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**COMMUNICATION**

Andrea M. Armani *et al.*  
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electric field sensing and modulation

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for electric field sensing and modulation†Yingmu Zhang,<sup>a</sup> Jinghan He,<sup>b</sup> Patrick J. G. Saris,<sup>a</sup> Hyun Uk Chae,<sup>c</sup> Subrata Das,<sup>ib</sup>  
Rehan Kapadia<sup>c</sup> and Andrea M. Armani<sup>id</sup> \*<sup>abc</sup>Cite this: *J. Mater. Chem. C*, 2022,  
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Nonlinear optical organic molecules have advanced a wide range of fields spanning from integrated photonics to biological imaging. With advances in molecular design, an emerging application is multifunctional nonlinear organic materials. Unlike conventional molecules which simply emit light through single or multi photon processes, multifunctional materials can perform multiple tasks, such as modulating the optical signal or reporting an electric field intensity. In this work, we report a multifunctional organic molecular device with electric field 'sense-and-modulate' capability. This is achieved by combining two distinct functional modules. The electric field-reporting module relies on a photo-induced electron transfer (PeT) dye, tetraphenylethylene (TPE) as a two-photon (2p) imaging agent; while the electric field-modulating module relies on the organic photoconductor, naphthalimide (NAI). To reduce cross-talk between the two modules, they are separated by a long alkyl chain. The photophysical properties and photoconductivity of the probe molecule are studied in a range of solvents and in the solid state, and the results agree with the density functional theory predictions. Specifically, 2p excitation is demonstrated, and the photoconductivity is rapid and reversible. The entire system is optically controlled, including signal read-out, and the two modules can be operated simultaneously or individually.

## Introduction

Whether organic or inorganic, fluorescent molecules and particles have been a critical tool of discovery and an innovation

platform for nearly 200 years. For example, fluorophores have revealed cell structure and function, been used as indicators for detecting toxins, and have served as the active layer in many optoelectronic devices.<sup>1–6</sup> This prior work inspired the concept of an organic molecular device, or a device comprising a single or cluster of organic molecules. Initial research in this area leveraged existing probe molecules to demonstrate biodetection (eg molecular beacons) and biomolecular motors based on optical or mechanical responses.<sup>7–11</sup> Given their length scale, the potential applications of molecular devices extend far beyond biological systems into optoelectronics and quantum devices. For example, recent work has explored stimulated emission materials, organic solid-state lasers, and frequency-upconversion.<sup>12–18</sup> However, to create a molecular device capable of performing multiple functions requires combining several functional groups into a single molecular construct.

One of the key challenges in designing multifunctional molecules is reducing the cross-talk between the different excitation mechanisms. This type of interference can reduce the overall system efficiency and performance.<sup>19,20</sup> With recent advances in machine learning and computational design algorithms, material chemists are able to optimize molecular designs accelerating experimental efforts.<sup>21–23</sup>

In the present work, we develop and characterize an all-optical multifunctional molecular device designed to simultaneously sense and modulate electric fields. The molecule, named as NAI-TPE-PyS, is comprised of two non-interacting modules that are tethered together by a non-interacting spacer (Fig. 1). The sensor module is derived from tetraphenylethylene (TPE), which is a two-photon (2p) fluorophore that is excited in the near-IR (NIR). The modulator module is an organic photoconductor, naphthalimide (NAI), whose resistivity can be tuned using an ultraviolet optical source. The non-interacting alkyl chain does not affect the photophysical properties of the modules but rather reduces Dexter energy transfer between modules by physical spacing.

The TPE and NAI modules were judiciously chosen for low overlap of their absorption and emission spectra to reduce

<sup>a</sup> Mork Family Department of Chemical Engineering and Materials Science, University of Southern California, Los Angeles, California 90089, USA.  
E-mail: armani@usc.edu

<sup>b</sup> Department of Chemistry, University of Southern California, Los Angeles, California 90089, USA

<sup>c</sup> Ming Hsieh Department of Electrical and Computer Engineering-Electrophysics, University of Southern California, Los Angeles, California 90089, USA

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Fig. 1 Chemical structure of NAI-TPE-PyS with TPE-module in red and NAI-module in blue.

