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Rh(III)-catalyzed building up of used heterocyclic cations:
facile access to white-light-emitting materials

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Rh(III)-catalyzed building up of used heterocyclic cations: facile access to white-light-emitting materials†

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The first example of rhodium-catalyzed nondirected C–H activation/annulation reactions for the construction of fused heterocyclic cations is reported herein with excellent regioselectivity. Deuterium-labeling experiments indicated that the C(sp³)–H bond cleavage of the *N*-methyl group might be the rate-limiting step during the reaction process. This protocol provides an opportunity to rapidly access highly π -conjugated fused heterocyclic cations, which opens up a new avenue for efficient screening of single-molecule white-light-emitting materials, pure red-light-emitting materials, and π -conjugated radical materials. Importantly, novel white-light-emitting materials exhibited distinct anti-Kasha dual-emission and could rapidly be fabricated into robust organic and low-cost white light-emitting diodes.

Introduction

Fused heterocyclic cations have attracted considerable attention in scientific and engineering arenas due to their charming thermal, mechanical, optical, magnetic, electronic, and electrochemical properties.^{1–17} As such, fused heterocyclic cation materials have been widely applied in fluorescent bioimaging reagents, DNA intercalators, photosensitizers, optical devices, and electronic devices.^{1–17} Traditional synthetic methods for fused heterocyclic cations often suffer from tedious multiple-step syntheses or the use of toxic or environmentally harmful chemicals, and are thus environmentally and economically unbeneficial.^{7,18–21} Doubtlessly, the development of more atom- and step-economic syntheses of fused heterocyclic cation materials is highly desirable.

In the past decade, transition metal-catalyzed C–H functionalization has emerged as a powerful synthetic method and has aroused great interest in chemistry research because this strategy usually uses lower toxic reagents and can offer more atom- and step-economic syntheses of organic functionalization molecules than traditional synthetic methods, and could even enable unprecedented synthetic transformations.^{22–30} In

recent years, transition metal-catalyzed C–H functionalization has been developed as a straightforward and efficient approach to constructing fused heterocyclic skeletons (Fig. 1).^{25–50} However, rhodium-catalyzed nondirected C–H activation/annulation to build fused heterocyclic cations has not been reported and remains a challenge since the regioselectivity is always an issue in the reactions with multiple analogous C–H bonds of (hetero)arenes. Here, we have deliberately designed substrates that enable straightforward and effective construction of pyridophenothiazin/phenoxazin/phenoselenazin/phenazin-12-iums *via* the rhodium-catalyzed nondirected C–H activation/annulation (Fig. 1).

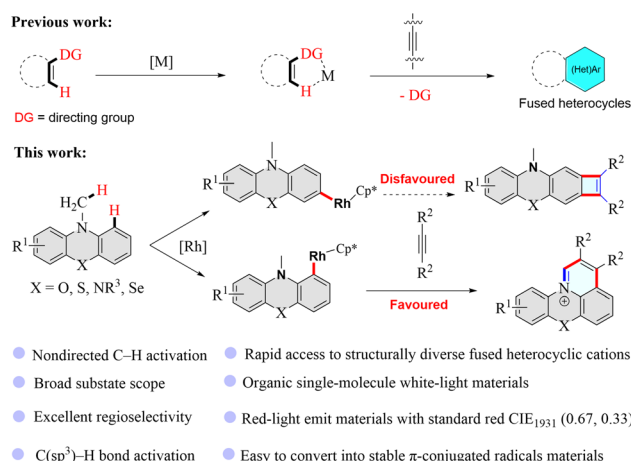


Fig. 1 Metal-catalyzed C–H activation/annulation with alkynes to construct π -conjugated heterocycles. DG = directing group.

