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Molecular engineering and electrolyte optimization strategies for enhanced performance of Ru(II) polypyridyl-sensitized DSSCs

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Dye-sensitized solar cells (DSSCs) are a leading third-generation solar cell technology due to their low cost, ease of fabrication, and tunable photoelectrochemical properties. Among DSSC components, the photosensitizer plays a crucial role in light absorption and charge generation, with Ru(II)-polypyridyl complexes standing out due to their superior photostability, broad absorption spectra, and efficient charge injection. This review provides a comprehensive analysis of molecular engineering strategies for Ru(II)-polypyridyl photosensitizers, emphasizing ligand modifications to design and develop novel Ru(II) photosensitizers with prolonged excited-state lifetimes, reduced charge recombination, enhanced light-harvesting capabilities, and improved overall solar-to-power conversion efficiency (PCE). In addition, cyclometallated polypyridyl Ru(II) complexes are explored as promising alternatives to Ru(II) complexes incorporating labile thiocyanate (SCN) ligands for DSSCs, which offer improved stability. The relationship between the molecular structure of Ru(II) photosensitizers and their photovoltaic characteristics is analyzed by examining key factors that influence their photovoltaic performance, including light-harvesting efficiency, fine-tuning ground and excited state oxidation potentials (GSOP/ESOP), extending excited state lifetimes, and minimizing charge recombination. Additionally, the impact of co-adsorbents, electrolyte additives, and interfacial engineering on DSSC performance is explored. Emphasis is placed on optimizing redox electrolytes beyond conventional iodide/triiodide (I^-/I_3^-) systems to minimize energy loss and enhance PCE. By carefully considering those challenges, this review lays the groundwork for the rational design of next-generation DSSCs that are more efficient, stable, and commercially viable.

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1. Introduction

The growing global reliance on electronic devices, combined with the rapid depletion of traditional fossil fuel reserves,

emphasizes the importance of developing sustainable energy solutions.^{1–4} As energy demand grows, experts are actively studying renewable energy solutions that provide both efficiency and environmental benefits. Solar energy stands out as

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a particularly promising option, considering that the daily solar radiation reaching the Earth's surface greatly surpasses global energy consumption needs. The amount of solar radiation reaching the Earth's surface greatly exceeds global energy demand. On average, 173 000 terawatts (TW) of solar radiation continuously strikes the Earth,⁵ while global electricity demand is approximately 3.0 TW.⁶ A moderately efficient solar cell array (~10% efficiency) covering a limited portion of the Earth's surface could generate substantial electricity, reducing dependence on fossil fuels and lowering greenhouse gas emissions.⁷ This led to the invention of new photovoltaic (PV) technologies, such as dye-sensitized solar cells (DSSCs), organic photovoltaics (OPVs), quantum dots solar cells (QDSCs), and as a third-generation alternative, perovskite solar cells (PSCs).^{8–11} DSSCs have drawn intensive attention and have the potential to replace silicon-based technology due to their low-cost, lightweight, facile solution processability, and superior photovoltaic performance.¹² Since their inception by O'Regan and Grätzel in 1991,¹³ DSSCs have undergone significant advancements, particularly in their key components including photoanode (TiO₂ or ZnO semiconductor coated FTO), photosensitizers

(dyes), electrolyte (I[−]/I₃[−] redox couple) and counter electrode (Pt coated FTO).^{14,15} Hence, DSSC researchers are still searching for novel and more efficient electrolytes,^{16–18} photosensitizers^{19,20} and/or semiconductors that would enhance the efficiency of DSSCs. The organic photovoltaics (OPVs), especially DSSCs, are a promising alternative to silicon-based solar cells with unique advantages such as low cost, ease of manufacturing, high surface to weight ratio, and flexibility.^{21–24}

Photosensitizers play a pivotal role in DSSCs, directly influencing their efficiency. Scientists have developed numerous highly efficient photosensitizing dyes, achieving a power conversion efficiency (PCE) stabilized at approximately 13% under one sun illumination (AM 1.5G sunlight, 100 mW cm^{−2}).²⁵ Since the discovery of DSSCs by O'Regan and Grätzel in 1991,¹³ several photosensitizers have been used throughout the last 34 years, including both metal-free and metal-based photosensitizers.^{26,27} Among these classes metal-based photosensitizers specially, Ru(II)-polypyridyl complexes have established themselves as the benchmark due to their exceptional photophysical and electrochemical properties. These complexes have broad absorbance spectra that extend into the near-



Fig. 1 The key optimization strategies of the DSSCs components including modifications to electrolytes, Ru(II)-polypyridyl dyes, co-adsorbents, counter electrodes, and additives.



infrared region, high molar extinction coefficients, and long-lived excited states, which all contribute to effective electron injection into the semiconductor.²⁸ Moreover, Ru(II)-based photosensitizers exhibit outstanding photostability and redox reversibility, resulting in longer device lifetimes and improved stability under operational conditions. While metal-free photosensitizers, such as porphyrins, indoline, carbazole, triphenylamine derivatives, thiophene, indoline, BODIPY, and phthalocyanine dyes, have attained competitive PCEs, they frequently suffer from limitations such as dye aggregation, a shorter absorption range, and stability concerns.²⁹ In contrast, Ru(II) complexes, which include highly investigated photosensitizers such as **N3**, **N719**, **C101**, and **C106**, have consistently exhibited superior performance, producing high PCEs while maintaining strong long-term stability. Although Ru-based photosensitizers are more expensive, their effectiveness and endurance make them vital for high-performance DSSCs and a preferred choice for commercial applications.²⁸

Herein, we present a comprehensive analysis of molecular engineering options for improving the overall performance of DSSCs, with a special emphasis on Ru(II)-polypyridyl photosensitizers. It highlights significant changes in ligand design that improve light absorption, excited-state durations, and charge transfer qualities. Beyond Ru(II)-polypyridyl photosensitizer modifications, the review investigates critical DSSC components such as electrolytes, co-adsorbents, and counter electrodes, as shown in Fig. 1. In addition, alternative redox mediators beyond the traditional iodide/triiodide (I^-/I_3^-) system are examined, with a focus on charge recombination and device stability. Similarly, the utilization of diverse co-adsorbent architectures to avoid dye aggregation, improve dye adsorption, and reduce charge recombination is investigated. The study also assesses different counter-electrode materials to replace the standard platinum-based material, providing cost-effective and efficient solutions. By addressing these combined molecular engineering and material optimization

strategies, this work aims to offer valuable insights into the rational design of next-generation DSSCs with superior efficiency, long-term stability, and commercial viability.

2. Composition of DSSCs

A DSSC comprises several key components that work in unison to efficiently convert sunlight into electrical energy. These essential elements include photoanode, a sensitizing dye, a redox electrolyte, and a counter electrode, each playing a vital role in the device's photovoltaic performance. As shown in Fig. 2, the working and counter electrodes are commonly composed of transparent conductive glass covered with fluorine-doped tin oxide (FTO).^{30,31} The photoanode is made up of an FTO substrate coated with a layer of mesoporous titanium dioxide (TiO_2), which acts as an effective surface for dye adsorption. Meanwhile, the counter electrode is often coated with a platinum catalyst. A redox electrolyte, typically comprising the I^-/I_3^- redox pair, is injected between the two electrodes to enhance charge transport.³² When the DSSC is exposed to sunlight, the photosensitizer absorbs photons ($E = h\nu$), resulting in electron excitation and subsequent injection into TiO_2 's conduction band. A schematic diagram showing the main components of DSSCs as well as the flow of electrons inside the cell has been illustrated in Fig. 2.

2.1. Working electrode (photoanode)

The photoanode, serving as the working electrode in a typical dye-sensitized solar cell (DSSC), consists of three main components: a transparent conductive glass (TCG), a mesoporous semiconductor layer, and an adsorbed dye. TCGs are typically composed of thin films of doped metal oxide semiconductors, supported on either glass or flexible substrates. Among these, fluorine-doped tin oxide (FTO) is the most used due to its high thermal stability, excellent electrical conductivity, and optical transparency.^{33,34} More recently, alternative

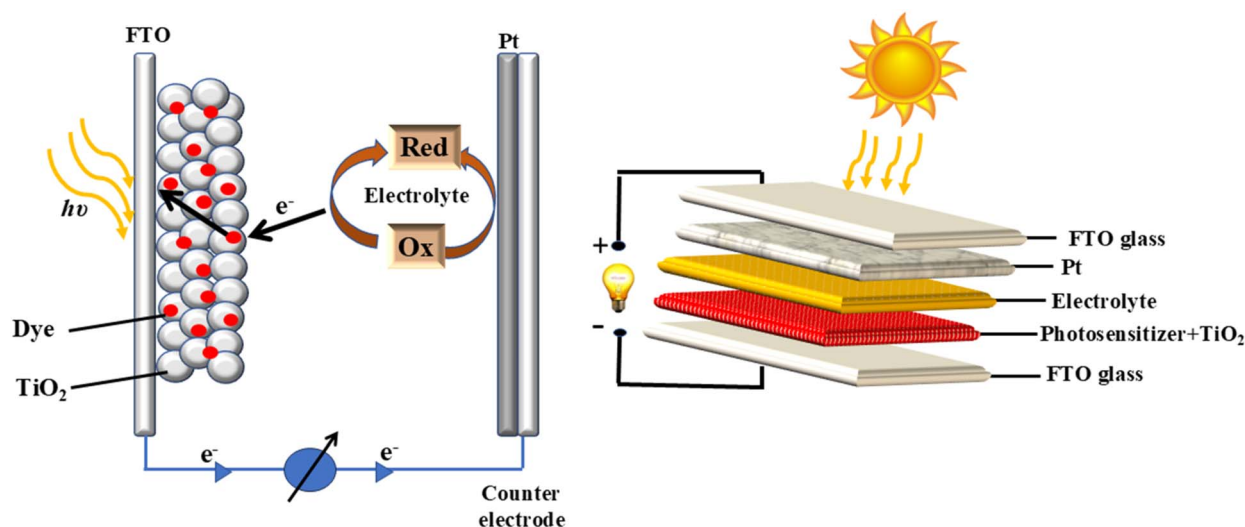


Fig. 2 A schematic diagram showing the main components of DSSCs and the flow of the current inside the cell.

materials such as graphene-based electrodes have been explored for better compatibility with flexible substrates.^{35,36}

The semiconductor layer, typically an n-type metal oxide, is deposited on top of the TCG. The primary role of the semiconductor oxide layer is to provide a strong bond with the photosensitizer and facilitate efficient electron transport. The ideal characteristics of a semiconductor oxide layer for DSSCs include: (1) high surface area for enhanced dye adsorption and efficient light harvesting (2) minimal photon loss and high transparency in the visible spectrum (3) a conduction band (CB) lower than the lowest unoccupied molecular orbitals (LUMOs) of the dye, allowing effective electron injection (4) high electron mobility for efficient charge transport (5) chemical stability against redox electrolytes to reduce electron recombination (6) presence of hydroxyl groups or surface defects to enable strong dye bonding (7) ease of fabrication, stability, affordability, and environmental friendliness.

Titanium dioxide (TiO₂) in its anatase crystalline form remains the most widely used semiconductor due to its superior photovoltaic efficiency, abundance, electron affinity, dye loading, surface area and cost-effective synthesis. Compared with other transition metal oxides, which makes it the most suitable choice as photoanode for DSSCs. This material was first employed in DSSCs by O'Regan and Grätzel (1991),¹³ and despite extensive research on alternatives such as zinc oxide (ZnO),^{37,38} Nb₂O₅,³⁹ SrTiO₃,⁴⁰ Zn₂SnO₄,⁴¹ tin oxide (SnO₂),⁴² and TiO₂ continues to deliver the highest efficiencies.⁴³ Beyond conventional TiO₂, researchers have developed modified semiconductor oxides by doping them with transition metals, alkaline earth metals, rare earth metals, and non-metals to improve performance.^{44,45} Selenium nanoparticle films have also been applied over TiO₂ layers to create cost-effective solar cells.⁴⁶ Despite these advancements, TiO₂ continues to outperform other materials due to its ideal electronic properties for DSSCs.

A variety of TiO₂ nanostructures have been investigated for DSSCs, including one-dimensional (1D) nanostructures such as nanotubes,⁴⁷ nanorods,⁴⁸ nanowires,⁴⁹ and nanosheets.⁵⁰ These 1D nano structures are characterized by their excellent light-scattering ability and efficient electron transport, however their limitation lies in their relatively low surface area, which reduces dye loading.⁵¹ Furthermore, mesoporous TiO₂ is widely used to optimize the balance between dye loading and electron mobility due its small pore sizes and large surface area.⁵² Three common crystalline phases of TiO₂ have been studied: anatase (tetragonal), rutile (tetragonal), and brookite (orthorhombic). Anatase is preferred for DSSC applications due to its superior electron transport properties, while rutile is valued for its stability. Brookite is less common but may offer benefits in charge separation under specific conditions.⁵³ Mixed-phase structures, particularly anatase–rutile combinations, can optimize performance in DSSCs by leveraging the strengths of each phase.⁵⁴

High-efficiency DSSCs typically incorporate two layers of mesoporous TiO₂ deposited on the FTO glass to enhance light absorption and scattering that help to improve overall light harvesting. The first layer, responsible for light absorption, is made of anatase TiO₂ nanoparticles ranging from 15 to 20 nm

in diameter and has a thickness of ~10 μm. The second layer, designed to improve light scattering, is formed of bigger anatase particles (200–400 nm) with a thickness of ~3 μm.⁵⁵ These porous layers not only provide a huge surface area for dye adsorption, but they also promote electrolyte diffusion, resulting in efficient charge transport. To further enhance device performance, a thin TiO₂ coating (~1 nm) is deposited *via* aqueous TiCl₄ treatment that help to facilitate charge injection and enhance the electron lifetime by minimizing surface impurities on TiO₂ and improving dye adsorption.³² This process, often achieved through screen printing or doctor blading,⁵⁶ forms a highly porous layer with an extremely large surface area, approximately 1000 times larger than the electrode's actual area, facilitating sufficient dye-sensitizer loading. The electrode is immersed in the photosensitizer solution to enable covalent bonding between the TiO₂ and the dye through an anchoring group attached to the photosensitizer molecule.

