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ARTICLE

A zinc-responsive fluorophore based on 5'-(*p*-hydroxyphenyl)-pyridylthiazole

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The synthesis and photophysical properties of an ion-sensitive fluorescent compound are described, featuring 5'-(*p*-hydroxyphenyl)pyridylthiazole (HPPT) as a highly emissive fluorophore. Density functional theory (DFT) calculations indicate that HPPT is capable of intramolecular charge transfer (ICT), with further polarization upon complexation with Zn(II). A 4'-picolyloxy-HPPT derivative was prepared and determined to form a 1:1 complex with Zn(NO₃)₂ in acetonitrile, with a Stokes shift of 137 nm (6323 cm⁻¹) and a 67-nm bathochromic shift in emission relative to the neutral ligand, and a fluorescence quantum yield (Φ_{PL}) of 92%. An X-ray crystal structure of the HPPT-Zn(II) complex confirmed a tridentate structure with a seven-membered chelate ring, and the picolyloxy unit rotated out of plane.

Introduction

Chemically responsive fluorescent dyes are useful as probes for biological imaging and as components in the design of molecular logic gates.^{1,2,3,4,5} In many cases, molecular recognition events can be transduced by fluorescence using photoinduced electron transfer (PET) or intramolecular charge transfer (ICT) mechanisms.^{6,7} The latter has been broadly applied toward the design of ratiometric fluorescence sensing modalities that can quantify changes in the concentration of specific analytes.

ICT is commonly observed in π -conjugated systems with electron-rich and electron-poor (donor and acceptor) units at opposite ends of the molecule. If the excited-state dipole is proximal to an ion binding site, then ICT efficiency can be modulated in a supramolecular fashion. For example, acceptor units that can bind metal ions may experience an increased electric polarization upon chelation, causing a bathochromic shift in photoemission (Figure 1a). Ratiometric fluorescence sensing has been applied successfully toward the detection of transition-metal ions including Zn(II), an essential cofactor in many biological functions.^{8,9,10} The Zn(II) ion is compatible with excited-state processes, and can be incorporated into the design of fluorophores with large Stokes shifts.^{11,12,13,14}

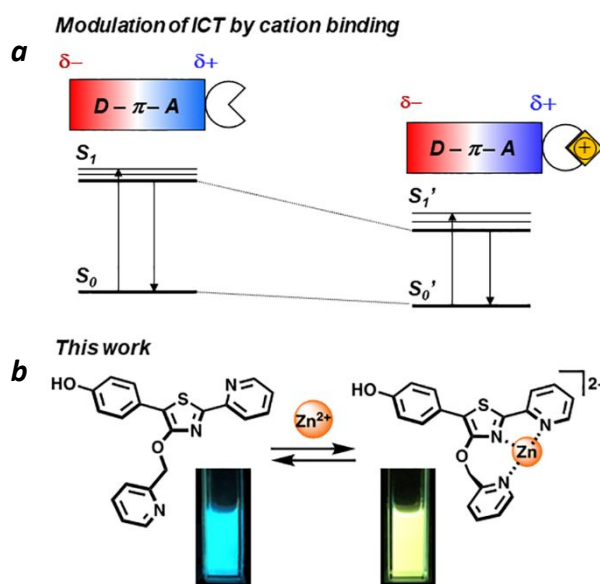


Figure 1. (a) Intramolecular charge transfer (ICT) within a π -conjugated donor–acceptor system, modulated by ion binding; (b) a Zn(II)-binding fluorophore using 5'-(*p*-hydroxyphenyl)pyridylthiazole (HPPT) as an ICT platform.

Several Zn(II)-binding fluorophores are based on well-established dye scaffolds such as BODIPY,^{1,15} xanthene,^{3,16} and coumarin,^{3b,11} which all have high quantum yields (Φ_{PL}) but produce small Stokes shifts upon metal ion binding. Fluorophores containing quinoline units have been shown to respond to Zn(II) with Stokes shifts of over 200 nm (8500 cm⁻¹), but are limited by their low Φ_{PL} .¹³ These earlier examples set the stage for developing new classes of ion-responsive fluorophores, featuring supramolecular binding motifs that have direct impact on their excited-state dipoles.

