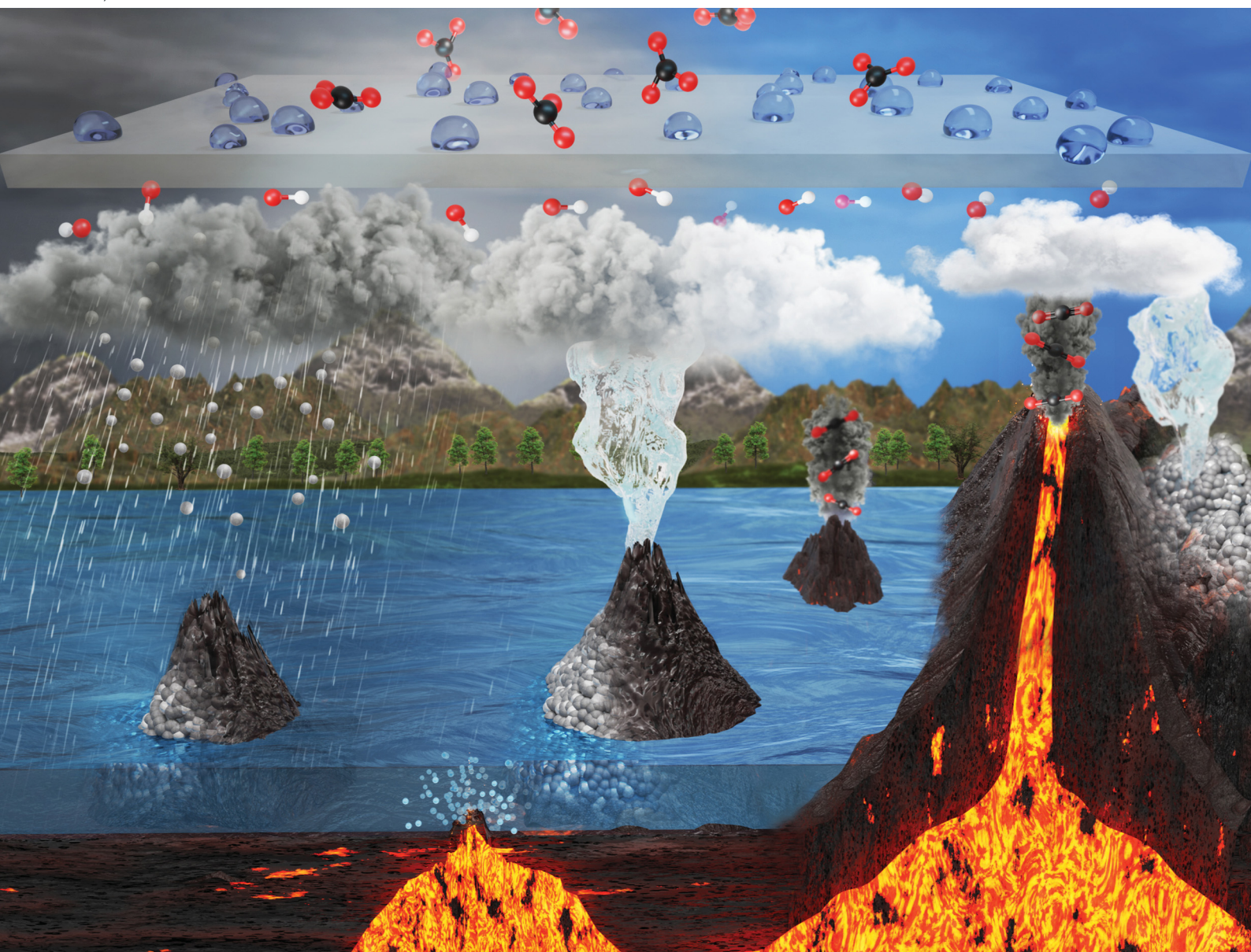


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PAPER

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Breaking the current limitation of electrochemical CO₂
reduction *via* a silica-hydroxide cycle

Cite this: *Energy Environ. Sci.*, 2024, 17, 6215Breaking the current limitation of electrochemical CO₂ reduction *via* a silica-hydroxide cycle†Chulwan Lim,^{‡ab} Sangkuk Kim,^{‡ac} Ji Hwan Song,^{‡a} Man Ho Han,^{id a} Young-Jin Ko,^a Kwan-Young Lee,^{id b} Jae-Young Choi,^{id de} Woong Hee Lee^{id *a} and Hyung-Suk Oh^{id *ade}

Alkaline local pH during a vigorous electrochemical CO₂ reduction reaction (CO₂RR) can improve the activity and selectivity of CO₂RR. However, it also leads to an alkalinity problem in that hydroxide ions obstruct the mass transfer of CO₂ to the active site, thereby limiting the current density. In this study, we introduce a silica-hydroxide cycle, which moderates the local pH by redistributing hydroxide ions, analogous to the carbonate-silicate cycle responsible for the drawdown of atmospheric CO₂ on Earth. In the membrane electrode assembly (MEA) of a CO₂ electrolyzer, SiO₂ undergoes weathering due to the high local pH and consequently consumes OH⁻, reducing the pH within the MEA. The dissolved silicate ions move to the membrane and are almost regenerated to SiO₂ with release of OH⁻. Geological and spectral observations suggest that the silica-hydroxide cycle reduces the local pH thereby enhancing mass transfer of CO₂, breaking the limitation of current density for CO₂RR. Our work proposes new chemical approaches to increase current density, mainly improved by physical methods, and contributes valuable insight for improving a variety of electrochemical systems.

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Broader context

The electrochemical CO₂ reduction reaction (CO₂RR) is viewed as a promising approach to attain carbon neutrality while generating valuable fuels and chemicals. Enhancing the current density in CO₂RR presents a significant challenge for the efficient production of chemicals on a meaningful scale, which is essential for economic viability. Recently, various mechanical approaches have enhanced the performance of CO₂RR through enhancing the mass transfer of CO₂, playing a pivotal role in improving the current density. Nevertheless, in the high current density region, the hydroxide ions created during the reaction form a local alkaline environment, which converts CO₂ to carbonate ions, hindering the mass transfer of CO₂ to the catalyst surface and limiting the current density for CO₂RR. This study delves into the role of SiO₂ in the electrochemical CO₂RR regarding the silica-hydroxide cycle, which moderates the local pH by redistributing hydroxide ions. This mechanism draws a parallel with the silicate-carbonate cycle, which plays a crucial role in the natural sequestration of atmospheric CO₂ on Earth, thus offering a novel avenue for breaking the current density limitation of CO₂RR through controlled pH management.

