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Conjugated Microporous Polymers with Excellent Electrochemical Performance for Lithium and Sodium Storage

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Conjugated microporous polymers are explored as high-performance electrode materials for Lithium and Sodium storage, which exhibit high specific capacity, superior cycle stability and remarkable rate capability. The excellent electrochemical performance can be attributed to their conductive frameworks, plentiful redox-active units, high specific surface area and homogenous microporous structure.

In efforts to address the energy crisis and environmental issues, clean and sustainable energy systems have been investigated, such as solar cells, fuel cells and rechargeable batteries. Currently, lithium ion batteries (LIBs) dominate the portable consumer electronic market due to their high energy density.[1,2] But because of the high cost and limited resources of lithium (Li), it is urgent to find a substitute for Li to meet the demands for rechargeable batteries. Recently, sodium ion batteries (SIBs) have been recognized as a low cost alternative to LIBs for next generation battery systems and large scale energy storage devices due to the natural abundance of sodium (Na) and the similar chemical property of Na and Li.[3-5] However, the commonly used electrode materials in state-of-the-art LIBs and SIBs contain transitional metal,[6-8] which is resource-limited and non-environmental friendly. As an alternative, organic electrode materials have attracted more-and-more interest in recent years due to their low cost, high electrochemical performance, resource sustainability, environmental friendliness, structure diversity and flexibility.[9-11] During the past decades, a large variety of organic electrode materials have been reported for capacitors[11,12] and rechargeable batteries, such as conjugated carboxylates,[13-15] conjugated polymers[16,17] and polynitroxide radicals,[18] demonstrating the potential of organic electrodes for the next generation of green, sustainable and versatile energy storage devices.

In recent years, conjugated microporous polymers (CMPs), combining π-conjugated skeletons with permanent nanopores, have received increasing interest as an important branch of organic porous polymers.[19,20] Compared with other inorganic or inorganic-organic hybrid porous materials, CMPs have intrinsic advantages such as high degree of π-conjunction, homogenous microporous structure, ultrahigh specific surface area, diversity and flexibility in the molecular design. Because of its unique features, CMPs have shown great potential in gas absorption,[21,22] gas separation,[23,24] heterogeneous catalysis[25,26] and so on. However, there have been only a few reports of CMPs applied for Li or Na storage. Recently, Jiang and co-workers have applied CMPs for supercapacitors[27] and cathodes of LIBs.[28] Sakaushi and co-workers applied bipolar porous polymeric frameworks for all-organic[29] and Na-organic energy storage devices,[30] exhibiting excellent electrochemical performance. These results indicate that CMPs have enormous potential as green sustainable and flexible electrode materials for next generation energy storage devices.

In this work, we synthesize the polymer 4,7-dicarbazyl-[2,1,3]-benzothiadiazole (PDCzBT) by $FeCl_3$ oxidation coupling polymerization and investigate its electrochemical performance for Li and Na storage. The prepared PDCzBT shows ultrahigh specific surface area that provides abundant active sites for storage reaction, and uniform microporous structure that is beneficial for the rapid transport of electrons and ions. The skeleton of PDCzBT is built with plentiful redox-active units, which provide abundant energy storing modules. PDCzBT delivers a n-type reversible redox behavior with the reaction of Li^+/Na^+ doping/dedoping into/from the polymer chains. The assembled batteries based on PDCzBT electrodes exhibit excellent electrochemical performance for Li and Na storage, including high specific capacity, outstanding cycle stability and superior rate performance. These results indicate that CMPs is a promising electrode material with high capacity, outstanding rate capability and long cycle life for next generation energy storage devices.

Experimental

Samples preparation

Tetrahydrofuran (THF) and chloroform were purified by distillation prior to use. 2,1,3-benzothiadiazole (99%), carbazole (96%), cuprous oxide (96%), ferric chloride (97%), dimethylacetamide (DMAc, 99.8%), and other solvents were purchased from J&K or Aldrich and used as received.

