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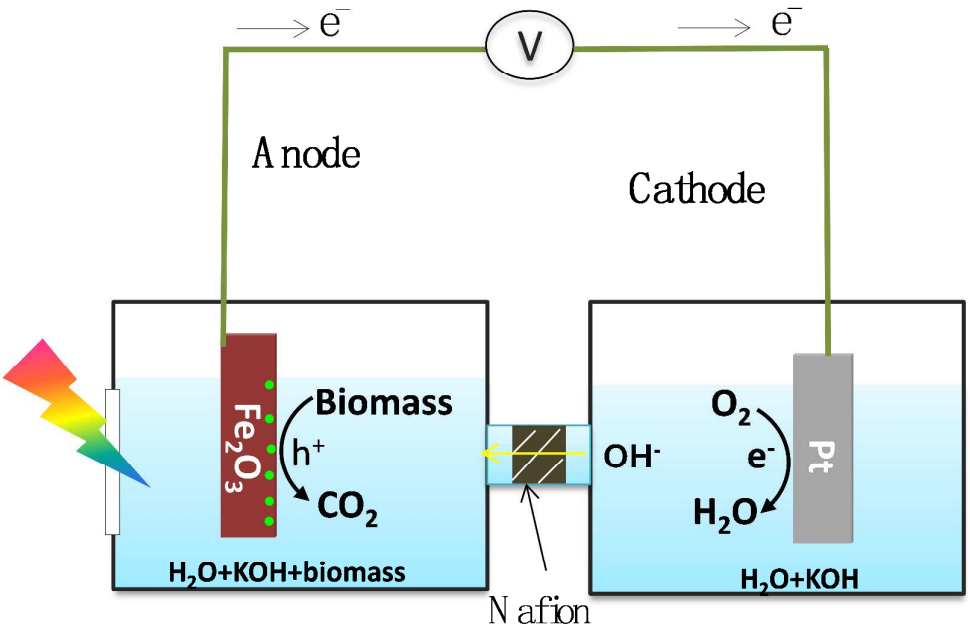
Abstract

Photofuel cells have been demonstrated to be a promising strategy for generating electricity using biomass. Here we present a photofuel cell with a visible light α-Fe_2O_3 based photoanode that can be directly powered by varieties of biomasses, such as methanol, glycerol, glucose, cellubiose and starch. The photocurrent density and power density of the photofuel cell are significantly enhanced by loading cocatalyst (metal hydroxide, e.g. $Ni(OH)_2$) on the α-Fe_2O_3 photoanode. The power density of the photofuel cell powered by glucose is enhanced over two times, from 0.082 mA cm^{-2} for α-Fe_2O_3 to 0.18 mW cm^{-2} for $Ni(OH)_2$/α-Fe_2O_3 photoanode.



Transition Metals (Ni, Fe, and Cu) Hydroxides Enhanced α-Fe_2O_3 Photoanode-based Photofuel Cell

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Introduction

The utilization of abundant biomass as an energy source can prevent environmental pollution and reduce the dependence on fossil resource.[1, 2] The technological challenge is to sustainably capture the biomass energy using green process. Direct alcohol fuel cells (DAFCs) with potentials of high power density and pollution-free, offer a possible solution to this problem. [3, 4] Although DAFCs powered by low molecular weight alcohols (usually methanol or ethanol) have been demonstrated to be practically feasible, biomass-derived compounds such as glucose, polysaccharide, and cellulose have not been reported as fuels for DAFCs. This is due to the low activity of catalysts in the activation of the C–C bonds and oxidation of biomass at low temperatures.[5, 6] Biofuel cells can be powered by various biomasses at low temperatures; however, low electric power output, limited lifetime and rigorous reaction conditions seriously hinder their applications.[7, 8]

Photofuel cell (PFC) mainly consists of a semiconductor photoanode, a cathode, electrolyte and fuels.[9-14] With PFC technology, biomass as fuel can be oxidized on photoanode under light irradiation and oxygen is reduced on counter electrode through an external circuit, generating electricity. In essence, PFC is different from fuel cells reported previously, and it widens the range of biomass fuels for electricity generation. As it has been proved, various biomasses such as glycerol, glucose, saccharides, proteins, ammonia, *etc*. can be used as fuels for PFCs with TiO_2 photoanode. Previous work demonstrates that PFCs can convert biomass into electricity at ambient, even at low temperatures. However, TiO_2 and WO_3 with band

gaps of 3.2 eV and 2.8 eV, respectively, only use a small proportion of solar irradiation.[15] Thereby, semiconductors with wide range of light absorption are highly desired for PFCs.

Hematite (α-Fe_2O_3) with optical band gap of 2.1 eV is a promising material for photoanode, due to its abundance, ecofriendly, and photochemical stability in basic electrolytes. [16-18] But the short excited state lifetime (~1 ps) and the small hole diffusion length (~2 to 4 nm) of hematite significantly hinder its efficiency in charge separation and collection. To overcome these problems, various strategies have been adopted, including reducing the size,[19, 20] doping [21-23] and loading cocatalyst [24-26], *etc.* Among them, loading cocatalyst is the most efficient way for photoelectrochemical water splitting. However, α-Fe_2O_3 loaded with appropriate cocatalyst which can distinctly enhance biomass oxidation has not been reported.

