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Photoinduced electron-driven proton transfer from water to an N-heterocyclic chromophore: nonadiabatic dynamics studies for pyridine–water clusters

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It has been found in recent molecular beam experiments that pyridine molecules photoexcited at 255 nm can abstract hydrogen atoms from hydrogen-bonded water molecules in pyridine–water clusters, resulting in pyridinyl-hydroxyl radical pairs. The reaction could only be detected for clusters containing at least four water molecules. To provide insight into the mechanisms of this reaction, we performed *ab initio* excited-state trajectory surface-hopping dynamics simulations for two pyridine–water complexes, containing one and four water molecules, respectively, using the second-order algebraic-diagrammatic-construction (ADC(2)) electronic-structure method. A computationally efficient surface-hopping algorithm based on the Landau–Zener formula has been used to evaluate the transition probability between electronic states. The formation of the pyridinyl radical via an electron-driven proton transfer (EDPT) process from water to pyridine is confirmed by the simulations. The analysis of the competing excited-state reaction mechanisms up to 500 fs reveals that adiabatic relaxation to local minima of the $S_1(n\pi^*)$ potential-energy surface is the dominant channel in both clusters, followed by internal conversion to the electronic ground state via so-called ring-puckering conical intersections. The efficiency of the latter contribution is weakly dependent of the size of the clusters. The EDPT reaction occurs on the fastest time scales (faster than 200 fs) with a branching ratio of several percent. It is found to be four times more efficient in the pyridine–(H₂O)₄ cluster than in the pyridine–H₂O cluster, which is qualitatively consistent with the experimental observations. A detailed understanding of the photoinduced reaction mechanisms in complexes of N-heterocyclic chromophores with water molecules is of relevance for future systematic knowledge-based developments of optimized materials for photocatalytic water splitting with sunlight.

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1. Introduction

Photoinduced electron-driven proton-transfer (EDPT) reactions,¹ a subset of so-called proton-coupled electron-transfer (PCET) reactions,² are of great importance in the photochemistry of hydrogen-bonded complexes, being relevant for the function of various sunscreens,^{3,4} the photostability of nucleobases,^{5,6} and for artificial photosynthesis.^{7–11} The most extensively studied EDPT reactions are excited-state chromophore-to-solvent hydrogen-transfer reactions, in which a hydrogen atom is transferred from a photoexcited chromophore to solvent

molecules, such as ammonia or water.^{12–17} Extensive experimental and theoretical investigations have shown that these reactions can take place in clusters consisting of several water molecules and an acidic aromatic chromophore, such as phenol,^{14,17} adenine¹⁸ or imidazole.^{19,20} It was found that the number of solvent molecules in micro-solvated systems can have a significant influence on the rate constant of the EDPT reaction.^{19–21}

Alternatively, an H-atom can be transferred from a solvent molecule to a photoexcited chromophore. While limited experimental evidence exists for such photoreactions, extensive computational studies have shown that heterocyclic aromatic chromophores, such as pyridine,^{22,23} triazine,²⁴ heptazine²⁵ or adenine,¹⁸ can upon photoexcitation abstract a hydrogen atom from a water molecule by a low-barrier H-atom transfer reaction, resulting in the formation of a reduced chromophore and a hydroxyl radical. By the absorption of a second photon, the H-atom can be photodetached from the reduced chromophore.

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Thereby the chromophore is recovered and the water molecule is split with two photons into a hydrogen atom and a hydroxyl radical. Recently, this biphotonic homolytic water-splitting reaction was discovered in a molecular beam experiment for photo-excited hydrogen-bonded pyridine–water clusters by Jouvet and coworkers.²⁶ In this experiment, the formation of the pyridinyl neutral radical was detected in pyridine–water clusters upon 255 nm irradiation, which confirms that pyridine (Py) can serve as a photocatalyst for the splitting of water molecules by UV light. An interesting observation made in this experiment was that at least four water molecules were necessary for the photoinduced water-splitting reaction to be detectable.²⁶ To find an explanation for this observation, we recently explored the photoinduced EDPT reaction from water to pyridine in $\text{Py}-(\text{H}_2\text{O})_n$ ($n = 1-4$) clusters.²⁷ It was found that for $\text{Py}-(\text{H}_2\text{O})_n$ clusters with $n = 1-3$, the hydrogen bond between pyridine and the nearest water molecule is significantly weakened in the S_1 excited state of $n\pi^*$ character, which is unfavorable for the EDPT reaction. In the $\text{Py}-(\text{H}_2\text{O})_4$ cluster, on the other hand, the S_1 excited state is of $\pi\pi^*$ character and the hydrogen bond with the nearest water molecule is strengthened. Moreover, determination of the reaction barriers showed that the $\text{Py}-(\text{H}_2\text{O})_4$ cluster has the lowest barrier for the EDPT reaction among these clusters.