Synthesis of 4,7-dicarbazyl-[2,1,3]-benzothiadiazole (DCzBT). To a mixture of 2,1,3-benzothiadiazole (1.470 g, 5.0 mmol) and carbazole (2.508 g, 15.0 mmol), cuprous oxide (2.862 g, 20.0 mmol), 10 mL DMAc was added. The suspension was heated to 160 ℃ and refluxed for 24 h under nitrogen. After cooling to room temperature, the mixture was filtered to remove the cuprous oxide and then the

filtrate was poured in water. The suspension was filtered and washed with water and ethanol to get the red brown solid. Final product DCzBT was purified with silica gel column chromatography (510 mg, 21 %). ^{1}H NMR (600 MHz, $CDCl_3$): δ (ppm) = 8.22 (d, 4H), 8.07 (s, 2H), 7.45 (m, 4H), 7.37 (m, 4H), 7.33 (d, 4H). ^{13}C NMR (150 MHz, $CDCl_3$) : δ (ppm) = 152.64, 141.02, 129.72, 127.75, 126.09, 124.09, 120.81, 120.58, 110.43.

Synthesis of polymer 4,7-dicarbazyl-[2,1,3]-benzothiadiazole (PDCzBT). The solution of monomer DCzBT (200 mg, 0.43 mmol) dissolved in 30 mL of anhydrous chloroform was dropwise transferred to a suspension of ferric chloride (920 mg, 3.44 mmol) in 20 mL of anhydrous chloroform. The solution mixture was stirred for 24 h at room temperature under nitrogen protection, and then 100 mL of methanol was added to the above reaction mixture. The resultant precipitate was collected by filtration and washed with methanol and concentrated hydrochloric acid solution. After extracted in a Soxhlet extractor with methanol for 24 h, and then with THF for another 24 h extraction, the desired polymer was collected and dried in vacuum oven at 80 ℃ overnight. Brown powder, yield in 96%. Anal. Calcd. for $C_{30}H_{18}N_4S$: C, 77.23; N, 12.01; H, 3.89; S, 6.87. Found: C, 77.20; N, 12.12; H, 3.94; S, 6.62.

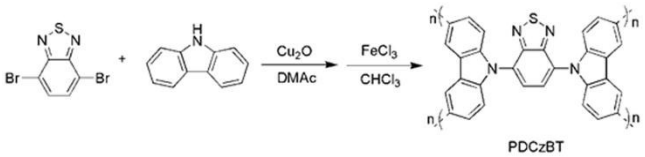
Characterization

Morphological information was obtained using field emission scanning electron microscopy (FESEM, HITACHI S-4800) and transmission electron microscopy (TEM, HITACHI H-7650). ^{1}H and ^{13}C NMR spectra were measured by the Bruker Avance III model 600 MHz using $CDCl_3$ as solvent. ^{13}C CP/MAS solid-state NMR measurement was carried out on a Bruker Avance III model 400 MHz NMR spectrometer at a MAS rate of 5 kHz. FT-IR spectrum was collected in attenuated total reflection (ATR) mode on a Thermo Nicolet 6700 FT-IR Spectrometer. Thermo gravimetric analysis (TGA) was carried out using a SDT Q600 (V20.9 Build 20) with a temperature ramp of 10 ℃ min^{-1} from 20 ℃ to 800 ℃ under N_2 atmosphere. Nitrogen adsorption/desorption measurements were performed at 77 K using a Quantachrome Autosorb gas-sorption system.

