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[Bmpy] or [Bmim]: which is better for H₂ sensing?

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Ionic liquids (ILs) have been found to be a good type of electrolyte material to fabricate highly sensitive H₂ sensors, accredited to their minimal vapor pressure and array of tunable physicochemical properties. Of the two IL molecules commonly used, [Bmpy][NTf₂] and [Bmim][NTf₂], experimental results reveal that [Bmim][NTf₂] exhibits a higher ionic diffusivity and conductivity than [Bmpy][NTf₂]. However, recent hydrogen sensing tests demonstrate that [Bmpy][NTf₂] based sensors are more sensitive instead. Until now, this seemingly contradictory phenomenon has lacked a reasonable explanation because of the spatial and temporal limitations of current experimental techniques. Thus, molecular dynamics (MD) simulations were used in this work to examine the electric double layer (EDL) structure and H₂ diffusion in the EDL for the two IL species. With the use of multiple descriptors like IL number distributions, orientation distributions, etc., the electrolyte|electrode heterostructure can be categorized into three distinct regions: the 1st EDL, the 2nd EDL, and the bulk phase. The self-diffusion coefficients of IL cations and anions for each region are then calculated and compared, which is, as per our knowledge, the first time that the diffusion-related differences in the different regions of the electrolyte|electrode interphase have been addressed. As compared to [Bmim], [Bmpy] cations demonstrate a more scattered orientation distribution within the 1st EDL, which allows more H₂ transport pathways to the electrode and thus leads to a higher possibility of H₂ redox reaction. Furthermore, H₂ molecules show a slightly higher bulk solubility and higher probability density in the 1st EDL of the positive electrode (PE) in [Bmpy][NTf₂] than in [Bmim][NTf₂]. Collectively, these results provide insights into why [Bmpy][NTf₂] is a more sensitive electrolyte material than [Bmim][NTf₂].

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1. Introduction

Hydrogen is a clean and powerful energy resource with a high energy density of 120–142 MJ kg^{−1}.¹ Thus, it has long been considered a promising energy source to confront the pollution and global warming challenges caused by traditional fossil fuels. In fact, it has been extensively utilized in various applications, like automobile fuel cells and satellite energy supplies.^{2,3} However, H₂ has a low ignition energy (0.017 mJ),⁴ and even a minor H₂ leak at a volumetric ratio as low as 4% in air can lead to an explosion. Such safety concerns are one of the key challenges that hinder the large-scale application of H₂ energy. Thus, hydrogen sensing technology with high sensitivity/selectivity/reliability, short response time, and preferably low cost is greatly needed to detect the possible H₂ leakage during the stage of H₂ production, transportation, storage, and other hydrogen-related applications.

Various types of hydrogen sensors have been developed based on different detection mechanisms: catalytic, thermal

conductivity, electrochemical, resistance, work function based (MOX/FET based⁵), etc.⁴ Among them, electrochemical (EC) hydrogen sensors are of particular interest accredited to their high selectivity and sensitivity, real-time measurement capability and, more importantly, low-power consumption, enabling their long-term operation. In addition, based on different operating mechanisms, the EC hydrogen sensors can be further classified into the following three sub-types: conductometric (change in resistance), potentiometric (development of a potential), and amperometric (generation of redox current) sensors.^{6,7} In a conductometric sensor, the design of such hydrogen sensors is based on the careful choosing of the liquid electrolyte so that conductivity changes in the sensor could be observed once H⁺ is generated from H₂ molecules in the case of H₂ leakage, wherein H⁺ has a much higher conductivity than that of other ion species in the electrolyte.⁸ In a potentiometric sensor, there are usually at least two electrodes, one sensing electrode (SE) and one reference electrode (RE), where the potential difference between SE and RE is measured for different hydrogen concentrations based on the Nernst equation.⁴ A potentiometric sensor has little to no dependence on sensor size and geometry, which is beneficial for sensor miniaturization. In addition, the detected potential has a logarithm relationship with the gas concentration, enabling its

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capability of measuring gas concentrations spanning a wide range,⁹ extending beyond 10 orders of magnitude.^{7,8} However, this also indicates a lower accuracy at high gas concentrations.⁴ The third type of EC sensor, the amperometric sensor, detects current signals instead that are caused by redox reactions taking place at a sensing electrode. An amperometric sensor should consist of at least two electrodes, the working (WE) or sensing electrode (SE), and the counter electrode (CE), and operate under an externally applied voltage. In real-world applications, a reference electrode (RE) is also commonly included to help control the thermodynamic potential of the WE during sensing.⁸ While in operation, H₂ gas is first dissolved into the electrolyte placed between the electrodes, and then adsorbed onto the electrode's surface. At the WE (or SE, positive electrode), the adsorbed H₂ is oxidized into H⁺ protons, and this process is also commonly referred to as the hydrogen oxidation reaction (HOR), known as one of the simplest but also most important reactions in the EC field.^{10,11} On the other hand, oxygen is reduced and combines with protons at the CE, yielding water molecules as the final product. In an amperometric sensor, the sensing current is often proportional to the H₂ concentration, not only enabling the highly accurate measurement of H₂ concentration, but also ensuring that sensitivity remains stable across a wide range of concentrations.^{4,7,8,12}

Given the prior introduction to the three types of EC hydrogen sensors, the amperometric sensor is favored in our study, because of its high sensitivity with excellent measuring accuracy, low H₂ concentration detect limit (ppm or even ppb level), and being more widely available commercially.^{4,13} In this kind of sensor, platinum (Pt) metal is favored as the electrode material, because it is one of the most efficient electrocatalysts that could facilitate hydrogen adsorption and hydrogen-related redox reactions (HOR) by reducing the activation energy,^{14–16} although the use of other metal materials like palladium (Pd) has also been reported.¹⁷ For the electrolyte material, a commonly used aqueous electrolyte is sulphuric acid.^{18,19} However, such traditional liquid electrolytes have a limited operative potential window, and suffer from solvent volatility issues that may impair the sensing accuracy and thus shorten their service life.^{20,21} To resolve these challenges, ionic liquids (ILs), consisting of organic cations and anions, are utilized as electrolyte materials for H₂ sensors, accredited to their negligible vapor pressure coupled with a series of unique physicochemical properties, *e.g.*, large electrochemical window, low melting temperature, high ion conductivity and thermal stability, *etc.*^{2,22} In addition, their physicochemical properties could be

tuned with ease by utilizing different cation and anion combinations, making them a promising type of electrolyte material for H₂ sensors.⁶

Of all the ILs, two have been utilized and proven effective as electrolyte materials in H₂ sensors, namely, 1-butyl-1-methylpyrrolidinium bis(trifluoromethanesulfonyl)imide ([Bmpy]⁺[NTf₂][−], denoted as [Bmpy][NTf₂] thereafter), and 1-butyl-3-methylimidazolium bis(trifluoromethanesulfonyl)imide ([Bmim]⁺[NTf₂][−], denoted as [Bmim][NTf₂] hereafter).^{2,11,23–26} The structures of the two ILs are shown in Fig. 1. Both types of IL cation molecules comprise a 5-membered ring, with the [Bmpy] ring being not-flat and the [Bmim] ring being planar due to π - π interactions.

Recent hydrogen sensing experiments show that a [Bmpy][NTf₂] based sensor is apparently more sensitive than one based on [Bmim][NTf₂].^{11,25} However, previous experiments revealed that both the cation and anion in [Bmpy][NTf₂] demonstrate smaller ionic diffusivity and conductivity than in [Bmim][NTf₂].^{27,28} If the HOR is limited by the H₂ diffusion process to the WE (or SE), these results appear contradictory, and a reasonable guess would be maybe it is the electrolyte|electrode region instead of the bulk IL that makes the difference: it is easier for H₂ to transport through the electrolyte|electrode region and reach the WE (or SE, positive electrode) in the [Bmpy][NTf₂] case than in the [Bmim][NTf₂] case. In the meantime, near the electrolyte|electrode interface, an electrical double layer (EDL) structure will be formed. But unlike in the aqueous solution, IL molecules could be strongly adsorbed onto metal surfaces, which is then followed by one or more layers of ion pairs, making the EDL in IL more than one layer thick.²⁹ To date, there have been many experimental and simulation studies on the structures and diffusion related properties of ILs, in both the bulk IL and electrolyte|electrode interphase. For example, experimentally, the pulsed-field-gradient spin-echo (PGSE) NMR method could be used to measure the self-diffusivity of IL cations and anions. The complex impedance measurements could be used to measure the bulk ionic conductivity and the differential capacitance (DC) of the EDL.^{27,28,30} The quasi-elastic neutron scattering (QENS) is employed to investigate the ion dynamics.^{31–33} The atomic force microscopy (AFM) force spectroscopy and *in situ* scanning tunneling microscopy (STM) can be used to study the EDL structure.³⁴ Meanwhile, there are also many simulation studies on the various aspects of the electrolyte|electrode interface, for example, number density distribution, charge density distribution, differential capacitance (DC), integral capacitance (IC), electric potential profiles, molecule orientations, *etc.*^{35–42}

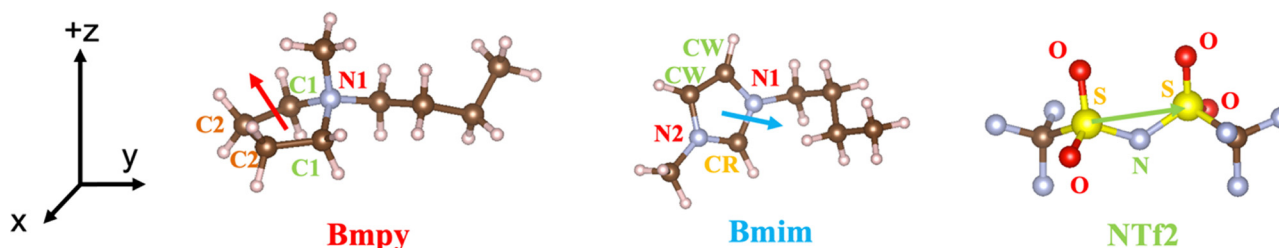


Fig. 1 IL molecule orientation vector.