While these static excited-state electronic-structure calculations provide a plausible explanation why at least four water molecules are necessary for the EDPT reaction in pyridine–water complexes, zero-point energy and finite-temperature effects as well as dynamical phenomena, which are not taken into account in these static calculations, may be important in the EDPT reaction. In the present study, *ab initio* on-the-fly nonadiabatic dynamics simulations have been performed to further unveil the underlying mechanisms of the EDPT reaction in pyridine–water clusters. The relatively small pyridine–water clusters serve herein as model systems which allow relatively inexpensive explorations of these phenomena with first-principles methods. These studies are intended to pave the way for simulations of the photochemistry of larger chromophores, such as heptazine (tri-s-triazine)^{25,28} or trianisoheptazine²⁹ in aqueous environments. These chromophores are potential photocatalysts for water splitting with visible light.^{29,30}

2. Computational methods

The ground-state equilibrium geometries of $\text{Py}-(\text{H}_2\text{O})_n$ ($n = 1$ and 4) complexes were optimized with second-order Møller–Plesset perturbation theory (MP2)³¹ with the cc-pVDZ basis set. For the $\text{Py}-(\text{H}_2\text{O})_4$ complex, there exist numerous equilibrium structures with comparable energies. We selected a low-energy structure for this cluster such that the hydrogen bonds among the water molecules are optimally saturated. The excited-state calculations were performed with the second-order algebraic-diagrammatic-construction (ADC(2)) method^{32,33} using the cc-pVDZ basis set. The TURBOMOLE program package³⁴ was used for the electronic-structure calculations.

The absorption spectra of the pyridine–water complexes were simulated using the nuclear ensemble method proposed by Barbatti and coworkers.³⁵ A set of 2000 initial conditions were generated from the harmonic-oscillator thermal Wigner distribution at room temperature. Vertical excitation energies and oscillator strengths of the five lowest-lying excited states were computed for each initial geometry. The lines were broadened by a Lorentzian function with a width of 0.05 eV and summed up to provide the spectrum.

The nonadiabatic dynamics was simulated with a trajectory surface-hopping (TSH) method which is based on the Landau–Zener (LZ) formula.^{36–38} In this method, the transition probability between adiabatic electronic states for a given classical trajectory is determined from the local shape of the adiabatic potential energies near surface crossings. This is possible because the electronic mixing angle (and therefore the nonadiabatic electronic coupling element) in the vicinity of a two-state avoided crossing is uniquely determined by the shape of the adiabatic potential-energy functions. The explicit construction of diabatic potential-energy surfaces is therefore not required in this method nor is the time-dependent electronic wave function computed. Therefore, the “lack-of-decoherence” problem, which arises in the widely used fewest-switches surface-hopping (FSSH) method of Tully³⁹ does not apply to these calculations. The LZ-based surface-hopping (LZSH) method is computationally more efficient than the fewest-switches algorithm, since the computation of derivative couplings, which are sharply peaked at conical intersections, is avoided.

Briefly, the absolute value of the energy gap ΔV_{jk}^a between the energy of the current adiabatic electronic state j and another adiabatic electronic state k is monitored at each time step of the trajectory. If a local minimum of the energy gap is attained, the hopping probability is evaluated according to the adiabatic formulation of the Landau–Zener formula^{36–38}

$$P_{\text{LZ}}^{j \rightarrow k} = \exp \left(-\frac{\pi}{2\hbar} \sqrt{\frac{\Delta V_{jk}^a(R(t^*))^3}{\left. \frac{d^2 \Delta V_{jk}^a(R(t))}{dt^2} \right|_{t=t^*}}} \right) \quad (1)$$

As in the fewest-switches algorithm,³⁹ a pseudorandom number uniformly generated from (0,1) is compared with $P_{\text{LZ}}^{j \rightarrow k}$ to determine whether the trajectory hops to state k or continues in the current state j .