In this work, we present a PFC with α-Fe_2O_3 based photoanodes that directly converts biomass-derived compounds, such as methanol, glucose, glycerol, cellubiose, and starch. To improve the PFC efficiency and the stability of α-Fe_2O_3 photoanode, transition metal hydroxides $Ni(OH)_2$, $Fe(OH)_3$, and $Cu(OH)_2$ are loaded as biomass oxidation cocatalysts. It was found that the cocatalysts can remarkably enhance the photoresponse and stability of bare α-Fe_2O_3 for biomass oxidation, which demonstrates an example for prompting PFC by loading efficient cocatalyst on photoanode.

Experimental

Materials

All chemicals were analytical grade and were used as purchased. Solutions were prepared using high purity water (Millipore Milli-Q purification system, resistivity > 18 MΩ·cm). The FTO (fluorine-doped tin oxide) conductive glass was purchased from Nippon Sheet Glass Company 5 (Japan) and was ultrasonic cleaned with acetone, ethanol and deionized water for 20 min each in sequence prior to use.

Preparation of the Fe_2O_3 films

α-Fe_2O_3 films were deposited on a F-doped SnO_2 (FTO) substrate electrode using a modified chemical bath deposition method reported in elsewhere.[27] Specifically, 0.2 mol L^{-1} $FeCl_3$ ($FeCl_3 \cdot 6H_2O$, ≥ 99%, Shanghai Chemical) aqueous solution (10 mL) and 0.3 mol L^{-1} NH_2CONH_2 (98%, Shanghai Chemical) were mixed in a 50 mL glass beaker and heated at 100 °C for 4 h. FTO was placed vertically in these beaker with the conducting edge facing the wall of the beaker. After the reaction, the film formed on FTO was thoroughly rinsed in high purity water and annealed at 500 °C for 3 h to get the desired phase. Finally, the prepared sample was further annealed at 750 °C for 10 min.

Fabrication of $M(OH)_x/Fe_2O_3$ photoanode

$Ni(OH)_2$ was deposited onto α-Fe_2O_3 by a successive ionic layer adsorption and reaction method. In a typical synthesis, α-Fe_2O_3 electrodes were dipped into 0.1 mol L^{-1} $Ni(NO_3)_2$ (≥ 98.5%, Shanghai Chemical) solution for 40 s, following by drying with compressed air. Then the electrodes were dipped into 1 mol L^{-1} KOH solution for another 40 s and dried with compressed air. And then the prepared sample was

washed with ethanol and dried at 60 °C for 1 h in air. $Fe(OH)_3/\alpha$-Fe_2O_3 and $Cu(OH)_2/\alpha$-Fe_2O_3 samples were prepared by the same method without further heat treatment

Characterizations of the electrodes

The prepared samples were characterized by X-ray powder diffraction (XRD) on a Rigaku D/Max-2500/PC powder diffractometer using Cu Kα radiation (operating voltage: 40 kV, operating current: 20 mA, scan rate: 5° min^{-1}). The UV−visible diffuse reflectance spectra were recorded on a UV-visible spectrophotometer (JASCO V-550) and calibrated by Kubelka-Munk method. The morphologies of the electrodes were examined by a Quanta 200 FEG scanning electron microscope (SEM) equipped with an energy dispersive spectrometer (accelerating voltage of 20 kV). Transmission electron microscopy (TEM) images were taken on a Tecnai G^2 Spirit (FEI company) using an accelerating voltage of 120 kV. High-resolution transmission electron microscopy (HRTEM) images were performed on Tecnai G^2 F30 S-Twin (FEI company) with an accelerating voltage of 300 kV. The liquid products quantitative analysis was carried out using HPLC (Agilent 1200) with refractive index (RI) detector and ultraviolet (UV, λ=210 nm) detector for arabinose, erythrose, glyceraldehyde, glycolaldehyde, glycollate and formate. Reactant and products were separated through an ion exclusion column (Alltech OA-1000) heated at 55 ^oC. The eluent was a solution of H_2SO_4 (0.005 mol L^{-1}). Products were identified by comparison with standard samples, which were obtained from Sigma–Aldrich.

Photoelectrochemical and electrochemical measurements

All the photoelectrochemical measurements were carried out in a three-electrode cell with a flat quartz window to facilitate illumination of the photoelectrode surface. The working electrode is α-Fe_2O_3, and $Hg/HgCl_2$ (saturated KCl) and Pt plate (2 cm × 4 cm) were used as reference and counter electrode, respectively. A Nafion membrane was used to prevent the crossover between the anode and cathode. The illumination source was a 300 W Xe arc lamp and the light intensity at the surface of the electrodes was 300 mW cm^{-2}. The electrochemical measurements were performed on a CHI 760D electrochemical workstation (CHI, Shanghai) at room temperature. The electrolyte is 1 mol L^{-1} KOH solution with/without 0.025 mol L^{-1} glucose (or methanol (10 vol%), glycerol (0.025 mol L^{-1}), cellubiose (0.025 mol L^{-1}) and starch (1 wt%).