We have identified 5'-phenyl-2'-(2-pyridyl)thiazole (PPT) as a promising scaffold for the rational design of ratiometric

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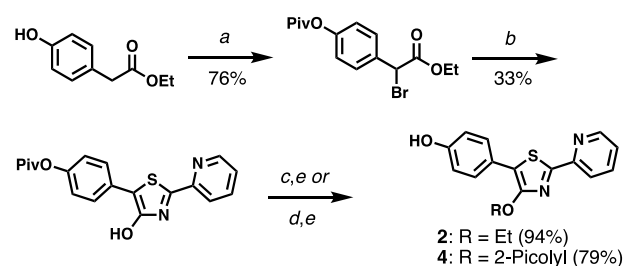
fluorescent probes, using the ICT approach (**Figure 1b**). The bidentate cleft of the pyridylthiazole unit can be fashioned into a receptor for transition-metal ions, providing a useful mechanism for lowering the energy of its acceptor state.^{17,18} Earlier studies on the chemistry and photophysical properties of 2'-(2-pyridyl)thiazoles by Yan¹⁹ and Beckert²⁰ have shown these compounds to be capable of supporting high Φ_{PL} and large Stokes shift due to structural changes in their excited states, with considerable tolerance for chemical modification. Their studies provide an excellent starting point for designing PPT derivatives whose fluorescence outputs can be tuned by well-defined supramolecular interactions. In this paper, we introduce 5'-(*p*-hydroxyphenyl)pyridylthiazole (HPPT) as a highly emissive ICT-type fluorophore, and a Zn(II)-binding HPPT derivative that exhibits a 67-nm bathochromic shift in emission upon complexation, with 92% Φ_{PL} and a Stokes shift of 137 nm (6323 cm⁻¹).

Results and discussion

DFT calculations. To determine the feasibility of generating excited-state ICT within the PPT scaffold, we first conducted density functional theory (DFT) calculations to evaluate the effects of *p*-substitution on the phenyl ring of the PPT unit, with and without Zn(II) complexation.²¹ HPPT was selected as an ICT-type fluorophore due to the strong polarization of its molecular orbital (MO) states by the *p*-hydroxyl group, relative to that of unsubstituted PPT (**Figure 2** and ESI). Coordination between the pyridylthiazole unit and Zn(NO₃)₂ further increased polarization of the LUMO coefficients as expected, with the decrease in energy band gap (ΔE_{g}) being most pronounced for HPPT. We also performed DFT calculations on

pyridylthiazole derivatives bearing an picolyloxy group (4'-*O*-Pic-PPT **3** and 4'-*O*-Pic-HPPT **4**), a well-known ancillary ligand for stabilizing Zn(II) binding, and confirmed that its presence had a negligible effect on the spatial distribution of MO coefficients (see ESI).

Synthesis. PPT derivatives with a 4'-hydroxyl on the central thiazole ring were prepared using Hantzsch conditions,²² similar to those described by Beckert and coworkers (**Scheme 1**).^{23,24} In the case of HPPT, ethyl *p*-hydroxyphenylacetate was converted to a pivaloyl ester then brominated at the methylene position, and condensed with 2-pyridylthioamide by microwave heating to yield the corresponding 4'-hydroxythiazole in 25% yield over three steps. *O*-Alkylation with ethyl iodide or picolyl bromide followed by methanolysis produced 4'-*O*-Et-HPPT (**2**) or 4'-*O*-Pic-HPPT (**4**) respectively. PPT derivatives lacking the *p*-hydroxyl group, namely 4'-*O*-Et-PPT (**1**) and 4'-*O*-Pic-PPT (**3**), were prepared according to literature procedure and alkylated as described above.²⁴



Scheme 1. Synthesis of HPPT derivatives **2** and **4**. (a) (i) PivCl, DMAP, pyridine; (ii) NBS, AIBN, TBT, 80 °C. (b) 2-pyridylthioamide, EtOH, 90 °C (μW). (c) EtI, K₂CO₃, acetone. (d) 2-PicBr-HBr, K₂CO₃, acetone. (e) K₂CO₃, MeOH. See ESI for reagent abbreviations.

