

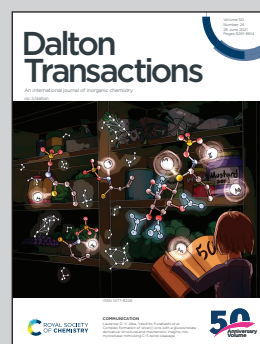


Showcasing research from Professor Ola Nilsen's laboratory, Department of Chemistry, University of Oslo, Norway.

Quinizarin: a large aromatic molecule well suited for atomic layer deposition

The optically active hydroxyanthraquinones have proven well suited as volatile and reactive molecules for the formation of hybrid films by atomic/molecular layer deposition. We use it as a tool for designing optical coatings by exploiting extended aromaticity. One of these molecules, the orange dye molecule quinizarin, is combined with trimethylaluminium to produce a deep pink coating with strong optical absorption. Its low deposition temperature of 175 °C allows coating of even fragile and temperature-sensitive organic substrates, illustrated by coating a self-dead spider.

As featured in:



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8307Quinizarin: a large aromatic molecule well suited
for atomic layer deposition†Per-Anders Hansen  and Ola Nilsen  *

Atomic layer deposition (ALD) is a remarkable synthesis tool due to the vast array of materials that can be deposited and the complexity of structures that can be designed. The low-temperature layer-by-layer approach even allows organic and inorganic components to be combined as hybrid or composite materials. The technique is then called molecular layer deposition (MLD). This opens the door for deposition of advanced optical materials using highly absorbing aromatic molecules. Unfortunately, most large aromatic molecules are difficult to sublime or have insufficient reactivity. This is a major barrier for ALD when designing with the use of organic components for dye-sensitized solar cells, luminescence, visible light photochemistry, chemical sensors and organic electronics. In this work, we introduce a well-known orange dye molecule, quinizarin. This molecule has a large conjugated aromatic system with strong absorption of visible light and shows strong luminescence both in solutions and as a complex together with aluminium ions. Interestingly, quinizarin also shows surprisingly good properties for film deposition due to reactive –OH groups and low sublimation temperature (130 °C). Strongly coloured pink hybrid films were deposited with trimethylaluminium and quinizarin at 175 °C with a growth rate of 0.28 nm per cycle. These films were not luminescent although their optical absorption spectra are similar to those of the corresponding solution. An attempt was made to dilute quinizarin through partial replacement with pentaerythritol as a multilayer structure or simultaneous co-pulsing, although this also did not produce luminescent films. The low sublimation temperature, good reactivity and large conjugated system of quinizarin open the way for exploration of solid-state hybrid and organic films based on this molecule along many different technological pathways.

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Introduction

Atomic layer deposition (ALD) is a remarkable synthesis tool capable of producing a vast array of materials from inorganic,¹ quaternary² and even quinary compounds,³ nanostructures,⁴ organic–inorganic hybrids,⁵ crystalline MOF films⁶ and pure organic compounds.⁷ ALD is also often named molecular layer deposition (MLD) when organic molecules are introduced as a building block into organic–inorganic hybrid or purely organic film materials. The low temperature and layer-by-layer approach limits the interdiffusion between layers and enables a wide range of combinations of different material classes at the sub-nanometer scale. This gives rise to completely new possibilities in designing with combinations of material properties⁸ and to control optical and electronic interactions.⁹ While ALD normally coats all surfaces similarly, area-selective deposition is possible and is an expanding field.¹⁰

Due to this flexibility in material choices and design possibilities, it is not surprising that ALD has been explored for a wide range of advanced applications like optical temperature sensors,¹¹ oxide electronics,¹² and up¹³ and down¹⁴ conversions, in particular when the addition of organic building blocks opens up many new technological possibilities.¹⁵ However, the range of organic components to choose from as precursors in ALD has so far been limited to quite small molecules since they need to be delivered through the gas phase.

Reviews on ALD and MLD chemistries^{7,16} reveal that organic precursors are quite limited beyond a single aromatic ring or a short, conjugated chain. A few examples exist where larger conjugated systems like naphthalene and biphenyl⁶ are used, showing near-UV activity and even visible absorption.¹⁷ These optically absorbing dye precursors require sublimation temperatures of around 250 °C or above, approaching their thermal stability. This illustrates well the difficulty in balancing Vis and NIR active optical properties with ALD/MLD compatibility.

Quinizarin (1,4-dihydroxyanthraquinone) is a well-known dye molecule, but its excellent properties are less known as an ALD/MLD precursor. This molecule belongs to the anthraquinone class of organic dyes, which show a range of optical, elec-

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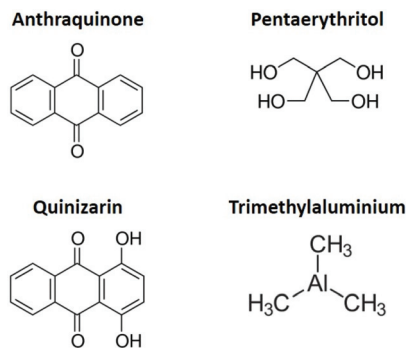


Fig. 1 Anthraquinone, the parent molecule of quinizarin, and the three precursor molecules used in this work.