However, all these aforementioned experimental and simulation studies could still not explain the seemingly contradictory phenomenon – why [Bmpy][NTf₂] based H₂ sensors are more sensitive than [Bmim][NTf₂] based H₂ sensors. There is a gap between the IL|Electrode EDL study and the H₂ sensor observation. In addition, in the current hydrogen sensor experiments, although the HOR could be studied by cyclic voltammograms (CVs),^{24,43} the *in situ* H₂ diffusion process to the WE (or SE), particularly through the electrolyte|electrode interphase, is still difficult to examine directly with the experimental techniques available today. Thus, this motivated us to do a comparative study of the EDL structure of the two IL species and find out how the structures affect the H₂ sensibility. By using molecular dynamics (MD) simulations, we can investigate the EDL structure and the H₂ diffusion process to the WE at the atomic level, which is hard to accomplish in experiments. This work could also provide guidance on designing better electrolyte materials for hydrogen sensors in the future.

The manuscript is organized as follows. First, the ionic diffusivity and conductivity for cations and anions in [Bmpy][NTf₂] and [Bmim][NTf₂] ILs are calculated from our simulations and compared with experimental results to validate the reliability of our simulation approach. Then, we study the H₂ diffusivity and solubility in these two bulk ILs. Second, we investigate the electrode|electrolyte|electrode heterostructure using various descriptors, including number density distributions, molecule orientations, and molecule displacements along the z-axis, which allows us to gain a good understanding of the structure of the whole electrochemical cell and divide the cell into five regions. Third, we examine the diffusivity of ions and the probability of H₂ in each of these five regions. Finally, we discuss how these structures and dynamics properties obtained from the simulations are relevant to H₂ sensing.

2. Methods

2.1. Force field (FF) or interatomic potential (IP) parameters

Two types of ILs were studied in the comparative study, [Bmpy][NTf₂] and [Bmim][NTf₂]. For both ILs, the Canongia Lopes & Padua (CL&P) FF^{44–47} was adopted because of its good accuracy, compatibility, and transferability. CL&P FF was developed based on the OPLS-AA⁴⁸/AMBER⁴⁹ framework and is well-tailored for the IL species, making it one of the most widely used FF for IL simulations. Meanwhile, a charge scaling factor of 0.8 was applied to all IL atoms, as suggested by many papers,^{50,51} to improve the predictions on thermodynamic and dynamic properties significantly. The Pt and H₂ FF parameters were extracted from the work of Heinz *et al.* and Wang *et al.*,^{52,53} respectively.

The geometric mixing rules (OPLS default mixing rules) are applied for the nonbonded Lennard-Jones (LJ) interactions between unlike atoms. Meanwhile, intramolecular nonbonded interactions are only counted for atoms that are three or more bonds apart. And specifically, the 1–4 interactions are scaled down by a factor of 0.5.

2.2. Molecular dynamics (MD) simulation

MD simulations were performed using the large-scale atomic/molecular massively parallel simulator (LAMMPS)^{54,55} or GROMACS.⁵⁶

For the NPT ensemble, the Nose–Hoover thermostat and barostat^{57,58} were used with time constants of 0.1 ps and 1.0 ps in LAMMPS. Canonical sampling through velocity rescaling (CSVR) thermostat and barostat⁵⁹ of time constants of 0.1 ps and 1 ps were used in GROMACS. The LJ interactions are truncated at 12 Å, and the long-range Coulombic interactions beyond 12 Å were handled *via* the particle–particle–particle–mesh (PPPM) method with a relative precision of 10^{−6}.⁶⁰ All C–H bonds in IL cations and the H–H bonds in the inserted H₂ molecules were constrained using the SHAKE algorithm⁶¹ in LAMMPS and LINCS algorithm⁶² in GROMACS. The equations of motion were integrated utilizing the velocity-Verlet algorithm⁶³ with a 1 fs time-step for LAMMPS and the stochastic integrator with a 2 fs time-step for GROMACS.

For the bulk ILs, 500 ion pairs were packed inside a cubic box of 70 Å (x-axis) × 70 Å (y-axis) × 70 Å (z-axis) using the software Aten.⁶⁴ The NPT ensemble MD run was conducted using LAMMPS for 200 ps at 300–420 K, with a pressure of 1 bar applied in all directions isotropically. The ILs were further equilibrated for 50 ps with the NVE ensemble, followed by 30 ns production runs. For the electrochemical cell with electrodes, 1000 IL ion pairs were first packed inside a box of 53 Å (x-axis) × 53 Å (y-axis) × 550 Å (z-axis). The NPT ensemble MD run was performed with LAMMPS for 600 ps at 340 K, applying 1 bar of pressure only along the z-axis. Each Pt electrode consists of a 3-layer Pt(111) slab, with a dimension of 53 Å (x-axis) × 53 Å (y-axis), and each layer consists of 418 Pt atoms. The Pt layer in direct contact with the ionic liquid was mobile and uniformly charged, while the two other layers were frozen and neutral. Three values of the total electrode charge Q_e were tested for the innermost layer of each Pt electrode: $\pm 2e$, $\pm 5e$, $\pm 8e$. A 10 ns NVE ensemble MD simulation was then performed at 340 K to equilibrate the heterostructure, with periodic boundary conditions only along the x and y directions. At last, 40 H₂ particles were randomly placed into the heterostructure. The structure was first equilibrated for 50 ps in the NVE ensemble, and the final MD run was carried out in an NVE ensemble for 40 ns at 340 K. For each test case, three parallel MD runs were performed. The data generated from the 40 ns NVE MD run are used to study the system properties and particle dynamics. The atomic and molecule Center-of-Mass (COM) coordinates are recorded every 10 ps.

The solubility of H₂ was obtained from the alchemical free energy calculation method. Molecular dynamics simulations were performed with GROMACS at 270, 300, 340, and 400 K. The system (500 ion pairs and 1 H₂) was equilibrated with the NPT ensemble for 3 ns. The sampling of the free energy was performed with the NPT equilibration run of 3 ns, followed by the NPT production run of 12 ns (270 K), 10 ns (300 K), and 6 ns (340, 400 K). The solvation energy, *i.e.*, excess chemical potential, of H₂ was obtained with the BAR method.⁶⁵

2.3. Data analysis

The diffusivity of cation and anion species and the total conductivity were obtained from the incoherent mean square displacement (MSD) and charge-weighted MSD, respectively, based on the Einstein–Helfand formula.^{60,66} To obtain the molecule number density distribution, we sliced the simulation



cell along the z direction into multiple bins with a width of 0.1 Å and averaged the number over all the parallel MD runs. It should be noted that we used the atomic coordinate instead of the COM coordinate to identify the location of a molecule for all the following distribution calculations, in order to avoid the effect of the molecular orientations caused by their bulky size. For IL cations, the center N atom (labeled as “N1”, also the atom closest to COM in a cation molecule) in the [Bmpy] & [Bmim] ring were used, while for [NTf₂] anions, the center N atom is used, as indicated in Fig. 1.

The orientation of an IL molecule was defined by a vector, as shown in Fig. 1. For an IL cation molecule, we used the normal vector of a cation ring to indicate its direction. For an IL anion [NTf₂], we connect its two sulfur atoms to form the direction vector. Since the Pt electrodes are placed along the xy -plane, and the z -axis corresponds to the normal vector of the Pt electrode, the angles between an IL direction vector and [001] vector ($+z$) were calculated to obtain the IL molecule orientation. For IL cations ([Bmpy] or [Bmim]), an angle value of 0° and 180° means that the cation ring stays parallel to the electrode surface, while for IL anions ([NTf₂]), an angle value of 90° means that the linear anion molecular stay parallel to the electrode surface. The orientation distribution is subsequently derived in a manner similar to that of the number distribution.

3. Results and discussion

3.1. Bulk ionic liquids

3.1.1. FF validation: self-diffusion & conductivity. First, to validate our simulation model, we calculated the self-diffusivity of cations/anions and total conductivity of the bulk IL from 300 to 420 K. They were plotted as the inverse of temperature and then compared with experimental results,^{27,28} as shown in Fig. 2. The calculated IL diffusivity (D) values have an order of magnitude from 10^{-8} to 10^{-6} cm² s⁻¹, and conductivity (σ) values

have an order of magnitude from 10^{-3} to 10^{-2} S cm⁻¹. Their values align closely with the available experimental data (dashed lines) across the investigated temperature range. Meanwhile, in both ILs, cations show a higher self-diffusivity value than the anions, in agreement with various previous reports.^{67–69}

The self-diffusion coefficient of H₂ in bulk ILs was also investigated, as shown in Fig. 3. It shows that H₂ molecules move at 10^{-5} to 10^{-4} cm² s⁻¹, which is about one order of magnitude faster than IL ions. Similar to the IL case, H₂ demonstrates a higher diffusivity in [Bmim][NTf₂] than in [Bmpy][NTf₂]. The temperature dependence of H₂ can be fitted to an Arrhenius relationship with activation energies of 0.16 and 0.21 eV for [Bmim][NTf₂] and [Bmpy][NTf₂], respectively, while the temperature dependence of ions in Fig. 2 resembles the Vogel–Tammann–Fulcher (VTF) relationship.^{70–72} H₂ diffusivity from experimental and computational efforts for other imidazolium-NTf₂ ionic liquids are also

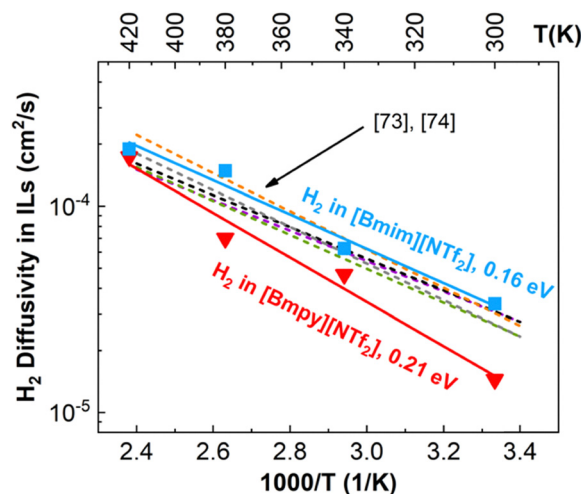


Fig. 3 Calculated self-diffusivity of H₂ in ILs, with reference data shown as dashed lines.^{73,74}

