

## PAPER

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9291A tetranuclear Pr–W heterometal cluster-imbedded  
antimotungstate for the catalytic synthesis of  
benzimidazoles†Zhijian Zheng,<sup>a</sup> Song Jiang,<sup>‡a</sup> Xuejiao Chen,<sup>‡a</sup> Haoqi Liu,<sup>a</sup> Yufeng Liu,<sup>a</sup>  
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A novel organic–inorganic Pr–W heterometal cluster-imbedded antimotungstate,  $[(\text{CH}_3)_2\text{NH}_2]_4\text{Na}_{20.5}\{[(\text{OAc})\text{Pr}_4\text{W}_6\text{O}_{12}(\text{tar})_3](\text{SbW}_9\text{O}_{33})_4(\text{H}_2\text{O})_{10}\}\text{Cl}_{0.5}\cdot\text{ca. } 40.3\text{H}_2\text{O}$  (**Pr-1**), was synthesized via a one-step *in situ* assembly strategy. **Pr-1** features a distinctive tetranuclear Pr-based organic–inorganic hybrid structure stabilized by polycarboxylic acid ligands, which efficiently regulate the hydrolysis of  $\text{Pr}^{3+}$  to facilitate the assembly of Pr–W heterometallic clusters. Notably, this structure includes abundant Lewis acid sites from  $\text{Pr}^{3+}$ , making it highly promising for catalytic applications. **Pr-1** has been successfully demonstrated as an environmentally friendly, heterogeneous catalyst in synthesizing benzimidazoles from *o*-phenylenediamines and benzaldehydes. The high catalytic activity, excellent structural stability, and recyclability over five cycles highlight the potential of **Pr-1** as an efficient and sustainable catalyst for N-containing heterocyclic organic synthesis.

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## Introduction

Polyoxometalates (POMs) are an important class of metal–oxygen clusters with excellent oxidative decomposition resistance and thermal stability, enabling extensive applications in pharmaceuticals, catalysis, materials science *etc.*<sup>1–8</sup> Currently, lanthanide metal ions (Ln), owing to their flexible coordination capabilities and high coordination numbers, are increasingly utilized as connection units in the synthesis of novel POMs. The incorporation of Ln ions not only facilitates the assembly of POM building blocks into more intriguing structures but also endows POMs with unexpected properties.<sup>9–13</sup> Therefore, Ln-containing POMs (Ln-POMs) have become an important branch of POM chemistry.<sup>14–20</sup> Among Ln-POMs,  $\text{Pr}^{3+}$  ions can act as metal cores, which play a crucial role in connecting polyoxoanions to generate multinuclear

metal clusters, giving rise to diverse and novel structures.<sup>21–25</sup> However, the application of  $\text{Pr}^{3+}$  ions in constructing novel POMs still suffers from challenges. On one hand, the large spatial steric hindrance of  $\text{Pr}^{3+}$  becomes a stumbling block for the effective construction of the corresponding Pr-containing POMs (Pr-POMs). On the other hand, the oxophilic nature of  $\text{Pr}^{3+}$  rapidly interacts with oxygen-rich POM building blocks to form amorphous precipitates rather than well-defined crystals during the synthesis of Pr-POMs. This complicates the determination of the crystal structure of the target product, highlighting the necessity of developing effective strategies for the synthesis of Pr-POMs.

The introduction of polycarboxylic acid ligands into the reaction for the purpose of regulating the hydrolysis of Ln ions has proven to be a highly effective strategy.<sup>9,26–30</sup> In the context of such reactions,  $\text{Pr}^{3+}$  initially coordinates with organic ligands to form multinuclear Pr oxide clusters, which subsequently connect the generated  $\text{WO}_6$  units to produce novel Pr–W heterometallic clusters by adjusting the pH. Therefore, these heterometallic clusters attract surrounding polyoxoanions, consequently leading to the formation of diverse structures. Besides, the metal ions in the clusters furnish abundant Lewis acid activation sites, which exhibit substantial potential within the domain of catalytic organic reactions.

