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659Photoelectrochemical water-splitting over a
surface modified p-type Cr₂O₃ photocathode†Keita Sekizawa,  * Keiichiro Oh-ishi  and Takeshi Morikawa 

Cr₂O₃ is a p-type semiconductor with a negative conduction band minimum position suitable for photocathodic H₂ generation. Therefore, Cr₂O₃ is a candidate photocathode material for photoelectrochemical (PEC) water-splitting. However, Cr₂O₃ has not yet been applied for the purpose of H₂ generation because the efficiency and stability of the photocurrent generated by a Cr₂O₃ electrode are poor, due to high defect and vacancy concentrations. In the present work, the Cr₂O₃ surface was modified with n-type TiO₂ after which Pt particles were deposited to catalyse H₂ production. The TiO₂ overlayer passivated the Cr₂O₃ surface states that otherwise cause deleterious interactions with the Pt particles. This layer also improved charge separation from the conduction band of Cr₂O₃ to the Pt co-catalyst, by forming a p–n junction. As a result of the TiO₂ insertion, the cathodic photocurrent resulting from light absorption by Cr₂O₃ was enhanced and stabilized. This represents the first-ever use of Cr₂O₃ as a light-absorbing material in a multi-layered electrode to accomplish PEC water-splitting for H₂ generation.

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Introduction

H₂ has attracted considerable interest as a clean energy carrier that does not produce CO₂, and is expected to play an important role in meeting growing worldwide energy requirements.^{1–3} Photoelectrochemical (PEC) water-splitting into H₂ and O₂ using semiconductor electrodes is a promising approach to the direct conversion of solar energy into H₂.^{4–6} In 1972, Honda and Fujishima first reported the PEC decomposition of water to H₂, utilizing an n-type TiO₂ electrode in conjunction with UV illumination and the application of an external electrical bias.⁷ After this pioneering work, numerous semiconductor photocatalysts and photoelectrodes were developed.^{8–17} It is widely recognized that most semiconductors require the addition of a sacrificial electron donor or the application of an external voltage, as the edges of the conduction band and valence band typically do not straddle the water reduction and oxidation potentials. As a means of promoting water splitting without an external electrical bias while separately generating H₂ and O₂ at different electrodes, PEC devices containing electrically connected p-type and n-type semiconductor photoelectrodes have been designed.^{18–22} Metal oxides are suitable semiconductor materials because they are abundant and relatively in-

expensive. Since metal oxide semiconductors predominantly exhibit n-type characteristics, because they readily form oxygen vacancies, there have been many reports regarding the use of n-type semiconductor photoanodes to generate oxygen. In contrast, there are only a few native p-type semiconductor oxides, because the cation vacancies or oxygen interstitials that produce p-type conductivity are not easily generated.^{23,24} Moreover, utilizing photocathodes for H₂ evolution requires that the conduction band minimum of the cathode material be more negative than the H⁺/H₂ reduction potential (0 V vs. reversible hydrogen electrode; RHE). The oxide p-type semiconductors meeting this requirement are limited, with examples being copper oxides^{25,26} and iron oxides.^{27–29}

Cr₂O₃ is one of the few oxides exhibiting p-type conductivity.^{30,31} Standard Cr₂O₃ has a hexagonal corundum structure.^{32,33} In the unit cell of non-defective Cr₂O₃, two-thirds of the octahedral oxygen sites are occupied by Cr while the remaining one-third are unoccupied (Fig. S1†). Consequently, removing Cr atoms from the cell generates vacancies. Moreover, the migration of Cr to the pre-existing empty sites (typically interstitial sites in nearby locations) results in Cr Frenkel defects, while the diffusion of O atoms generates O vacancies. These defects and vacancies are known to determine the appearance of p-type and n-type characteristics.^{34–38} The defect formation energy values calculated using first-principles DFT+U by Diawara *et al.*³⁸ suggest that the stability of defects follows the order: Cr Frenkel defect > Cr vacancy > O vacancy. Therefore, native Cr₂O₃ tends to form metal-deficient sites that result in p-type conductivity. The p-type conductivity of Cr₂O₃ has been applied to produce

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† Electronic supplementary information (ESI) available: An image of the crystal structure, UV/visible absorption spectra, STEM-EDX images and PEC properties. See DOI: 10.1039/C9DT04296B



hole transfer layers inserted between a back contact and a semiconductor photoabsorber, such as CuO ²⁶ or Fe_2O_3 .³⁹ Moreover, the conduction band minimum of Cr_2O_3 is located at a much more negative level than the redox potential for water reduction.^{40,41} Even so, to the best of our knowledge, there have been no reports of the application of Cr_2O_3 as a photoabsorber in photocathodes for H_2 generation. This may be due to the ready formation of various defects and vacancies in Cr_2O_3 . Defects, vacancies and surface dangling bonds form states within the bandgap that can serve as recombination centres for photogenerated charges (Scheme 1a). In addition, a high density of bandgap states will transition Cr_2O_3 into a conductor.^{42,43} However, if photogenerated charges can be prevented from transferring to such bandgap states, it may be possible to apply Cr_2O_3 as a photocathode for H_2 evolution. Furthermore, since band gap narrowing *via* doping with specific ions has already been realized,^{44–47} Cr_2O_3 also has the potential to become a visible-light responsive photocathode.