Electrochemical measurements

The electrochemical experiments were performed in 2032 coin-type cells. The working electrodes were prepared by mixing 60 wt % PDCzBT with 30 wt % Super P and 10 wt % polytetrafluoroethlyene (PTFE) binders. The obtained electrode samples were rolled into slices and cut into square pieces of 1 cm × 1 cm (the areal loading of the total materials is about 3 mg/cm^2 and the areal loading of the active materials is around 1.8 mg/cm^2), then pasted on a stainless steel current-collector under a pressure of 15 MPa, followed with dried in vacuum oven at 120 ℃ for 4 h, used as the working electrode. Pure Li (or Na) foil was used as the counter electrode, which was separated from the working electrode by a Celgard 2500 polymeric separator. The electrolyte was 1 M $LiPF_6$ in ethylene carbonate (EC)/dimethyl carbonate (DMC)/diethyl carbonate (DEC) (1:1:1, v/v/v) containing 5% (by volume) vinylene carbonate (VC) or 1M $NaClO_4$ in EC/DMC (1:1, v/v) containing 2% (by volume) fluoroethylene carbonate (FEC). The cells were assembled in an argon-filled glovebox with the concentrations of moisture and oxygen at less than 1ppm. The galvanostatic charge/discharge cycling performance was measured using a LAND battery testing system. Cyclic voltammetry (CV) was performed using an IM6 electrochemical workstation at a scan rate of 0.2 mV/s. The electronic conductivity of PDCzBT was carried out by linear voltage scanning method (The detailed measuring method is shown in supplementary information). The capacity was calculated based on the mass of PDCzBT.

Results and discussion



Scheme 1. Synthetic route of conjugated microporous polymer PDCzBT.

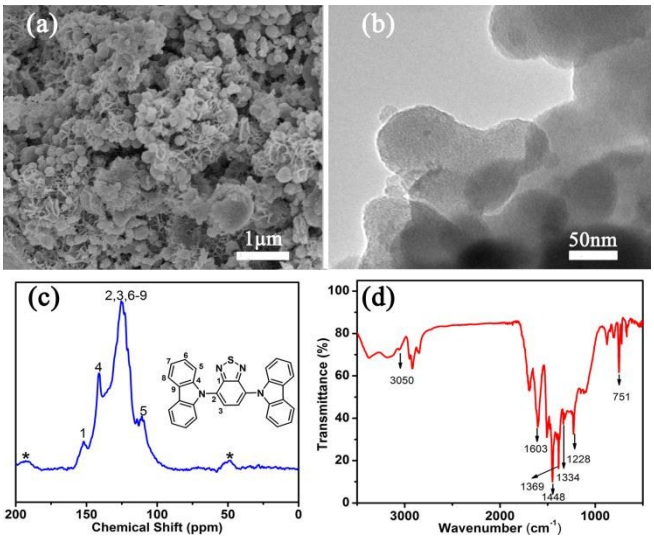


Fig. 1 (a) The SEM image, (b) the TEM image, (c) ^{13}C CP/MAS solid NMR, and (d) FT-IR spectrum of PDCzBT.

The polymer 4,7-dicarbazyl-[2,1,3]-benzothiadiazole (PDCzBT) was synthesized by $FeCl_3$ oxidation coupling polymerization[23, 31] as shown in Scheme 1. The SEM image (Fig. 1a) shows that the PDCzBT consists of relatively uniform sub-micron spheres. As can be seen in Fig. 1b, the TEM images indicate the porous structure of PDCzBT. The molecular structure of PDCzBT was assessed with ^{13}C CP/MAS solid nuclear magnetic resonance (NMR), as shown in Fig. 1c. The characteristic peak at 152 ppm corresponds to the carbon in the –C=N groups. The peak at 141 ppm is ascribed to carbons connecting to the N atom in the carbazole groups. The signal at 110 ppm and the broad peaks from 120 to 129 ppm are assigned to the other carbon atoms of the aromatic rings in the polymers. A more detailed analysis of the structure (monomer and polymer) was confirmed by the FT-IR spectra (Fig. S1 and Fig. 1d). The spectra reveal the following absorption peaks: the bands around 751 cm^{-1} are due to ring deformations of aromatic structure; the peaks at 1228, 1334, 1369 cm^{-1} may be the stretching frequencies of the structure such as N-S-N, C-N and N-S moiety; the bands around 1448~1603 cm^{-1} should be the double bonds (C=C and C=N) stretching vibration in the structure; the peak around 3049 cm^{-1} is the stretching frequency of the C-H bonds in the aromatic structure. The thermostability of the PDCzBT was also investigated by thermo gravimetric analysis (TGA) as it is crucial to the safety of rechargeable batteries. The resultant PDCzBT exhibited a high thermal stability without decomposing up to 350 ^{o}C (Fig. S2).

