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The high persistence of PFAS is sufficient for their management as a chemical class

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Per- and polyfluoroalkyl substances (PFAS) are a class of synthetic organic substances with diverse structures, properties, uses, bioaccumulation potentials and toxicities. Despite this high diversity, all PFAS are alike in that they contain perfluoroalkyl moieties that are extremely resistant to environmental and metabolic degradation. The vast majority of PFAS are therefore either non-degradable or transform ultimately into stable terminal transformation products (which are still PFAS). Under the European chemicals regulation this classifies PFAS as very persistent substances (vP). We argue that this high persistence is sufficient concern for their management as a chemical class, and for all “non-essential” uses of PFAS to be phased out. The continual release of highly persistent PFAS will result in increasing concentrations and increasing probabilities of the occurrence of known and unknown effects. Once adverse effects are identified, the exposure and associated effects will not be easily reversible. Reversing PFAS contamination will be technically challenging, energy intensive, and costly for society, as is evident in the efforts to remove PFAS from contaminated land and drinking water sources.

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Environmental significance

Recent scientific publications and regulatory actions suggest that PFAS should be managed as a class. This Perspective article supports a broad scope in restricting the use of PFAS in society based on the extremely high environmental persistence of all PFAS. Our previous work (cited below) has highlighted the particular concern of high environmental persistence and noted that high persistence is the root cause of most of the serious cases of environmental contamination in the last 50 years. The field of green chemistry also recognizes the particular issue with persistence and recommends that chemical manufacturers incorporate “design for degradation”. We are of the opinion that highly persistent PFAS are incompatible with green chemistry principles and future visions of sustainable development. Phasing out “non-essential” uses of PFAS from society and replacing them with suitable alternatives will benefit environmental and human health, now and into the future. Unfortunately, as PFAS are non-mineralizable, a reduction of environmental exposure to PFAS will be a slow process even if their uses are phased out.

Perspective article

The national authorities of five European states have agreed to prepare a joint REACH restriction proposal to limit the risks to

human and environmental health associated with per- and polyfluoroalkyl substances (PFAS).¹ The exact scope of the restriction proposal for PFAS will be determined during the development phase of the project, but it will initially aim for a broad inclusion of substances (the aim is to include “substances that contain at least one aliphatic –CF₂– or –CF₃ element”)² and uses to be restricted. A recent commentary by Kwiatkowski *et al.*³ outlined the scientific basis for managing PFAS as a chemical class. This was an important contribution towards efforts to regulate^{4,5} and group^{6,7} PFAS. In this Perspective article, we would like to voice our support for a broad scope in the restriction proposal for PFAS, based largely on concerns regarding their high persistence. A precedent for this exists in the recent action by the European Union (EU) to restrict intentionally added microplastics based largely on concerns regarding their high persistence.⁸

PFAS are a class of fluorinated synthetic organic substances with diverse properties that have been used in a wide variety of industrial and consumer applications.^{6,9} The Organisation for

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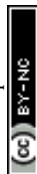
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Economic Co-operation and Development (OECD)/United Nations Environment Programme (UNEP) compiled a list of PFAS with over 4700 CAS numbers that may have been on the global market.¹⁰ More recently, the United States Environmental Protection Agency (US EPA)¹¹ has compiled a list of more than 8000 PFAS entries.^{12,13} As noted in the Supplement to the recent paper by Johnson *et al.*,¹⁴ the PFAS class comprises distinct substances with very different structures and properties: high-molecular-weight polymers and high-molecular-weight non-polymers; neutral, anionic, cationic and zwitterionic substances; solids, liquids, and gases; highly reactive and non-reactive (inert) substances; soluble and insoluble substances; and volatile and involatile substances. In the environment, some PFAS are mobile¹⁵ and others immobile, and some bioaccumulate while others do not.

PFAS display a wide range of environmental and human health hazards. Some PFAS are considered of low health concern (*e.g.*, certain fluoropolymers,¹⁶ perfluoroalkanes¹⁷ and perfluoroalkyl-*tert*-amines^{3,17}), whereas others (*e.g.* long-chain perfluoroalkyl carboxylic (PFCAs) and perfluoroalkane sulfonic (PFSAs) acids) are associated with adverse health effects in humans and wildlife at current environmental exposure levels.¹⁸ Other PFAS that are structurally similar to PFCAs and PFSAs (*e.g.* the perfluoroalkylether carboxylic (PFECAs) and sulfonic acids (PFESAs), perfluoroalkyl dicarboxylic acids (PFdiCAs), perfluoroalkyl phosphonic (PFPAAs) and phosphinic acids (PFPIAs) and perfluoroalkane sulfinic acids (PFSiAs)) with similarly long perfluoroalkyl(ether) chains to long-chain PFCAs and PFSAs are expected to be of equivalent concern.^{7,19} Even if some PFAS are considered of low health concern, there may be starting materials, breakdown products and/or other PFAS by-products of higher concern released during their lifecycle (*e.g.* in the case of certain fluoropolymers²⁰) or they may be of high climate/environmental concern (*e.g.* in the case of perfluoroalkanes²¹ and perfluoro-*tert*-amines²²).