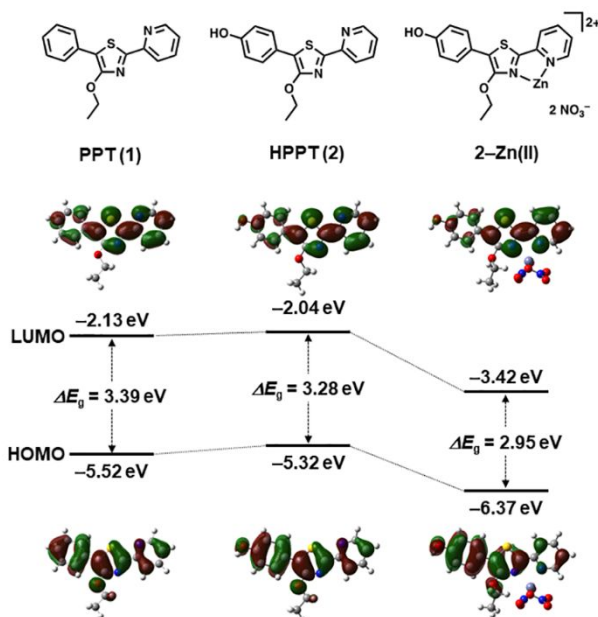


Figure 2. DFT calculations of energy bands and spatial distributions of molecular orbital coefficients (HOMO and LUMO) for 4'-*O*-Et-PPT (**1**), 4'-*O*-Et-HPPT (**2**), and **2-Zn(II)**. See ESI for additional DFT calculations.

X-ray crystal structure determination. The crystal structure of compound **4** revealed the π -conjugated chromophore to be nearly planar (phenyl–thiazolyl dihedral angle = 12.15°; pyridyl–thiazolyl dihedral angle = 7.33°), with the nitrogen atoms adopting a transoid geometry to minimize steric interactions between rings. In contrast, the 4'-*O*-picolyl group is nearly perpendicular with respect to the thiazole ring (**Figure 3a**). The crystal structure of complex **4-Zn(II)**, which was prepared by mixing Zn(NO₃)₂ with **4** in equimolar ratio, confirmed the tridentate nature of Zn coordination by the 4'-*O*-picolyl unit and the two pyridylthiazole nitrogens in cisoid geometry (**Figure 3b**). The bite angles for N1–Zn–N2 and N2–Zn–N3 are 77.8° and 100.2°, respectively. Notably, there are no significant differences in the distances between Zn–N bonds (Zn–N1 bond length = 2.121 Å; Zn–N2 = 2.125 Å; Zn–N3 = 2.110 Å), despite the fact that the pyridine ring in the picolyloxy arm is rotated out of plane by 26.4° relative to the thiazole. The seven-membered chelate ring is uncommon,²⁵ but may be a useful binding motif in the design of other coordination complexes.

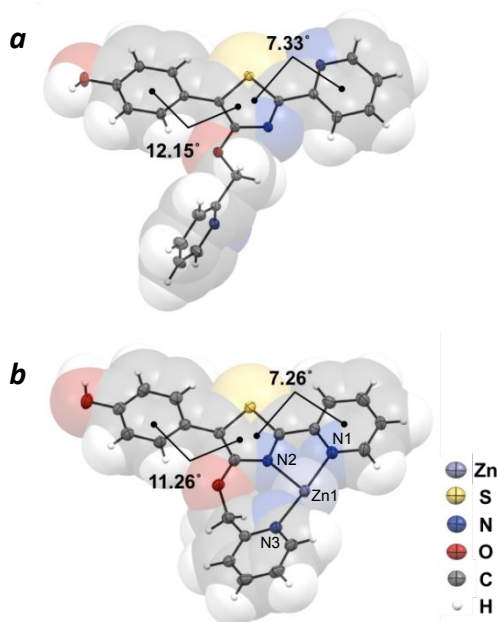


Figure 3. X-ray crystal structures of (a) 4'-O-Pic-HPPT (**4**) and (b) 4-Zn(II). Nitrate counterions are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

Photophysical properties. Absorbance and photoluminescence (PL) spectra of 4'-O-Pic-PPT **3** and 4'-O-Pic-HPPT **4** were acquired in solvents of varying polarity to determine the extent of electric polarization within HPPT (**Figure 4**). HPPT **4** has a strong absorption band near 380 nm ($\epsilon_M = 2 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$, CH_3CN) corresponding with its π - π^* transition, and is redshifted by 11 nm relative to that of PPT **3** ($\lambda_{\text{max}} = 369 \text{ nm}$). **4** also exhibits solvatochromism in its PL spectra, with a shift in peak emission (λ_{em}) from 461 nm in toluene (ϵ_d 2.38) to 487 nm in ethanol (ϵ_d 24.5), indicating increased stabilization of its excited-state dipole by a polar environment (see ESI).²⁶ Similar phenomena were observed with 4'-O-ethyl derivative **2**, validating HPPT as a donor-acceptor ICT system (**Table 1** and ESI). The Φ_{PL} for all PPT derivatives in acetonitrile are well over 90%, and correlate well with previous studies.^{20,23} Optical band gaps (E_g) were estimated from the absorption edge to be 3.0 eV for **1** and **3** and 2.9 eV for **2** and **4**, indicating minimal perturbation of the π -conjugated system by the 4'-O-alkyl unit.