tronic, photocatalytic¹⁸ and red-ox properties.¹⁹ Quinizarin shows both strong visible absorption, offering large possibilities for chemical modification, and, the key point in this work, exceptionally good properties as an ALD/MLD precursor: it sublimes well at 130 °C and has good reactivity with two hydroxyl groups, and it has a rather low cost. As a metal-organic hybrid, it is explored in supercapacitors as metal-organic frameworks²⁰ and dye sensitization of TiO₂²¹ and can couple to semiconducting nanoparticles to create new low-energy luminescent states.²² Anthraquinone, quinizarin and its derivatives have also been explored for a diverse range of uses such as eco-friendly replacement of cobalt in Li-ion batteries.^{23,24} In addition to red-ox activity, electronic properties and strong visible absorption, quinizarin and Al³⁺ ions form highly luminescent and strongly coloured complexes and polymeric structures in solutions.²⁵ Similar properties are observed for Ga³⁺ and In³⁺.²⁶ The flexibility in chemical modifications and suitability as an ALD/MLD precursor allow for many exciting new avenues to be explored in the design of hybrid, organic and nanocomposite materials.

In this work, we introduce quinizarin (H₂Qz) as an organic precursor in ALD/MLD and describe the deposition of Al₂Qz₃ aluminium-quinizarin films. The growth, composition and long-term stability are explored, alongside their optical properties. Finally, as Al₂Qz₃ complexes in solution show strong visible luminescence while their dried powder counterparts do not, we attempted a “solid-state dilution” by replacing quinizarin with pentaerythritol (H₄Pe), forming Al₄Pe₃. Al₄Pe₃ is optically passive and transparent, and its function is to increase the distance between the quinizarin molecules in the film material. With this in mind, the exploration of Al₄Pe₃ is limited to growth parameters and optical properties. The precursors used in this work and the parent anthraquinone is shown in Fig. 1.

Experimental

The films were deposited in an F-120 research-type ALD-reactor (ASM Microchemistry Ltd) at 175 °C. Trimethyl aluminium (TMA) (Witco GmbH, 98%), Al(acac)₃ (Sigma Aldrich, 99.999%), quinizarin (Sigma Aldrich, 96%) and pentaerythritol

(Sigma Aldrich, 98%) were used as precursors. TMA was introduced from an external bubbler kept at room temperature, while Al(acac)₃, quinizarin and H₄Pe were sublimed at 130, 130 and 160 °C, respectively. Pulse parameters (Al/purge/anion/purge) for TMA and H₂Qz cycles were 1/1.5/3/1.5 s, for TMA and H₄Pe were 1/1/5/1 s, and for Al(acac)₃ and H₂Qz were 2/1/2/1 s. For the Al(acac)₃ and H₂Qz depositions, both 175 and 225 °C were used as deposition temperatures. These parameters were used for the data presented in this work, while other parameters were explored.

p-Type Si(100) substrates cut into 0.5 × 4 cm² strips were used for all depositions. These were placed some 5 cm apart in the gas inlet and exhaust sides of the deposition chamber to monitor thickness gradients (see the photo in Fig. 6). The native oxide on the silicon substrates was not removed but was measured before deposition for every substrate. In addition, some depositions included polished steel substrates for reflection mode FTIR and amorphous silica substrates for UV-Vis measurements. All substrates were dry wiped, and dust was removed using pressurized air before deposition.

In situ quartz crystal microbalance (QCM) analyses were conducted using a 6 MHz AT-cut quartz crystal. The crystal was mounted in a home-made holder and was used to monitor the mass increase, which is proportional to the change in the frequency,²⁷ during the deposition to determine saturation conditions for pulse and purge parameters. The signal was recorded using a Colnatec Eon-LT system and processed by averaging over 16 consecutive ALD cycles. The temperature was stabilized for 90 minutes before any experiments were conducted to ensure a stable temperature and response from the QCM crystals.

The crystallinity of the samples was determined with a Bruker D8 Discovery X-ray diffractometer, using CuKα₁ radiation and a Ge(111) monochromator. Film thickness and refractive index were determined with a J. A. Woollam alpha-SE spectroscopic ellipsometer in the 380–900 nm range. The ellipsometry data were modelled using a Cauchy model. FTIR measurements were performed in the reflection mode at 75° incident angle using a Bruker Vertex 70 spectrometer equipped with a Pike VeeMAX III specular reflection accessory. All data were acquired using 64 scans in the wavenumber range of 4000–370 cm^{−1} and at 2 cm^{−1} resolution. A spectrum obtained from the uncoated electropolished steel substrate was used as the background during the measurements and subtracted from the spectra of the films. The FTIR spectra of the organic precursors quinizarin and pentaerythritol were obtained from the online database Wiley SpectraBase.^{28,29} UV-Vis transmittance and reflectance spectra were collected with a Shimadzu UV-3600 UV-Vis-NIR spectrophotometer. Absorbance spectra were calculated from transmittance and reflectance obtained while using an integrating sphere. Luminescence (PL) measurements were performed with a 280 nm diode and a UV lamp (254 and 365 nm) as the excitation source and a USB2000+ spectrometer from OceanOptics for detection.