It must be noted that our current understanding of PFAS toxicity and other hazards is based on a relatively small number of compounds, and little is known about the properties and behaviour of most of the thousands of PFAS as individual chemicals or as the much more commonly present complex mixtures. It is likely that additional environmental and human health concerns will be identified as research continues into this class of chemicals.¹⁹ New findings will undoubtedly be made in toxicology that will change the current hazard and risk perspective of certain PFAS. For example, over the last two decades advances in the understanding of the human toxicology of perfluorooctanoic acid (PFOA) has caused US state level drinking water guidelines to be steadily lowered (from 100 000 ng L⁻¹ in West Virginia in 2002 down to 10 ng L⁻¹ in multiple states in 2019).²³

The diversity in PFAS structures, properties and behaviour is often used by the PFAS manufacturing industry^{24,25} to argue that PFAS cannot be treated as a single class. We acknowledge the diversity of PFAS in terms of properties, behaviour, hazards and risks, and that statements such as “all PFAS are bioaccumulative and toxic” are overgeneralized and debatable. However, despite their diversity, PFAS do share one common

structural feature that makes them highly problematic, namely the presence of perfluoroalkyl moieties, resulting in their shared resistance to environmental and metabolic degradation.¹⁹ The vast majority of PFAS are either non-degradable or degrade to form terminal products which are still PFAS (*e.g.* *n*:2 fluorotelomer alcohols – C_{*n*}F_{2*n*+1}CH₂CH₂OH – degrade ultimately to perfluoroalkyl carboxylic acids – C_{*n*}F_{2*n*+1}COOH).¹⁹ The extreme resistance to degradation of the perfluoroalkyl moiety is due to: (i) the strength of the C–F bond, which is the strongest single bond in organic chemistry,²⁶ (ii) multiple C–F bonds (“perfluorination”) on the same geminal carbon, leading to the shortening (*i.e.*, increased strength) of the C–F bond,²⁶ (iii) the strength of the C–F bond and high electronegativity of fluorine prevents nucleophilic attack²⁶ and (iv) the strong electron withdrawing effect of the fluorines in perfluoroalkyl moieties also strengthens the skeletal bonds in the carbon chain.²⁷ For example, the C–C bond in hexafluoroethane is 29 kJ mol⁻¹ stronger than that in ethane.²⁸ The high stability of perfluoroalkyl moieties is the reason that many of the applications of PFAS are for uses in high temperature and/or corrosive environments.²⁷

The phrasing “the vast majority of PFAS can be classified as highly persistent” was used in the preceding paragraph because there are a few specific structural combinations that may result in specific groups of PFAS that are mineralizable. Merck, for example, have published a patent²⁹ for structural combinations of novel fluorosurfactants that they believe may lead to the development of biodegradable commercial products. In this patent,²⁹ multiple structural combinations are included that contain per- and polyfluoroalkyl moieties connected to a heteroatom, namely O, S or N (*i.e.*, end groups such as CF₃O, (CF₃)₂N, CF₃S, and many more). To date, only the novel fluorosurfactant 10-(trifluoromethoxy)decane-1-sulfonate, which is a PFAS, has been shown to be mineralizable.³⁰ Furthermore, independent of Merck’s efforts, a novel fluorosurfactant, 3-hydroxy-2-(trifluoromethyl)propanoic acid,³¹ developed for the emulsion polymerization of vinylidene fluoride (VDF) has also been reported to be easily degradable. It is emphasized that these structural combinations that result in non-persistent PFAS are rare.

Environmental persistence is often expressed by half-lives of a chemical in various environmental media based on the kinetics of dominant abiotic and biotic degradation processes.³² Although the concept of environmental persistence is simple to understand, its assessment for a given chemical in a given environmental medium can sometimes be a challenging task because of the difficulty in measuring and estimating environmental degradation half-lives for chemicals.³³ But this challenge does not exist in the case of PFAS because scientists have reached consensus that the vast majority of PFAS are resistant to complete mineralization under natural conditions.¹⁹ Indeed, to our knowledge, PFAS are the most environmentally persistent substances among organic chemicals, and have thus earned the moniker in the popular press of “forever chemicals”.³⁴ Therefore, under the REACH definition of persistence, which includes persistent transformation products, the vast majority of PFAS are very persistent (vP).³⁵ Unfortunately, not all jurisdictions



include persistent transformation products in their definition of persistence, and this has given the manufacturing industry some basis to claim that not all PFAS are persistent by neglecting their transformation products.

In a recent paper by Cousins *et al.*,³⁶ the concerns regarding the high persistence of chemicals were explained. In summary, these concerns are:

- The continuous release of highly persistent chemicals will necessarily lead to widespread, long-lasting, and increasing contamination.
- Increasing concentrations will result in increased exposures and therefore increased probabilities for known and unknown health effects, be it by individual PFAS and/or in a mixture with other substances.
- Once adverse effects are identified, it will be technically challenging, energy intensive, and thus costly, to reverse the chemical contamination and therefore effects. Even with extensive efforts, removal will be only partially effective, and some level of contamination and some adverse effects will be irreversible.

