS₂,

molybdenite, which is mined in numerous pan-global locations and purified by comminution, flotation and leaching.^{20,21} In the years 2006 - 2011, demand has surged to 264,000 tonnes per year of molybdenum metal, an increase of 25%.^{22,23} The demand of the mineral is driven by metallurgical, green- and now nanomaterial-based applications.²⁰

Herein we discuss the feasibility of MoS₂ nanosheet production by exfoliation of molybdenite ores, as exemplified by samples obtained from three locations (Queensland AUS [QL], New South Wales AUS [NSW] and Telemark NOR [TM]). The crystalline molybdenite samples were obtained untreated from the selected minerals (see Table S1 for purity analysis). The potential to obtain MoS₂ nanosheet from these sources by employing the Coleman method of MoS₂ exfoliation is detailed in this report.^{24,25} The nanomaterials produced have been subject to a comparison of exfoliation efficiency, as well as compositional and structural features, relative to the nanosheets produced from MoS₂ powder documented in the literature.²⁴

Results and Discussion

50 mL solutions of *N*-methylpyrrolidone (NMP) containing flakes of the selected molybdenite (375 mg, 7.5 mg/mL concentrations) were bath ultrasonicated for 48 hours, with aliquots obtained at various time intervals. Unexfoliated

material was removed from the aliquots by centrifugation (45 minutes at 1500 rpm). UV-visible spectra of the supernatants showed two peaks with maxima at 628 and 674 nm, attributed to the characteristic A and B direct excitonic transitions of MoS₂ respectively (Figures 1, S1a and S2a).²⁴

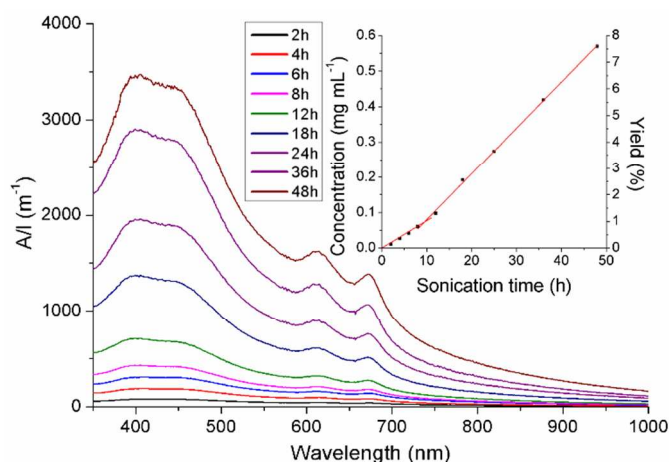


Figure 1. UV-Vis absorption spectrum stack plot of MoS₂ dispersions obtained from the **QL** concentrate at different time intervals. The inserted graph shows the concentration of the MoS₂ dispersion from the **QL** concentrate, relative to sonication time.

The UV-Vis data was modified to remove the background noise associated with the scattering exponent of the MoS₂ nanosheets.²⁴ The intensity of the feature at *ca.* 674 nm was then used as a spectroscopic handle to estimate the concentrations of the MoS₂ dispersions produced.²⁵ The final concentrations of the exfoliates were found to be 0.57, 0.42 and

0.84 mg mL⁻¹, from **QL**, **NSW** and **TM** precursors, respectively. These values obtained are considerably lower than the concentration of 1.79 mg mL⁻¹ produced from the exfoliation of bulk-MoS₂ powder, which was run alongside the molybdenite samples. In all cases, plots of concentration versus sonication times (insert in Figure 1, figures S1b and S2b) reveal at least two kinetic regimes of MoS₂ nanosheet production, as the exfoliation rates (k_{exf}) in the initial 12 hours of sonication are noticeably slower than those observed after this period. Using the values of k_{exf} obtained after this initial 12 hour period, it was clear to see the selection of the concentrate results in different exfoliation rates (k_{exf} between 9.4 and 20.7 $\mu\text{g mL}^{-1} \text{h}^{-1}$; Table S2). The rates of exfoliation were also found to be significantly lower than those observed for the exfoliation of bulk-MoS₂ powder (37.1 $\mu\text{g mL}^{-1} \text{h}^{-1}$). Stability investigations show that there is a very small (< 4 %) reduction in the intensity of the A and B bands over a period of 4 weeks; indicating that the dispersions are highly stable (Figure S3).

