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Enhancement of catalytic centres by RuO₂ addition to CuFe₂O₄ cathode catalyst for rechargeable lithium–air batteries: influence of CO₂ on Li–O₂ battery performances†

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Herein, the oxygen reduction reaction and oxygen evolution reaction (ORR/OER) kinetics of the inverse-spinel CuFe₂O₄ catalyst was enhanced *via* the addition of a very low quantity of RuO₂. It was found that minimal addition of RuO₂ resulted in an improvement in the limiting current density and onset potential, lower Tafel slope and good stability for the ORR/OER. Additionally, the CuFe₂O₄ cathode catalyst with the optimal RuO₂ content resulted in an outstanding Li–O₂ battery capacity of 14 250 mA h g^{−1}. Given that the presence of CO₂ poses a major challenge in achieving Li–air batteries at a practical level, the performance of the optimized catalyst under a strained Li–air condition and in pure CO₂ atmosphere (Li–CO₂ battery) was analyzed to understand its CO₂ tolerance and stability. It is crucial to understand the capability of the catalyst to decompose Li₂CO₃ formed as a stable discharge product from CO₂, which generally clogs the pores of the cathode catalyst. Thus, *in situ* impedance analysis and *ex situ* XRD technique were applied to decipher the fate of CO₂ in the reactions of Li–air/Li–CO₂ batteries. Moreover, stabilization to prevent the decomposition of the electrolyte was achieved in the presence of CO₂.

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1. Introduction

Currently, lithium–air batteries (LABs) are considered potential post-lithium-ion battery technology, having been realistically proven and thoroughly explored on the laboratory scale.¹ However, the commercialization of LABs has been impeded by a few key challenging issues such as dendrite growth, carbonate accumulation, electrolyte evaporation, electrolyte decomposition, and wide electrochemical potential window of the discharge product, which result in a poor cycling stability.^{2–4} The electrolyte decomposition is primarily due to the formation of corrosive reactive intermediates such as superoxide (O₂[−]) in the oxygen reduction reaction (ORR), resulting in electrolyte instability.^{5,6} In addition, the CO₂-sensitive nature of LABs results in the permanent accumulation of Li₂CO₃, a discharge product in the cathode pores, which deteriorates the performance of the battery in a short time.^{7,8} Two other factors generally found in any type of metal battery are the formation of dendrites, which pierce the separator, causing the eventual short circuiting of the battery, and poor energy efficiency

performance.⁹ The poor performance of Li–O₂ batteries is due to the limited decomposition of the discharge-product (Li₂O₂), leading to high oxygen evolution reaction (OER) overpotentials, necessitating a high charging potential to be applied to the battery.^{10,11} In the case of Li–air batteries, CO₂ is unavoidable given that it is present in the atmosphere at about 0.04%. Interestingly, it has been noticed that the presence of CO₂ is helpful in handling the superoxide. In the presence of CO₂, a more stable intermediate (C₂O₆^{2−}) was detected instead of the superoxide in an ethylene carbonate-diethyl carbonate-based (EC-DEC) electrolyte.¹² The Li–air battery showed a threefold increase in capacity in the presence CO₂ in the same work. It was later confirmed that the EC-DEC combination electrolyte is not suitable for Li–air batteries given that the irreversible consumption of this electrolyte occurs in the battery reactions. Thus, the reaction mechanism and intermediates formed in LABs are heavily dependent on the electrolyte used. For example, in a high donor number (HDN) DMSO-based electrolyte, the superoxide was converted to more stable intermediate species.¹³ Similarly, in a low donor number (LDN) electrolyte of tetraethylene glycol dimethyl ether (TEGDME), the presence of CO₂ improved the stability of the anode by forming a protective Li₂CO₃ film on it and that of the electrolyte by capturing the aggressive O₂[−].¹⁴ In general, the presence of CO₂ in Li–air batteries is found to be beneficial to address some of their challenges but at the cost of the accumulation of the Li₂CO₃ in

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the pores of the cathode, affecting their cycle life. Thus, regarding practical Li–air batteries, Li_2CO_3 re-oxidation is a prerequisite to achieve the aforementioned benefits of the presence of CO_2 . The decomposition of Li_2CO_3 can be achieved by using better electrocatalysts, providing a low polarization between the discharge and charge voltages. To achieve low polarization, there are reports of using external field-assisted strategies such as magnetic/force coupling and photo-assisted systems.^{15,16}

There have been reports on studies of spinel oxide cathode catalysts in Li–air batteries, exhibiting satisfactory performances. However, mostly these works neglected the influence CO_2 on the battery performances. An early work in this direction reported an initial capacity of about 3000 mA h g^{-1} using the CoMn_2O_4 spinel oxide at 0.2 mA cm^{-2} , but the capacity was essentially irreversible.¹⁷ Recently, spinel Mn_3O_4 /carbon composite nanofibers have shown an excellent catalytic performance for OER, which led to reduced overpotentials for discharge–charge reactions with better cycle stability in a hybrid electrolyte based Li– O_2 battery.¹⁸ Similarly, an Mn_3O_4 /rGO nanocomposite-based Li– O_2 battery exhibited enhanced ORR activity and delivered an initial discharge capacity of $16000 \text{ mA h g}^{-1}$.¹⁹ Moreover, in a hierarchical free-standing-type spinel MnCo_2O_4 electrode, a substantial capacity of $10520 \text{ mA h g}^{-1}$ was achieved at an applied 100 mA g^{-1} current density with a polarization voltage gap of 0.65 V .²⁰ Also, coprecipitation-synthesized spinel MnCo_2O_4 nanospheres exhibited superior ORR/OER electro-catalytic activities. Moreover, in an Li– O_2 battery, this MnCo_2O_4 cathode catalyst showed a polarization voltage of 0.85 V , while delivering a discharge capacity of 8518 mA h g^{-1} at 100 mA g^{-1} , with stable 20 charge–discharge cycles.²¹ A cobalt and iron-based spinel oxide in N-doped graphene ($\text{Co}[\text{Co}_x\text{Fe}]_4/\text{NG}$) showed a discharge capacity of $13312 \text{ mA h g}^{-1}$ at 50 mA g^{-1} .²² In a report, the crystal structure of LaSrNiO was changed from perovskite to spinel phase by doping different amounts of Sr, which in turn promoted both the ORR/OER activities of the catalyst. This Sr-doped LaSrNiO catalyst provided a discharge capacity of 9000 mA h g^{-1} in an Li– O_2 battery.²³ In our recent study, the $\text{Co}/\text{CoFe}_2\text{O}_4$ spinel oxide was tested as an air-breathing electrode and exhibited a capacity of 4320 mA h g^{-1} at 100 mA g^{-1} in an Li– O_2 battery.²⁴ The bimetallic spinel NiFe_2O_4 catalyst was also able to provide substantial capacities of 12460 and 3450 mA h g^{-1} at 100 and 1500 mA g^{-1} , respectively, suggesting its robust rate-capability.²⁵ The potentials of the spinel oxide catalyst in an Li–air battery was also evident in another recent work, where the surface e_g occupancy of the Ni–Co spinel oxide structures was optimized, supplying a capacity of $13759 \text{ mA h g}^{-1}$ with a lower charging overpotential owing to its finer OER performance.²⁶

Here, in this work, the performance of the inverse spinel oxide CuFe_2O_4 was examined for ORR/OER followed by the optimization of its performance *via* the addition of the minimal quantity of RuO_2 , promoting the reactions by the increasing the number of catalytic centres. The application of CuFe_2O_4 has been reported elsewhere, which achieved an inferior capacity as low as 677 at 0.1 mA cm^{-2} .²⁷ In contrast, the present work achieved an improved capacity, which was further enhanced by

the addition of a small amount of RuO_2 to CuFe_2O_4 , resulting in an increase in the number of catalytic centres. Furthermore, the presence of the unavoidable CO_2 in the practical Li–air battery was found to be beneficial to the electrolyte stability over the optimized catalyst. Also, the ability of the catalyst to decompose the discharge products such as $\text{Li}_2\text{C}_2\text{O}_4/\text{Li}_2\text{CO}_3$ during charging was analyzed systematically in the Li–air and Li– CO_2 batteries. The products formed on the catalyst pores during the discharge reactions over multiple cycles were subjected to *in situ* impedance analysis and *ex situ* X-ray diffraction analysis to understand the detailed charge–discharge mechanism and influence of CO_2 on the Li–air battery performances.

2. Experimental

2.1 Synthesis of $\text{RuO}_2@/\text{CuFe}_2\text{O}_4$ electrocatalyst

The $\text{RuO}_2@/\text{CuFe}_2\text{O}_4$ electrocatalyst was synthesized in a three-step process. The sol–gel synthesis of CuFe_2O_4 was followed by probe-sonication with RuO_2 to form a homogenous composite. Later on, the composite was vacuum-heated overnight at $120 \text{ }^\circ\text{C}$ to achieve improved crystallinity. Initially, $4 \text{ mmol Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ and $8 \text{ mmol Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ were mixed in deionized water (20 mL) and stirred thoroughly on a magnetic stirrer. Subsequently, this nitrate solution was added to a glycine solution, which served as a gelling agent. This was accompanied by a change in the color of the solution, indicating the complex formation of metal-ions with glycine. The solution was kept stirring overnight at $30 \text{ }^\circ\text{C}$, generating a thick gel. On further heating of the gel at about $160 \text{ }^\circ\text{C}$, an autocombustion reaction occurred, resulting in the formation of the CuFe_2O_4 powder.