Coordination of Zn(II) to 4'-O-Pic-HPPT **4** led to striking changes in emission colour with retention of high Φ_{PL} (**Table 1**). The λ_{em} of **4** shifted by 67 nm toward the red upon Zn(II) binding, from 472 to 539 nm. A 22-nm redshift in absorbance ($\Delta\lambda_{\text{max}}$) from 380 to 402 nm was also observed, with well-defined isosbestic points at 305 and 393 nm (**Figure 5a**). Stoichiometric titration of $\text{Zn}(\text{NO}_3)_2$ into a CH_3CN solution of **4** produced linear changes in both absorption and emission peak intensities, with no further change after one molar equivalent (**Figure 5b**). The association constant (K_a) for Zn^{2+} binding by **4** was determined to be $9.6 \times 10^7 \text{ M}^{-1}$ (see ESI), higher than that reported for some similar compounds in CH_3CN .¹⁸ In comparison, coordination of Zn(II) to 4'-O-Pic-PPT **3** induced much smaller redshifts in emission ($\Delta\lambda_{\text{em}} = 22 \text{ nm}$) and absorbance ($\Delta\lambda_{\text{max}} = 14 \text{ nm}$), as well as a slight loss in Φ_{PL} and a lower K_a value (see ESI).

No redshifts were observed for 4'-O-ethyl derivatives **1** or **2**, meaning that the pyridyl nitrogen in the 4'-O-picolyl group was necessary for Zn(II) binding, and could adopt an appropriate geometry for stable complex formation. Lastly, exchanging counterions from nitrate to perchlorate or triflate had little effect on the photophysical properties of **4** (see ESI).

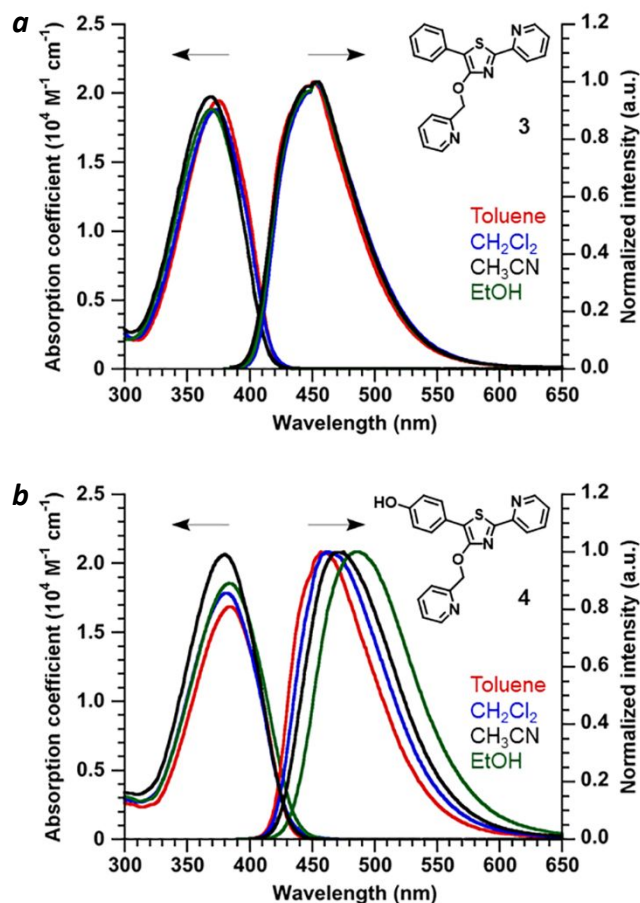


Figure 4. (a) Optical absorption and normalized PL spectra of 4'-O-Pic-PPT **3** and (b) 4'-O-Pic-HPPT **4**. Spectra recorded at 30 μM and 20 $^\circ\text{C}$.

