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Structure and chemistry of surface-doped Pt:SnO₂ gas sensing materials

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Surface-doped Pt:SnO₂ was synthesized by impregnation of calcined SnO₂ made by an aqueous sol-gel route. The structure of the introduced Pt-dopant and its behaviour during gas exposure were examined by in-situ and operando X-ray absorption spectroscopy. The results reveal that Pt forms a nano-sized PtO₂ phase, which was not found for bulk and surface doped materials, studied previously. In a comparative investigation of undoped and Pt-doped SnO₂ gas sensors the performance and the surface chemistry were investigated, the latter one using operando FT-IR spectroscopy. The results prove a strong influence of the different Pt structures on the surface chemistry of SnO₂, providing the basis for an understanding on the varying sensor performance of differently synthesized Pt:SnO₂ gas sensing materials.

1. Introduction

Gas sensors based on semiconducting metal oxides, first and foremost SnO₂, are widely used for the detection of flammable and toxic gases such as CO, H₂ or CH₄.^{1,2} Under typical gas sensing conditions - high background humidity and presence of interfering gases - pure SnO₂ materials show a strongly decreased sensor signal, sensitivity and stability. In order to overcome this disadvantage noble metals like Pd³⁻¹² or Pt^{3,5,13-18} are added to SnO₂, changing its surface chemistry and/or electrical properties.^{17,18} Previous studies have shown, that already small differences in the introduction method of the noble metal, (doping before or after calcination, calcination temperature or noble metal concentration) lead to very different dopant structures and thus gas sensing properties.^{3,5,19,20} However, these results were obtained by ex-situ techniques, like photoelectron spectroscopy or electron microscopy, and the derived structural parameters can thus not be directly applied to explain the effect of noble metal doping during gas sensing.^{21,22} An improved approach to

understand the role of noble metals in gas sensing materials is the use of operando spectroscopic techniques which allow the examination of gas sensors under operating conditions. X-ray absorption spectroscopy (XAS) has been proven to be a suitable operando techniques to study structural parameters of even low concentrated noble metals (0.2 %wt) in gas sensing materials, by interpretation of the extended X-ray absorption fine structure (EXAFS), and the behaviour of the noble metal in various atmospheric conditions, mainly obtained by the analysis of the X-ray absorption near edge structure (XANES).²³⁻²⁵ Complementary to the structural evaluation the surface chemistry of gas sensing layers can be studied by operando diffuse reflectance infrared Fourier-transform spectroscopy (DRIFTS), revealing surface species actively interacting with or caused by the atmosphere, such as carbonylic species on the noble metal surface, carbonates, hydroxyls or M-O overtones on SnO₂, which are considered to be fundamental for understanding the gas reception mechanism.^{9,22,26-28} The potential of these methods was demonstrated when examining SnO₂:Pt which was doped before the final calcination steps.²⁰ The structural evaluation by XAS showed that Pt replaces Sn⁴⁺ ions in the SnO₂ lattice and that it is atomically distributed on the surface and in the bulk of the whole material.²⁹ As observed by DRIFTS the atomically distributed Pt⁴⁺ sites change the chemistry of hydrogen containing gases (H₂, H₂O) and hence improve the CO sensing performance of this material.¹³ Furthermore, it was shown that Pt⁴⁺ in the bulk of the material has a strong influence on the electrical properties of the SnO₂ and consequently the improved gas sensing properties were traced back to changes in the chemistry and the electronic properties of the material.³⁰ This raises the question which effect Pt

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exhibits on SnO₂ if it is only present on the surface of the material.

2. Experimental

2.1. Sample preparation

Undoped tin dioxide was synthesized by an aqueous sol-gel route starting with SnCl₄ (Merck, purified by distillation). The precipitated solid was separated, washed several times, dried (120 °C), calcined at 1000 °C for 8 h under air and milled afterwards. The Pt loadings were introduced either before the calcination step (gel-impregnation, GI) or after the calcination step (powder-impregnation, PI) as an aqueous solution of H₂PtCl₆ (PtCl₄, Aldrich, ≥99.99 % trace metal basis) added to the SnO₂ suspension and stirred for 48 h at room temperature. The PI material was additionally thermally treated after the impregnation (450°C, 1h). The prepared doped materials aimed at a nominal Pt loading of 0.2 % wt. Gas sensors were made by screen printing a paste, made from undoped or Pt-doped SnO₂ powders and an organic binder (propandiol), on alumina substrates equipped with interdigitated electrodes and a backside heating meander.³¹ Pt-electrodes and Pt-heaters were used for gas sensor characterization by DC-resistance measurements and IR spectroscopy. For the X-ray spectroscopic analysis of the Pt-doped sample alumina substrates with Au-electrodes and an Ag/Pd-alloy-heater were used to avoid interference of Pt signals (absorption/fluorescence) from the substrates during XAS measurements.