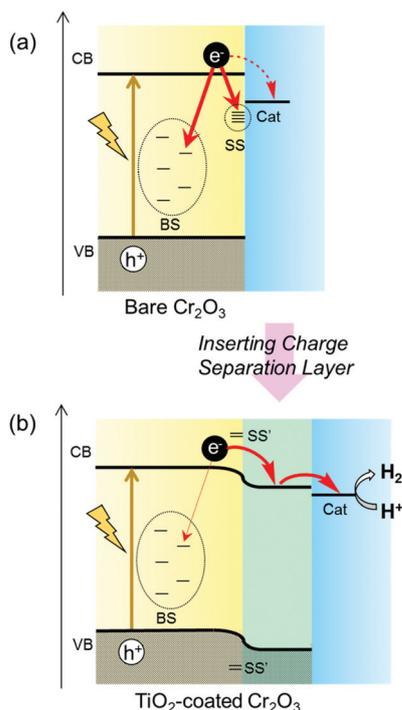
Recently, it has been reported that various benefits can be obtained by the surface treatment of photoelectrodes.^{48,49} Specifically, such processing can protect the unstable surface of the photoabsorber from photocorrosion by preventing contact with the electrolyte. Furthermore, this treatment can passivate surface states. When the atomic orbitals of surface states couple with those of deposits, the coupled surface states can be moved away from the band gap because of the for-

mation of bonding and antibonding states⁴⁹ (shown as SS' in Scheme 1b). Moreover, the formation of a heterojunction by surface treatment is expected to modify band bending due to the diffusion of charge carriers between two semiconductors having different Fermi levels. This modification should enhance both photogenerated electron–hole separation and carrier transfer. Therefore, in the present work, the Cr_2O_3 surface was modified with n-type TiO_2 to ensure both charge separation and passivation. Subsequently, Pt particles were deposited on the TiO_2 layer to catalyse H_2 production⁵⁰ (Scheme 1b).

Results and discussion

Structural and optical properties of Cr_2O_3

Cr_2O_3 electrodes were prepared by the radio frequency (RF) reactive magnetron sputtering of Cr targets in an Ar/O_2 (9 : 1) plasma to give a thickness of 60 nm on a transparent conductive oxide (TCO; SnO_2 -coated indium tin oxide (ITO)), followed by annealing at 823 K. As shown in Fig. 1 (blue line), the X-ray diffraction (XRD) patterns of the resulting Cr_2O_3 on TCO exhibit peaks at 33.8° , 36.4° and 55.3° , which are assignable to the (104), (110) and (116) diffractions of Cr_2O_3 (eskolaite), respectively. All other peaks are attributed to SnO_2 or ITO, while there are no peaks due to impurities. The X-ray photoelectron spectroscopy (XPS) data obtained from the resulting Cr_2O_3 layer are presented in Fig. 2a, and are similar to those produced by the reference Cr_2O_3 (Fig. 2c). The peaks at binding energies of 576.5 and 586.1 eV are ascribed to the Cr $2p_{3/2}$ and $2p_{1/2}$ signals of Cr(III) oxide, respectively. Although the Cr $2p_{3/2}$ peak includes a shoulder, this peak shape is typical of Cr_2O_3 and is attributed to the multiplet structure of Cr(III) oxide.⁵¹ Pt was subsequently deposited on the electrodes as a co-catalyst for H_2 evolution from water by sputtering, to a nominal thickness of 1 nm. Scanning transmission electron microscopy (STEM) combined with energy-dispersive X-ray spectroscopy (EDX) for elemental mapping was used to assess cross-sections of the Pt/ Cr_2O_3 sample, as shown in Fig. 3a. In



Scheme 1 Schematic diagram of the expected electron transfer process in (a) bare Cr_2O_3 and (b) TiO_2 -coated Cr_2O_3 photocathodes. BS and SS indicate bulk states and surface states, respectively. The former are generated by defects and vacancies, while the latter result from surface dangling bonds.

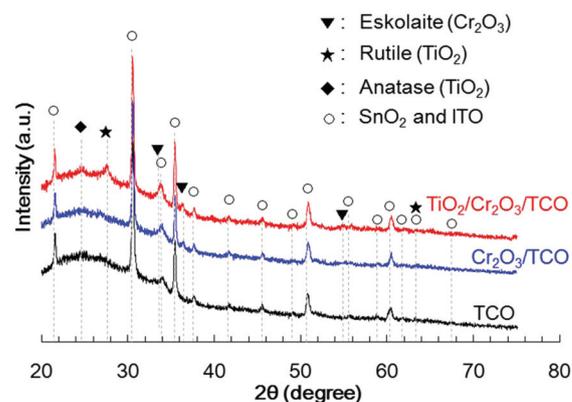


Fig. 1 XRD patterns obtained from TCO (black line), Cr_2O_3 /TCO (blue line) and TiO_2 / Cr_2O_3 /TCO (red line).



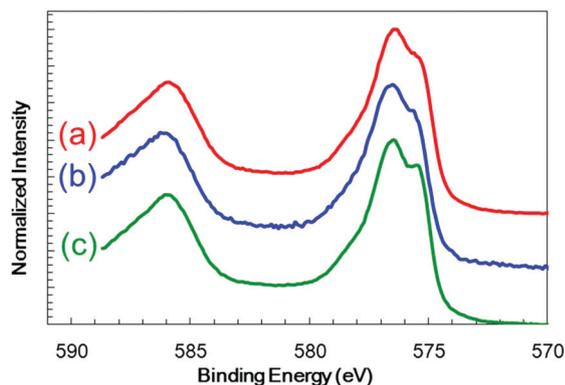


Fig. 2 XPS spectra obtained from a Cr₂O₃ electrode (a) before and (b) after photoelectrolysis for 1 h and from (c) the reference Cr₂O₃.