2.2. Counter electrode

The counter electrode (CE) is a crucial component in DSSCs. Its primary role is to facilitate the transfer of electrons from the external circuit to the electrolyte, thereby regenerating the redox couple. For optimal performance, CE should possess excellent catalytic properties for the reduction of triiodide ions, alongside high electrical conductivity. Platinum is commonly used for the construction of CEs, typically deposited on FTO substrate using techniques such as screen printing or doctor blading. However, to lower costs and address the issue of platinum corrosion due to the iodide/triiodide electrolyte, researchers are investigating alternative carbon-based materials. Materials such as graphite, graphene, carbon nanotubes, and carbon black are gaining attention for their low cost, high surface area, superior catalytic activity, excellent electrical conductivity, and resistance to corrosion from the iodide/triiodide electrolyte, making them effective for charge transport.⁵⁷

2.3. Photosensitizers

Photosensitizers are one of the critical components in DSSCs. Typically, a photosensitizer is an organometallic⁵⁸ or organic molecule⁵⁹ that harvests solar photons and converts them into electrical current. Its photochemical, photophysical, and molecular electronic characteristics significantly influence the electron transfer processes and solar spectrum absorption at the semiconductor interface. One of the most widely used organometallic photosensitizers is Ru(II)-polypyridyl complexes.⁶⁰ These consist of a central Ru(II) ion coordinated with ancillary ligands that possess at least one anchoring group and a light-harvesting group. When the photosensitizer is photoexcited, an electron moves from the metal center to the π*-orbital, which is often found on the ligand with the anchoring group. This ensues by electron injection from the excited photosensitizer state into TiO₂'s conduction band (CB) across timescales ranging from femtoseconds to picoseconds.⁶¹ Ru(II)-polypyridyl complexes have demonstrated superior performance in DSSCs in terms of both overall conversion efficiency and stability.^{62,63} DSSCs based on Ru(II)-polypyridyl



photosensitizers have achieved a certified power conversion efficiency (PCE) of 12.3% under standard illumination (AM 1.5G, 100 mW cm⁻²). In lab-scale devices, a record PCE of 14.3% has been reported under similar conditions.⁶⁴ The current commercially relevant photosensitizers, such as **N3**, **N719**, and black dye (**N749**), are prone to degradation over time,⁶³ Fig. 3. This degradation is at least partly due to the loss of monodentate, labile NCS-ligands.⁶⁵

Most Ru(II)-polypyridyl complexes exhibit strong MLCT characteristics within the visible region, typically ranging from 400 nm to 600 nm. These complexes possess thermodynamically favorable ground and excited state potentials, along with extended excited state lifetimes. All these criteria make Ru(II)-polypyridyl complexes versatile and powerful compounds used in a variety of high-tech applications. Their unique properties allow them to play crucial roles in medical treatments,^{66,67} catalysis,⁶⁸ sensing,⁶⁹ electronics,⁷⁰ imaging,⁷¹ and energy conversion.⁶⁵

2.3.1. Requirements of photosensitizers. An ideal photosensitizer for DSSCs should meet several key criteria. It should efficiently harvest light across the entire visible spectrum and extend into the near-infrared region. The presence of anchoring groups *e.g.* (–COOH, –OH, –NO₂, –H₂PO₃, –SO₃H, *etc.*) is essential for strong binding to the semiconductor oxide surface, enabling strong electronic coupling with TiO₂ conduction band. Effectively inject electrons into the semiconductor oxide upon photoexcitation with an excited state energy level surpassing the CB edge of the semiconductor oxide and possess sufficiently excited state lifetime for electron injection.⁷² Additionally, the photosensitizer should possess redox potential higher than that of the electrolyte/hole conductor, with reversible and stable oxidation and reduction processes.⁷³ Finally, we should maintain photo and thermal stability to avoid its degradation and remain inert to prevent side reactions with the electrolyte for at least 20 years.⁶¹ While researchers have developed various photosensitizers over the past three decades, including metal complexes, porphyrins, phthalocyanines, and metal-free organic photosensitizers, none currently fulfill all requirements, facing challenges such as low efficiency, limited extension coefficient, scalability issues, and poor long-term stability.

3. Working principles of DSSCs

A DSSC operates through the following steps as shown in Fig. 4. First, when the photosensitizer absorbs the energy of the incident light, an electron is excited from the GSOP to ESOP, as illustrated in step (1). If the thermodynamic driving force is sufficient, the electron is injected from the ESOP of the photosensitizer into the CB of the TiO₂, as shown in step (2). The injected electron to the CB may move into the external circuit if no recombination happened. The electron travel through the external circuit (*e.g.* light bulb) to the platinum counter electrode and enters the cell *via* reduction of the I₃[–]/I[–] couple at the Pt/electrolyte interface, as shown in step (3). The reduction of the I₃[–]/I[–] couple subsequently helps to regenerate the dye by reducing it (step 4). All the four steps (1–4) are the recommended main steps that should happen in the DSSC to enhance the PCE of the cell. While steps 5–7 represent electron recombination and are not recommended to happen in the cell because it helps to decrease the cell efficiency. One of the

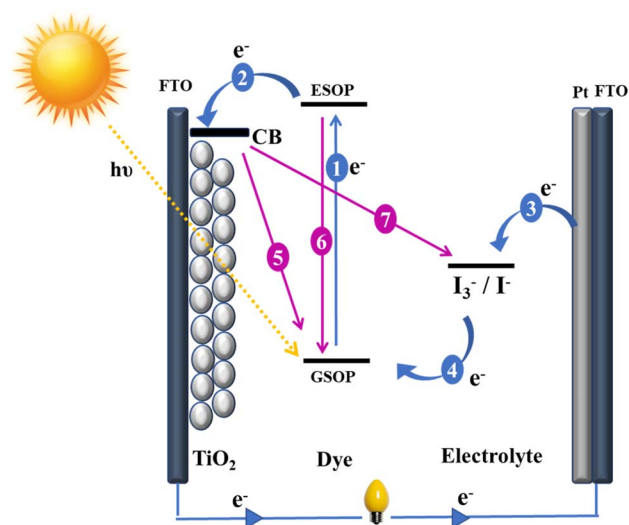


Fig. 4 Working mechanism of DSSCs. The diagram shows the electron flow and energy transfer processes, where blue arrows indicate the desired pathways for efficient energy conversion, and purple arrows highlight charge recombination losses.

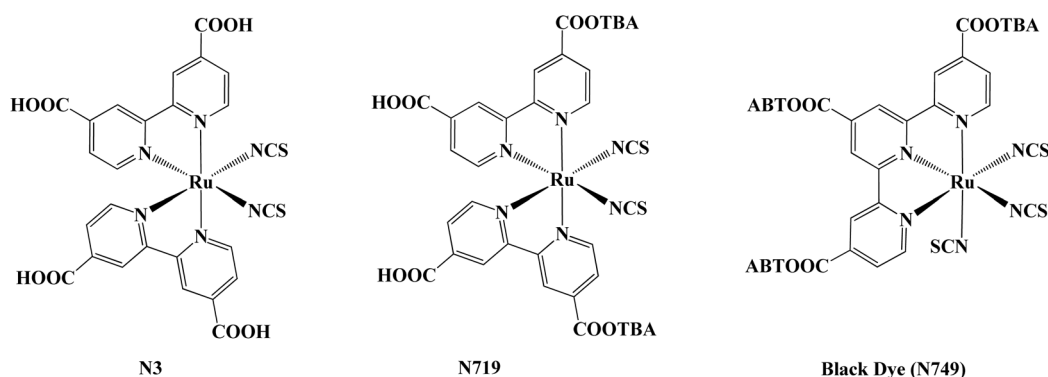


Fig. 3 Chemical structures of benchmark **N3**, **N719**, and black dye (**N749**) photosensitizers.

undesired steps is step 5 that represents recombination of electrons from the CB to the photosensitizer to regenerate the dye ground state. Also, relaxation may happen *via* non-radiative decay processes through electron transfer from the ESOP to the GSOP, as illustrated in step 6. Finally one of the common recombination is the reduction of the redox mediator (I_3^-/I^-) by electrons in the conduction band of TiO_2 ,⁷⁴ as illustrated in step 7.

The solar spectrum at air mass 1.5, which represents hemispherical solar spectral irradiance, is distributed as follows: approximately 5% in the ultraviolet range (300–400 nm), 43% in the visible spectrum (400–700 nm), and 52% in the near-IR region (700–2500 nm).⁷⁵ Consequently, an optimal photosensitizer for single-junction photovoltaic cells should exhibit broad absorption across the entire solar spectrum, particularly between 400 nm and 920 nm, to maximize light-to-electricity conversion. Effective light harvesting in the visible and near-IR regions requires that the photosensitizer's ESOP remains above the CB edge of TiO_2 , enabling efficient electron injection. Additionally, the photosensitizer's GSOP must be sufficiently negative compared to the electrolyte to ensure proper regeneration of the oxidized photosensitizer.⁷⁶

The schematic diagram in Fig. 4 illustrates the thermodynamically favorable interfacial electron transfer process in a DSSC. When the photosensitizer absorbs a photon, it is excited to a higher energy state (step 1). This is followed by electron injection into the conduction band (CB) of TiO_2 (step 2), while the oxidized photosensitizer is simultaneously reduced by the I_3^-/I^- redox couple. An ideal $Ru(II)$ -polypyridyl photosensitizer should have an ESOP higher than that of the TiO_2 -CB to ensure efficient electron injection. Additionally, its GSOP should be lower than the redox potential of the electrolyte to enable rapid regeneration of the photosensitizer by electron donation from I_3^-/I^- . A lower ESOP value for the dye can decrease electron injection efficiency and enhance recombination reactions, which can adversely affect the performance of the DSSC.

4. Electronic transitions in $Ru(II)$ complexes

Ruthenium(II)-polypyridyl complexes, are widely used as photosensitizers in DSSCs due to their strong light absorption, efficient charge transfer, and long excited-state lifetimes. $Ru(II)$ -polypyridine complexes exhibit unique absorption characteristics due to their d^6 electronic configuration and the nature of their polypyridine ligands. The nitrogen atoms in these ligands provide strong σ -donor capabilities, while the aromatic rings contribute π -donor and π^* -acceptor orbitals, leading to complex electronic interactions.⁷⁷ The absorption of light in these complexes results in various electronic transitions, which are crucial for their photovoltaic performance. As illustrated in Fig. 5, the molecular orbital (MO) diagram for a typical octahedral $Ru(II)$ complex consists of σ_L orbitals (strong bonding, ligand-centered), π_L orbitals (bonding, ligand-centered), π_M orbitals with t_{2g} symmetry (nonbonding, metal-centered), σ_M^*

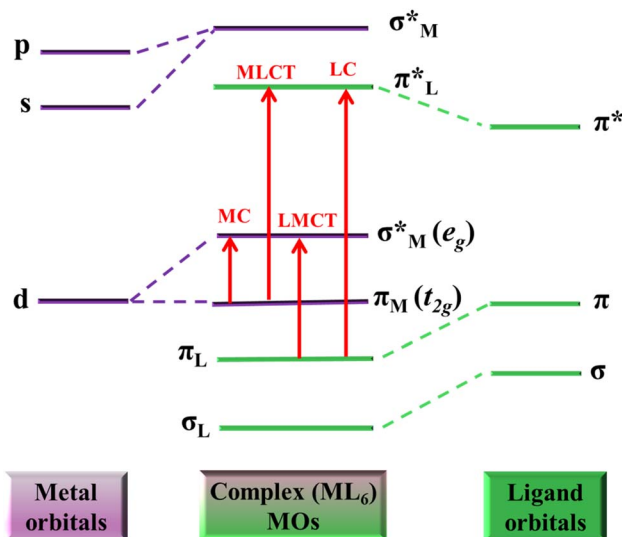


Fig. 5 Molecular orbital diagram for the octahedral metal complex showing the main four types of transitions.

orbitals with e_g symmetry (antibonding, metal-centered), π_L^* orbitals (antibonding, ligand-centered), and strongly antibonding σ_M^* orbitals (metal-centered).