Foerster resonant energy transfer. However, the precise chemical structure of the molecule is optimized using density functional theory as part of this work. The tether enforces colocalization of the modules to reduce the possibility for bulk phase separation of the modules, which could occur for a simple blend of the two components under the processing conditions for various applications. The entire system is optically controlled, including signal read-out, and the two modules can be operated simultaneously or individually (Fig. 1).

The TPE sensor module relies on the ultrafast photo-induced electron transfer (PET) process to allow large changes in the electric field to be detected. Unlike voltage sensitive dyes,<sup>24–28</sup> PeT molecules show superb signal-to-noise ratio, sensitivity, and response speed.<sup>29–33</sup> This performance is inherent to the ultrafast PeT process, which is an intramolecular electron transfer through its donor-spacer-acceptor backbone upon exposure to excitation sources, competing with the radiative fluorescent emission.<sup>34,35</sup> In addition, TPE exhibits a unique behavior known as aggregation-induced emission (AIE). In an AIE material, the emission signal in the condensed state is increased due to restricted intramolecular rotation (RIR).<sup>36,37</sup> Among PET and AIE materials, TPE is unique because it exhibits both behaviors.

As mentioned, NAI is a photoconductor.<sup>38–41</sup> This type of molecule allows the resistivity to be modulated with light. Although organic photoconductors have been widely used in optoelectronics,<sup>42–47</sup> their integration into a multifunctional molecule is underexplored.<sup>48</sup> Moreover, an organic molecular device with both electric field sensor and modulator capabilities has yet to be demonstrated.

## Results and discussions

### Molecular design and computational modelling

The molecular device is comprised of two covalently coupled modules: a tetraphenylene (TPE)-based PeT dye connected with naphthalimide (NAI)-based photoconductor.

As mentioned, the NAI “modulate” module was chosen due to its extensive use as an organic photoconductor and was incorporated without structural modification. The “sense” module of NAI-TPE-PyS was designed based on TPE but was modified with PyS to create TPE-PyS, representing the donor-spacer-acceptor system, as required by a PeT dye.<sup>49</sup>

To assist in this design, the reporter part was modelled *in silico* based on a truncated structure, namely, Me-TPE-PyS (Fig. 2a and b). The alkyl linkers and peryleneimide moieties of NAI-TPE-PyS are separated from electronic communication with TPE-module and therefore were replaced with methyl groups to reduce the computational cost of the modelling. The ground state equilibrium geometry and frontier orbital densities of Me-TPE-PyS were calculated by density functional theory (DFT) in the gas phase at the B3LYP/6-31g\* level of theory. With TPE as an electron donating group and PyS as an electron accepting one, the HOMO density resides on the TPE moiety with some extension onto the phenylene linker (Fig. 2c, red solid). The LUMO density resides on the pyridinium and benzylidene moieties with some extension onto the phenylene linker (Fig. 2c, blue mesh).

The excited states were calculated by TD-DFT with a conductor-like polarizable continuum medium (dielectric = 78.39, water) at the CAM-B3LYP/6-311++G\*\* level of theory. Fig. 2d–f shows the natural

