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Nano onions based on an amphiphilic $\text{Au}_3(\text{pyrazolate})_3$ complex†‡

Atena B. Solea,^a Davide Dermutas,^a Farzaneh Fadaei-Tirani,^a Luigi Leanza,^b Massimo Delle Piane,^b Giovanni M. Pavan^{b,c} and Kay Severin^{b,*a}

Multilayer vesicles with an onion-like architecture can form by self-assembly of organic amphiphiles such as dendrimers, small-molecule surfactants, and block copolymers. Thus far, there are limited reports about multilayer vesicles based on coordination compounds. Herein, we show that nano onions are obtained by aggregation of an amphiphilic $\text{Au}_3(\text{pyrazolate})_3$ complex in aqueous solution. The nanostructures were characterized by cryogenic and transition electron microscopy, dynamic light scattering, and energy-dispersive X-ray analysis. Control experiments with analogous Ag(I) and Cu(I) complexes revealed the importance of Au(I) for the formation of well-defined nano onions. A structurally related Au(I) complex without solubilizing polyethylene glycol side chains was analyzed by single-crystal X-ray diffraction. In the solid state, columns of offset, π -stacked $\text{Au}_3(\text{pyrazolate})_3$ complexes are observed, but short intermolecular Au...Au contacts were not found. Molecular dynamics simulations provided further insights into the aggregation process in aqueous solution, supporting the formation of nano-onion structures through lateral interactions between stacked complexes.

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Introduction

The integration of inorganic coordination compounds into soft matter systems presents an exciting frontier in materials science, offering new avenues for tailoring the mechanical, optical, and functional properties of soft materials.¹ Trinuclear gold(I) complexes with bridging pyrazolate ligands represent interesting building blocks in this context. Initially reported by Bonati in 1974,² these compounds display good chemical stability.^{3,4} Structural modifications can be readily achieved by variation of the pyrazolate ligand.⁴ Moreover, $\text{Au}_3(\text{pyrazolate})_3$ complexes are prone to aggregate *via* aurophilic interactions.^{4,5} This characteristic provides a unique opportunity to control the assembly of such complexes. It is also worth noting that

aggregates of $\text{Au}_3(\text{pyrazolate})_3$ complexes are often luminescent.⁴

Several studies have shown that $\text{Au}_3(\text{pyrazolate})_3$ complexes can form liquid-crystal phases.⁶ Serrano and co-workers have examined pyrazolate ligands with di- and tri-*n*-decyloxyphenyl substituents in position 3 and 5.^{6e,f} The resulting Au trimers were found to form columnar mesophases, some of which displayed good stability at room temperature. Concurrent investigations by Kim *et al.* demonstrated that a single long alkyl chain in 4-position of the pyrazolate ligand can be sufficient for the formation of metallomesogens.^{6d}

The assembly of $\text{Au}_3(\text{pyrazolate})_3$ complexes with functionalized benzyl substituents in 4-position was studied by Aida and co-workers.^{6b,7,8} They observed the formation of luminescent organogels⁷ or fibers,⁸ depending on the nature of the substituents (Fig. 1). Phase transitions of the organogels were found

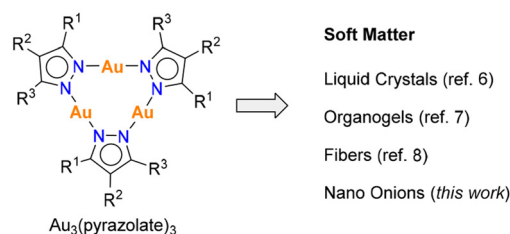


Fig. 1 $\text{Au}_3(\text{pyrazolate})_3$ complexes as building blocks for the construction of soft matter.

^aInstitut des Sciences et Ingénierie Chimiques, École Polytechnique Fédérale de Lausanne (EPFL), 1015 Lausanne, Switzerland. E-mail: kay.severin@epfl.ch

^bDepartment of Applied Science and Technology, Politecnico di Torino, Torino 10129, Italy

^cDepartment of Innovative Technologies, University of Applied Sciences and Arts of Southern Switzerland, Lugano, Viganello 6962, Switzerland

†Computational materials and data pertaining to the study conducted herein (input files, trajectories, etc.) are available at: <https://zenodo.org/doi/10.5281/zenodo.13828919>.