Four primary electronic transitions define the absorption properties of $Ru(II)$ -polypyridyl complexes.⁷⁷ The first transition called metal-centered (MC) which occurs between different d orbitals of the $Ru(II)$, typically transitions from the lower-energy t_{2g} orbitals to the higher-energy e_g orbitals.⁷⁸ The second transition is Metal-to-Ligand Charge Transfer (MLCT) which involve the excitation of an electron from a metal-centered orbital (typically Ru d-orbitals) to a ligand-based π^* orbital (usually polypyridyl ligands like bipyridine or phenanthroline). MLCT transitions occur in the visible to near-IR region (400–700 nm), making them essential for light absorption in DSSCs.^{64,77} The third is Ligand-to-Metal Charge Transfer (LMCT) transitions that occur when an electron is transferred from ligand-based orbitals to the Ru-based orbitals. Typically occurs when the metal is in a high oxidation state and can appear as intense bands in the UV (below 400 nm) but are less common in $Ru(II)$ -polypyridine complexes.^{79,80} The fourth transition is Ligand-Centered (LC), also known as π - π^* transitions which occur within the ligand orbitals. Usually, they appear as intense bands in the UV region.⁸¹ Additionally, an uncommon transition, Ligand-to-Ligand Charge Transfer (LLCT), can occur when electrons are transferred between different ligands within the same complex.⁸²

5. Molecular engineering for the optimization of DSSCs

5.1. Evolution of $Ru(II)$ photosensitizers

In 1993, Nazeeruddin and Grätzel introduced a photosensitizer known as *cis*-dithiocyanato-bis(2,2',4,4'-bipyridyl-4,4'-dicarboxylate) ruthenium(II), commercially named N3. DSSCs of N3 photosensitizer achieved a notable solar-to-electric power conversion



efficiency of 10%.^{83,84} Further, the deprotonation of carboxylic protons in the **N3** photosensitizer to form **N719** resulted in a considerable shift in oxidation and reduction potentials towards more negative values and a significant increase in overall conversion efficiency to be 11.18%.^{83,85}

Studies suggest that the fully protonated **N3** photosensitizer transfers most of its protons to the TiO₂ surface, increasing its positive charge. This enhances electron injection efficiency, resulting in a higher short-circuit photocurrent density (J_{SC}) of 18.20 mA cm⁻², compared to 17.73 mA cm⁻² for **N719**.⁸⁶ However, surface protonation causes a positive shift in the TiO₂ conduction band, leading to a reduction in open circuit photovoltage (V_{OC}). As a result, **N3** exhibits a lower V_{OC} (0.72 V) than the deprotonated **N719** (0.846 V). Extensive research on the impact of protonation on photovoltaic performance has shown that a singly protonated photosensitizer is the most effective in maximizing DSSC efficiency compared to its multiple protonated counterparts.⁶³ Furthermore, in 2001, M. Nazeeruddin introduced a Ru(II)-based **N749** photosensitizer known as (black dye) which achieved solar-to-power conversion efficiency of 10.4%, J_{SC} of 20.53 mA cm⁻², and V_{OC} of 0.72 V under standard one sun illumination.⁸⁷ It was suggested that the presence of three thiocyanato (S=C=N-) and terpyridine tricarboxylic acid ligands caused a significant red shift in the metal-to-ligand charge transfer (MLCT) band by decreasing the π^* energy levels of the terpyridine ligands. This shift enabled the photosensitizer to harvest photons across the entire visible spectrum and extend into the near-infrared region of the solar spectrum, thus creating a panchromatic photosensitizer.⁸⁷

An alternative strategy for improving the light-harvesting capabilities of Ru(II)-polypyridyl photosensitizers involves extending the conjugation of the bipyridine ligand. This modification adjusts the GSOP of the photosensitizer, thereby enhancing its absorption characteristics. In this regard, photosensitizers such as **Z-910** and **K-19** have been designed by substituting the carboxylic anchoring groups on one of the bipyridyl ligands with extended conjugated moieties such as methoxy-3-vinylbenzene and hexoxy-3-vinylbenzene, respectively (Fig. 6). These photosensitizers exhibit broader MLCT bands and significantly higher molar extinction coefficients than the widely used **N719**. The presence of methoxy-3-vinylbenzene and hexoxy-3-vinylbenzene moieties help to increase the conjugation in the complex that significantly enhances their ability to absorb light and causes red shift in their absorption spectrum due to the improved electron-donating ability, ultimately boosting the photovoltaic performance of the cell-sensitized with **Z-910** and **K-19**. The π -conjugated photosensitizers **Z-910** and **K-19** achieved remarkable PCEs (light-to-power conversion efficiencies) of 10.2% and 7.1%,^{88,89} respectively. Furthermore, **K-19** demonstrates exceptional thermal stability, attributed to its long alkyl chain.⁸⁹ In 2006, K. Jiang introduced **HRS-1**, a ruthenium(II) complex featuring a hydrophobic moiety (2-hexyl-5-vinylthiophene) that helps to extend the π -conjugation inside the complex. The absorption spectrum of the photosensitizer (**HRS-1**) demonstrated a notable increase (33%) in the molar extinction coefficient and MLCT red shifted by 10 nm relative to **N719**.⁹⁰ Comparative results under identical conditions demonstrated

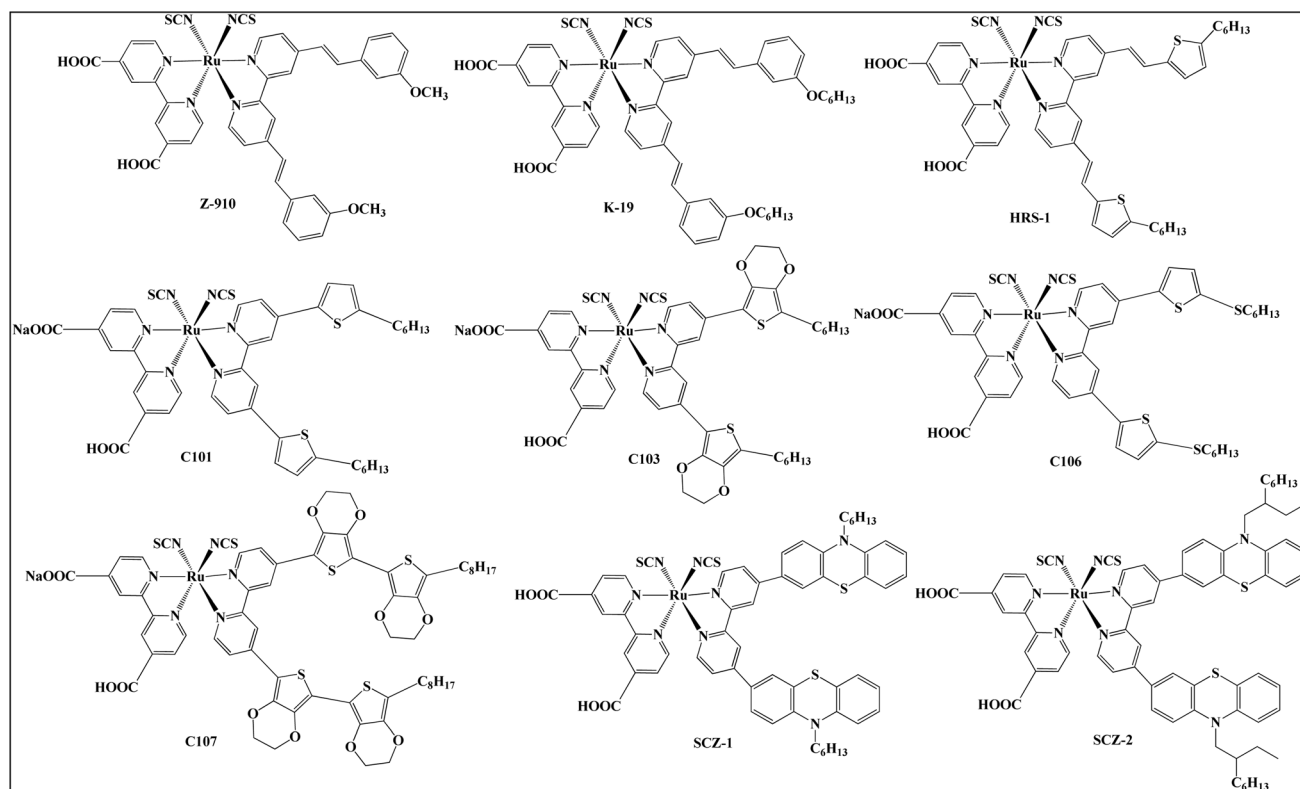


Fig. 6 The molecular structures of the Ru(II) complexes photosensitizers.



that **HRS-1** achieved an overall conversion efficiency of 9.5% compared to 8.9% for **N719**. The high efficiency and stability of **HRS-1** are attributed to the strong electron-donating thienyl groups in the 4,4',9-di(hexylthienylvinyl)-2,2'-bipyridyl ligand.⁹⁰ Furthermore, **HRS-1** had a greater short-circuit current density (J_{SC}) than **N719**-sensitized DSCs. This improvement is most likely attributable in part to the enhanced molecular extinction coefficient and red-shifted absorption. In addition, the photosensitizer **C101**, a homolog of **HRS-1** with the ethylene groups between the EDOT and bipyridyl units deleted (as shown in Fig. 6), displayed a 10.33% increase in solar-to-power conversion efficiency. This enhancement is related to the enhanced molar extinction coefficients produced by extending the π -conjugation of the bipyridyl ligands. This boosts the optical absorptivity of the mesoporous titania film and the charge collection yield in dye-sensitized solar cells (DSSCs).^{91,92} Furthermore, replacing the hexyl chain substituent in **C101** with thiohexyl chains resulted in **C106**, a novel heteroleptic Ru(II)-polypyridyl photosensitizer with a higher power conversion efficiency of 10.57% to 11.4%. The advantage of thioalkyl chains over simple alkyl chains is not yet explained.⁹²

Additionally, Q. Yu *et al.* reported two heteroleptic Ru(II) complexes with electron-rich ligands, **C103** and **C107**. The study found that the ruthenium complex **C107** had a similar absorption profile to its analogue **C103**, but with a red shift and a greater molar extinction coefficient ($27.4 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ at 559 nm) than **C103** ($20.5 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ at 550 nm). As a result, **C107** had a higher PCE of 10.7% compared to 10.4% for **C103**.⁹³ The red shift and enhanced molar extinction coefficient of **C107** are due to its longer and bigger ligand, which increases electron-donating capacity. This powerful electron-donating ligand destabilizes the metal's GSOP by transferring electron density to the metal center. Consequently, the overall gap between the GSOP-metal and the ESOP-ligand decreases, leading to a red shift in the absorption spectrum.⁹¹ Finally, Z. She *et al.* reported two Ru(II) complexes, **SCZ-1** and **SCZ-2**, featuring a phenothiazine-modified bipyridine as an ancillary ligand. These photosensitizers (**SCZ-1** and **SCZ-2**) exhibit enhanced light-harvesting capacity due to higher molar extinction coefficients of $1.77 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$ and $1.66 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$, respectively, compared to $1.27 \times 10^4 \text{ M}^{-1} \text{ cm}^{-1}$

for the reference (**N719**). Under AM 1.5G irradiation, **SCZ-1** and **SCZ-2** DSSCs achieved PCEs of 10.4% and 10.2%, respectively, surpassing **N719**'s efficiency of 9.9%. The higher PCEs of **SCZ-1** and **SCZ-2** are primarily attributed to the increased J_{SC} resulting from the improved absorption coefficient. The photovoltaic characteristics of **SCZ-1** and **SCZ-2** is comparable, suggesting that the difference in alkyl chains (hexyl or 2-ethyloctyl) on the phenothiazine moiety does not significantly impact their light-absorption ability. Both **SCZ-1** and **SCZ-2** photosensitizers showed similar V_{OC} values to that of **N719**, but with higher J_{SC} values. The dye loading amounts on TiO_2 are lower for **SCZ-1** and **SCZ-2** compared to **N719**, yet their higher extinction coefficients and light-harvesting efficiency contribute to the enhanced J_{SC} .⁹⁴ The molecular structures of **Z-910**, **K-19**, **HRS-1**, **C101**, **C106**, **C103**, **C107**, **SCZ-1** and **SCZ-2** are depicted in Fig. 6 as well as its photovoltaic characteristics is tabulated in Table 1.

5.2. Thiocyanate free photosensitizers

Thiocyanate-free Ru(II) complexes for DSSCs are ruthenium-based photosensitizers that do not include thiocyanate (SCN) ligands in their molecular structure as illustrated in Fig. 7. These complexes typically employ other ligands to achieve the desired photophysical and electrochemical properties necessary for efficient light absorption and electron transfer processes in DSSCs.