Benzimidazoles represent a significant category in N-containing heterocyclic compounds due to their presence in a variety of natural products and extensive applications in medicinal chemistry.<sup>31,32</sup> Therefore, benzimidazoles display a

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wide array of pharmacological activities, encompassing anti-fungal, antitumor, anti-tuberculosis, anti-inflammatory, and antibacterial properties.<sup>33–36</sup> Currently, the principal approaches for the synthesis of benzimidazoles mainly rely on catalysis utilizing transition metals, noble metals, lanthanide metals, and other Lewis acids accompanied by notable limitations.<sup>37,38</sup> As a result, the exploration of efficient and environmentally friendly methods for the synthesis of benzimidazole derivatives holds considerable significance.

In contrast to conventional catalysts mentioned above, Pr-POMs possess distinct advantages such as excellent catalytic activity, easy separation and recycling, and environmental friendliness.<sup>39–43</sup> Herein, a novel organic–inorganic Pr–W heterometal cluster-embedded antimonotungstate,  $[(\text{CH}_3)_2\text{NH}_2]_4\text{Na}_{20.5}\{[(\text{OAc})\text{Pr}_4\text{W}_6\text{O}_{12}(\text{tar})_3](\text{SbW}_9\text{O}_{33})_4(\text{H}_2\text{O})_{10}\}\text{Cl}_{0.5}\cdot\text{ca. } 40.3\text{H}_2\text{O}$  (**Pr-1**) has been synthesized *via* a one-step *in situ* assembly strategy. **Pr-1** contains an unprecedented tetranuclear Pr organic–inorganic hybrid structure,  $\{(\text{OAc})\text{Pr}_4\text{W}_6(\text{tar})_3\}$  (HOAc =  $\text{CH}_3\text{COOH}$ ,  $\text{H}_3\text{tar}$  = tartaric acid). The tetrameric polyoxoanion is aggregated by four  $\{\text{SbW}_9\}$  units encapsulating a  $\{\text{Pr}_4\text{W}_6\}$  unit, which is further expanded to a one-dimensional chain constructed by  $\text{Na}^+$  cations. As a heterogeneous catalyst, **Pr-1** could efficiently catalyze the condensation cyclization of *o*-phenylenediamine with benzaldehyde to produce benzimidazole. Moreover, **Pr-1** still retains high catalytic activity even after five cycles, demonstrating excellent stability.

## Experimental

### Synthesis of Pr-1

$\text{Na}_2\text{WO}_4\cdot 2\text{H}_2\text{O}$  (6 mmol, 1.9791 g),  $\text{SbCl}_3$  (0.8 mmol, 0.1825 g),  $(\text{CH}_3)_2\text{NH}\cdot\text{HCl}$  (12 mmol, 0.9785 g), and DL-tartaric acid (0.3 mmol, 0.0451 g) were sequentially dissolved in a 0.5 M NaAc/HAc buffer solution at pH = 4.8. The solution was adjusted to pH = 4.5 using 6 M HCl and stirred for 10 min.  $\text{PrCl}_3\cdot 6\text{H}_2\text{O}$  (0.6 mmol, 0.2132 g) was subsequently added to the solution, and the pH was adjusted to 4.0 after the solids were completely dissolved. The clear green solution was stirred for another 10 min before being filtered. The green filtrate was left to evaporate at room temperature. Colorless stick crystals were collected after about six weeks (yield: 35% based on  $\text{PrCl}_3\cdot 6\text{H}_2\text{O}$ ). Elemental analysis (%) for  $\text{C}_{22}\text{H}_{140.6}\text{N}_4\text{Cl}_{0.5}\text{Na}_{20.5}\text{O}_{214.3}\text{Pr}_4\text{Sb}_4\text{W}_{42}$ : found (calc.): C, 2.05 (2.01); H, 1.01 (1.08), N, 0.39 (0.43); Na, 3.62 (3.58); Pr, 4.28 (4.29); W, 58.72 (58.71). FT-IR ( $\text{cm}^{-1}$ ): 3388 (m), 3155 (m), 2794 (w), 1611 (m), 1538 (m), 1464 (m), 1407 (m), 1370 (m), 1169 (w), 1136 (w), 1071 (w), 1017 (w), 936 (s), 833 (vs), 763 (vs), 680 (vs), 598 (vs).