X-ray photoelectron spectroscopy (XPS) was performed with a Thermo Scientific Theta Probe Angle-Resolved X-ray photo-



electron spectrometer system using a standard monochromated Al K α source ($h\nu = 1486.6$ eV). The spectra were acquired using a pass energy of 200 eV for survey spectra and 50 eV for detailed peak scans with an analysis spot of roughly 400 μm .

Results and discussion

Growth

The Al_2Qz_3 films showed a growth rate of around 0.28 nm per cycle. The films showed a thickness gradient of about 33% reduction in the growth rate over 5 cm. This gradient was reproducible and persisted when the pulse and purge times or the precursor temperature was changed. For Al_4Pe_3 , the growth rate and gradient were 0.90 nm per cycle and 14.3%, respectively.

Fig. 2 shows the QCM response as a function of pulsing times. The observed frequency change is proportional to the mass change of the deposited material. The TMA pulse saturates at 0.2 s. For the organic pulses, quinizarin does not show saturation even after 20 s pulses while pentaerythritol saturates at 10 s. During each precursor pulse, the expected byproduct is methane from reactions between aluminium's methyl groups and the organic precursors' hydroxyl groups. For reference, the masses of the reactants and leaving molecule are: TMA = 144.18, quinizarin (H_2Qz) = 240.21, pentaerythritol (H_4Pe) = 136.15 and CH_4 = 16.04 g mol $^{-1}$.

Fig. 3 shows the frequency change during each pulse and purge step for rather long pulse and purge times (2.7 s/10 s/20 s/10 s for TMA/purge/organic/purge). The TMA pulse shows an immediate and rapid mass increase before reaching a plateau at about 1 s, while the organic precursors show a much slower and longer increase in addition to a short delay in the reaction. The mass gains during each half-cycle (pulse and purge) are $\Delta f_{\text{Al}} = 7.39$ Hz for TMA and $\Delta f_{\text{Qz}} = 10.38$ Hz for H_2Qz for Al_2Qz_3 growth, and $\Delta f_{\text{Al}} = 2.16$ Hz for TMA and $\Delta f_{\text{Pe}} = 12.27$ Hz for H_4Pe for Al_4Pe_3 growth.

While the QCM signal does not directly reveal the chemistry taking place, it can be used to evaluate possible hypotheses.

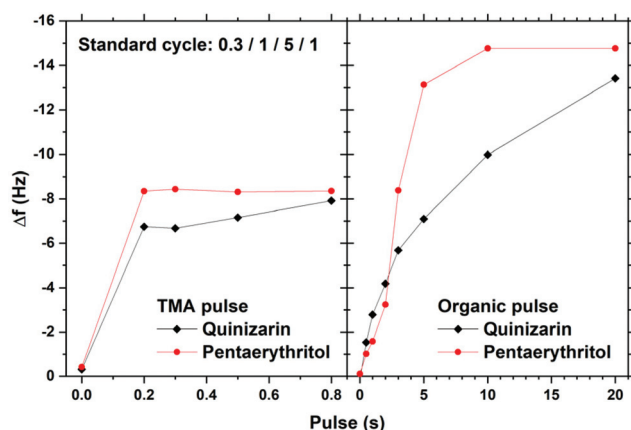


Fig. 2 QCM characterization of growth saturation as a function of pulse durations for Al_2Qz_3 and Al_4Pe_3 .

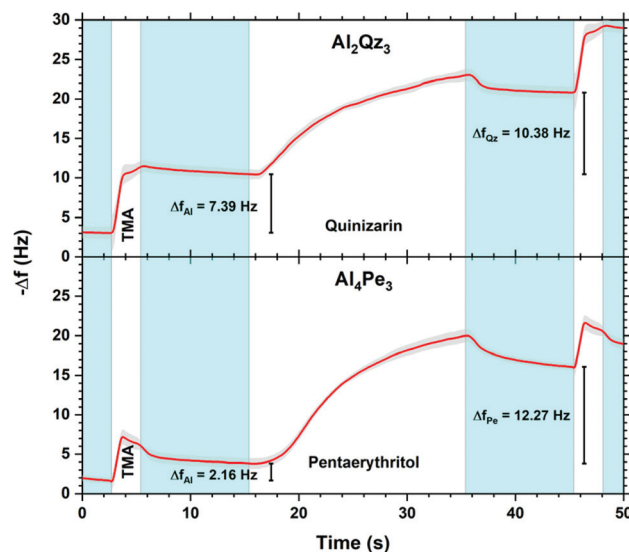
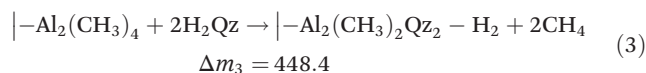
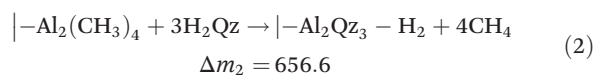
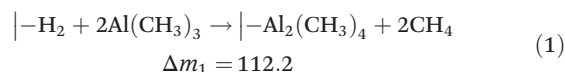


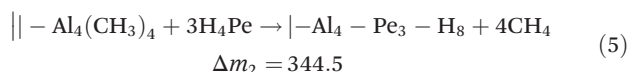
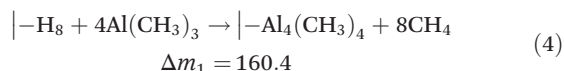
Fig. 3 Detailed QCM characterization of the growth during pulsing and purging. Pulses are marked in white and named with each respective precursor while nitrogen purges are marked in blue.