‡Electronic supplementary information (ESI) available: Synthetic procedures, experimental and simulation details. CCDC 2357277. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d4nr03901g>



A

$\text{AuCl}(\text{SMe}_2)$
THF, NEt_3

$\text{Au}_3(\text{L1})_3$ (85%)

PEG³

B

$\text{AuCl}(\text{SMe}_2)$
THF, NEt_3

$\text{Au}_3(\text{L2})_3$ (95%)

A

50 nm

B

100 nm

C

100 nm

D

Gray Value

Distance / nm

$d = 4.0 \pm 0.2$ nm

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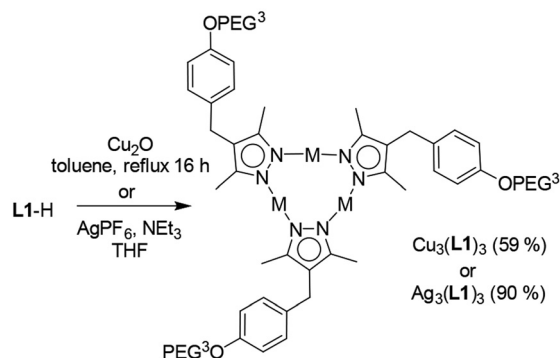
Alongside the spherical aggregates, we were able to detect sheet-like structures (Fig. 2B). The latter are likely precursors for the multilayer vesicles.

CryoEM images of freshly prepared samples showed a larger proportion of sheets and nano onions at incipient stages (see the ESI, Fig. S36†). Time-dependent zeta potential measurements showed an increase of the potential from -25.3 ± 0.8 mV to around -15 mV in the first five hours, followed by small changes within the next 24 hours (see the ESI, Fig. S28†). We assume that the increase of the zeta potential is linked to the transformation of sheet structures into nano onions.

In order to examine if the presence of Au(I) was essential for the aggregation process, we synthesized the analogous copper and silver complexes $\text{Cu}_3(\text{L1})_3$ and $\text{Ag}_3(\text{L1})_3$ (Scheme 2). The syntheses proceeded smoothly, and the formation of the complexes was confirmed by NMR spectroscopy and HRMS.

DLS analyses of aqueous solutions containing $\text{Cu}_3(\text{L1})_3$ or $\text{Ag}_3(\text{L1})_3$ revealed the presence of aggregates with an average size of around 200 nm and 158 nm for $\text{Cu}_3(\text{L1})_3$ and $\text{Ag}_3(\text{L1})_3$, respectively (see the ESI, Fig. S31 and S32†). CryoEM images of $\text{Cu}_3(\text{L1})_3$ samples displayed spherical aggregates with a multi-layered architecture. However, the orientation of the layers was less ordered when compared to what was observed for $\text{Au}_3(\text{L1})_3$, and one could distinguish multiple domains of parallel sheets within one aggregate (Fig. 3A and B). In the case of the Ag(I) complex $\text{Ag}_3(\text{L1})_3$, the structures of the aggregates were even less defined. Sheets could be identified, but only few of them were arranged in a parallel fashion (Fig. 3C and D). The results imply that the self-assembly process is influenced by the metal ion, with the presence of Au(I) being crucial for the formation of well-defined nano onions.

Attempts to analyze $\text{Au}_3(\text{L1})_3$ by single crystal X-ray diffraction (XRD) were unfortunately not successful. While we were able to obtain single crystals of the complex, the quality of the diffraction data was poor. We suspected that the flexible polyethylene glycol side chains in $\text{Au}_3(\text{L1})_3$ were hampering a structural analysis. Therefore, we synthesized the known complex $\text{Au}_3(\text{L3})_3$ ⁸ with simple benzyl side chains in position 4 of the bridging pyrazolate ligands (Scheme 3).



Scheme 2 Synthesis of the complexes $\text{Cu}_3(\text{L1})_3$ and $\text{Ag}_3(\text{L1})_3$.