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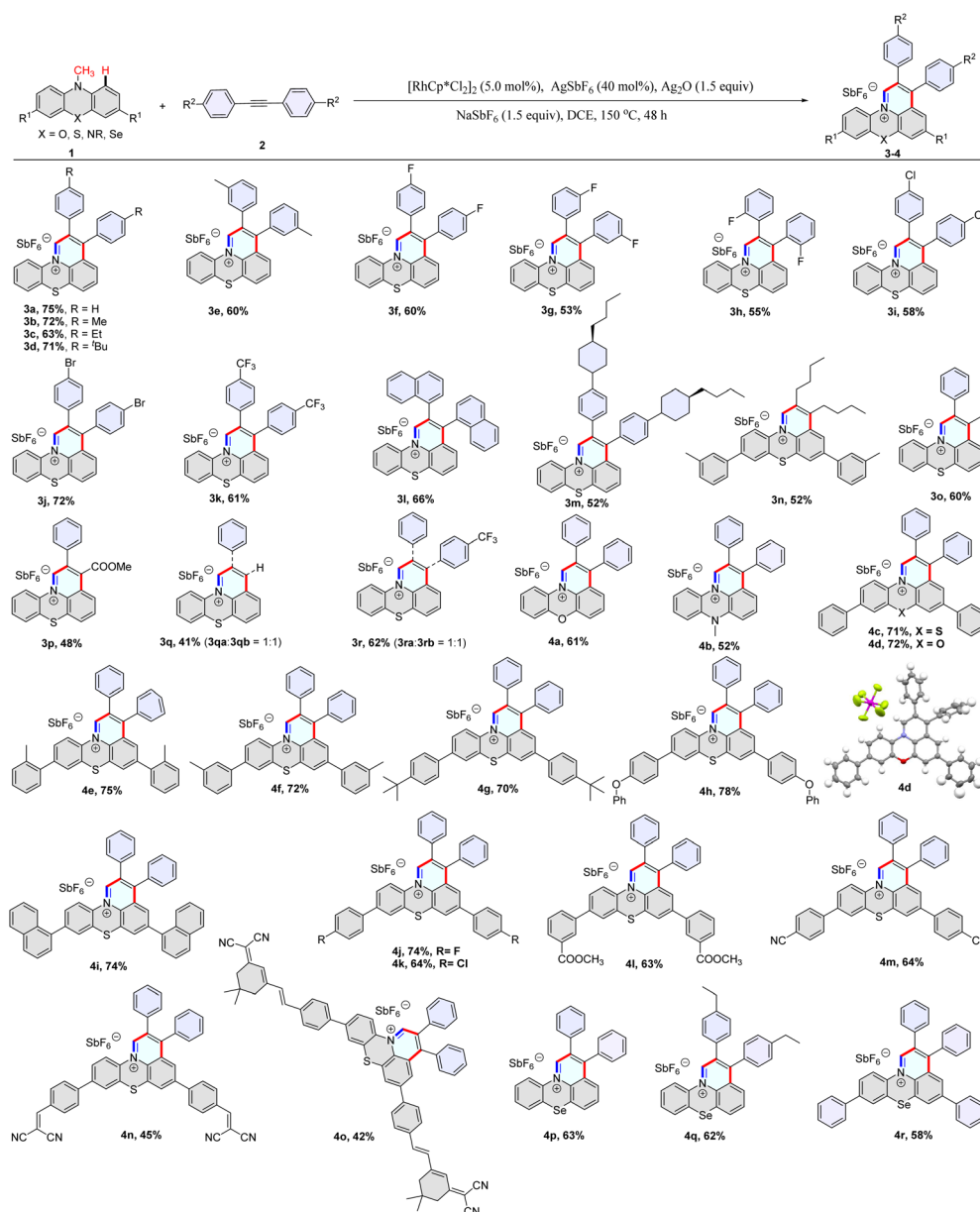
Results and discussion