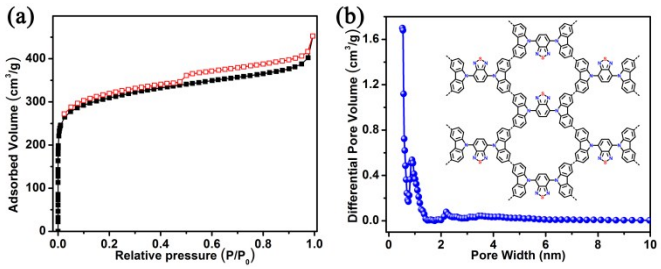


Fig. 2 (a) N_2 adsorption/desorption isotherm and (b) pore-size distribution of PDCzBT (the inset is its structural representation).

The pore structure of the polymer was evaluated by nitrogen adsorption-desorption isotherm at 77 K. The curve depicted in Fig. 2a displays a type Ⅰ isotherm with a rapid uptake at low relative pressure according to IUPAC classifications,[32] which indicating the microporous nature of PDCzBT. There is gradual rise and small hysteresis at relatively high pressure, consistent with the presence of interparticular void, which could be ascribed to porosities existing between the highly aggregated nanoparticles.[33] The calculated Brunauer-Emmett-Teller (BET) specific surface area and the total pore volume (V_{tol}) of PDCzBT is 1166 m^2 /g and 0.7 cm^3 /g, respectively. The surface and pore volume are comparable or higher than previously reported CMPs.[27-30] The micropore volume $V_{0.1}$ is 0.45 cm^3/g, and simultaneously the $V_{0.1}/V_{tol}$ value is 0.64, which means that micropores are dominant in PDCzBT. This is also certified by the pore size distribution (Fig. 2b). It reveals that PDCzBT has a dominant ultramicropore width at 0.52 and 0.86 nm, indicating the relatively uniform micropores in the polymer. The inherent ultramicropore and large pore volume are accessible to Li and Na ions transport and its high surface area endows PDCzBT abundant active sites for Li and Na storage.

Electrochemical performance of PDCzBT for Li storage

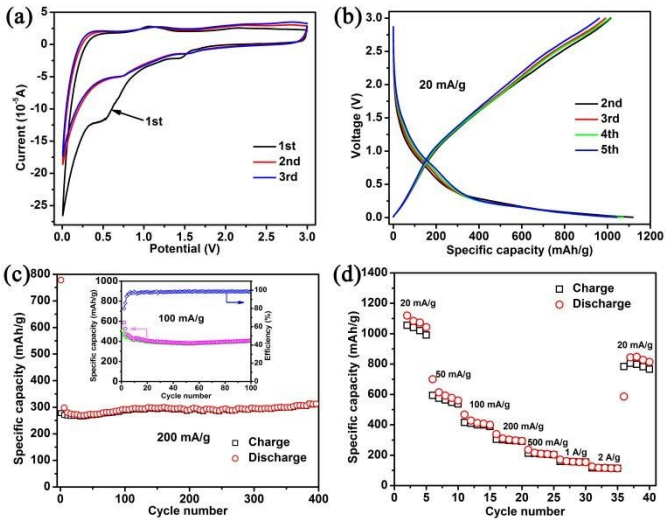


Fig. 3 Electrochemical performance of PDCzBT electrode for LIBs. (a) Cyclic voltammogram (CV) profiles. (b) Galvanostatic charge/discharge profiles at a current density of 20 mA/g. (c) Cycle performance at a current density of 200 mA/g, and the inset is cycle performance and coulombic efficiency under 100 mA/g. (d) Rate performance at varied current density ranging from 20 to 2000 mA/g.

