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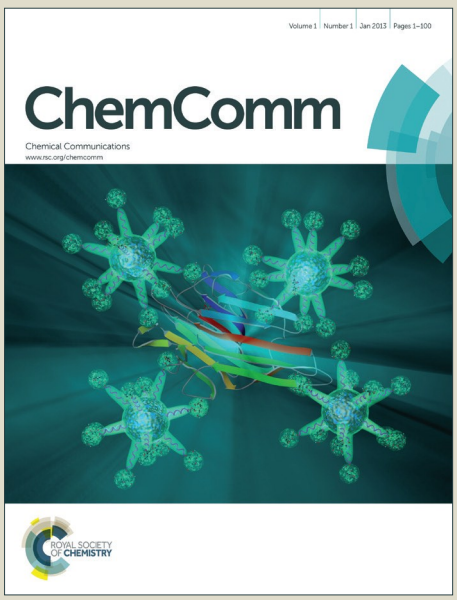

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Photoinduced electron transfer in porous organic salt crystals impregnated with fullerenes†

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Porous organic salt (POS) crystals composed of 9-(4-sulfophenyl)anthracene (SPA) and triphenylmethylamine (TPMA) were impregnated with fullerenes (C_{60} and C_{70}), which were arranged in one dimensional close contact. POS crystals of SPA and TPMA without fullerenes exhibit blue fluorescence due to SPA, whereas the fluorescence was quenched in POS with fullerenes due to electron transfer from the singlet excited state of SPA to fullerenes.

Metal–organic frameworks (MOFs) have merited increasing attention as an interesting platform to design and develop artificial photosynthetic systems, because MOFs can contain light-harvesting and charge-separation units as well as catalytic units in a single crystal, providing the structural organization to integrate these units of artificial photosynthesis into a single crystal.[1] There are many examples of MOFs in which energy and electron transfer have been investigated by fixing the distance and orientation between chromophores, which were precisely determined by single crystal X-ray crystallography.[2-10]

As compared with MOFS, porous organic salts (POSs) have advantage in terms of synthetic versatility, easier processing, more flexibility and better workability.[11-15] In particular, POSs composed of ammonium cations and sulfonate anions have enabled various systematic design of resultant porous structures by changing the combination of ionic components using strong intermolecular interactions such as hydrogen bonds and electrostatic interactions.[16,17] However, there has been no report on photoinduced electron transfer in POS crystals.

We report herein the first example of efficient photoinduced electron transfer in POS crystals composed of 9(4-sulfophenyl)anthracene (SPA) and triphenylmethylamine (TPMA) impregnated with fullerenes (C_{60} and C_{70}) as revealed by femtosecond laser flash photolysis and electronic paramagnetic resonance (EPR) measurements. Fullerenes were chosen as electron acceptors, because they are known to undergo efficient electron-transfer reduction with small reorganization energies.[18] The combination of SPA and TPMA was chosen in order to make enough space for accommodation of fullerenes in the porous structure.

SPA was synthesized according to the literature (see Electronic Supplementary Information (ESI†)).[19] The organic salt composed of SPA and TPMA yielded pale yellow needle crystal of SPA/TPMA suitable for X-ray crystallographic analysis from an *o*-dichlorobenzene solution. The X-ray crystal structure of SPA/TPMA in Fig. 1a revealed that the crystal had a porous structure, which was hierarchically constructed (Fig. S1 in ESI†). Four SPA molecules and four TPMA molecules assembled into tetrahedral supramolecular cluster via cubic-like charge-assisted hydrogen bond.[20] Anthracene moieties of SPA were jutting out in the tetrahedral direction and located around core structure. The tetrahedral clusters accumulated along c-axis by CH-π interaction between anthracene moieties and trityl group to form linear column structure differently from previous diamondoid structures.[21] The

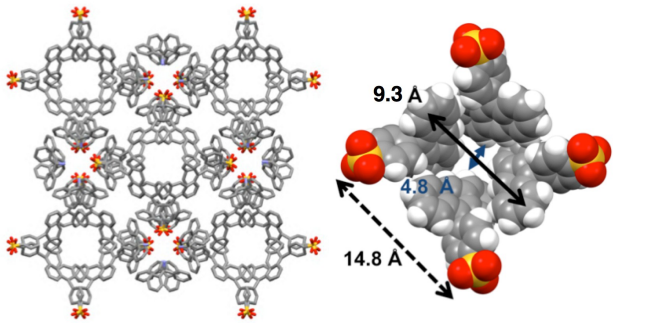

Fig. 1 Crystal structure of SPA/TPMA (left panel). The void space in SPA/TPMA (right panel).