Table 1. Photophysical properties of PPT derivatives (in CH_3CN)

Compound	Abs (λ_{max}) (nm)	$\epsilon_M \times 10^4$ ($\text{M}^{-1} \text{ cm}^{-1}$)	λ_{em} (nm) ^a	ν_{St} (cm^{-1}/nm)	Φ_{PL} (%) ^b
1 (4'-O-Et-PPT)	371	1.90	454	4928/ 83	96
2 (4'-O-Et-HPPT)	383	1.77	476	5101/ 93	96
3 (4'-O-Pic-PPT)	369	1.97	454	5074/ 85	94
4 (4'-O-Pic-HPPT)	380	2.07	472	5129/ 92	92
3-Zn(II)	383	1.82	476	5101/ 93	87
4-Zn(II)	402	1.85	539	6323/ 137	92

^a $\lambda_{\text{ex}} = 370 \text{ nm}$ for **1** and **3**, 380 nm for **2** and **4**, 383 nm for **3-Zn(II)**, and 402 nm for **4-Zn(II)**. ^b Φ_{PL} determined using an integrating sphere.

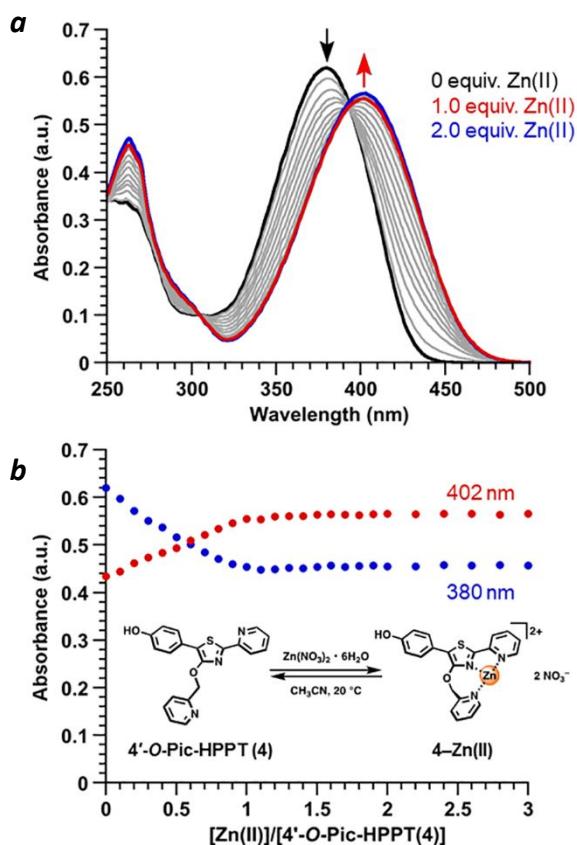


Figure 5. (a) Spectral changes in the absorbance of **4** (30 μM) upon titration of $\text{Zn}(\text{NO}_3)_2 \cdot 6 \text{H}_2\text{O}$ (3 mM in CH_3CN). (b) Stoichiometry of **4-Zn(II)** complex formation in CH_3CN using the molar ratio method, based on absorbance maxima.

With regard to visual perception, HPPT derivatives produce photons that are easily detected by the human eye. The photoluminescence of **4** ($\lambda_{\text{ex}}/\lambda_{\text{em}} = 393/472 \text{ nm}$) corresponds with chromaticity coordinates (0.15,0.22) in Commission Internationale de L'Eclairage (CIE) 1931 colour space (**Figure 6**).²⁷ Formation of the **4-Zn(II)** complex results in a peak emission at 539 nm, which shifts the CIE coordinates to (0.34,0.56) within the green region. To translate these CIE coordinates into an optical readout of Zn(II) binding, the ratio of emission intensities at 539 and 472 nm (F_{539}/F_{472}) increases by x-ray crystallography and by the absence of a Zn-induced shift in emission for 4'-O-ethyl HPPT **2** in solution. The *p*-hydroxyphenyl unit in Pic-HPPT **4** is responsible for a 67-nm shift in emission upon Zn(II) binding, compared with a 22-nm shift in emission for PPT **3**. The **4-Zn(II)** complex has a relatively large Stokes shift of 137 nm or 6323 cm^{-1} , which bodes well for optical sensing and imaging applications. Investigations on the metal-ion binding properties of HPPT derivatives in aqueous solutions are in progress, and will be reported in due course.