In all molybdenite sources, both large flakes from the crystalline aggregates and the exfoliated nanosheets exhibit two Raman peaks that correspond to the A_{1g} and E_{2g} modes (Figure 2a-c). These peaks are known to exhibit a well-defined dependence on layer *z*-thickness and thus can be used as a way of estimating nanosheet thickness.²⁶ All of our exfoliations gave few-layer MoS₂ nanosheets as the predominant product. Nanosheets from the **TM** dispersion showed the two modes separated by 19 cm⁻¹, similar to the peak separation observed for single-layered MoS₂ nanosheets exfoliated from MoS₂ powder.^{24,27} It was expected that exfoliation of the **QL** and **NSW** concentrates also produced single-layered 2D-MoS₂, however, these were not detected in Raman spectra obtained

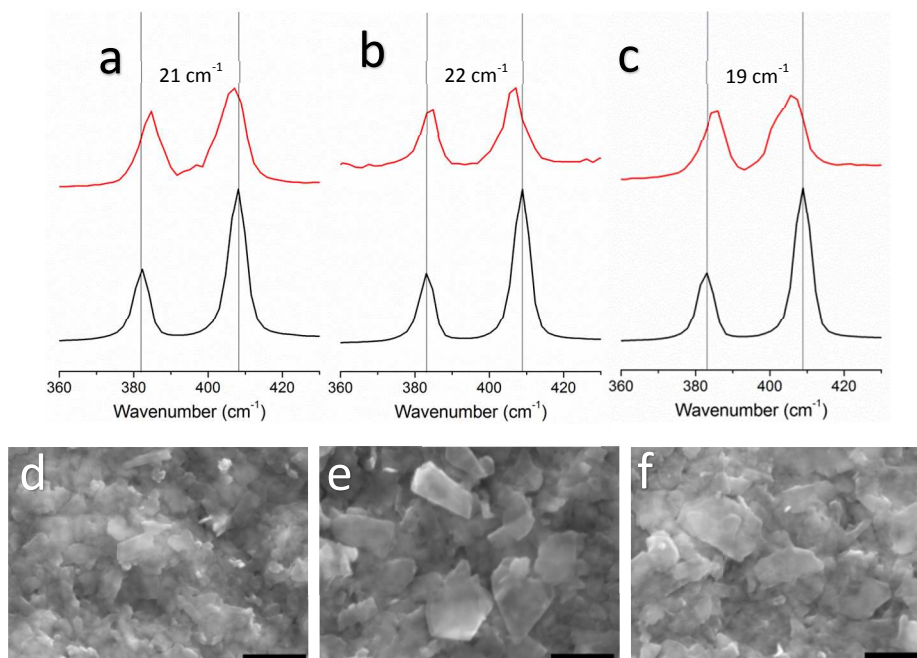


Figure 2. (a-c) Selected range (360-430 cm⁻¹) Raman spectra for **QL**-, **NSW**- and **TM**-sourced materials, respectively. In each spectra, a comparison between the bulk material (black) and the exfoliated nanosheets (red, separation of peaks = 26 cm⁻¹) can be made. (d-f) SEM images of the MoS₂ thin films on a PVDF membrane from the (d) **QL**-, (e) **NSW**- and (f) **TM**-sourced dispersions. Scale bars set at 500 nm.

from either exfoliated sample, presumably due to the smaller lateral dimensions of the nanosheets. Statistical analysis (sample size = 50, figure S4) of randomly selected flakes indicates that more than 80% of the flakes dispersed contain less than 6 stacked monolayers. The remaining areas were covered with thicker material whose thickness could not be determined by Raman spectroscopy.²⁶

Membrane-supported thin films of MoS₂ nanosheets were prepared from the **QL**, **NSW** and **TM** dispersions by vacuum filtration^{24,28} and imaged by scanning electron microscopy (SEM). All films produced were found to have similar surface morphology (Figures 2d-f and S5); the surface of the films consist of partially-stacked 2D-nanosheets with lateral dimensions between 200-600 nm.

Nanosheets were deposited onto lacey carbon grids for (scanning) transmission electron microscope ((S)TEM) imaging. The resulting HAADF STEM images (Figure 3) and TEM images (Figure S6) show faceted morphologies and a range of irregular shapes. The lateral dimensions of the nanosheets ranged from 200 nm to greater than 1 μm . Selected area diffraction (SAED) patterns (Figure S6) confirmed that in all samples, most flakes consisted of single crystals and that all of the flakes had the expected hexagonal crystal structure. Atomic resolution, high-angle annular dark-field (HAADF) STEM imaging (Figures 2 and S7) confirmed that all flakes were highly crystalline and largely defect-free. The edge structure of flakes observed in high magnification HAADF STEM images can be used to estimate flake thickness.²⁴ The majority of the flakes were found to be 1-10 trilayers thick, although flake thicknesses varied for different areas within individual flakes.