Later, a finite quantity of finely ground CuFe_2O_4 powder and 2, 5 and 7 wt% RuO_2 were added to a beaker containing 100 mL deionized water. Then, this mixture was subjected to probe sonication for 1 h for homogenous mixing. The resulting dispersion was centrifuged and collected. Finally, the sample was vacuum dried $120 \text{ }^\circ\text{C}$ overnight. The obtained samples were named RCFO-2, RCFO-5 and RCFO-7 according to the 2, 5 and 7 wt% of RuO_2 added to the pristine CuFe_2O_4 , respectively. The pristine CuFe_2O_4 without the addition of RuO_2 was labelled as CFO.

2.2 Material and electrochemical characterization

The different RCFO composite samples were characterized by diverse techniques for a thorough understanding. The structural identification of the samples was performed using a $\text{Cu K}\alpha$ radiation configured X-ray diffractometer (XRD, Rigaku, SmartLab) at 5° min^{-1} scan rate. The chemical compositions of the samples were evaluated *via* X-ray photoelectron spectroscopy (XPS, Thermo Fisher, ESCALAB Xi⁺). Further, the morphological feature of the optimal sample was probed using a scanning electron microscope (VEGA3TESCAN) together with energy-dispersive X-ray analysis (EDAX, Bruker Nano D-12480).

2.3 ORR/OER/ CO_2 RR kinetics measurement on the catalysts

The generated catalysts were subjected to ORR/OER linear sweep analysis (LSV) to screen the best-performing catalyst on

a rotating ring-disc electrode (RRDE), where the disc and ring were made of glassy carbon and platinum, respectively. Each of the RCFO composite catalysts was made into catalyst ink using Super P carbon in a solvent mixture of deionized water, isopropyl alcohol (150 μL each) and 5 wt% Nafion binder (10 μL) solution. The composite-catalyst and conductive carbon (Super P) were mixed in a ratio of 7 : 3 in each case. Finally, the mixture was subjected to bath-sonication to obtain the catalyst inks. The RRDE having a collection efficiency of 0.37 was polished and prepared for the analysis by drop-casting the ink (6 μL) on its polished disc (0.13 cm^2), followed by drying overnight. Later, the catalyst-coated RRDE in an O_2/Ar -saturated 0.1 M KOH electrolyte was subjected to cyclic voltammetry and linear sweep voltammetry (CV and LSV) to examine its ORR/OER characteristics in a three-electrode arrangement. The arrangement was comprised of platinum counter and KCl-saturated Ag/AgCl reference electrodes alongside the RRDE working electrode. The ORR and OER studies were executed at the sweep rate of 10 mV s^{-1} in the potential window of 0 to -1.0 and 0 to 1 V vs. Ag/AgCl, respectively. The stability of the RCFO catalysts was investigated by means of 250 cycles of LSV technique.

The aforementioned electrochemical examinations were executed in a three-electrode framework coupled with a Biologics SP-150 electrochemical workstation. The potentials measured during these studies were against the Ag/AgCl scale, which was switched to the more broadly accepted reversible hydrogen electrode (RHE) scale according to eqn (1), as shown below:

$$E_{\text{RHE}} = E_{\text{Ag/AgCl}} + 0.098 + 0.059 \text{ pH} \quad (1)$$

where E_{RHE} is the potential denoted in the RHE scale and $E_{\text{Ag/AgCl}}$ denotes the measured potential on the Ag/AgCl scale. The pH of the 0.1 M KOH was 14 and that of the 0.1 M KHCO_3 was 8.4 and 6.8 under Ar and CO_2 -saturated conditions, respectively.

2.4 Fabrication and testing of Li– O_2 , Li–air and Li– CO_2 batteries

After the initial ORR/OER performance screening of the catalysts, the RCFO-5 catalyst-coated electrode was employed as the Li– O_2 battery cathode for testing under a pure O_2 atmosphere maintained in a custom-made acrylic box to understand the ORR/OER activity in the real-life battery, and in a mixture of O_2 and CO_2 (90 : 10 v/v), corresponding to an Li–air battery. In the case of the Li– O_2 batteries, the ORR/OER activities correspond to Li_2O_2 formation during discharging and its decomposition while charging, respectively. The impact of CO_2 on the battery performance of the catalyst was further analyzed by running the battery in a pure CO_2 atmosphere, corresponding to an Li– CO_2 battery. All these batteries in the form of CR-2032 coin cells were assembled inside a Nichwell glovebox (α -1500u) with a moisture and O_2 content below 1 ppm. The cathode of the battery was made by coating a slurry of the RCFO-5 catalyst on a circular disc (diameter = 16 mm) of hydrophobic carbon paper, keeping the mass loading about 1 mg. The cathode slurry was made by grinding the RCFO-5 catalyst, conductive Super P carbon and the binder (poly(vinylidene fluoride) PVDF) in *N*-methyl-2-

pyrrolidone solvent (NMP) in a wt% ratio of 70 : 20 : 10. The coated and dried cathode was separated from the lithium chip anode by a Whatman glass microfiber filter paper in the coin cell battery. The electrolyte employed in the batteries was lithium trifluoromethane sulfonate in tetraethylene glycol dimethyl ether (1 M LiCF_3SO_3 in TEGDME). The fabricated batteries were subjected to open circuit voltage (OCV) decay employing a Neware battery analyzer (CT-4008-5V50mA) interfaced with a computer. The galvanostatic charge–discharge (GCD) analysis of the batteries was performed in the potential region of 2–4.5 V vs. Li/Li^+ . The cyclic voltammograms (CV) were also obtained in pure O_2 , O_2/CO_2 (10 : 90 v/v) and pure CO_2 atmospheres in the same potential region. An electrochemical workstation (Biologics BCS-805) was employed to carry out these electrochemical tests at room temperature. The required gas atmosphere was maintained in a custom-made acrylic box, having the coin cells placed in it, maintaining the gas flow at ~ 1 psi. The cycling stability of the batteries was investigated using GCD cycling with the capacity limited to 1000 mA h g^{-1} at 100 mA g^{-1} current density in the required gaseous atmosphere. Moreover, the electrochemical impedance analyses of the cathodes in the frequency range of 0.1 Hz to 10 kHz were carried at various of discharge depth using a workstation.

3. Results and discussion

3.1 Crystal structure and morphology

The formation of CuFe_2O_4 and its composites using the solgel synthetic route followed by sonication and vacuum-heating was confirmed *via* XRD assessment. The XRD patterns of the composites having distinct ratios of RuO_2 and CuFe_2O_4 were recorded, as shown in Fig. 1. The ensemble of Bragg peaks observed at 2θ values of 18.3°, 30.1°, 35.5°, 37.1°, 43.1°, 53.6°, 57.1°, 62.7°, 65.9° and 67.0° are indexed to the cubic inverse spinel CuFe_2O_4 (JCPDS # 98-10-5232). An average crystallite size of 18.3 nm was calculated based on the Bragg peaks of the CuFe_2O_4 according to the Scherrer equation, given in eqn (2) as follows:

$$\text{Crystallite size } (D) = K\lambda/\beta \cos \theta \text{ \AA} \quad (2)$$

where K represents the shape factor having a dimensionless value of 0.9, λ is the wavelength of X-ray used (0.15418 nm), β stands for the fitted full width at half maximum (radian) of the Bragg peak and θ indicates the diffraction angle of the Bragg peak. Although the XRD data show that the major phase is CuFe_2O_4 , the synthetic process yielded a minor metallic Cu phase in the sample due to the reduction reaction facilitated by glycine as the reducing agent. The introduction of the minimal quantity of RuO_2 is reflected in the recorded XRD pattern. The (110) plane of RuO_2 is visible at around 28° in all the composite samples.²⁸ The intensity of the RuO_2 Bragg peaks became more pronounced with an increase in the amount of RuO_2 added to the composites. The successful synthesis of CuFe_2O_4 was further verified by Raman spectroscopy. Its Raman profile, as shown in Fig. S1,† reveals peaks at 190, 290, 460, 530, and 690 cm^{-1} , which correspond to the T_{2g} (1), E_g , T_{2g} (2), T_{2g} (3),

and A_{1g} vibrational modes of CuFe_2O_4 , respectively.²⁹ However, the Raman instrument could not detect the presence of RuO_2 , which may be due to its tiny amount in the composite.