2.2. General experimental setup

For all experiments similar basic setups were used, allowing to control the atmospheric composition, the temperature of the sensor and the readout of the sensor response.²² Gases were mixed using home-made gas mixing stations based on mass flow controllers (Bronkhorst EL-Flow select). Test gases mixtures (H₂ in synthetic air or He) were added to a carrier gas stream (synthetic air or He). Humidity levels were dosed by using evaporators filled with deionized water. The sensors were heated by applying voltage and current to the backside heaters using a DC-powder supply (Agilent E 3630A) and adjusting the exact values according to the sensor's temperature calibration. All experiments were conducted at 300 °C if not specified differently. To characterize the sensor responses as such and during all operando spectroscopic experiments the resistance was measured using a digital multimeter (Keithley 2000). All experiments were performed by measuring one sensor at a time to avoid downstream effects, which are expected for Pt-doped SnO₂ gas sensors.¹⁵

2.3. Operando X-ray absorption spectroscopy

In situ and operando high-energy-resolution fluorescence detected (HERFD) X-ray absorption spectroscopy measurements were performed at beamline ID26 at the European Synchrotron Radiation Facility (ESRF, Grenoble,

France). At this experimental station the beam was provided by a set of mechanically independent undulators. The incident energy was selected by a Si (111) double crystal monochromator. The beam size on the sample had an area 0.2 x 0.5 mm² (vertical x horizontal) and the energy calibration was performed using a Pt foil. The spectra were recorded in high-energy-resolution fluorescence detection by means of an X-ray emission spectrometer. The sample, analyzer crystals and avalanche photodiode were displaced in a vertical Rowland geometry. The Pt L₃-edge HERFD spectra were measured recording the intensity of the Pt Lα₁ emission line (9442 eV) as a function of the incident energy. The emission energy was selected using the <660> reflection of four spherically bent Ge crystal analyzers. The broadening of the elastic peak was 1.8 eV. The sensors were mounted in a special home-made cell which allowed recording both HERFD spectra and electrical resistance as function of the atmosphere composition and temperature.²⁴ XAS data analysis was carried out using Athena and Artemis software of the IFEFFIT package.³² The spectra were energy calibrated, background subtracted and then normalized. The structural parameters were obtained by adjusting theoretical structural models, *ab initio* calculated phase and amplitude functions using the FEFF 6.0 code³³, to the experimental data. The EXAFS spectra in k space were Fourier Transformed between 3.0-11.5 Å⁻¹ and the data fitting was performed in R space between 1.0-2.2 Å for the first coordination shell, and between 1-4 Å when further shells were considered. The number of fitted parameters was always lower than the number of independent ones.

The ratio between the oxidized and reduced Pt species were estimated by linear combination fitting (LCF) of operando XANES spectra using Athena IFEFFIT.³² The linear combination fitting was performed in the spectral range of -20 and 70 eV referred to the absorption edge. This procedure allowed tracking the proportion of oxidized and reduced Pt species. Since the references for the linear combinations were the first spectrum of the series and a spectrum of a Pt foil (both recorded at room temperature) the calculated ratios are relative values.

2.4. Operando DRIFT spectroscopy

Operando FT-IR spectroscopy was done on a Bruker Vertex70v spectrometer equipped with an external high performance globar, a KBr beamsplitter and a mid-band MCT detector. All samples were measured in diffuse reflectance geometry using a mirror optic (Praying Mantis, Harrick) with a home-made operando cell placed inside, which allows simultaneous resistance and spectroscopic measurements.⁹ Single channel spectra were recorded averaging 512 scans with a spectral resolution of 1 cm⁻¹. The as obtained single channel spectra were processed with Bruker's OPUS software (version 7.2). Absorbance spectra were calculated as apparent absorbance using a single channel spectrum of the sensor before test gas exposure as reference.^{9,28}

3. Results

3.1. Gas sensing performance

Figure 1 shows the sensor signals to CO of undoped and Pt-doped materials in dry and humid air. Considering first the sensor performance in dry air (empty symbols) the undoped material shows the highest sensor signals to CO, followed by Pt:SnO₂ GI, while Pt:SnO₂ PI shows the lowest sensor signals. However, in the presence of water vapour (filled symbols) the situation changes: The undoped material shows a substantial decrease of the sensor signals compared to dry air, while the sensor signals of Pt:SnO₂ PI are strongly increased. The Pt:SnO₂ GI sensor shows only a minor increase, compared to the one observed for Pt:SnO₂ PI. In order to understand these findings, the differences in the structure of the platinum dopant and the resulting changes of the materials chemistry have to be studied.