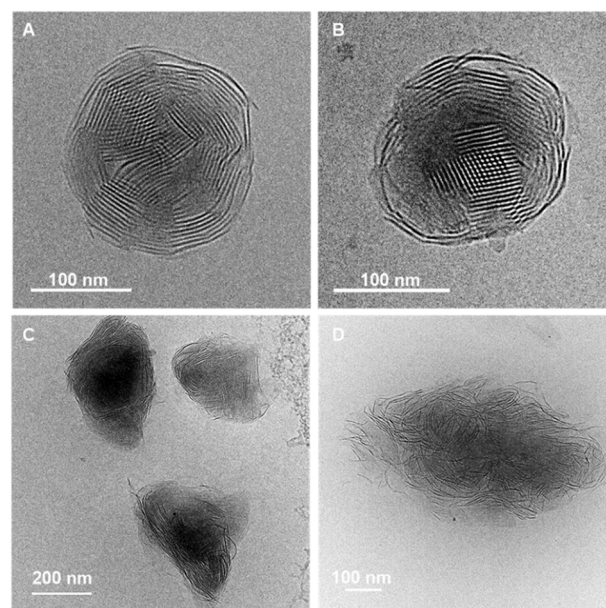
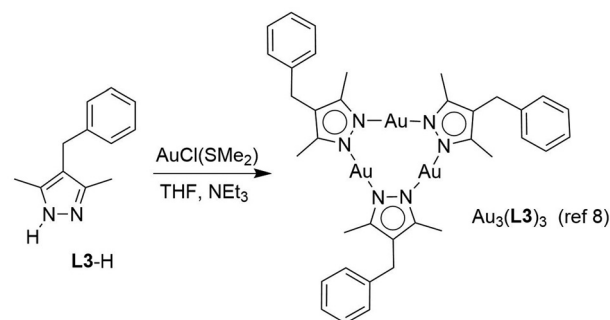


Fig. 3 CryoEM images of the aggregates obtained from $\text{Cu}_3(\text{L1})_3$ (A and B) or from $\text{Ag}_3(\text{L1})_3$ (C and D).



Scheme 3 Synthesis of the complex $\text{Au}_3(\text{L3})_3$.

Single crystals of $\text{Au}_3(\text{L3})_3$ were obtained by slow vapor diffusion of pentane into a CHCl_3 solution of the complex and an XRD analysis was performed (Fig. 4). The central $\text{Au}_3(\text{pyrazolate})_3$ core of complex $\text{Au}_3(\text{L3})_3$ is planar, and the lengths of the Au–N bonds (1.993–2.007 Å) are within the expected range. The benzyl side chains are not in the same plane as the pyrazolate heterocycles, with two benzyl groups pointing ‘up’ and one ‘down’. The individual $\text{Au}_3(\text{pyrazolate})_3$ trimers are arranged in an offset, π -stacked fashion. The coplanar arrangement of the gold trimers gives rise to columns of $\text{Au}_3(\text{pyrazolate})_3$ complexes, which are separated by closely packed benzyl groups. The closest intermolecular contact between neighboring Au atoms is 3.6672(4) Å. This value is smaller than the sum of van der Waals radii for Au(I) (3.80 Å),^{5b} but longer than the intermolecular Au...Au contacts in previously reported $\text{Au}_3(\text{pyrazolate})_3$ complexes (~ 3.4 Å).⁴ It is worth noting that the presence of short Au...Au contacts in $\text{Au}_3(\text{pyrazolate})_3$ complexes is often associated with solid-state