### Synthesis of benzimidazoles

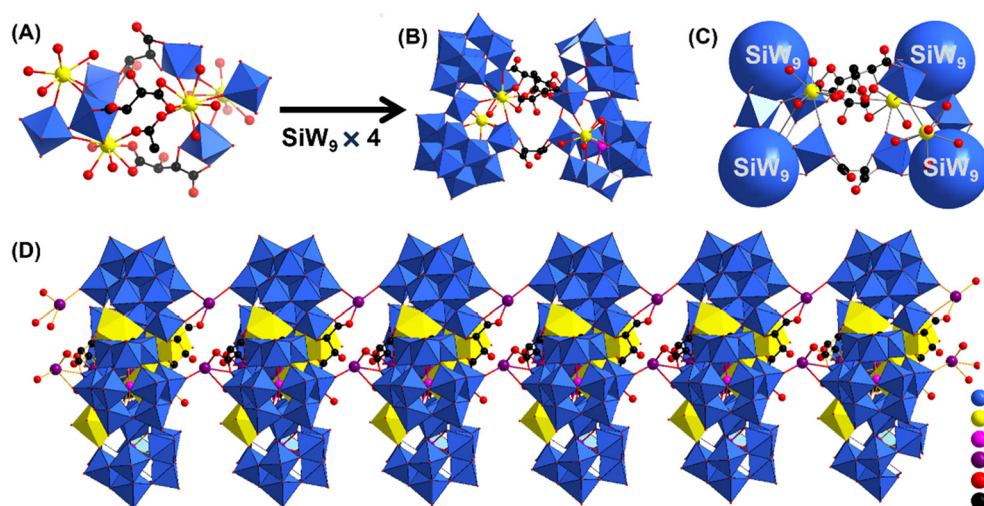
Benzaldehyde (**1a**, 0.2 mmol), *o*-phenylenediamine (**2a**, 0.2 mmol), **Pr-1** (0.2 mol%), and EtOH (1 mL) were added into a 4 mL reaction vial. Then the reaction was carried out in screw cap vials with a Teflon seal at 100 °C for 1 h under an

oxygen atmosphere. After the reaction, the mixture was purified by column chromatography (petroleum ether/EtOAc) or recrystallization to afford the desired products. After the reaction was completed, the catalyst was separated from the reaction solution by filtration and then washed three times with EtOAc and EtOH, respectively, and finally the recovered catalyst was refreshed by drying in a vacuum drying oven at 50 °C for 3 h.