The electrochemical performance in terms of lithium storage was evaluated using 2032 coin-type half cells. As shown in the cyclic voltammogram (CV) profiles (Fig. 3a), during the first cathodic scan, an irreversible band appeared at 1.0-0.5 V, corresponding to the electrochemical decomposition of electrolyte for the formation of solid electrolyte interface (SEI) film on the electrode surface. The CV bands at low potential region of 0-0.5 V, emerged as a pair of asymmetric redox peaks with a larger cathodic branch than anodic branch, reflecting the doping-dedoping reaction of Li^+ into/from the polymer chains.[16,30] The reaction mechanism can be based on the equation (1). However, the initial coulombic efficiency is as low as 49.1 % (Fig. S3), which is ubiquitous in carbon-based electrodes especially in polymer electrodes.

$$PDCzBT + xLi^+ + xe^- \leftrightarrow Li^+_xPDCzBT^{x-} \quad (1)$$

Fig. 3b displays the typical charge-discharge curves of the coin cells cycled in the n-dopable potential region, which is consistent with the CV curves. The large part of specific capacity (> 70%) in the region below 0.5 V (Fig. S4), corresponds to Li^+ doping into the polymer chains.[16,30] The specific capacity above 0.5 V might be attributed to the Li^+ absorbing on the surfaces/interfaces of PDCzBT.[34] The detailed Li storage mechanism can be further studied by the CVs (Fig. S5) at different scan rates,[35] which demonstrate that it is a mixed process involving both Li^+ doping into the polymer chains and absorbing on the surfaces/interfaces. It is worth noting that at a current density of 20 mA/g the reversible capacity could be up to 1042 mAh/g after 5 cycles as shown in Fig. 4d, which is higher than other organic materials reported previously.[13,16,29] The high specific capacity may be ascribed to the high specific surface area and inherent homogeneous microporous structure of PDCzBT, which endow PDCzBT abundant active sites for Li storage. It can be seen from Fig. 3c that the batteries with PDCzBT based electrode also exhibit outstanding cycle stability, with a reversible capacity of 404 mAh/g achieved after 100 cycles at a current density of 100 mA/g and a high coulombic efficiency of > 98% from 10^{th}. The PDCzBT electrodes exhibit a specific energy of 224.4 Wh/kg based on the mass of the polymer. Even at a high current density of 200 mA/g, the as-prepared PDCzBT electrodes deliver a moderate specific capacity of 312 mAh/g after 400 cycles.

For the LIBs application, one of challenging problems is the limited rate performance at high charge/discharge rates. The porous structures have been demonstrated to be beneficial for improving the rate performance of LIBs. The unique structure of PDCzBT with numerous homogeneous micropore is helpful for Li ions diffusion in the polymer chains, thus making the PDCzBT suitable as Li storage materials even at high charge/discharge rates. As shown in Fig. S8, the I–V curve of the PDCzBT is linear, which exhibits Ohmic behavior. The conductivity is calculated as 6.83×10^{-4} S/cm, which is comparable with silicon, and demonstrates that the PDCzBT exhibits semiconducting properties and the framework of PDCzBT is helpful for the faster movement of electrons. It can be seen from Fig. 4d, PDCzBT demonstrates a superior rate capability. When the current density increased to 500, 1000, 2000 mA/g, the reversible capacity can still retain 215, 161 and 117 mAh/g, respectively. All these results indicate that PDCzBT is a promising electrode

material with high capacity, outstanding rate capability and long cycle life for Li storage devices.