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† Electronic Supplementary Information (ESI) available: Experimental and spectral data. See DOI:

column structures were bundled by CH-π interactions between anthracene moieties to give a bowl-like void space among the columns. The void spaces were connected to lead to channel type porous structure having bottle-neck (Fig. S2 in ESI†). Recrystallization solvents were incorporated there but disordered.

The void space of SPA/TPMA from *o*-dichlorobenzene with the diameter of 9.3 Å (Fig. 1b) is capable to accommodate fullerenes. Recrystallizations of SPA/TPMA with C_{60} and C_{70} gave purple-red and red-brown needle crystals, respectively. X-ray studies revealed that the porous organic salts composed of SPA and TPMA were impregnated with fullerenes as shown in Fig. 2. The crystallographic parameters of SPA/TPMA, SPA/TPMA/C_{60} and SPA/TPMA/C_{70} are shown in Table S1 (ESI†). The void space is enlarged by impregnation of fullerenes to be 10.8 Å and 11.1 Å in SPA/TPMA/C_{60} and SPA/TPMA/C_{70}, respectively (Fig. S3 in ESI†), because of comparatively weak interaction among the columns. This indicates the flexibility of the POS crystals. Moreover, the bottleneck in SPA/TPMA/C_{60} and SPA/TPMA/C_{70} is also expanded to be 6.6 Å and 7.2 Å, respectively. Consequently, the distance between fullerenes is getting closer in the porous structures.

The SPA/TPMA crystal exhibits fluorescence due to the SPA moiety under photoirradiation (Fig. 3).[22] In contrast, the SPA/TPMA/C_{60} and SPA/TPMA/C_{70} exhibited no fluorescence probably because of electron transfer from the singlet excited state of SPA to the fullerene. The occurrence of the photoinduced electron transfer in SPA/TPMA/C_{60} crystals was confirmed by femtosecond laser flash photolysis measurements. The observed transient absorption spectra are shown in Fig. 4, where the transient absorption band at 580 nm of SPA ($^1SPA^*$) was observed upon femtosecond laser excitation and the decay is accompanied by increase in the new absorption band at 700 nm due to the radical cation of SPA ($SPA^{\bullet+}$).[23] The concomitant formation of $C_{60}^{\bullet-}$ was also observed at 1030 nm.[18] Thus, photoinduced electron transfer

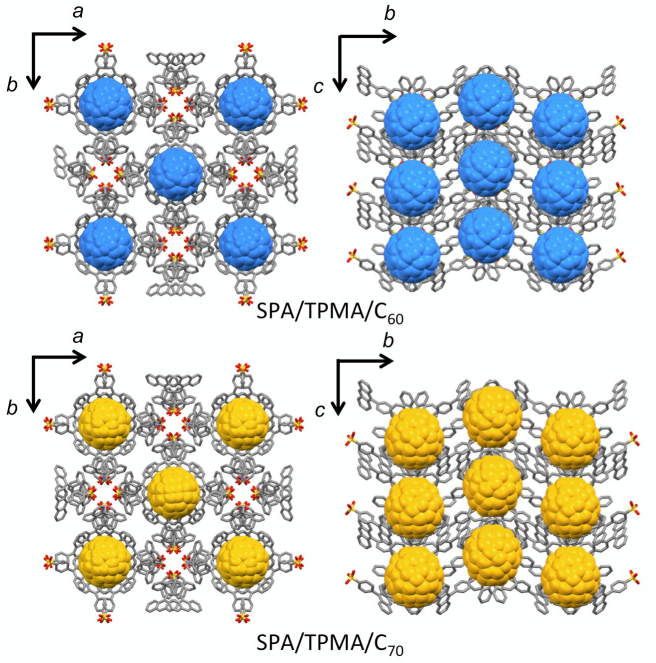

Fig. 2 X-ray crystal structures of SPA/TPMA/C_{60} (upper panels) and SPA/TPMA/C_{70} (lower panels).

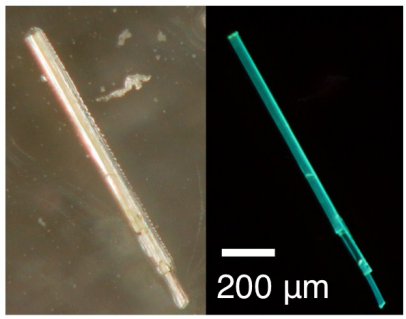

Fig. 3 Colour change of a SPA/TPMA crystal from *o*-dichlorobenzene under photoirradiation.

from $^1SPA^*$ to C_{60} occurred in SPA/TPMA/C_{60} to produce the charge-separated (CS) state ($SPA^{\bullet+}$/TPMA/$C_{60}^{\bullet-}$).