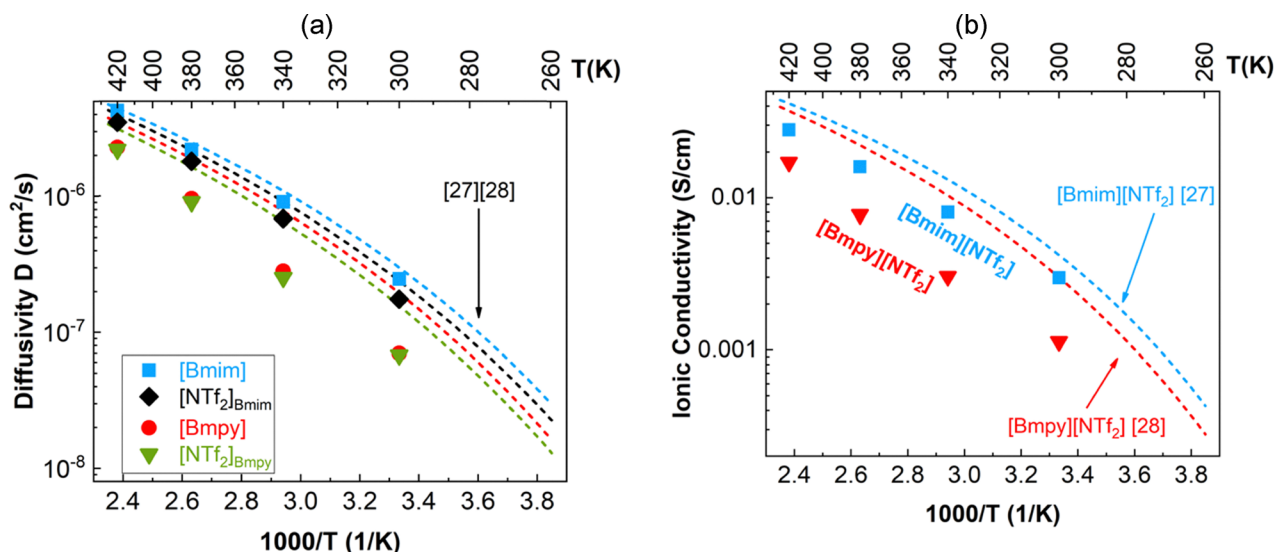


Fig. 2 Calculated (a) self-diffusivity and (b) ionic conductivity at 300–420 K, with experimental data shown as dashed lines.^{27,28}



shown for comparison.^{73,74} Both these literature values and activation energies agree with our simulations.

These results confirmed the reliability of our simulation approach. Furthermore, they suggest that cations/anions and H₂ both move faster in the bulk [Bmim][NTf₂] than in the bulk [Bmpy][NTf₂]. However, recent H₂ sensor experiments reveal that [Bmpy][NTf₂] is more sensitive in sensing H₂ concentration than [Bmim][NTf₂].^{11,25} Such seemingly contradictory results indicate that the point where it truly makes a difference is probably at the IL|Electrode interface instead of in the bulk IL region, or the H₂ solubility.

3.1.2. H₂ solubility. The calculated excess chemical potential μ_{ex} , i.e., solvation free energy, of H₂ is shown in Table 1. The solvation free energy in [Bmpy][NTf₂] is slightly lower than that in [Bmim][NTf₂] at lower temperatures but becomes similar at high

temperatures. The relationship between μ_{ex} and concentration c ⁷⁵ is:

$$\mu_{\text{ex}} = -RT \ln \left(RT \frac{c}{p} \right) \quad (1)$$

where T is the absolute temperature, R is the gas constant, and p is the gas pressure. The concentration can be related to the molar fraction x as:

$$x = \frac{cM}{\rho} \quad (2)$$

where ρ is the mass density and M is the molar mass. For 0.1 MPa H₂, the values of c and x are listed in Table 1. It appears that the solubility of H₂ in either ionic liquid is low, as expected from the H₂ non-polarity. Experimentally reported values are 0.77 and 0.86 mM⁷⁶ at 298 K for [Bmim][NTf₂] and [Bmpy][NTf₂], while we have 0.57 and 0.62 mM at 300 K. Considering that experimental data were extrapolated from 10.1 MPa H₂ to 0.1 MPa, we think there is reasonable agreement. Our data indicate that H₂ solubility is slightly higher in [Bmpy][NTf₂] than in [Bmim][NTf₂] at lower temperatures.

3.2. Electrical double layer (EDL) at the IL|electrode interface

3.2.1. Number (N) density distributions. The molecule number (N) distributions were acquired for IL species ([Bmim], [NTf₂] in [Bmim][NTf₂] labeled as [NTf₂]_{Bmim}, [Bmpy], [NTf₂] in [Bmpy][NTf₂] labeled as [NTf₂]_{Bmpy}) and they were plotted for regions within 30 Å near both the Pt negative electrode (NE) and positive electrode (PE), as shown in Fig. 4. First, near the NE

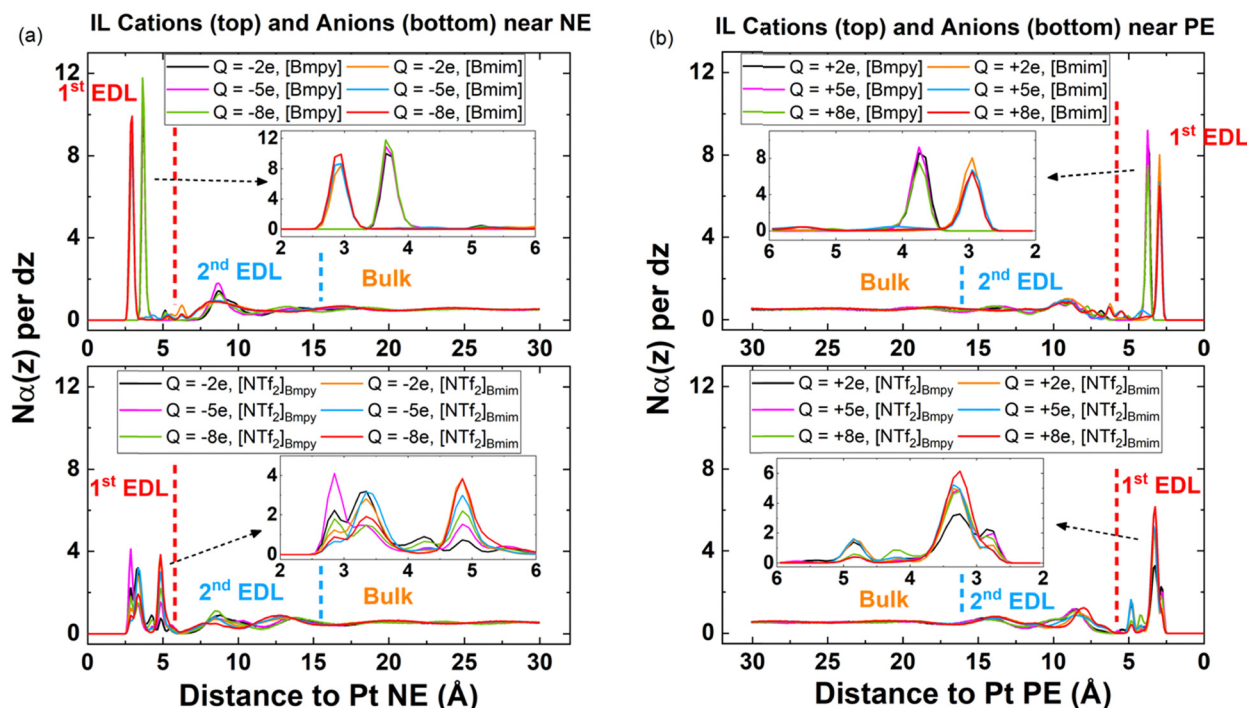


Fig. 4 Number density distribution for (a) IL cations (top row) and anions (bottom row) near Pt NE; (b) IL cations (top row) and anions (bottom row) near Pt PE.



region, both cations display very sharp peaks in the N distributions, with [Bmpy]'s peak occurring at ~ 3.7 Å and [Bmim]'s peak at ~ 3.0 Å (Fig. 4a, top). In the meantime, the peak heights of both cations rise with increasing electrode charges (inset), which is reasonable, as more IL cations will get attracted to the Pt surface by the increasing negative charges in the NE. Peak locations remain almost unchanged with NE charges. For anions near the NE region (Fig. 4a, bottom), it becomes much more complicated. Not only will the anions experience the coulombic repulsion from the NE, but they will also be attracted by the cations strongly adsorbed by the NE, and essentially, it is a balance of these two coulombic effects. The anion N distributions close to the electrode region (within 6 Å) could be approximately categorized into three sub-peaks, and they are located roughly at 2.8 Å, 3.3 Å, and 4.8 Å away from the electrode, as illustrated in the inset of Fig. 4a bottom. These multiple sub-peaks correspond to various IL molecular orientations (Fig. 7), as will be explained in more detail afterward. Intuitively, there is an expectation that with more charges on the NE, the anions tend to be repelled away from the electrode surface due to the increasing coulombic repulsion. However, if we examine the electrode charge transition from $-2 e$ to $-5 e$, the first two anion sub-peaks closest to the electrode actually go up, *i.e.*, the first sub-peak of $[\text{NTf}_2]_{\text{Bmpy}}$ (at 2.8 Å, magenta line) and the second sub-peak of $[\text{NTf}_2]_{\text{Bmim}}$ (at 3.3 Å, blue line), as depicted in the inset of Fig. 4a bottom. Although this may seem counterintuitive, it is not unusual, as more cations will be attracted to the NE when transitioning from $-2 e$ to $-5 e$, and the coulombic attraction effect from the adsorbed cations outperforms the repulsion effect from the NE. This could be indicative of the overscreen effect commonly observed in electrochemical ionic liquid systems when IL ions adhere to electrode interfaces, as described in numerous studies,^{77–79} where the initial layer of adsorbed counterions overbalances the charges on the neighboring electrode. When the NE charge increases from $-5 e$ to $-8 e$, both of the first two anion sub-peaks (2.8 Å & 3.3 Å) $[\text{NTf}_2]_{\text{Bmpy}}$ and $[\text{NTf}_2]_{\text{Bmim}}$ decrease in height, while the third sub-peak (4.8 Å) height rises, indicating that some anions closest to the electrode are repelled farther away due to the rising NE coulombic repulsion.