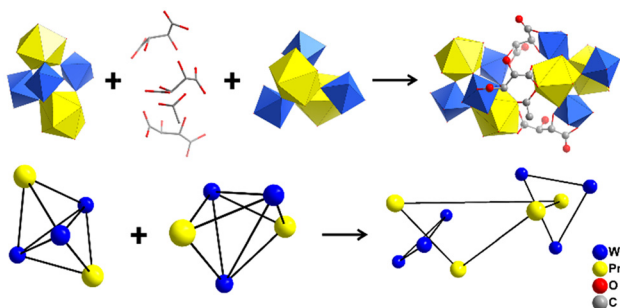
## Results and discussion

### Crystal structure

Single-crystal X-ray diffraction analyses reveal that **Pr-1** crystallizes in the triclinic space group  $P\bar{1}$  (Table S1†). The average bond valence sum (BVS) values for W, Pr, and Sb atoms are 5.97 (within the range of 5.53–6.28), 3.30 (within the range of 3.26–3.34) and 2.80 (within the range of 2.77–2.85), indicating that they are in +6, +3 and +3 oxidation states, respectively (Table S2†). In the polyanionic skeleton of **Pr-1**, four  $\{\text{SbW}_9\text{O}_{33}\}$  units are interconnected *via* a unique organic–inorganic hybrid linking fragment  $(\text{OAc})\text{Pr}_4\text{W}_6(\text{tar})_3$  (Fig. 1A and S1†), which consists of a single acetic anion, three tartrate anions, four  $\text{Pr}^{3+}$  cations and six  $\{\text{WO}_6\}$  octahedra. The organic–inorganic hybrid fragment features two sets of corner-sharing W and Pr centers, denoted as  $\{\text{Pr}_2\text{W}_3\text{O}_{29}\}$ , which are interconnected by three tartrate anions and one acetic anion (Fig. 2). This linkage is facilitated by a Pr–O–C–O–Pr bridge emanating from the acetic anion, along with one W–O–C–O–W bridge and two Pr–O–C–O–W bridges established by the tartrate anions. In detail, two of the four praseodymium ions, specifically Pr1 and Pr3, exhibit a distorted mono-capped square-antiprism geometry, being coordinated by nine oxygen atoms (Fig. S2†). In this configuration, the Pr1 ion is bonded to five  $\mu_2$ -O atoms (O7 and O8 from the  $\{\text{SbW}_9\text{O}_{33}\}$  subunit and O3, O35, and O68 from the  $\{\text{WO}_6\}$  fragment) with the bond distances ranging from 2.436 Å to 2.581 Å. Additionally, it forms a bond with one water molecule (O1W) at a distance of 2.589 Å, one oxygen atom (O1) from the acetic acid ligand at a distance of 2.354 Å, and two oxygen atoms (O1L and O2L) from the tartaric acid ligand with distances of 2.572 Å and 2.587 Å. On the other hand, the remaining two praseodymium ions, Pr2 and Pr4, adopt an eight-coordinate distorted double-capped triangular prism geometry (Fig. S2†). In this arrangement, Pr2 is coordinated by five  $\mu_2$ -O atoms (O39 and O40 from the  $\{\text{SbW}_9\text{O}_{33}\}$  subunit and O4, O6, and O67 from the  $\{\text{WO}_6\}$  fragment) with the bond distances varying from 2.323 Å to 2.495 Å. Furthermore, it forms bonds with three water molecules (O2W, O3W, and O4W) with distances ranging from 2.452 Å to 2.589 Å. It is noteworthy that the tetrameric polyanions (Fig. 1B and C) are organized into a one-dimensional (1D) architecture through the interconnection of W–O–Na–O–W and W–O–Na–O–C bridging units (Fig. 1D), in which the six-coordinate  $\text{Na}^+$  cation is coordinated by O7L, O17L, O9W, O12W, O134 and the Cl1 anion with a Na–Cl bond length of



**Fig. 1** (A) The organic–inorganic hybrid fragment in **Pr-1**; (B) the polyanionic skeleton of **Pr-1**; (C) the simplified schematic depiction of **Pr-1**; and (D) the one-dimensional framework of **Pr-1** constructed by  $\text{Na}^+$  cations.



**Fig. 2** Schematic representation of the polyhedron and the topology of organic–inorganic hybrid fragments in the polyanionic skeleton of **Pr-1**.

2.718 Å. These 1D structures are arranged in parallel along the *a*-axis direction (Fig. S2†).

In terms of packing efficiency, the unit cell exhibits a void volume of approximately 14.6%, as calculated using the solvent-accessible surface. This significant void contributes to highly disordered solvent molecules (Fig. S4†), which are removed using the SQUEEZE routine, and the molecular formula is finally determined through comprehensive consideration of crystallographic data, elemental analyses, and thermogravimetric measurements.