Wang *et al.* designed and synthesized thiocyanate-free Ru(II) complexes comprises from 4,4'-dicarboxylic acid-2,2'-bipyridine together with two functionalized pyridyl azolate ancillary ligands consisting of pyrazolate groups namely TFRS-4 provides $V_{OC} = 0.75 \text{ V}$, and a high PCE of 10.2%.⁹⁵ Wu *et al.*, modifying TFRS-4 by introducing two *trans*-oriented quinolinyl pyrazolate ancillaries afforded TFRS-52. The introduced *tert*-butyl group having quinoline in 6th position of TFRS-52 achieved improved $V_{OC} = 860 \text{ mV}$, and power conversion efficiency of 10.1%. The improved high open-circuit voltage is based on the increase of *tert*-butyl group the upward shift in conduction band edge and higher recombination resistances were noticed.⁹⁶

Chou *et al.* designed and synthesized a heteroleptic tridentate ancillary ligand to replace traditional thiocyanates (SCN). This ancillary ligand is used to synthesize thiocyanate free Ru(II) complexes to replace SCN-based photosensitizers

Table 1 Photovoltaic properties (J_{SC} , V_{OC} , FF and PCE) of Ru(II) based-dyes for n-type DSSCs

No.	Photosensitizer code	J_{SC} (mA cm^{-2})	V_{OC} (V)	FF (%)	PCE (%)	Reference
1	N3	18.20	0.720	73.00	10.00	84
2	N719	17.73	0.846	72.00	11.18	85
3	N749	20.53	0.720	70.40	10.40	87
4	Z-910	7.20	0.777	76.40	10.20	88
5	K-19	13.20	0.718	74.60	7.10	89
6	HRS-1	20.00	0.680	69.00	9.50	90
7	C101	17.75	0.749	77.70	10.33	92
8	C103	18.35	0.760	74.80	10.40	93
9	C106	18.28	0.749	77.20	10.57	92
10	C107	19.18	0.739	75.10	10.70	93
11	SCZ-1	19.85	0.761	68.80	10.40	94
12	SCZ-2	19.88	0.761	67.70	10.20	94



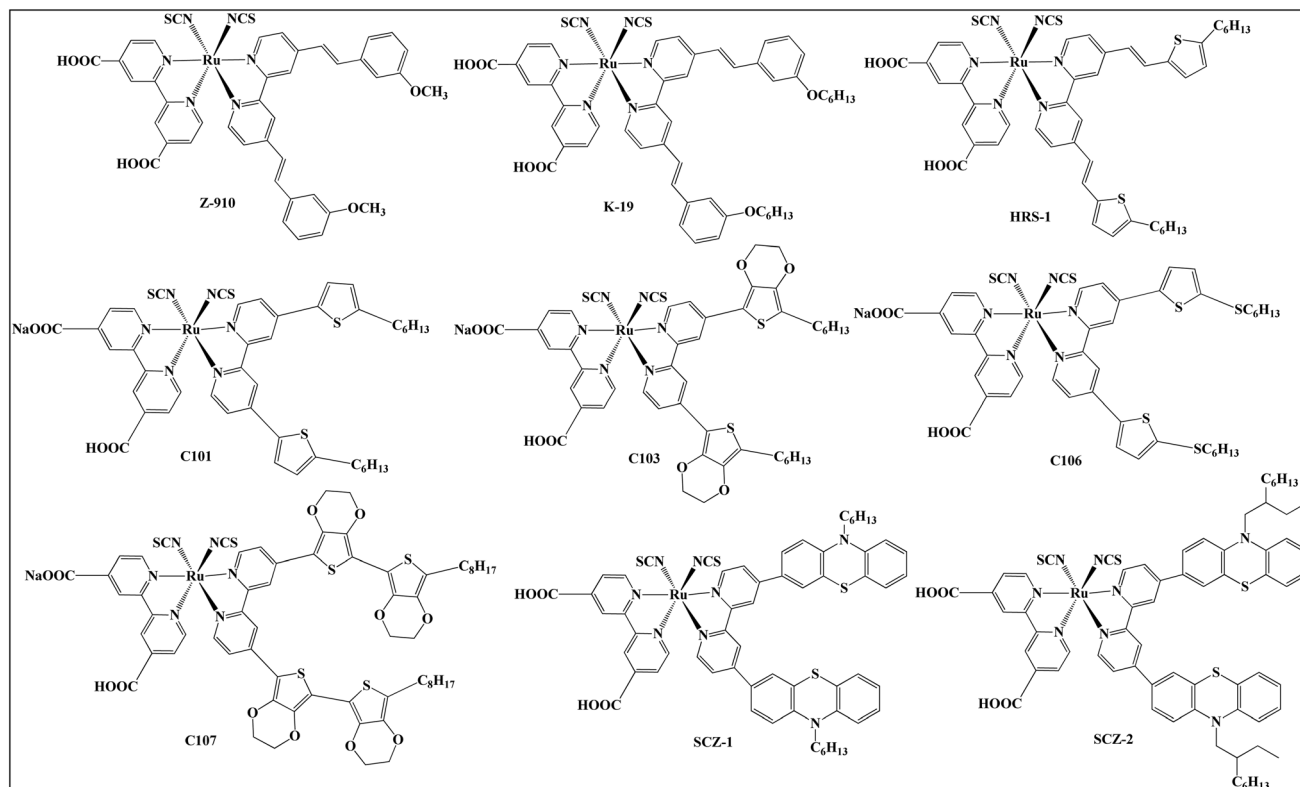


Fig. 7 Thiocyanates free Ru(II) polypyridyl heteroleptic complexes.

such as **N3** and **N719**, which have been widely used in DSSCs. This approach helps to enhance light-harvesting capabilities and stability by replacing (SCN) ligands, which can lead to isomerization issues and reduce the stability of the solar cells. The synthesized TF complexes (TF1-4) exhibit improved photophysical and electrochemical properties compared to **N749**. They show intense visible absorption bands around 510 nm and broad absorption in the longer-wavelength region at approximately 650 nm and 720 nm, attributed to MLCT and LLCT transitions.⁹⁷ The TF photosensitizers demonstrate higher PCEs compared to **N749**, with the highest PCE of 10.7% for TF-3. The V_{OC} of TF photosensitizers range between 0.76–0.79 V, which is significantly higher than that of **N749** (SCN-based) that achieved V_{OC} of 0.72 V. The improved performance of TF photosensitizers is attributed to better packing on the TiO_2 surface, which reduces charge recombination and enhances V_{OC} . Additionally, the SCN-free design improves the stability of the solar cells, as evidenced by long-term aging experiments.⁹⁷ Additionally, the negative pole of the dipole moment in complexes (TF1-4) is expected to be located closer to the TiO_2 surface, resulting in the uplift of the TiO_2 conduction band level and consequently enhanced V_{OC} . Overall, the TF photosensitizers offer a promising alternative to traditional ruthenium-based photosensitizers for DSSCs.

H. Cheema *et al.* reported thiocyanate-free ruthenium dye (HD-11) featuring a mono-dentate ligand of 3-(trifluoromethyl)pyrazole to replace the labile isothiocyanate ligand. HD-11 revealed 50 nm red shift in the low energy metal-to-ligand charge transfer (MLCT) absorption peak compared to **N719**. However, despite

this spectral shift, HD-11 demonstrated a lower PCE (5.2%) than **N719** under similar experimental conditions.⁹⁸

S. Ashraf *et al.* synthesized thiocyanate-free Ru(II) photosensitizers, SD-15 and SD-16, utilizing 4,4',5,5'-tetramethyl-1*H*,1'*H*-2,2'-bis-imidazole as an ancillary ligand for chelation. These complexes exhibited a red shift in the low-energy metal-to-ligand charge transfer (MLCT) absorption peak. However, under identical experimental conditions, SD-15 and SD-16 exhibited lower power conversion efficiencies (PCEs) of 3.32% and 1.06%, respectively. This decline in performance was attributed to a significantly reduced electron lifetime at the dye/ TiO_2 interface (0.71 ms), compared to 8.8 ms for **N719**, as determined by impedance measurements.⁹⁹

On the other hand, cyclometallating ligands are commonly used in Ru(II) complexes for DSSCs applications. In 2009, Bescho, Grätzel, and colleagues introduced an efficient Ru(II) dye (**1**) for DSSCs featuring a cyclometallated 2-phenylpyridine ligand, with power conversion efficiencies comparable to **N719**.¹⁰⁰ The complex (**1**) shows a red-shifted spectral response and higher external quantum efficiency, reaching over 80% at 600 nm and extending to 800 nm. This shift and enhanced absorption are due to the cyclometallated ligand, which destabilizes the HOMO more than the LUMO, leading to a conversion efficiency of 10.1% under standard sunlight. Moreover, Berlanguette reported a series of Ru photosensitizers with a substituted 3-(2'-pyridyl)-1,8-naphthalimide cyclometallated ligand, achieving high power conversion efficiencies (PCEs) in DSSCs. Substituting the pyridine ring with conjugated groups enhances molar absorption extinction coefficients, while the



Table 2 Photovoltaic properties of the thiocyanate free Ru(II) complexes

No.	Photosensitizer code	J_{sc} (mA cm ⁻²)	V_{oc} (V)	FF (%)	PCE (%)	Reference
1	TFRS-4	18.7	0.750	72.9	10.2	95
2	TFRS-52	16.3–16.8	0.832–0.860	72–78	10.1–10.88	96
3	TF1	18.22	0.740	0.676	9.11	97
4	TF2	20.00	0.790	0.665	10.50	97
5	TF3	21.39	0.760	0.660	10.70	97
6	TF4	20.27	0.770	0.675	10.50	97
7	HD-11	12.89	0.57	0.71	5.2	98
8	SD-15	10.20	0.58	0.56	3.32	99
9	SD-16	3.52	0.52	0.58	1.06	99
10	1	17.00	0.800	0.74	10.10	100
11	2	10.1	0.68	0.67	7.00	101
12	3	10.3	0.66	0.68	7.30	101

1,8-naphthalimide fragment reduces electron density on the metal, keeping the Ru(III)/Ru(II) redox couple more positive than 0.8 V *versus* NHE. This maintains effective dye regeneration by the iodide-based redox mediator. The dye platform allows for modifications that improve light absorption and reduce recombination, achieving PCEs of 7% for complex (2) and 7.3% for complex (3).¹⁰¹ The molecular structures of all thiocyanate free Ru(II) complexes are illustrated in Fig. 7 and Table 2.

5.3. Excited state lifetime optimization

An efficient photosensitizer must have suitable excited state properties, as an electron injection from the photosensitizer's excited state to the semiconductor oxide is essential for the dye-sensitization process. The photosensitizer's excited state should last long enough to allow for proper energy transfer and electron injection. To increase the excited state lifetime of Ru(II)-polypyridyl complexes, specific ligand structural changes are necessary, such as adding electron-donating or withdrawing groups, extending π -conjugation, and utilizing stiff ligands. These changes improve the electronic properties of the complex, stabilize the excited state, and reduce non-radiative decay processes.¹⁰² To achieve efficient electron injections, the excited state lifetime should be long enough to allow for considerable electron transfer processes before the photosensitizer returns to the ground state. The photosensitizer's molecular structure affects its lifetime, which can be increased by adding suitable ligands.¹⁰³ One effective strategy for increasing the excited state lifetime of Ru(II) complexes, particularly the terpyridine photosensitizers [Ru(tpy)₂]²⁺, is to modify the ligands structure. Attaching strong electron-drawing substituents such as SO₂Me to the 4'-position of terpyridine ligands can boost room temperature excited state durations to 25 ns, as illustrated in complex (4) in Fig. 8. Also, attaching electron-donating groups can destabilize metal-based GSOPs, leading to non-radiative decay, but they are less effective than electron-withdrawing groups.¹⁰²

A key strategy for enhancing the excited-state lifetime in tridentate-based Ru(II) photosensitizers involves ligand modifications to reduce non-radiative decay pathways.^{104,105} One effective approach is incorporating tridentate ligands with additional nitrogen atoms, which help stabilize the acceptor

orbitals of the metal-to-ligand charge transfer (MLCT) excited state. For instance, complexes featuring the 2-aryl-4,6-bis(2-pyridyl)-s-triazine have demonstrated longer excited-state lifetimes at room temperature compared to conventional terpyridine-based systems. This improvement is attributed to the incorporation of triazine rings and the near-planar conformation of the non-coordinating ring, which lowers the MLCT state's energy and create a substantial energy gap to the metal-centered (MC) state, thereby suppressing non-radiative decay.¹⁰⁵ Another promising strategy involves the design of dinuclear complexes incorporating planar bridging ligands. These structures enable extensive electron delocalization within the acceptor ligand upon MLCT excitation, reducing Franck-Condon factors associated with non-radiative decay and significantly prolonging the excited-state lifetime. For example, photosensitizer (5) is a binuclear Ru(II) complex featuring a large, planar bistridentate bridging ligand exhibits exceptionally long-lived emission and a high quantum yield in the near-infrared region, positioning it among the most stable near-IR-emitting Ru(II) complexes.¹⁰⁴ The extension of excited-state lifetimes in these complexes is further influenced by electronic interactions between the metal centers and ligands. In dinuclear systems, strong electronic coupling between the two metal centers enhances the stability of the monooxidized species, contributing to prolonged emission lifetimes. A notable example is a tridentate Ru(II) complex incorporating N-heterocyclic carbene (NHC) ligands, designated as (6), which has demonstrated promising photophysical properties which is attributed to the electron-rich nature of NHC ligands.¹⁰⁶ The chemical structures of complexes (5 and 6) are illustrated in Fig. 8.