We set out our investigation by using 10-methyl-10*H*-phenothiazine (**1a**) and 1,2-diphenylethyne (**2a**) as model substrates in the presence of $[\text{RhCp}^*\text{Cl}_2]_2/\text{AgSbF}_6$, Ag_2O , and NaSbF_6 in DCE at 120 °C for 48 hours, which afforded 2,3-diphenylpyrido[3,2,1-*kl*]phenothiazin-12-ium **3a** in 44% yield (Table S1,† entry 1). After the optimization of the reaction conditions [see Tables S1 and S2 in the ESI†], we were able to improve the yield to 75% under reaction conditions comprising $[\text{RhCp}^*\text{Cl}_2]_2$ (5 mol%), AgSbF_6 (40 mol%), Ag_2O (1.5 equiv.) and NaSbF_6 (1.5 equiv.) in DCE at 150 °C under N_2 for 48 h (Table S1,† entry 12). It is

noteworthy that this C–H activation/annulation reaction showed high regioselectivity at the methyl and C₁ positions, and no C₃-alkenylation product was detected.

Subsequently, a library of structurally diverse fused heterocyclic cations was synthesized using phenothiazines, phenoxazines, phenoselenazines, and phenazine compounds with alkynes under the optimized reaction conditions (Table 1, **3a–3n** and **4a–4r**). Diaryl alkynes with electron-neutral, electron-rich, and electron-deficient groups on the phenyl ring were successfully annulated and gave moderate to good yields of desired products (**3a–3m**). Dialkyl acetylene was also annulated with 10-methyl-10*H*-phenothiazine and provided **3n** in 52%

Table 1 Rhodium-catalyzed construction of fused heterocyclic cations^a



^a Reaction conditions: **1** (0.1 mmol), **2** (0.15 mmol), $[\text{Cp}^*\text{RhCl}_2]_2$ (5.0 mol%), AgSbF_6 (40 mol%), Ag_2O (1.5 equiv.), NaSbF_6 (1.5 equiv.), DCE (1 mL), at 150 °C, 48 h under a N_2 atmosphere.

yield (Table 1). Unsymmetrical alkynes were used to annulate with 10-methyl-10*H*-phenothiazine, giving moderate yields (Table 1, **3o–3r**). The $\text{PhC}\equiv\text{CH}$ and 1-(phenylethynyl)-4-(trifluoromethyl)benzene reacted with compound **1a** to give two regioisomeric products with a ratio of 1 : 1 (Table 1, **3q–3r**). To our delight, prop-1-yn-1-ylbenzene and methyl 3-phenylpropiolate underwent annulation with **1a** to generate products **3o** and **3p** with excellent regioselectivity (Table 1, Fig. S20 and 21†). It should be mentioned that various functional groups such as fluorides, chloride, bromide, phenoxy, cyano, and carboxylic ester were well tolerated under the reaction conditions (**3f–3j**, **4h**, **4j–4n**), which provides an ample opportunity for further manipulation of the initial products. Derivatives of dicyanoisophorone (DCI) tend to possess ultra-fast intramolecular charge transfer and were used as fluorescence probes with large Stokes shifts. Notably, the DCI unit was well tolerated in the reaction, and the corresponding product was isolated in moderate yield (**4o**). Furthermore, the structure of **4d** was confirmed by single-crystal X-ray analysis (Fig. S18†).

To gain mechanistic insight into the C–H activation/annulation process, a series of control experiments were performed. In the H/D exchange experiment, 10% and 67% of deuterium incorporation were found at the C_1 and C_3 -positions of 10-methyl-10*H*-phenothiazine (**1a**) in the presence of CD_3OD under varied reaction times (Scheme S2 and Fig. S1†), whereas there is no H/D exchange at the methyl position of 10-methyl-10*H*-phenothiazine (**1a**). These results indicated that the C–H bond activation step might be reversible, the C–H bonds at the C_1 and C_3 -positions of **1a** have similar reactivity, and the reaction might begin with the metalation of the C_1 or C_3 -position. Subsequently, the kinetic isotope effect (KIE) was investigated, and two different KIE values were obtained (Schemes S3 and S4, and Section S5 of the ESI†). A significant KIE value ($k_{\text{H}}/k_{\text{D}} = 2.4$) between **1a** and $[\text{D}]_3\text{-1a}$ with **2a** was observed, suggesting that the $\text{C}(\text{sp}^3)\text{–H}$ bond cleavage of the methyl group of **1a** might be involved in the rate-limiting step. Furthermore, the radical scavenger 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) was added to the reaction and showed a negligible effect, which ruled out a radical pathway (Table S1,† entry 18).