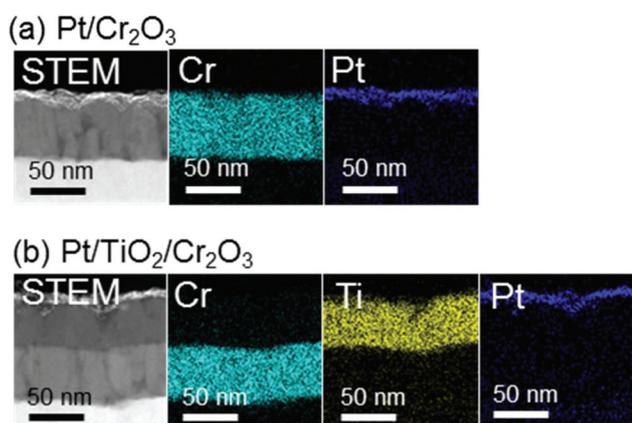


Fig. 3 Cross-sectional STEM images and STEM-EDX elemental maps for (a) Pt/Cr₂O₃ and (b) Pt/TiO₂/Cr₂O₃.

both electrodes, a flat Cr₂O₃ layer was observed on the TCO film.

UV/visible absorption spectra were acquired by a transmission method for this electrode and are shown in Fig. 4 as a blue line. Although an interference spectrum was included in the original spectrum, the absorption edge is evident at 410 nm, in agreement with a Cr₂O₃ reference bulk powder (Fig. S2†) and previous reports^{52–54}

PEC properties of the Cr₂O₃ electrodes

Fig. 5a and b present the current–potential curves obtained from the Cr₂O₃-based electrodes and from a TiO₂ electrode in a 0.5 M Na₂CO₃:NaHCO₃ (1:1) buffer electrolyte (pH 9.7) under chopped AM 1.5 irradiation (1 Sun; 100 mW cm⁻²), with irradiation of the deposited semiconductor side. The bare Cr₂O₃ electrode (Fig. 5a, black line) generated a cathodic photocurrent in the negative voltage range below +1.1 V vs. RHE, although it also produced a dark current. The photocurrent density of this sample at +0.1 V, after removing the dark current density from the current density produced during light irradiation, was $-15 \mu\text{A cm}^{-2}$. This cathodic photocurrent indicates that the Cr₂O₃ electrode exhibited p-type conductivity

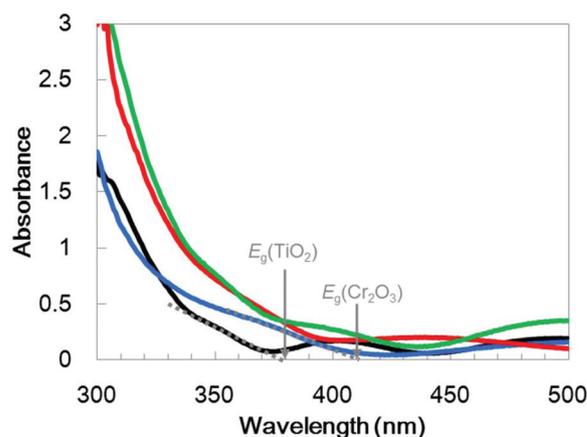


Fig. 4 UV/visible absorption spectra for TiO₂/Cr₂O₃ (red line), Cr₂O₃ (blue line) and TiO₂ (black line) on TCO and the sum of the Cr₂O₃ and TiO₂ spectra (green line). E_g indicates the absorption edges for the band-gap excitation.

and could possibly be utilized as a photocathode for water splitting. However, a large cathodic dark current was apparent between +1.1 and +1.5 V, including a peak of $-240 \mu\text{A cm}^{-2}$ at +1.08 V. This significant dark current indicates that electrons leaked to defect states and surface states in the band gap. Upon applying a constant potential of +0.11 V vs. RHE, the dark current was decreased due to electron charging, while the cathodic photocurrent was observed to decay from 8.1 to $0.9 \mu\text{A cm}^{-2}$ during 1 sun irradiation for 3 min (Fig. 5c, black line). After 1 h of photoelectrolysis, the dark current was slightly increased but the difference in the photocurrent in the current–potential curve was negligible (Fig. S3†). In addition, when the chemical state of the Cr₂O₃ surface was examined by XPS, there were almost no differences before and after the photoelectrolysis (Fig. 2a and b, respectively), suggesting no self-reduction of Cr₂O₃ by the photogenerated electrons. Therefore, Cr₂O₃ itself was stable during photoelectrolysis but photoelectrons could be trapped in defects and surface states, resulting in photocurrent decay.

Pt was deposited on Cr₂O₃ as a co-catalyst for H₂ evolution, by sputtering. The current–potential curve produced by the resulting Pt/Cr₂O₃ exhibited a new and significant dark current between 0.5 and 0.9 V, including a peak of $-700 \mu\text{A cm}^{-2}$ at +0.63 V. Moreover, the photocurrent was negligible after Pt loading (Fig. 5a, green line). Although the dark current was reduced by applying a constant potential of +0.11 V vs. RHE under dark conditions, the photocurrent was smaller than that obtained from pure Cr₂O₃ (Fig. 5c, green line). The large dark current in the current–potential curve suggests that new band-gap states were generated within the bandgap of Cr₂O₃ following Pt deposition. The appearance of these deleterious band gap states can possibly be ascribed to the coupling of the atomic orbitals of dangling atoms at the Cr₂O₃ surface with those of deposited Pt atoms. The resulting bonding and anti-bonding orbitals would generate deleterious band-gap states,⁵⁵ leading to electron leakage and recombination of photo-gener-