Another approach is to use strong σ -donating ligands, such as carbenes, which are effective electron donors and form strong bonds with the ruthenium metal center. This destabilizes the lowest unoccupied molecular orbital (LUMO) and increases the energy of metal-centered (MC) states. Duati and colleagues reported heteroleptic photosensitizers (7–9) incorporating mixed ligands, with one being 2,2';6',2''-terpyridine (tpy) and the other being one of the following: 2,6-bis([1,2,4]triazol-3-yl)pyridine or 2,6-bis(5-phenyl-[1,2,4]triazol-3-yl)pyridine, or 2,6-bis([1,2,3,4]tetrazol-5-yl)pyridine. These



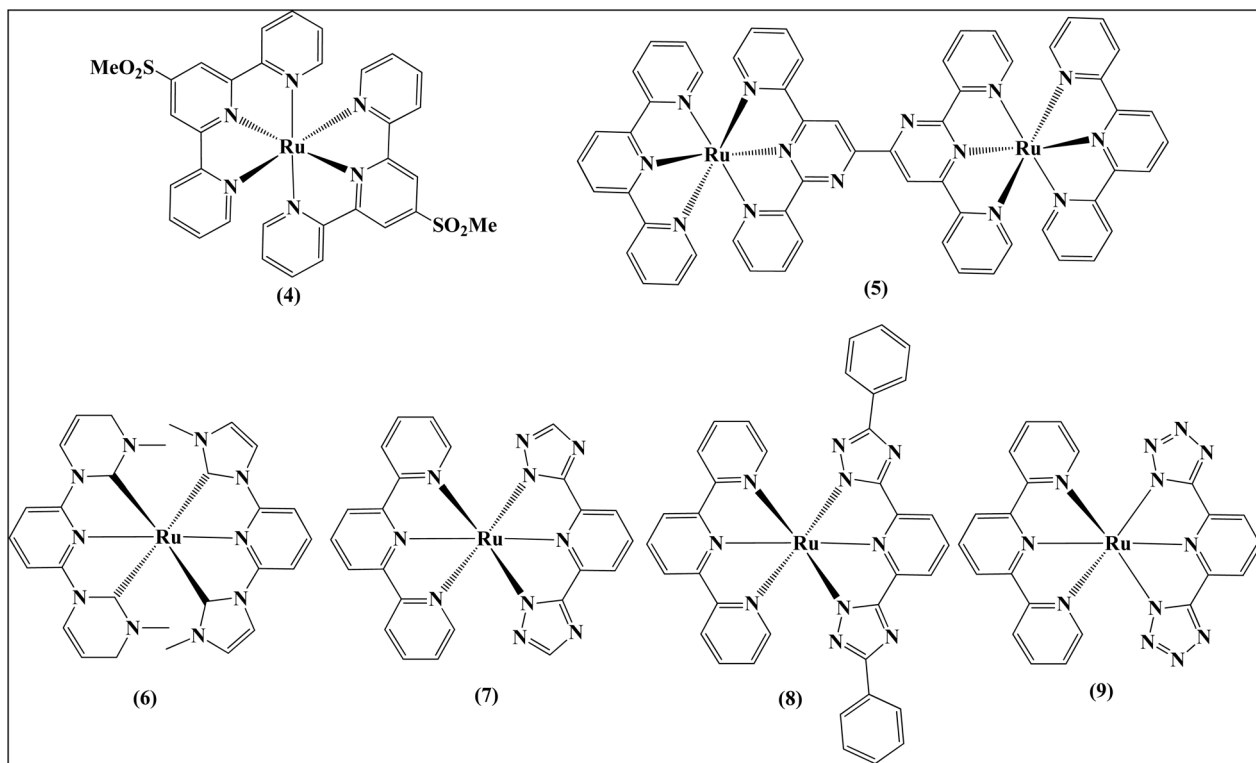
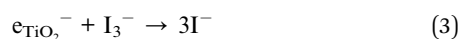
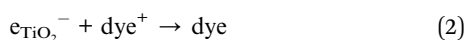
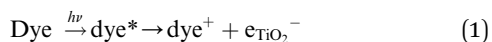


Fig. 8 Tridentate heteroleptic Ru(III) complexes (4–9).

ligands are powerful σ -donors, as illustrated in Fig. 8, and they destabilize the ground state, lowering the energy gap between the metal LUMO and ligand HOMO. Consequently, the photosensitizer emits at a lower energy around 700 nm, with a prolonged excited state lifetime of 77 ns. However, protonation of these triazole rings quenches the excited state by decreasing the electron-donating ability, thereby reversing the process.¹⁰⁷ Cyclometalated ligands, being strong σ -donors, destabilize the HOMO of the metal, reducing the gap between the metal HOMO and ligand LUMO, which significantly decreases the excited state lifetime.^{108,109}

5.4. Strategies to minimize charge recombination reactions

Electron recombination at the semiconductor/dye/electrolyte interface is harmful to the photovoltages in DSSCs, leading to a significant reduction in the device's solar-to-power conversion efficiency.¹¹⁰ Electron recombination occurs when electrons deviate from the normal electron injection process (eqn (1)) and can recombine with oxidized dye molecules (eqn (2)) or with the electrolyte (I_3^-) (eqn (3)).



Therefore, minimizing charge recombination between TiO_2 and oxidized photosensitizer molecules and/or redox couples in

DSSCs is crucial for maximizing device solar-to-power conversion efficiency. Fig. 4 provides an overview of the kinetics, including electron injections, recombination reactions, and dye regeneration for each process.

In traditional DSSCs, the rate of electron recombination is significantly slower than the electron injection process from the excited photosensitizer into the CB of the semiconductor (TiO_2). However, due to the relatively sluggish diffusion through the nanoporous TiO_2 , electrons tend to accumulate in the semiconductor/electrolyte interface. This proximity increases the likelihood of recombination with oxidized dye molecules or redox species within the electrolyte.¹¹¹ Dye regeneration efficiency is primarily governed by two key factors. The first is the molecular structure of the photosensitizer, which plays a crucial role in determining its ground-state oxidation potential (GSOP), directly impacting the regeneration process. The second factor is the redox potential of the electrolyte, which must be appropriately aligned to facilitate efficient electron donation for regenerating the oxidized dye. To sustain a strong photocurrent, the regeneration process must outpace recombination reactions.¹¹² Therefore, minimizing charge recombination between TiO_2 and oxidized photosensitizer molecules and/or redox couples in DSSCs is crucial for maximizing device solar-to-power conversion efficiency. Fig. 4 provides an overview of the kinetics, including electron injections, recombination reactions, and dye regeneration for each process.

The effect of photosensitizer molecular structure on recombination reactions was studied. Research indicates that adding electron-donating groups to photosensitizers and separating



dye cation centers from the TiO_2 surface will greatly reduce charge recombination,¹¹³ resulting in long-lasting charge separation at the dye/ TiO_2 contact.¹¹⁴ Haque *et al.* confirm this by using transient absorption spectroscopy to investigate the charge-recombination dynamics of three Ru-based multifunctional photosensitizers (**10–12**) with varying triphenylamine

antennas (Fig. 9). Photosensitizers (**10** and **11**) comprise electron-donating units with low molecular weight: triphenylamine (TPA) and *N,N'*-diphenyl-*N,N'*-bis(3-methylphenyl)-[1,1'-biphenyl]-4,4'-diamine (TPD), respectively. In contrast, photosensitizer (**12**) consists of poly(vinyl triphenylamine) chains (polyTPA) with around 100 repeat units linked to one of the

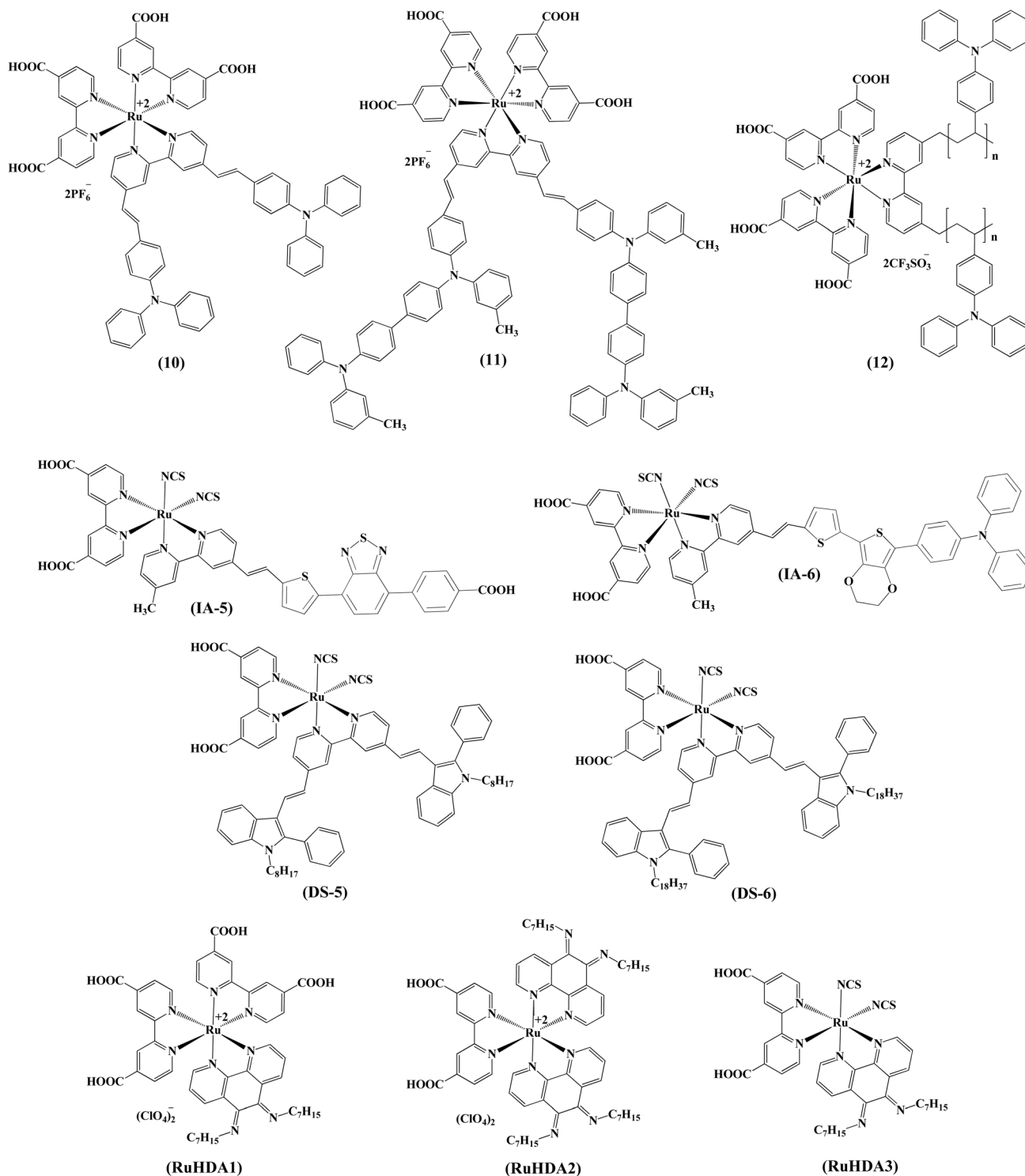


Fig. 9 Ru(II) bipyridyl photosensitizer (**10–12**) incorporating different triphenylamine antenna and phenanthroline-based Ru(II) complexes (**RuHDA1–3**).



Ru(II) core's bipyridine units. The recombination kinetics of these complexes vary significantly depending on the physical separation between the dye cation core and the TiO₂ semiconductor oxide surface. The photosensitizer with the larger antenna had a substantially longer recombination half-time. Complexes **10**, **11**, and **12** exhibited recombination half-times of 350 μ s, 5 ms, and 4 s, respectively. This shows that the spatial separation of the dye cation from the TiO₂ surface, accomplished through the multistep migration of the dye moiety away from the TiO₂ interface, is the key factor controlling charge transfer at the dye/TiO₂ interface.¹¹⁴

Another strategy is to use monosubstituted ancillary ligands, which have been shown to effectively suppress recombination processes. Abdellah *et al.* synthesized two high molar extinction coefficient monosubstituted-bipy Ru(II) complexes, **IA-5** and **IA-6**,¹¹⁵ as illustrated in Fig. 9. These complexes use an electron acceptor ancillary ligand in **IA-5** and an electron donor ancillary ligand in **IA-6** to study how these ligands and the quantity of anchoring groups (COOH) affect the photovoltaic characteristics of DSSCs. The results showed that the highly conjugated and strong donor auxiliary ligand in **IA-6** reduced charge recombination and increased overall power conversion efficiency (PCE) to 7.81% at a V_{OC} of 0.69 V. In contrast, the electron-accepting auxiliary ligand in **IA-5** produced a lower PCE of 6.20% and a V_{OC} of 0.62 V, outperforming the reference **N719**.¹¹⁵

Another way is to incorporate long aliphatic chains and bulky groups into the photosensitizer's auxiliary ligand to reduce recombination losses. Ashraf *et al.* created two heteroleptic polypyridyl Ru(II) complexes, **SD-5** and **SD-6**, with electron-donating *N*-alkyl-2-phenylindole moieties in the auxiliary ligand, as illustrated in Fig. 9. In DSSCs, **SD-6** had an overall efficiency of 8.14% with a V_{OC} of 0.675 V, beating **SD-5**, which had a PCE of 4.99% and a V_{OC} of 0.59 V. **SD-6** outperforms **SD-5** because of its long alkyl chains ($-C_{18}H_{37}$), which limit dye aggregation and suppress charge recombination, resulting in greater PCE and V_{OC} values.¹¹⁶ The addition of indole moieties with extended alkyl chains to **SD-6** increased performance compared to **N719**.¹¹⁶

On the other hand, Pashaei *et al.* conducted a similar investigation, synthesizing three phenanthroline-based Ru(II) complexes (**RuHDA1-3**) with long-chain auxiliary ligands and varied types and numbers of ancillary and anchoring ligands as illustrated in Fig. 9.¹¹⁷ The study revealed that **RuHDA3**, which has one auxiliary ligand and one anchoring ligand with two carboxylic acid groups and two NCS groups, is the most effective photosensitizer. **RuHDA3** achieved a V_{OC} of 0.65 V, compared to 0.58 V for **RuHDA1** and 0.55 V for **RuHDA2**, and a PCE of 6.11%, vs. 4.41% and 3.40% for **RuHDA1** and **RuHDA2**, respectively. Time-Resolved Absorption Spectroscopy (TAS) was utilized to investigate electron transfer kinetics in DSSCs and to link molecular structure to regeneration and recombination lifetimes. The findings demonstrated that balancing auxiliary and anchoring ligands is critical for reducing recombination losses and increasing DSSC efficiency.¹¹⁷ The final strategy involves treating photosensitizer-coated TiO₂ electrodes with pyridine derivatives, providing a simple and effective method for

increasing DSSC performance. This technique lowers recombination rates and raises open-circuit voltage, resulting in a considerable boost in overall DSSC efficiency. Huang *et al.* found that treating **N3** photosensitizer-coated TiO₂ electrodes with pyridine derivatives like 4-*tert*-butylpyridine, 2-vinylpyridine, and poly(2-vinylpyridine) considerably improved photovoltaic performance. The V_{OC} exhibited an approximate 28% increase, reaching 0.73 V, while the PCE improved by 29.3%, achieving 7.5% compared to untreated electrodes. This enhancement in photovoltaic performance is ascribed to the pyridine derivatives, which may reduce the recombination rate.¹¹⁸