Based on these results and previous reports,^{35–50} a plausible catalytic cycle is proposed (Fig. 2). First, catalyst $[\text{Cp}^*\text{RhCl}_2]_2$ reacts with AgSbF_6 to offer the highly electrophilic $[\text{Rh}^{\text{III}}\text{Cp}^*]$ species. In path A, the substrate **1a** reacts with $[\text{Rh}^{\text{III}}\text{Cp}^*]$ to form the intermediate **IA** after a reversible C–H bond cleavage process at the C_1 -position. Next, the alkyne **2a** coordinates with the rhodium intermediate **IA**, and subsequently inserts into the Rh–C bond to provide the intermediate **IIA**. Then, the intermediate **IIA** undergoes rhodium-catalyzed $\text{C}(\text{sp}^3)\text{–H}$ bond activation to form the seven-membered rhodacycle intermediate **IIIA**. Finally, dehydrogenation of **IIIA** afforded the intermediate **IV**, followed by a reductive elimination process to release the desired product **3a**. The formed rhodium(I) species is oxidized to the $[\text{Rh}^{\text{III}}\text{Cp}^*]$ by Ag_2O . Furthermore, the generation of all the intermediates **IA–IV** was supported by high-resolution electrospray ionization mass spectrometry (HRMS) analysis (Fig. S3–S6†). In path B, substrate **1a** reacts with $[\text{Rh}^{\text{III}}\text{Cp}^*]$ to generate the intermediate **IB** after a reversible C–H bond cleavage

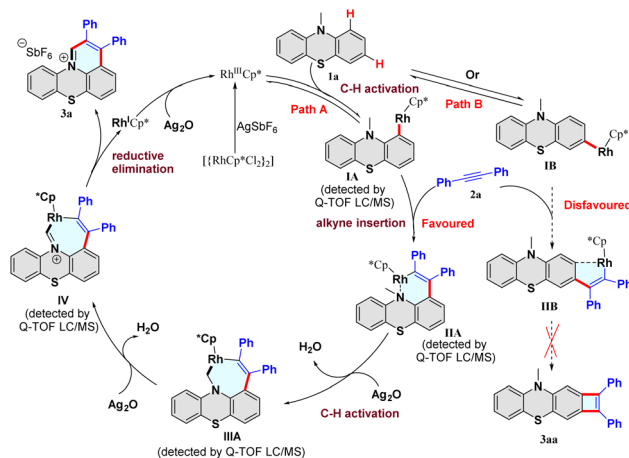


Fig. 2 Proposed mechanism.

process at the C_3 -position. Then, alkyne inserts into the **IB** to form the intermediate **IIB**. However, forming the rigid four-membered product **3aa** is difficult due to its thermodynamic instability. Notably, we did not observe the four-membered compound **3aa** in the reaction. Consequently, path B is disfavoured and could be excluded from the reaction. Hence, the C–H activation/annulation reaction showed high regioselectivity at the methyl and C_1 positions, and no C_3 -position product was detected.