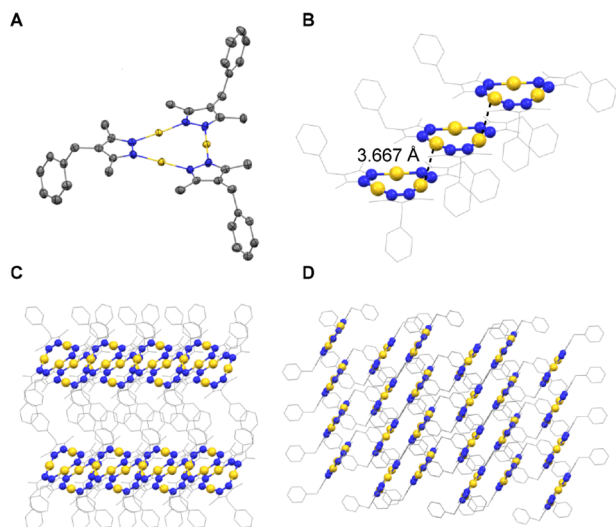


Fig. 4 Molecular structure of $\text{Au}_3(\text{L3})_3$ in the crystal with ellipsoids at 50% probability (A). Packing of adjacent $\text{Au}_3(\text{L3})_3$ complexes (B). View of the packing inside the crystal (C and D). Color code: Au: yellow, N: blue, C: grey. Hydrogen atoms are omitted for clarity.

luminescence.⁴ Solid $\text{Au}_3(\text{L3})_3$, as well as the parent complex $\text{Au}_3(\text{L1})_3$, were found to be non-emissive.

The structure of $\text{Au}_3(\text{L3})_3$ provides hints about possible arrangements of the gold trimers in the nano onions. The high-contrast areas are expected to contain the gold complexes. A gold-rich layer could form by a parallel arrangement of columns made from π -stacked $\text{Au}_3(\text{pyrazolate})_3$ complexes. The interface with the water-rich layer would then be provided by the PEG side chains. The nano onions were found to be non-emissive. Therefore, we assume that close $\text{Au}\cdots\text{Au}$ contacts are not present in the aggregates. The crystal structure of $\text{Au}_3(\text{L3})_3$ provides evidence that stacks of $\text{Au}_3(\text{pyrazolate})_3$ complexes can form without strong aurophilic interactions. One should note that in aqueous solution, the interaction between the gold trimers is expected to be reinforced by hydrophobic interactions.

To assess the potential of nano onions for the uptake of lipophilic molecules, we conducted uptake experiments using Nile Red. Nile Red's solvatochromic properties allow for easy monitoring of encapsulation in the nano onions. In water, the poorly soluble dye exhibited weak emission with a maximum at 660 nm (see ESI, Fig. S38[†]). Once encapsulated within the nano onions, the emission maximum shifted to 636 nm and its intensity significantly increased, indicating successful encapsulation.

To complement the experimental observations, MD simulations were employed to investigate the self-assembly of $\text{Au}_3(\text{L1})_3$ complexes at the molecular level (for details, see the ESI,† section 9).¹⁹ A coarse-grained (CG) model was developed based on an all-atom (AA) representation of the $\text{Au}_3(\text{L1})_3$ complex to enable simulations over larger time and length scales (Fig. 5A). Initial simulations were conducted with 1000 monomers randomly dispersed in the simulation box,

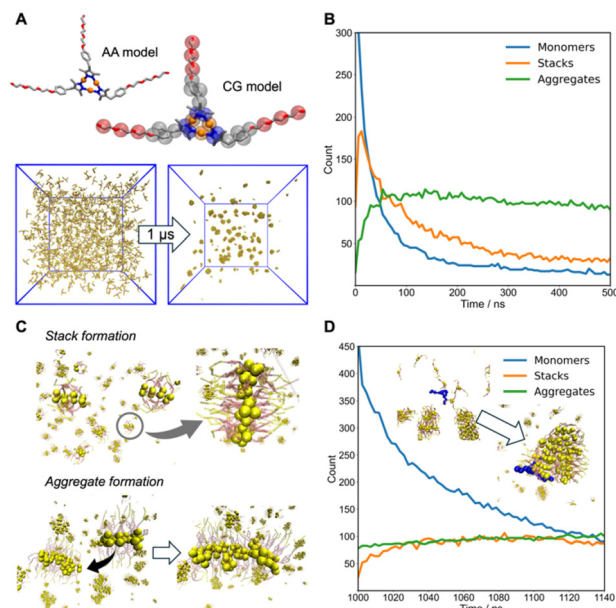


Fig. 5 All-atom (AA) and coarse-grained (CG) models of the $\text{Au}_3(\text{L1})_3$ complex used in the simulations (A). Evolution of assemblies over time in a 1 μs (CG time) simulation (B). Stack formation and aggregation (C). Behavior of assemblies after the introduction of 500 additional monomers (D).