The TMA molecule should react with the Qz surface with at least one of its methyl groups. Likewise, H_2Qz should react with at least one of TMA's remaining methyl groups. We use the term "should" because TMA is a strong Lewis acid, and can coordinate directly to Lewis bases without any reaction taking place, like in the pyridine–TMA complex. However, this is unlikely here due to the reactivity of the hydroxyl groups in Qz. Still, unreacted methyl groups may remain in our film after a complete cycle. If so, this will have a notable effect on the overall stoichiometry and we expect post-reactions with ambient air. If all methyl groups react, an overall stoichiometry of Al_2Qz_3 is expected, while $\text{Al}(\text{CH}_3)\text{Qz}$ should form if only two methyl groups react. The latter is expected to be converted to $\text{Al}(\text{OH})\text{Qz}$ upon reaction with moist air. These two products can be formed by combinations of the following three half-reactions:



The mass gains during H_2Qz exposure (eqn (2) and (3)) should be about 6 and 4 times larger than those during the TMA half-cycle (eqn (1)). The observed ratio in the frequency change is on the other hand closer to 1.4 (Fig. 3). This indicates that even less Qz is deposited per Al atom than what is expected from simply balancing the valency of Al^{3+} and Qz^{2-} , and that neither of the most plausible reaction mechanisms are close to describing the real growth.

A similar analysis of the QCM results for the Al_4Pe_3 film gave $\Delta f_{\text{Al}} = 2.16$ Hz for TMA and $\Delta f_{\text{Pe}} = 12.27$ Hz for H_4Pe , resulting in an observed mass ratio of *ca.* 6 for $\Delta f_{\text{Pe}}/\Delta f_{\text{Al}}$. On the basis of the overall stoichiometry of Al_4Pe_3 , we combined the hypothetical reactions in eqn (4) and (5) and obtained an expected $\Delta f_{\text{Pe}}/\Delta f_{\text{Al}}$ value of 2.1. This indicates that a relatively large proportion of the $-\text{OH}$ groups remains unreacted during growth, and hence, the film has an excess of H_xPe . A large amount of $-\text{OH}$ in the deposited films will also be shown by FTIR analysis later (Fig. 5).



An Al_2Qz_3 film was investigated with XPS to determine its chemical bonding and stoichiometry. The C 1s peak is shown in Fig. 4, while the survey scan and Al 2s and O 1s peaks are shown in ESI Fig. 4–6.† Fig. 4 shows 3 different carbon species, with a makeup of 75% C–C and about 12–13% each of C–O and C=O. This matches well with quinizarin, having 10 out of 14 carbon atoms that are not bonded to oxygen, ideally giving 71.4% C–C. As the film material shows distinct C–O and C=O species, these two oxygen atoms must bond differently to Al^{3+} . A bidentate binding to both oxygen atoms would show up as a single carbon species in XPS instead of two distinct species. The C=O oxygen may still participate in coordinating Al^{3+} . The Al 2p peak shows that there is only one major Al species, indicating that all Al atoms are bonded similarly. The single O 1s peak indicates that all oxygen atoms are bonded similarly as well, *i.e.* to both Al and C instead of only to Al as in Al_2O_3 . The O 1s peak also does not have the discrete energy

found for Al_2O_3 . Thus, there is no indication of Al_2O_3 in the deposited material.

The atomic ratios are determined to be 5.5% Al, 67% C, and 27.5% O, giving an overall stoichiometry of $\text{AlC}_{12.2}\text{O}_5$. The two ideal reactions eqn (1) + eqn (2) (Al_2Qz_3) and eqn (1) + eqn (3) ($\text{Al}(\text{CH}_3)\text{Qz}$) give a stoichiometry of $\text{AlC}_{21}\text{O}_6$ and $\text{AlC}_{15}\text{O}_4$, respectively. The measured stoichiometry is close to $\text{Al}(\text{CH}_3)\text{Qz}$, indicating a substantial amount of unreacted Al–CH₃ groups. As mentioned, we expect that these methyl groups react with moisture upon air exposure to form Al–OH giving the composition $\text{Al}(\text{OH})\text{Qz}$, or $\text{AlC}_{14}\text{O}_5$ which is fairly close to the measured value. Thus, both XPS and QCM indicate that substantial amounts of Al–CH₃ groups stay unreacted during film growth.

Both QCM and XPS indicate an over-stoichiometry of Al, even compared to $\text{Al}(\text{OH})\text{Qz}/\text{Al}(\text{CH}_3)\text{Qz}$. XPS shows that the carbonyl C=O bond is intact in the film. Given that Al^{3+} is bonded to the hydroxy group and possibly only coordinated to the carbonyl oxygen, there is a possibility that some of the carbonyl groups contribute by coordinating additional TMA molecules. TMA can both coordinate to and react with carbonyl groups.³⁰ Dandley *et al.* found that TMA would interact with carbonyl groups to form both $\text{C}=\text{O}\cdots\text{Al}(\text{CH}_3)_3$ adduct structures and covalent Al–O bonds at different temperatures in ALD-like vapor phase chemistry.³³ This can offer a potential explanation for the apparent over-stoichiometry of Al. However, it will require further work to elucidate this effect and the details of the growth mechanism in the TMA– H_2Qz system.