Organic fluorescence materials, especially single-molecule-white-light emitting materials, and red-light emitting materials, have attracted considerable attention in recent years due to their applications in organic light-emitting diodes, field-effect transistors, security systems, fluorescent markers, bioprobes, etc.^{51–61} However, discovering single-molecule white-light materials is a challenging task because white-light-emission involves a broad emission covering the whole visible range (400–700 nm), which is difficult for ordinary organic fluorescent molecules due to their intrinsic limitation of photophysical properties. Therefore, organic single-molecule white-light materials are scarce, and only a few have been developed so far. Compounds with anti-Kasha systems have great potential for preparing single-molecular white-light-emitting materials because of their double emissions with a relatively short wavelength (blue-light component) from a high-lying excited state emission ($S_n \rightarrow S_0$, $n \geq 2$), and a relatively long wavelength (orange-light component) from a low-lying excited state emission ($S_1 \rightarrow S_0$).^{56–58}

We further measured the photophysical properties of the constructed library of pyrido-phenothiazin/phenoxazin/phenoselenazin/phenazin-12-iums, and the corresponding absorption and emission maxima are summarized in Table S3 and Fig. S8–11.† To our delight, these compounds exhibited distinct dual emissions with a relatively short blue emission (anti-Kasha emission) and a relatively long orange emission wavelength (Kasha emission), which thus efficiently covers the whole visible range (400–700 nm). Fortunately, nine fused heterocyclic cations **3d**, **3e**, **3i**, **4f**, **4g**, **4h**, **4p**, **4q**, and **4r**





Fig. 3 Fluorescence properties. Fluorescence emission spectra of **3d**, **3e**, **3i**, **4f–4h**, **4p–4r** in CH_2Cl_2 and **4f*** in the PMMA film.

displayed white light emissions in dichloromethane with Commission Internationale de l'Eclairage (CIE) coordinates of **3d** (0.28, 0.20), **3e** (0.30, 0.29), **3i** (0.24, 0.21), **4f** (0.33, 0.32), **4g** (0.32, 0.35), **4h** (0.32, 0.32), **4p** (0.25, 0.22), **4q** (0.40, 0.32), and **4r** (0.30, 0.23), respectively (Fig. 3, 4 and S12[†]). Compound **4f** displayed dual emission with blue emission at approximately 470 nm and orange emission at 610 nm with a fluorescence quantum yield of 10% in dichloromethane (Fig. 3). It is worth pointing out that **4f** emits bright white light with CIE coordinates of (0.33, 0.32), very close to those of pure white light (CIE: 0.33, 0.33). Moreover, **4f** exhibits white-light emission in polymethyl methacrylate (PMMA) film ($c = 0.005$ wt%) with CIE coordinates of (0.31, 0.32) and a fluorescence quantum yield of

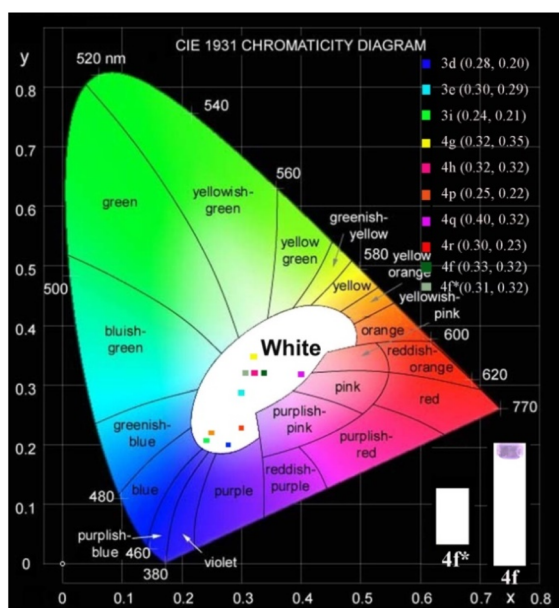


Fig. 4 Chromatic coordinates. Commission Internationale de l'Eclairage (CIE) coordinates of **3d** (0.28, 0.20), **3e** (0.30, 0.29), **3i** (0.24, 0.21), **4g** (0.32, 0.35), **4h** (0.32, 0.32), **4p** (0.25, 0.22), **4q** (0.40, 0.32), and **4r** (0.30, 0.23) in CH_2Cl_2 , respectively.