The rate constant of photoinduced electron transfer from $^1SPA^*$ to C_{60} was determined from the fast decay of absorbance at 600 nm due to $^1SPA^*$ to be 7.3×10^{11} s^{-1} (Fig. 5a). The decay time profile of absorbance at 700 nm due to $SPA^{\bullet+}$ exhibited two components as shown in Fig. 5b, where the fast component afforded the rate constant of 7.3×10^{11} s^{-1}, which agrees with the rate constant of electron transfer from $^1SPA^*$ to C_{60}, and the slower component afforded the rate constant of 1.9×10^{10} s^{-1}. This value is larger than the decay rate constant of $C_{60}^{\bullet-}$ (5.7×10^9 s^{-1}) determined from the slower decay of absorbance at 1030 nm (Fig. 5c). The faster decay of $SPA^{\bullet+}$ than the decay of $C_{60}^{\bullet-}$ indicates that electron transfer from TPMA to $SPA^{\bullet+}$ occurred to produce the CS state ($SPA/TPMA^{\bullet+}/C_{60}^{\bullet-}$), which decayed with the rate constant of 5.7×10^9 s^{-1} and the CS lifetime of 180 ps. Electron transfer from TPMA to $SPA^{\bullet+}$ is indeed energetically feasible, because the one-electron oxidation potential of TPMA (1.20 V *vs*. SCE) is lower than that of SPA (1.40 V *vs*. SCE) as shown by their cyclic voltammograms (Fig. S5 in ESI†).

Similarly a transient absorption spectrum of the CS state of SPA/TPMA/C_{70} was observed by femtosecond laser flash photolysis measurements (Fig. S6 in ESI†). The rate constant of photoinduced electron transfer from $^1SPA^*$ to C_{70} was determined from the fast decay of absorbance at 600 nm due to $^1SPA^*$ to be 2.6×10^{11} s^{-1}. The CS lifetime of $SPA/TPMA^{\bullet+}/C_{70}^{\bullet-}$ was also determined from the slower decay of absorbance at 1030 nm due to $C_{70}^{\bullet-}$ to be 4.4×10^9 s^{-1}, which corresponds to the lifetime of 230

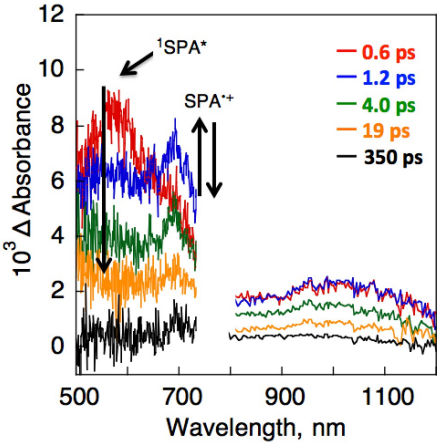

Fig. 4 Transient absorption spectra of SPA/TPMA/C_{60} dispersed in KBr pellet recorded at 0.6, 1.2, 4.0, 19 and 350 ps after femtosecond laser excitation at 393 nm.

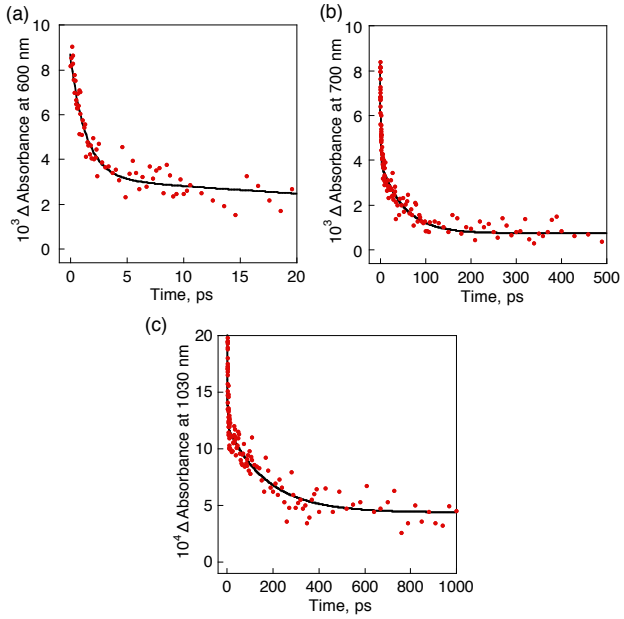

Fig. 5 Decay time profiles of absorption at (a) 600 nm, (b) 700 nm and (c) 1030 nm of SPA/TPMA/C_{60} in KBr pellet.