A photofuel cell was constructed using a two-compartment quartz cell with a Nafion membrane as separator. α-Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes were used as anodes and a Pt wire was used as cathode, while a Nafion membrane was served as separator. The distance between the anode and cathode is 10 cm. The photovoltaic performance of the cells was measured with a Keithley 2400 source measure unit, irradiated by AM 1.5 (100 mW cm^{-2}).

The energy efficiency (PEC) and fill factor (FF) of such a device can be estimated through Equation (1) and Equation (2), respectively.

$$PEC = P_{Out}/P_{h\nu} \times 100\% \quad (1)$$

$$FF = (V_{max} \times I_{max})/(V_{oc} \times I_{sc}) \quad (2)$$

P_{Out} and $P_{h\nu}$ represent the output electrical power and the imput photochemical

enengy, respectively; I_{sc} is the measured short circuit current; and V_{oc} is the specified open current voltage. I_{max} and V_{max} correspond to the current and voltage at the maximum power point.

Results and Discussions

The XRD patterns of Fe_2O_3 film was shown in Fig. 1a. The sample shows diffraction peaks with 2θ at 35.6°and 64.0°, which correspond to the indices of (110), and (300) planes (PDF NO. 840306) of hematite phase. The UV-Visible spectrum (Fig. 1b) gives the bands of α-Fe_2O_3 at 410 nm and around 600 nm, which corresponds to the direct transition of O^{2-} 2p → Fe^{3+} 3d and the transition of the spin-forbidden Fe^{3+} 3d → 3d, respectively. [28] Fig. 1c and 1d illustrate the top-view analysis and morphology of α-Fe_2O_3 film characterized by SEM. The film is consisted of α-Fe_2O_3 nanorod particles, with the diameter of 20-30 nm and the length of 60-100 nm. The cross analysis of the film shows a thickness of about 300 nm with a uniform and continuous morphology.

Fig. 2 shows the dark and photocurrent densities of an α-Fe_2O_3 photoanode under light illumination. The dark response is negligible up to 0.6 V *vs.* SCE for both 1 mol L^{-1} KOH electrolyte and 1 mol L^{-1} KOH containing glucose. Above 0.6 V *vs.* SCE, an anodic current formed due to water or glucose oxidation. It suggested that α-Fe_2O_3 photoanode has equal oxidation ability for glucose and water under these conditions. Whereas in KOH, the onset potential is shifted to -0.43 V *vs.* SCE and the photocurrent is due to water oxidation. When glucose is added to aqueous KOH electrolyte, the photo-onset potential shifts to a lower value of -0.6 V *vs.* SCE and the

photocurrent density obviously increases. These results indicate that glucose is preferentially oxidized over water, and more photo-generated electrons and holes are used for O_2 reduction and glucose oxidation. The results also suggest that α-Fe_2O_3 photoanode can efficiently oxidize biomass in photochemical process, which means that the sunlight utilization for PFC could be extended to the visible light region.

To improve the activity and stability of α-Fe_2O_3 photoanode, transition metal hydroxides $Ni(OH)_2$, $Fe(OH)_3$, and $Cu(OH)_2$ were deposited on α-Fe_2O_3 film by a successive ionic layer adsorption and reaction method. The SEM elemental mapping was carried out to detect the state of the $Ni(OH)_2$ modifying on Fe_2O_3. As shown in Fig. S1, Ni is distributed randomly on Fe_2O_3 and the mass ratio of Ni to Fe is ca. 0.24 detected by EDS. To further study the microstructure of Ni(OH)2/α-Fe2O3, HRTEM analyses were conducted. HRTEM images of a representative $Ni(OH)_2$/α-Fe_2O_3 photoanode are shown in Fig. 3a and 3b. The HRTEM image confirms the Ni species present on the edge of α-Fe_2O_3. The magnified HRTEM image in Fig. 3b exhibits fringes with lattice spacing of *ca.* 0.37 nm, which corresponds to the (012) plane of α-Fe_2O_3. While the lattice fringes with a spacing of 0.218 and 0.154 nm is consistent with the (103) and (300) planes of the $3Ni(OH)_2 \bullet 2H_2O$ (α-$Ni(OH)_2$). The EDS linescan performed across the α-Fe_2O_3 nanorod shows the distinct spatial profiles of Fe and Ni, confirming Ni species being deposited on Fe_2O_3 surface (Fig. S2, ESI†). These results reveal that α-$Ni(OH)_2$ nanoparticles adhere to the surface of α-Fe_2O_3, forming heterojunctions between the two components.