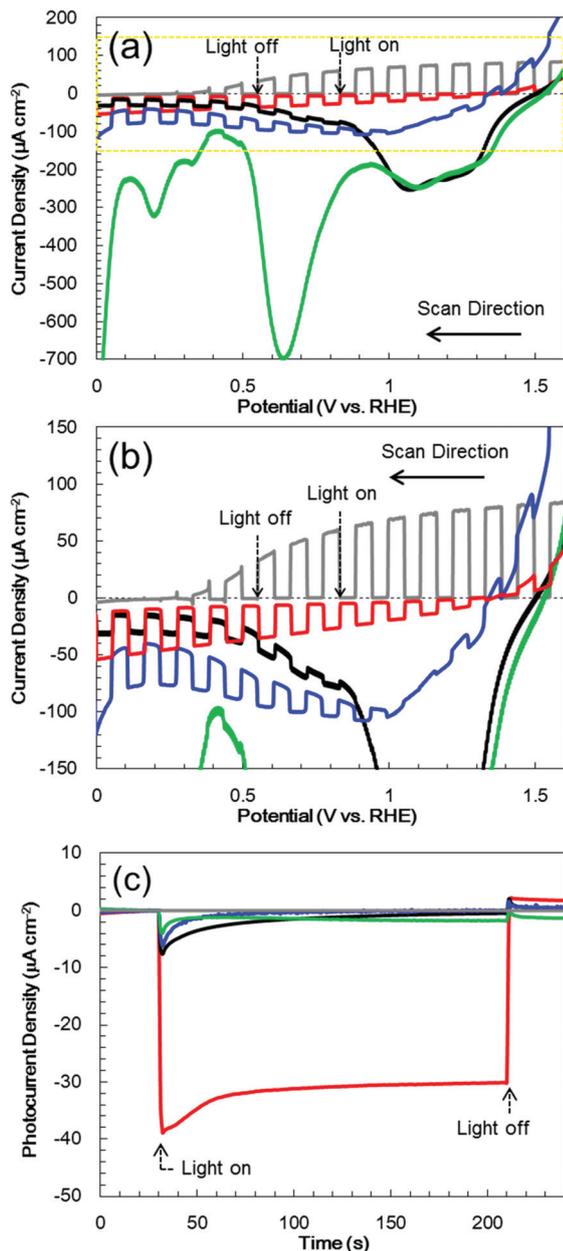


Fig. 5 (a) Photocurrent–potential characteristics under chopped light irradiation, (b) enlargement of the area in the yellow dotted line in (a), and (c) photocurrent transients at +0.11 V vs. RHE in a 0.5 M $\text{Na}_2\text{CO}_3:\text{NaHCO}_3$ (1:1) buffer electrolyte (pH 9.7) under one sun (100 mW cm^{-2} , AM 1.5) illumination for the bare Cr_2O_3 (black line), Pt/ Cr_2O_3 (green line), TiO_2 (grey line), $\text{TiO}_2/\text{Cr}_2\text{O}_3$ (blue line) and Pt/ $\text{TiO}_2/\text{Cr}_2\text{O}_3$ (red line).

ated electrons and holes. As a result, the addition of Pt nanoparticles to the Cr_2O_3 surface diminished the photocurrent without acting as a co-catalyst for H_2 evolution.

Coating with TiO_2 as passivation and charge-separation layer

TiO_2 overlayers with thicknesses ranging from 10 to 120 nm were subsequently deposited by sputtering a TiO_2 target in an Ar/O_2 (4:1) plasma onto Cr_2O_3 . This was followed by post-

annealing at 748 K. The XRD patterns generated by $\text{TiO}_2/\text{Cr}_2\text{O}_3/\text{TCO}$ (Fig. 1, red line) exhibit peaks attributable to anatase (101), and rutile (110) and (310) diffractions at 24.8° , 27.8° and 63.4° , respectively. In the STEM-EDX images for the Pt/ $\text{TiO}_2/\text{Cr}_2\text{O}_3$ electrode (Fig. 3b), the TiO_2 layer completely covered the Cr_2O_3 surface, preventing contact with both the Pt and the electrolyte. The UV/visible absorption spectrum of the $\text{TiO}_2/\text{Cr}_2\text{O}_3$ electrode was essentially consistent with the sum of the spectra of the TiO_2 and Cr_2O_3 electrodes (Fig. 4, grey dotted line). The absorption edge of the TiO_2 electrode (Fig. 4, black line) was observed at 380 nm, consistent with a TiO_2 reference bulk powder (Fig. S2†) and the previous report.⁵⁶ These data indicate that simulated solar radiation imparted to the TiO_2 side should be absorbed by both TiO_2 and Cr_2O_3 .

In photocurrent-potential curves, the surface of Cr_2O_3 was subsequently coated with a 60 nm thick TiO_2 layer. The resulting $\text{TiO}_2/\text{Cr}_2\text{O}_3$ photocathode generated a cathodic photocurrent below +1.1 V (Fig. 5b, blue line). The photocurrent at +0.1 V was $-35 \mu\text{A cm}^{-2}$, and so exceeded that obtained from the bare Cr_2O_3 ($-15 \mu\text{A cm}^{-2}$). Although the TiO_2 electrode (without Cr_2O_3) exhibited an anodic photocurrent at values more positive than +0.3 V (Fig. 5b, grey line), the photocurrent of $\text{TiO}_2/\text{Cr}_2\text{O}_3$ cannot be explained as a combination of those for TiO_2 and Cr_2O_3 . The cathodic photocurrent produced by $\text{TiO}_2/\text{Cr}_2\text{O}_3$ was also larger than that for Cr_2O_3 . The anodic photocurrent produced by $\text{TiO}_2/\text{Cr}_2\text{O}_3$ was smaller and appeared at a more positive potential than that for TiO_2 . This positive shift of the onset potential for the anodic photocurrent can be explained by the formation of a p–n junction. The Fermi level of the n-type TiO_2 would be expected to be shifted positively following the formation of a junction with the p-type Cr_2O_3 because the Fermi levels of TiO_2 and Cr_2O_3 would be equalized by carrier diffusion. Moreover, the Cr_2O_3 layer would be expected to inhibit photogenerated electron transfer from TiO_2 to TCO, because the conduction band minimum of Cr_2O_3 is higher than that of TiO_2 . Hence, $\text{TiO}_2/\text{Cr}_2\text{O}_3$ generated a cathodic photocurrent. Moreover, the dark current produced by $\text{TiO}_2/\text{Cr}_2\text{O}_3$ decreased, from -240 to $-100 \mu\text{A cm}^{-2}$, by the TiO_2 coating, indicating a decrease in the electron transfer to the band-gap states. Since the recombination of photo-generated electrons and holes at the band-gap states was suppressed, the photocurrent was enhanced by the TiO_2 coating. However, at a constant potential of +0.11 V vs. RHE, the photocurrent also decayed after irradiation (Fig. 5c, blue line).