The accuracy of the LZSH method has been critically examined in comparison with the FSSH method and with exact nonadiabatic quantum dynamics calculations for a few benchmark model systems. These calculations showed that the LZSH method provides a qualitatively correct description for few-state few-mode models of the excited-state nonadiabatic dynamics of the ammonia cation,³⁷ phenol³⁸ and pyrazine.⁴⁰ The *ab initio* direct-dynamics implementation of the LZSH method has recently been applied to the investigation of the radiationless decay dynamics of the $S_2(\pi\pi^*)$ excited state of pyrazine.⁴⁰

In the simulations, Newton's equations for the nuclear motion were integrated with a time step of 0.5 fs using the



velocity-Verlet algorithm. The trajectories were propagated up to 500 fs. Since the ADC(2) excitation energies are determined by the diagonalization of a Hermitian matrix,³² the adiabatic energies exhibit proper behavior in the vicinity of conical intersections between excited states and hopping probabilities between excited states can be evaluated according to eqn (1). This is not the case for S_1 – S_0 conical intersections, since the single-reference character of the electronic ground-state wave function necessarily breaks down in the vicinity of S_1 – S_0 conical intersections. Therefore, the propagation of a trajectory was terminated when the energy gap between the S_1 surface and the S_0 surface dropped below 0.2 eV. These events were counted as radiationless decay to the electronic ground state. Possible back reactions from the S_0 surface to the S_1 surface and possible H-atom back-transfer reactions on the S_0 surface are therefore neglected.

3. Results and discussion

3.1. Absorption spectrum

The ground-state equilibrium geometries of the Py–H₂O and Py–(H₂O)₄ clusters are shown in Fig. 1(a) and (b), respectively. The pyridine–water hydrogen-bond length (R_{NH}) of the Py–(H₂O)₄ cluster is 0.23 Å shorter than the hydrogen-bond length of the Py–H₂O cluster, which indicates that the strength of the hydrogen bond is substantially increased in the larger cluster.

A comparison of the computed absorption spectra of the Py–H₂O and Py–(H₂O)₄ clusters with the experimental spectrum of pyridine in aqueous solution^{41,42} is shown in Fig. 2. For the sake of comparison, the maxima of the spectra are normalized to unity. The peak of the absorption band of the Py–(H₂O)₄ cluster is red-shifted by 0.12 eV with respect to the peak of the spectrum of the Py–H₂O cluster. This is due to a redshift of the bright $^1\pi\pi^*$ state of Py with increasing number of water molecules. While the $^1n\pi^*$ state of pyridine is optically dark, it can borrow intensity from the bright $^1\pi\pi^*$ state in the Py–H₂O cluster, since the $^1\pi\pi^*$ and $^1n\pi^*$ state are nearly degenerate in this cluster.^{22,27} As the number of water molecules increases from one to four, the stabilization of the n orbital by the surrounding water molecules leads to a further increase of the excitation energy of the $^1n\pi^*$ state, which removes the near-degeneracy of



Fig. 2 Absorption spectra of the Py–H₂O cluster (blue) and the Py–(H₂O)₄ cluster (red), simulated at the ADC(2)/cc-pVDZ level. The black line is the experimental spectrum of pyridine in aqueous solution.³⁸ The initial conditions for excited-state dynamics simulations of both clusters are sampled from an energy window of 5.06 ± 0.05 eV (marked in grey).

the $^1\pi\pi^*$ and $^1n\pi^*$ states and results in a lower oscillator strength of the $^1n\pi^*$ state. The ADC(2) method predicts the band maximum of the Py–(H₂O)₄ cluster at 5.06 eV, which is blue-shifted by 0.2 eV with respect to the band maximum of the experimental spectrum.⁴¹

3.2. Nonadiabatic dynamics simulations

The initial conditions for excited-state dynamics simulations of the Py–H₂O and Py–(H₂O)₄ clusters were sampled from an energy window of 5.06 ± 0.05 eV to mimic the gas-phase experiment in which the laser frequency was chosen near the peak of the absorption spectrum of pyridine.²⁶ A stochastic algorithm, using the oscillator strength as the criterion for the transition probability was used, which results in 144 and 65 trajectories started in the S_1 and S_2 states, respectively, of the Py–H₂O cluster. In the Py–(H₂O)₄ cluster, 180 and 58 trajectories are initiated in the S_1 and S_2 states, respectively. The calculated time-dependent population probabilities of the ground state (with zero initial population) and the three lowest excited adiabatic electronic states of the Py–H₂O and Py–(H₂O)₄ clusters are shown in Fig. 3. Overall, the electronic population dynamics of the two clusters is similar. The S_2 state population exhibits a few weak recurrences and then decays to a low value within 100 fs. The population of the S_1 state attains a maximum of 0.95 at 50 fs and then decays with a long decay constant of 4.3 ps (2.5 ps) for the Py–H₂O (Py–(H₂O)₄) cluster. The population of the S_3 state (a higher $^1n\pi^*$ state of pyridine) rises to a maximum value of 0.1 (0.05) within the first 10 fs for the Py–(H₂O)₄ (Py–H₂O) cluster and then quickly decays to zero. The ground-state population increases monotonously with the same time constant as the S_1 decay. The computed excited-state lifetime of the Py–(H₂O)₄ cluster is (presumably fortuitously) almost identical with the experimental result of 2.2 ps for pyridine in aqueous solution.⁴¹