Fig. 4a shows the linear sweep voltammetric (LSV) curves of α-Fe_2O_3 and

$Ni(OH)_2/\alpha$-Fe_2O_3 photoanode in 1 mol L^{-1} KOH electrolyte with glucose under chopped light illumination. After loading $Ni(OH)_2$ cocatalyst, the photocurrent for glucose oxidation on the α-Fe_2O_3 electrode is significantly increased. The $Ni(OH)_2$ film itself shows no photoresponse for glucose oxidation (Fig. S3, ESI†). While in 1 mol L^{-1} KOH electrolyte without glucose, the photocurrent curve of $Ni(OH)_2/\alpha$-Fe_2O_3 electrode is different from that of α-Fe_2O_3 electrode (Fig. S4, ESI†), which may be due to the oxidation of $Ni(OH)_2$ in the photoelectrochemical process. This suggests that $Ni(OH)_2$ acts as the active site for glucose oxidation.

The enhanced photocurrent observed for the $Ni(OH)_2/\alpha$-Fe_2O_3 sample in Fig. 4b is possibly due to that $Ni(OH)_2$ traps photo-generated holes, forming Ni^{3+} species, then efficiently catalyze the glucose oxidation, and simultaneously Ni^{3+} is reduced back to Ni^{2+}.[29-33] Namely, the cocatalyst $Ni(OH)_2$ can enhance the photo-generated charges separation of α-Fe_2O_3 and provide the catalytic sites for glucose oxidation.

Some metal components (Fe and Cu) with the multiple oxidation states were demonstrated to be excellent electrocatalysts for glucose sensor. [34, 35] So, $Fe(OH)_3$ and $Cu(OH)_2$ were also used to modify α-Fe_2O_3 photoanode for glucose oxidation. As shown in Figs. 4c and 4d, The major phenomenological observation is that modification of α-Fe_2O_3 with $Fe(OH)_3$ and $Cu(OH)_2$ can also obviously enhance the photocurrent density. The above results show that $Fe(OH)_3$ and $Cu(OH)_2$ play the similar role to that of $Ni(OH)_2$. Namely, they can also improve the hole transfer from Fe_2O_3 to glucose and provide active sits for glucose oxidation.

The α-Fe_2O_3 based photoanodes with cocatalysts, like $Ni(OH)_2$ reported in this

study can photooxidize various biomass molecules in combination with an O_2-reducing cathode. Fig. 5 shows the photocurrent density generated by methanol, glycerol, cellubiose and starch. The photocurrent densities of $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes can reach 4.7 and 5.0 mA cm^{-2} at 0 V *vs.* SCE fueled by methanol and glycerol, respectively. Whereas cellubiose and starch with relative big molecular weight, give photocurrent densities respectively about 4.0 and 3.5 mA cm^{-2} at 0 V *vs.* SCE with $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes. It was noted that photocurrent density produced by starch was lower than that obtained from cellubiose solution under similar conditions, probably because of the slow hydrolysis of starch. The results illustrate that all these biomass materials are promising fuels for PFC and loading cocatalyst is an efficient method for improving the PFC performance.

A simple PFC consisting of α-Fe_2O_3 (or $Ni(OH)_2/\alpha$-Fe_2O_3) photoanode and Pt plate cathode in glucose solution were fabricated. Fig. 6 illustrates the voltage-current density and power density-current density curves, and Table 1 displays the photovoltaic parameters for both PFCs. The device with α-Fe_2O_3 photoanode yields an open-circuit voltage (V_{oc}) of 0.38 V, a short-circuit current (I_{sc}) of 1.17 mA cm^{-2}, and a fill factor (FF) of 0.18, resulting in an overall power conversion efficiency (PCE) of 0.082%. In contrast, the photovoltaic device containing the $Ni(OH)_2/\alpha$-Fe_2O_3 photoanode presents V_{oc} of 0.43 V, the I_{sc} of 1.97 mA cm^{-2} and FF of 0.21. As a result, the $Ni(OH)_2/\alpha$-Fe_2O_3 based cell reaches a conversion efficiency of 0.18%, outperforming that of the Fe_2O_3-based device. As shown in power density-current density curves, the device with α-Fe_2O_3 photoanode yields a

maximum power density of 0.082 mA cm^{-2}, while the maximum power density (P_{max}) of $Ni(OH)_2/\alpha$-Fe_2O_3 is 0.18 mA cm^{-2}. These results reveal that with α-Fe_2O_3 based PFC, biomass can be directly converted into electrical energy and the efficiency of can be enhanced by modifying α-Fe_2O_3 photoanode with cocatalysts.

The stability of α-Fe_2O_3 based photoanodes and the intermediates in the reaction were tested in 1 mol L^{-1} KOH electrolyte containing glucose. Fig. 7 illustrates the amperometric I–t curves of α-Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 electrodes under continuous illumination at 0 V *vs.* SCE. The photocurrent of bare α-Fe_2O_3 after 4 h illumination is 86% of the initial value, but in the case of $Ni(OH)_2/\alpha$-Fe_2O_3 electrode the photocurrent can be maintained by 96%. Additionally, XPS results (Fig. S5, ESI†) shows that the valence state of Ni are maintained during the reaction in 4 h, suggesting that Ni^{2+} in $Ni(OH)_2/\alpha$-Fe_2O_3 electrode can be rapidly regenerated. It suggests that cocatalyst is rather effective for improving the steady state of α-Fe_2O_3 photoanode.