### Catalytic activity for benzimidazole synthesis

To evaluate the catalytic performance of **Pr-1**, a model reaction was conducted using benzaldehyde (**1a**, 0.2 mmol) and *o*-phenylenediamine (**2a**, 0.2 mmol) in ethanol (EtOH) under an oxygen atmosphere (Table 1). Product **3a** was obtained with a 45% yield when catalyzed by **Pr-1** (0.2 mol%) at room temperature for 0.5 h. To enhance the reaction yield, various solvents including dimethyl carbonate (DMC), propylene carbonate (PC), isopropanol (*i*-PrOH), and ethyl acetate (EA) were investigated. Although the reaction could proceed in these sol-

**Table 1** Optimization of the reaction conditions<sup>a</sup>

| Entry | Solvent | Temperature (°C) | Time (h) | Yield <sup>b</sup> [%] |
|-------|---------|------------------|----------|------------------------|
| 1     | EtOH    | rt               | 0.5      | 45                     |
| 2     | EA      | rt               | 0.5      | 19                     |
| 3     | DMC     | rt               | 0.5      | 18                     |
| 4     | PC      | rt               | 0.5      | 29                     |
| 5     | EtOH    | 60               | 0.5      | 71                     |
| 6     | EtOH    | 80               | 0.5      | 80                     |
| 7     | EtOH    | 90               | 0.5      | 85                     |
| 8     | EtOH    | 100              | 0.5      | 89                     |
| 9     | EtOH    | 100              | 1        | 93                     |
| 10    | EtOH    | 100              | 1.25     | 93                     |

<sup>a</sup> Reaction conditions: benzaldehyde **1a** (0.2 mmol), *o*-phenylenediamine **2a** (0.2 mmol), **Pr-1** (0.2 mol%), solvent (1 mL),  $\text{O}_2$ , rt, 0.5 h.

<sup>b</sup> The reaction yields were determined by gas chromatography using biphenyl as an internal standard.

vents, the highest yield was achieved in EtOH. Additionally, the reaction was significantly influenced by temperature, with the yield of **3a** increasing as the temperature rose, reaching an optimal 89% yield at 100 °C. Finally, extending the reaction time to 1 h resulted in a maximum yield of 93%.

To investigate the substrate scope of the catalytic system, various aldehydes and phenylenediamines were employed to synthesize benzimidazoles under optimized conditions (Table 2). Benzaldehydes bearing electron-donating substituents such as  $-\text{Me}$ ,  $-\text{iPr}$ , and  $-\text{OMe}$  were successfully converted into the corresponding benzimidazoles with yields of 89%, 88%, and 90%, respectively (**3b–3d**). The electron-withdrawing groups including  $-\text{F}$ ,  $-\text{Cl}$ , and  $-\text{Br}$  were also well tolerated, affording the desired products in good yields (**3e–3i**). The electronic properties of substituents have little effect on the

**Table 2** Studies on the substrate scope of the catalytic system<sup>a</sup>

| 1               | 2               | 3           |
|-----------------|-----------------|-------------|
|                 |                 |             |
|                 |                 |             |
| 10 mmol, 1.06 g | 10 mmol, 1.08 g | 1.67 g, 86% |
|                 |                 |             |
| 3a, 91%         | 3b, 89%         | 3c, 88%     |
|                 |                 |             |
| 3d, 90%         | 3e, 88%         | 3f, 91%     |
|                 |                 |             |
| 3g, 85%         | 3h, 79%         | 3i, 87%     |
|                 |                 |             |
| 3j, 83%         | 3k, 85%         | 3l, 83%     |
|                 |                 |             |
| 3m, 85%         | 3n, 89%         | 3o, 81%     |