6. The role of co-adsorbents

The utilization of co-adsorbents plays a crucial role in enhancing the performance of DSSCs by mitigating recombination processes, preventing aggregation of the photosensitizer, and forming a dense protective layer on TiO₂.¹¹⁹ The co-adsorbent enhances device efficiency in two ways. First, due to its small size, it acts as a filler, occupying the interspaces between the photosensitizer molecules to form a compact layer on the TiO₂ surface. Second, it reduces aggregation among photosensitizer molecules, which can lower electron injection efficiency and increase electron recombination reactions at the TiO₂/electrolyte interface. Co-adsorbents used with the photosensitizer significantly improve the device's open-circuit voltage (V_{OC}).¹²⁰ The selection of appropriate co-adsorbents is critical for improving the efficiency of DSSCs. An ideal co-adsorbent should possess a molecular structure that prevents competitive adsorption between dye molecules and minimizes dye aggregation on the TiO₂ surface. Compounds such as chenodeoxycholic acid (CDCA) and hexadecylmalonic acid (HDMA) have demonstrated effectiveness in achieving uniform dye distribution, thereby maximizing light absorption. Additionally, co-adsorbents should form a compact monolayer on the TiO₂ surface to mitigate electron recombination with I^-/I_3^- .¹²¹ Anchoring groups, including phosphonate and carboxylate functional groups, enhance adsorption stability by strongly binding to the TiO₂ surface.¹²² Furthermore, co-adsorbents with hydrophobic chains can reduce dye aggregation through repellent interactions, contributing to improved efficiency.¹²³ Optimizing the concentration of co-adsorbents is essential to achieving optimal DSSC performance.¹²⁴ Although a high molar extinction coefficient is not a primary requirement, co-adsorbents indirectly enhance light harvesting by facilitating optimal dye distribution. By fine-tuning these properties, co-adsorbents play a key role in maximizing the overall efficiency of DSSCs.

Various co-adsorbents have been utilized to enhance the performance of DSSCs. Lim *et al.* demonstrated that incorporating stearic acid as a co-adsorbent with the **N719** photosensitizer markedly enhances photovoltaic performance. Solar devices showed a notable 25% increase in short circuit current (J_{SC}) and solar-to-power conversion efficiencies compared to control devices. This improvement is attributed to stearic acid's low dipole moment and high solubility, which slows down the

photosensitizer anchoring rate during competitive adsorption. Consequently, a dense and strongly bonded dye layer form on the TiO_2 surface, reducing photosensitizer aggregation and leading to enhanced device performance.¹²⁵ Song *et al.* introduced triaryl amine-based co-adsorbent, 4-(bis(9,9-dimethyl-9H-fluoren-2-yl)amino)benzoic acid (HC-acid), designed to enhance the efficiency of DSSCs. HC-acid serves as an alternative to deoxycholic acid (DCA), offering dual functions: preventing π - π stacking of organic dye molecules and enhancing light harvesting at shorter wavelengths. When used with the organic dye NKX2677, HC-acid significantly improves the solar cell's efficiency, achieving a PCE of 9.09% under AM 1.5G conditions. This represents a 38% increase in efficiency compared to NKX2677 alone, attributed to a 20% increase in J_{SC} and an 11% increase in V_{OC} .¹²⁶ Han *et al.* introduced innovative donor-acceptor type co-adsorbents (Y1 and Y2). These co-adsorbents consist of three units: an electron-donating group, a π -spacer,

and an electron-accepting group. They effectively address challenges such as competitive light absorption by I^-/I_3^- , prevent dye aggregation, and minimize charge recombination. The co-adsorbent Y1, with strong absorption around 390 nm, effectively restored the dip in the IPCE spectrum caused by I^-/I_3^- in black dye-sensitized solar cell. In addition, the use of Y1 as co-adsorbent led to an increase of approximately 20 mV in open-circuit voltage (V_{OC}) and an enhancement in short-circuit current density (J_{SC}), resulting in a power conversion efficiency (PCE) of 11.28%, compared to 10.70% in the absence of the co-adsorbent. Meanwhile, Y2 provided a comparatively lower performance boost. The butyloxyl chains in Y1 helped prevent dye aggregation and reduce recombination, contributing to improved performance.¹²⁷ Finally, Wang *et al.* discovered that incorporating 1-decylphosphonic acid as a co-adsorbent alongside the Ru(II)-bipyridyl complex (Z907) significantly enhanced the stability, short-circuit photocurrent density (J_{SC}),

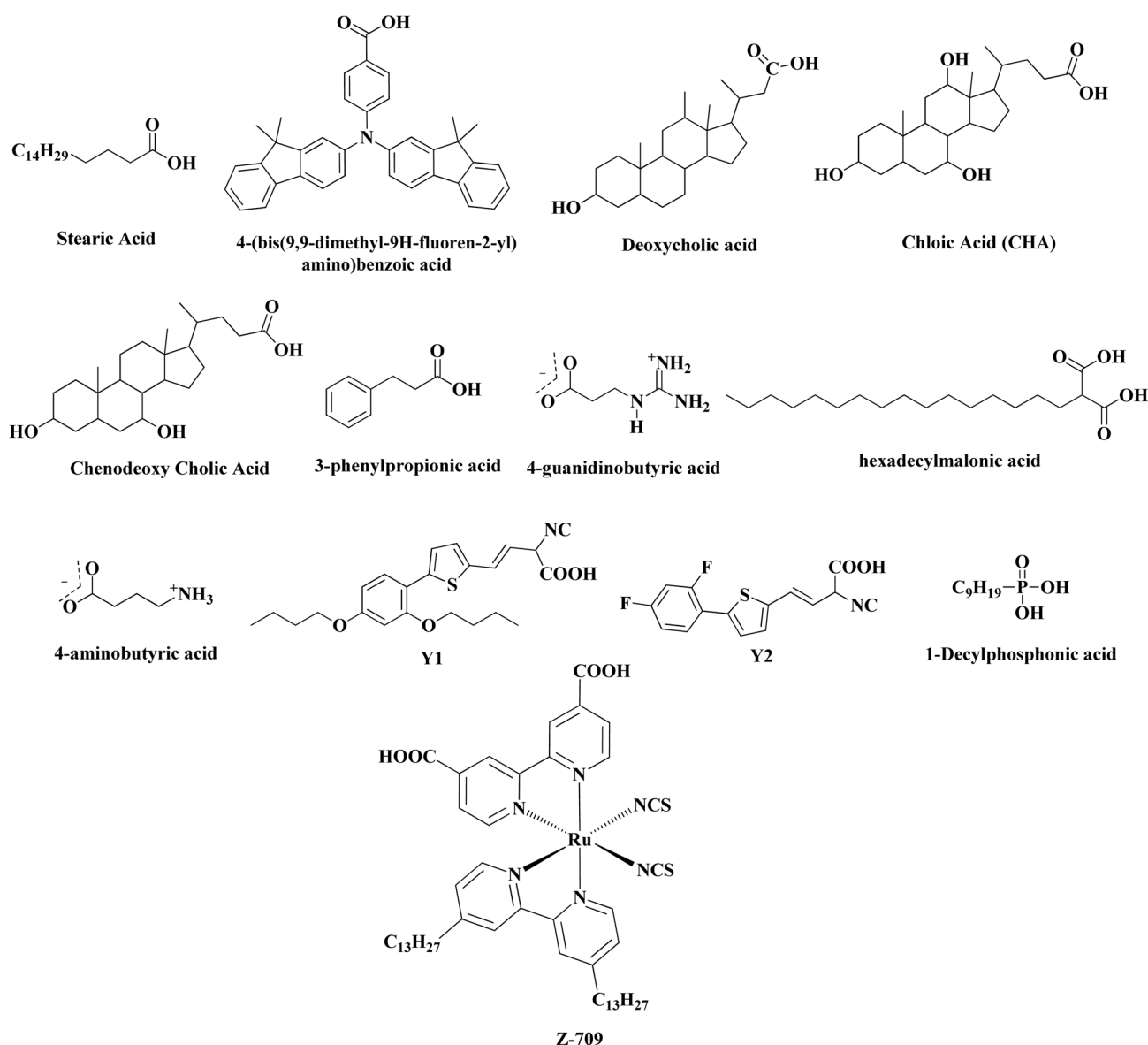


Fig. 10 The molecular structures of the well-known co-adsorbents used in DSSCs, and Ru(II)-bipyridyl complex (Z907).

and overall efficiency of the device. DPA helps to improve conversion efficiency from 6.8% without DPA to 7.3% as well as helps to improve the cells stability under thermal stress because DPA strongly binds to the oxide surface through P–O–metal bonds, reducing hydrophilic sites available for water adsorption.¹²⁸ The chemical structure of the Ru(II)-bipyridyl complex (Z907) is depicted in Fig. 10.

The molecular structures of some co-adsorbents employed in DSSCs to enhance their performance such as stearic acid, 4-(bis(9,9-dimethyl-9H-fluoren-2-yl)amino)benzoic acid, **Y1**, **Y2**, 1-decylphosphonic acid, deoxycholic acid,¹²⁹ chenodeoxycholic acid,¹³⁰ 3-phenylpropionic acid,¹³¹ 4-guanidinobutyric acid,¹³² hexadecylmalonic acid,¹³³ and 4-aminobutyric acid¹³⁴ are listed in Fig. 10. In conclusion, although co-adsorbents play a beneficial role in DSSCs, their improper application can disrupt protective layers and enhance recombination processes, ultimately compromising the performance of the solar cell, especially in the absence of dye aggregation.¹³⁵

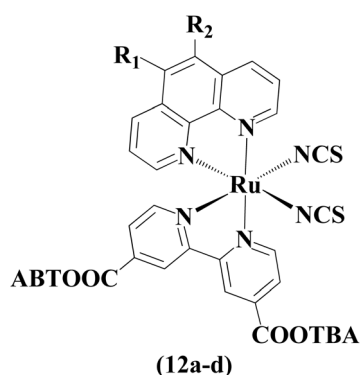


Fig. 11 Molecular structure of Ru(II)-polypyridyl photosensitizers (**12a–d**). Where: $R_1 = H$, $R_2 = H$ (**12a**); $R_1 = CH_3$, $R_2 = CH_3$ (**12b**); $R_1 = H$, $R_2 = NH_2$ (**12c**); $R_1 = H$, $R_2 = NO_2$ (**12d**).

7. Redox electrolyte function in DSSCs

Redox electrolytes are crucial components of DSSCs, significantly impacting both the device's performance and long-term stability. In DSSCs, the redox couple acts as a redox mediator in liquid or gel electrolytes, while in solid electrolytes, the hole transport material (HTM) serves as the hole mediator.¹³⁶ The role of electrolytes becomes crucial following photon absorption, as the dye, upon excitement, swiftly injects electrons into the CB of TiO_2 , leading to its oxidation. The oxidized dye is then regenerated by receiving electrons from the reduced species in the electrolyte. The oxidized species of the redox mediator migrate towards the counter electrode, whereas the reduced species migrate from the counter electrode to the oxidized dye, primarily by diffusion. In DSSCs, the electrolyte plays a dual role: regenerating the oxidized dye and replenishing the reduced species of the redox couple at the CE.

7.1. Electrolyte (I^-/I_3^-) drawbacks

The interaction between photosensitizers and electrolytes can enhance charge recombination reactions, thereby diminishing the performance of DSSCs. Reynal *et al.* investigated a series of heteroleptic Ru(II)-polypyridyl complexes (**12a–d**) with various substituents, including $-H$, $-CH_3$, $-NH_2$, and $-NO_2$, as illustrated in Fig. 11. These photosensitizers were found to significantly reduce the open-circuit voltage (V_{OC}) when used with the I^-/I_3^- electrolyte, particularly in complexes **12c** and **12d**, which incorporate $-NH_2$ and $-NO_2$ groups on the phenanthroline ligand. These groups caused significant changes in the interfacial charge transfer processes, thereby limiting device performance. The V_{OC} was notably decreased for complexes **12c** and **12d**, with values of 0.48 V and 0.44 V, respectively, compared to 0.67 V for complexes **12a** and **12b**. The decline in

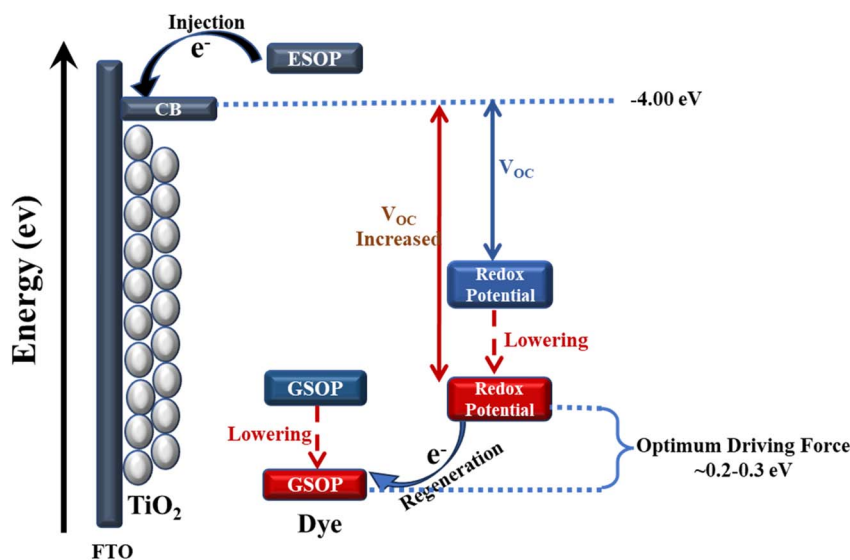


Fig. 12 Shows a strategy for boosting the V_{OC} by concurrently moving the redox couple potential and the dye's ground-state oxidation potential (GSOP) to higher positive values, resulting in an ideal driving force for rapid dye regeneration.