The products from glucose with $Ni(OH)_2/\alpha$-Fe_2O_3 photoanode was detected by HPLC after 6 h reaction. These intermediates are identified as arabinose, erythrose, glyceraldehyde, glycolaldehyde, glycollate and formate, and the carbon balance is ca. 71% (Table S1 and Fig. S6, ESI†). The carbon loss of 29% is mainly attributed to the release of CO_2, which are produced from the cleavage of C-C bond of glucose and its derivatives. The result implies that α-Fe_2O_3 based photoanodes can efficiently cleave the C-C bond in biomass and oxidize it to CO_2 in the PFC at low temperature.

Conclusion

The PFCs consisting of α-Fe_2O_3 based photoanode and an O_2-reducing cathode was fabricated. The sunlight utilization of PFCs was extended to the visible light region. We firstly reported that transition metal hydroxides (Ni, Fe and Cu) are excellent cocatalysts of α-Fe_2O_3 for biomass oxidation. The performance of PFC can be obviously enhanced by loading these cocatalysts on α-Fe_2O_3 photoanode. Moreover, compared with traditional DAFCs, PFCs can efficiently break the C-C bond of the biomass, and thus PFCs can be directly powered with natural biomass molecules.

Acknowledgments

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Figures and Tables

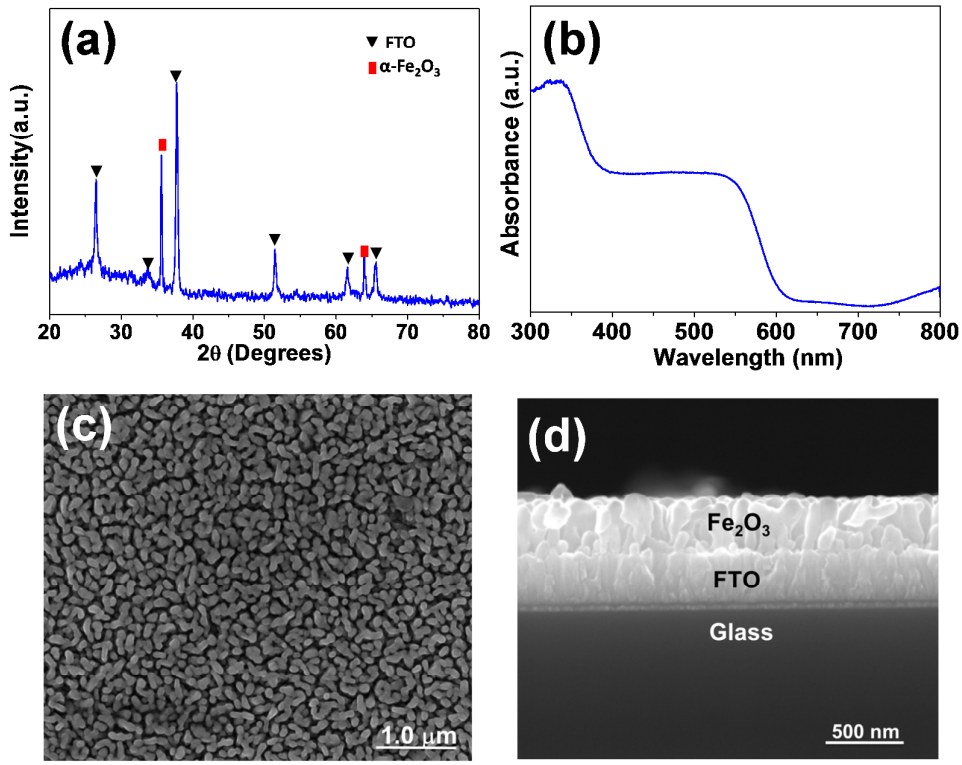

Fig. 1 (a) XRD pattern and (b) UV-vis absorption spectra of of Fe_2O_3 film. (c) Top view and (d) side view SEM images of Fe_2O_3 film.