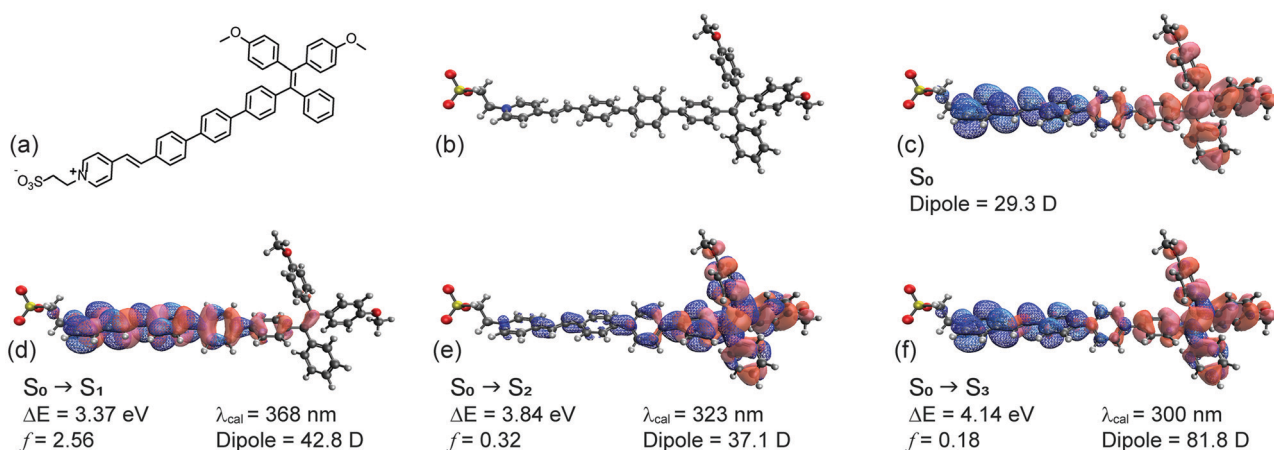


Fig. 2 TPE-PyS model compound: (a) chemical structure; (b) optimized ground state geometry; (c) HOMO (red solid) and LUMO (blue mesh) iso-surfaces ( $0.05 \text{ \AA}^{-3}$ ); natural transition orbital iso-surfaces ( $0.05 \text{ \AA}^{-3}$ ) for hole (red solid) and electron (blue mesh) wavefunctions: (d)  $S_1$ , (e)  $S_2$ , and (f)  $S_3$ .





transition orbitals for the lowest three possible singlet states along with corresponding transition properties. Accordingly, the first singlet excited state ( $S_1$ ) has a highly allowed vertical transition with an energy of 3.37 eV (368 nm) (Fig. 2d). The dipole moment of  $S_1$  (42.8 D) is 18.9 D more than that of the ground state (29.3 D), suggesting that it may be moderately solvatochromic. This transition is characterized by a  $\pi$ - $\pi^*$  transition on the pyridinium and phenylene linker, corresponding to electron transfer from lower lying occupied orbitals to the LUMO. The second singlet excited state ( $S_2$ ) has a vertical transition energy of 3.84 eV (323 nm) (Fig. 2e). The dipole moment of  $S_2$  (37.1 D) is 7.8 D more than that of the ground state (29.3 D), suggesting that it may be weakly solvatochromic. This transition is characterized by a localized  $\pi$ - $\pi^*$  transition on the TPE moiety, corresponding to electron transfer from the HOMO to higher lying unoccupied orbitals. The third singlet excited state ( $S_3$ ) has a vertical transition energy of 4.14 eV (300 nm) (Fig. 2f). The dipole moment of  $S_3$  (81.8 D) is 52.5 D more than that of the ground state (29.3 D), suggesting that it may be strongly solvatochromic. This transition corresponds to the HOMO to LUMO transition, representing charge transfer from TPE to PyS. Notably, based on this analysis, solvatochromicity, which is an indicator of voltage sensitivity, is anticipated.

## Synthesis

The synthetic route of NAI-TPE-PyS is summarized in Scheme 1, and the synthesis details are included in the (Scheme S1 and Fig. S1-S14, ESI†). Using a Suzuki coupling reaction, a TPE unit, 4,4'-(2-(4-bromophenyl)-2-phenylethene-1,1-diyl)-diphenol, was conjugated with a phenylenevinylene spacer at the bromine terminal to form HO-TPE-CHO. The hydroxide terminals were then connected to two alkylated NAI moieties *via* Williamson etherification. Finally, a Knoevenagel condensation of the aldehyde terminal with an electron-withdrawing pyridinium inner salt, 3-(4-methylpyridin-1-ium-1-yl)propane-1-sulfonate, yielded NAI-TPE-PyS as an orange powder with a 32% overall yield.