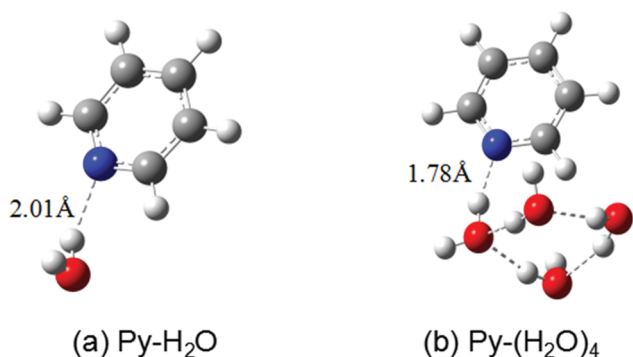


Fig. 1 Ground-state equilibrium geometries of the pyridine–(H₂O) cluster (a) and the pyridine–(H₂O)₄ cluster (b) calculated with the MP2/cc-pVDZ method.

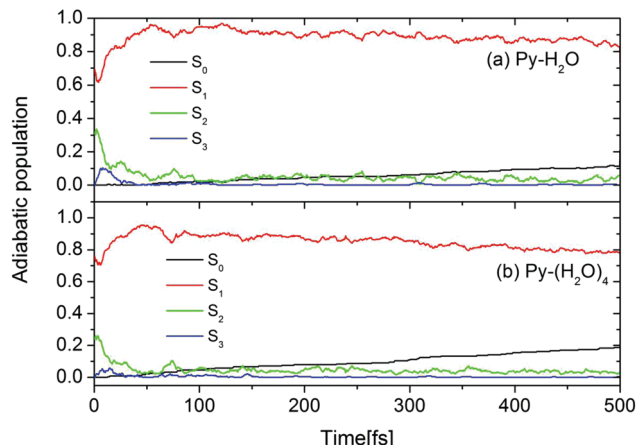


Fig. 3 Population probabilities of the electronic ground state and three lowest adiabatic excited electronic states for the Py-H₂O (a) and Py-(H₂O)₄ (b) clusters.

The trajectories were analyzed with respect to the reaction mechanism they represent. The fractions of trajectories corresponding to the different radiationless deactivation mechanisms for the Py-H₂O and Py-(H₂O)₄ clusters are shown in Fig. 4(a) and (b), respectively. At the end of the simulations, relaxation to local minima of the S₁(nπ*) surface of pyridine is the dominant channel for both clusters (88.9% and 81.9%, respectively). Most of these trajectories show a migration of the water molecules towards one of faces the pyridine ring, while the hydrogen bond between the N-atom of pyridine and the nearest water molecule weakens or breaks. 10.1% and 13.9% of trajectories return to the ground state through conical intersections of the ring-puckering type for the Py-H₂O and Py-(H₂O)₄ clusters, respectively. S₁-S₀ conical intersections of pyridine have been located and discussed in ref. 43 and 44. They result from the lowering of the energy of the diabatic ¹ππ* state and the increase of the energy of the S₀ state of pyridine upon out-of-plane distortion of the nitrogen atom. These ring-puckering conical intersections are well known for benzene^{45,46} as well as for N-heterocycles, for example pyrazine,⁴⁶ or pyrimidine bases, such as uracil or thymine.⁴⁷



Fig. 4 Branching ratios of different excited-state deactivation channels for the Py-H₂O cluster (a) and the Py-(H₂O)₄ cluster (b) within the first 500 fs after excitation.

Increasing the number of water molecules seems to have a moderate effect on the efficiency of the ring-puckering internal conversion pathway.