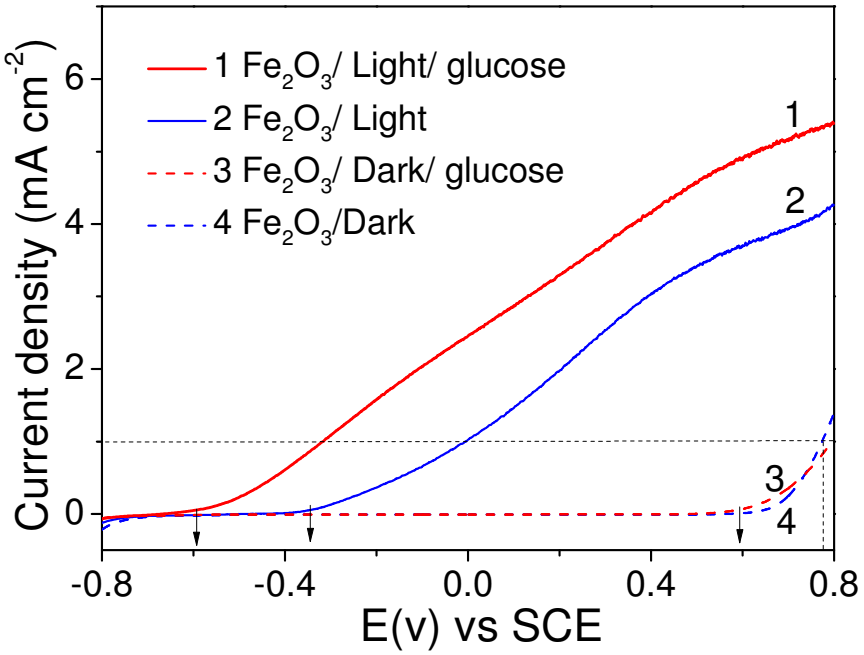

Fig. 2 Dark and photocurrent densities for α-Fe_2O_3 photoanode in 1 mol L^{-1} KOH electrolyte and in 1 mol L^{-1} KOH electrolyte containing 0.025 mol L^{-1} glucose; Light source: 300 W Xe lamp; Scanning rate: 20 mV s^{-1}.

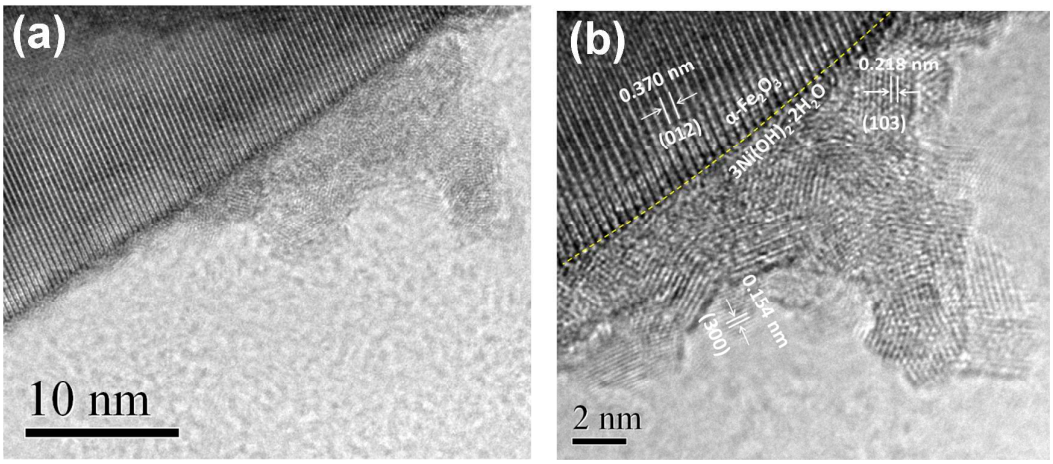

Fig. 3 (a) HRTEM image of $Ni(OH)_2/\alpha$-Fe_2O_3 and (b) the magnified HRTEM image of the selected frame from image (a).