V_{OC} was ascribed to ineffective electron regeneration from I^-/I_3^- , primarily due to charge recombination caused by the quenching of the excited-state photosensitizers (12a-d) by the I^-/I_3^- .¹³⁷ This highlights the strong correlation between open-circuit voltage (V_{OC}) and the molecular structure of photosensitizers. The molecular structure greatly determines the strength of recombination processes between the photosensitizer and the I^-/I_3^- electrolyte.^{138,139} Therefore, to prevent the redox pair from interacting with the adsorbed

photosensitizer on the TiO_2 surface, it's important to reduce complex formation between the photosensitizer and the electrolyte through careful molecular engineering of the photosensitizers.

7.2. Developed strategy for efficient redox couples

In DSSCs, liquid electrolytes act as liquid redox systems, with the redox couple dissolved in aqueous or organic solvent media.

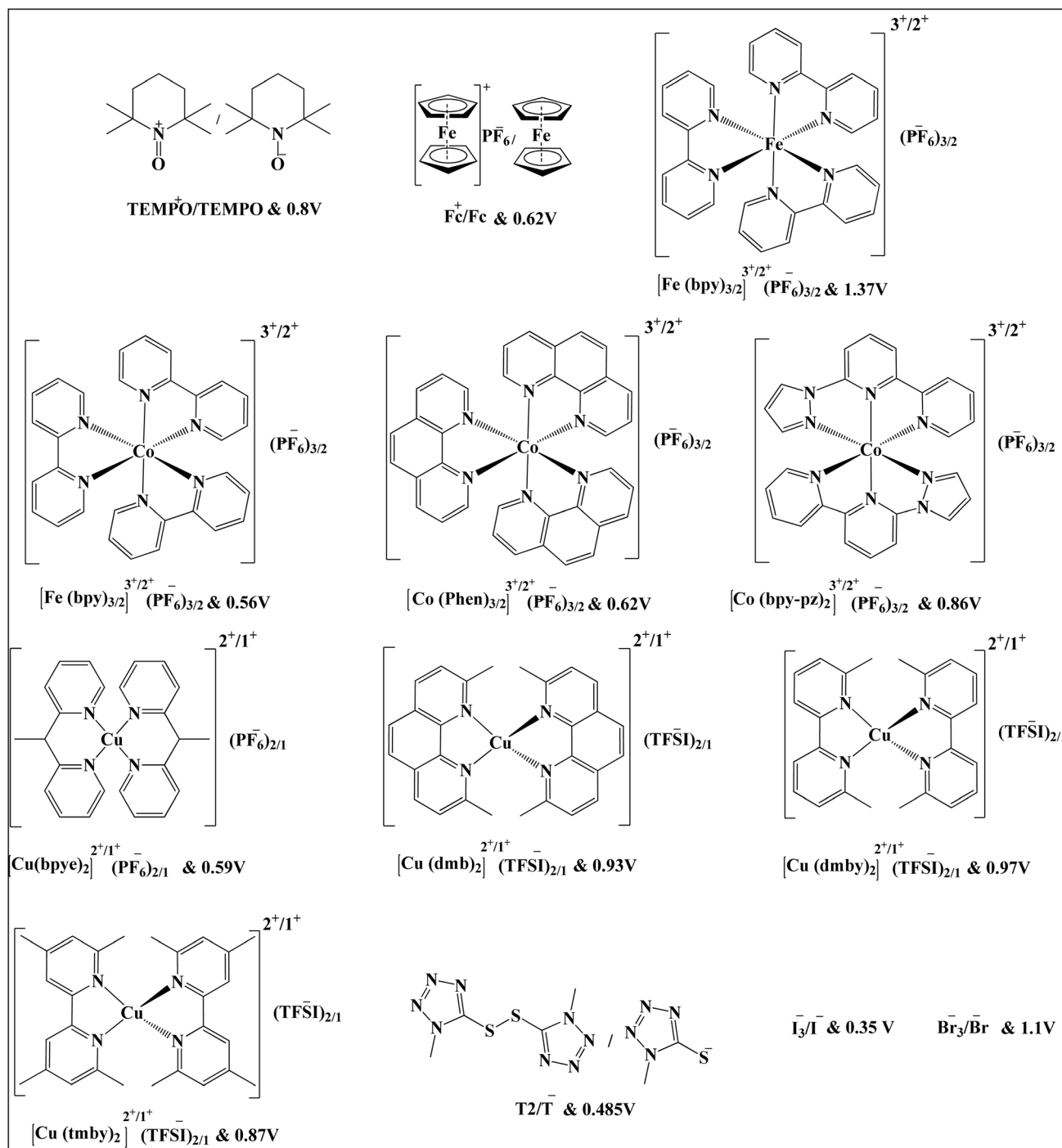


Fig. 13 Chemical structures of the most common redox couples used in DSSCs along with their standard redox potential. Where: TEMPO is 2,2,6,6-tetramethyl-1-piperidinyloxy & T₂ (dimer of 5-mercapto-1-methyltetrazole ion) & T⁻ (5-mercapto-1-methyltetrazole ion).



The iodide/triiodide (I^-/I_3^-) redox couple is the most often utilized electrolyte in DSSCs. However, it restricts the possible open-circuit voltage (V_{OC}) to roughly 0.70 to 0.80 V due to high energy loss during the photosensitizer regeneration process, which is a key disadvantage of current DSSC technology. Replacing the I^-/I_3^- electrolyte with more efficient redox couple can boost the device's photovoltage. To enhance the photovoltage, alternative redox couples with improved efficiency can replace the I^-/I_3^- system. However, for optimal photocurrent generation, the regeneration of the photosensitizer by the redox mediator must occur at a faster rate than the back-transfer of electrons from TiO_2 to the oxidized dye.¹⁴⁰ Ensuring a sufficient driving force for dye regeneration is crucial to prevent recombination of injected electrons with the oxidized dye. If the driving force is inadequate, recombination can reduce photovoltaic performance by decreasing both J_{SC} and V_{OC} . Conversely, a strong driving force can limit V_{OC} . Thus, efficient dye regeneration requires an ideal driving force of roughly 20–25 kJ mol⁻¹. This can be achieved by selecting a redox couple matched to the specific photosensitizer.¹⁴¹ Therefore, to improve the open-circuit potential and photocurrent of DSSCs, innovative redox mediators with a higher redox potential while preserving the ideal driving force for dye regeneration are required. The general technique for raising V_{OC} entails concurrently lowering the dye's ground state oxidation potential (GSOP) level and the redox shuttle's reduction potential to a higher potential (Fig. 12). This technique optimizes the driving force for dye regeneration, which varies depending on the redox couple used.

Since the discovery of DSSCs, various redox couples have been explored and used as potential redox shuttles in DSSCs, including: (1) halogenated redox couples such as I^-/I_3^- ¹⁴² and Br_3^-/Br^- ;¹⁴³ (2) electrolytes incorporating transition-metal complexes, such as Co^{3+}/Co^{2+} ,¹⁴⁴ Cu^{2+}/Cu^+ ,¹⁴⁵ Fe^{3+}/Fe^{2+} ,¹⁴⁶ and ferrocenium/ferrocene (Fc^+/Fc);¹⁴⁷ and (3) organic radical-based redox couples like TEMPO/TEMPO⁺¹⁴⁸ and T_2/T^- .¹⁴⁹ The chemical structures and the standard redox potentials of some developed redox shuttles used in DSSCs are illustrated in Fig. 13.

7.3. Electrolyte working principle

To better understand the working principles of electrolytes in DSSCs, the most effective redox couples, including I^-/I_3^- , cobalt complex electrolytes (Co^{2+}/Co^{3+}), and copper complex electrolytes (Cu^{2+}/Cu^+) were explored in depth, as shown in Fig. 14. Among these, the I^-/I_3^- redox couple is recognized for its outstanding efficiency and remains the most widely used electrolyte since its inception, primarily due to its favorable standard redox potential of approximately 0.35 V relative to the normal hydrogen electrode (NHE). This optimal potential of I^-/I_3^- enhances the driving force for dye regeneration across most photosensitizers. For example, the oxidation potential of I^-/I_3^- relative to the oxidation potential of $Ru(dcbpy)_2(NCS)_2$ photosensitizer such as **N3** or **N719**, which is around 1.1 V vs. NHE makes the driving force of regeneration of 0.75 V.¹⁵⁰ This substantial driving force (0.75 V) promotes efficient dye regeneration while mitigating the recombination of injected

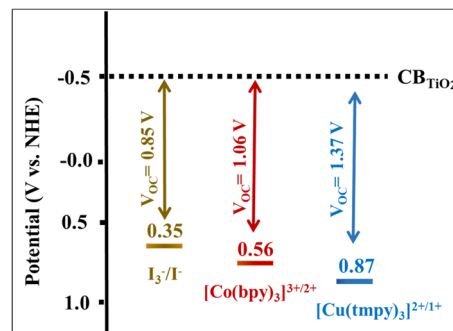


Fig. 14 Redox potential values and the corresponding V_{OC} values of the most common and efficient redox couples used in DSSCs, including I_3^-/I^- , $[Co(bpy)_3]^{3+/2+}$, and $[Cu(tmpy)_3]^{2+/1+}$.

electrons in TiO_2 with the electrolyte, ultimately enhancing the J_{SC} of the device.

Additionally, I^-/I_3^- redox electrolytes absorb light in the blue area of the visible spectrum, competing with dye absorption and resulting in decreased photocurrent.¹⁵¹ The corrosive behavior of I^-/I_3^- redox electrolytes, especially towards metal current collectors like silver (Ag), provides considerable hurdles for scaling up DSSCs to commercial modules.¹⁴⁸ Studies on metal thin films in the presence of I^-/I_3^- redox electrolytes demonstrate that metals like Ag, Au, and Al are extremely vulnerable to corrosion, while Pt, Ti, and Ni exhibit lesser corrosiveness.¹⁵² The I^-/I_3^- redox pair has substantial problems such as competing light absorption, metal corrosion, and limited photovoltage. To solve these difficulties, an appropriate alternative to the I^-/I_3^- redox pair is required.

To replace the I^-/I_3^- electrolyte, researchers have explored and utilized several redox couples that are less corrosive and possess suitable redox potentials to achieve high V_{OC} with promising solar-to-power conversion efficiencies.^{153,154} Saap *et al.* achieved a 12.3% efficiency in lab-scale DSSCs using the **N3** sensitizer under 1 sun illumination by employing cobalt complexes with substituted polypyridine ligands as potential alternatives to the volatile and corrosive iodide/triiodide (I^-/I_3^-) redox couple, the chemical structures of the developed polypyridyl cobalt(II/III) redox mediators incorporating are illustrated in Fig. 15. These cobalt complexes exhibit extinction coefficients around 10^2 M⁻¹ cm⁻¹ in the visible spectrum, minimizing interference with the light-harvesting capability of the photosensitizer.¹⁵⁵ Cobalt complexes offer distinct advantages over the I^-/I_3^- system, as the redox potential of the cobalt mediator can be fine-tuned by the ligands surrounding cobalt center. Additionally, bulky groups can be introduced as insulating spacers between the ligands, reducing recombination reactions between the I_3^- and titania, which ultimately leads to an increase in the V_{OC} . However, the best-performing mediator, based on tris(4,4'-di-*tert*-butyl-2,2'-dipyridyl)cobalt(II/III), achieved efficiency of up to 80% when compared to the standard iodide/triiodide mediator.

In conclusion, cobalt (Co^{2+}/Co^{3+}) mediator is useful since it is commercially available, easy to synthesize, and has nonvolatile and noncorrosive electron-transfer properties in DSSCs.



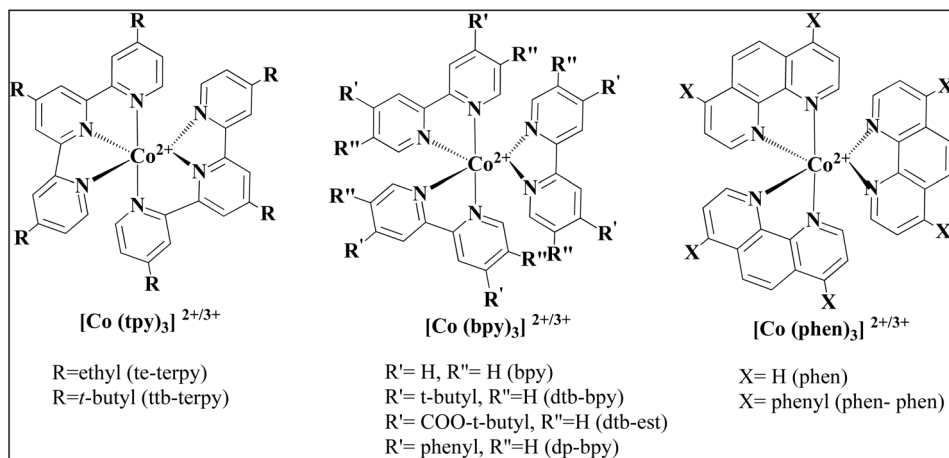


Fig. 15 Molecular structure of the terpyridine, bipyridine, and phenanthroline complexes of cobalt(II) mediator used in DSSCs with N3 photosensitizer.