Fig. 5 shows the FTIR reflectance spectra of the Al_2Qz_3 and Al_4Pe_3 films on polished steel substrates, in addition to the absorbance spectra of H_2Qz and H_4Pe powders in KBr pellets. The broad absorption band from the $-\text{OH}$ groups at 3300 cm^{-1} is visible in all spectra. The two peaks around 1600 cm^{-1} can be assigned to the carbonyl C=O stretching of quinizarin. Xuan *et al.* placed the carbonyl C=O stretching mode at 1630 cm^{-1} .³¹ ESI Fig. 3† shows these spectra zoomed in the $1800\text{--}500\text{ cm}^{-1}$ range. It is seen that the two carbonyl peaks are shifted to lower energies from 1630 and 1590 in the H_2Qz powder to 1610 and 1580 cm^{-1} in the Al_2Qz_3 films.

Al_2Qz_3 from $\text{Al}(\text{acac})_3$

In an attempt to test an alternative Al-precursor, we used the same Al-source, $\text{Al}(\text{acac})_3$, used in the solution-based work by Quinti *et al.*²⁵ $\text{Al}(\text{acac})_3$ has previously been used for ALD to grow Al_2O_3 .³² Unfortunately, despite several attempts, we did not obtain any film when combined with H_2Qz . There may be many reasons for this, one of which is that bidentate acac and bidentate Qz form nearly identical bonds with Al^{3+} , giving very little energy incentive for a reaction. A statistical ligand exchange between two isoenergetic ligands should still be possible if the reaction energy barrier was low. However, it seems that a solvent is necessary for this exchange to take place.

Optical properties

The Al_2Qz_3 films were deeply pink to the eye, similar to $\text{H}_2\text{Qz}-\text{Al}^{3+}$ solutions. Fig. 6 shows the UV-Vis transmission measure-

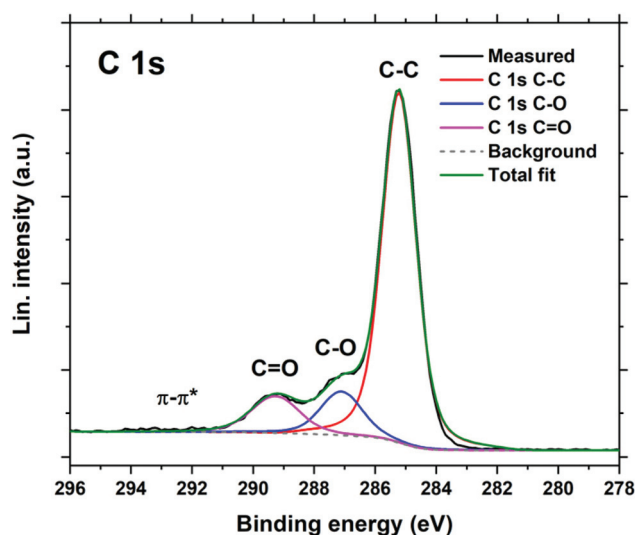


Fig. 4 XPS C 1s peak, showing contributions from three types of carbon species: C–C, C–O and C=O.



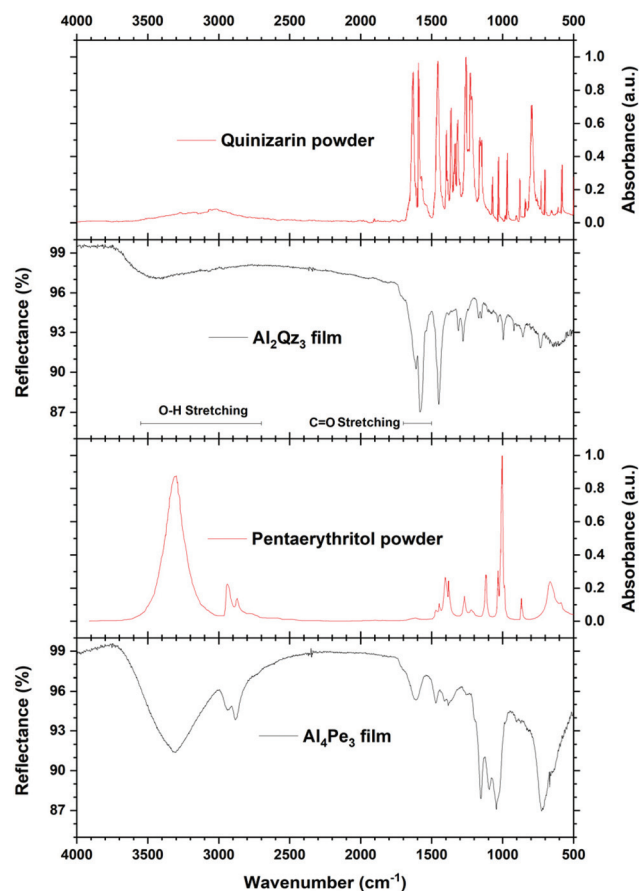


Fig. 5 FTIR data of Al_2Qz_3 and Al_4Pe_3 films on steel substrates. Quinizarin (H_2Qz) and pentaerythritol (H_4Pe) powders in KBr pellets are shown for reference, obtained from the online database Wiley SpectraBase.^{28,29}