Near the PE region, the N distributions of cations exhibit similar characteristics as those in the NE regions. That is, there exists one sharp peak for both [Bmpy] and [Bmim] when in the closest vicinity of the PE, with [Bmpy]'s peak located at ~ 3.8 Å and [Bmim]'s peak located at ~ 2.9 Å, as exhibited in Fig. 4b, top. With growing charges on the PE, the peak heights of [Bmpy] and [Bmim] decrease, and the cations' peak heights in the PE region are always lower than their counterparts in the NE region, indicating that some cations are repelled away from the electrode due to the PE coulombic repulsion. For anions (Fig. 4b, bottom), similar to the NE case, multiple sub-peaks show up in the vicinity of the PE region, *e.g.*, at ~ 2.8 Å, ~ 3.3 Å, and ~ 4.8 Å. The locations of these sub-peaks are similar for $[\text{NTf}_2]_{\text{Bmpy}}$ and $[\text{NTf}_2]_{\text{Bmim}}$, irrespective of the electrode charge variations. Furthermore, there exists one dominating sub-peak located at ~ 3.3 Å for both IL species, and the height of this sub-peak increases with more charges in the PE, accredited to the increasing electrode coulombic attraction.

To sum up, it is clear that [Bmim] molecules stay closer to both electrodes (NE, PE) than [Bmpy] molecules, possibly due to two reasons: (1) there are two nitrogen atoms in the [Bmim] ring while only one in the [Bmpy] ring, leading to a higher van der Waals attraction between the [Bmim] ring and Pt atoms; (2) the [Bmim] ring is aromatic flat, while the [Bmpy] ring is non-flat. For anions, the N distributions look quite similar for $[\text{NTf}_2]_{\text{Bmim}}$ and $[\text{NTf}_2]_{\text{Bmpy}}$ with no significant difference, as the peaks are found at the same locations in both IL cases.

Despite the difference caused by the IL ion species or electrode charges, the N distributions do demonstrate some shared characteristics at both electrodes. When they are close to electrodes, the N distributions display sharp peaks extending to 6–7 Å away from the electrode. As will be shown later, this corresponds to the 1st electrical double layer or 1st EDL. Further into the bulk, the N distributions still show fluctuations, but they are more like waves instead of peaks. This region might extend to 16–17 Å towards the bulk, though no strict boundary exists. As will be shown later, this corresponds to the 2nd electrical double layer (2nd EDL). Beyond that, the distributions become relatively flat with minimal fluctuations, indicating the simulation cell length is sufficient in the z -axis, as observed in the simulation work of other solid/liquid interfaces.^{80,81}

3.2.2. Molecule orientation. Using the vectors we defined in Section 2.3, we can investigate the IL molecule orientations, presented in heat maps. In heat maps, the x -axis represents the distance to the Pt electrode, the y -axis refers to the angles between the IL molecule vector and the Pt electrode normal vector (as illustrated in Fig. 1), and the possibility of an orientation occurrence in space is indicated by color temperature: a higher orientation possibility is reflected by the red color, while a lower possibility is expressed in the blue and darker colors. Fig. 5 gives an example for [Bmpy], $[\text{NTf}_2]_{\text{Bmpy}}$ in a region 15 Å within the Pt NE electrode. We see discrete color ribbons and spots that are primarily red (*i.e.*, very high probability) inside the 1st EDL. This high probability corresponds to the sharp peaks in Fig. 4. When it comes to the 2nd EDL, the molecule orientations become more uniformly distributed as the isolated color stripes and spots become connected, and the red color cools down into cyan and yellow (lower probability). Again, this corresponds to the waves observed in Fig. 4.

For [Bmpy] inside the 1st EDL, there exist two types of characteristic molecular orientations. The first type corresponds to the two red spots in the top (close to 180°) and bottom (close to 0°), and the IL cation rings are in parallel with the electrode surface at such orientations. Meanwhile, at this configuration, the distributions appear as spots with confined y -space, indicating a small angular movement. The second type relates to the two orange/yellow ribbons close to 90° . At this configuration, IL cation rings stay at a nearly perpendicular direction to the electrode surface, with a large angular movement indicated by the elongated angle distributions (ribbons instead of spots). Similarly for $[\text{NTf}_2]_{\text{Bmpy}}$, two types of characteristic orientations exist: the first type relates to an orientation parallel to the electrode surface, with small angular movement, indicated by the red spots close to 90° . The second type corresponds to the



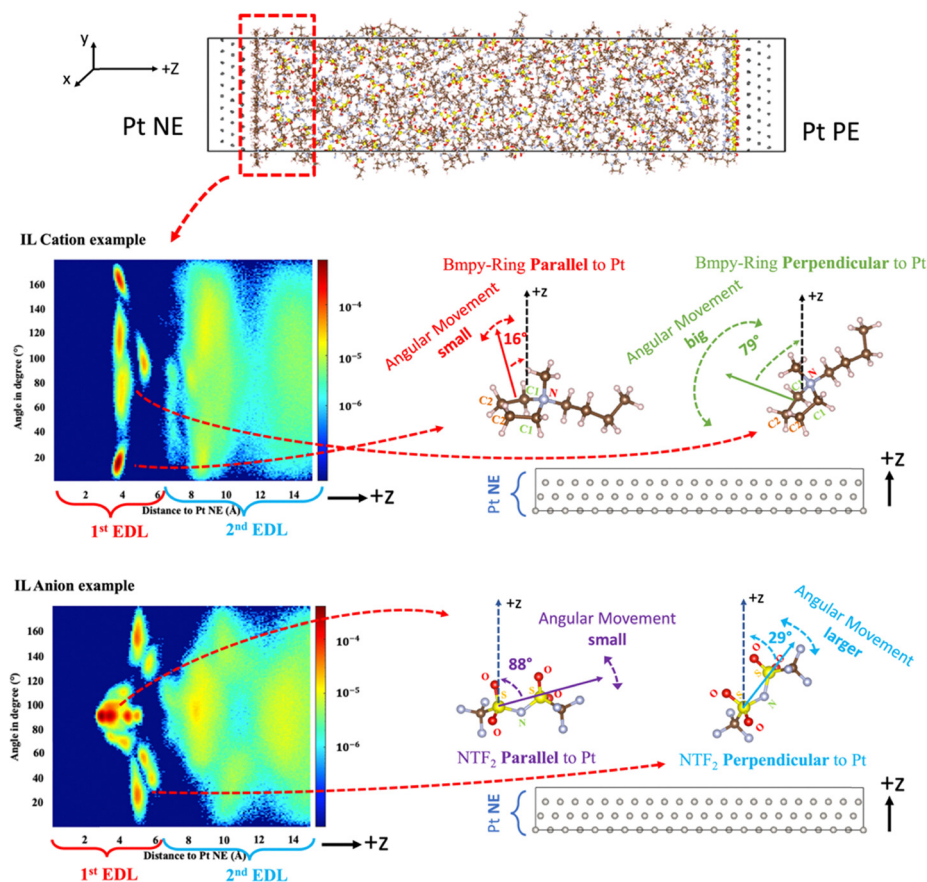


Fig. 5 Schematic of IL orientation distribution: using [Bmpy] and [NTf₂]_{Bmpy} near the NE region (Pt Q = −8 e) as an example. For [Bmpy], an angle close to 0° or 180° indicates that the cation ring tends to stay parallel to the electrode, and the angular movement of this molecule is usually small. If the angle of a cation is close to 90°, the molecule tends to remain perpendicular to the electrode, usually with big angular movement. For [NTf₂]_{Bmpy}, an angle close to 90° suggests a parallel configuration with small angular movement, while an angle near 0° or 180° indicates a perpendicular configuration with larger angular movement.

orientation almost perpendicular to the electrode surface with larger angular movement, as indicated by the two orange/yellow ribbons located close to 0° and 180°. In addition, for both IL cations and anions, the parallel orientation dominates as compared to the perpendicular orientation, as indicated by their color difference, especially in the case of using a logarithmic scale for the color legend.

The complete orientation distribution comparison between two ILs is shown below for cations and anions under various electrode charge conditions at both electrodes (Fig. 6–9). For cations in the NE region (Fig. 6), the orientation distributions in general change little with the electrode charges. There are also some changes with respect to the small yellow color stripes of angles ranging from ~70° to ~110°, at 5–6 Å away from the Pt surface, which correspond to the tiny shoulder peaks in the *N* distributions as shown in the top of Fig. 4a. In the meantime, for cations in the PE region (Fig. 8) with the charge value going up, there are more yellow color stripes that show up in the region at 4–7 Å to the electrode surface, and the color stripes in the 2nd EDL turned yellower, especially at 8–10 Å to the electrode surface. This is in line with expectations, as more cations are repelled away from the PE, and thus granted more

freedom of movement, consistent with the decreasing sharp peak heights ([Bmpy] peak: at ~3.8 Å, [Bmim] peak: at ~2.9 Å) in the *N* distributions results (Fig. 4b, top). Overall, in the closest vicinity to the electrode (within 4 Å) under each charge condition, most cations stay parallel to the electrode surface (~0/180°) in both IL species, but the cation orientations in [Bmpy] show a more scattered distribution than [Bmim]. In a H₂ sensor, the hydrogen molecules need to first be transported through the 1st EDL of PE to reach the electrode, before the redox reaction can happen. A cation ring staying parallel to the electrode surface tends to have a high possibility of blocking the H₂ transport pathway to the Pt surface than a perpendicular configuration. Thus, the more scattered orientation distribution in [Bmpy][NTf₂] makes it easier for H₂ to reach both electrodes, and thus a higher redox reaction possibility than that in [Bmim][NTf₂]. In addition, the closer distance of red spots in [Bmpy] compared with [Bmim] is consistent with the number (*N*) distribution in Fig. 4.