Introduction

The electrochemical CO₂ reduction reaction (CO₂RR) is viewed as a promising solution for achieving carbon neutrality while producing valuable fuels and chemicals.^{1–3} Increasing the current density for CO₂RR poses a significant challenge for the mass production of chemicals while ensuring economic feasibility.^{4,5} Recently, various mechanical approaches, including gas-phase CO₂, device design, gas diffusion electrode (GDE), membrane electrode assembly (MEA), pressurized system and nanosized catalyst, have successfully enhanced the performance of CO₂RR *via* enhancing mass transfer of CO₂, playing a pivotal role in improving the current density.^{6–17} Under high current density regime coming from a recently developed system, an intriguing phenomenon, called local pH,

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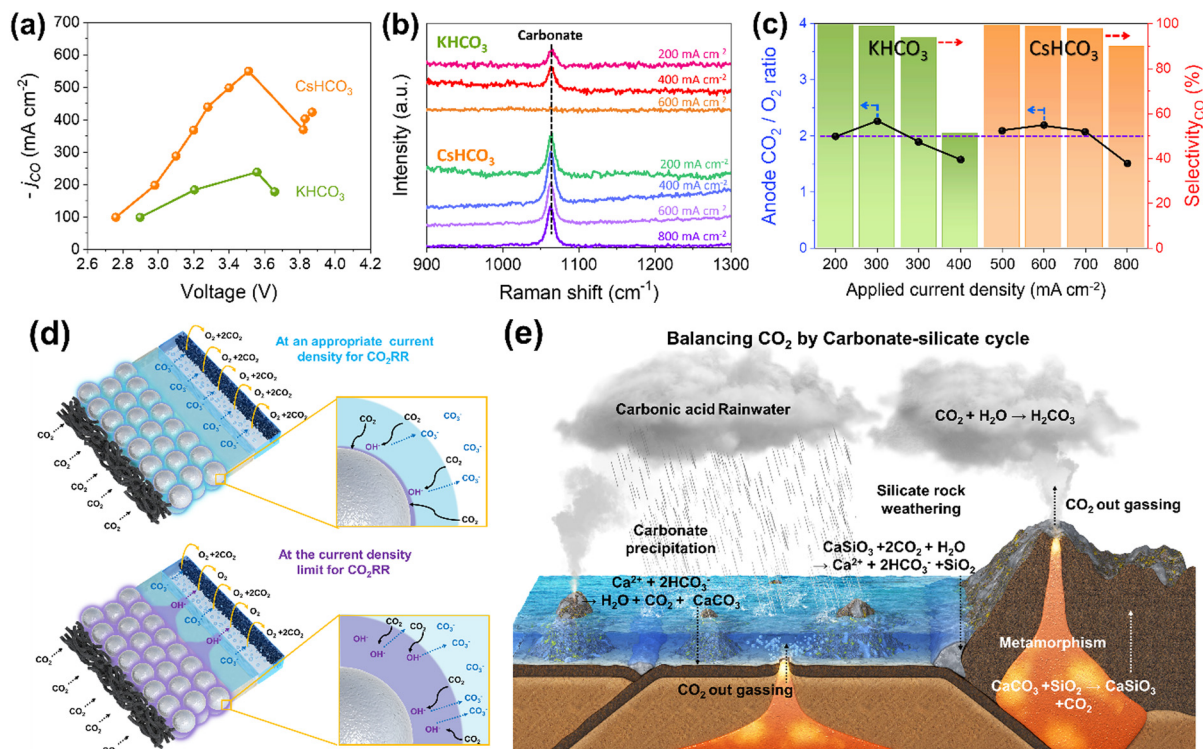


Fig. 1 Alkalinity issue in CO₂RR and carbonate-silicate cycle. (a) CO partial current density versus cell voltage in a zero-gap electrolyzer using Ag black with 0.1 M KHCO₃ and 0.1 M CsHCO₃. (b) *In situ/operando* Raman spectroscopy results with 1 M KHCO₃ and CsHCO₃ electrolytes. (c) CO₂/O₂ ratio at the anode in a zero-gap electrolyzer with 0.1 M KHCO₃ and 0.1 M CsHCO₃. (d) A schematic depiction of the alkalinity problem in a zero-gap CO₂ electrolyzer. (e) A schematic representation of the carbonate-silicate geochemical cycle on Earth.

clearly appeared for CO₂RR. The hydroxide ions generated from the reaction create a local alkaline environment at the electrode, promoting CO₂RR *via* reducing the hydrogen evolution reaction (HER) and modulating binding energies for the intermediates.^{18–24} Nevertheless, some studies addressing local pH have indicated an alkalinity problem in that hydroxide ions block the mass transfer of CO₂ to the catalyst by converting carbonate, limiting the current density for CO₂RR.^{25,26} The alkalinity problem of CO₂RR can be validated using alkali cation effects. Numerous experimental and theoretical investigations reveal that a larger size of alkali cation exhibits smaller hydration number at the Helmholtz layer and lower pK_a values near the electrode, mitigating the local pH phenomenon.^{27–29}

The current density of CO₂RR according to alkali cation was assessed in a CO₂ MEA electrolyzer using Ag black sprayed GDE (Fig. 1a and Fig. S1–S3, ESI[†]). CsHCO₃ electrolyte demonstrated a significantly higher maximum current density for CO₂RR than KHCO₃ electrolyte. To delve into the pH reaction environment of the cathode in the MEA electrolyzer, *in situ/operando* Raman spectroscopy and gas analysis of the anode were carried out (Fig. 1b and c). In the Raman spectra for the KHCO₃ electrolyte, the carbonate peak disappeared at current densities exceeding 400 mA cm⁻², while the carbonate peak persisted at a current density of 800 mA cm⁻² using CsHCO₃ electrolyte. The *in situ/operando* Raman spectroscopy focused on the catalyst surface. Thus, the disappeared carbonate peak indicates that

the catalyst surface possesses high alkalinity with almost no carbonate ions resulting from reaction with OH⁻ and CO₂ gas. These observations indicate that the larger size of Cs cations alleviates the local pH phenomenon at high current densities. The gas ratio of CO₂ and O₂ for the anode side decreased below 2 was observed at 400 mA cm⁻² for KHCO₃ and 800 mA cm⁻² for CsHCO₃. The changes of gas ratio by current density are well explained by the schematic illustration in Fig. 1d. The ratio of CO₂ and O₂ is close 2 in the current density region showing high faradaic efficiency (FE) for CO, indicating the main carbonate ion transfer from cathode to anode.^{30,31} The decreased CO₂ at the anode shows transfer of hydroxide ions from the cathode, suggesting a highly alkaline reaction environment at the cathode sufficient to prevent hydroxide ion contact with CO₂. Additionally, as shown in the *in situ/operando* Raman results in Fig. 1b, the mass transfer of CO₂ to the catalyst layer is impeded by high local pH, further limiting the current density for CO₂RR. This suggests that, besides physical properties that can be improved by mechanical methods, the chemical reaction environment also plays a significant role in current limitations. Thus, chemical approaches, such as cation effects, are essential for enhancing current density in CO₂RR. The key for a chemical strategy is alleviating local pH by rebalancing CO₂ and hydroxide ion concentration at the cathode.