Following the above trials, Pt nanoparticles were deposited on $\text{TiO}_2/\text{Cr}_2\text{O}_3$ as a co-catalyst. The resulting Pt/ $\text{TiO}_2/\text{Cr}_2\text{O}_3$ specimen exhibited a cathodic photocurrent below +1.4 V (Fig. 5b, red line), with an onset potential more positive than that of $\text{TiO}_2/\text{Cr}_2\text{O}_3$. In addition, the dark current was diminished by the Pt loading. At a constant potential of +0.11 V vs. RHE, the cathodic photocurrent was larger and more stable than those of the other electrodes (Fig. 5c, red line). Thus, in contrast to the results obtained from the direct deposition of Pt on Cr_2O_3 (Fig. 5a, green and black lines), there was no decrease in the photocurrent and an increase in the dark current. These results demonstrate that the TiO_2 overlayer pas-



sivated the Cr_2O_3 surface states to prevent the generation of new band-gap states *via* the coupling of Cr_2O_3 and Pt. As a result, the dark current due to electron leakage to band-gap states was decreased. The photocurrent was also enhanced, because the photoelectrons generated in Cr_2O_3 were able to flow to the Pt nanoparticles *via* the TiO_2 layer, after which they were consumed for H_2 evolution.

Effect of the TiO_2 overlayer thickness

In order to optimize the thickness of the TiO_2 layer on the Cr_2O_3 electrode, Pt/ TiO_2 / Cr_2O_3 electrodes having TiO_2 thicknesses of 10, 30, 60, 80 and 120 nm were prepared by adjusting the deposition time in the TiO_2 sputtering. UV/visible absorption spectra acquired by a transmission method for these electrodes are shown in Fig. S4.† It can be seen that the absorptions in $\lambda < 380$ nm increase in proportion to the thickness of TiO_2 because the absorption edge of TiO_2 is 380 nm (Fig. 4). The photocurrent transients of these specimens were assessed under irradiation with simulated solar radiation (100 mW cm^{-2} , AM 1.5) for 3 min while applying a potential of +0.11 V *vs.* RHE (Fig. 6). Although the photoelectrode with a 10 nm-thick TiO_2 layer underwent a rapid decay, the other photoelectrodes showed stable photocurrents. The highest photocurrent value was observed for a thickness of 30 nm, and the photocurrents decreased with increases in the TiO_2 thickness.

To investigate the effect of the TiO_2 thickness in more detail, incident photon to current conversion efficiency (IPCE) spectra were obtained by irradiation with monochromatic light using band pass filters. Fig. 7 shows the IPCE spectra acquired from Pt/ TiO_2 / Cr_2O_3 specimens with various TiO_2 layer thicknesses. The IPCE spectra of the electrodes with more than 30 nm thick TiO_2 layers increased in intensity at 400 nm, showing a good correlation with the absorption spectrum of Cr_2O_3 (Fig. 4, blue line). Although the IPCE values at wave-

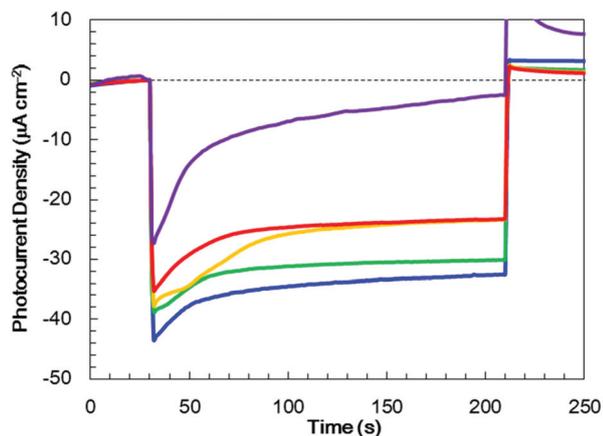


Fig. 6 Photocurrent transients for Pt/ TiO_2 / Cr_2O_3 specimens with various TiO_2 layer thicknesses (10 nm; purple line, 30 nm; blue line, 60 nm; green line, 80 nm; yellow line and 120 nm; red line) at +0.11 V *vs.* RHE in a 0.5 M Na_2CO_3 : NaHCO_3 (1:1) buffer electrolyte (pH 9.7) under one sun (100 mW cm^{-2} , AM 1.5) illumination. The thickness of the Cr_2O_3 layer was fixed at 60 nm.

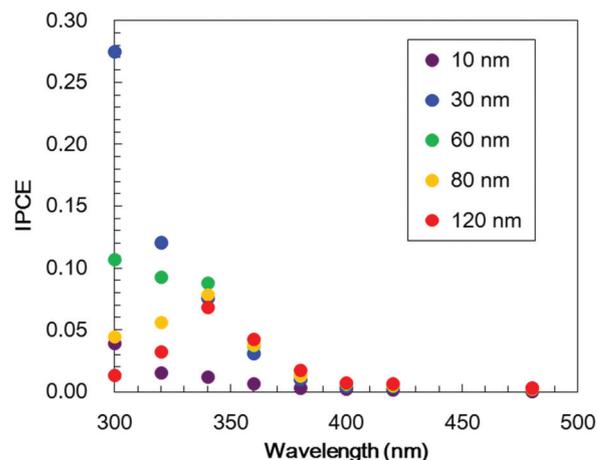


Fig. 7 IPCE spectra obtained from Pt/ TiO_2 / Cr_2O_3 electrodes at various wavelengths using different TiO_2 layer thicknesses (10 nm; purple, 30 nm; blue, 60 nm; green, 80 nm; yellow and 120 nm; red) at +0.11 V *vs.* RHE in a 0.5 M Na_2CO_3 : NaHCO_3 (1:1) buffer electrolyte (pH 9.7) under monochromatic illumination. The thickness of the Cr_2O_3 layer was fixed at 60 nm.