Electrochemical performance of PDCzBT for Na storage

Recently, SIBs have been recognized as a promising alternative to current LIBs for next generation battery systems and large scale energy storage devices due to the availability of sodium (Na) and the similar chemical property of Na and Li.[3-5] However, many traditional inorganic intercalation materials used in LIBs are usually composed of a 3D rigid network and relatively small tunnel size, which is suitable for diffusion of Li^+ but not for Na^+.[15,36] The larger ionic radius of Na^+ than Li^+ results in a larger strain during insertion/extraction process, which leads to the disintegration of 3D rigid network. In contrast, the ion radius has very little effect on the electrochemical performance of n-type organics, mainly due to their flexible framework. Organic materials can accommodate large Na ions reversibly without much spatial hindrance, thus facilitating to achieve a fast kinetics for Na ions insertion and extraction reactions.[14,15,37] In addition, organic materials are environmentally friendly, resource sustainability and structure diversity. Consequently, organic materials might be a good candidate as electrode materials for Na-based energy storage devices.

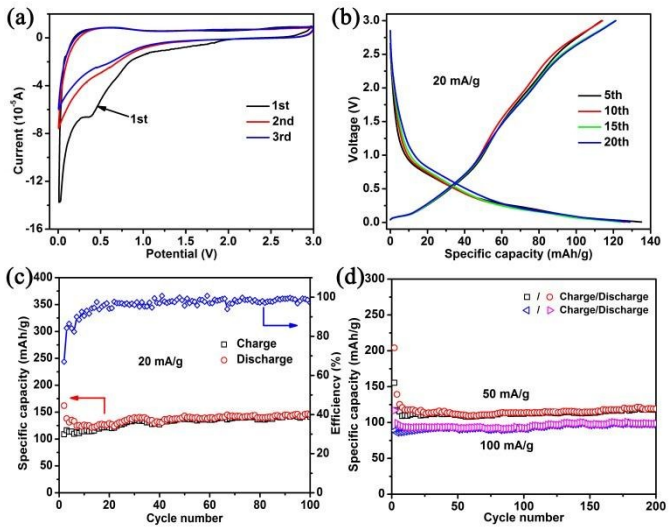


Fig. 4 Electrochemical performance of PDCzBT electrode for SIBs. (a) Cyclic voltammogram (CV) profiles. (b) Galvanostatic charge/discharge profiles at a current density of 20 mA/g. (c, d) Cycle performance at a current density of 20, 50 and 100 mA/g.

In this work, we apply PDCzBT as electrode materials for SIBs and investigate its Na storage behavior. It can be seen from Fig. 4a and b, the CV curves and charge/discharge profiles are similar to the curves for lithium storage, only with lower voltage and capacity, which is due to the difference in the thermodynamics and kinetics for the insertion of Li and Na ions. Because of the larger ionic radius of Na^+ (102 pm) than Li^+ (76 pm), the kinetics of Na insertion and extraction are more sluggish compared to Li.[4,38] Fig. 4b displays a typical n-doping charge/discharge profiles, and the Na storage mechanism is also an absorption and insertion process. The insertion process can be based on the equation (2), corresponding to the Na^+ doping/dedoping into/from the polymers chains. Fig. 4c and d show that the assembled SIBs exhibited a moderate specific capacity and remarkable cycle stability, with a reversible capacity of 145 mAh/g achieved after 100 cycles at a current density of 20 mA/g and a high coulombic efficiency of > 96% from 15th. The difference between Na storage and Li storage might be attributed to the larger ionic radius and more sluggish kinetics of Na^+ than Li^+, which results in a lower doping and absorbing amount of Na^+. The specific explanation is further being studied. At a high current density of 50 and 100 mAh/g, the reversible capacity up to 119 and 99 mAh/g (Fig. 4d) could be obtained after 200 cycles respectively without any capacity reduction from 15th cycle, demonstrating the outstanding electrochemical performance of PDCzBT for Na storage. However, the initial coulombic efficiency is very low at 29.8 % (Fig. S7), and it still needs numerous efforts to study deeply and improve its performance.

$$PDCzBT + xNa^+ + xe^- \leftrightarrow Na^+_xPDCzBT^{x-} \quad (2)$$

The excellent electrochemical performance of PDCzBT for Li and Na storage can be ascribed to four aspects: (1) conductive CMPs frameworks that is conductive to electronic fast transport, (2) abundant redox-active units of its skeletons that provide sufficient energy storing modules, (3) high specific surface areas that endows it more active sites for storage reaction, (4) inherent well-developed microporous structure that allow for the fast transport of ions.