## Spectroscopic characterization

Due to the combination of the hydrophobic NAI terminals and the hydrophilic pyridinium inner salt terminal, NAI-TPE-PyS can be readily dissolved in polar aprotic solvents like DMSO and DMF as well as in low-polarity lipophilic solvents like DCM

and chloroform. However, it shows low solubility in protic solvents like water or PBS buffer. Accordingly, the steady-state photophysical properties of NAI-TPE-PyS were analysed both in solutions with different polarity and in solid state thin films using ultraviolet-visible (UV-vis) and photoluminescence (PL) spectroscopy.

TPE is a moderate electron donor, and the pyridinium inner salt is a strong electron acceptor. Thus, the interaction between these compounds will lower the HOMO-LUMO energy gap, facilitating the intramolecular charge transfer (CT) process. As a result, the normalized absorption spectra of the molecule exhibit a broad profile tailing to 500 nm (Fig. S15, ESI†). Detailed photophysical data of NAI-TPE-PyS in selective solvents are shown in Fig. S16 and in Table S1 (ESI†). At shorter wavelengths, around 300–380 nm, the absorption band can be assigned to the  $n$ - $\pi^*$  and  $\pi$ - $\pi^*$  transitions of NAI- and the elongated TPE-conjugation moieties which overlaps with the absorption of NAI-TPE-CHO precursor. The absorption band at wavelengths longer than 380 nm, which is not present in the NAI-TPE-CHO precursor, can be assigned to the CT transition from TPE to pyridinium inner salt.

The two modules of NAI-TPE-PyS can be activated under independent optical control (Fig. 3). When excited at 350 nm, NAI-TPE-PyS displays characteristic narrowed emission of naphthalene derivatives centered at 400 nm, which indicates that the emission mainly stems from NAI-based module. When excited at 390 nm, NAI-TPE-PyS emits orange luminescence with the maximum peak at 585 nm dominated by the emission of TPE-based PeT dye, while no emission of NAI-module was observed. Consequently, the dual-component compound can be controlled by two wavelengths, where 390 nm is used to excite the TPE-module for voltage imaging and 350 nm is used to drive the photo-conductive NAI-module for voltage modulation. The difference in energy between these wavelengths (*circa* 400 meV) is enough to prevent crosstalk between the two modules.

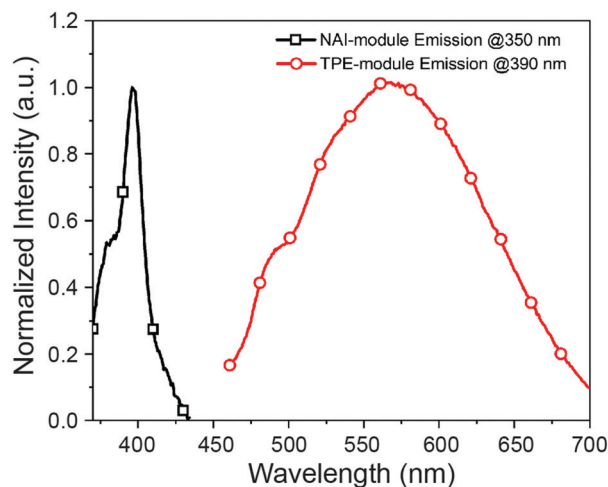
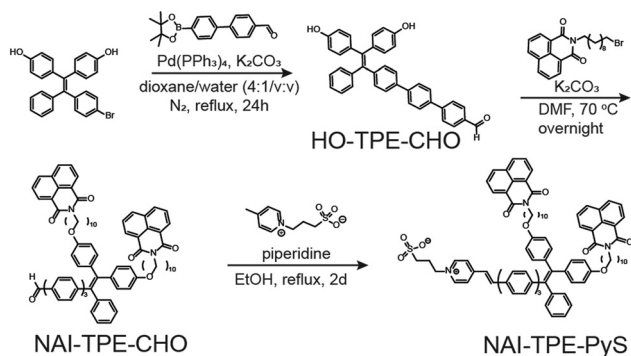


Fig. 3 Emission spectra of NAI-TPE-PyS at  $\lambda_{\text{ex}}$  = 350 nm (black squares) and  $\lambda_{\text{ex}}$  = 390 nm (red circles) in PBS buffer (concentration: 10  $\mu$ M, containing 1% DMSO).