In nature, Earth's CO₂ is balanced by various carbon cycles including the carbonate-silicate cycle.^{32–35} This



geochemical process redistributes CO₂ in the atmosphere to the oceans and land by rainwater, silicate weathering, carbonate precipitation and sedimentation, resulting in drawdown of atmospheric CO₂ (Fig. 1e) The stored CO₂ in land is redistributed to the atmosphere by volcanism. In this study, we aim to break the current density limitation for CO₂RR

using silica oxide *via* a silica-hydroxide cycle, similar to the carbonate-silicate cycle. The unique silica properties allow distribution of the high OH⁻ concentration at the catalyst surface to the other cathode side, reducing the elevated local pH at the catalyst surface and facilitating mass transfer of CO₂ for CO₂RR.

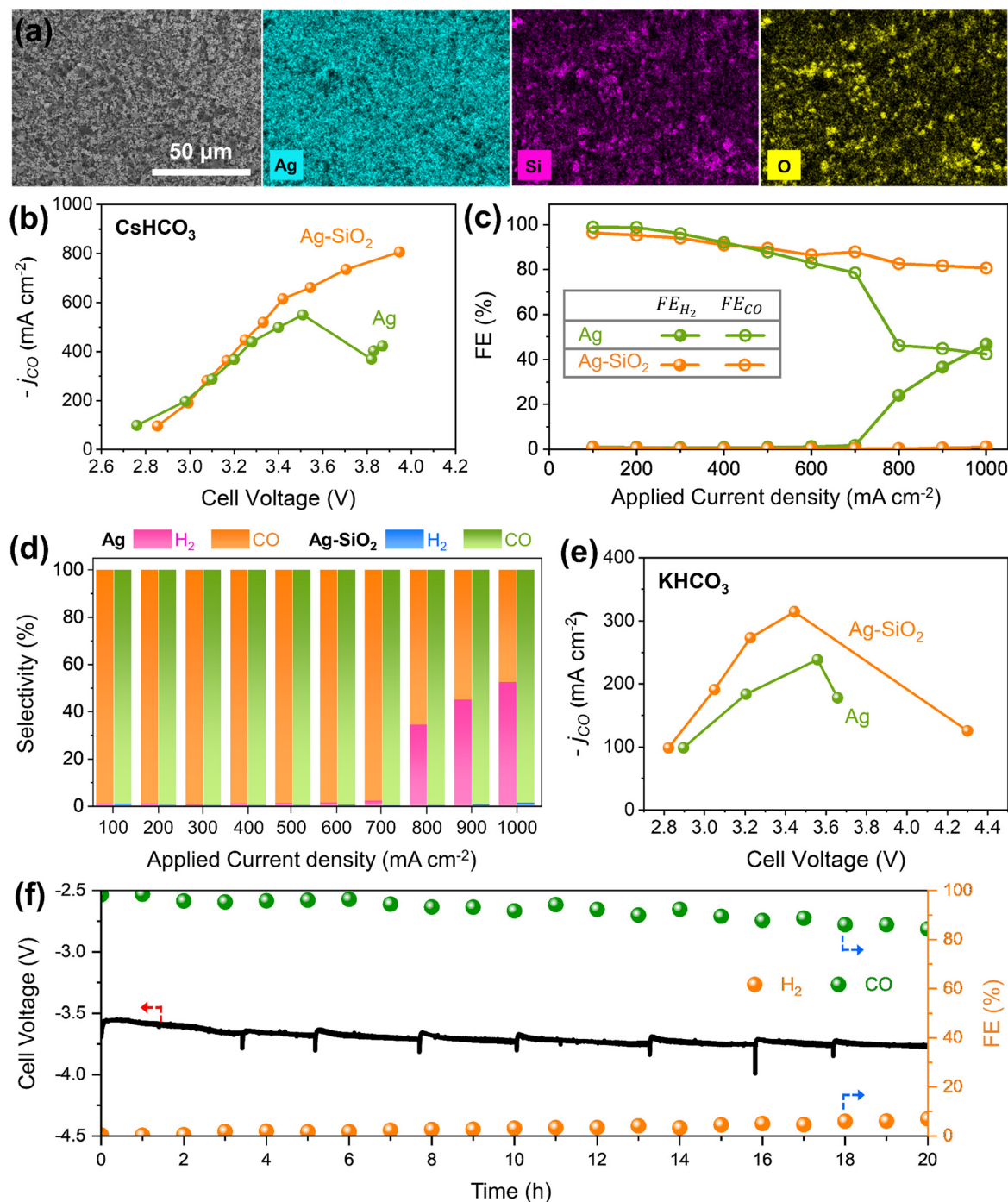


Fig. 2 Impact of SiO₂ in MEA CO₂ electrolyzer, alongside CO₂RR performance comparisons between Ag black and Ag-SiO₂. (a) SEM image and EDS elemental mapping of the Ag-SiO₂ electrode. (b) CO partial current density, (c) H₂ and CO faradaic efficiency, and (d) H₂ and CO selectivity for both Ag black and Ag-SiO₂ using 0.1 M CsHCO₃. (e) CO partial current density of Ag black and Ag-SiO₂ with 0.1 M KHCO₃. (f) Durability test for Ag-SiO₂ in the zero-gap electrolyzer using 0.1 M CsHCO₃ at 700 mA cm⁻² for 12 h.