3.2. Structural investigations

As basis for understanding the effect of Pt surface doping the structure, oxidation state and dispersion of Pt need to be unravelled. For GI materials a detailed XAS investigation has reported by M. Hübner et al.²⁵, showing that Pt replaces Sn⁴⁺ ions in the SnO₂ lattice and that it is distributed in the whole material.²⁵ Similar results were recently reported by N. Murata et al.¹⁶, finding a up to 10 %at of Pt incorporated in SnO₂. Hence, the following section focuses on the PI material. Figure 2 presents the Fourier-transform of ex-situ recorded EXAFS spectra of PtO₂ (bulk) reference and Pt:SnO₂ PI. The peaks shown in figure 2 are related to the backscattering of the nearest neighbours and thus the atomic radial distribution function (not phase-shifted) around the Pt absorber atoms. The visual inspection already reveals that the first coordination shell around the Pt atoms of PtO₂ and Pt:SnO₂ PI is the same, while the outer backscattering contributions (2nd and 3rd peak) are dissimilar. This indicates that the Pt dopant in SnO₂ forms a PtO₂-like phase which has not completely the same structure like the bulk oxide (PtO₂). Quantitative structural information was obtained by fitting theoretical calculated spectra to the experimental ones; the structural parameters are presented in table 1 and the refined spectra of the experimental data and the reference materials are shown in figure 3 and ESI 1 to 2, respectively. The first hypothetical structural model considered that the chemical environment of Pt atoms in Pt:SnO₂ PI is similar to one found in PtO₂, this approach was satisfactory to describe the first coordination shell; however it failed to fit to the outer shells. The Pt atoms in SnO₂:Pt presented 6.0 oxygen atoms at 2.02 Å, it means that coordination number and distance fits to the values observed for PtO₂ bulk reference. Interestingly, the mean-square deviation of interatomic distances (σ^2) is considerably higher for the sensing material than for the reference bulk oxide. This parameter accounts for the structural disorder, at low temperatures it expresses that static disorder of the pair distribution function and at high temperatures it reveals the thermal disorder summed to the static one. As long as all ex-situ measurements were carried out at room temperature, the high σ^2 suggest that Pt atoms in the sensor powder are in

disordered environment that could result from the low particle size. To describe the outer shells FEFF 6.0 generated paths from PtO₂ and SnO₂ models were used. A constraint was applied to the second coordination forcing the number of neighbours to be ≤ 6 like it was found for PtO₂. The analysis shows that Sn atoms are required to fit the model to the experimental spectra. Pt:SnO₂ PI presents Pt neighbours at 3.14±0.05 Å and Sn atoms at 3.63±0.03 Å. This differs from the previous study where Pt was introduced before the calcination step: In this study the 2nd coordination shell contained only Sn atoms; this was interpreted as an atomic distribution of Pt, replacing Sn atoms in the SnO₂ lattice.²⁵ The EXAFS analysis of Pt:SnO₂ PI has to be interpreted as the formation of a nano-sized PtO₂ phase on the surface or inside of the SnO₂ material. To determine the distribution of the Pt dopant temperature programmed reduction (TPR) was performed on a stepwise heated sensor in a 0.5 % vol. H₂/He atmosphere, the recorded XANES spectra (figure 4a) were quantitatively analysed by LCF analysis, as shown in figure 4b. The Pt:SnO₂ PI material is already reduced at room temperature in 0.5 % vol. H₂/He; the stepwise increase of the sensor's temperature leads to a further reduction of the Pt, which ends in the formation of a mainly metallic Pt at 300 °C. Switching from the reducing atmosphere to synthetic air causes a re-oxidation of the Pt at 300 °C. Since the re-oxidation is a slow process it did not reach the initial Pt composition in the time allotted to the re-oxidation experiment (90 minutes). In contrast Pt:SnO₂ GI is only partially reduced under much harsher conditions (2 % vol. H₂/He at 600°C, see ESI 3). The clear difference in the reducibility of Pt for the two samples points out that for Pt:SnO₂ PI the Pt is accessible to gases, i.e. only present at the surface and not incorporated in the SnO₂ lattice. The Pt oxidation state strongly depends on the composition of the atmosphere, i.e. the Pt actively interacts with reducing and oxidizing gases.

The structural evaluation presented above revealed that powder-impregnation leads to the formation of a nano-sized PtO₂ phase mostly present at the surface of SnO₂ (figure 5a), which strongly differs from the structure Hübner et al.²⁵ proposed for gel-impregnated materials, for which Pt is atomically dispersed in the SnO₂ bulk and surface (figure 5b). A similar occurrence of atomically distributed Pd or a separate oxide phase is observed when comparing Pd-doped SnO₂ materials made by gel-impregnation or flame spray pyrolysis (dopant introduced before/during the grain formation) with powder-impregnated materials (dopant introduced after the grain formation).^{4,24}

3.3. Operando XAS

The fundamental models for the enhanced gas sensing performance of noble metal loaded SnO₂ are closely linked to the oxidation state of the noble metal: For metallic clusters a chemical sensitization by a spill-over mechanism is expected, while for partial oxidized clusters - being reduced during test gas exposure - an electronic sensitization by the alignment of the Fermi-levels is a commonly accepted model.^{17,18} Hence,

the oxidation state of Pt under operation conditions is of fundamental interest to understand the role of Pt for gas sensing. Since the TPR experiments showed that PtO₂ phase is easily reduced and re-oxidized at 300 °C the question raises which oxidation state of Pt is dominant under operation conditions in the absence and presence of H₂. Figure 6 shows XANES recorded from heated sensors in dry air and dry air containing 50 ppm H₂. For both materials the presence of H₂ leads to a slight decrease of the white line (first absorption feature above the Pt L₃ edge), which indicates a weak reduction of the Pt, but for none of the materials a shift or strong decrease of the white line, which indicates the formation of metallic Pt, is observed. In the case of Pt:SnO₂ PI the lack of a considerable reduction can be explained with the high excess of O₂ – 20.5 % O₂ in comparison with 50 ppm H₂ – which favours the re-oxidation and prevents the formation of a metallic Pt phase under operating conditions. As Pt:SnO₂ GI is not reducible even at higher H₂ concentrations and in the absence of oxygen, it is expected that it is not reduced by very low H₂ concentration in air. The XAS spectra shown in figure 6 do not show any differences between Pt:SnO₂ PI and GI, i.e. for both materials oxidized Pt dominates under operating conditions.