lengths longer than 340 nm were similar for each electrode, the IPCE values below this point became smaller as the TiO_2 layer thickness exceeded 30 nm. These results suggest that light absorption by Cr_2O_3 was inhibited by the TiO_2 overlayer. As an example, as shown by the black line in the absorption spectrum in Fig. 4, a TiO_2 layer having a thickness of 60 nm had an absorbance greater than 1 below 320 nm. Thus, more than 90% of the light below this wavelength could not reach Cr_2O_3 . The observation that the IPCE spectra were affected by the TiO_2 thickness clearly indicates that the cathodic photocurrent originated from the photoexcitation of Cr_2O_3 , and that light absorption by TiO_2 decreased the photocurrent. However, the IPCE values for the photoelectrode with a 10 nm-thick TiO_2 layer were lower than those of the other electrodes at almost all wavelengths. Although a 10 nm-thick TiO_2 layer covered the surface of Cr_2O_3 (Fig. S5†), the thickness of TiO_2 may be insufficient for the formation of a p-n junction structure. As a result of the trade-off between the light filtering effect of TiO_2 and the formation of a charge separation structure, the optimal TiO_2 film thickness with regard to generating a large photocurrent is believed to be 30 nm.

PEC water splitting

To confirm PEC H_2 evolution, the gaseous products were analysed by gas chromatography after 1 h of irradiation with AM 1.5 light while applying a bias of +0.11 V *vs.* RHE (Table 1). The Pt/ TiO_2 (60 nm)/ Cr_2O_3 specimen generated a stable photocurrent over this time span (Fig. S6†), with a charge flow of 110 mC cm^{-2} . During this trial, 510 nmol cm^{-2} of H_2 was produced (entry 1 in Table 1), equivalent to 89% faradaic efficiency. Control experiments without Pt or without both Pt and TiO_2 did not produce H_2 (entries 2 and 4, Table 1), while a test without TiO_2 gave only a very small amount of this product (entry 3 in Table 1). Thus, even if light is absorbed by Cr_2O_3 ,



Table 1 The amounts of charge and H₂ obtained from a 1 h PEC reaction^a

Entry	Photoelectrode	Charge mC cm ⁻²	H ₂ nmol cm ⁻²
1 ^b	Pt/TiO ₂ /Cr ₂ O ₃	109.8	510
2	Cr ₂ O ₃	13.3	nd
3 ^b	TiO ₂ /Cr ₂ O ₃	1.7	nd
4	Pt/Cr ₂ O ₃	10.5	2

^a AM 1.5 light irradiation with applying bias at +0.11 V vs. RHE. ^b The thickness of TiO₂ layers was 60 nm.

PEC water splitting does not proceed without TiO₂ and Pt. Therefore, it is considered that PEC water splitting progresses due to the generation of photoexcited electrons in Cr₂O₃ in response to light absorption. These electrons are transferred to the Pt *via* the TiO₂ charge-separation layer.

Conclusion

The generation of H₂ *via* solar PEC water-splitting by Cr₂O₃ was successfully realized following surface modification of the oxide. Bare Cr₂O₃ was found to produce a significant dark current but a minimal photocurrent. The addition of Pt further deteriorated the performance of this material. Therefore, neither bare Cr₂O₃ nor Pt/Cr₂O₃ provided H₂ by PEC water-splitting. Applying a TiO₂ overlayer with subsequent Pt deposition to fabricate a Pt/TiO₂/Cr₂O₃ electrode resulted in an enhanced photocurrent and realized solar PEC H₂ generation. The dark current almost disappeared following this surface treatment, suggesting that TiO₂ passivated the Cr₂O₃ surface. In this layered electrode, a photocathode current was induced in response to light absorption by Cr₂O₃, while the TiO₂ layer acted as a filter to inhibit this light absorption. Nevertheless, since the photocurrent was enhanced by the presence of a TiO₂ coating, it is evident that TiO₂ also acted as a charge separation layer. Thus, within the Pt/TiO₂/Cr₂O₃ electrode, TiO₂ accumulated photoexcited electrons generated by light absorption by Cr₂O₃ while blocking undesired interactions between Cr₂O₃ and Pt. As a result, H₂ generation was realized through water-splitting. In future, the use of this TiO₂ layering technique with a doped Cr₂O₃ could lead to practical visible-light-driven PEC water-splitting systems.

Experimental section

Materials

Tin(IV) oxide/ITO double-layered TCO coated glass (Geomatec Co., Ltd) was used as substrates for the preparation of the electrodes. Commercially available Cr, TiO₂ and Pt sputtering targets and TiO₂ powder were obtained from the Kojundo Chemical Laboratory Co. A reference sample of Cr₂O₃ for the XPS analysis and UV/Visible diffuse reflectance spectroscopy was obtained from Nacalai Tesque, Inc.

Preparation

A Cr₂O₃ film was deposited by RF reactive magnetron sputtering of a Cr target with an Ar/O₂ (9 : 1 v/v) plasma. After deposition, the electrodes were annealed at 823 K under an N₂/O₂ (4 : 1 v/v) gas flow for 2 h. Subsequently, a TiO₂ layer was deposited on the surface of the Cr₂O₃ thin film by RF reactive magnetron sputtering of a TiO₂ target with an Ar/O₂ (4 : 1 v/v) plasma. The deposited electrodes were then annealed at 823 K under an O₂ gas flow for 2 h. The Pt cocatalyst was loaded onto either TiO₂-coated or bare Cr₂O₃ specimens by RF magnetron sputtering. The amount of Pt applied was adjusted to obtain an average thickness of 1 nm.