For the Py-(H₂O)₄ cluster, about 4.2% of the trajectories of the simulations up to 500 fs reach a conical intersection with ground state *via* the EDPT photorelaxation pathway, which is about four times more than found for the Py-H₂O cluster. The EDPT pathway is thus not the dominant pathway for the radiationless deactivation of photoexcited pyridine-water clusters. It should be kept in mind, however, that the branching ratio for the EDPT process is systematically underestimated by the present simulations due to the omission of proton tunneling in the classical trajectory calculations.

Fig. 5 shows a histogram of the excited-state lifetimes, defined as the time when the trajectory hits an S₁/S₀ conical intersection. The numbers of trajectories returning to the ground state by the EDPT and ring-puckering channels are represented by the black and red histograms, respectively. Most EDPT events occur within the first 200 femtoseconds, while most of the ring-puckering events take place after 300 fs. All EDPT events in the Py-H₂O cluster occur within the first 100 fs, while in the Py-(H₂O)₄ cluster some EDPT events take place after 200 fs. The surrounding water molecules in the Py-(H₂O)₄ cluster stabilize the hydrogen bond of the central water molecule with pyridine, which reduces the probability of dissociation of the pyridine-water hydrogen bond.

This interpretation is confirmed by Fig. 6, which shows the time evolutions of the N-H hydrogen-bond distance for all trajectories of both clusters. During the first 100 femtoseconds, the N-H bond distance is seen to increase significantly faster for the Py-H₂O cluster than for the Py-(H₂O)₄ cluster. For both clusters, the pyridine-water hydrogen bond for most trajectories either weakens or breaks within 100 fs. The trajectories undergoing the EDPT reaction in the Py-(H₂O)₄ cluster roam around conical intersection seams for several vibrational periods of the mobile H-atom before they cross to the electronic ground state (indicated by trajectories with $R_{\text{NH}} < 1.5$ Å in Fig. 6(b)).



Fig. 5 Distributions of excited-state lifetimes, defined as the time when trajectories hit a conical intersection with the electronic ground state, for the Py-H₂O cluster (a) and the Py-(H₂O)₄ cluster (b). Black: deactivation by EDPT; red: deactivation *via* ring-puckering conical intersections.





Fig. 6 Time evolution of the N–H hydrogen-bond distance between pyridine and water for the Py–H₂O cluster (a) and the Py–(H₂O)₄ cluster (b).

Similar N–H vibrational oscillations are not observed in the Py–H₂O cluster (Fig. 6(a)).

The time-dependent potential energies of the ground state and two lowest excited states are shown in Fig. 7(a) for a typical trajectory undergoing the EDPT reaction in the Py–(H₂O)₄ cluster. At $t = 0$, the trajectory is started in the adiabatic S_1 state which is the locally-excited $1\pi\pi^*$ state in the Franck–Condon region. At 52.5 fs, the energy gap between the S_1 and the S_2 states reaches a local minimum and a hopping takes place from the S_1 to the S_2 state. The evolution in the S_2 state leads to a pronounced destabilization of the electronic ground state. The time evolutions of the S_2 – S_1 and S_1 – S_0 energy gaps are displayed in Fig. 7(b). After about 30 fs propagation in the quasi-degenerate S_1 and S_2 states, the trajectory hits a conical intersection with the ground state. In this region, the S_2 – S_1 as well as S_1 – S_0 energy gaps are

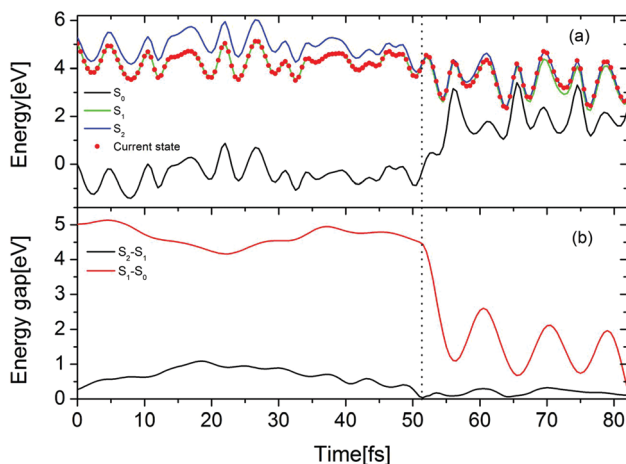


Fig. 7 (a) Time evolution of the potential energies of the electronic ground state (black) and two lowest excited states (green and blue) for a selected trajectory, computed at ADC(2)/cc-pVDZ level. The red dot indicates the currently populated state. (b) Time evolution of the energy gaps between the S_2 and S_1 states (black) and S_1 and S_0 states (red). The vertical dotted line marks the time when the trajectory hops from the S_1 to the S_2 state.

small (0.11 eV for S_2 – S_1 and 0.29 eV for S_1 – S_0), which implies that this conical intersection may be a three-state conical intersection.