Scheme 1 The synthesis scheme of NAI-TPE-PyS.



To investigate the predicted solvatochromism of the TPE-PyS module, the emission was characterized in several solvents. When excited at 390 nm in PBS, the NAI-TPE-PyS emitted at 582 nm (Fig. 4). The photoluminescent quantum yield in an aqueous solution (10  $\mu$ M, containing 1% DMSO) was measured to be 0.02 (2%). When the NAI-TPE-PyS was a solid, the quantum yield increased to 8%, providing preliminary evidence of AIE behavior.

This emission wavelength blue-shifts when the solvent is changed to DCM (Fig. 4). To further confirm this behavior, the emission was characterized over a range of solvents with varying polarities from chloroform to water, and a linear Lippert–Mataga correlation was established by plotting Stokes shifts ( $\nu_{\text{abs}} - \nu_{\text{em}}$ ) against solvent orientation polarizability ( $\Delta f$ ) (Fig. S19, ESI†).<sup>50</sup> From the slope of the Lippert–Mataga plot, the increase of dipole moment ( $\mu_e - \mu_g$ ) from ground state to excited state was calculated to be 29.8 D. Therefore, using the estimation of  $\mu_e$  from the DFT model (59.1 D), the ground state dipole ( $\mu_g$ ) is calculated to be 29.3 D. This large value indicates that the CT-component of the excited state will be stabilized in polar media.<sup>51,52</sup>

To further investigate the CT process, the related molecule NAI-TPE-CHO was synthesized and characterized. The maximum emission peak of this molecule was at 420 nm, representing a large bathochromic shift of 162 nm from the NAI-TPE-PyS compound. This difference can be ascribed to the donor-spacer-acceptor structure. Such interplay also explains the relatively weak emission of NAI-TPE-PyS compared to other published TPE-based AIEgens,<sup>53</sup> as non-radiative deactivation of the CT state competes with the radiative process.

Based on the combination of the computational and experimental results, we conclude that the reporter module will be a

voltage sensitive dye. Computationally, the reporter has the proper ordering of excited state energies and dipole moments to facilitate voltage dependence. Spectroscopically, the synthesized dye displays the characteristic solvatochromism associated with voltage sensitive fluorophores due to sensing of the dielectric environment. We expect that it will demonstrate voltage sensitive luminescence when the molecule's motion is restricted, for example, if the molecule is inserted and aligned in a cellular membrane.

### Aggregation induced emission properties

To rigorously investigate the AIE behavior of the TPE core, two approaches were used to initiate aggregation in a controlled manner. First, poor or low efficacy solvents were used. In our studies, DMSO was an ideal solvent for NAI-TPE-PyS, and THF was an antisolvent for NAI-TPE-PyS. Thus, by using a mixture of DMSO:THF, aggregates were induced. Second, the concentration of the NAI-TPE-PyS molecule in water was systematically increased, resulting in aggregation or micelle formation.

The DMSO solution of NAI-TPE-PyS (20  $\mu$ M) exhibited weak fluorescence centered at 545 nm (Fig. S20, ESI†). As the concentration of THF was increased, the emission intensity increased, due to the formation of aggregates. Finally, at a THF fraction of 99%, the emission intensity of the mixture reached a maximum which was 3-fold higher than that of the initial DMSO solution (Fig. S20, ESI†). Based on the results in Fig. S20 (ESI†), the introduction of the alkyl NAI terminals did not eliminate the AIE response. However, it should be noted that, unlike other previously reported TPE-based AIEgens whose emission experiences significant enhancement upon aggregation in solution, the emission intensity of NAI-TPE-PyS at high THF fraction merely increases three times. This could be attributed to the long alkyl chains of the NAI-TPE-PyS interfering with the aggregation process, ultimately limiting the size of the aggregates.<sup>51</sup> Using dynamic light scattering (DLS), the diameter of the nanoaggregates was measured to be approximately 500 nm (Fig. S21, ESI†). Similar increases in emission intensity were observed as the concentration of NAI-TPE-PyS in water was increased from 0.01 to 80  $\mu$ M (Fig. S20, ESI†).