The XPS technique is a powerful analytical tool, which was used to extract fine information about the elements and their elemental environments in the RCFO composites. Also, it was employed to corroborate the successful incorporation of RuO_2 in the CuFe_2O_4 matrix. The three variations of the RCFO composites were subjected to XPS analysis. In particular, the XPS deep scanning was concentrated around the binding energy value where the Ru 3d peak appeared and the deconvoluted data are shown in Fig. 2(a)–(d). For comparison, deep-scan was also performed on the pristine CFO sample. The Ru 3d spectra of all three RCFO composites exhibit peaks of Ru $3d_{5/2}$ (280.8 eV) together with its satellite peak (282.4 eV) and Ru $3d_{3/2}$ (285 eV) with its satellite (286.6 eV).³⁰ It is known that the X-rays irradiated in this binding energy region are also capable of emitting C 1s electron together with the Ru 3d electrons, as reflected in the deconvoluted scan profiles. The peak observed in the deconvoluted spectra at about 284.9 eV matches C 1s of the C=C bond and the peak at about 288.4 eV corresponds to the C=O carbonyl group.³¹ The Ru 3d peaks ($3d_{3/2}$ or $3d_{5/2}$ and satellite peaks) are clearly missing in the C 1s XPS profile of the pristine CFO catalyst, as displayed in Fig. 2(d), unlike in the RuO_2 -embedded composites. The Ru 3d peak showed a high intensity for all the RuO_2 -embedded composites, thus confirming the

successful incorporation of RuO_2 in CuFe_2O_4 . The XPS technique was also exploited to determine the chemical environments of the elements present in the pristine CuFe_2O_4 sample. The survey spectrum, as illustrated in Fig. S2(a),† verifies the existence of Cu, Fe, O and C elements in the pristine CuFe_2O_4 catalyst. Moreover, the deconvoluted spectra of Cu 2p shown in Fig. S2(b)† display two notable peaks at 933.8 eV and 953.5 eV, corresponding to Cu $2p_{3/2}$ and Cu $2p_{1/2}$, respectively, together with their satellite peaks. These peak positions are consistent with the Cu^{2+} oxidation state.³² The Fe 2p spectra shown in Fig. S2(c)† have two main peaks at 710.5 eV (Fe $2p_{3/2}$) and 724 eV (Fe $2p_{1/2}$) together with two satellite peaks corresponding to Fe^{3+} . Each of these main peaks could be deconvoluted into two sets corresponding to Fe^{3+} occupying the tetrahedral and octahedral (t_d & o_d , respectively) positions of the inverse spinel, as seen in Fig. S2(c).†³³ Similarly, the O 1s spectra, as illustrated in Fig. S2(d),† were deconvoluted to lattice oxygen and surface-adsorbed oxygen, matching the binding energies of 529.6 and 531.1 eV, respectively.³² The C 1s profile in Fig. S2(e)† indicates the presence of C=C, C–O, and C=O bonds corresponding to the deconvoluted peaks at 284.8, 286.0, and 288.4 eV, respectively. Thus, the XPS analyses clearly confirm the presence of RuO_2 on CuFe_2O_4 in the RuO_2 -embedded CuFe_2O_4 samples. The presence of Fe in the O_d and t_d positions was also confirmed.

Scanning electron microscopy (SEM) was employed to explore the morphological features of the composite samples. A representative composite, RCFO-5, was observed under various magnifications, as illustrated in Fig. 3(a) and (b). It was observed that this sample has a uniformly distributed microporous structure with crystallite grains having a size of about ~500 nm. The SEM images at higher magnifications revealed that the individual grains of CuFe_2O_4 have an average size of about 500 nm. It was also visible that some smaller grains are distributed on these CuFe_2O_4 grains, indicating the embedded RuO_2 grains. In the SEM images, as shown in Fig. S3,† of the pristine CuFe_2O_4 , its surface consists of uniform grains and pores. Also, CuFe_2O_4 aggregates with a large size were seen. Here, the smaller grains of RuO_2 were absent, unlike in the case of the RCFO-5 sample, where distinct tiny grains of RuO_2 were visible. Furthermore, the quantitative elemental composition of the RCFO-5 composite was examined by EDAX analysis. The EDAX profile shown in Fig. 3(d) projects the elemental weight percentages of Cu, Fe, O and Ru as 24.58%, 12.62%, 62.41% and 0.39%, respectively. Thus, the above-discussed XRD, SEM/EDAX and XPS results confirmed that RuO_2 was embedded in the CuFe_2O_4 matrix. It should be noted that the magnetic nature of the sample prevented morphological examination using high-resolution transmission electron microscopy (HR-TEM).

3.2 Electrocatalytic activities towards ORR and OER kinetics

A rechargeable Li–air battery works on the principle of utilizing the energy released from ORR during the cell-discharge, forming lithium oxide as the discharge product. This oxide-product is oxidized by means of OER back to metallic lithium and molecular oxygen upon charging, making them ready for discharge once again. Thus, a bifunctional catalyst for use in Li–

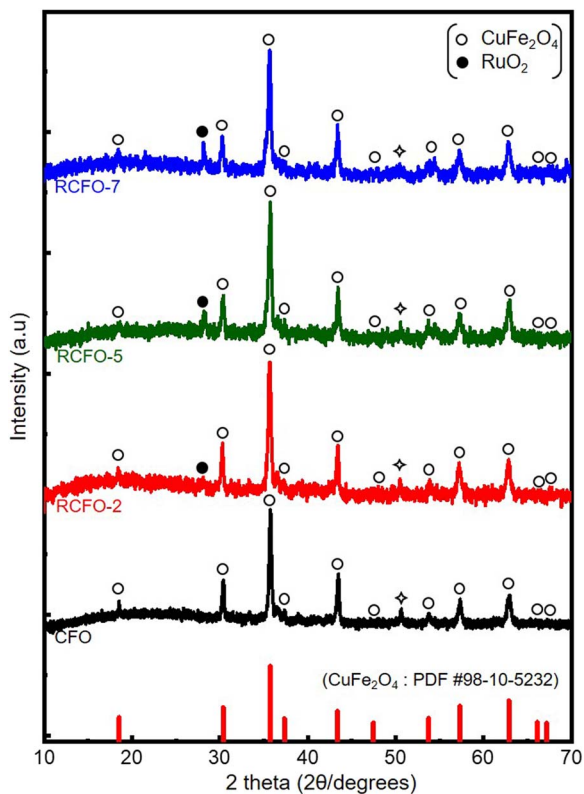


Fig. 1 XRD patterns recorded for pristine CuFe_2O_4 (CFO) and various RuO_2 - CuFe_2O_4 (RCFO-2, 5, 7) composite samples. The \diamond symbol represents the minor phase elemental copper.

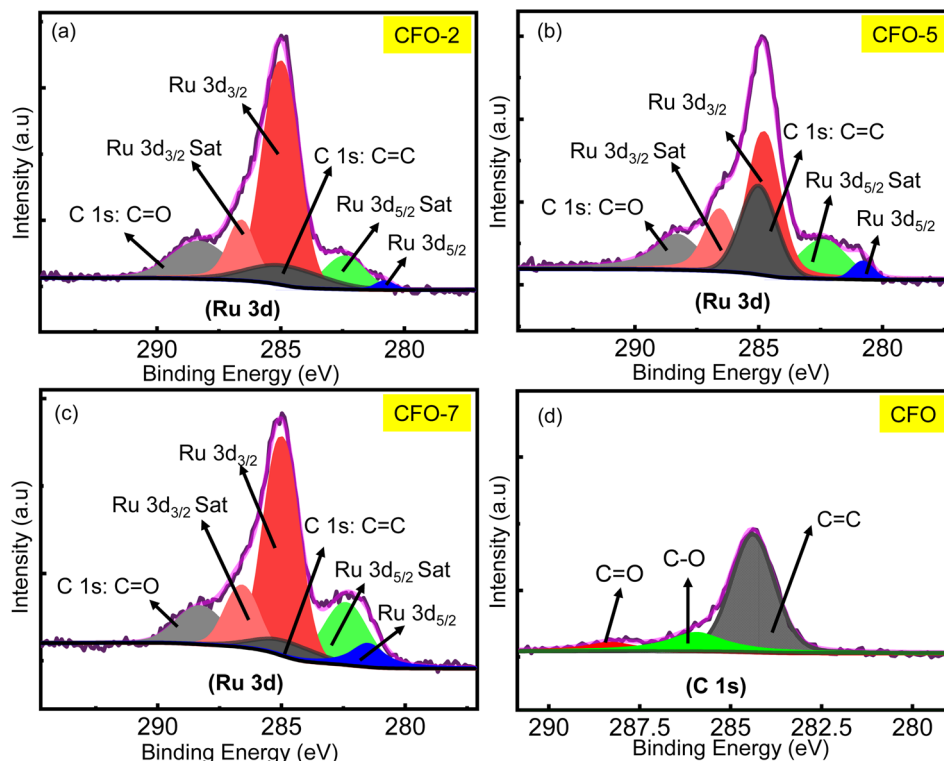
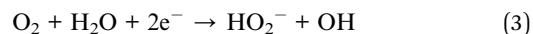


Fig. 2 (a–d) Deconvoluted Ru 3d and C 1s XPS spectra recorded for the RCFO composites and the pristine CFO samples.

air batteries must effectively catalyze the ORR and OER. Consequently, the catalytic kinetics of the synthesized CFO and the RCFO composite catalysts were thoroughly explored towards ORR/OER on a rotating ring (Pt) disc (glassy-carbon-surface) electrode. The glassy carbon of the disc was coated with the catalyst ink made from each of the samples and subjected to cyclic voltammetry (CV) studies. The ORR CV analyses at a scan rate of 10 mV s^{-1} were carried out in 0.1 M KOH electrolyte in the potential range of $0\text{--}1 \text{ V vs. RHE}$. These CV studies were conducted under Ar -saturated and O_2 -saturated KOH conditions to extract the exact ORR activities, as shown in Fig. 4(a–d). In the case of all the catalysts, a peak in the potential region of around $0.65\text{--}0.7 \text{ V vs. RHE}$ emerged in the O_2 -saturated electrolyte, corresponding to the ORR. These peaks are copiously significant compared to the current response registered in the Ar -saturated electrolyte, implying the pronounced ORR activity of the catalysts. Among the catalysts, the CV profile recorded for the RCFO-5 composite catalyst noticeably exhibited the highest increased current response for the ORR, indicating its superiority. It can be noticed that all the composite catalysts containing RuO_2 exhibited an enhanced current response, indicating the positive influence of introducing RuO_2 in the CuFe_2O_4 catalyst. The dominant ORR activity of the RCFO-5 composite catalyst was further analyzed in the following RRDE studies. The LSV profiles of the catalysts were recorded on the RRDE at 400 rpm at a scan rate of 10 mV s^{-1} and denoted as disc-currents in Fig. 5(a). The onset potential and the limiting current densities of these catalysts were extracted from the LSV profiles and plotted in Fig. 5(b). It can be observed that the

limiting current density and the onset potential of the RCFO-5 catalyst are the highest among the composite catalysts, confirming its robustness. Additionally, it is noticeable that the pristine CuFe_2O_4 had the lowest limiting current density and worst onset potential. The best catalyst showed an onset potential of 0.77 V and a limiting current density of 5.5 mA cm^{-2} . Moreover, the Tafel plots of these four catalysts were also constructed from their LSV profiles and are depicted in Fig. 5(c). Notably, the lowest Tafel slope of 111 mV dec^{-1} was achieved on the RCFO-5 catalyst compared to RCFO-2 (115 mV dec^{-1}) and RCFO-7 (140 mV dec^{-1}), and all the composites showed an improved Tafel slope compared to the pristine CFO (156 mV dec^{-1}). Thus, the reaction kinetic metrics of the RCFO-5 catalyst were further analyzed. The catalysis of the ORR on the catalysts can proceed either through a $4e^-$ or $2e^-$ transfer mechanism, as shown below:



The mechanism involving $2e^-$ transfer proceeds *via* sluggish kinetics given that it involves a peroxide intermediate, while the $4e^-$ transfer pathway results in direct H_2O formation. To estimate the amount of H_2O_2 formed during the ORR, a fixed potential of 0.6 V vs. RHE was applied to the Pt ring electrode, while running the LSV technique on the catalyst-coated disc of the RRDE. The H_2O_2 formed on the disc will be subjected to

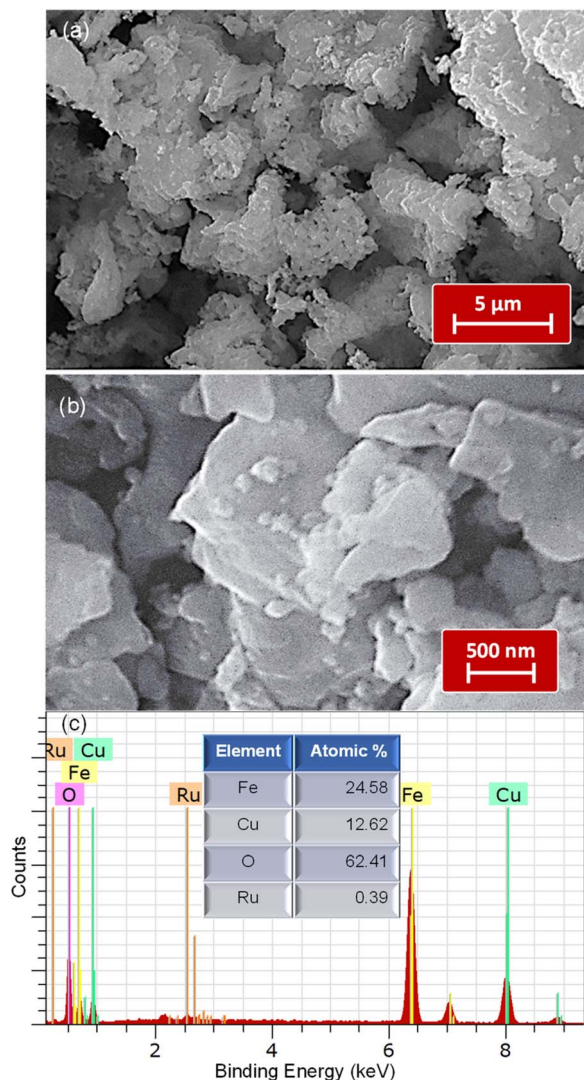


Fig. 3 (a and b) SEM images at different magnifications and (c) EDAX profile recorded for the RCFO-5 composite sample. Inset: Quantitative composition of the various elements analyzed from the EDAX.

oxidation on the Pt ring at this potential and the current response detected by the ring will be proportional to the amount of the H_2O_2 formed. All the generated catalysts including RCFO-5 showed no significant ring current response, as evidently seen in Fig. 5(a), suggesting the preferred $4e^-$ transfer pathway. According to the ring current obtained in the LSV profile of the RCFO-5 catalyst, the quantitative measurement of the H_2O_2 and number of electrons transferred (n) during the ORR was carried out using the following eqs:³⁴

$$n = \frac{4i(d)}{i(d) + \frac{i(r)}{N}} \quad (5)$$

$$\text{H}_2\text{O}_2\% = \frac{200 \frac{i(r)}{N}}{i(d) + \frac{i(r)}{N}} \quad (6)$$

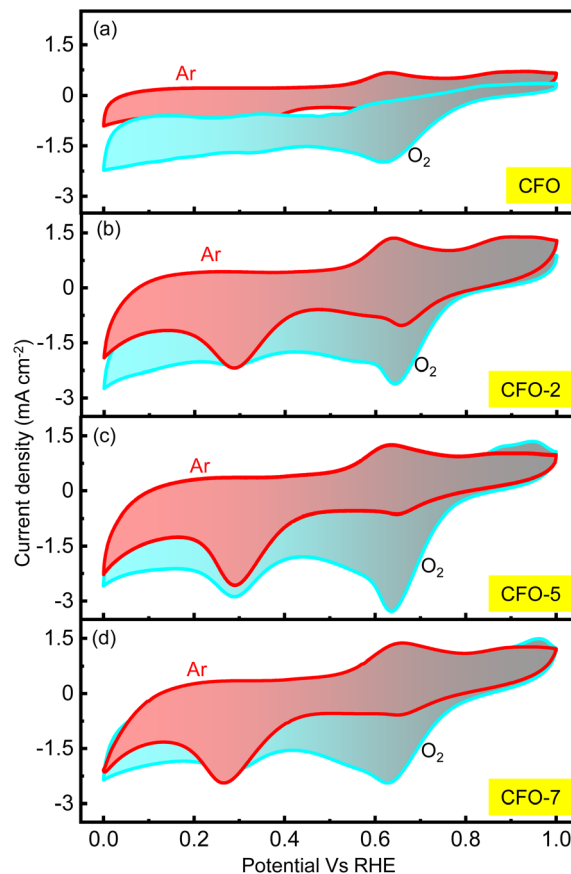


Fig. 4 (a–d) ORR CV profiles in O_2 - and Ar-saturated 0.1 M KOH electrolytes recorded for CFO and the composite catalysts at the scan rate of 10 mV s^{-1} .

where the disc- and ring-current are represented by $i(d)$ and $i(r)$, respectively, and $N = 0.37$ is a unitless value known as the RRDE collection efficiency. Based on these equations, a plot was constructed, as depicted in Fig. 6(a). It can be seen that the number of electrons transferred is close to 4 and the percentage of the H_2O_2 formed is as low as 1%. Moreover, the Koutecky–Levich (K–L) plots of the RCFO-5 catalyst were constructed using the K–L equation in eqn (7) to substantiate the $4e^-$ transfer during the ORR, as follows:

$$\frac{1}{i} = \frac{1}{i_L} + \frac{1}{i_K} = \frac{1}{B\omega^{1/2}} + \frac{1}{i_K} \quad (7)$$

where i , i_K and i_L are the total current density, kinetic and limiting current densities (mA cm^{-2}). ω is the speed at which the electrode is rotated (rad s^{-1}) and the constant B is the Levich constant, presenting information about the dissolved O_2 concentration, diffusion coefficient, and kinematic viscosity. To construct the K–L plots, the LSV plots at different electrode rotation speeds (400 to 1600 rpm) were obtained at a scan rate of 10 mV s^{-1} , as shown in Fig. 6(b). An increase in the current density was seen as the rotation speed increased, indicating a higher rate of the ORR. This trend is expected due to the improved mass transportation of the dissolved O_2 molecules under the hydrodynamic force during rotation. It was noted that

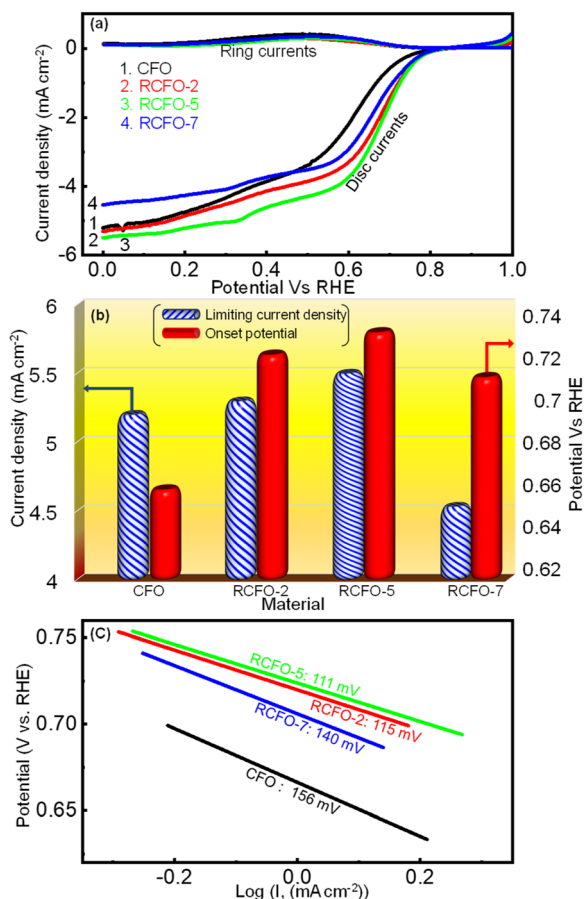


Fig. 5 (a) ORR ring-disc LSV profiles obtained for various catalysts in 0.1 M KOH electrolyte at the scan rate of 10 mV s⁻¹, (b) comparison of ORR onset potential and limiting current densities and (c) Tafel plots obtained for the various catalysts.

the ORR proceeded at mixed-diffusion rates at higher potentials and in a kinetically limited manner in the low potential range of 0–0.5 V given that the current response in this region is nearly stationary. Subsequently, the K–L plots were plotted, as shown in Fig. 6(c), at different potentials of 0.4, 0.5, 0.6 and 0.7 V. The obtained K–L plots were seen to be linear and parallel to each other, as observed for first-order ORR reaction kinetics with 4e⁻ transfer.