Energy-dispersive X-ray (EDX) spectroscopy measurements, performed in the SEM and STEM, were used to obtain compositional information from both the thin films and individual flakes. Quantitative analysis gave the expected Mo:S ratio of 1:2 in all cases. STEM-EDX mapping revealed calcium rich regions in all samples, believed to be nanometre scale calcium carbonate aggregates on the flake's surface (Figure S8-11). Nanosheets prepared from the **NSW** source showed significantly more calcium contamination than nanosheets made from the other sources. SEM-EDX analyses did not show any significant quantities of calcium, but did detect other elements: all samples were found to contain non-negligible amounts of iron, in addition to bismuth found in **QL** and **NSW** thin films (Figures S8-10).

Conclusions

We have shown that untreated molybdenite mineral samples can be used as a precursor for exfoliation to produce few-layered MoS₂ nanosheets that are highly crystalline, and largely defect-free, with small amounts of aggregates on the surface of the sheets. The selection of molybdenite source from various locations worldwide appears to have little effect on the purity or crystallinity of the nanosheets produced. In addition thin films

comprised of stacked MoS₂ nanosheets have been produced using a simple filtration method. The relatively high purity of the nanosheets and thin films produced by the methodology described herein gives hope that a cost-effective approach for the production of MoS₂ for components in electrical circuitry using relatively unrefined material is possible. The electronic properties of the nanosheets produced are currently the subject of further work.

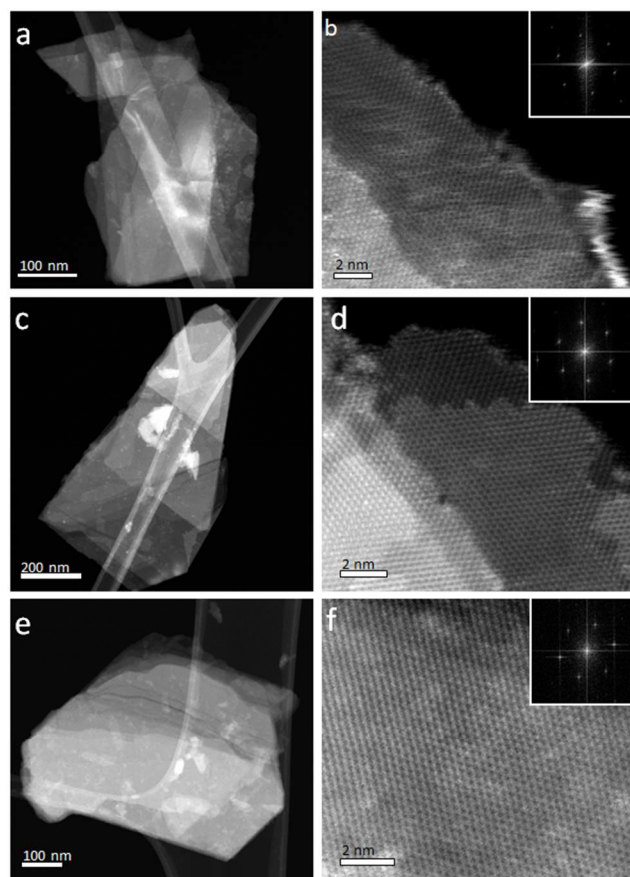


Figure 3. HAADF STEM images of a representative selection of nanosheets obtained from **QL** (a, b), **NSW** (c, d) and **TM**-sources (e, f). Fourier transforms for each of the lattice resolution images are inset for b, d and f and illustrate the expected single crystal lattice symmetry.

Experimental Section

Background Information

The molybdenite aggregates were obtained from selected molybdenite-containing mines: Toowong, Queensland (AUS); Kingsgate, New South Wales (AUS) and Siljord, Telemark (NOR). Compositional analysis of the aggregates were obtained by a Chemeca Camebax SX100 microsonde Electron Probe Micro-Analyser (EPMA), running a 20 nA beam at 20 keV with a 1.5 μm spot size on polished sections. The measurements were compared to elemental standards: **Ni**, **Co**, **Fe**, **Re**, **W** and **Bi** metals, **MoS₂**, **CuFeS₂**, **ZnS** and **ZnSe**.

Molybdenum(IV) sulfide powder and *N*-methylpyrrolidone (NMP) were purchased from Sigma-Aldrich, which were used as supplied.