Results and discussion

Effects of SiO₂ on current density for CO₂RR

An Ag–SiO₂ electrode was fabricated on a GDE by spraying ink of a mixture of SiO₂ nanopowder and commercial Ag black (Fig. S4–S6, ESI†). Fig. 2a illustrates SEM and SEM-EDX mapping images of the Ag–SiO₂ electrode, indicating that Ag black and SiO₂ nanoparticles are well dispersed on the GDE substrate. The effects of SiO₂ for CO₂RR performance were studied by using an anion exchange membrane (AEM) zero-gap electrolyzer with 0.1 M CsHCO₃ anolyte and humidified CO₂ gas (Fig. 2b, d and Fig. S7, ESI†).^{36,37} At an applied current density below 600 mA cm⁻², both the Ag and Ag–SiO₂ electrodes show similar CO₂RR performance, with a high FE_{CO} exceeding 90%. However, above an applied current density of 700 mA cm⁻², the FE_{H₂} of the Ag electrode drastically increased with rising current density, resulting in a maximum current density for CO of 549 mA cm⁻². On the other hand, the FE_{CO} for the Ag–SiO₂ electrode remained at over 80% and the selectivity for CO (H₂ and CO, SE_{CO}) was nearly 100% at an applied current density until 1 A cm⁻² while suppressing the HER. The difference of FE_{CO} and SE_{CO} at a high current density region would come from production of formate (Fig. S8, ESI†). The maximum current density of the Ag–SiO₂ electrode for CO reached a value of 806 mA cm⁻², implying that additional SiO₂ to the Ag black

electrode contributed to breaking the current density limitation for CO₂RR. Exploring the effects of varying SiO₂ ratios indicates the necessity for an optimal amount of SiO₂ (Fig. S9, ESI†). To confirm the physical effects, an Ag–TiO₂ electrode was fabricated using similar size of TiO₂ instead of SiO₂ (Fig. S10, ESI†). Ag–TiO₂ exhibits a similar limiting current density to Ag for CO₂RR. The results of Ag–TiO₂ and the comparable ECSA value for both the Ag and Ag–SiO₂ electrodes show that the impact of SiO₂ does not originate from physical properties (Fig. S11 and S12, ESI†). A similar phenomenon of breaking the maximum current density limit for CO₂RR was observed using 0.1 M KHCO₃ anolyte at lower current densities (Fig. 2e and Fig. S13, ESI†). To validate the SiO₂ effects for high current density, chronopotentiometry tests were conducted at a current density of 700 mA cm⁻² where the effect of SiO₂ is seen (Fig. 2f and Fig. S14, ESI†). FE_{CO} of the Ag electrode was close to 80% and rapidly decreased after 3 h. In contrast, the Ag–SiO₂ electrode demonstrated stable cell voltage and consistent FE_{CO} above 90% for 12 h, suggesting that the effects of SiO₂ are not an instant phenomenon and could serve as a sustained strategy.

The origin of SiO₂ effects in breaking current density limitation

To elucidate the origins of SiO₂ effects that overcome the current density limitation of CO₂RR, the cathode reaction

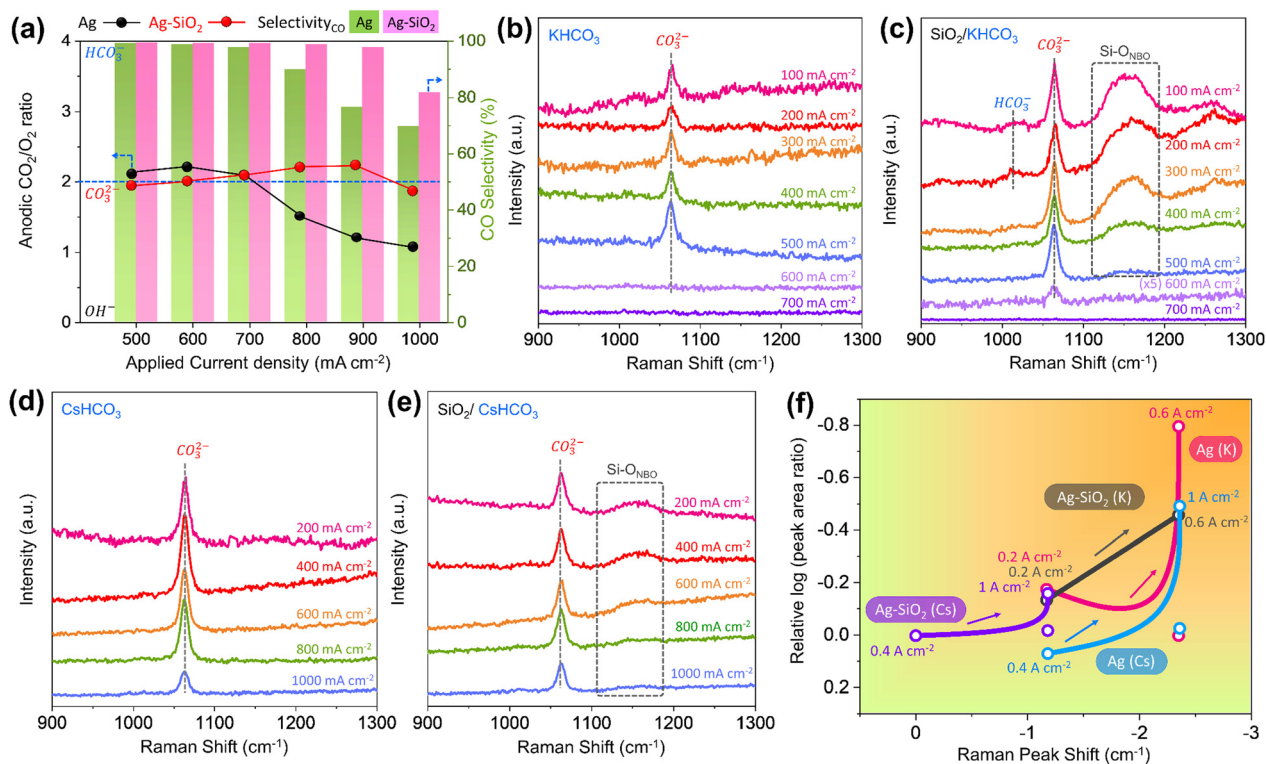


Fig. 3 The analyses of the role of SiO₂ in modulating the cathode reaction environment. (a) CO₂/O₂ ratio at the anode in a zero-gap electrolyzer using Ag–SiO₂ with 0.1 M CsHCO₃. (b) *In situ/operando* Raman spectroscopy results for the Ag electrode in 1 M KHCO₃ electrolyte. (c) *In situ/operando* Raman spectroscopy results for the Ag–SiO₂ electrode in 1 M KHCO₃ electrolyte. (d) *In situ/operando* Raman spectroscopy results for the Ag electrode in 1 M CsHCO₃ electrolyte. (e) *In situ/operando* Raman spectroscopy results for the Ag–SiO₂ electrode in 1 M CsHCO₃ electrolyte. (f) Summary of *in situ/operando* Raman spectroscopy results using carbonate Raman peak shifts and carbonate peak intensity ratios for both Ag and Ag–SiO₂ electrodes in KHCO₃ and CsHCO₃ electrolytes. Reference points for peak shift and intensity: KHCO₃, 100 mA cm⁻²; CsHCO₃, 200 mA cm⁻².



environment in the MEA electrolyzer was investigated using gas analysis of the anode and *in situ/operando* Raman spectroscopy using a GDE. For the Ag-SiO₂ electrode, the gas composition ratio of CO₂ to O₂ remains near 2 within a current density range of 500–1000 mA cm⁻². However, for the Ag electrode, this ratio drops to almost 1 when the current density exceeds 800 mA cm⁻² (Fig. 3a). This suggests that SiO₂ mitigates the local pH effects stemming from high current densities.