Moreover, the OER kinetics of CFO and the composite catalysts were also evaluated utilizing the LSV technique. The LSV profiles in the potential region of 1–2 V vs. RHE, as shown in Fig. 7(a), were obtained at a sweep rate of 10 mV s⁻¹ in 0.1 M KOH electrolyte saturated with argon gas. The onset potentials and the limiting current densities of CFO and the composite catalysts were extracted from these LSV profiles and plotted in Fig. 7(b). It is known that a lower onset potential for OER indicates a better catalyst. According to Fig. 7(b), it can be seen that the onset potentials of the composites moved to lower potentials as the amount of RuO₂ increased in the composites although the difference between the RCFO-5 and RCFO-7 catalysts was marginal. Notably, as seen in the ORR analyses, the RCFO-5 catalyst exhibited superior OER activity on further

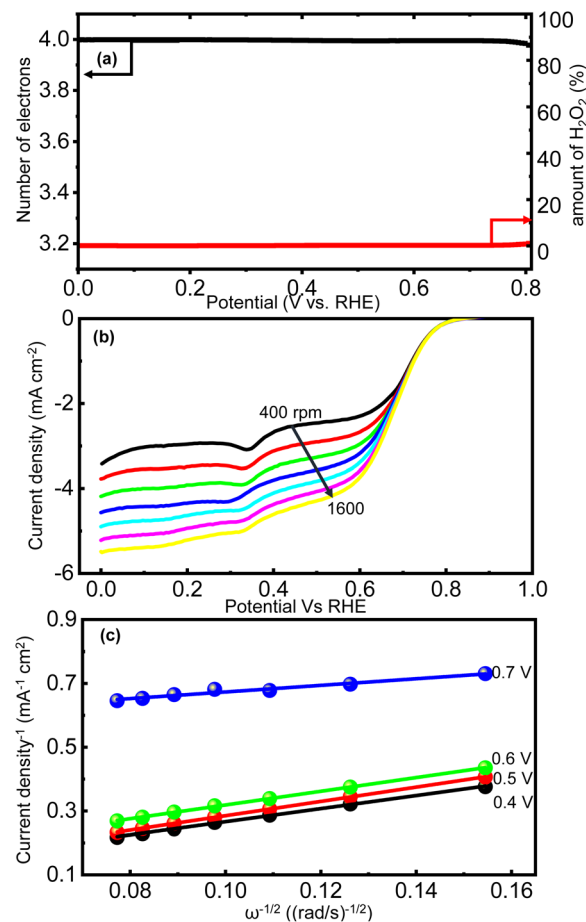


Fig. 6 (a) Quantification of the H₂O₂ formed and the number of electrons transferred in ORR, (b) ORR LSV profiles at varying electrode rotation speeds of RRDE at a sweep rate of 10 mV s⁻¹, and (c) K–L plots at multiple potentials derived from the speed-dependent LSV plots of the RCFO-5 composite electrocatalyst recorded in 0.1 M KOH.

analyses. For example, the limiting current density achieved on the RCFO-5 catalyst (21 mA cm⁻²) was greater than that on the other composite catalysts, as shown in Fig. 7(b). Moreover, the Tafel slopes of the catalysts were calculated from the constructed Tafel plots, as shown in Fig. 7(c). The RCFO-5 catalyst delivered the lowest Tafel slope of 160 mV dec⁻¹ compared to the RCFO-2 (162 mV dec⁻¹) and RCFO-7 (166 mV dec⁻¹) catalysts. In all these analyses, the RuO₂ composite exhibited improved metrics compared to the pristine CuFe₂O₄ catalyst. The presence of RuO₂ with its loading as low as 2 wt% could reduce the Tafel slope of CuFe₂O₄ (230 mV dec⁻¹), indicating the effectiveness of embedding the RuO₂ in the CuFe₂O₄ matrix for enhanced OER performance.

The stability tests for CFO and the RCFO composite catalysts towards ORR and OER were conducted by running LSV for 250 cycles at a scan rate of 10 mV s⁻¹. The obtained LSV plots for the first and 250th ORR cycles are shown in Fig. 8(a–d). The RCFO-5 composite catalyst was found to be the most stable, given that its LSV curves obtained in the first and 250th cycles overlapped. Similarly, in the OER stability tests, the RCFO-5 composite catalyst maintained a very high limiting current density of 17

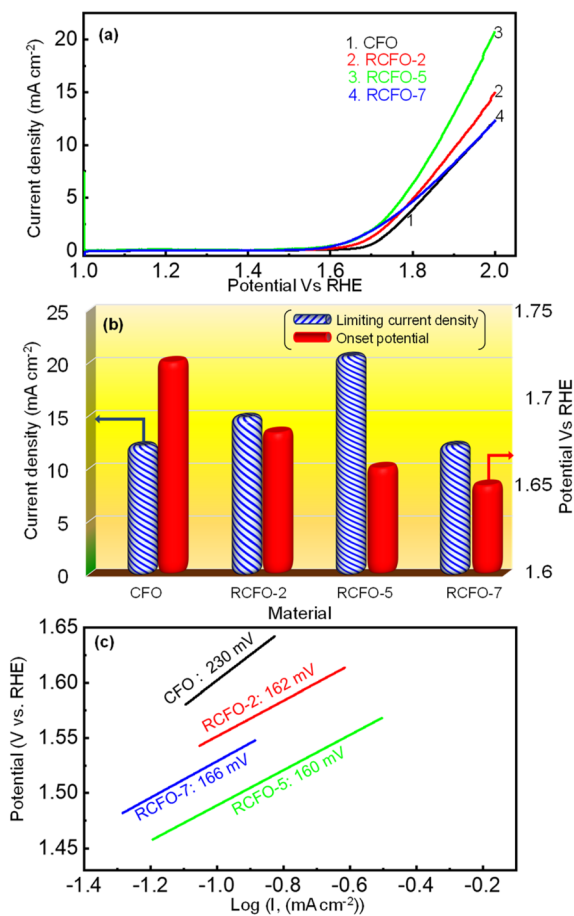


Fig. 7 (a) OER LSV profiles of the composite catalysts and the pristine CFO catalyst recorded at 10 mV s^{-1} , (b) comparison of OER onset potential and limiting current densities and (c) OER Tafel plots with slope recorded for the various catalysts in 0.1 M KOH .

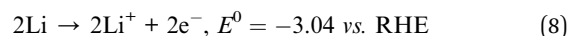
mA cm^{-2} at the end of the 250th cycle. In contrast, the CFO catalyst could only achieve a limiting current density of about 12 mA cm^{-2} at the beginning, which deteriorated to as low as 4 mA cm^{-2} in the 250th cycle. Although the other RuO_2 -embedded composites showed better activity than the pristine CFO catalyst, they did not match the performance of the RCFO-5 catalyst. These results are confirmed by the LSV profiles from the stability tests, as shown in Fig. 8(e–h).

3.3 Performances of RCFO-5 cathode catalyst in Li–O₂/Li–air/Li–CO₂ batteries

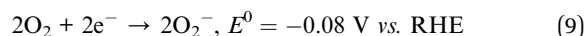
Given that the RCFO-5 composite catalyst showed superior ORR and OER activities in the RRDE experiments, the composite catalyst was further applied in real-life Li–O₂, Li–air and Li–CO₂ battery studies. Initially, the enhanced electrochemical performance of the RuO_2 -based composite catalysts compared to the pristine catalyst in the Li–O₂ battery was evaluated by recording a GCD profile at a current rate of 500 mA g^{-1} . The result, as displayed in Fig. S4,† shows that the RCFO-5 catalyst exhibits lowest overpotential in the charge–discharge profiles, confirming its enhanced RRDE performance. The effect of introducing the RuO_2 in CuFe_2O_4 improved the electrochemical

performance of the Li–O₂ battery, as observed in the reduction of the overpotential (Fig. S4†) for the Li–O₂ battery. Each of these batteries in the form of CR-2032 coin cells were initially tested for their ability to hold the open circuit voltage (OCV) when opened to the respective atmospheres of pure O₂, mixed atmosphere of O₂/CO₂ (10%) and pure CO₂. The purpose of performing the additional tests on the Li–O₂ battery was to examine the influence of CO₂ its performance, which led to the optimization of the practicality of the Li–O₂ battery. The required environment of these gases was generated in a home-made container sealed with an air-tight balloon. A pictorial representation of this container is shown in the inset of Fig. 9(a). The obtained self-decay profiles recorded for the Li–O₂, Li–air and Li–CO₂ batteries showed their capability to hold a stable voltage for more than six days, as shown in Fig. 9, demonstrating their potential in practical applications. The inset of Fig. 9(c) shows the observed OCV of the Li–air battery in a digital multimeter. Thus, further studies were conducted using each battery. To estimate the capacity that can be derived from these batteries having the RCFO-5 cathode catalyst, rate capability tests were conducted by means of the galvanostatic charge–discharge (GCD) technique at varying current densities. To derive any capacity from a rechargeable Li–O₂ battery, a potential cathode catalyst must catalyze the reduction of atmospheric O₂ during the discharge and oxidation of Li_2O_2 (discharge-product) to metallic lithium and molecular O₂ during charging. The reactions happening upon the aforementioned discharging of the Li–O₂ battery are given below, whereas the reactions in the reverse direction will be forced to happen upon charging:

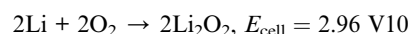
At anode:



At cathode:



Overall reaction:



Thus, to estimate the capacity of the cell at various current rates, the Li–O₂ battery having the RCFO-5 catalyst-coated cathode and a lithium chip as the anode was subjected to discharging and charging in the potential window of 2.0–4.5 V vs. Li/Li⁺. The cell was initially forced to discharge at the current densities of 200, 300 and 500 mA g^{-1} , and then charged at the same current rate, and the obtained GCD curves are depicted in Fig. 9(b). At the current density of 200 mA g^{-1} , an excellent discharge capacity of $14\,250 \text{ mA h g}^{-1}$ was achieved. Furthermore, even at higher current densities of 300 and 500 mA g^{-1} , capacities as high as 6000 and 1500 mA h g^{-1} were attained, respectively. These discharge curves have the advantageous trait of steady discharge voltage of roughly 2.5 V over most of the capacity. To decompose the discharge product Li_2O_2 on charging, a high voltage beyond 4.0 V is generally required due the stable and insulative nature of the discharge product Li_2O_2 . Here, on the

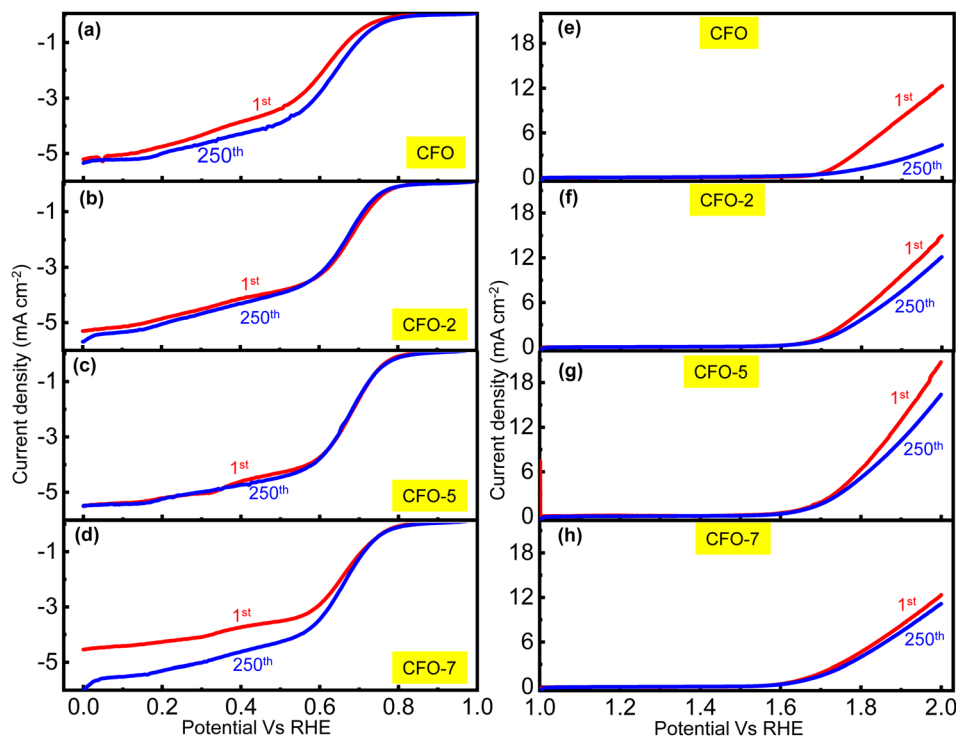


Fig. 8 (a–d) ORR cycle-life data and (e–h) OER cycle-life analysis of the CFO and the composite catalysts in 0.1 M KOH electrolyte recorded at 10 mV s⁻¹ scan rate.

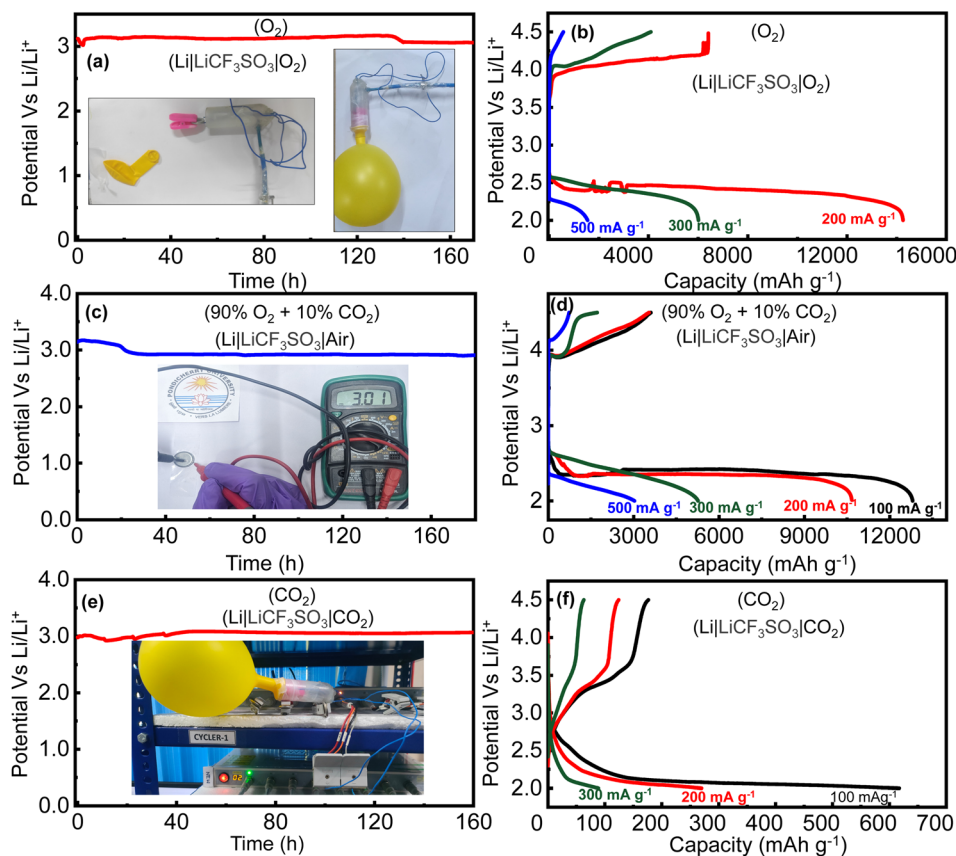


Fig. 9 (a, c, e) OCV decay profiles of the Li–O₂, Li–air and Li–CO₂ batteries with inset showing the CR2032 Li–air coin cell connected to a voltmeter displaying the cell voltage and (b, d, f) rate capability analyses of the Li–O₂, Li–air and Li–CO₂ batteries using the RCFO-5 cathode catalyst.

RCFO-5 composite catalyst, the charge voltage plateau was observed at the relatively low voltage of 4.0 V. According to the above-mentioned study, it was determined that the Li–O₂ battery (note: pure O₂ as the cathode) using the developed composite cathode catalyst showed a satisfactory performance, given that a high discharge capacity was obtained. However, the use of a pure O₂ atmosphere as the cathode in an Li–O₂ battery is impractical given that the ambient environment only has 21 vol% O₂ together with CO₂ and other gases.

Thus, the battery tests conducted on the RCFO-5 composite catalyst in a pure O₂ atmosphere (Li–O₂ battery) was impractical compared to the more attractive open-air system known as the Li–air battery. The major challenge here is to determine the impact of the CO₂ present in the atmosphere on the battery. Although the presence of CO₂ was reported to have some advantages in the Li–air battery,¹⁴ the decomposition of Li₂CO₃ formed on discharging is a major challenge in the cathode catalyst. Furthermore, the huge band gap and the thermodynamic stability of Li₂CO₃ make it difficult to re-oxidize on charging. Although the Earth's atmosphere only has 0.4 vol% CO₂, a strained atmosphere of 10 vol% CO₂ and 90 vol% O₂ was selected to explore the influence of CO₂ on the performance of the Li–air battery. Thus, initially rate capability tests by means of the GCD technique at 100, 200, 300 and 500 mA g⁻¹ were carried out, as shown in Fig. 9(d). The discharge capacities of 12 800, 10 680, 5200 and 3000 mA h g⁻¹ were obtained at these current densities, respectively. These capacities derived from the Li–air battery were lower compared with that of the Li–O₂ battery at the same current density (Fig. 9(b)). This reduction can only be attributed to the presence of CO₂ in the given gas atmosphere given that it was the only difference compared to the conditions in Fig. 9(b).