3.4. Operando DRIFTS

Pt-based catalysts are widely used for the catalytic oxidation of CO.^{34,35} Oxidized^{36–40} as well as reduced^{35,41–44} Pt species are reported to be catalytically active, whereupon the latter one is found to be more reactive.^{41,42} Thus it is reasonable to assume that Pt, in either form, chemically sensitizes the SnO₂ material and that such a chemical sensitization should result in a changed surface chemistry, as it was demonstrated for Pt:SnO₂ GI.¹³

Figure 7 shows DRIFT spectra of both Pt-doped materials recorded during CO exposure in dry and humid air. In agreement with previous works undoped SnO₂ shows no carbonyl species under operation conditions (spectra not shown)⁴⁵. During CO exposure both Pt-doped samples (figure 7a and b) show the formation of different carbonyl species; due to the absence of carbonyls on undoped SnO₂, these carbonyls are identified as platinum carbonyl species (Pt-CO). On both materials only carbonyl species assigned to oxidized platinum are observed, while no carbonyl bands on metallic Pt can be assigned.⁴⁶ This supports the findings by operando XAS, which did not show the formation of metallic Pt under operation conditions. A closer look reveals, that Pt:SnO₂ PI shows two carbonyl species at 2148 and 2131 or 2135 cm⁻¹ (figure 7 a), while Pt:SnO₂ GI shows only one carbonyl species at 2124 cm⁻¹ (figure 7 b). Again this is in line with the structure proposed for both materials: In case of Pt:SnO₂ PI the presence of PtO₂ clusters offers a higher heterogeneity of Pt-sites – e.g. edges and planar surface regions – available for CO adsorption; which is in line with the appearance of several different carbonyl species. In case of Pt:SnO₂ GI the atomically distributed Pt should be found in a homogeneous surrounding,

which is in line with the appearance of a single carbonyl species.

Since the carbonyls on Pt:SnO₂ PI appear at higher wavenumbers than the carbonyl species on Pt:SnO₂ GI (see figure 7) an incorporation of Pt into the SnO₂, as it is observed for Pt:SnO₂ GI, can be ruled out; again this is in line with XAS results. Consequently, the occurrence of carbonyl species on the Pt-doped samples point out that the sensitization caused by Pt-doping is related to changed chemistry, which is linked closely to the structure of the Pt doping. Regarding the effect of humidity on the carbonyl species the following was observed. On Pt:SnO₂ PI the presence of water vapour strongly decreased the intensity of the Pt-CO bands, indicating a lower CO adsorption on the PtO₂ phase, which results in a lower level of CO oxidation on the PtO₂ phase. On the opposite, for the Pt:SnO₂ GI the intensity of the Pt-CO band is increased by the humidity, indicating a higher CO adsorption and hence an increased CO oxidation activity of the atomic Pt sites (in humid conditions). These findings suggest that the differences in the sensor performance (figure 1) are most likely related to the different chemical activity of the Pt sites.

4. Discussion

The Pt-structures, revealed by XAS, create new reaction sites in addition to the Sn-O-Sn sites present on undoped (pristine) SnO₂:

For Pt:SnO₂ GI the atomically distributed Pt offers new reactive surface sites (Pt-O-Sn), which separate the oxidation of CO from interfering compounds like water vapor.¹³ While the pristine SnO₂ surface is deactivated by water vapour⁴⁵, the Pt-O-Sn sites on Pt:SnO₂ remain active maintaining the sensing activity (figure 1). The increase of the carbonyl species in humid air can be explained by a decreased level of CO oxidation onto the (pristine) SnO₂ surface, which causes a shift for the surface reaction to the atomic Pt sites. This compensation of the decreased activity of SnO₂ surface explains the independence of the CO sensor signal from the water vapour in the background (figure 1).

For Pt:SnO₂ PI three different reactive sites are possible: the SnO₂ surface (Sn-O-Sn), the PtO₂ surface (Pt-O-Pt) or the boundary of the two oxides (Pt-O-Sn) and the gas phase, the so called 3-phase-boundary. The latter one, the Pt-O-Sn site at the 3-phase-boundary, may differ from the Pt-O-Sn sites found for atomically distributed Pt.

Regarding the situation in dry air the operando DRIFT spectrum (figure 7a) suggests that the reaction is mainly taking place at the PtO₂ surface (Pt-O-Pt sites), indicated by the appearance of carbonyl species (2148 and 2131 cm⁻¹) and in line with the TPR XANES results, proving the feasible reduction and re-oxidation of the PtO₂ phase. Assuming that there is no electron exchange between SnO₂ and PtO₂, any reaction at the Pt-O-Pt sites does not change the resistance of SnO₂ and hence has no impact on the sensor signal. The CO oxidized on the PtO₂ phase is not anymore available for a reaction on the SnO₂ phase (Sn-O-Sn or Sn-O-Pt sites) and consequently the parallel reaction on the PtO₂ phase will decrease the sensor signals for

a certain concentration, as it is observed for Pt:SnO₂ PI in dry air (figure 1). In humid air the situation changes: the carbonyl bands decrease strongly, i.e. the Pt-O-Pt sites deactivated, and the reaction is shifted to the SnO₂ surface or the 3-phase-boundary. As the pristine SnO₂ surface is deactivated in humid air, the oxidation of CO in humid air on Pt:SnO₂ PI has to take place at the Pt-O-Sn sites, i.e. the 3-phase-boundary. A reaction at the 3-phase-boundary, which keeps the surface active in the presence of water vapour while the PtO₂ phase gets deactivated, gives an appropriate explanation for the high sensor performance of Pt:SnO₂ PI in humid air (figure 1).