allowing us to monitor their dynamic assembly over 1 μs (CG time). We classified the $\text{Au}_3(\text{L1})_3$ complexes into three categories based on the spatial arrangement and orientation of their cores: monomers (isolated molecules), stacks (clusters of monomers whose Au atom planes are aligned and in sufficient proximity for stacking interactions), and aggregates (clusters that can contain multiple stacks or non-stacked monomers). This classification allowed us to track the dynamic behaviour of these assemblies over time (Fig. 5B), capturing the rapid decrease in the number of monomers, which assemble into stacks (Fig. 5C) within the first 50 ns (CG time). Beyond this point, the number of stacks gradually decreases over time, as these assemble into larger aggregates, primarily through lateral interactions between the aliphatic chains (Fig. 5C). This observation aligns closely with the XRD analysis of $\text{Au}_3(\text{L3})_3$, which revealed columns of offset π -stacked trimers in the solid state. Simulations confirm that similar stacking occurs in aqueous solution, where stacks of 2–4 monomers form. The stacks in solution remain self-limited in size and further assemble into larger aggregates through lateral interactions rather than forming extended fibers. This suggests that the flexible PEG chains in $\text{Au}_3(\text{L1})_3$ play a key role in limiting the stack size and promoting lateral growth into larger aggregates, as observed in the nano onion structures. After 1 μs of simulation, an additional 500 monomers were introduced to the system. Fig. 5D shows a continued, rapid decline in the monomer count, while the number of stacks and aggregates remains relatively stable. The stability in the number of stacks implies that their formation becomes limited once larger aggregates are present, and the primary mode of monomer



incorporation shifts toward expanding the aggregates through lateral interactions between the aliphatic chains, contributing to their growth in size rather than in number. This lateral growth mechanism indicates that, over extended time scales, the aggregates can develop into sheet-like structures, consistent with the experimental cryoEM observation of layered nano onions.

Conclusions

We have synthesized an $\text{Au}_3(\text{pyrazolate})_3$ complex with three PEG side chains. In aqueous solution, this complex was found to form multilayer vesicles with an onion-like architecture. Thus far, there are only limited reports about the formation of nano onions from metalloamphiphiles.^{13–15} To the best of our knowledge, the use of $\text{Au}(\text{I})$ complexes is unprecedented in this context.

Control experiments with analogous $\text{Ag}(\text{I})$ and $\text{Cu}(\text{I})$ complexes highlighted the crucial role of $\text{Au}(\text{I})$ in the formation of well-defined nano onions. A structurally related $\text{Au}(\text{I})$ complex, lacking solubilizing PEG chains, was examined by single-crystal XRD. In the solid state, columns of offset, π -stacked $\text{Au}_3(\text{pyrazolate})_3$ complexes were observed. A similar stacking of gold-trimers could provide the molecular basis for the formation of nano onions. To support this hypothesis, MD simulations were conducted. The simulations revealed that $\text{Au}_3(\text{L1})_3$ complexes first assemble into small stacks, which then aggregate *via* lateral interactions between the side chains, suggesting a lateral growth mechanism in the formation of the larger multilayer vesicles.

Overall, our results provide further evidence for the potential of $\text{Au}_3(\text{pyrazolate})_3$ complexes as building blocks in molecular nanoscience.

Author contributions

A.B.S. and K.S. initiated the study, A.B.S. and D.D. performed the experiments and analyzed the data, L.L., M.D.P. and G.M. P. performed the modeling work, F.F.-T. collected and processed the X-ray data, and A.B.S. and K.S. co-wrote the manuscript. All authors discussed the results and commented on the manuscript.

Data availability

Synthetic procedures, experimental and simulation details. Computational materials and data pertaining to the study conducted herein (input files, trajectories, *etc.*) are available at: <https://zenodo.org/records/13828920>.

Conflicts of interest

There are no conflicts to declare.

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