For anions in the NE region (Fig. 7), closest to the electrodes are two red spots centered at 90°. They are located at around a 3 Å distance from the NE surface, with a gap of ~0.4 Å between the two spots. The centers of the two spots look maroon, suggesting a high probability of occurrence. They correspond to the first two



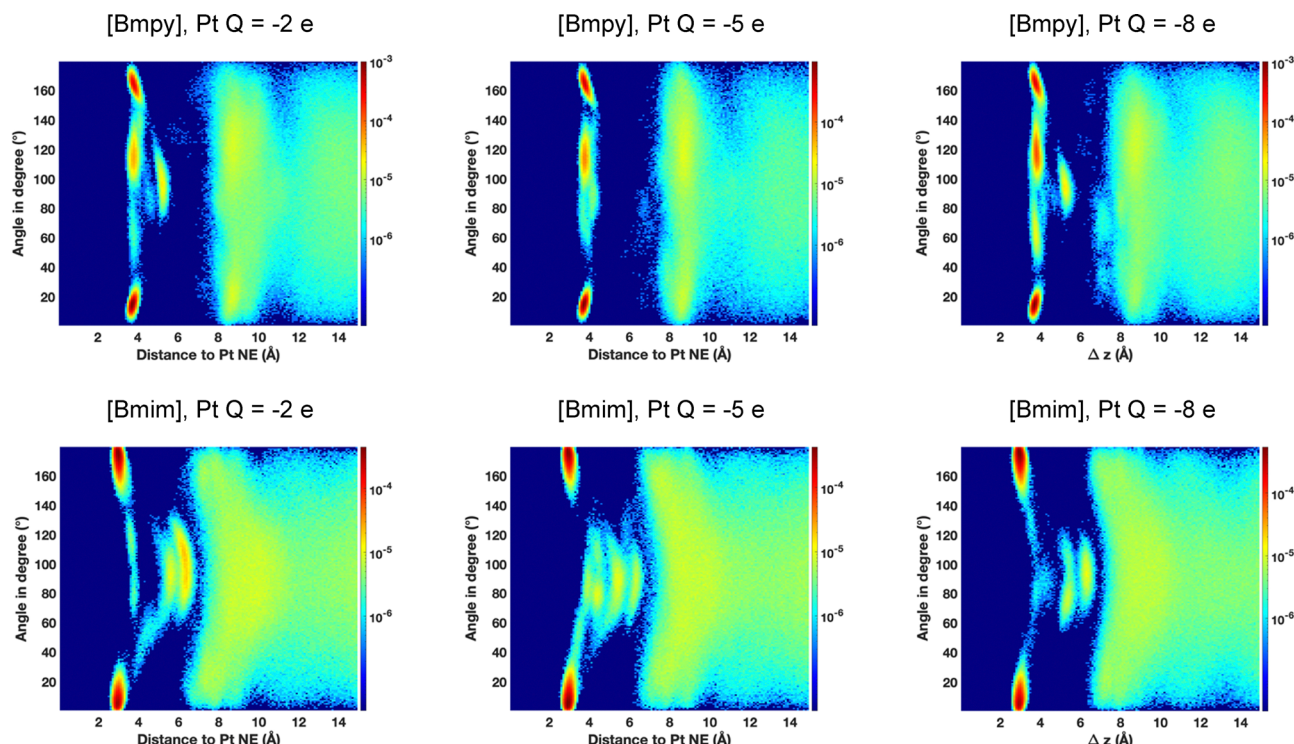


Fig. 6 IL cation orientation distribution near the Pt NE region: [Bmpy] (top row) vs. [Bmim] (bottom row), under different Pt electrode charge conditions.

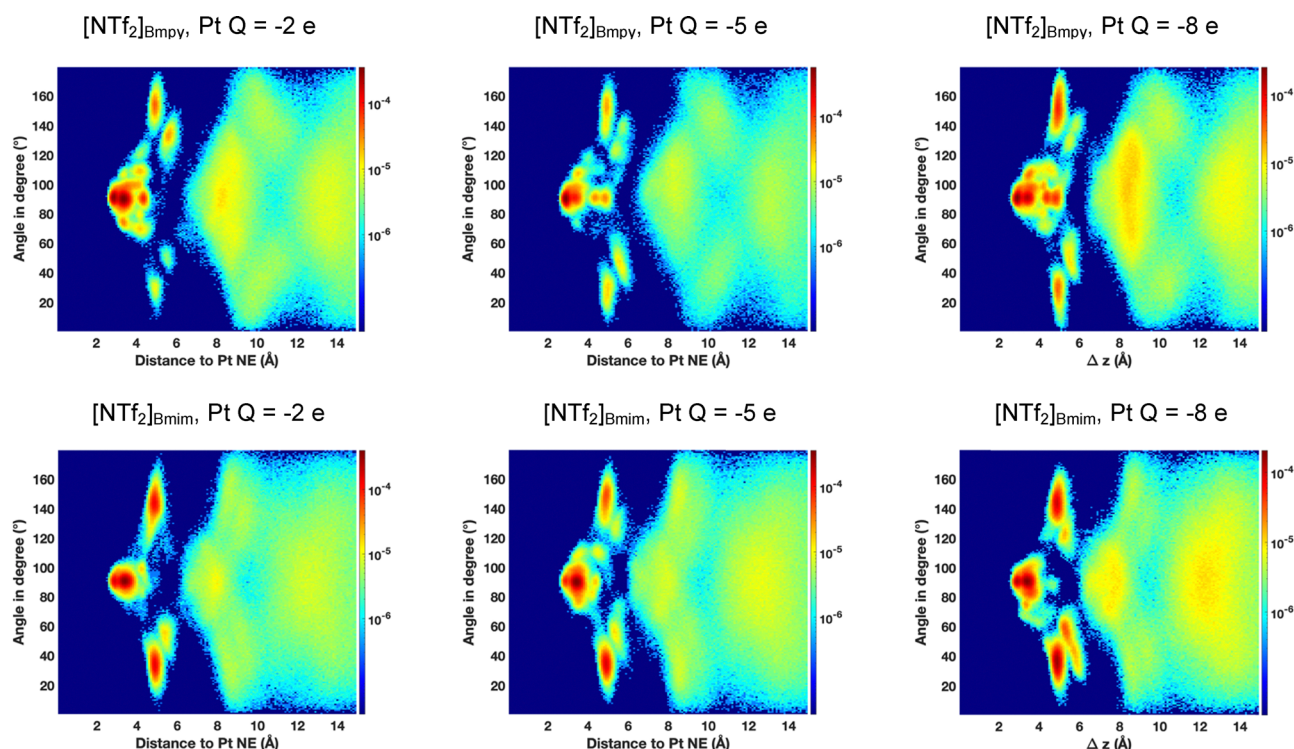


Fig. 7 IL anion orientation distribution near the Pt NE region: [NTf₂]_{Bmpy} (top row) vs. [NTf₂]_{Bmim} (bottom row), under different Pt electrode charge conditions.

sub-peaks at 2.8 Å and 3.3 Å in the *N* distributions as shown in the bottom of Fig. 4a, and the associated anions prefer staying parallel to the electrode surface with small angular movement.

When farther away from the NE, there exist multiple orange/yellow ribbons in the region of 4–6 Å away from the electrode, which are symmetrical with respect to 90° degrees. All these



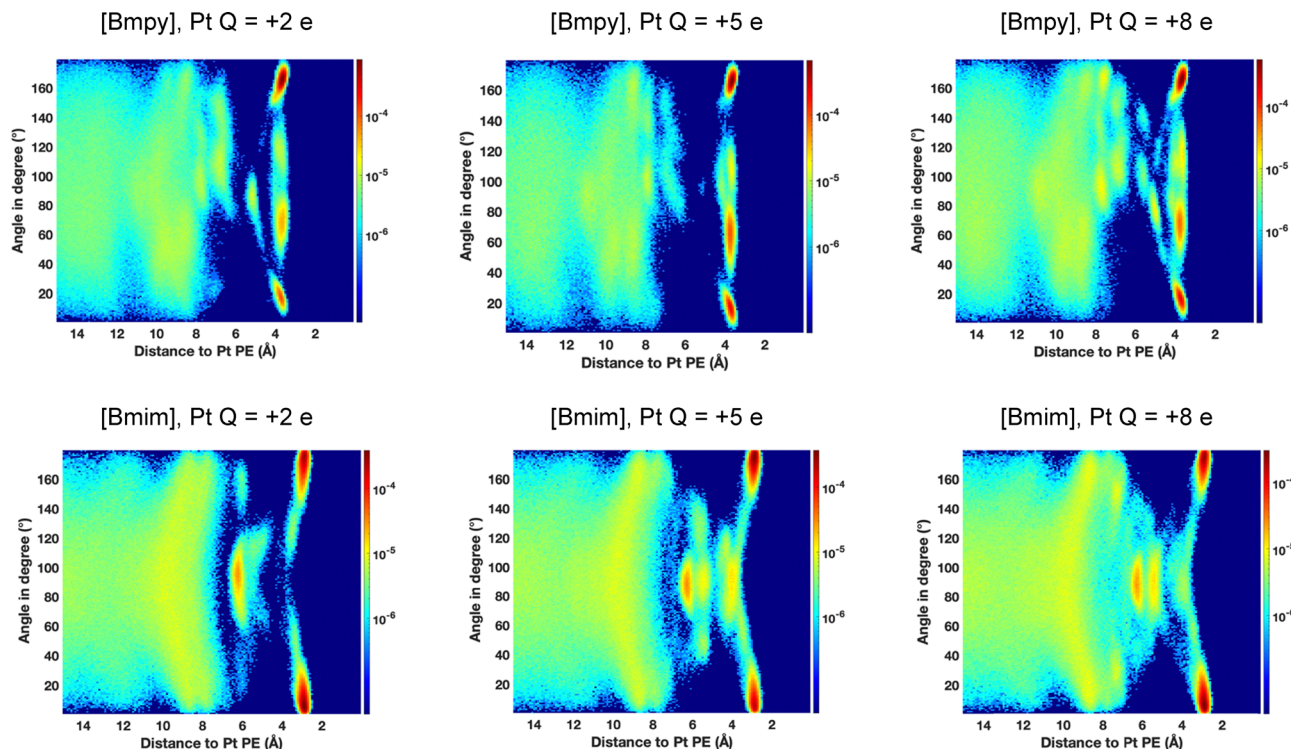


Fig. 8 IL cation orientation distribution near the Pt PE region: [Bmpy] (top row) vs. [Bmim] (bottom row), under different Pt electrode charge conditions.