5. Conclusions

By the use of complementary operando spectroscopic techniques – XAS and DRIFTS – and the direct comparison with analogous data of a related material, structure and surface chemistry of surface doped Pt:SnO₂ were revealed and provided the basis for better understanding of the CO gas sensing mechanism of Pt doped SnO₂. In contrast to atomically distributed Pt (gel-impregnation) it was found that introduced after the calcination step (powder-impregnation) Pt forms a separate oxide phase, which offers additional reaction sites, which are not electronically coupled to the SnO₂. Hence, the decreased sensor signals in dry air are caused by a parallel reaction on the PtO₂ phase with no effect onto the electrical properties of SnO₂. Nevertheless, the presence of PtO₂ clusters at the surface of SnO₂ prevents a complete deactivation of the SnO₂ surface in humid air, resulting in a higher CO sensing performance in the presence of water vapour.

This work demonstrated the fundamental role of structure and distribution of noble metals in oxide materials for the surface chemistry, which has a direct link to the materials functionality - in this case, gas sensing. Furthermore, it has been shown that the used of complementary experimental techniques and the comparison of rational designed materials is a powerful method to study the structure-function relationships of functional materials, such as gas sensors or catalysts.

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References

- 1 A. Gurlo, N. Barsan and W. Udo, in *Metal Oxides - Chemistry and Applications*, ed. J. L. G. Fierro, CRC Press, 2006, pp. 683–738.
- 2 R. Jaaniso and O. Kiang Tan, *Semiconductor Gas Sensors*, Woodhead Publishing Limited, 2013.
- 3 A. Cabot, A. Diéguez, A. Romano-Rodríguez, J. R. Morante and N. Barsan, *Sensors Actuators B Chem.*, 2001, **79**, 98–106.
- 4 D. Degler, H. W. P. De Carvalho, U. Weimar, N. Barsan, D. Pham, L. Mädler and J.-D. Grunwaldt, *Sensors Actuators B Chem.*, 2015, **219**, 315–323.

- 5 A. Diéguez, A. Vilà, A. Cabot, A. Romano-Rodríguez, J. R. Morante, J. Kappler, N. Barsan, U. Weimar and W. Göpel, *Sensors Actuators B Chem.*, 2000, **68**, 94–99.
- 6 M. Epifani, J. Arbiol, E. Pellicer, E. Comini, P. Siciliano, G. Faglia and J. R. Morante, *Cryst. Growth Des.*, 2008, **8**, 1774–1778.
- 7 G. Ghiotti, A. Chiorino, G. Martinelli and C. Carotta, *Sensors Actuators B Chem.*, 1995, **25**, 520–524.
- 8 F. Gyger, A. Sackmann, M. Hübner, P. Bockstaller, D. Gerthsen, H. Lichtenberg, J.-D. Grunwaldt, N. Barsan, U. Weimar and C. Feldmann, *Part. Part. Syst. Charact.*, 2014, **31**, 591–596.
- 9 S. Harbeck, A. Szatvanyi, N. Barsan, U. Weimar and V. Hoffmann, *Thin Solid Films*, 2003, **436**, 76–83.
- 10 R. G. Pavelko, H. Daly, M. Hu, C. Hardacre and E. Llobet, *J. Phys. Chem. C*, 2013, **117**, 4158–4167.
- 11 T. Wagner, M. Bauer, T. Sauerwald, C. Kohl and M. Tiemann, *Thin Solid Films*, 2011, **520**, 909–912.
- 12 A. Marikutsa, M. Rumyantseva and A. Gaskov, *J. Phys. Chem. C*, 2015, **119**, 24342–24350.
- 13 K. Großmann, S. Wicker, U. Weimar and N. Barsan, *Phys. Chem. Chem. Phys.*, 2013, **15**, 19151–8.
- 14 I. Kocemba and J. Rynkowski, *Sensors Actuators, B Chem.*, 2011, **155**, 659–666.
- 15 L. Mädler, T. Sahm, A. Gurlo, J.-D. Grunwaldt, N. Barsan, U. Weimar and S. Pratsinis, *J. Nanoparticle Res.*, 2006, **8**, 783–796.
- 16 N. Murata, T. Suzuki, M. Kobayashi, F. Togoh and K. Asakura, *Phys. Chem. Chem. Phys.*, 2013, **15**, 17938.
- 17 N. Yamazoe, *Sensors Actuators B Chem.*, 1991, **5**, 7–19.
- 18 N. Yamazoe, Y. Kurokawa and T. Seiyama, *Sensors Actuators B Chem.*, 1983, **4**, 283–289.
- 19 J. Kappler, N. Barsan, U. Weimar, A. Diéguez, J. L. Alay, Romano-Rodríguez, J. R. Morante and W. Göpel, *Fresenius. J. Anal. Chem.*, 1998, **361**, 110–114.
- 20 A. Cabot, J. Arbiol, J. R. Morante, U. Weimar, N. Barsan and W. Göpel, *Sensors Actuators B Chem.*, 2000, **70**, 87–100.
- 21 A. Gurlo and R. Riedel, *Angew. Chem. Int. Ed. Engl.*, 2007, **46**, 3826–48.
- 22 N. Barsan, D. Koziej and U. Weimar, *Sensors Actuators B Chem.*, 2007, **121**, 18–35.
- 23 M. Hübner, D. Koziej, J.-D. Grunwaldt, U. Weimar and N. Barsan, *Phys. Chem. Chem. Phys.*, 2012, **14**, 13249–13254.
- 24 D. Koziej, M. Hübner, N. Barsan, U. Weimar, M. Sikora and J.-D. Grunwaldt, *Phys. Chem. Chem. Phys.*, 2009, **11**, 8620–8625.
- 25 M. Hübner, D. Koziej, M. Bauer, N. Barsan, K. Kvashnina, M. D. Rossell, U. Weimar and J.-D. Grunwaldt, *Angew. Chem. Int. Ed. Engl.*, 2011, **50**, 2841–4.
- 26 D. Koziej, K. Thomas, N. Barsan, F. Thibault-Starzyk and U. Weimar, *Catal. Today*, 2007, **126**, 211–218.
- 27 D. Koziej, N. Barsan, V. Hoffmann, J. Szuber and U. Weimar, *Sensors Actuators B Chem.*, 2005, **108**, 75–83.
- 28 K. Großmann, R. G. Pavelko, N. Barsan and U. Weimar, *Sensors Actuators B Chem.*, 2012, **166-167**, 787–793.
- 29 M. Hübner, D. Koziej, M. Bauer, N. Barsan, K. Kvashnina, M. D. Rossell, U. Weimar and J.-D. Grunwaldt, *Angew. Chemie*, 2011, **123**, 2893–2896.
- 30 M. Hübner, N. Barsan and U. Weimar, *Sensors Actuators B Chem.*, 2012, **171-172**, 172–180.
- 31 N. Barsan and U. Weimar, *J. Phys. Condens. Matter*, 2003,