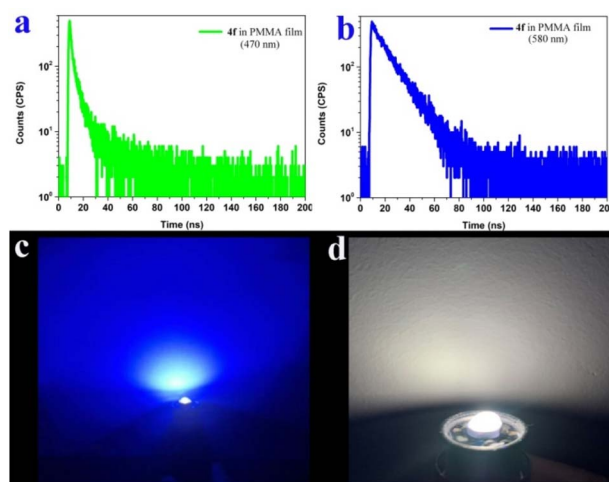


Fig. 5 Transient emission spectra and luminescence image. (a) and (b) Transient emission spectra of **4f** in the PMMA film at room temperature. (c) Luminescence image of a commercially available UV lamp. (d) Luminescence image of a UV chip coated with the **4f** film (0.005 wt% in PMMA) when the LED is turned on.

20% (Fig. 3 and 4). In addition, the excited-state lifetimes of **4f** in the dichloromethane solution and in the PMMA film were tested at room temperature (Table S4 and Fig. S14[†]). It is observed that the radiative excitons in **4f** are short-lived components with nanosecond order (Fig. 5a and b), and no long-lifetime fluorescence components exist, which ruled out the delayed fluorescence or phosphorescence emission. The thermogravimetric analyzer measurement demonstrates that **4f** is thermally stable (Fig. S15[†]).

White light-emitting diodes (LEDs) have received much attention and have been widely used in solid-state lighting displays and illumination.^{61–63} Currently, most commercial white light-emitting diodes are made by coating green or inorganic phosphors on LED chips.^{61–63} To further prove the potential application of our white-light materials, we coated the **4f** film (0.005 wt% in PMMA) on a commercially available UV chip, and bright white light could be obtained when the LED is turned on (Fig. 5c and d). This work unlocks an opportunity to rapidly fabricate robust organic and low-cost white LEDs.

The excitation-wavelength-dependent fluorescence experiments and theoretical calculations further demonstrated that **4p** and **4f** possess an anti-Kasha dual-emission character.^{56–58,64,65} The relative intensity of dual emission highly depends on the excitation wavelength. For **4p**, lower energy excitations result in a slight red-shift at shorter wavelengths, and the intensity of longer wavelengths decreases and then increases (Fig. 6a). For **4f**, higher energy excitations resulted in enhanced emissions at shorter wavelengths (Fig. 6c). This experiment implied that the two fluorescence bands were from different excited states: the white light emission of **4f** with the blue-light component from a high-lying excited state emission ($S_2 \rightarrow S_0$), and the orange-light component from a low-lying excited state emission ($S_1 \rightarrow S_0$) (Fig. 6c). The internal conversion (IC) from the S_2 to the S_1 state is comparatively slow