To gain further insight into the EDPT process in the Py–(H₂O)₄ cluster, the time-dependent energies of the occupied frontier Hartree–Fock orbitals are shown in Fig. 8(a) for a selected trajectory which was initiated in the $S_1(\pi\pi^*)$ state. The dot marks the orbital which carries the hole in the excited electronic state. In the lower part of the figure, snapshots of the hole-carrying orbital are shown at specific times. The photoexcitation of pyridine to the $1\pi\pi^*$ state generates a hole in the π orbital of pyridine and the hole amplitude is localized on the pyridine ring (0 fs). The energy of the π orbital of pyridine crosses the energies of the accidentally nearly degenerate n (on pyridine) and p_z (on the hydrogen-bonded water molecule) orbitals at 52.5 fs. The trajectory hopping from the S_1 state to the nearly degenerate S_2 state is accompanied by a hole transfer from the π orbital to the n orbital. The amplitude of the latter is distributed almost equally among the pyridine ring and the hydrogen-bonded water molecule (see Fig. 8, 52.5 fs). As time proceeds, the amplitude becomes fully localized on the $p_{x/y}$ orbital of the water molecule (see Fig. 8, 55.0 fs). Beyond 55 fs, the hole remains localized on the water molecule and oscillates between the $p_{x/y}$ and p_z orbitals (see Fig. 8, 60.0 fs). Once the hole is localized on the water molecule, the proton involved in the hydrogen bond with the N-atom of pyridine is driven away for the oxygen atom of water and attaches to the N-atom of the negatively charged pyridine, which completes the EDPT reaction. The time evolution of the N–H and O–H distances is shown in Fig. 8(b). For the initial 50 fs, the N–H distance decreases gradually with periodic oscillations with a period of 10 fs.



Fig. 8 (a) Time evolution of the orbital energies and (b) time-dependent N–H and O–H bond lengths for a selected trajectory. The circle in (a) indicates the current hole-carrying orbital. Three orbitals cross in energy at 52.5 fs which is marked with yellow star. The lower part shows hole orbitals at selected times.



The O–H bond length is equal to the N–H hydrogen-bond length at 52.5 fs. As time proceeds, the O–H bond breaks and a covalent N–H bond is formed. It can be seen that the formation of the covalent N–H bond is nearly coincident with the hole transfer from the π orbital to the n orbital shown in Fig. 8(a).

4. Conclusions

In the present work, we performed the first excited-state non-adiabatic dynamics simulations for photoexcited pyridine–water clusters, using a recently developed surface-hopping method based on the Landau–Zener formula at PES crossings and the ADC(2) electronic-structure model. To obtain insight into the dependence of the dynamics on the size of the clusters, the reaction dynamics of two clusters, containing one and four water molecules, respectively, was considered. Recent benchmarking studies have shown that this surface-hopping method can provide qualitatively correct descriptions of the nonadiabatic excited-state dynamics.⁴⁰ A limitation of the simulations at the present stage is the neglect of proton tunneling effects, which may lead to a systematic underestimation of the branching ratio of the EDPT reaction.

The dynamics simulations for the two selected Py–water clusters show that relaxation to local minima of the $S_1(n\pi^*)$ PES is the dominant reaction channel. On the S_1 surface, a rearrangement of the clusters takes place by an extension of the hydrogen bond with the N-atom of Py and the formation of hydrogen bonds with the π -system of the aromatic ring. Internal conversion to the electronic ground state *via* S_1 – S_0 conical intersections of the ring-puckering type is found with a branching ratio of about 10%, which does not strongly depend on the size of the clusters. The EDPT process, that is, H-atom transfer from the N-hydrogen-bonded water molecule to the N-atom of pyridine, is the fastest process with a low branching ratio (1%) for the Py–H₂O cluster. This branching ratio increases to about 4% in the Py–(H₂O)₄ cluster. The increased efficiency of the EDPT reaction in the Py–(H₂O)₄ cluster is qualitatively consistent with the observation that clusters with at least four water molecules are necessary for the experimental detection of the excited-state H-atom transfer reaction.²⁶ The relatively efficient excited-state deactivation of Py *via* ring-puckering conical intersection is a loss process which is undesirable for photocatalytic water splitting. Out-of-plane deformations in $\pi\pi^*$ excited states are suppressed, however, in fused aromatic ring systems, such as, for example, heptazine. Heptazine-based polymeric materials, such as graphitic carbon nitride,⁴⁸ indeed show excellent photostability.