anions in the region, pertaining to various orientations, contribute together to the formation of the third sub-peak (4.8 Å) in N distributions (Fig. 4a, bottom). In addition, when the charge on NE changes from $-5 e$ to $-8 e$, such orange stripes at 4–6 Å expand both in size and intensity for $[\text{NTf}_2]_{\text{Bmpy}}$ and $[\text{NTf}_2]_{\text{Bmim}}$, indicating that more anions are repelled into this region from the inner interfacial region (<4 Å). For anions in the PE region (Fig. 9), more anions get attracted to the electrode with more charges on PE, and the isolated yellow stripes at 4–6 Å diminish. This is anticipated, as when anions are drawn to the electrode, there are more restrictions on their movement, and thus more tendency for them to stay parallel ($\sim 90^\circ$) to the electrode surface. Besides that, under each charge condition at both electrodes, the orientation distributions of $[\text{NTf}_2]_{\text{Bmim}}$ and $[\text{NTf}_2]_{\text{Bmpy}}$ are mostly similar. Inside the 1st EDL, almost all $[\text{NTf}_2]$ anions stay parallel to the electrode surface on the PE side, while on the NE side, there is a much higher possibility for $[\text{NTf}_2]$ to stay non-parallel to the Pt surface, although the parallel paradigm still dominates. This is understandable, as the positive charges in the PE strongly attract the $[\text{NTf}_2]$ anions, while the negative charges in the NE side may repel $[\text{NTf}_2]$ out, allowing more freedom of the anion orientation. When it comes to the 2nd EDL, like the IL cations, the molecule orientation becomes more randomly distributed in all test cases, while there is indeed a slightly higher chance of $[\text{NTf}_2]$ staying parallel to the Pt surface than other orientations.

Considering the results from both the N and orientation distributions, the difference in using various electrode charge values will only lead to changes in magnitude, but the EDL heterostructures and their main structural features remain

unchanged. No matter which electrode charge value is used, the electrolyte|electrode interphase could always be classified into three regions: The 1st EDL stays closest to the electrode and displays sharp peaks in N distributions with a predominate parallel molecule orientation. The 2nd EDL stays beyond 6–7 Å way from the electrode, and the bulk region stays even further away. Thus, for the subsequent analysis of the EDL structure, we will only base them on the case where the electrode charge is $\pm 8 e$.

3.2.3. Molecule displacement. In the previous two sections examining the number/orientation distribution, the subject was the probability of molecular properties at fixed space. To gain a supplementary perspective, we focused on the molecules instead and observed how they transverse the space. Specifically, we first selected IL molecules of interest with the following criteria: if the IL molecules once reach within 15 Å of the Pt electrode over the entire simulation run, they are selected. This ensures that we can retrieve all IL molecules that once stayed close to the IL|Electrode interface, and the remaining IL molecules are considered as staying in the bulk region. Then, we studied molecule positions along the z -axis across the whole simulation, as the z -direction is where we are really interested. An example is given for [Bmpy] of the PE region in Fig. 10. The x -axis represents the molecule ID (we assigned an ID for each molecule group for easier analysis afterwards), and thus each symbol represents a molecule. Since there are 1000 cations and 1000 anions and there are three parallel runs, the ID ranges from 1 to 3000. Meanwhile, the y -axis represents the mean distance to the electrode for an IL molecule. The error bar is represented by $\pm \text{std}$, where std is the standard deviation.



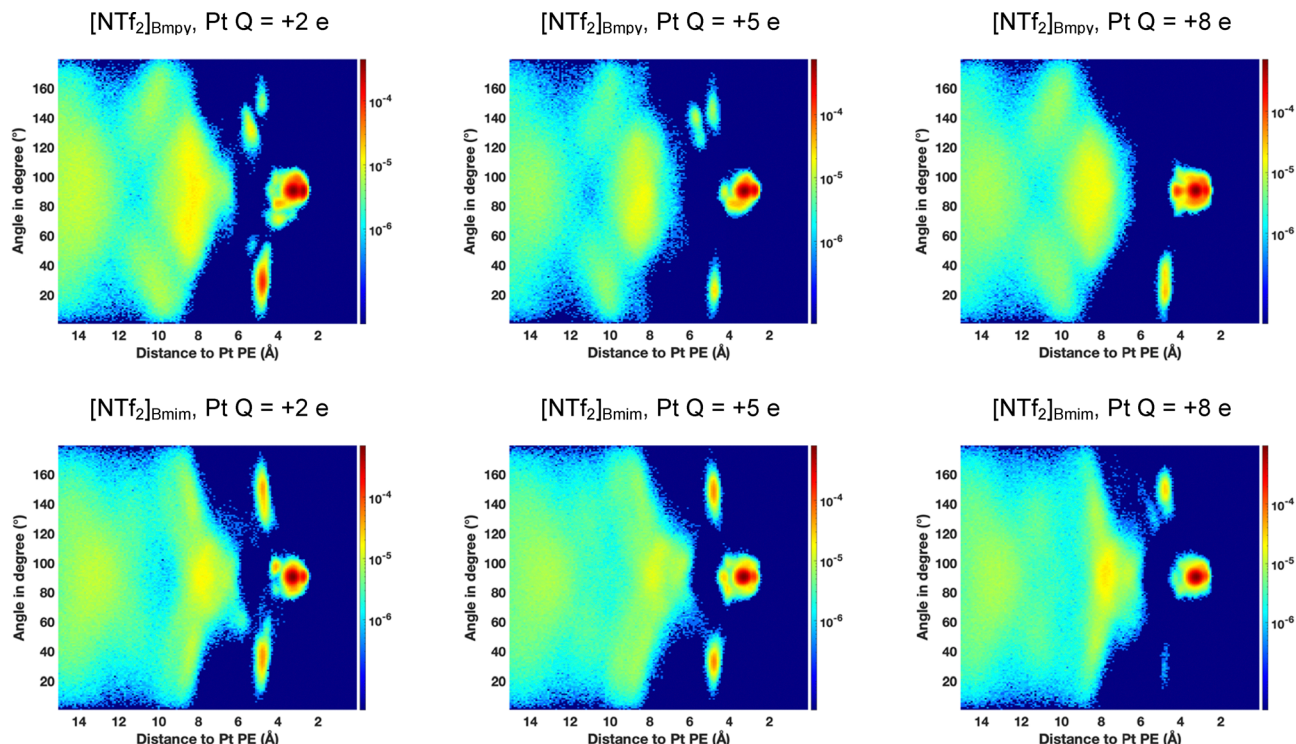


Fig. 9 IL anion orientation distribution near the Pt PE region: $[\text{NTf}_2]_{\text{Bmpy}}$ (top row) vs. $[\text{NTf}_2]_{\text{Bmim}}$ (bottom row), under different Pt electrode charge conditions.

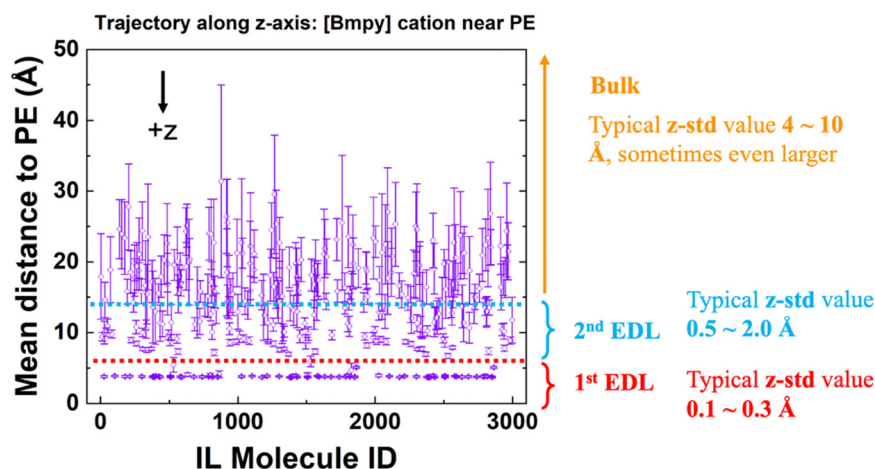


Fig. 10 IL molecular displacements along the z-axis: using [Bmpy] near the PE region (Pt Q = +8 e) as an example.

Fig. 10 also shows that the heterostructure could be divided into three regions depending on the distance to the Pt electrodes, similar to the findings from the number/orientation distribution. When closest to Pt electrodes, the error bar of the molecules is quite small, with typical std values of $\sim 0.1\text{--}0.3$ Å, meaning that the molecular movement along the z-axis is highly restricted. This corresponds to the 1st EDL with a boundary located around 6 Å away from the electrode. Farther away from the electrode, there are many IL molecules exhibiting larger error bar lengths, with the typical std values becoming 0.5–2.0 Å. This shows that molecules were granted greater

freedom of movement along the z-axis at a greater distance from the electrode but were still somewhat limited. This matches the 2nd EDL in the N distributions. As compared to the 1st EDL|2nd EDL boundary, the boundary between the 2nd EDL and bulk region becomes vague, but it is still reasonable to say that it is around 15.0 Å combining the findings from the number/orientation distribution. Beyond the 2nd EDL, the error bar of an IL molecule appears noticeably larger, with std values becoming 4–10 Å, or even larger. This suggests that the molecular displacement become completely unconstrained along the z-axis in the bulk region.



3.2.4. Structure of the whole simulation cell. With results from the number/orientation distribution and the molecule displacement, we can now depict the whole picture of the whole simulation cell, as shown in the schematic below. Pointing towards +z, the Pt-Electrode|IL|Pt-Electrode heterostructure could be divided into five regions as shown in Fig. 11: (1) 1st EDL in NE (NE-1L), (2) 2nd EDL in NE (NE-2L), (3) Bulk, (4) 2nd EDL in PE (PE-2L), and (5) 1st EDL in PE (PE-1L).