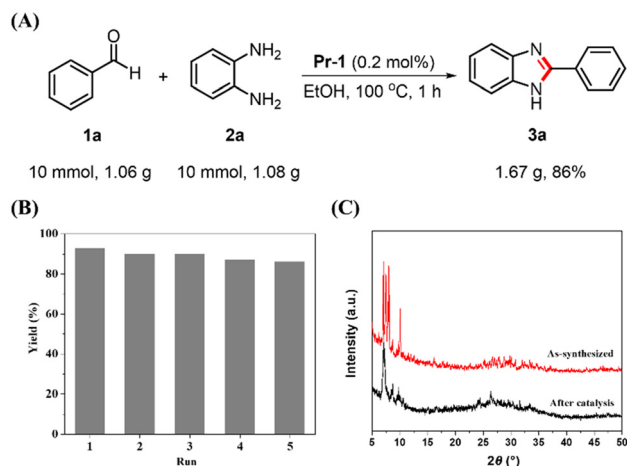
<sup>a</sup> Reaction conditions: benzaldehyde **1** (0.2 mmol), *o*-phenylenediamine **2** (0.2 mmol), **Pr-1** (0.2 mol%), EtOH (1 mL), O<sub>2</sub>, 100 °C, 1 h.

reaction yield. Notably, the reaction yields of **3f** and **3g** were higher than that of **3h**, indicating that steric hindrance had a certain effect on the reaction efficiency. Furthermore, this catalytic system was also compatible with aliphatic and heterocyclic aldehydes, affording the corresponding products **3j** and **3k** in 83% and 85% yields, respectively. Subsequently, we also explored the substrate range of *o*-phenylenediamines. The *o*-phenylenediamines containing -F and -Cl groups could react smoothly with benzaldehyde, producing the corresponding products (**3l**, **3m**) in 83% and 85% yields, respectively. The disubstituted *o*-phenylenediamines with 5,6-dimethyl and 5,6-dichlorine groups were also successfully converted into the corresponding products in good yields (**3n**, **3o**).

Finally, a scaled-up reaction and cycling experiments were conducted to evaluate the practicality of this method. 10 mmol of **1a** and 10 mmol of **2a** reacted smoothly under optimal conditions, yielding the desired product **3a** in 86% yield, indicating its potential for larger-scale synthesis (Fig. 3A). Reusability tests revealed that the catalytic performance of **Pr-1** remained consistent after five runs, indicating that the catalytic activity of **Pr-1** remained stable during the catalytic process (Fig. 3B). Furthermore, no significant changes in the PXRD spectra before and after catalysis further confirm the stability of the **Pr-1** structure (Fig. 3C).

## Conclusions

In summary, a novel Pr-W heterometallic cluster-embedded antimotungstate, **Pr-1**, has been successfully synthesized.

**Fig. 3** (A) Gram-scale reaction; (B) cycling experiments; and (C) PXRD patterns.

Polycarboxylic acid ligands were introduced to modulate the hydrolysis of Pr<sup>3+</sup>, effectively overcoming its oxophilic nature and preventing the formation of amorphous precipitates. As an environmentally friendly and heterogeneous catalyst, **Pr-1** exhibited exceptional performance in the synthesis of benzimidazoles from *o*-phenylenediamines and benzaldehydes. It exhibited high efficiency and stability, maintaining its catalytic activity over five cycles. This endows **Pr-1** with significant potential for sustainable organic synthesis. This work highlights the promising role of Ln-POMs in advancing environmentally friendly and economically viable catalytic processes, paving the way for future exploration in POM chemistry.

## Author contributions

Zhijian Zheng, Yayu Dong, and Xiaopeng Sun wrote the draft; Haoqi Liu collected the data; Zhijian Zheng, Song Jiang and Xuejiao Chen performed the IR, TG and XRD experiments; Song Jiang and Yufeng Liu investigated the catalytic activities; and Guoping Yang contributed to conceiving the idea of the study. All authors contributed to the writing and revisions.

## Data availability

The data supporting this article have been included as part of the ESI.† Crystallographic data for **Pr-1** have been deposited at the CCDC under 2373492.†

## Conflicts of interest

There are no conflicts to declare.

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