The mechanism of the EDPT reaction mechanism has been visualized in the Hartree–Fock orbital picture. Photoexcitation of an electron from the π orbital to the π^* orbital in the pyridine chromophore in Py–water complexes initially creates an exciton (electron–hole pair). When the OH distance of the N-hydrogen-bonded water molecule extends (that is, the NH distance decreases) during the dynamics, the energy of the π orbital decreases, while the energies of the n orbital of pyridine and

the p orbitals of the water molecule increase. When two (or more) of these orbitals become degenerate, a hole transfer from pyridine to the nearest water molecule takes place, resulting in a negatively charged pyridine ring and a positively charged water molecule. This charge separation provides the driving force for the proton transfer from the water molecule to the pyridine ring, resulting in the pyridinyl-hydroxyl biradical.

The first *ab initio* study of the dynamics of hydrogen abstraction from a water molecule by a photoexcited organic chromophore was performed for the cluster of adenine with a single water molecule by Došlić and coworkers.⁴⁹ In the present study, the first *ab initio* excited-state dynamics simulations have been performed for pyridine–water clusters which reveal the effect of an increasing number of water molecules on the photoinduced water-splitting reaction. In future extensions of the present work, the effects of proton tunneling should be quantified. It should also be interesting to perform simulations for photoexcited aromatic chromophores in a liquid water environment at finite temperature with *ab initio* (DFT) molecular dynamics simulations or QM/MM simulation methods.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 A. L. Sobolewski and W. Domcke, *J. Phys. Chem. A*, 2007, **111**, 11725.
- 2 D. R. Weinberg, C. J. Gagliardi, J. F. Hull, C. F. Murphy, C. A. Kent, B. C. Westlake, A. Paul, D. H. Ess, D. G. McCafferty and T. J. Meyer, *Chem. Rev.*, 2012, **112**, 4016.
- 3 J. E. A. Otterstedt, *J. Chem. Phys.*, 1973, **58**, 5716.
- 4 L. A. Baker, B. Marchetti, T. N. V. Karsili, V. G. Stavros and M. N. R. Ashfold, *Chem. Soc. Rev.*, 2017, **46**, 3770.
- 5 A. L. Sobolewski and W. Domcke, *Europhys. News*, 2006, **37**, 20.
- 6 A. L. Sobolewski and W. Domcke, *Phys. Chem. Chem. Phys.*, 2004, **6**, 2763.
- 7 J. Hemminger, G. Crabtree and M. Kastner, *Basic Energy Sciences Advisory Committee*, US Department of Energy, 2008.
- 8 M. Pagliaro, A. G. Konstandopoulos, R. Ciriminna and G. Palmisano, *Energy Environ. Sci.*, 2010, **3**, 279.
- 9 G. F. Moore and G. W. Brudvig, *Annu. Rev. Condens. Matter Phys.*, 2011, **2**, 303.
- 10 D. Gust, T. A. Moore and A. L. Moore, *Faraday Discuss.*, 2012, **155**, 9.
- 11 M. D. Kärkäs, O. Verho, E. V. Johnston and B. Åkermark, *Chem. Rev.*, 2014, **114**, 11863.