ARTICLE

Journal Name

- 15, R813–R839.
- 32 B. Ravel and M. Newville, *J. Synchrotron Radiat.*, 2005, **12**, 537–41.
- 33 J. J. Rehr and R. C. Albers, *Rev. Mod. Phys.*, 2000, **72**, 621–654.
- 34 A. Russell and W. S. Epling, *Catal. Reviews Sci. Eng.*, 2011, **53**, 337–432.
- 35 G. Ertl, *Angew. Chemie - Int. Ed.*, 2008, **47**, 3524–3535.
- 36 M. D. Ackermann, T. M. Pedersen, B. L. M. Hendriksen, O. Robach, S. C. Bobaru, I. Popa, C. Quiros, H. Kim, B. Hammer, S. Ferrer and J. W. M. Frenken, *Phys. Rev. Lett.*, 2005, **95**, 255505.
- 37 E. M. C. Alayon, J. Singh, M. Nachtegaal, M. Harfouche and J. A. van Bokhoven, *J. Phys. Conf. Ser.*, 2009, **190**, 012152.
- 38 B. L. M. Hendriksen and J. W. M. Frenken, *Phys. Rev. Lett.*, 2002, **89**, 046101.
- 39 J. Singh, M. Nachtegaal, E. M. C. Alayon, J. Stötzl and J. a. Van Bokhoven, *ChemCatChem*, 2010, **2**, 653–657.
- 40 E. M. C. Alayon, J. Singh, M. Nachtegaal, M. Harfouche and J. A. van Bokhoven, *J. Catal.*, 2009, **263**, 228–238.
- 41 A. Boubnov, A. Gänzler, S. Conrad, M. Casapu and J.-D. Grunwaldt, *Top. Catal.*, 2013, **56**, 333–338.
- 42 A. M. Gänzler, M. Casapu, A. Boubnov, O. Müller, S. Conrad, H. Lichtenberg, R. Frahm and J.-D. Grunwaldt, *J. Catal.*, 2015, **328**, 216–224.
- 43 J. Wintterlin, S. Våkening, T. V. W. Janssens, T. Zambelli and G. Ertl, *Science (80-.)*, 1997, **278**, 1931–1934.
- 44 F. J. Gracia, L. Bollmann, E. E. Wolf, J. T. Miller and a. J. Kropf, *J. Catal.*, 2003, **220**, 382–391.
- 45 D. Degler, S. Wicker, U. Weimar and N. Barsan, *J. Phys. Chem. C*, 2015, **119**, 11792–11799.
- 46 E. Ivanova, M. Mihaylov, F. Thibault-Starzyk, M. Daturi and K. Hadjiivanov, *J. Mol. Catal. A Chem.*, 2007, **274**, 179–184.