Furthermore, cobalt polypyridine complexes are attractive as redox shuttles due to their low visible light absorption, low metal corrosion, outer-sphere one-electron redox chemistry, and higher positive redox potential.¹⁵⁶ Cobalt ($\text{Co}^{2+}/\text{Co}^{3+}$) electrolytes with bulky groups have much lower ion mobility during diffusion than iodine ions.¹⁵⁷

7.4. Electrolyte additives

In DSSCs, various nitrogen-containing heterocyclic derivatives such as imidazole, triazole, benzimidazole, and pyridine are commonly used as electrolyte additives to enhance photovoltaic performance.¹⁵⁸ The effects of these additives are well-documented and include modifying the redox couple

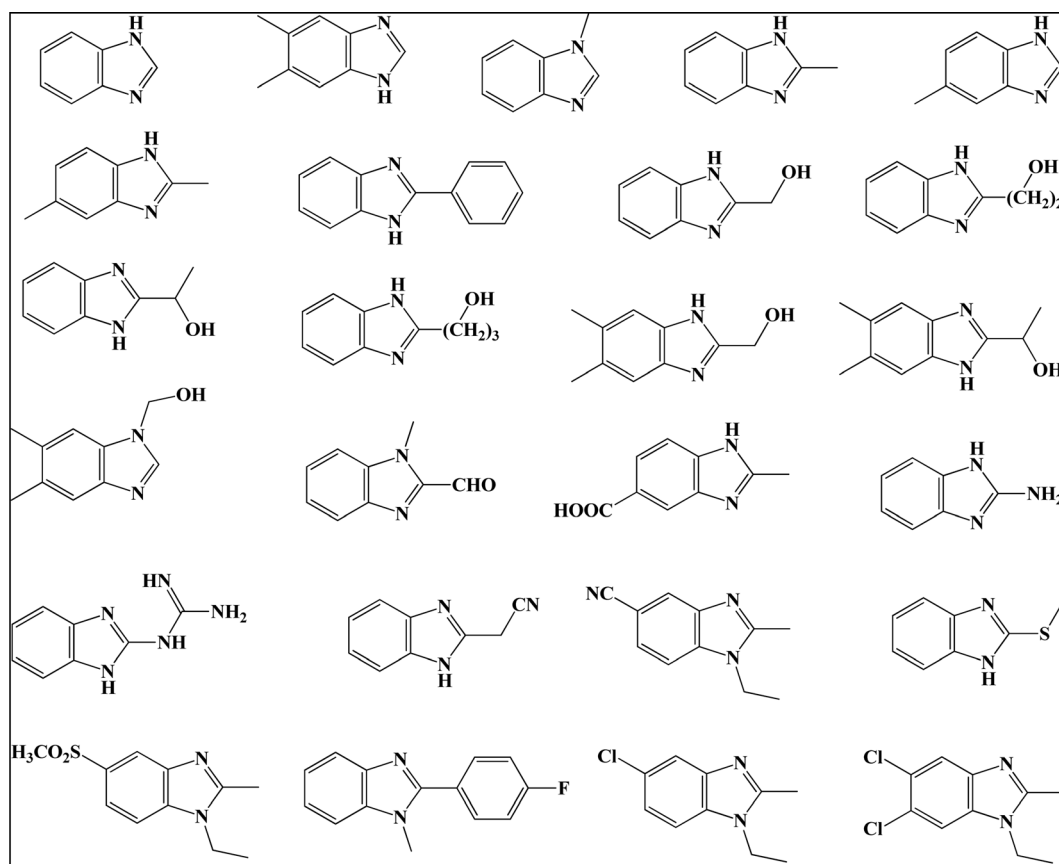


Fig. 16 Molecular structures of benzimidazole derivatives as additives in electrolyte.



potential, shifting the conduction band (CB) level of TiO_2 , forming a compact layer on the TiO_2 surface, reducing dye aggregation, and blocking the TiO_2 surface from the electrolyte. These additives adsorb onto the TiO_2 surface, elevating the conduction band level and increasing V_{OC} . The extent of this shift largely depends on the electron-donating power, charge density, and donating capability of the additives. However, while raising the TiO_2 conduction band level increases V_{OC} , it also reduces electron injection efficiency from the photosensitizers to TiO_2 , resulting in a decrease in the short-circuit current density (J_{SC}).¹⁵⁹ Kusama *et al.* investigated the impact of benzimidazole additives on DSSCs performance based on **N3** photosensitizer and an I^-/I_3^- redox electrolyte in acetonitrile. Their findings revealed that the addition of benzimidazoles enhances the V_{OC} and fill factor (ff) but reduces the J_{SC} . For example, 2-amino-1-methylbenzimidazole yielded the highest V_{OC} of 0.83 V but the lowest J_{SC} of 10.3 mA cm^{-2} . Computational calculations show that higher partial charges of nitrogen atoms in benzimidazoles increase V_{OC} but decrease J_{SC} . Smaller benzimidazole molecules and those with a larger dipole moment difference from acetonitrile improves V_{OC} and J_{SC} , respectively. The highest efficiency of 7.6% was achieved with 5-chloro-1-ethyl-2-methylbenzimidazole and 2-(1-hydroxyethyl) benzimidazole under 100 mW cm^{-2} illumination. These effects are attributed to interactions between benzimidazoles and the TiO_2 , altering dark currents and the flatband potential of TiO_2 .¹⁶⁰ The molecular structures of the benzimidazole-based additives used in electrolytes are listed in Fig. 16.

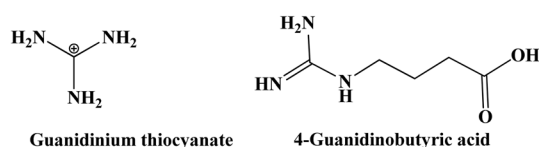


Fig. 17 Chemical structure of guanidinium cation and 4-guanidinobutyric acid additives.

Kopidakis *et al.* explores the effect of guanidinium as an adsorbent on the performance of DSSCs. Specifically, it examines how guanidinium influences recombination and band-edge movement in these cells. Guanidinium added to the electrolyte was found to slow down recombination by a factor of about 20, which would typically increase the V_{OC} . However, guanidinium also causes the TiO_2 band edges to shift downward by 100 mV, which would decrease V_{OC} . Despite this opposing effect, the overall impact of guanidinium results in a net improvement in V_{OC} of about 20 mV due to the dominant effect of reduced recombination. This improvement is attributed to the dominant effect of reduced recombination over the unfavorable band-edge shift. The downward band-edge shift is likely due to the positive charge of the guanidinium cation interacting with the TiO_2 surface, leading to a buildup of positive surface charges.¹⁶¹ Furthermore, Zhang *et al.* investigated the impact of 4-guanidinobutyric acid (GBA) on the performance of DSSCs fabricated with an amphiphilic ruthenium photosensitizer (**K-19**) and GBA. Their study demonstrated a notable increase in the V_{OC} by $\sim 50 \text{ mV}$ compared to cells without GBA. The inclusion of GBA caused a shift in the CB of TiO_2 to a more negative potential, thereby reducing the back-reaction between electrons in the TiO_2 conduction band and I_3^- in the electrolyte. This modification resulted in enhanced stability and efficiency, achieving a PCE of $\sim 8\%$ under simulated sunlight.¹⁶² The chemical structures of the guanidinium cation and 4-guanidinobutyric acid are presented on Fig. 17.

Furthermore, Kusama *et al.* reported that using 10 different aminotriazole derivatives as additives improved the V_{OC} and solar-to-power conversion efficiency (η) of DSSCs, although it reduced the J_{SC} . The highest efficiency (η) of 7.6% was achieved by adding 3-amino-1H-1,2,4-triazole to the electrolyte mixture.¹⁵⁹ The molecular structures of the ten aminotriazole derivatives additives used as electrolyte to additives to improve the V_{OC} of DSSCs are illustrated in Fig. 18.

Nazeeruddin *et al.* explored the addition of 4-*tert*-butylpyridine and 3-methyl-2-oxazolidinone in conjunction to enhance the performance of DSSCs based on **N3** photosensitizer.

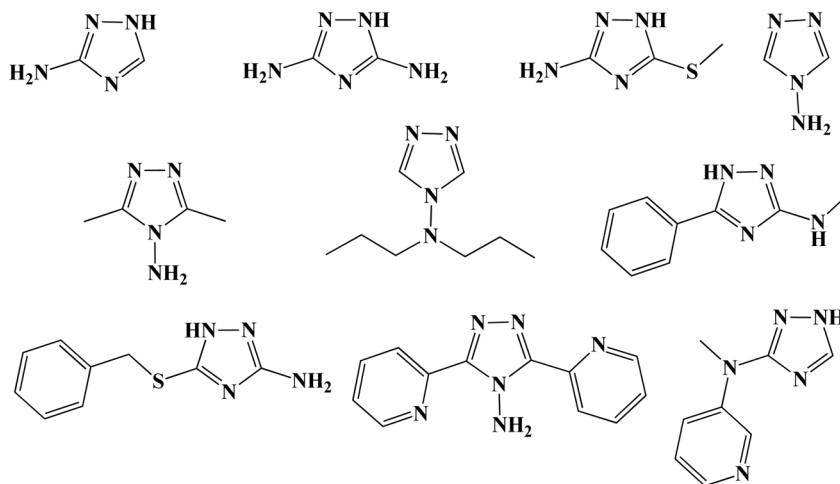


Fig. 18 Molecular structures of aminotriazole derivatives used as additives in electrolyte.

Initially, without treating the N3-covered film with 4-*tert*-butylpyridine, the V_{OC} was 0.38 V, and the PCE was approximately 3.7%. After treating the dye-coated TiO_2 film with 4-*tert*-butylpyridine, V_{OC} increases to 0.66 V, and the PCE improved to about 8.5%. The high V_{OC} is attributed to the dark current suppressing at the semiconductor/electrolyte junction. This suppression occurs because 4-*tert*-butylpyridine adsorbs onto the TiO_2 surface, blocking surface states that facilitate the reduction of I_3^- by conduction band electrons, thus reducing recombination and enhancing efficiency. Further improvement is achieved by using a mixture of acetonitrile and 3-methyl-2-oxazolidinone (90/10, v/v) as the electrolyte solvent, which increases V_{OC} to 0.72 V and maintains a high PCE of $\sim 10\%$.¹⁶³ The addition of 3-methyl-2-oxazolidinone to the electrolyte improves the overall stability and efficiency of the cell by optimizing the electrolyte's properties, such as viscosity and polarity, which can enhance ion mobility and reduce recombination losses. Finally, Afrooz *et al.* investigated the use of diethyl oxalate (DEOX) as an effective additive in dye-sensitized solar cells (DSSCs) with iodide/triiodide (I^-/I_3^-) electrolytes. They examined the impact of DEOX and 4-*tert*-butylpyridine (TBP) additives on DSSC performance by comparing various electrolyte compositions. Their findings revealed that electrolytes containing both DEOX and TBP additives exhibited the best photovoltaic performance. The addition of DEOX led to increased short-circuit current density and open-circuit voltage, resulting in improved overall efficiency. They proposed that DEOX forms a complex with iodine molecules, influencing the electron transfer processes in the cell.¹⁶⁴

8. Conclusion

This paper offers a comprehensive review of optimization strategies for dye-sensitized solar cells (DSSCs), highlighting recent advancements across key components, including Ru(II) photosensitizers, electrolytes, co-adsorbents, and additives. The review underscores the critical role of modifying Ru(II) photosensitizers to enhance excited state lifetimes and minimize charge recombination. By incorporating electron-donating or withdrawing groups and utilizing ligands with extended π -systems, significant improvements in electron injection efficiency and reduced non-radiative decay have been achieved, leading to enhanced photogenerated current and overall cell performance. Structural modifications of Ru(II) photosensitizers concluded that the most effective photosensitizer is one with an architecture consisting of one auxiliary ligand and one anchoring ligand, incorporating two carboxylic acid groups and two thiocyanate (SCN) groups. Additionally, the development of thiocyanate-free Ru(II) complexes and cyclometallated polypyridyl Ru(II) complexes has shown promising potential in enhancing the long-term stability of DSSCs. Furthermore, charge recombination is addressed not only through structural modifications of Ru(II) photosensitizers and careful engineering of the semiconductor/electrolyte interface but also through the exploration of efficient electrolytes. These electrolytes are designed to outperform conventional I^-/I_3^- electrolytes, which can cause corrosion, light absorption competition with the

photosensitizer, suppress the open-circuit voltage (V_{OC}) to about 0.70 to 0.80 V, and decrease overall power conversion efficiency. Various alternative redox couples have been encouraged such as transition-metal complexes (*e.g.*, Co^{3+}/Co^{2+} , Cu^{2+}/Cu^+). These alternatives aim to optimize the redox potential and improve overall DSSC performance. Finally, the development of convenient co-adsorbents has been instrumental in preventing photosensitizer aggregation, thereby improving device stability and efficiency. In summary, these advancements represent a significant and forward-thinking approach to optimizing DSSCs. Continued refinement of molecular designs, reduction of charge recombination, and innovation in co-adsorbents and redox electrolytes are essential for achieving higher efficiencies and enhancing commercial viability. This review not only highlights the substantial progress made but also sets the stage for future breakthroughs that will further advance DSSC technology and its application in the renewable energy sector.

Data availability

No new data were created or analyzed in this study. Data sharing is not applicable to this article.

Conflicts of interest

The author has no conflict to declare.

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