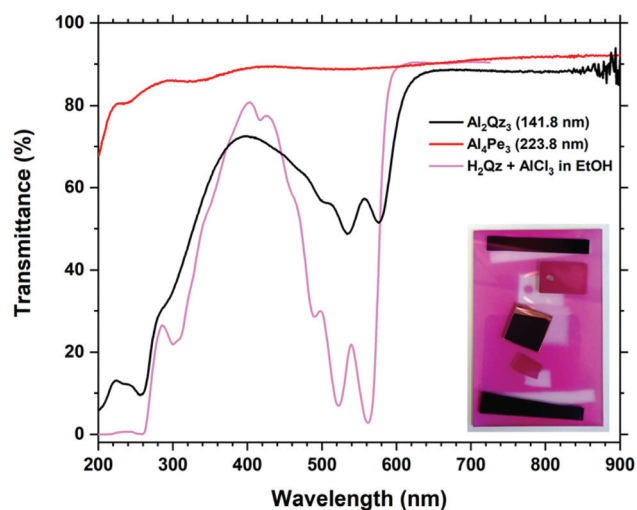


Fig. 6 UV-Vis transmission measurements of Al_2Qz_3 and Al_4Pe_3 films, both on silica substrates, and a luminescent solution of H_2Qz and AlCl_3 in ethanol. The photo inset shows the samples for this deposition. The gas inlet and exhaust are at the bottom and top, respectively. From the bottom, the samples are Si, silica, steel, sintered alumina (not used in this article) and Si.

ments of the Al_2Qz_3 and Al_4Pe_3 films on silica substrates, in addition to an ethanolic solution of AlCl_3 and H_2Qz for comparison of absorption peaks. The peak positions and shapes of Al_2Qz_3 are very similar to those of the solution, indicating that the film material is similar in optical properties to the complexes or polymeric structures formed in solution.²⁵ These peaks are distinctly different from those of a solution containing only H_2Qz without any Al^{3+} as shown in ESI Fig. 7.† The films show a blue shift in the UV range and a red shift in the visible range. The two strongest peaks in solution are shifted from 522.2 and 562.0 nm to 534.6 and 576.2 nm in the film, giving a red shift of the absorbance when Al is bonded to Qz. The Al_4Pe_3 film shows no absorption. Interference due to the thickness of the film is seen as waves, and the apparent drop in transparency below 240 nm is due to the silica substrate (Fig. 6).

Fig. 7 shows the absorbance and extinction coefficient obtained from the Al_2Qz_3 films. The optical model used for modelling the ellipsometry data is a general oscillator model with 5 Gaussian absorption peaks in the visible range. It is seen that the optical modelling fits very well with the measured absorption spectrum, showing the model's validity. The model fit to the ellipsometry data is very good, as can be seen in ESI Fig. 2.†

Dilution with pentaerythritol

The films of Al_2Qz_3 were not luminescent as deposited. This is likely caused by concentration quenching, as also the dried powder from highly luminescent Al^{3+} + quinizarin solutions is non-luminescent. To overcome this, we alloyed Al_2Qz_3 with Al_4Pe_3 in 1 : 8 and 1 : 19 ratios, effectively diluting the concentration of quinizarin molecules, which are the only optically active species. Based on the growth rate of Al_4Pe_3 , the separation of each Al_2Qz_3 cycle by 8 and 19 Al_4Pe_3 cycles leads to an approximate Qz-Qz layer separation of 7.2 and 17.1 nm, respectively. This should ensure that no Qz-Qz inter-layer

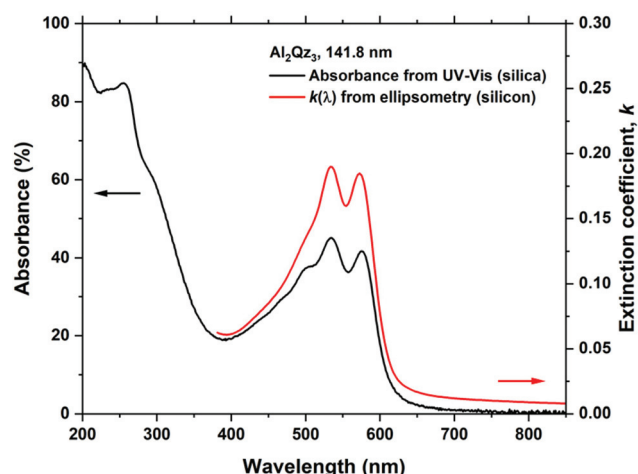


Fig. 7 Comparison of absorbance, obtained from transmittance and reflectance measurements on silica substrates, and the general oscillator optical model of the film material based on 5 Gaussian absorption peaks in the visible range.



interaction can cause luminescence quenching. Intra-layer or in-plane Qz–Qz interactions from Qz deposited during a single H₂Qz pulse are still possible. Despite such layered dilution, these films were not luminescent.