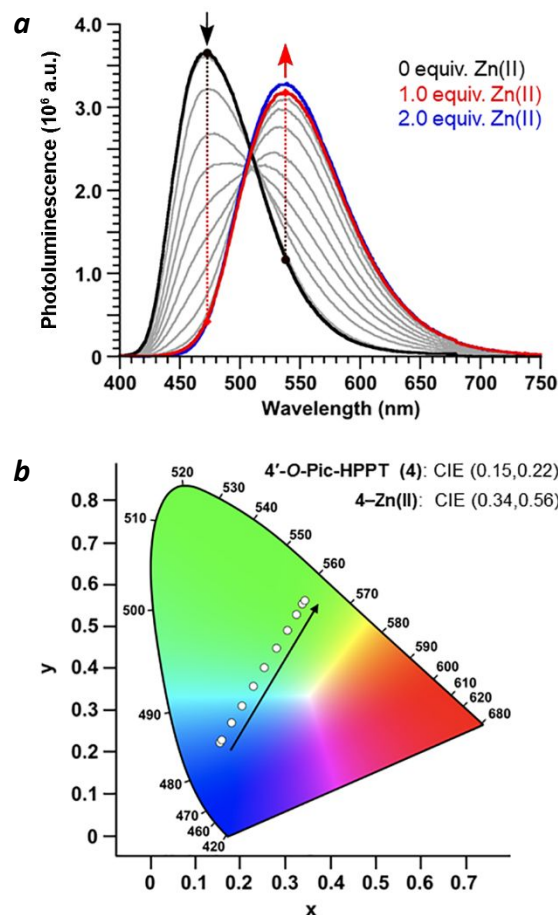


Figure 6. (a) Spectral changes in the photoluminescence of **4** (30 μM in CH_3CN) upon titration of $\text{Zn}(\text{NO}_3)_2 \cdot 6 \text{H}_2\text{O}$ (3 mM in CH_3CN) and excitation at the isosbestic point ($\lambda_{\text{ex}} = 393 \text{ nm}$). (b) CIE chromaticity diagram of **4** showing the progression in emitted light with $\text{Zn}(\text{NO}_3)_2$ addition.

Conclusions

5'-(*p*-Hydroxyphenyl)pyridylthiazole is a new ICT-type fluorophore that can be configured for a strongly emissive response to Zn(II) ions. The HPPT platform is easy to synthesize and amenable to derivatization, particularly at the 4' position. The 4'-O-picolyl arm stabilizes the coordination of Zn(II), as shown by x-ray crystallography and by the absence of a Zn-induced shift in emission for 4'-O-ethyl HPPT **2** in solution. The *p*-hydroxyphenyl unit in Pic-HPPT **4** is responsible for a 67-nm shift in emission upon Zn(II) binding, compared with a 22-nm shift in emission for PPT **3**. The **4-Zn(II)** complex has a relatively large Stokes shift of 137 nm or 6323 cm^{-1} , which bodes well for optical sensing and imaging applications. Investigations on the metal-ion binding properties of HPPT derivatives in aqueous solutions are in progress, and will be reported in due course.

Experimental

Chemical reagents and supplies obtained from commercial vendors were used as supplied unless otherwise noted. *N*-bromosuccinimide (NBS) was recrystallized from water as a white

crystalline solid. Microwave reactions were performed using a CEM Discovery microwave reactor. ^1H and ^{13}C NMR spectra were collected at room temperature on a Varian INOVA 300, Bruker AV400 HD, or AV800 HD spectrometer. NMR samples were prepared in chloroform-*d*, acetone-*d*₆, or dimethyl sulfoxide-*d*₆, with chemical shifts (δ) reported in ppm relative to residual solvent resonances as internal standards. Electrospray ionization mass spectra were obtained using an Advion Expression compact spectrometer. Full details on the synthesis and characterization of **1–4** and their intermeidates are described in the ESI.

Absorption spectra were acquired on a Varian Cary50 Bio UV-Vis spectrophotometer using a 1-cm quartz cuvette. Photoluminescence spectra and absolute quantum yields (Φ_{PL}) were measured using an Edinburgh Instruments FLS 980 spectrometer with an integrating sphere accessory. Chromaticity (CIE) coordinates were analysed with F980 software. Quantum chemical calculations were performed using the hybrid density functional theory (DFT) functional (Gaussian 16).²¹ Geometries of PPT derivatives were optimized at the B3LYP level using the 6-31+G(d,p) basis set for the ground state. Geometries of PPT–Zn(II) complexes were optimized at the B3LYP/Gen level using the 6-31+G(d,p) basis set for H, C, N atoms and LANL2DZ basis set for the Zn(II) atom in the gas phase. Crystallographic data reported in this paper are tabulated in the supplementary materials and archived at the Cambridge Crystallographic Data Centre under reference numbers CCDC 1960137 and 1960138.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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¹ (a) A. Loudet, K. Burgess, BODIPY Dyes and Their Derivatives: Syntheses and Spectroscopic Properties. *Chem. Rev.* 2007, **107**, 4891; (b) J. Han, K. Burgess, Fluorescent Indicators for Intracellular pH. *Chem. Rev.* 2010, **110**, 2709.