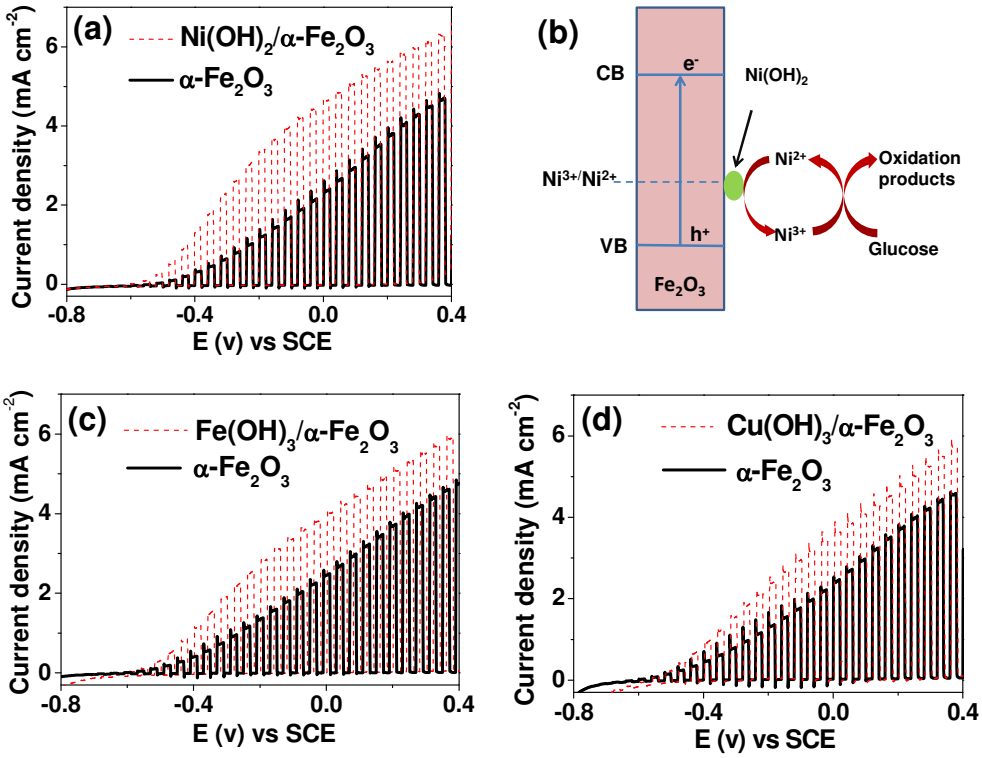

Fig. 4 (a) LSV curves of Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes under chopped light illumination; (b) The possible process of glucose oxidation over $Ni(OH)_2/Fe_2O_3$ photoanode; (c) and (d) LSV curves of α-Fe_2O_3 photoanodes modified with $Fe(OH)_3$ and $Cu(OH)_2$; Reaction condition: 1 mol L^{-1} KOH aqueous solution with 0.025 mol L^{-1} glucose; Light source: 300 W Xe lamp; Scanning rate: 20 mV s^{-1}.

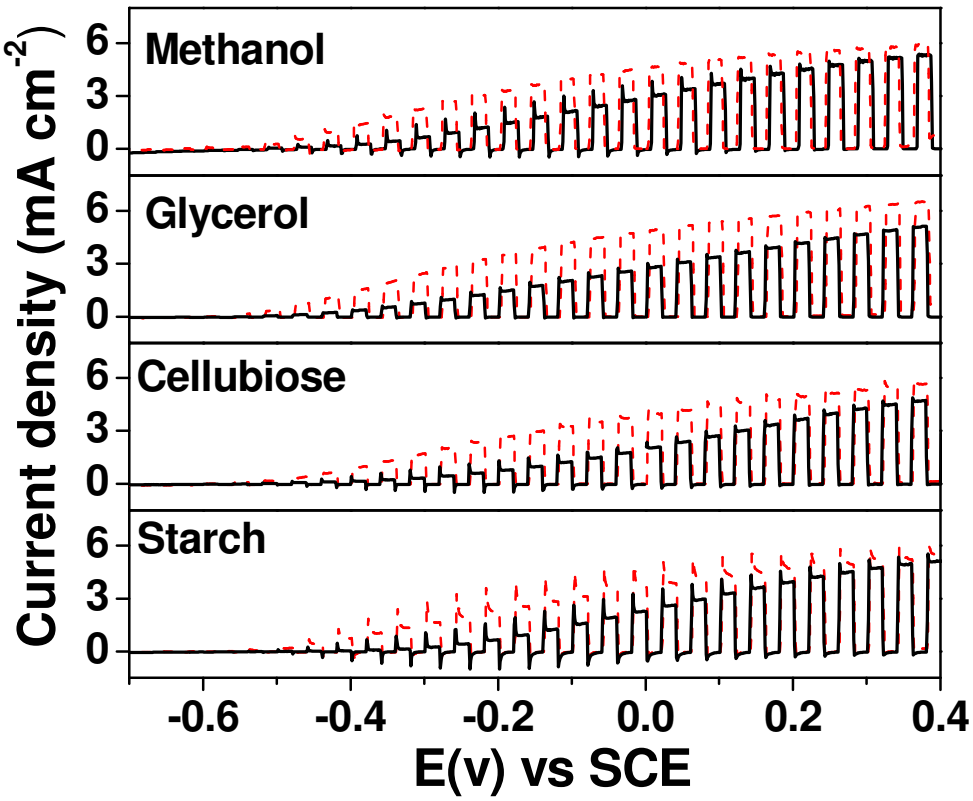

Fig. 5 LSV curves of α-Fe_2O_3 (solid) and $Ni(OH)_2$/α-Fe_2O_3 (dashed) photoanodes in 1 mol L^{-1} KOH electrolyte with different biomass-derived compounds under chopped light illumination; methanol (10 vol%); glycerol (0.025 mol L^{-1}); cellubiose (0.025 mol L^{-1}) and starch (1 wt%). Light source: 300 W Xe lamp; Scanning rate: 20 mV s^{-1}.