As the in-plane concentration of quinizarin resulting from a single Al₂Qz₃ cycle may be high enough to cause concentration quenching, we attempted to reduce the in-plane concentration of Qz by pulsing in H₄Pe just before the H₂Qz pulse. The idea was to occupy some of the TMA-terminated surface sites with H₄Pe, resulting in a reduced amount of Qz being deposited for each H₂Qz pulse. The regular H₂Qz pulse in Al₂Qz₃ depositions was replaced with an H₄Pe pulse immediately followed by an H₂Qz pulse in 3 s/1 s and 2 s/2 s ratios. The films resulted in strong gradients in colour, from pink to transparent. Although the gradients are usually not desired, this meant that the concentration of Qz gradually reduced to zero through the sample chamber, and this sample could thus be used in a combinatorial manner. Unfortunately, no part of these films showed any luminescence. A more in-depth investigation into controlling the in-plane and inter-plane concentration of Qz is outside the scope of this work. We believe that proper control of minimum Qz–Qz distances to avoid concentration quenching could lead to the formation of luminescent Al₂Qz₃ films. This path has previously been explored in other systems, like Al³⁺ doped ZnO. Learning from these prior works can be a way of finding a solution also for this hybrid system.

Stability

The rapidly increasing number of different hybrid materials that are being produced by ALD and MLD necessarily includes film materials that are air-stable, air-sensitive and anywhere in between. In this work, it is the optical properties which are in focus. Regarding stability in air, we investigated how around 7.5 months of air exposure affected the Al₂Qz₃ and Al₄Pe₃ film thicknesses in addition to the optical absorption and refractive index.

The Al₂Qz₃ films show very good stability towards air exposure, with only a 6% increase in the film thickness after 239 days (7 months, 24 days). These materials' $n(\lambda)$ and $k(\lambda)$ values were nearly unchanged as given in ESI Fig. 1,† apart from a slight reduction in $n(\lambda)$ ($n(632.8\text{ nm}) = 1.75 \rightarrow 1.73$) and a very slight sharpening of the visible absorption peaks in $k(\lambda)$. Al₄Pe₃, on the other hand, showed more prominent swelling with a thickness increase of 22.5%. The reduction in $n(\lambda)$ was similarly minute to that for Al₂Qz₃, $n(632.8\text{ nm}) = 1.57 \rightarrow 1.55$. Neither the Al₂Qz₃ nor Al₄Pe₃-diluted films showed any evolving luminescence after this air exposure.

Conclusion

In this work, we have obtained deep pink coloured films from TMA and quinizarin with a stoichiometry close to that of Al(OH)Qz, in addition to transparent and optically inactive films from TMA and pentaerythritol by ALD/MLD. Both QCM and XPS show an under-stoichiometry from ideal Al₂Qz₃. This is

interpreted as the presence of unreacted Al–CH₃ groups that likely oxidize upon air exposure. An absorbance of 40% in the 500–600 nm range is obtained for a 140 nm thick film. Both Al₂Qz₃ and Al₄Pe₃ films are quite air-stable, showing only minor expansion and a slight reduction in the refractive index following long-term (>7 months) air exposure. Al₄Pe₃ shows clear saturation during both TMA and organic pulse, resulting in a growth rate of 0.90 nm per cycle. Al₂Qz₃, on the other hand, does not show complete saturation during the organic pulse even for 20 s pulses. When deposited without saturation conditions, Al₂Qz₃ and Al₄Pe₃ show gradients in the 14–33% range over a distance of 5 cm. The Al₂Qz₃ films are non-luminescent, even though similar Al–Qz complexes in solution are highly luminescent. Most likely, the high density of quinizarin in the solid films causes Qz–Qz interactions and concentration quenching. We attempted to “dilute” quinizarin with Al₄Pe₃ to mitigate such concentration quenching but unfortunately did not obtain luminescent films. This work establishes quinizarin and anthraquinone derivatives as strongly absorbing visible dyes that are surprisingly well suited for ALD/MLD. This class of dye molecules includes a large range of molecules that are quite similar to quinizarin which may be similarly suited, but with a vast range of different optical and electronic properties.

Author contributions

Per-Anders Hansen: conceptualization, methodology, investigation, writing – original draft, visualization, and funding acquisition. Ola Nilsen: resources, writing – review & editing, supervision, and funding acquisition.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 V. Miikkulainen, M. Leskela, M. Ritala and R. L. Puurunen, *J. Appl. Phys.*, 2013, **113**, 021301.