Conclusions

In summary, PDCzBT with large surface area and uniform microporous structure was facilely prepared by $FeCl_3$ oxidation coupling polymerization and applied as electrode materials for LIBs and SIBs. The high surface area and plentiful redox-active units provide abundant active sites and energy storing modules. The homogenous microporous structure is beneficial for the fast transport of electrons and ions. The Li/Na storage mechanism of PDCzBT is an absorption and insertion process. The assembled batteries based on PDCzBT electrodes exhibit excellent electrochemical performance for Li and Na storage, including high specific capacity, outstanding cycle stability and superior rate performance. These noticeable results demonstrate the enormous potential of CMPs as green, sustainable, flexible and high-performance electrode materials for next generation energy storage devices.

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Notes and references

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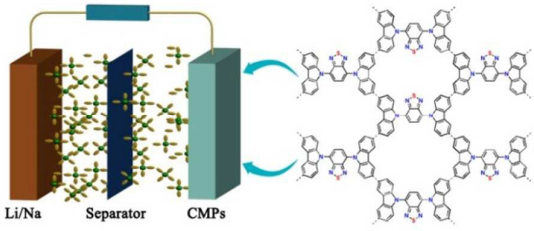
1. M. Armand and J. M. Tarascon, *Nature*, 2008, **451**, 652-657.
2. V. Etacheri, R. Marom, R. Elazari, G. Salitra and D. Aurbach, *Energy Environ. Sci.*, 2011, **4**, 3243-3262.
3. S.-W. Kim, D.-H. Seo, X. Ma, G. Ceder and K. Kang, *Adv. Energy Mater.*, 2012, **2**, 710-721.
4. V. Palomares, P. Serras, I. Villaluenga, K. B. Hueso, J. Carretero-Gonzalez and T. Rojo, *Energy Environ. Sci.*, 2012, **5**, 5884-5901.
5. M. D. Slater, D. Kim, E. Lee and C. S. Johnson, *Adv. Funct. Mater.*, 2013, **23**, 947-958.
6. X. X. Li, F. Y. Cheng, B. Guo and J. Chen, *J. Phys. Chem. B*, 2005, **109**, 14017-14024.
7. M. S. Whittingham, *Chem. Rev.*, 2004, **104**, 4271-4301.
8. J. Chen and F. Cheng, *Acc. Chem. Res.*, 2009, **42**, 713-723.
9. B. Dunn, H. Kamath and J.-M. Tarascon, *Science*, 2011, **334**, 928-935.
10. X.-P. Gao and H.-X. Yang, *Energy Environ. Sci.*, 2010, **3**, 174-189.
11. Z. Song and H. Zhou, *Energy Environ. Sci.*, 2013, **6**, 2280.
12. X. Feng, Y. Liang, L. Zhi, A. Thomas, D. Wu, I. Lieberwirth, U. Kolb and K. Muellen, *Adv. Funct. Mater.*, 2009, **19**, 2125-2129.
13. M. Armand, S. Grugeon, H. Vezin, S. Laruelle, P. Ribiere, P. Poizot and J. M. Tarascon, *Nat. Mater.*, 2009, **8**, 120-125.
14. L. Zhao, J. Zhao, Y.-S. Hu, H. Li, Z. Zhou, M. Armand and L. Chen, *Adv. Energy Mater.*, 2012, **2**, 962-965.
15. Y. Park, D. S. Shin, S. H. Woo, N. S. Choi, K. H. Shin, S. M. Oh, K. T. Lee and S. Y. Hong, *Adv. Mater.*, 2012, **24**, 3562-3567.