- 12 O. David, C. Dedonder-Lardeux and C. Jouvet, *Int. Rev. Phys. Chem.*, 2002, **21**, 499.
- 13 A. L. Sobolewski, W. Domcke, C. Dedonder-Lardeux and C. Jouvet, *Phys. Chem. Chem. Phys.*, 2002, **4**, 1093.
- 14 A. L. Sobolewski and W. Domcke, *J. Phys. Chem. A*, 2001, **105**, 9275.
- 15 A. L. Sobolewski and W. Domcke, *Chem. Phys. Lett.*, 2000, **321**, 479.
- 16 A. L. Sobolewski and W. Domcke, *Chem. Phys. Lett.*, 2000, **329**, 130.
- 17 W. Domcke and A. L. Sobolewski, *Science*, 2003, **302**, 1693.
- 18 X. Wu, T. N. V. Karsili and W. Domcke, *ChemPhysChem*, 2016, **17**, 1298.
- 19 R. Szabla, R. W. Góra, M. Janicki and J. Šponer, *Faraday Discuss.*, 2016, **195**, 237.
- 20 R. Szabla, J. Šponer and R. W. Góra, *J. Phys. Chem. Lett.*, 2015, **6**, 1467.
- 21 R. Szabla, J. E. Šponer, J. Šponer, A. L. Sobolewski and R. W. Góra, *Phys. Chem. Chem. Phys.*, 2014, **16**, 17617.
- 22 X. Liu, A. L. Sobolewski, R. Borrelli and W. Domcke, *Phys. Chem. Chem. Phys.*, 2013, **15**, 5957.
- 23 X. Liu, A. L. Sobolewski and W. Domcke, *J. Phys. Chem. A*, 2014, **118**, 7788.
- 24 J. Ehrmaier, M. J. Janicki, A. L. Sobolewski and W. Domcke, *Phys. Chem. Chem. Phys.*, 2018, **20**, 14420.
- 25 J. Ehrmaier, T. N. V. Karsili, A. L. Sobolewski and W. Domcke, *J. Phys. Chem. A*, 2017, **121**, 4754.
- 26 N. Esteves- López, S. Coussan, C. Dedonder-Lardeux and C. Jouvet, *Phys. Chem. Chem. Phys.*, 2016, **18**, 25637.
- 27 X. Pang, J. Ehrmaier, X. Wu, C. Jiang, W. Xie, A. L. Sobolewski and W. Domcke, *Chem. Phys.*, 2018, **515**, 550.
- 28 J. Ehrmaier, W. Domcke and D. Opalka, *J. Phys. Chem. Lett.*, 2018, **9**, 4695.
- 29 E. J. Rabe, K. L. Corp, A. L. Sobolewski, W. Domcke and C. W. Schlenker, *J. Phys. Chem. Lett.*, 2018, **9**, 6257.
- 30 W. Domcke, J. Ehrmaier and A. L. Sobolewski, *ChemPhotoChem*, 2019, **3**, 10.
- 31 C. Møller and M. S. Plesset, *Phys. Rev.*, 1934, **46**, 618.
- 32 J. Schirmer, *Phys. Rev. A: At., Mol., Opt. Phys.*, 1982, **26**, 2395.
- 33 C. Hättig, *Adv. Quantum Chem.*, 2005, **50**, 37.
- 34 TURBOMOLE, v. 6.6, 2014; a development of the University of Karlsruhe and Forschungszentrum Karlsruhe 1989–2007; TURBOMOLE GmbH since 2007.
- 35 M. Barbatti, M. Ruckebauer, F. Plasser, J. Pittner, G. Granucci, M. Persico and H. Lischka, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2014, **4**, 26.
- 36 A. K. Belyaev and O. V. Lebedev, *Phys. Rev. A: At., Mol., Opt. Phys.*, 2011, **84**, 014701.
- 37 A. K. Belyaev, W. Domcke, C. Lasser and G. Triglia, *J. Chem. Phys.*, 2015, **142**, 104307.
- 38 W. Xie and W. Domcke, *J. Chem. Phys.*, 2017, **147**, 184114.
- 39 J. C. Tully, *J. Chem. Phys.*, 1990, **93**, 1061.
- 40 W. Xie, M. Sapunar, M. Sala, N. Došlić and W. Domcke, *J. Chem. Phys.*, in revision.
- 41 H. H. Perkampus, *UV-Vis Atlas of Organic Compounds*, VCH, 1992.
- 42 M. Chachisvilis and A. H. Zewail, *J. Phys. Chem. A*, 1999, **103**, 7408.
- 43 A. L. Sobolewski and W. Domcke, *Chem. Phys. Lett.*, 1991, **180**, 381.
- 44 P. C. Varras, P. S. Gritzapis and K. C. Fylaktakidou, *Mol. Phys.*, 2018, **116**, 154.
- 45 I. J. Palmer, I. N. Ragazos, F. Bernardi, M. Olivucci and M. A. Robb, *J. Am. Chem. Soc.*, 1993, **115**, 673.
- 46 A. L. Sobolewski, C. Woywod and W. Domcke, *J. Chem. Phys.*, 1993, **98**, 5627.
- 47 R. Improta, F. Santoro and L. Blancafort, *Chem. Rev.*, 2016, **116**, 3540.
- 48 Y. Wang, X. Wang and M. Antonietti, *Angew. Chem., Int. Ed.*, 2012, **51**, 68.
- 49 S. Chaiwongwattana, M. Sapunar, A. Ponzì, P. Decleva and N. Došlić, *J. Phys. Chem. A*, 2015, **119**, 10637.