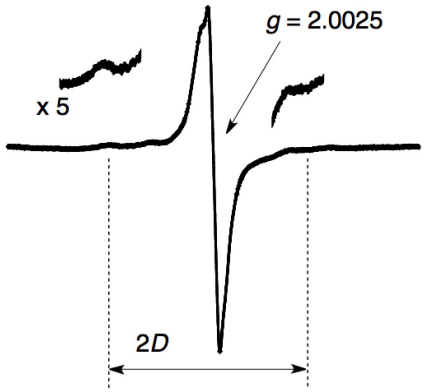

Fig. 6 ESR spectrum observed upon photoirradiation (λ > 340 nm) of crystals of SPA/TPMA/C_{60} at 77 K.

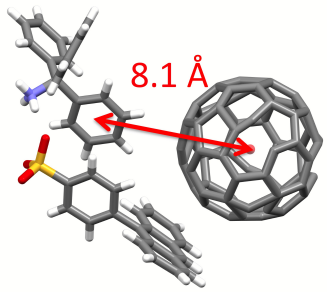

Fig. 7 The centre-to-centre distance between TPMA and C_{60} in the SPA/TPMA/C_{60} crystal.

ps.

Upon photoirradiation (λ > 340 nm) of single crystals of SPA/TPMA/C_{60} in an ESR cavity at 77 K, we could observe a signal at g = 2.0025 as shown in Fig. 6. This signal can be assigned to the superposition of $TPMA^{\bullet+}$ and $C_{60}^{\bullet-}$. In addition, we could observe a signal due to a triplet CS state, for which the zero-field splitting (D) value was determined to be 49 G. On the basis of this value, the distance between $TPMA^{\bullet+}$ and $C_{60}^{\bullet-}$ is estimated to be 8.3 Å,[24] which is in good agreement with that found in the crystal structure (8.1 Å in Fig. 7).

In conclusion, porous organic salt (POS) crystals composed of 9-(4-sulfophenyl)anthracene (SPA) and triphenylmethylamine (TPMA) were successfully impregnated with fullerenes (C_{60} and C_{70}) to reveal the X-ray crystal structures. Photoexcitation of POSs impregnated fullerenes resulted in efficient electron transfer from $^1SPA^*$ to fullerenes, followed by hole transfer from $SPA^{\bullet+}$ to TPMA to produce the final CS states ($SPA/TPMA^{\bullet+}/C_{60}^{\bullet-}$ and $SPA/TPMA^{\bullet+}/C_{70}^{\bullet-}$), which decayed with lifetimes of 180 and 230 ps, respectively. This work has demonstrated the excellent ability of POSs impregnated with fullerenes for efficient photoinduced electron transfer. The tunability and crystalline structures of POSs impregnated with various electron acceptors provide a promising platform to develop artificial photosynthetic systems.

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22 UV-vis and fluorescence spectra of SPA/TPMA in PhCN are shown in Fig. S4 in ESI.

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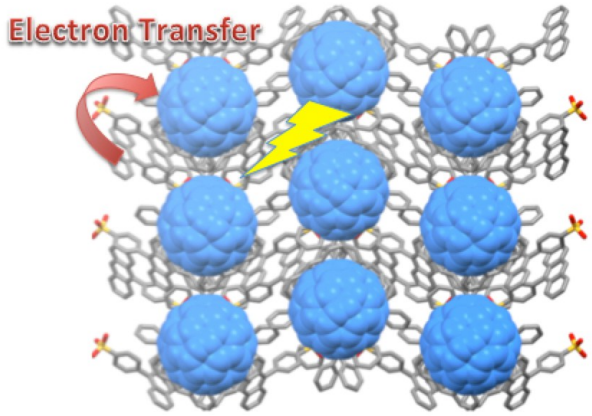
24 The distance (r, Å) between the two radical centres was estimated by the following equation: $r = [(2.78 \times 10^4)/D]^{1/3}$.

TOC Graphic



Porous organic salt (POS) crystals composed of 9-(4-sulfophenyl)anthracene and triphenylmethylamine were impregnated with fullerenes (C_{60} and C_{70}), resulting in fluorescence quenching by photoinduced electron transfer.