In the 1st EDL at both electrodes (NE-1L and PE-1L), the IL molecular movement along the z-axis is highly restricted with little displacement. Most molecules tend to stay parallel to the electrode surface, and their angular movement is constrained. The 1st EDL extends about 6 Å away from the electrode surface in our test, although this boundary value might vary for different types of IL. Beyond the 1st EDL extending further into the bulk (6–16 Å), is where the 2nd EDL (NE-2L, PE-2L) is located, where the IL molecular movement along the z-axis is still limited, but with more freedom as compared to that in the 1st EDL. Meanwhile, the IL molecules show a mostly random orientation distribution, with slight preference of staying parallel to the electrode surface, and their angular movement is unconstrained as compared to that in the 1st EDL. With such characteristics, it is hard to set a strict boundary for the 2nd EDL with the bulk phase, like the 1st EDL|2nd EDL boundary. The 2nd EDL|Bulk boundary is located around 16 Å away from electrode surface, but could be longer depending on the IL species, temperature, *etc.* Beyond the 2nd EDL is the bulk region located in the middle of the simulation cell, where the effect of the electrodes diminishes and can be considered as neglected. In the bulk region, the IL molecular movement becomes large and unconstrained. Meanwhile, the molecule orientations become randomly distributed completely with large, unconstrained angular movement.

This kind of 3-region and 2-double-layer structure phenomenon aligns with previously reported studies.^{29,34,82–85} According to

previous studies,^{83,84} the IL|Electrode interfacial region could be split into three discrete regions: (1) the boundary/Stern (surface-adsorbed) layer, corresponding to our 1st EDL; (2) the transition (near-surface) zone, corresponding to our 2nd EDL; and (3) the bulk phase, which could be confirmed by experiments like atomic force microscopy (AFM) measurement.^{34,85}

3.2.5. Diffusion of the whole simulation cell. With the classification of these five different regions, we then extracted IL molecules that belong to each of the regions, and computed their self-diffusivity (*D*) values for both cations (Fig. 12a) and anions (Fig. 12b) at 340 K along the *x/y* axis and *z*-axis in each region. The IL molecule categorization criterion is set to be below this to avoid possible data pollution, as it is possible that an IL molecule's trajectory may span multiple regions during a simulation run. Given the previous observation on the region boundaries, if the center-of-mass (COM) of a molecule remains within 5.5 Å of the NE or PE surface throughout the simulation, this molecule is considered to belong to region NE-1L or PE-1L. If a molecule's COM resides 6.5–16 Å distant from the NE or PE surface over the course of the simulation, then this molecule is designated as belonging to region NE-2L or PE-2L. If a molecule's COM stays beyond 17 Å away from both electrodes all the time, it will be assigned to the bulk region. It should be noted that 1 Å of gap is reserved between different regions to make the IL categorization easier.

In each figure, the blue and red dashed lines refer to the IL molecule *D* values in the pure [Bmpy][NTf₂] and [Bmim][NTf₂] case (without adding Pt electrodes and hydrogen particles), extracted from Fig. 2a. The calculated *D* values in the 2nd EDLs & 1st EDLs look symmetrical at the NE side and PE side, with respect to the bulk phase. For both IL cations and anions in either [Bmpy][NTf₂] or [Bmim][NTf₂], the *D* values in the 2nd EDL are 1–2 orders of magnitude smaller than that in the bulk phase, and the *D* values in the 1st EDL decrease further by 1–2

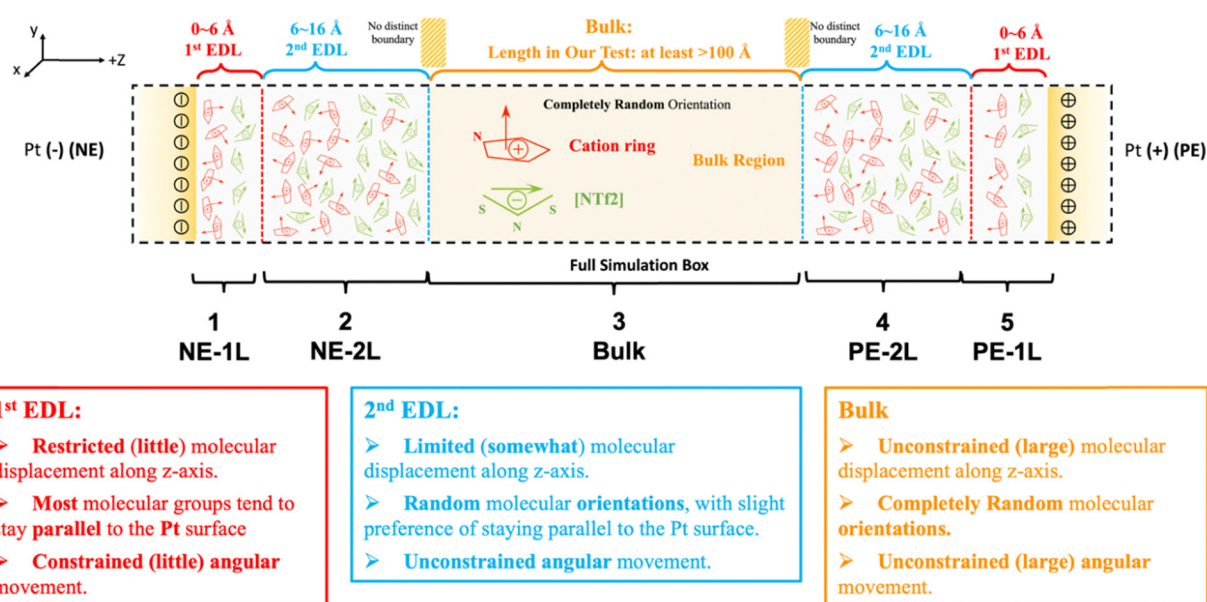


Fig. 11 Schematic of the whole simulation cell based on the results of the number distribution, molecule orientation, and molecule displacement.

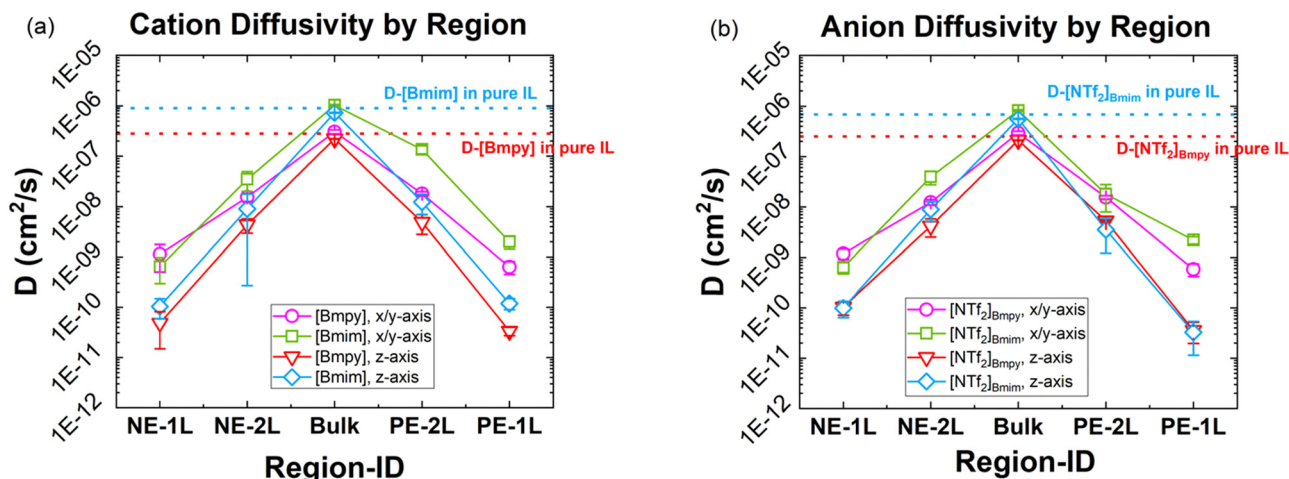


Fig. 12 Self-diffusivity values of IL (a) cations and (b) anions along the x/y axis and z-axis in the five regions (Pt $Q = \pm 8 e$).

orders of magnitude as compared to the 2nd EDL. In addition, the D values along the z-axis are always smaller than their counterparts along the x/y axis. For example, in the 1st EDL, the difference is at least 1 order of magnitude. Such trends are as expected, as the Pt electrode is placed on the xy-plane. The anisotropy becomes reduced in the 2nd EDL and negligible in the bulk. D values in the bulk region are consistent with those in bulk ILs (dashed lines) from Fig. 2c.