² H. Kobayashi, M. Ogawa, R. Alford, P. L. Choyke, Y. Urano, New Strategies for Fluorescent Probe Design in Medical Diagnostic Imaging. *Chem. Rev.*, 2010, **110**, 2620.

³ (a) X. Chen, T. Pradhan, F. Wang, J. S. Kim, J. Yoon, Fluorescent Chemosensors Based on Spiroring-Opening of Xanthenes and Related Derivatives. *Chem. Rev.* 2012, **112**, 1910; D. Cao, Z. Liu, P. Verwilst, S. Koo, P. Jangjili, J. S. Kim, W. Lin, Coumarin-Based Small-Molecule Fluorescent Chemosensors. *Chem. Rev.* 2019, **119**, 10403.

⁴ J. Chan, S. C. Dodani, C. J. Chang, Reaction-based small-molecule fluorescent probes for chemoselective bioimaging. *Nat. Chem.* 2012, **4**, 973.

⁵ X. Li, X. Gao, W. Shi, H. Ma, Design Strategies for Water-Soluble Small Molecular Chromogenic and Fluorogenic Probes. *Chem. Rev.*, 2014, **114**, 590.

⁶ (a) R. A. Bissell, A. P. de Silva, H. Q. N. Gunaratne, P. L. M. Lynch, G. E. M. Maguire, K. R. A. S. Sandanayake, Molecular fluorescent signalling with 'fluor-spacer-receptor' systems: approaches to sensing and switching devices via supramolecular photophysics. *Chem. Soc. Rev.*, 1992, **21**, 187. (b) A. P. de Silva, H. Q. N. Gunaratne, T. Gunnlaugsson, A. J. M. Huxley, C. P. McCoy, J. T. Rademacher, T. E. Rice, Signaling Recognition Events with Fluorescent Sensors and Switches. *Chem. Rev.*, 1997, **97**, 1515.

⁷ B. Valeur, I. Leray, Design principles of fluorescent molecular sensors for cation recognition. *Coord. Chem. Rev.*, 2000, **205**, 3.

⁸ Z. Xu, J. Yoon, D. R. Spring, Fluorescent chemosensors for Zn²⁺. *Chem. Soc. Rev.*, 2010, **39**, 1996.

⁹ (a) E. M. Nolan, S. J. Lippard, Small-Molecule Fluorescent Sensors for Investigating Zinc Metalloneurochemistry. *Acc. Chem. Res.* 2009, **42**, 193; (b) E. Tomat, S. J. Lippard, Imaging mobile zinc in biology. *Curr. Opin. Chem. Biol.*, 2010, **14**, 225.

¹⁰ K. P. Carter, A. M. Young, A. E. Palmer, Fluorescent Sensors for Measuring Metal Ions in Living Systems. *Chem. Rev.* 2014, **114**, 4564.

¹¹ N. C. Lim, J. V. Schuster, M. C. Porto, M. A. Tanudra, L. Yao, H. C. Freake, C. Brückner, Coumarin-Based Chemosensors for Zinc(II): Toward the Determination of the Design Algorithm for CHEF-Type and Ratiometric Probes. *Inorg. Chem.* 2005, **44**, 2018.

¹² K. Komatsu, Y. Urano, H. Kojima, T. Nagano, Development of an Iminocoumarin-Based Zinc Sensor Suitable for Ratiometric Fluorescence Imaging of Neuronal Zinc. *J. Am. Chem. Soc.* 2007, **129**, 13447.