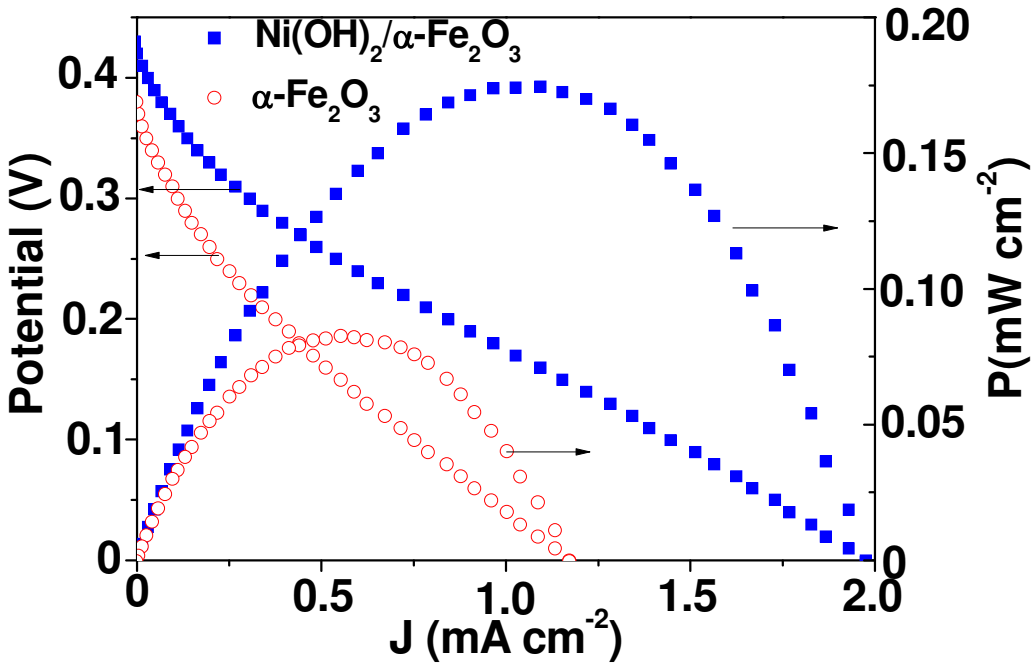

Fig.6 I-V and I-P curves of PFCs for α-Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes operating with light irradiation of in 1 mol L^{-1} KOH electrolyte containing 0.025 mol L^{-1} glucose, under irradiation with the AM 1.5G simulated solar light (100 mW·cm^{-2}).

Table 1 Photovoltaic parameters of the light-assisted glucose fuel cells consisting of α-Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes.

Anode	Intensity (mW cm^{-2})	V_{OC} (V)	I_{SC} (mA cm^{-2})	FF	P (mW cm^{-2})	PEC (%)

α-Fe_2O_3	100	0.38	1.17	0.18	0.082	0.08
$Ni(OH)_2/\alpha$-Fe_2O_3	100	0.43	1.98	0.21	0.18	0.18

See in Fig.6 for the test condition.

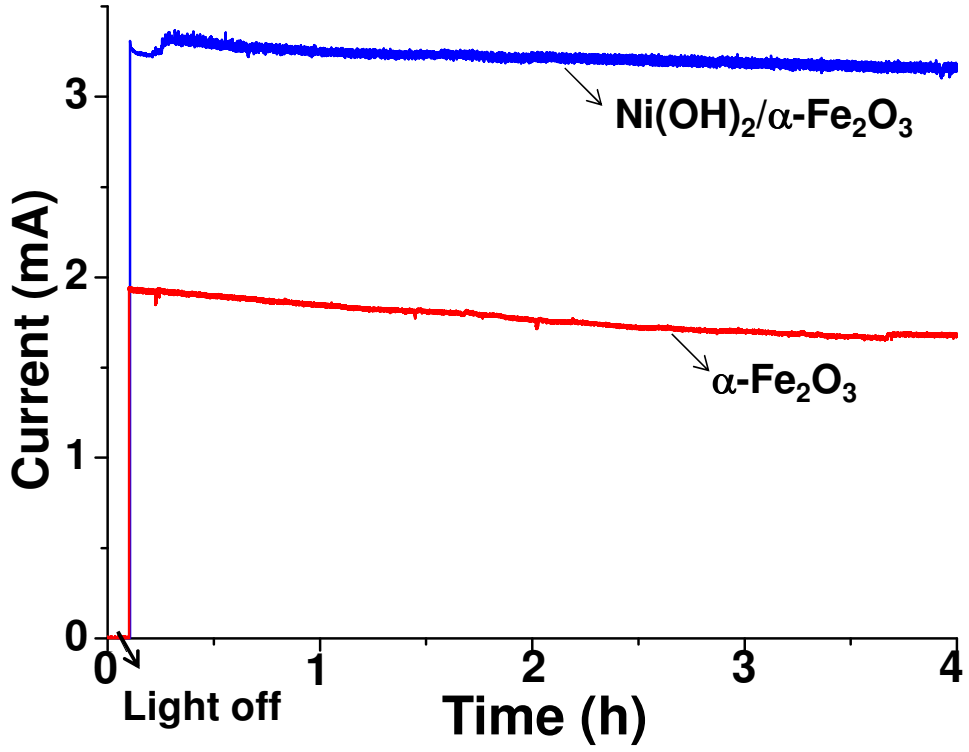

Fig. 7 Amperometric I–t curves of α-Fe_2O_3 and $Ni(OH)_2/\alpha$-Fe_2O_3 photoanodes at 0 V *vs.* SCE in a 1 mol L^{-1} KOH electrolyte containing 0.025 mol L^{-1} glucose. Light source: 300 W Xe lamp. The area of α-Fe_2O_3 film is 1.58 cm^2.