It was observed that the Li–air battery (10% CO₂ + 90% O₂) delivered a diminished capacity compared to that of the Li–O₂ battery (100% O₂) using the same composite cathode catalyst. Given that the only difference here was the introduction of 10% CO₂ together with 90% O₂, it had a clear influence on the discharge capacity of the Li–air battery. To examine the impact of CO₂ on the activity of the catalyst, further analysis on the battery working in a pure CO₂ atmosphere was separately performed. Thus, rate capability tests on the Li–CO₂ (Li||CO₂) battery by means of GCD analyses were executed at different current densities. The obtained GCD profiles at current densities of 100, 200 and 300 mA g⁻¹ are shown in Fig. 9(f), achieving the capacities of 620, 270, and 90 mA h g⁻¹, respectively. It was noticed that the discharge and charge capacities were notably lower, with a large difference between them, indicating that the catalyst was not able to oxidize or decompose the Li₂CO₃ discharge product completely. It should be noted that the Li–CO₂ battery (pure CO₂ on the cathode atmosphere) is also not a practical battery in the aforementioned atmosphere given that the ambient air contains mostly 21 vol% O₂, not a 100% CO₂. This battery may be preferred where a pure CO₂ atmosphere is present such as in Mars or where CO₂ is prevalent such as industrial furnaces.³⁵

A similar performance trend in the cyclic stability test was observed when moving to the Li–CO₂ battery from the Li–O₂

battery. The cycling stability, a practically more important feature of a battery, was analyzed by running multiple cycles of GCD at a current rate of 100 mA g⁻¹ with a charge–discharge capacity cutoff at 1000 mA h g⁻¹. As can be observed in Fig. 10(a), the Li–O₂ battery (100% O₂) demonstrated almost stable charge–discharge capacities, with a minor decrease in the capacity seen in the 10th cycle. Moreover, the overpotential measured in the Li–O₂ battery, which is the difference in voltage between the charging plateau and discharging plateau, is lower when compared with that of the Li–air battery. The overpotential of the Li–O₂ battery (100% O₂) at the capacity of 200 mA h g⁻¹ was measured to be 1200 mV (Fig. 10(a)). In contrast, the capacity of the Li–air battery (10% CO₂ + 90% O₂) began to fade after a few cycles, as observed in Fig. 10(b), because of the Li₂CO₃ buildup in the pores of the cathode catalyst, making it difficult for the catalyst to break it down over the cycles. By the 10th cycle, the capacity diminished to as low as 477 mA h g⁻¹. Moreover, the polarization overpotential measured in the Li–air battery (Fig. 10(b)) was as high as 1800 mV, which is higher than the polarization voltage observed for the Li–O₂ battery (Fig. 10(a)). Thus, the catalyst needed a much higher input voltage to decompose the discharge product in the Li–air battery. Notably, the cyclic stability of the RCFO-5-based Li–CO₂ battery (100% CO₂) was also tested by running multiple cycles of GCD at 100 mA g⁻¹, as depicted in Fig. 10(c). Although the cell could achieve the limiting high capacity of 1000 mA h g⁻¹ for up to three cycles, the charging capacity was phenomenally low, and the discharge capacity diminished simultaneously, indicating that the catalyst could not completely oxidize the carbonate/oxalate-based discharge product. Consequently, rapid fading of the cell capacity occurred, and by the 10th cycle, the capacity of the Li–CO₂ battery plummeted to as low as 50 mA h g⁻¹. This implies that under a pure CO₂ environment, the Li–CO₂ battery performance was severely affected due to the formation of Li₂CO₃, which requires a high decomposition potential, clogging the catalyst pores, and thereby retarding the CO₂ uptake.

3.4 Mechanism of Li–O₂/Li–CO₂ batteries

A deeper understanding of the reaction mechanism and the electrolytic stability in the Li–O₂ and the Li–CO₂ batteries were acquired by analysing the CV profiles recorded at 0.1 mV s⁻¹ on the Li–O₂ and Li–CO₂ batteries. Here, the CV profile in an atmosphere of Ar was also recorded to understand the stability of the electrolyte in the absence of any active species such as O₂ and CO₂. The obtained profiles are given in Fig. 10(d). The CV profiles recorded on the Li–O₂ battery demonstrated an onset potential of about 2.8 V vs. Li/Li⁺ on the cathodic sweep of the O₂ reduction reaction, which is close to the theoretical potential, as shown in eqn (10). The oxidative peak of the CV profile was observed at about 4 V, consistent with the previously observed charging plateau potential (Fig. 9(b)). At a higher potential above 4.2 V, there was a sharp increase in the current density, which can be correlated with the electrolytic decomposition promoted by the aggressive reaction intermediate such as superoxide (O₂⁻) during the ORR. When these observations

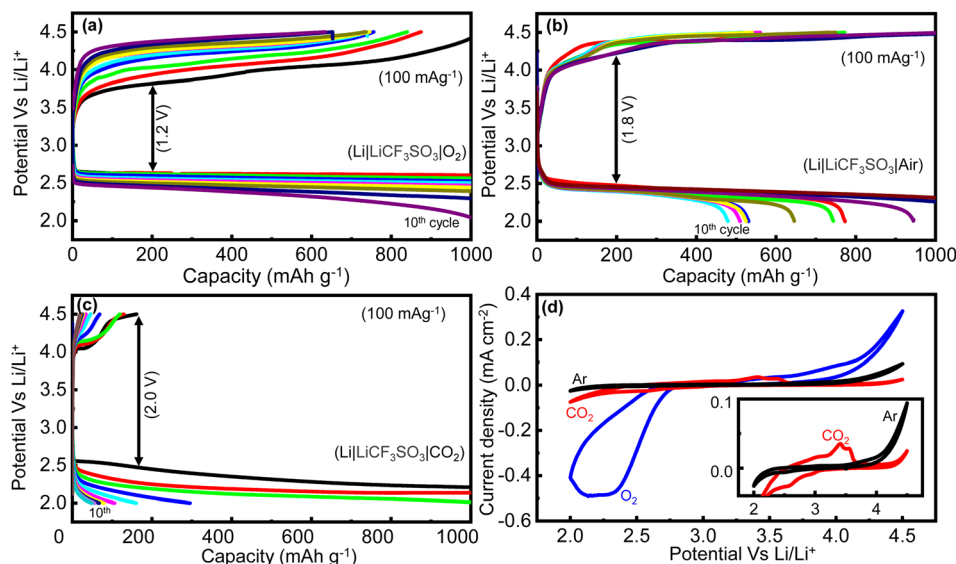
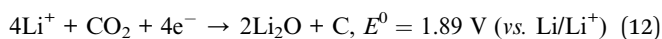
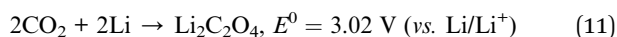
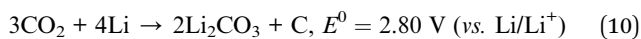


Fig. 10 (a–c) The cycle life performance of the Li–O₂, Li–air and the Li–CO₂ batteries at 100 mA g^{−1} using the RCFO-5 the cathode catalyst and (d) CV profiles recorded in various atmospheres of Ar, O₂ and CO₂ at a scan rate of 0.1 mV s^{−1}. Inset: Zoomed-in CV profile of Li–CO₂ battery in comparison with its performance in an inert Ar-atmosphere.

are compared with the CV profile obtained in an Ar atmosphere, the intermediate (O₂[−])-promoted electrolyte decomposition was more evident. Specifically, the non-negligible current spike observed at higher potentials (>4.2 V) even in the absence O₂ reveals the instability of the electrolyte at higher voltages, which was amplified in the presence of O₂ due to the formation of the intermediate.

Further information about the electrolyte stability and reaction mechanism was gathered from the CV profile of the Li–CO₂ battery. For better visibility, the CVs recorded in the CO₂ and Ar atmospheres are given in the inset of Fig. 10(d). The reduction of CO₂ on the RCFO-5 catalyst in the Li–CO₂ battery proceeds through multiple steps at the potentials of 2.9, 2.6, and around 2.0 V vs. Li/Li⁺. Although the theoretical potential of Li₂CO₃ formation is 2.8 V, the sluggish kinetics of the reaction will make the actual reaction potential lower.³⁶ Thus, the peak appearing at around 2.6 V is designated to the reduction reaction to form Li₂CO₃ according to eqn (10). The less obvious peak located at 2.9 V is attributed to the formation of Li₂C₂O₄ according to eqn (11). It is known that Li₂C₂O₄ as a discharge product is a preferred product in the Li–CO₂ battery.³⁶ The lower peak appearing at around 2.0 V is correlated with the formation of Li₂O according to eqn (12).



The oxidation peaks mainly appeared at around 3.4 V, corresponding to the decomposition of Li₂C₂O₄ and 3.6 V for the decomposition of Li₂CO₃.³⁶ Another interesting observation that can be made here is the electrolyte protection due to the

presence of CO₂, as reported elsewhere.¹⁴ The current density related to the electrolyte decomposition at higher voltages (>4.2) observed in the Ar atmosphere was more suppressed in the presence of CO₂ (Fig. 10(d) inset). Additionally, this suppressed peak corresponding to the decomposition of the electrolyte in the Li–CO₂ battery shifted to a higher voltage 4.4 V, suggesting that the electrolyte is more safeguarded in the presence of CO₂.