¹³ (a) L. Xue, H.-H. Wang, X.-J. Wang, H. Jiang, Modulating Affinities of Di-2-picolylamine (DPA)-Substituted Quinoline Sensors for Zinc Ions by Varying Pendant Ligands. *Inorg. Chem.*, 2008, **47**, 4310; (b)

- L. Xue, C. Liu, H. Jiang, A ratiometric fluorescent sensor with a large Stokes shift for imaging zinc ions in living cells. *Chem. Commun.*, 2009, 1061.
- ¹⁴ D. Bourassa, C. M. Elitt, A. M. McCallum, S. Sumalekshmy, R. L. McRae, M. T. Morgan, N. Siegel, J. W. Perry, P. A. Rosenberg, C. J. Fahrni, Chromis-1, a Ratiometric Fluorescent Probe Optimized for Two-Photon Microscopy Reveals Dynamic Changes in Labile Zn(II) in Differentiating Oligodendrocytes. *ACS Sensors*, 2018, **3**, 458.
- ¹⁵ Y. Wu, X. Peng, B. Guo, J. Fan, Z. Zhang, J. Wang, A. Cui, Y. Gao, Boron dipyrromethene fluorophore based fluorescence sensor for the selective imaging of Zn(II) in living cells. *Org. Biomol. Chem.* 2005, **3**, 1387.
- ¹⁶ S. C. Burdette, G. K. Walkup, B. Spingler, R. Y. Tsien, S. J. Lippard, Fluorescent Sensors for Zn²⁺ Based on a Fluorescein Platform: Synthesis, Properties and Intracellular Distribution. *J. Am. Chem. Soc.* 2001, **123**, 7831.
- ¹⁷ (a) J. C. Jeffery, C. R. Rice, L. P. Harding, C. J. Baylies, T. Riis-Johannessen, Ligand Reprogramming in Dinuclear Helicate Complexes: A Consequence of Allosteric or Electrostatic Effects? *Chem. Eur. J.*, 2007, **13**, 5256; (b) S. Bullock, L. J. Gillie, L. P. Harding, C. R. Rice, T. Riis-Johannessen, M. Whitehead, Isomeric pyridylthiazole donor units for metal ion recognition in bi- and tri-metallic helicates. *Chem. Commun.*, 2009, 4856.
- ¹⁸ M. H. Zheng, X. Hu, M. Y. Yang, J. Y. Jin, Ratiometrically Fluorescent Sensing of Zn(II) Based on Dual-Emission of 2-Pyridylthiazole Derivatives. *J. Fluor.*, 2015, **25**, 1831.
- ¹⁹ M.-H. Zheng, J.-Y. Jin, W. Sun, C.-H. Yan, A new series of fluorescent 5-methoxy-2-pyridylthiazoles with a pH-sensitive dual-emission. *New J. Chem.*, 2006, **30**, 1192.
- ²⁰ U.-W. Grummt, D. Weiss, E. Birckner, R. Beckert, Pyridylthiazoles: Highly Luminescent Heterocyclic Compounds. *J. Phys. Chem. A*, 2007, **111**, 1104.
- ²¹ Gaussian 16, Revision A.03, M. J. Frisch et al, Gaussian, Inc., Wallingford CT, 2016.
- ²² A. Hantzsch, Untersuchungen über Azole. *Justus Liebigs Ann. Chem.*, 1889, **250**, 257.
- ²³ (a) E. Täuscher, D. Weiß, R. Beckert and H. Görls, Synthesis and Characterization of New 4-Hydroxy-1,3-thiazoles. *Synthesis*, 2010, **10**, 1603; (b) S. Wolfram, H. Würfel, S. H. Habenicht, C. Lembke, P. Richter, E. Birckner, R. Beckert, G. Pohnert, A small azide-modified thiazole-based reporter molecule for fluorescence and mass spectrometric detection. *Beilstein J. Org. Chem.*, 2014, **10**, 2470.
- ²⁴ M. Kaufmann, M. L. Hupfer, T. Sachse, F. Herrmann-Westendorf, D. Weiß, B. Dietzek, R. Beckert, M. Presselt, Introducing double polar heads to highly fluorescent Thiazoles: Influence on supramolecular structures and photonic properties. *J. Colloid Interface Sci.*, 2018, **526**, 410.
- ²⁵ D. Aguila, E. Escribano, S. Speed, D. Talancon, L. Yerman, S. Alvarez, Calibrating the coordination chemistry tool chest: metrics of bi- and tridentate ligands. *Dalton Trans.*, 2009, 6610.
- ²⁶ P. Suppan and N. Ghoneim, *Solvatochromism*, Royal Society of Chemistry, Cambridge, 1997.
- ²⁷ T. Smith and J. Guild, The C.I.E. colorimetric standards and their use. *Trans. Opt. Soc.*, 1931, **33**, 73-134.