To more directly observe the reaction environment of the cathode, *in situ/operando* Raman spectroscopy was carried out using a customized GDE Raman cell (Fig. S15, ESI†).³⁸ To estimate the pH of the cathode according to current density, we constructed two calibration curves for ranges of 9–13 and 13–14 (Fig. S16–S20, ESI†). Even at a low current density of 100 mA cm⁻², the bicarbonate peak for all *in situ/operando* Raman

spectra nearly disappears, signifying pronounced local pH effects in the GDE.²³ In a KHCO₃ electrolyte, the carbonate peak disappears for the Ag electrode at a current density of 600 mA cm⁻² (Fig. 3b). Conversely, for the Ag-SiO₂ electrode, a minor carbonate peak persists at the same current density, indicating that SiO₂ mitigates local pH effects, reminiscent of cation effects (Fig. 3c). The Ag-SiO₂ electrode also exhibited a similar trend in CsHCO₃ electrolyte. The carbonate peak for the Ag electrode is lower than that for the Ag-SiO₂ electrode at a current density of 1 A cm⁻² (Fig. 3d and e).

For the Ag-SiO₂ electrode, peaks corresponding to Si-O_{NBO} (where NBO is non-bridging oxygen) were observed at a current density between 100 and 400 mA cm⁻², hinting at the onset of SiO₂ degradation in alkaline conditions (Fig. S21, ESI†).^{39,40} For intuitive representation of local pH under various conditions,

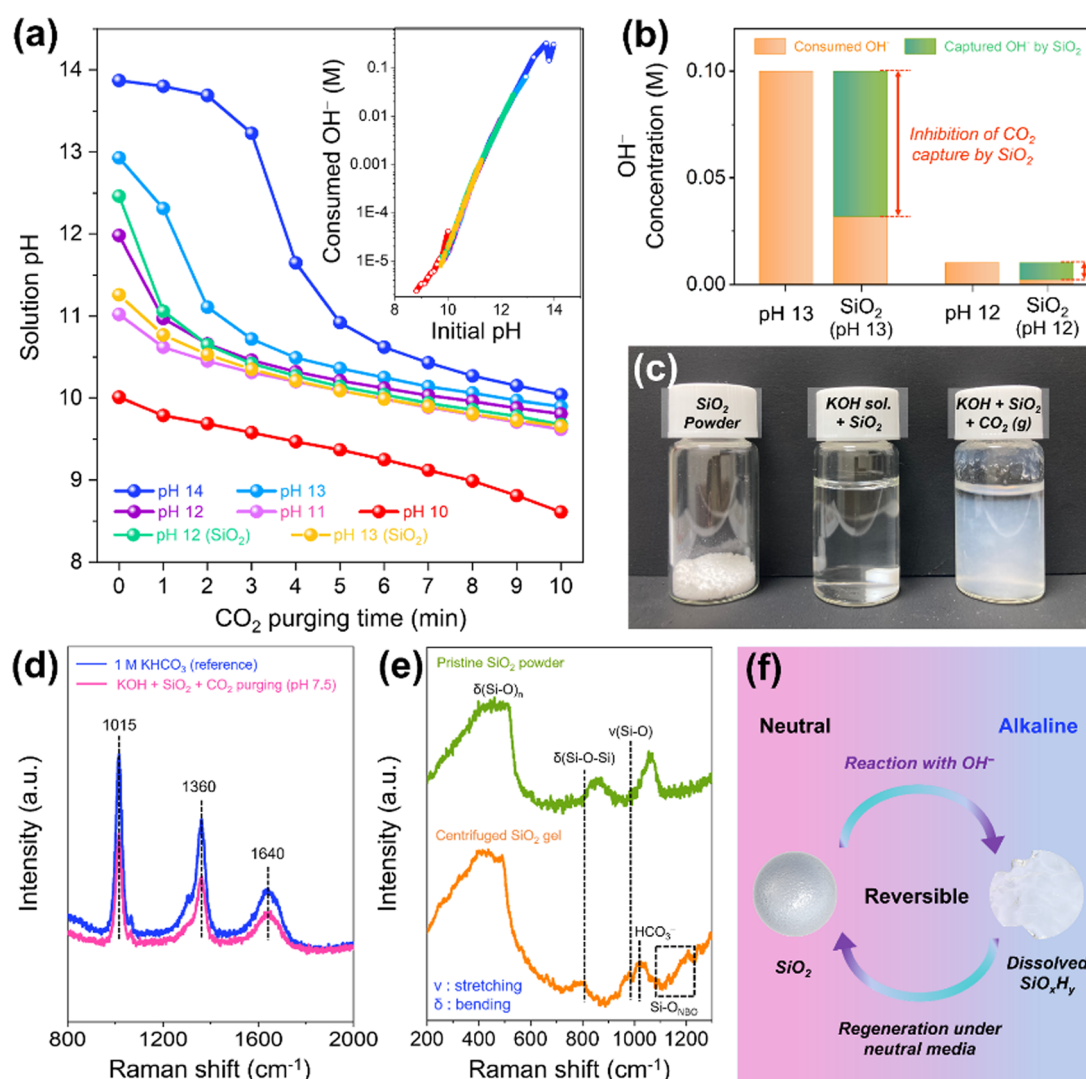


Fig. 4 Analysis of SiO₂ in CO₂RR media. (a) Change in solution pH for various initial pH solutions under CO₂ purge, plotted against CO₂ purging time (Inset: Quantity of OH⁻ consumed based on current pH). (b) Variation in OH⁻ concentration during CO₂ purging, attributed to OH⁻ capture due to the presence of SiO₂. (c) Images of pristine SiO₂ powder, SiO₂-dissolved KOH solution, and an opaque gel obtained within the SiO₂-dissolved KOH solution after CO₂ purging. (d) Raman spectra comparing 1 M KHCO₃ solution (reference), and the CO₂-purged, SiO₂-dissolved KOH solution (containing the opaque gel). (e) Raman spectra comparing pristine SiO₂ powder and centrifuged SiO₂ gel obtained from the opaque gel within the CO₂-purged solution. (f) A schematic illustrating the reversible behavior of SiO₂ species during the reaction.