### Nonlinear optical properties

In addition to single photon excitation processes, NAI-TPE-PyS can also undergo two-photon excitation processes. The emission decay lifetime in a DMSO solution under 750 nm excitation was measured using a confocal microscope-based Time-correlated single photon counting (TCSPC) setup. The TCSPC histogram and exponential decay fit are provided in Fig. S22 (ESI†). Fluorescence at wavelengths shorter than 550 nm were observed from NAI-TPE-PyS with a monoexponential excited state decay time of 160 ps. Using the photoluminescent quantum yield of 0.02 measured previously, the radiative and non-radiative rate constants are calculated to be  $1.25 \times 10^8 \text{ s}^{-1}$  and  $6.125 \times 10^9 \text{ s}^{-1}$  respectively, which is consistent with a weakly emitting organic fluorophore. These results confirm the predicted two-photon behavior of NAI-TPE-PyS using NIR

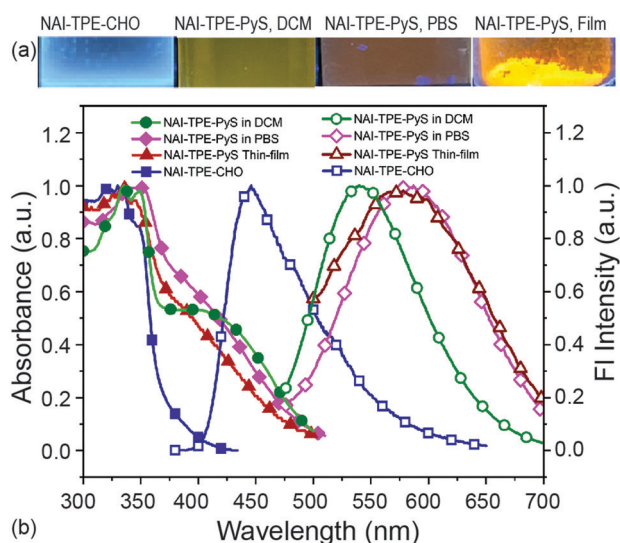
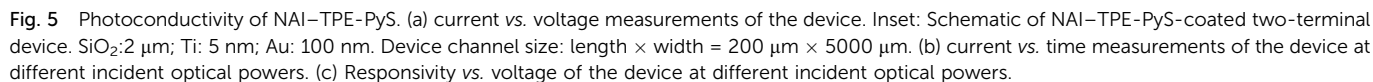


Fig. 4 Optical absorption and emission properties of NAI-TPE-PyS. (a) Images of fluorescence from NAI-TPE-CHO and NAI-TPE-PyS solutions and solid powder ( $\lambda_{\text{ex}} = 365 \text{ nm}$ ). (b) Normalized absorption spectra (solid symbols) and emission spectra (hollow symbols) of NAI-TPE-PyS in DCM or PBS or in solid state, and NAI-TPE-CHO at  $\lambda_{\text{ex}} = 390 \text{ nm}$ .





All commercially available chemical reagents and solvents were either purchased from VWR or Sigma Aldrich. Unless otherwise noted, they were used without further purification.  $^1\text{H}$  and  $^{13}\text{C}$ -NMR spectra were recorded on a Varian Mercury 400 MHz spectrometer with 96-spinner sampler changer using either deuterated chloroform or DMSO as solvent, as indicated. UV-Vis absorption spectra were measured on Beckman Coulter Life Science UV/vis spectrophotometer, DU 730 with a wavelength resolution of 2 nm. Steady-state fluorescence spectra were recorded on Horiba Scientific Fluoromax-4 spectrofluorometer with excitation slit width of 5 nm and emission slit width of 5 nm. Quantum yield was determined using a Quanta- $\phi$  integrating sphere attached to the Fluoromax-4. Time-resolved



photoluminescence (TRPL) measurements were performed with a standard confocal microscope-based Time-correlated single photon counting (TCSPC) setup.