The dispersions were prepared in a bench-top sonication bath (Elmason P 70H [HF power 820 W across four horns]). Centrifugations were carried out using Centurion 822 series (up to 15 mL dispersions) and Jouan B4i (up to 50 mL dispersions) centrifuges. UV-Vis absorption spectroscopy was performed with an Agilent 8453 spectrometer using a 1 cm glass cuvette. Accurate dilutions of the MoS₂ dispersions in NMP were produced to obtain suitable UV-Vis spectra. The subsequent data was then analysed after modification (implementing the Beer-Lambert Law) to represent the absorption intensities at the concentrations when the aliquots were first taken from the reaction.

Membrane-supported thin films of MoS₂ nanosheets suitable for Scanning Electron Microscopy (SEM) were produced by the dilution of 30 mL suspensions in *isopropanol* (270 mL), and filtration through a polyvinylidene fluoride membrane (PVDF, pore size 0.1 μm). The thin films were washed with *isopropanol* (3 × 50 mL) before drying in a vacuum chamber for 12 hours at ambient temperature. The thin films are carbon-coated using a Gatan Model 682 Precision Etching coating system before carrying out SEM and EDX analyses by the use of Philips XL 30FEG scanning electron microscope and a DX4 detector.

Raman spectra were acquired on a Renishaw 1000 system, with a solid state 50 mW 514.5 nm laser. The laser beam was focused onto the MoS₂ sample by a 100× objective lens. The scattered signal was detected by an air cooled CCD detector. Samples were prepared by dilution of the NMP dispersions with ethanol (1:40 v/v), before deposition onto a 300 nm SiO₂/Si substrate and drying under ambient conditions over 48 hours.

Samples were prepared for transmission electron microscope (TEM) imaging by drop casting a dilute solution of the MoS₂ dispersions (1 % solution in EtOH) onto holey carbon support films which were then washed with ethanol and air dried. Flake size and shape was assessed using low magnification bright field TEM in a Philips CM20 and a FEI Tecnai F20, both microscopes used a LaB₆ electron source and were operated at 200kV. High angle annular dark field (HAADF) scanning transmission electron microscope (STEM) was performed using a probe side aberration corrected FEI Titan G2 operated at 200kV. For HAADF imaging a convergence angle of 26 mrad, a HAADF inner angle of 52 mrad, and a probe current of ~200 pA were employed. Energy dispersive X-ray (EDX) spectroscopy was performed in the Titan using a Super-X EDX detector with a solid angle of collection of 0.7 sr. EDX quantification was performed using Bruker Esprit software, where the bremsstrahlung background was subtracted, peaks deconvoluted, and quantification performed using the Cliff-Lorimer approach. Fe, Cu, Zr, and C counts were known to originate from the microscope (Fe and Zr) and support film (Cu and C), so these elements were

included in the deconvolution procedure but excluded from the final quantification.

Purity of Molybdenite Aggregates

The molybdenite aggregates were analysed by Electron probe Microanalysis (EPMA) and found to be of high purities (> 99.5 %; Table S1). The trace element populations are low (below detection limits in many analysis spots), nor were there any major substitution. Bi was detected in the sample from **QL** where Bi₂S₃ was an associated phase. The high standard deviations reveal the high variations. The Re and W do appear in several points in the analyses above detection and are probably real.

Synthesis of MoS₂ Dispersions from the Aggregates

In a 60 mL screw-cap glass jar, NMP (50 mL) was added to 375 mg of the chosen molybdenite aggregate (manually broken up into approx. 10 flakes) before suspending the sealed jars in an operational ultrasonic bath for 48 hours. For concentration determination, 3 mL aliquots of the dispersions were removed at selected time intervals and diluted in NMP (3 mL) within a 15 mL centrifuge tube. Each aliquot was left to settle overnight before centrifugation at 1,500 rpm for 45 minutes. Two-thirds of the resulting supernatant was pipetted to provide a suspension of MoS₂ nanosheets in NMP (note: at 50% of its initial concentration). After 48 hours the dispersions that remained in the reaction vessels (approx. 25 mL) were decanted into 50 mL centrifuge tubes containing 20 mL NMP and left to settle overnight. Centrifugation at 1,500 rpm for 45 minutes followed by pipetting of 30 mL of the supernatant gave a suspension of MoS₂ nanosheets of suitable quantities for thin film production.

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Notes and references

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Electronic Supplementary Information (ESI) available: UV-Vis, Raman, SEM, HAADF STEM and EDX data of the nanosheets produced from **QL**, **NSW** and **TM**-sourced materials. See DOI: 10.1039/b000000x/

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Table of Contents Entry



This work raises the possibility that the unrefined mineral molybdenite could be used as a cost-effective source of MoS₂ nanosheets.