we summarize the *in situ/operando* Raman results based on two parameters: the log-transformed relative peak area ratio of the carbonate peak and its shift (Fig. 3f). The carbonate peak shift is expected to be related to the Stark effect or cation concentration.^{41–43} However, since there is no carbonate intermediate in the CO₂RR mechanism on the Ag surface, we assumed that this peak shift was caused by cation concentration. The calibration results unequivocally show that carbonate peak area, which is related to carbonate concentration, is correlated to pH and peak shift corresponds to cation concentration (Fig. S22 and

S23 and Note S2, ESI[†]). Based on calibration data, the estimated local pH results from *in situ/operando* Raman spectra are shown in Fig. S24 and Table S1 (ESI[†]). While the alkalinity of the Ag electrode increased with current density, the Ag–SiO₂ electrode in comparison retained its pH, underscoring the role of SiO₂ in the suppression of local pH effects.

Behavior of SiO₂ in CO₂RR system

In an effort to understand the effects of SiO₂, it is necessary to observe the behavior of SiO₂ in a medium that simulates the

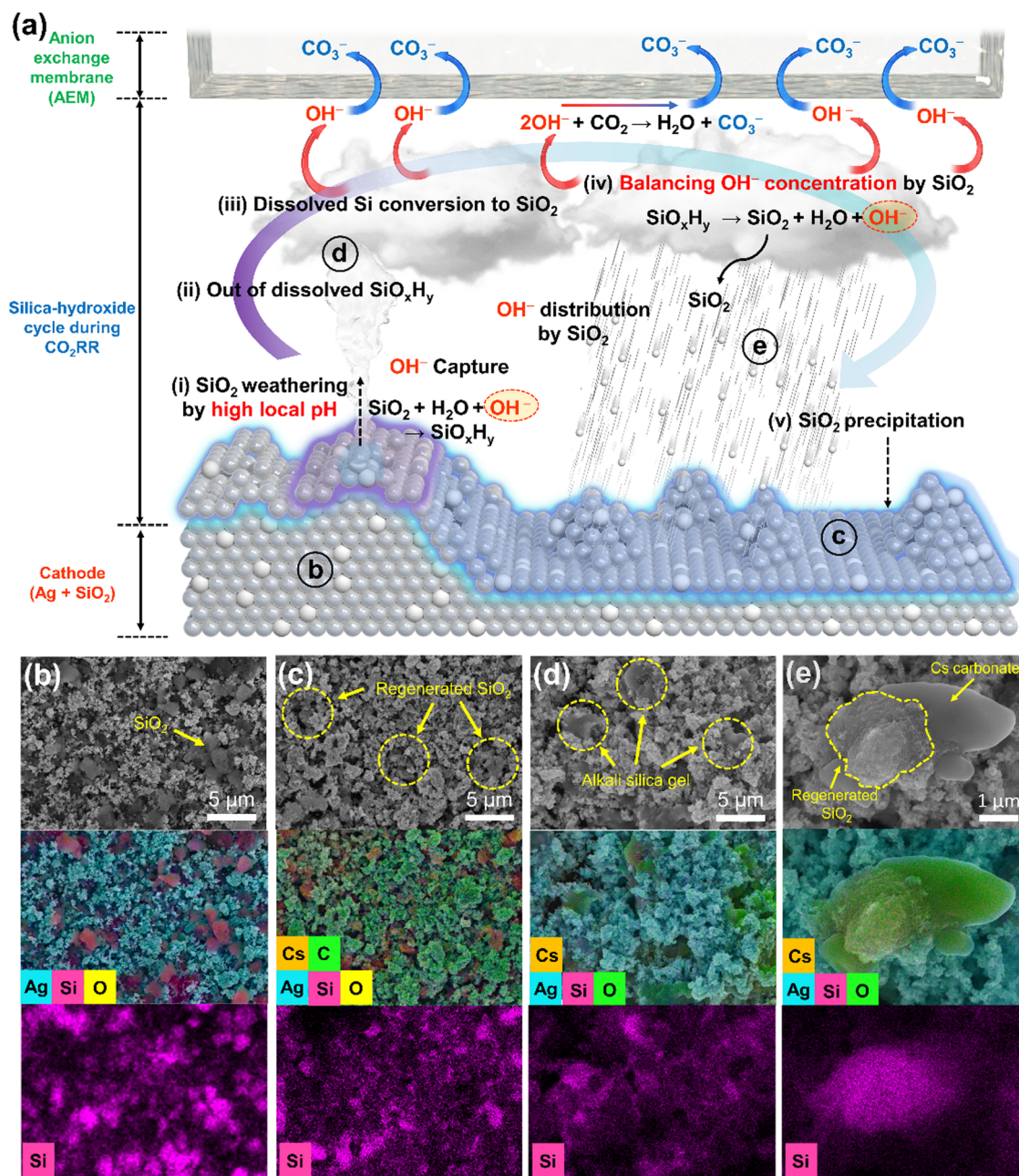


Fig. 5 Geological and spectral observations of silica-hydroxide cycle in MEA CO₂ electrolyzer. (a) A schematic representation of the carbonate-silicate geochemical cycle occurring during electrochemical CO₂ reduction reaction on Ag–SiO₂. SEM images and EDS elemental mapping of the Ag–SiO₂ electrode before (b) and after (c)–(e) reaction (c: 900 mA cm⁻² for 20 min; (d) and (e): 1 A cm⁻² for 20 min) in MEA CO₂ electrolyzer.