### Synthesis of HO-TPE-CHO

4'-(4,4,5,5-Tetramethyl-1,3,2-dioxaborolan-2-yl)-[1,1'-biphenyl]-4-carbaldehyde (2.3 mmol, 0.7 g), 4,4'-(2-(4-bromophenyl)-2-phenylethene-1,1-diyl)-diphenol (1.9 mmol, 0.84 g), Pd(PPh<sub>3</sub>)<sub>4</sub> (0.19 mmol, 200 mg) and K<sub>2</sub>CO<sub>3</sub> (7.6 mmol, 1.05 g) were added into a 100 mL round-bottom flask. The flask was fitted on the Schlenk line, vacuumed, and refilled with nitrogen iteratively three times. A mixing solvent (dioxane/water: 40 mL/10 mL) was bubbled with nitrogen for 30 min and then transferred to the flask through a canula. The mixture was then allowed to react for 24 hours at 100 °C. After cooling to the room temperature, the mixture was poured into water, and the pH was adjusted to 5. Then, the system was extracted with dichloromethane (DCM) and washed with water three times. The organic solvent was then removed and the crude solid recrystallized by hexane/DCM to give the pure product. <sup>1</sup>H NMR (400 MHz, Chloroform-*d*) δ 10.06 (s, 1H), 7.96 (d, *J* = 8.1 Hz, 2H), 7.79 (d, *J* = 8.4 Hz, 2H), 7.68 (s, 4H), 7.41 (d, *J* = 8.2 Hz, 3H), 7.12 (d, *J* = 2.2 Hz, 4H), 7.06 (s, 2H), 6.93 (dd, *J* = 16.6, 8.5 Hz, 4H), 6.59 (t, *J* = 9.0 Hz, 4H). <sup>13</sup>C-NMR (101 MHz, Chloroform-*d*): 191.91, 154.22, 154.14, 146.48, 144.10, 143.74, 140.85, 140.34, 138.79, 137.56, 136.40, 135.16, 132.78, 131.90, 131.42, 130.32, 127.75, 127.64, 127.46, 127.37, 126.22, 126.13, 114.68, 114.56.

### Synthesis of NAI-TPE-CHO

Into a 250 mL two-necked round-bottom flask was added K<sub>2</sub>CO<sub>3</sub> (248.8 mg, 1.8 mmol), 2-(10-bromodecyl)-1*H*-benz[*de*]isoquinoline-1,3(2*H*)-dione (500 mg, 1.2 mmol) and HO-TPE-CHO (163 mg, 0.3 mmol). The flask was vacuumed and purged with dry N<sub>2</sub> three times. Then DMF (15 mL) was added, and the reaction was stirred overnight at 70 °C. After cooling to room temperature, the mixture was poured into water, extracted with DCM, washed with distilled water several times, and dried with anhydrous magnesium sulfate. The crude product was purified by silica column chromatography with hexane and ethyl acetate (gradient to 1 : 1/v:v) as eluent to give NAI-TPE-CHO as a yellow viscous oil. <sup>1</sup>H NMR (400 MHz, Chloroform-*d*) δ 9.97 (s, 1H), 8.54–8.48 (m, 4H), 8.16–8.11 (m, 4H), 7.87 (d, *J* = 8.4 Hz, 2H), 7.74–7.63 (m, 7H), 7.60 (s, 4H), 7.32 (d, *J* = 8.4 Hz, 2H), 7.07–6.97 (m, 6H), 6.87 (dd, *J* = 17.3, 8.8 Hz, 4H), 6.60–6.51 (m, 4H), 4.12–4.07 (m, 4H), 3.79 (t, *J* = 6.5 Hz, 4H), 1.71–1.59 (m, 8H), 1.28 (d, *J* = 42.0 Hz, 24H). <sup>13</sup>C-NMR (101 MHz, Chloroform-*d*): 191.52, 164.19, 133.83, 132.61, 132.21, 131.56, 131.16, 130.30, 128.13, 127.79, 127.60, 127.42, 127.34, 126.09, 122.73, 113.90, 68.23, 40.48, 32.75, 32.64, 29.38, 29.32, 29.29, 29.28, 28.08, 28.04, 27.09, 26.02, 25.95, 25.78, 25.64.