Characterization

The crystal structures of the films were analysed using XRD (Ultima IV, Rigaku Co.) with Cu K α radiation, while the optical properties of the films were assessed by using UV/visible absorption spectroscopy (UV-3600, Shimadzu Co.). STEM together with EDX was performed using a JEM-2100F microscope (JEOL Co.). Samples for these observations were cut using a focused ion beam (FIB; NB5000, Hitachi High-Tech Co.). Prior to cutting the samples, the film surfaces were coated with carbon and Pt layers to protect the specimens from the FIB. XPS data were acquired using a PHI5000 VersaProbe (Ulvac-Phi Co.) spectrometer.

PEC measurements

The photocurrent density measurements were performed in a three-electrode configuration consisting of a silver–silver chloride (Ag/AgCl) electrode, a platinum wire electrode and a prepared photocathode as the working, reference and counter electrodes, respectively. In these trials, the exposed TCO layer of the photocathode was covered with silicone rubber and all potentials were converted to the RHE reference scale using the equation E (vs. RHE) = E (vs. Ag/AgCl) + 0.20 V + 0.059 V \times pH. A sealed Pyrex® glass cell was employed as the reactor and Ar-saturated aqueous 0.5 M NaHCO₃ : Na₂CO₃ (1 : 1; pH 9.7) solutions were employed as the electrolytes. An electrochemical analyser (ALS612E and ALSCHI-614C, ALS Co., Ltd) supplied the potential and frequency to the electrode. The electrode was irradiated with a light intensity equivalent to one sun (AM 1.5; 100 mW cm⁻²) using a solar simulator (HAL-320, Asahi Spectra Co.). Prior to exposure, the light intensity was adjusted with a CS-20 instrument (Asahi Spectra Co.). The sample irradiation was limited to an area of 10 \times 10 mm *via* a slit. Linear sweep voltammetry was conducted at a scan rate of 50 mV s⁻¹ under chopped light irradiation. The photocurrent densities in the photocurrent transient curves were converted by subtracting the dark current density values determined just prior to photo-irradiation from the current densities. The PEC water-splitting reaction, in conjunction with the application of an electrical bias, was conducted while measuring the photocurrent generated under continuous irradiation at a fixed electrode potential of -0.5 V vs. Ag/AgCl. After incubation for 30 min to allow the gaseous reaction products to equilibrate