16. L. Zhu, Y. Niu, Y. Cao, A. Lei, X. Ai and H. Yang, *Electrochim. Acta*, 2012, **78**, 27-31.
17. L. M. Zhu, A. W. Lei, Y. L. Cao, X. P. Ai and H. X. Yang, *Chem. Commun.*, 2013, **49**, 567-569.
18. T. Suga, S. Sugita, H. Ohshiro, K. Oyaizu and H. Nishide, *Adv. Mater.*, 2011, **23**, 751-754.
19. J.-X. Jiang, F. Su, A. Trewin, C. D. Wood, N. L. Campbell, H. Niu, C. Dickinson, A. Y. Ganin, M. J. Rosseinsky, Y. Z. Khimyak and A. I. Cooper, *Angew. Chem. Int. Ed.*, 2007, **46**, 8574-8578.
20. Y. Xu, S. Jin, H. Xu, A. Nagai and D. Jiang, *Chem. Soc. Rev.*, 2013, **42**, 8012-8031.
21. O. K. Farha, A. M. Spokoyny, B. G. Hauser, Y.-S. Bae, S. E. Brown, R. Q. Snurr, C. A. Mirkin and J. T. Hupp, *Chem. Mater.*, 2009, **21**, 3033-3035.
22. W. Lu, D. Yuan, J. Sculley, D. Zhao, R. Krishna and H.-C. Zhou, *J. Am. Chem. Soc.*, 2011, **133**, 18126-18129.
23. S. Qiao, Z. Du and R. Yang, *J. Mater. Chem. A*, 2014, **2**, 1877-1885.
24. Y. He, S. Xiang and B. Chen, *J. Am. Chem. Soc.*, 2011, **133**, 14570-14573.
25. P. Kaur, J. T. Hupp and S. T. Nguyen, *ACS Catal.*, 2011, **1**, 819-835.
26. Y. Xie, T.-T. Wang, X.-H. Liu, K. Zou and W.-Q. Deng, *Nat. Commun.*, 2013, **4**.
27. Y. Kou, Y. Xu, Z. Guo and D. Jiang, *Angew. Chem. Int. Ed.*, 2011, **50**, 8753-8757.
28. F. Xu, X. Chen, Z. Tang, D. Wu, R. Fu and D. Jiang, *Chem. Commun.*, 2014, **50**, 4788-4790.
29. K. Sakaushi, E. Hosono, G. Nickerl, H. Zhou, S. Kaskel and J. Eckert, *J. Power Sources*, 2014, **245**, 553-556.
30. K. Sakaushi, E. Hosono, G. Nickerl, T. Gemming, H. Zhou, S. Kaskel and J. Eckert, *Nat. Commun.*, 2013, **4**, 1485.
31. Q. Chen, M. Luo, P. Hammershoj, D. Zhou, Y. Han, B. W. Laursen, C.-G. Yan and B.-H. Han, *J. Am. Chem. Soc.*, 2012, **134**, 6084-6087.
32. A. R. Paniego, *Anales De Quimica Serie a-Quimica Fisica Y Quimica Tecnica*, 1989, **85**, 386-399.
33. J. Weber, J. Schmidt, A. Thomas and W. Boehlmann, *Langmuir*, 2010, **26**, 15650-15656.
34. G. Wang, X. Shen, J. Yao and J. Park, *Carbon*, 2009, **47**, 2049-2053.
35. Y. Yan, B. Hao, D. Wang, G. Chen, E. Markweg, A. Albrecht and P. Schaaf, *J. Mater. Chem. A*, 2013, **1**, 14507-14513.
36. K. M. Abraham, *Solid State Ionics*, 1982, **7**, 199-212.
37. W. Deng, X. Liang, X. Wu, J. Qian, Y. Cao, X. Ai, J. Feng and H. Yang, *Sci. Rep.*, 2013, **3**, 2671.
38. C. Zhu, X. Mu, P. A. van Aken, Y. Yu and J. Maier, *Angew. Chem. Int. Ed.*, 2014, **53**, 2152-2156.



Conjugated microporous polymers are explored as high-performance electrode materials for lithium and sodium storage, exhibiting excellent electrochemical performance.