### Synthesis of NAI-TPE-PyS

A mixture solution of NAI-TPE-CHO (300 mg, 0.25 mmol), 3-(4-methylpyridin-1-ium-1-yl)propane-1-sulfonate (53 mg, 0.25 mol), and piperidine catalyst (0.2 mL) was refluxed in 10 mL dry EtOH under N<sub>2</sub> for 48 hrs. The solution turned deep red. After cooling to room temperature, solvent was removed and the crude solid was

purified by column with eluent of DCM:MeOH (10:1/v:v) to give red solid. <sup>1</sup>H NMR (400 MHz, dmso) δ 9.21 (d, *J* = 6.9 Hz, 1H), 8.94 (d, *J* = 7.2 Hz, 1H), 8.47–8.41 (m, 10H), 8.29 (s, 2H), 8.20 (d, *J* = 7.0 Hz, 2H), 7.86–7.66 (m, 8H), 7.64 (d, *J* = 8.8 Hz, 1H), 7.51 (d, *J* = 8.5 Hz, 1H), 7.10 (dd, *J* = 14.4, 7.3 Hz, 4H), 7.03–6.95 (m, 6H), 6.84 (dd, *J* = 22.2, 10.3 Hz, 4H), 6.69–6.61 (m, 4H), 4.80 (t, *J* = 6.9 Hz, 2H), 4.20 (t, *J* = 7.0 Hz, 4H), 3.81 (t, *J* = 6.8 Hz, 4H), 2.45–2.41 (m, 2H), 2.22 (t, *J* = 6.9 Hz, 2H), 1.59–1.52 (m, 10H), 1.26–1.19 (m, 22H). <sup>13</sup>C NMR (101 MHz, cdcl<sub>3</sub>) δ 164.12, 163.99, 133.75, 132.84, 131.10, 131.10, 127.57, 127.09, 126.87, 122.68, 56.55, 40.44, 29.68, 29.43, 29.30, 28.07, 27.09, 26.41, 22.66.

### Photoconductivity measurement and fabrication

SiO<sub>2</sub>/Si (2 μm thermal SiO<sub>2</sub>) substrate was cleaned with acetone, IPA and D.I. water. A combination of photolithography, metal deposition, and lift-off were performed to pattern the electrical pad. Photoresist (AZ 5214) was spuncoat for 60 seconds at 3000 rpm, and a standard template mask with channel dimensions of 200 μm (length) × 5000 μm (width) was used as the electrode pattern. 5 nm Ti and 100 nm Au was deposited using e-beam evaporation (Temescal, SL1800), and the residual photoresist was lifted-off, completing the electrode fabrication. A thin film of NAI-TPE-PyS was spun-coat on the device from a 6 wt% chloroform solution for 30 s with 500 rpm. Dark and light current measurements were performed. Photoconductivity of the device and *I*-*t* measurements were characterized by a Semiconductor Parameter Analyzer (Keysight B1500a). UV light source was provided by DYMAX LED DX-100.

### Author contributions

Y. Z. synthesized molecule, performed chemical and photo-physical characterizations, and prepared the manuscript. J. H. assisted in the synthesis and manuscript writing. P. J. G. S. performed theoretical calculations, contributed to the initial molecule design, and assisted in preparing the manuscript. H. U. C. performed photoconductivity measurements and analyzed data. S. D. performed Time-resolved photoluminescence (TRPL) measurements. A. M. A. and R. K. supervised the research and manuscript writing. All authors have given approval to the final version of the manuscript.

### Conflicts of interest

There are no conflicts to declare.

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