reaction environment for a CO₂ MEA electrolyzer. To verify the neutralization of local pH by CO₂, we purged the CO₂ gas into solutions of various pH and measured the pH every minute (Fig. 4a). As the duration of the CO₂ purge increased, the solution's pH decreased, owing to the reaction of OH⁻ ions with CO₂, resulting in carbonate ions. The amount of consumed OH⁻ with CO₂ is proportional to pH of solution, highlighting the alkalinity problem that blocks the transfer of CO₂ to active sites due to the high local pH. When the SiO₂ nanopowder was mixed with alkaline solution of pH 12 and 13, SiO₂ dissolved, leading to a decrease in the solution's pH. Regardless of the presence of SiO₂, the rate of OH⁻ consumption by CO₂ was invariably proportional to the pH of the solution, implying that CO₂ gas capture by OH⁻ is suppressed by SiO₂ (Fig. 4b). Interestingly, suspensions were detected in SiO₂-dissolved solution following an extended CO₂ purge (Fig. 4c). The analysis of collected solution and precipitate (by centrifugation) after 2 h of CO₂ purge (the pH of solution reached 7.5) demonstrates separation of SiO₂ and KHCO₃ solution,⁴⁴ suggesting almost reversible behavior of SiO₂ (Fig. 4d, e and Fig. S21, S25, ESI[†]). When the SiO₂ nanopowder was mixed with solutions of 0.1 M KHCO₃ (pH 8.65) and 1 M KHCO₃ (pH 8.69), the pH remained almost unchanged at 8.70 (Fig. S26, ESI[†]). The presumed reaction involving SiO₂ is depicted in Fig. 4f. In alkaline media, SiO₂ dissolved and converted to silicate ions by reaction of OH⁻. These silicate ions then precipitate as SiO₂ under neutral conditions with release of OH⁻ due to low solubility.^{44–46} We hypothesize that this cycle involving SiO₂ and hydroxide also takes place in the CO₂ MEA electrolyzer.

The silica-hydroxide cycle in MEA CO₂ electrolyzer

The mechanism of breaking the current limitation by SiO₂ in a CO₂ electrolyzer is hypothesized as being due to the reversible cycle of SiO₂ between alkaline and neutral media, carrying OH⁻ ions (Fig. 5a). We postulate that this silica-hydroxide cycle decreases the local pH of the electrode surface by increasing hydroxide distribution, thereby enhancing mass transfer of CO₂ for CO₂RR. This process of silica-hydroxide cycle is similar to the silicate-carbonate cycle that maintains the CO₂ level in the atmosphere. When a vigorous CO₂RR occurs in the MEA CO₂ electrolyzer, the pH of catalyst surface is presumed to be that of an alkaline medium in the range of pH 11–14 by *in situ* operando Raman spectroscopy. Concurrently, the pH of the AEM is expected to be neutral owing to abundant transported carbonate ions, indicating a suitable environment for the reversible silica-hydroxide cycle.

To demonstrate the silica-hydroxide cycle in the MEA CO₂ electrolyzer, silica weathering, evidence of the silica-hydroxide cycle, was monitored using geological and spectral observations with SEM. The pristine Ag–SiO₂ electrode clearly delineates the boundary between SiO₂ and Ag nanoparticles (Fig. 5b and Fig. S27, ESI[†]). After CO₂RR, this boundary becomes less distinct in SEM images (Fig. 5c and Fig. S28, ESI[†]), and this blurring intensifies as the current density increases (Fig. S29–S33, ESI[†]). Nevertheless, SiO₂ is well dispersed in the electrode,

similar to before the reaction. Notably, no discernible difference in Ag nanoparticles dependent on SiO₂ is observed, suggesting that SiO₂ has no effect on the chemical state and structure of Ag (Fig. S34–S39, ESI[†]). These outcomes corroborate the hypothesis of silica weathering under high current densities. To further substantiate the presence of dissolved silicate ions and the regeneration of SiO₂, EDS mapping was employed to discern the state of SiO₂. When the dissolved silicate ions dried with alkali cations, hydrous alkali silica gel is formed.⁴⁷ This hydrous alkali silica gel is observed in certain regions of the electrode, signifying the presence of dissolved silicate ions (Fig. 5d and Fig. S40, ESI[†]). The accumulation of regenerated SiO₂ on the Ag electrode, the final step of the silica-hydroxide cycle, is clearly shown in Fig. 5e. SiO₂ appears on the Cs carbonate precipitate that is generated during or post-CO₂RR, implying regenerated SiO₂.

Under high cathodic currents, SiO₂ weathering is triggered by high local pH, leading to a decrease in pH by consuming OH⁻ at the catalyst surface. The dissolved silicate ions, not reacted with CO₂, move to the AEM side.^{48,49} The dissolved silicate ions are converted to SiO₂ with counterbalancing release of OH⁻ under neutral conditions at the AEM and the converted SiO₂ is redistributed on the electrode. However, the silica-hydroxide cycle is not completely reversible in the MEA electrolyzer owing to the AEM, which is highly permeable to anions and cations. SEM-EDS mapping cross-section images after reaction show that the majority of SiO₂ remains in the cathode, but a small amount of SiO₂ has been moved to the AEM and anode (Fig. S41, ESI[†]). To eliminate the regeneration step at the AEM, 1 M KOH was used for MEA test instead of 0.1 M CsHCO₃ at a current density of 700 mA cm⁻². ICP-OES analysis shows that 67.6% of SiO₂ is preserved on the cathode with the silica-hydroxide cycle but, without the regeneration step, 5.8% of Si remained on the cathode. These results show that the regeneration step, which is critical for reversibility of the silica-hydroxide cycle, functions in a real MEA electrolyzer. This silica-hydroxide cycle reduces the concentration of OH⁻ that captures CO₂, thereby enhancing mass transfer of CO₂ to catalyst active sites and breaking the limitation of current density.

Conclusions

Our study introduces a novel strategy to overcome the limitation of the current density for CO₂RR by a chemical cycle approach, without the need for physical modification. Hydroxide ions generated from CO₂RR elevate the local pH. While this can enhance selectivity, it also hinders the transfer of CO₂ by forming carbonate ions, thus imposing a current density limitation for CO₂RR. Therefore, controlling the local pH is crucial for achieving a high current density. On Earth, atmospheric CO₂ levels are reduced *via* carbon cycles, such as the silicate-carbonate cycle. This geological carbon cycle redistributes CO₂ from the atmosphere to rocks through silicate weathering. Similarly, the silica-hydroxide cycle derived using



SiO₂ causes a reduction in the local pH of the catalyst surface in the MEA CO₂ electrolyzer. SiO₂ weathering is triggered by high local pH, and silicate ions transport the hydroxide ions near the AEM without CO₂ capture at the catalyst surface and revert to SiO₂. This cycle of SiO₂ enhances the mass transfer of CO₂, breaking the current density limitation for CO₂RR. Considering commercialization, the future goals of the silica-hydroxide cycle should be focused on enhancing the reversibility in real MEA electrolyzers. This chemical cycle strategy also observed on Earth holds a wide-ranging applicability across various systems, offering fresh insights for augmenting their efficiency and breaking their limitations.