3.2.6. H₂ probability in the whole simulation cell. To understand the H₂ distribution in the electrochemical cell, we also calculated the probability density (\AA^{-3}) of H₂ molecules, P , showing up in the five different regions based on their distance to the electrode along the z-axis. This is done by first obtaining the number (N) density distribution for H₂ and then averaging it over all the bins (bin size: 53 Å (x-axis) × 53 Å (y-axis) × 0.1 Å (z-axis)) contained in a specific region like NE-1L. This directly reflects the possibility of a H₂ showing up in a specific region, and the results for all electrode charge conditions are shown in Fig. 13a. On the one hand, the probability density of H₂ molecules varies only slightly in the 2nd EDL as compared to the bulk region, regardless of the electrode charge conditions. On the other hand, when it comes to the 1st EDL of both electrodes, the H₂ probability density drops dramatically for all cases, indicating that the 1st EDL indeed behaves like a barrier limiting the H₂ access to the electrode because of its highly ordered structure. In the meantime, under each electrode charge condition, there is always a higher H₂ probability density in [Bmpy][NTf₂] than that in [Bmim][NTf₂] in the 1st EDL of both electrodes. To make a better comparison between the two IL species, we calculated the ratio (R) of the H₂ probability density in [Bmpy][NTf₂] to the probability density in [Bmim][NTf₂] for each region under all the electrode charge conditions by $R = \frac{P(\text{H}_2 \text{ in } [\text{Bmpy}][\text{NTf}_2])}{P(\text{H}_2 \text{ in } [\text{Bmim}][\text{NTf}_2])}$. The results in Fig. 13b show that the ratio of $P([\text{Bmpy}][\text{NTf}_2])$ to $P([\text{Bmim}][\text{NTf}_2])$ is clearly larger than 1.0 in the 1st EDL of both electrodes under all electrode charge conditions, while the ratios in the 2nd EDL and bulk are close to 1.0. This suggests

that the H₂ movement activity in both the bulk and 2nd EDL regions should be similar for the two IL species, and it should be the H₂ probability density in the 1st EDL that contributes to the different H₂ sensibility of the two IL species. In addition, among all regions, the PE-1L region captures the most interest, as this is where an H₂ molecule gets oxidized. It shows that H₂ molecules have a higher possibility of staying in the PE-1L in [Bmpy][NTf₂] than in [Bmim][NTf₂] under all electrode charge conditions. Among all the electrode charge conditions, the H₂ probability density difference between [Bmpy][NTf₂] and [Bmim][NTf₂] in PE-1L is the highest at $Q = +5 e$, while the lowest is at $Q = +8 e$. The H₂ probability density is affected by the molecular structure of cations and anions and their energetic interaction with H₂. Thus, it might be difficult to pinpoint exactly the atomic origin as to why the density difference between these two ionic liquids behaves this way. However, as discussed previously in the N distribution (Fig. 4b), increased electrode charge tends to repel cations and attract anions, indicating a diminishing difference in PE-1L between [Bmpy][NTf₂] and [Bmim][NTf₂] as they have the same anions. This explains why there is the lowest probability density difference at the highest charge condition ($Q = +8 e$).

Another way to look at H₂ probability is to examine the integrated N density distribution as shown in Fig. 14, using Pt $Q = +8 e$ as an example. There is no H₂ in the region within 2 Å away from the Pt PE electrode. This is understandable as both Pt and H atoms have their van der Waals sizes. Within PE-1L (circled), H₂ mostly stay in a narrow region of 2–3 Å, right on the Pt surface, and it appears that the probability of finding H₂ is higher in [Bmpy][NTf₂] than in [Bmim][NTf₂].

3.2.7. H₂ sensing. Examining the results we previously discussed together allows us to gain insight into the possible rate-limiting steps in H₂ sensing and why [Bmpy][NTf₂] is a better option than [Bmim][NTf₂]. As discussed previously, H₂ needs to be solvated in the ionic liquids and then move through the bulk and double layers to reach the Pt PE electrode. First, when it comes to solvation, H₂ has slightly higher solubility in [Bmpy][NTf₂] than [Bmim][NTf₂] (Table 1). As the non-polar H₂

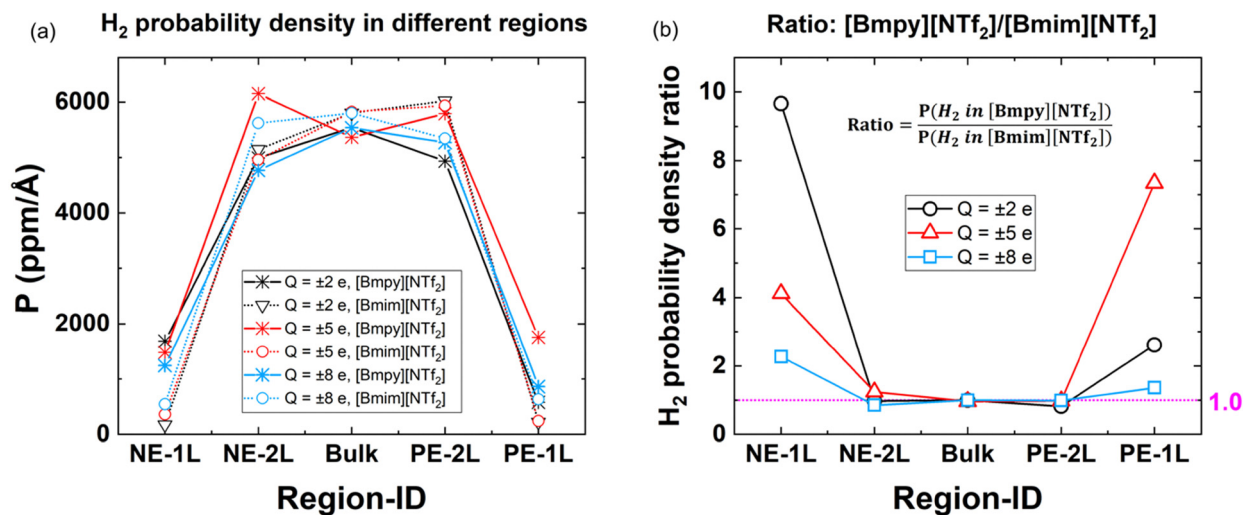


Fig. 13 (a) H₂ probability density (Å⁻¹) in different regions under different Pt electrode charge conditions: $Q = \pm 2$ e (black), $Q = \pm 5$ e (red), and $Q = \pm 8$ e (blue); (b) the ratio of H₂ probability density in [Bmpy][NTf₂] to probability density in [Bmim][NTf₂] for different regions under different Pt electrode charge conditions: $Q = \pm 2$ e (black), $Q = \pm 5$ e (red), and $Q = \pm 8$ e (blue).

overall has low solubility in these two ionic liquids, *i.e.*, on the ppm level even with 0.1 MPa of H₂ gas, the solvation of H₂ could be the rate-limiting step so a slight increase in its solubility has the potential to increase the performance. Second, H₂ seems to have high diffusivity, $\sim 10^{-5}$ cm² s⁻¹ at 300 K, and moves about two orders of magnitude faster than the bulk cations/anions (Fig. 2a and 3a) and even more than the cations/anions in the double layers (Fig. 12), suggesting H₂ diffusion is not likely to be the rate-limiting step. In other words, the diffusivity difference of ions between [Bmim][NTf₂] and [Bmpy][NTf₂] might not matter as the mass transport of H₂ is sufficiently fast. Third, we think that the H₂ probability (Fig. 13 and 14) could be used as a surrogate for the solubility for different regions in the simulation cell, and it shows that the H₂ solubility is higher in [Bmpy][NTf₂] than [Bmim][NTf₂] in the PE-1L region, a similar trend as in the bulk. The difference between two ionic liquids is

affected by the electrode charges. Finally, the orientation map (Fig. 6) suggests that the [Bmim] rings mostly lay flat on the Pt surface, potentially blocking the H₂ access and leading to inferior performance. In summary, we believe the higher H₂ solubility in the bulk and 1st double layer and cation orientation make [Bmpy][NTf₂] a better option than [Bmim][NTf₂].

Conclusions

In order to explain why [Bmpy][NTf₂] is a better electrolyte material than [Bmim][NTf₂] as observed in hydrogen sensing experiments, MD simulation was employed in this study to investigate the different structures in [Bmpy][NTf₂] and [Bmim][NTf₂] based hydrogen sensors. First, both IL cations and anions, along with the H₂ molecules, demonstrate a higher self-diffusivity in [Bmpy][NTf₂] than in [Bmim][NTf₂], excluding the possibility that it is the molecule diffusion in bulk IL that makes a difference. Second, the electrolyte|electrode interphase is investigated and compared between the two IL species. The heterostructure can be classified into three distinct regions depending on the distance to the electrode surface: the 1st EDL, the 2nd EDL, and the bulk phase, and their features can be well described using descriptors including IL number distributions, IL molecule orientation distributions, and their displacements along the *z*-axis. Furthermore, the whole simulation cell is divided into five distinct regions from NE to PE, with one 1st EDL and one 2nd EDL for each electrode, and the self-diffusivities of IL cations and anions are then calculated for each region, respectively. To the best of our knowledge, this is the first time of reporting the diffusion-related difference in the 1st and 2nd EDL, and the bulk phase, besides their structural difference like IL ion arrangements.^{29,34,82–85} The electrolyte|electrode interfaces of the two IL species are mostly similar except for the cations in the 1st EDL, in which [Bmpy] ions have a more scattered orientation distribution than [Bmim] ions. Considering that a parallel cation configuration has a higher chance of

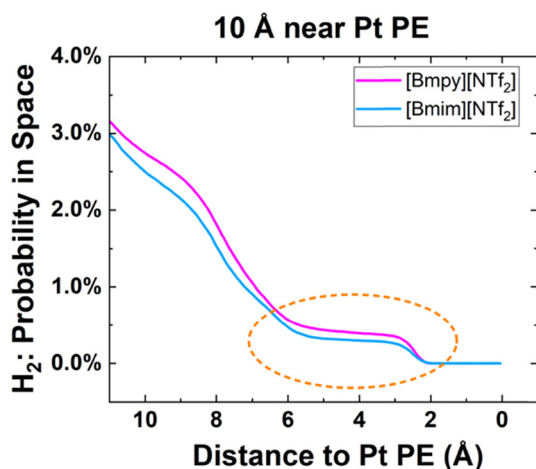


Fig. 14 Comparison of the integrated H₂ probability within 10 Å of the PE region (Pt $Q = \pm 8$ e): [Bmpy][NTf₂] vs. [Bmim][NTf₂].



obstructing the H₂ transport pathway to the Pt surface, and H₂ molecules need to first diffuse through the 1st EDL to reach the PE before the oxidation, more H₂ transport pathways to the Pt surface are made available in [Bmpy][NTf₂], resulting in a higher H₂ redox reaction possibility.

When it comes to the H₂ molecules, first, H₂ indeed demonstrates a slightly higher solubility in [Bmpy][NTf₂]. Additionally, as part of a comparative study, even with the same amount of H₂ molecules injected into the electrochemical simulation cell, our probability density results still show that there is a higher likelihood of finding H₂ molecules inside the 1st EDL of the PE in [Bmpy][NTf₂] than in [Bmim][NTf₂]. These results align with our observations on the IL molecule orientations and thereby altogether elucidate why [Bmpy][NTf₂] is a more sensitive electrolyte in hydrogen sensors than [Bmim][NTf₂].

Data availability

Input data for this article are available at https://github.com/yininghe/Simulation_Compare_2ILs.

Conflicts of interest

There are no conflicts to declare.

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