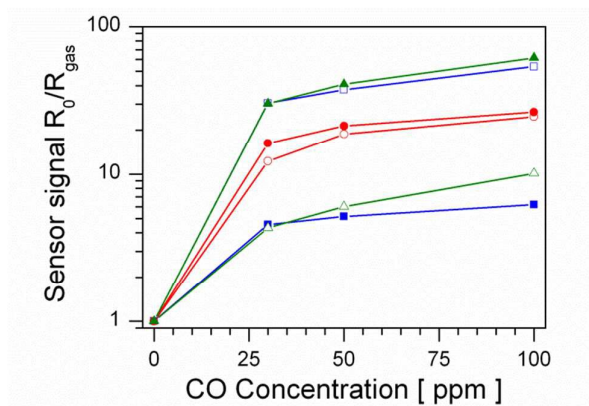


Figure 1 – Sensor signals of undoped SnO₂ (blue squares), Pt:SnO₂ GI (red circles) and Pt:SnO₂ PI (green triangles) in dry (empty symbols) and humid (50 % r.h. at 25 °C, filled symbols) air.

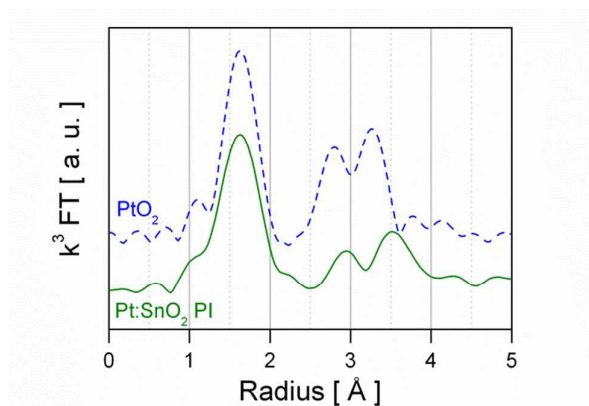


Figure 2 – k^3 weighted Fourier Transform of ex-situ recorded EXAFS spectra from reference PtO₂ (blue dashed line) and SnO₂:Pt (green straight line).

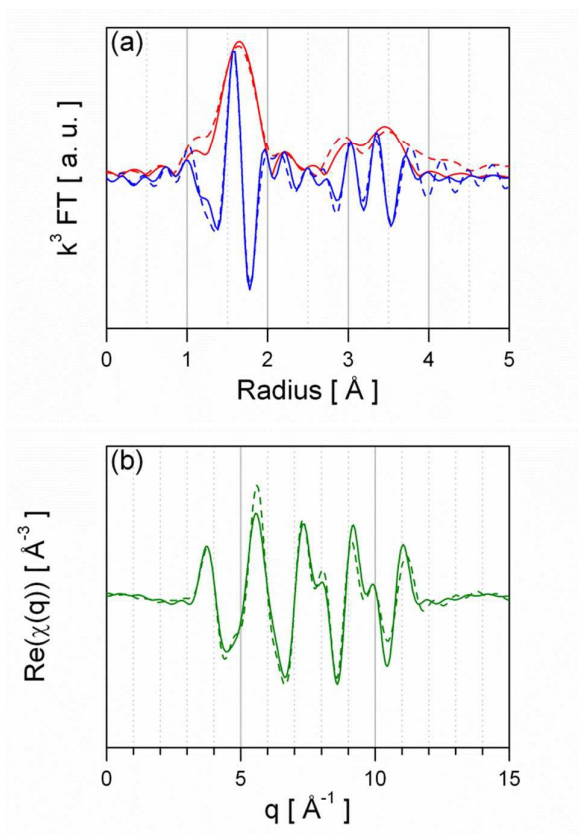


Figure 3 – Real part (a) and imaginary part (b) of the refined EXAFS fits of Pt:SnO₂ PI. The dashed line represents the experimental data and the solid line the fit function.

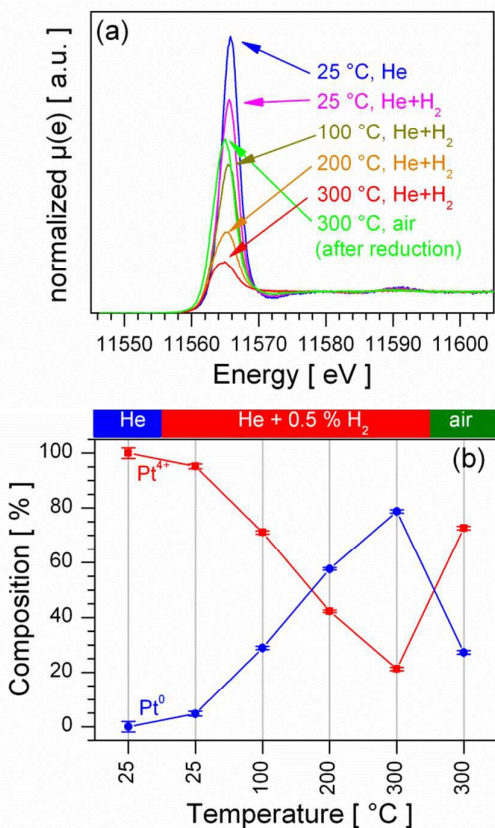


Figure 4 – TPR-XANES spectra recorded from Pt:SnO₂ PI sensor in different conditions (a) and the corresponding LCF calculations (b). In figure b the circles and line show the concentration of Pt⁴⁺, while the concentration of Pt⁰ is represented by the blue circles and line.