Experimental procedures

Materials and chemicals

Silver nanopowder (Alfa Aesar, APS 20–40 nm, 99.9%), silicon dioxide nanopowder (Sigma-Aldrich, 10–20 nm, 99.5% trace metal basis), titanium(IV) oxide nanopowder (Sigma-Aldrich, 21 nm primary particle size (TEM), ≥99.5% trace metals), Nafion 5 wt% ionomer solution (Sigma-Aldrich), isopropanol (DAEJUNG, 99.5%), and iridium(IV) oxide (Alfa Aesar, premium, 99.99%; Ir 84.5% min) were used as received, without further purification. Carbon paper (Fuel Cell Store, Sigracet 39BB) and platinumized titanium screen (Fuel Cell Store) were employed as substrates for cathodic and anodic electrodes, respectively. These substrates were cut to the desired size using a homemade punch. KHCO₃, which was used as the electrolyte, was procured from Sigma Aldrich (ACS reagent, >99.7%). The AEM used was sourced from Dioxide Materials (Sustainion X37-50 grade RT).

Preparation of Ag-based electrode for cathode

The Ag black, Ag–TiO₂ and Ag–SiO₂ electrodes were fabricated by spraying catalyst ink onto a gas diffusion layer (GDL) at 70 °C. For the Ag–SiO₂ catalyst ink, 90 mg of commercial silver nanopowder, 15 mg of silicon dioxide nanopowder, 120 mg of Nafion ionomer solution, and 3.5 mL of isopropanol were ultrasonically mixed. Ag–TiO₂ catalyst ink was fabricated with same method as for Ag–SiO₂ catalyst ink except for adding TiO₂ instead of SiO₂. The catalyst ink for Ag black was prepared using the same method as for Ag–SiO₂, but without the silicon dioxide nanopowder.

Preparation of Ir-based electrode for anode

The electrode for OER, employed as the anode, was fabricated by spraying catalyst ink onto a platinumized titanium screen using a 70 °C hot plate. The catalyst ink was prepared by mixing 120 mg of commercial iridium oxide (Alfa Aesar) with 120 mg of 5 wt% Nafion ionomer solution and 4 mL of isopropanol. The electrode had an area of 10 cm², and the iridium oxide loading was set at 2 mg cm⁻².

Single-cell electrochemical CO₂ reduction reaction test

MEAs were fabricated using the catalyst-coated electrode method for electrochemical reactions, and the geometric

electrode area was set at 10 cm². The fabricated electrodes, along with commercial IrO₂-sprayed platinumized titanium screen electrodes (Alfa Aesar, with a target loading of 2.0 mg cm⁻²), were employed as cathodes and anodes in all single-cell tests, respectively. The AEM (Dioxide Materials, X37-50 Grade RT) underwent pretreatment in a 1 M KOH solution for 48 h and was rinsed multiple times with deionized water before use. Following this, a 0.1 M KHCO₃ solution, acting as the electrolyte, was introduced to the anode side *via* a pump. Concurrently, 200 sccm of humidified CO₂ gas at 80 °C, heated using a mantle, was introduced to the cathode side. The electrochemical tests were conducted using a VSP potentiostat (BioLogic, VMP3B-20) equipped with a booster up to 20 A. The CO₂RR was conducted over 18 min for each applied current density.

A gas chromatograph (GC, Agilent 7890A) was employed for product gas analysis at the GC outlet. A water trap was interposed between the GC and the cathode outlet. Argon gas (99.999%) served as the carrier gas. The GC was fitted with a flame ionization detector (FID) to detect hydrocarbons like CO, CH₄, and C₂H₄, and a thermal conductivity detector (TCD) for hydrogen (H₂) gas detection. A methanizer was utilized to enhance CO detection before routing to the FID. Measurements of the product gases commenced 9 min after initiation of CO₂RR for each current density setting. The FEs of H₂ and CO were computed using the following equation:

$$FE_{\text{product}} (\%) = \frac{i_{\text{product}}}{i_{\text{total}}} \times 100 = \frac{V_{\text{product}} \times Q \times \frac{2Fp}{RT}}{i_{\text{total}}} \times 100$$

where Q = flow rate of the product: F = faradaic constant (96 485 C mol⁻¹): p = pressure: T = room temperature (298 K): R = ideal gas constant (8.314 J mol⁻¹ K⁻¹).

The partial current densities of the products were determined from the volume of a specific product, as indicated by the GC peak.

Physical characterization

The size distribution and microstructure of the Ag black and Ag–SiO₂ electrodes were examined using high-resolution transmission electron microscopy (HR-TEM, Titan at 300 kV, FEI Co., USA). HR-TEM images, high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) images, and energy-dispersive spectroscopy (EDS) mapping were performed with a Talos F200X system (FEI). Raman spectroscopy was employed to characterize various electrodes, powdered samples, and the prepared solutions, and was conducted at room temperature with a 532-nm laser (Renishaw). XPS (X-ray photoelectron spectroscopy, Nexsa, ThermoFisher Scientific) was conducted with a base pressure of 2×10^{-8} mbar and a monochromated Al K α (1486.6 eV) X-ray source. All XPS spectra were calibrated using the C 1s peak (284.8 eV) as a reference.

In situ/operando Raman spectroscopy

In situ/operando SERS (surface-enhanced Raman spectroscopy) was conducted using a standard three-electrode system, a custom-made *in situ* Raman cell (details provided in Fig. S14,



ESI[†]), and a laser with a 532-nm wavelength. The working electrode was fashioned by spraying catalyst ink (containing Ag and Ag-SiO₂ powders) onto a GDL (Sigracet 39BB, SGL Carbon) substrate, which amplified the Raman spectrum signal. A Pt wire and Ag/AgCl (3 M NaCl) served as the counter and reference electrodes, respectively. The electrolyte was either 1 M KHCO₃ or 1 M CsHCO₃, and gaseous CO₂ was purged onto the working electrode at a flow rate controlled with a ball flowmeter. The electrochemical experiments were controlled using a potentiostat (CompactStat, Ivium Technologies, Eindhoven, Netherlands). The thin electrolyte layer enveloping the electrode facilitated Raman spectroscopic observation of the working electrode's surface.

Author contributions

C. L., S. K. and J. H. S. designed/conducted the experiments, analyzed the data and wrote the manuscript. M. H. H. performed Raman analysis of Ag black and Ag-SiO₂. Y.-J. K. contributed to the electrochemical cell tests. K. Y. L. and J. Y. C. provided an idea for the electrochemical analysis. H.-S. O. and W. H. L. supervised the research and wrote the manuscript. All authors reviewed the manuscript.

Conflicts of interest

There are no conflicts to declare.

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