between the liquid and gas phases, these products were analysed by gas chromatography (GC-2014, Shimadzu Co.), employing a thermal conductivity detector and an active carbon column (Shincarbon ST, Shinwa Chemical Industries Co.) with Ar as the carrier gas. IPCE spectra were acquired under monochromatic light generated by a 300 W xenon lamp (MAX-303, Asahi Spectra Co.), using band-pass filters to obtain specific wavelengths.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- G. Glenk and S. Reichelstein, *Nat. Energy*, 2019, **4**, 216–222.
- D. Parra, L. Valverde, F. J. Pino and M. K. Patel, *Renewable Sustainable Energy Rev.*, 2019, **101**, 279–294.
- I. Staffell, D. Scamman, A. V. Abad, P. Balcombe, P. E. Dodds, P. Ekins, N. Shah and K. R. Ward, *Energy Environ. Sci.*, 2019, **12**, 463–491.
- T. Yao, X. An, H. Han, J. Q. Chen and C. Li, *Adv. Energy Mater.*, 2018, **8**, 1800210.
- C. Jiang, S. J. A. Moniz, A. Wang, T. Zhang and J. Tang, *Chem. Soc. Rev.*, 2017, **46**, 4645–4660.
- J. W. Ager, M. R. Shaner, K. A. Walczak, I. D. Sharp and S. Ardo, *Energy Environ. Sci.*, 2015, **8**, 2811–2824.
- A. Fujishima and K. Honda, *Nature*, 1972, **238**, 37–38.
- A. Kudo and Y. Miseki, *Chem. Soc. Rev.*, 2009, **38**, 253–278.
- R. Abe, *J. Photochem. Photobiol., C*, 2010, **11**, 179–209.
- X. B. Chen, S. H. Shen, L. J. Guo and S. S. Mao, *Chem. Rev.*, 2010, **110**, 6503–6570.
- K. Maeda and K. Domen, *J. Phys. Chem. Lett.*, 2010, **1**, 2655–2661.
- K. Maeda, *J. Photochem. Photobiol., C*, 2011, **12**, 237–268.
- F. E. Osterloh, *Chem. Soc. Rev.*, 2013, **42**, 2294–2320.
- T. Hisatomi, J. Kubota and K. Domen, *Chem. Soc. Rev.*, 2014, **43**, 7520–7535.
- A. Miyoshi, S. Nishioka and K. Maeda, *Chem. – Eur. J.*, 2018, **24**, 18204–18219.
- Y. Miseki and K. Sayama, *Adv. Energy Mater.*, 2019, **9**, 1801294.
- T. Morikawa, S. Sato, K. Sekizawa, T. Arai and T. M. Suzuki, *ChemSusChem*, 2019, **12**, 1807–1824.
- H. Yoneyama, H. Sakamoto and H. Tamura, *Electrochim. Acta*, 1975, **20**, 341–345.
- K. Ohashi, J. McCann and J. O. M. Bockris, *Nature*, 1977, **266**, 610–611.
- J. E. Turner, M. Hendewerk, J. Parmeter, D. Neiman and G. A. Somorjai, *J. Electrochem. Soc.*, 1984, **131**, 1777–1783.
- M. S. Prévot and K. Sivula, *J. Phys. Chem. C*, 2013, **117**, 17879–17893.
- K. Zhang, M. Ma, P. Li, D. H. Wang and J. H. Park, *Adv. Energy Mater.*, 2016, **6**, 1600602.
- A. Liu, H. Zhu and Y.-Y. Noh, *Mater. Sci. Eng., R*, 2019, **135**, 85–100.
- Z. Wang, P. K. Nayak, J. A. Caraveo-Frescas and H. N. Alshareef, *Adv. Mater.*, 2016, **28**, 3831–3892.
- A. Paracchino, V. Laporte, K. Sivula, M. Gratzel and E. Thimsen, *Nat. Mater.*, 2011, **10**, 456–461.
- S. N. F. M. Nasir, M. K. N. Yahya, N. W. M. Sopian, N. A. Ludin, M. A. Ibrahim, K. Sopian and M. A. M. Teridi, *RSC Adv.*, 2016, **6**, 56885–56891.
- S. Ida, K. Yamada, T. Matsunaga, H. Hagiwara, Y. Matsumoto and T. Ishihara, *J. Am. Chem. Soc.*, 2010, **132**, 17343–17345.
- K. Sekizawa, T. Nonaka, T. Arai and T. Morikawa, *ACS Appl. Mater. Interfaces*, 2014, **6**, 10969–10973.
- T. Morikawa, T. Arai and T. Motohiro, *Appl. Phys. Express*, 2013, **6**, 041201.
- A. Holt and P. Kofstad, *Solid State Ionics*, 1997, **100**, 201–209.
- H. Cao, X. Qiu, Y. Liang, M. Zhao and Q. Zhu, *Appl. Phys. Lett.*, 2006, **88**, 241112.
- L. W. Finger and R. M. Hazen, *J. Appl. Phys.*, 1980, **51**, 5362–5367.
- L. Pauling and S. B. Hendricks, *J. Am. Chem. Soc.*, 1925, **47**, 781–790.
- A. Holt and P. Kofstad, *Solid State Ionics*, 1999, **117**, 21–25.
- S. C. Tsai, A. M. Huntz and C. Dolin, *Mater. Sci. Eng., A*, 1996, **212**, 6–13.
- A. Holt and P. Kofstad, *Solid State Ionics*, 1994, **69**, 127–136.
- B. Medasani, M. L. Sushko, K. M. Rosso, D. K. Schreiber and S. M. Bruemmer, *J. Phys. Chem. C*, 2018, **122**, 12984–12993.
- F. Lebreau, M. M. Islam, B. Diawara and P. Marcus, *J. Phys. Chem. C*, 2014, **118**, 18133–18145.
- K. Sekizawa, K. Oh-ishi, K. Kataoka, T. Arai, T. M. Suzuki and T. Morikawa, *J. Mater. Chem. A*, 2017, **5**, 6483–6493.
- M. T. Greiner, M. G. Helander, W.-M. Tang, Z.-B. Wang, J. Qiu and Z.-H. Lu, *Nat. Mater.*, 2012, **11**, 76–81.
- Y. Xu and M. A. A. Schoonen, *Am. Mineral.*, 2000, **85**, 543–556.
- L. Farrell, K. Fleischer, D. Caffrey, D. Mullarkey, E. Norton and I. V. Shvets, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2015, **91**, 125202.
- E. Arca, K. Fleischer and I. V. Shvets, *Appl. Phys. Lett.*, 2011, **99**, 111910.
- H. Wang, Y. Zhang, L. Zhang, Y. Guo, S. Liu, F. Gao, Y. Han, G. Feng, X. Liang and L. Ge, *RSC Adv.*, 2016, **6**, 84871–84881.
- S. Goel, A. Kumar, J. K. Quamara and J. Kumar, *Adv. Sci. Lett.*, 2014, **20**, 1562–1566.
- J. J. Carey, M. Legesse and M. Nolan, *J. Phys. Chem. C*, 2016, **120**, 19160–19174.



- 47 E. Arca, K. Fleischer, S. A. Krasnikov and I. Shvets, *J. Phys. Chem. C*, 2013, **117**, 21901–21907.
- 48 D. Bae, B. Seger, P. C. K. Vesborg, O. Hansen and I. Chorkendorff, *Chem. Soc. Rev.*, 2017, **46**, 1933–1954.
- 49 N. Guijarro, M. S. Prevot and K. Sivula, *Phys. Chem. Chem. Phys.*, 2015, **17**, 15655–15674.
- 50 B. Xu, L. Tian, A. S. Etman, J. Sun and H. Tian, *Nano Energy*, 2019, **55**, 59–64.
- 51 M. C. Biesinger, B. P. Payne, A. P. Grosvenor, L. W. M. Lau, A. R. Gerson and R. S. C. Smart, *Appl. Surf. Sci.*, 2011, **257**, 2717–2730.
- 52 R. Cheng, B. Xu, C. N. Borca, A. Sokolov, C.-S. Yang, L. Yuan, S.-H. Liou, B. Doudin and P. A. Dowben, *Appl. Phys. Lett.*, 2001, **79**, 3122–3124.
- 53 M. M. Abdullah, F. M. Rajab and S. M. Al-Abbas, *AIP Adv.*, 2014, **4**, 027121.
- 54 C.-S. Cheng, H. Gomi and H. Sakata, *Phys. Status Solidi A*, 1996, **155**, 417–425.
- 55 R. Cohen, L. Kronik, A. Shanzer, D. Cahen, A. Liu, Y. Rosenwaks, J. K. Lorenz and A. B. Ellis, *J. Am. Chem. Soc.*, 1999, **121**, 10545–10553.
- 56 Y. Li, L. Zhang, W. Wu, P. Dai, X. Yu, M. Wu and G. Li, *Nanoscale Res. Lett.*, 2014, **9**, 270.