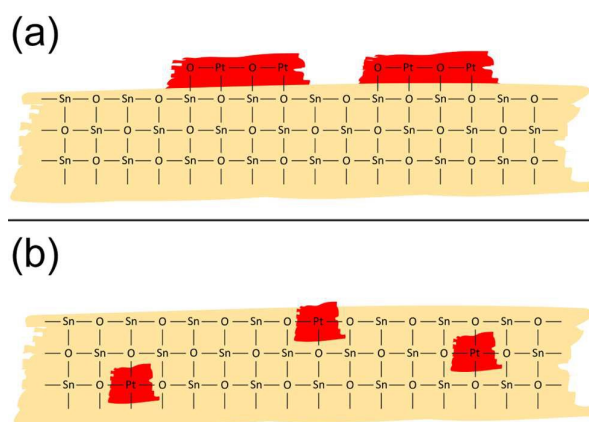


Figure 5 – Schematic representation of the Pt structure as proposed by EXAFS and XANES analysis for Pt:SnO₂ PI (a) and Pt:SnO₂ GI (b).

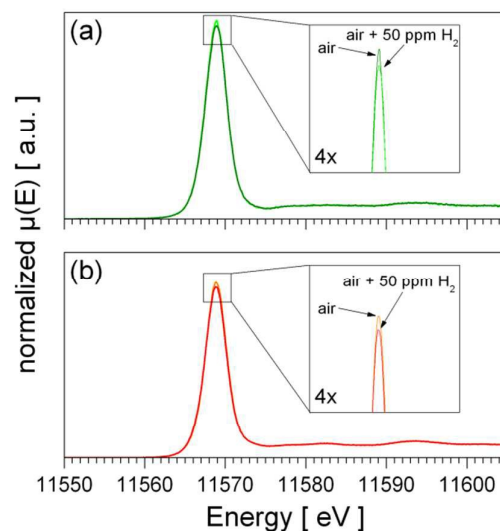


Figure 6 – Operando XANES spectra of Pt:SnO₂ PI (a) and Pt:SnO₂ GI (b) recorded at 300 °C in dry air and dry air plus 50 ppm H₂. The spectra shown in (b) is based on the raw data measured by M. Hübner et al.²⁵ and was treated similar to the data shown in (a).

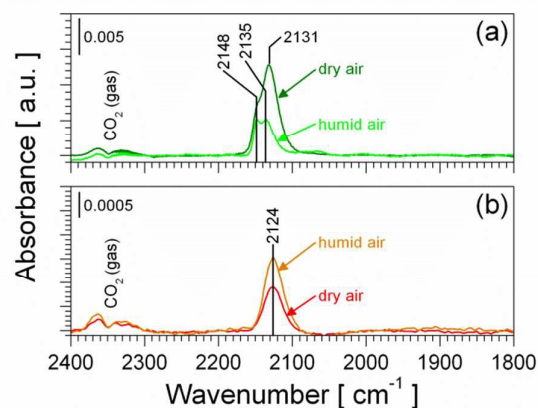


Figure 7 – operando DRIFT spectra of Pt:SnO₂ PI (a) and Pt:SnO₂ GI (b) recorded during the exposure of 100 ppm CO in dry air (reference: dry air) and humid air (reference: humid air) at 300 °C. The spectra shown in (b) are based on the experimental data previously published by K. Großmann et al.¹³

Table 1 – Parameters obtained from EXAFS analysis of the measured SnO₂:Pt PI and reference materials PtO₂ (ESI 1) and SnO₂ (ESI 2).

Atom	N	R [Å]	σ^2 [10 ⁻³ Å ²]	ρ [%]	
SnO ₂ :Pt	O	6.0 ^f	2.02±0.01 ^a	3.1±0.1 ^a	4.7
PI	Pt	2.4±1.4 ^c	3.14±0.05 ^a	4.8±3.0 ^a	
	Sn	3.6±1.4 ^c	3.63±0.03 ^a	4.8±3.0 ^a	
PtO ₂	O	6.0 ^f	2.02±0.01 ^a	2.7±0.9 ^a	4.5
	Pt	6.0 ^f	3.10±0.01 ^a	2.7±0.5 ^a	
SnO ₂	O	6.0 ^f	2.06±0.02 ^a	3.1±0.4 ^a	1.0
	Sn	2.0 ^f	3.21±0.02 ^a	5.2±0.2 ^a	
	O	4.0 ^f	3.60±0.02 ^a	4.1±3.0 ^a	
	Sn	8.0 ^f	3.73±0.02 ^a	4.0±0.2 ^a	

PtO₂: $\Delta E_0 = 12.9 \pm 1.0$ eV, SnO₂:Pt: $\Delta E_0 = 12.7 \pm 1.7$ eV SnO₂: $\Delta E_0 = 5.4 \pm 0.5$ eV;
 Sn K S_o²=1.07, Pt L₃ S_o²=0.83; a = adjusted; f = fixed; c = constrained