Thus, the reduced capacities and cycle-life stability of the Li–air and Li–CO₂ batteries can only be due to the incomplete decomposition of the discharge products (Li₂C₂O₄/Li₂CO₃) and not due to any electrolyte-related issues given that the electrolyte is more protected in the presence of the CO₂. The lack of the complete reoxidation of the lithium-carbonate/oxalate generated on the pores of the cathode during the discharge caused a huge resistive load on the cathode over multiple cycles. To verify this, electrochemical impedance spectroscopy studies at various discharge voltages for 1st and 5th cycles were performed. The Nyquist plots acquired from these measurements are given in Fig. 11(a and b). The obtained Nyquist plots were fit using an equivalent circuit, as presented in ESI Fig. S6,† and used to calculate the charge-transfer resistance (*R*_{ct}). The *R*_{ct} of the Li–CO₂ battery increased from 271 Ω at 3.1 V to 366 Ω at 2.0 V in the first discharge, as shown in Table S1.† The build-up of insulating Li₂CO₃ in the pores of the cathode as the cell discharges resulted in this increase in *R*_{ct}. However, more evidently, the *R*_{ct} at the end of 5th cycle increased further to a high resistance of 600 Ω at 2.0 V. These observations suggest that the discharge products were not completely decomposed and their accumulation increased over multiple cycles. Moreover, the presence of the carbonate discharge product remaining on the cathode catalyst after recharging was further confirmed by the postmortem XRD analysis. For this, the XRD pattern of the fresh electrode was recorded before the battery discharging. The electrode of another cell was also subjected to

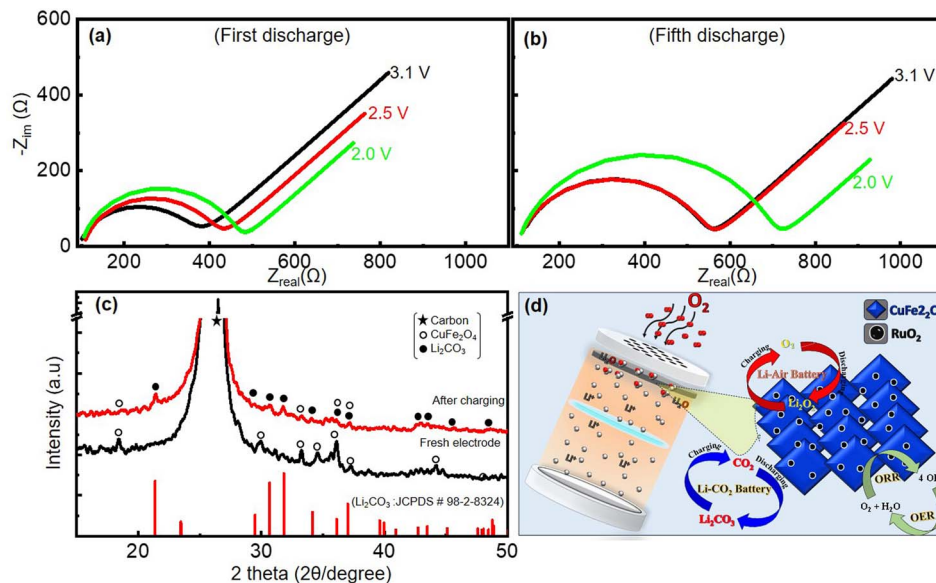


Fig. 11 Nyquist plots recorded in the frequency range of 0.1 Hz to 10 kHz at various stages of discharge in (a) 1st and (b) 5th cycles in a CO_2 atmosphere. (c) Postmortem analyses using XRD patterns of initial, discharged and charged RCFO-5 cathode catalyst and (d) schematic of all the catalytic reactions happening on the $\text{RuO}_2@ \text{CuFe}_2\text{O}_4$ catalyst used in the Li-air and Li- CO_2 batteries.

XRD after the first discharge followed by first charge. These two XRD patterns before and after the first discharge-charge are shown in Fig. 11(c). This electrode powder for the XRD analysis was collected by taking the electrode out of the dismantled

battery after charging the cell, followed by washing with ethanol and drying. The XRD pattern of the discharged electrode shows a peak corresponding to Li_2CO_3 even after the charging of the battery together with peaks of the coated-catalyst material and

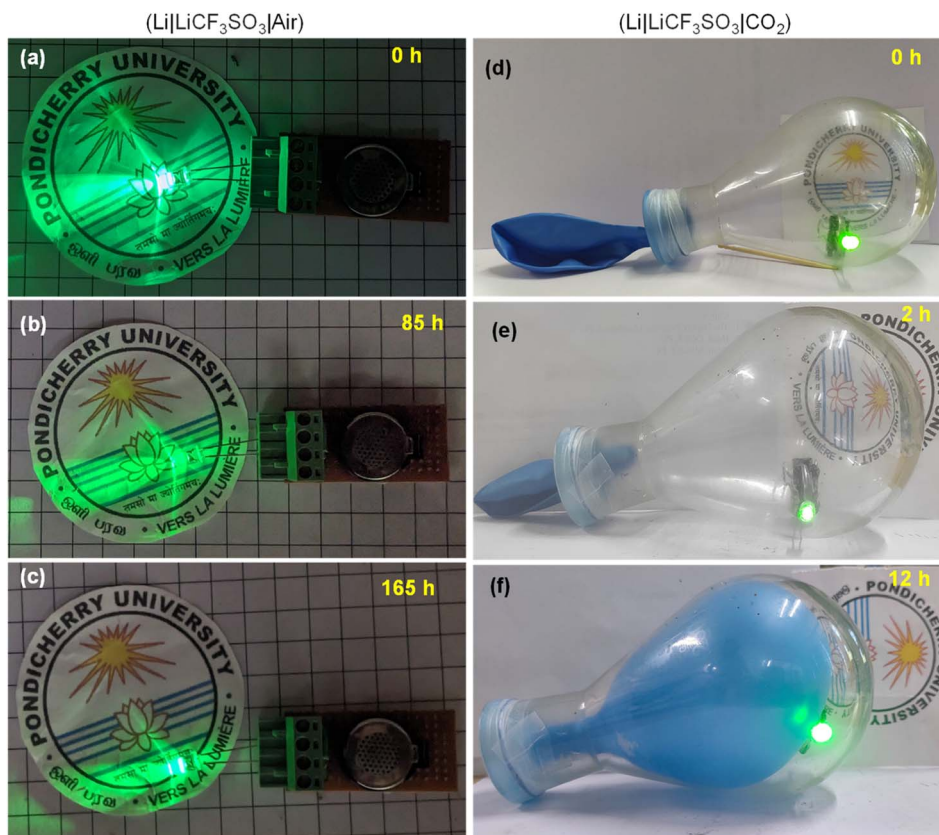


Fig. 12 (a-c) Timeline photographs showing the intensities of the glowing LED bulb, powered by the Li-air cell and (d-f) CO_2 utilization on discharge of Li- CO_2 battery while powering a commercial green LED in a pure CO_2 atmosphere maintained inside a round-bottom flask.

that of the carbon substrate used as the gas diffusion layer. Based on the aforementioned results, the overall reactions that are catalyzed on the RCFO-5 catalyst are schematically summarized in Fig. 11(d). It seems that the added RuO₂ acted as catalytic centres to promote the ORR/OER rates, resulting in high kinetics and discharge capacity. The enhanced reaction centres facilitated an improvement in the discharge capacity of the Li-O₂ battery.

3.5 Demonstration of prototype Li-air|Li-CO₂ batteries

The practical applications of the Li-air and Li-CO₂ batteries were demonstrated by lighting a 2.8 V commercial LED bulb in an open atmosphere and a closed container filled with CO₂, respectively. The Li-air battery could power the bulb for a long period and a timeline-photographic representation is given in Fig. 12(a-c). In the case of the Li-CO₂ battery, the CO₂ atmosphere was introduced in a 250 mL RB flask sealed with a balloon. The container was purged with pure CO₂ intentionally after placing the LED bulb connected to the Li-CO₂ battery inside. Within the 12 h of powering the LED bulb, the CO₂ gas inside the flask was almost utilized, as seen by the suction of the balloon inside the RB flask. On refilling the CO₂ gas, the powering of the LED bulb was noted, demonstrating the functional Li-CO₂ battery using the RuO₂@CuFe₂O₄ cathode catalyst. These observations are pictorially represented in Fig. 12(d-f). Thus, the addition of RuO₂ could increase the reaction centres, which facilitated high ORR/OER kinetics and high discharge capacity for the Li-air battery. The CO₂ had a diminishing effect on the Li-air battery performance but provided a safe cycling of the electrolyte.

4. Conclusions

According to this study, it can be concluded that a very small quantity of RuO₂ as low as 5 wt% embedded in the inverse spinel CuFe₂O₄ drastically improved the ORR/OER kinetics, imparting very good stability to the catalyst for the ORR/OER in 0.1 M KOH electrolyte. An improved current density response, onset potential and Tafel slope were achieved on the RuO₂-embedded CuFe₂O₄. The in-depth electrocatalytic investigations revealed a 4e⁻ transfer pathway with insignificant H₂O₂ formation in the ORR. A commendable OER performance with good stability and a low Tafel slope were achieved in the OER using this catalyst. The excellent ORR/OER performances were sustained on extending the study to an Li-air battery. The Li-O₂ battery operating in a pure O₂ atmosphere gave a high capacity of 14 250 mA h g⁻¹ at a current density of 200 mA g⁻¹. The impact of CO₂ on the performance of this enhanced battery catalyst was investigated *via* a wide variety of tests in Li-air and Li-CO₂ batteries including CV, GCD, postmortem XRD analyses and electrochemical impedance spectroscopy. The traces of Li₂CO₃ in the catalyst pores observed in the postmortem analysis of the electrode indicated that the RuO₂@CuFe₂O₄ catalyst could not completely re-oxidize the carbonate, which is consistent with the impedance observation of increasing

charge-transfer resistance over several charge-discharge cycles in a CO₂ atmosphere.

Data availability

The data that support the findings of this study are available from the corresponding author upon reasonable request.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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