It was argued by Cousins *et al.*³⁶ that high persistence alone should be established as a sufficient basis for management of a chemical, which was termed the “P-sufficient approach”. Logic dictates that evidence of bioaccumulation or some type of effect (biological or otherwise) is not required, given that the occurrence of an effect is of increasing probability and ultimately inevitable if emissions continue. Using current technology, it is not feasible to develop full hazard profiles for all PFAS. Managing these compounds as a highly persistent class is a clear benefit to society because otherwise a large number of resource-intensive and time-consuming toxicity and other tests would need to be conducted. On the basis of plausible reasoning, for continually released, highly persistent organic chemicals, a harmful effect (known or unknown) is likely in the future based on ever expanding exposures. This view may be controversial for some. For example, major regulatory bodies such as the US EPA are instead following the traditional risk assessment paradigm of testing as many PFAS as possible for adverse effects. Because traditional animal testing is costly and considered unethical, the US EPA is instead conducting a large number of high-throughput toxicity tests (*in vitro* assays) on 150 novel and emerging PFAS for multiple endpoints.³⁷ Such approaches are, however, uncertain and not comprehensive in their coverage of chemical or biological space and are not, therefore, reliably protective of human and environmental health. A preventative approach of not using highly persistent synthetic organic substances is more protective and also overall less costly for society, both in terms of fewer tests and reduction in externalized societal costs including the expected costs of health care, loss of biodiversity, loss of ecosystem services, loss of property value and remediation.

The arguments for not using highly persistent organic chemicals in society are not new and have been made for nearly 50 years, early on by Jensen in 1972³⁸ and Stephenson in 1977³⁹ and reformulated later in concepts such as “late lessons from early warnings”,⁴⁰ “planetary boundary threats from chemical pollution”,^{41–43} “poor reversibility” of exposure for persistent

pollutants^{44,45} and recently in the “P-sufficient approach”.³⁶ As well as not being new, the arguments have always been simple to understand for scientists and non-scientists alike, and have never been convincingly challenged. Arguments against the “P-sufficient” approach often point to inert substances such as some inorganic minerals that are also highly persistent and often in a non-bioaccessible form (*e.g.* sand and rocks). However, apart from their similar high persistence, PFAS are distinct from inert inorganic minerals in that PFAS (i) are synthetic organic substances and not naturally present (ii) continue to be manufactured, leading to ongoing accumulation in the environment, and (iii) are often bioaccessible. It is noteworthy that only about 20 biogenic organofluorine substances have been identified and PFAS are not among those listed.⁴⁶ Muir *et al.*⁴⁷ speculate that the substantial quantities of trifluoroacetic acid found in the deep ocean (estimate of 63 to 187 million tonnes) originate from deep-sea vents (*i.e.* a geogenic source), but this has not been proven.

There are finally encouraging signs that regulators are taking notice of the problems with high persistence (*e.g.* in the case of intentionally added microplastics) and that the European Green Deal will hopefully introduce regulatory provisions for highly persistent substances. It is further encouraging that in recent years concepts such as “green chemistry”⁴⁸ and “cradle-to-cradle”⁴⁹ are gaining momentum. For example, one of the 12 principles of green chemistry is: “design for degradation”,⁴⁸ *i.e.* chemical products should be designed so that at the end of their function they break down into innocuous transformation products and do not persist in the environment. Highly persistent substances such as PFAS are incompatible with green chemistry principles and future visions of sustainable development (*e.g.* the European Green Deal). The concept of essentiality has been put forward as a tool for guiding the phase-out of hazardous substances such as PFAS⁵⁰ and efforts have already begun to address this challenging, but not insurmountable, task in the EU and elsewhere. We are of the opinion that the EU restriction proposal¹ for PFAS should have a wide scope grounded in the one completely consistent feature of PFAS: their persistence, based on the arguments presented in this Perspective article.

Regulation alone is not sufficient to phase out PFAS. Innovation will also be needed by chemical and product manufacturers to develop safer chemical alternatives to PFAS and technological innovation that would make the function obsolete or the use of PFAS unnecessary. There is cause to be optimistic about the future, given the innovations and changes implemented in recent years (*e.g.* the transition to fluorine-free aqueous film-forming foams in South Australia⁵¹ and elsewhere, the wider availability of fluorine-free durable water repellent outdoor equipment,⁵² the banning of fluorinated ski waxes in international competition,⁵³ the phasing out of PFAS in cosmetics by multiple retailers,⁵⁴ the banning of PFAS in paper and board used in food contact materials in Denmark⁵⁵), but the work is far from finished.

In summary, although PFAS have diverse structures and behaviour, the high persistence of PFAS is sufficient cause of concern for all “non-essential” uses of PFAS to be phased out.



Unfortunately, there has until very recently been no legal precedent in any jurisdiction for regulation on high persistence alone. Promisingly, the idea of regulating highly persistent chemicals and intentionally added microplastics is being explored within the EU. We hope that this idea will spread around the world.

Conflicts of interest

There are no conflicts of interest to declare.

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