- 2 A. J. M. Mackus, J. R. Schneider, C. MacIsaac, J. G. Baker and S. F. Bent, *Chem. Mater.*, 2019, **31**, 1142–1183.
- 3 H. H. Sønsteby, O. Nilsen and H. Fjellvåg, *J. Vac. Sci. Technol., A*, 2016, **34**, 041508.
- 4 P.-A. Hansen, H. Fjellvåg, T. G. Finstad and O. Nilsen, *Chem. Vap. Deposition*, 2014, **20**(7–8–9), 274–281.
- 5 K. B. Klepper, O. Nilsen, P.-A. Hansen and H. Fjellvåg, *Dalton Trans.*, 2011, **40**, 4636–4646.
- 6 K. B. Lausund, M. S. Olsen, P.-A. Hansen, H. Valen and O. Nilsen, *J. Mater. Chem. A*, 2020, **8**, 2539–2548.
- 7 X. Meng, *J. Mater. Chem. A*, 2017, **5**, 18326–18378.
- 8 P.-A. Hansen, T. Zikmund, T. Yu, J. N. Kvalvik, T. Aarholt, Ø. Prytz, A. Meijerink and O. Nilsen, *Commun. Chem.*, 2020, **3**, 162.
- 9 P.-A. Hansen, C. S. Granerød, Ø. Prytz and O. Nilsen, *J. Lumin.*, 2019, **215**, 116618.
- 10 J. N. Kvalvik, J. Borgersen, P.-A. Hansen and O. Nilsen, *J. Vac. Sci. Technol., A*, 2020, **38**, 042406.
- 11 M. N. Getz, O. Nilsen and P.-A. Hansen, *Sci. Rep.*, 2019, **9**, 10247.
- 12 H. H. Sønsteby, E. Skaar, Ø. S. Fjellvåg, J. E. Bratvold, H. Fjellvåg and O. Nilsen, *Nat. Commun.*, 2020, **11**, 2872.
- 13 Z. Giedraityte, M. Tuomisto, M. Lastusaari and M. Karppinen, *ACS Appl. Mater. Interfaces*, 2018, **10**, 8845–8852.
- 14 M. N. Getz, P.-A. Hansen, O. S. Fjellvåg, M. A. K. Ahmed, H. Fjellvåg and O. Nilsen, *J. Mater. Chem. C*, 2017, **5**, 8572–8578.
- 15 H. Zhou and S. F. Bent, *J. Vac. Sci. Technol., A*, 2013, **31**, 040801.
- 16 P. Sundberg and M. Karppinen, *Beilstein J. Nanotechnol.*, 2014, **5**, 1104–1136.
- 17 M. Tuomisto, Z. Giedraityte, L. Mai, A. Devi, V. Boiko, K. Grzeszkiewicz, D. Hreniak, M. Karppinen and M. Lastusaari, *J. Lumin.*, 2019, **213**, 310–315.
- 18 J. Cervantes-González, D. A. Vosburg, S. E. Mora-Rodriguez, M. A. Vázquez, L. G. Zepeda, C. Villegas Gómez and S. Lagunas-Rivera, *ChemCatChem*, 2020, **12**, 3811–3827.
- 19 D. K. Palit, H. Pal, T. Mukherjee and J. P. Mittal, *J. Chem. Soc., Faraday Trans.*, 1990, **86**, 3861–3869.
- 20 S. Agrawal, S. M. Clarke, I. J. Vitorica-Yrezabal, C. Liu, W. Fang, P. T. Wood and D. Wright, *Mol. Phys.*, 2019, **117**, 3424–3433.
- 21 G. Ramakrishna, A. K. Singh, D. K. Palit and H. N. Ghosh, *J. Phys. Chem. B*, 2004, **108**, 4775–4783.
- 22 S. Rawalekar, S. Verma, S. Kaniyankandy and H. N. Ghosh, *Langmuir*, 2009, **25**, 3168–3172.
- 23 A. L. M. Reddy, S. Nagarajan, P. Chumyim, S. R. Gowda, P. Pradhan, S. R. Jadhav, M. Dubey, G. John and P. M. Ajayan, *Sci. Rep.*, 2012, **2**, 960.
- 24 V. Rajshekar Shetty, S. S. Gurukar, R. Marriappa, M. M. Kittappa and D. H. Nagaraju, *New J. Chem.*, 2015, **39**, 8534–8544.
- 25 L. Quinti, N. S. Allen, M. Edge, B. P. Murphy and A. Perotti, *J. Photochem. Photobiol., A*, 2003, **155**, 79–91.
- 26 L. Quinti, N. S. Allen, M. Edge, B. P. Murphy and A. Perotti, *J. Photochem. Photobiol., A*, 2003, **155**, 93–106.
- 27 G. Sauerbrey, *Z. Phys.*, 1959, **155**, 206–222.
- 28 <http://spectrabase.com/spectrum/6bxGJZ9BHEf>.
- 29 <http://spectrabase.com/spectrum/24A4lKktwTw>.
- 30 Y. Kwon, O. Scadeng, R. McDonald and F. G. West, *Chem. Commun.*, 2014, **50**, 5558–5560.
- 31 X. Xuan, X. Wang and N. Wang, *Spectrochim. Acta, Part A*, 2011, **79**, 1091–1098.
- 32 M. Nieminen, T. Sajavaara, E. Rauhala, M. Putkonen and L. Niinistö, *J. Mater. Chem.*, 2001, **11**, 2340–2345.
- 33 E. C. Dandley, C. D. Needham, P. S. Williams, A. H. Brozena, C. J. Oldham and G. N. Parsons, *J. Mater. Chem. C*, 2014, **2**(44), 9416–9424.