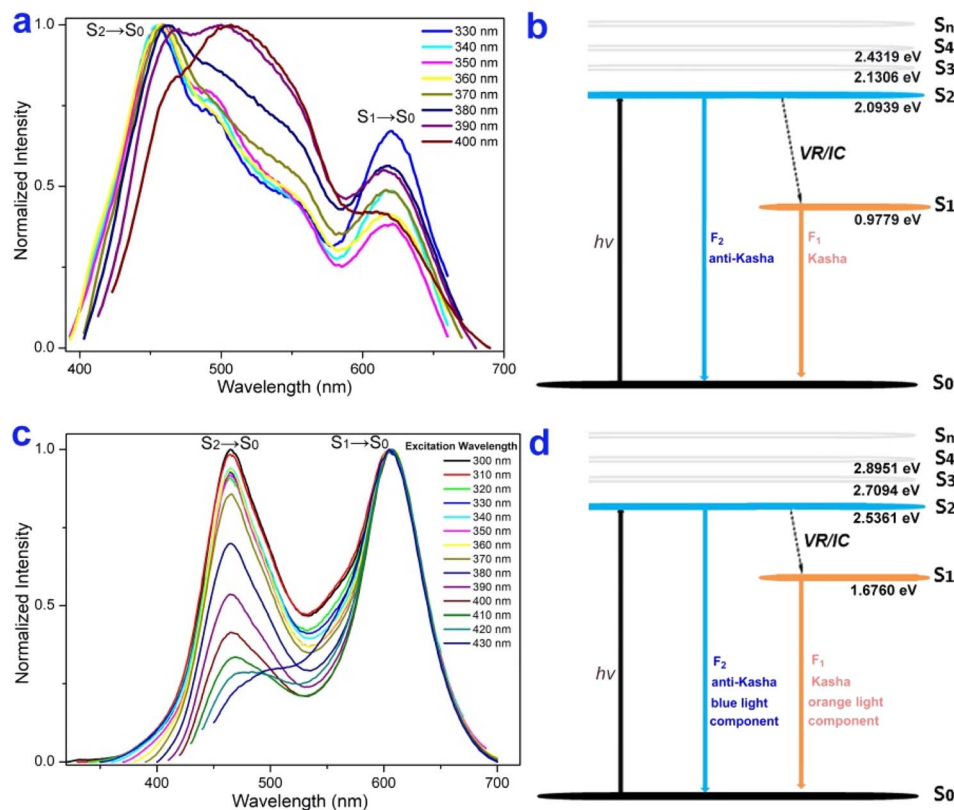


Fig. 6 Anti-Kasha dual-emission character. (a), and (c) Excitation-wavelength-dependent fluorescence spectra of **4p** and **4f** (1.0×10^{-6} M). (b), and (d) Jablonski diagram illustrating the anti-Kasha dual-emission mechanism of **4p** and **4f**.

due to the large energy gap $\Delta E(S_2 \rightarrow S_1)$ value (**4p**: 1.116 eV, **4f**: 0.6633 eV), and as a result, S_2 fluorescence can compete favorably with IC (Fig. 6b and d).

Furthermore, compounds **3d** and **4h** in solid films displayed only low energy emission (Kasha emission) due to strong π - π interactions. Compound **3d** showed red-light emission at 622 nm with CIE coordinates of (0.66, 0.33) (Fig. 7a and S13[†]). The **4h** exhibited pure red-light emission at 636 nm with a fluorescence quantum yield of 5% and CIE coordinates of (0.67, 0.33) (Fig. 7b), which is consistent with the National Television System Committee (NTSC) standard red CIE coordinates of (0.67, 0.33). The thermogravimetric analyzer measurement showed that **4h** is thermally stable (Fig. S15[†]).

π -conjugated radicals have recently attracted immense attention because of their excellent optical, electronic, and unique magnetic properties, and thus have a range of potential applications in nonlinear optics, spintronics, organic electronics, and energy storage devices.^{66–68} Herein, our newly developed fused heterocyclic cations are easy to convert into stable π -conjugated radicals. As illustrative examples, the cations **4c** and **4h** could be easily reduced to π -conjugated radicals by using NaI in an acetonitrile solution (see Section IX of the ESI[†]). The corresponding reduced products are electron paramagnetic resonance (EPR) active and offer g values of 2.0025 and 2.0023, respectively, which are almost equal to the g factor for a free electron that of the typical organic radicals ($g = 2.0023$) (Fig. 7c and d).

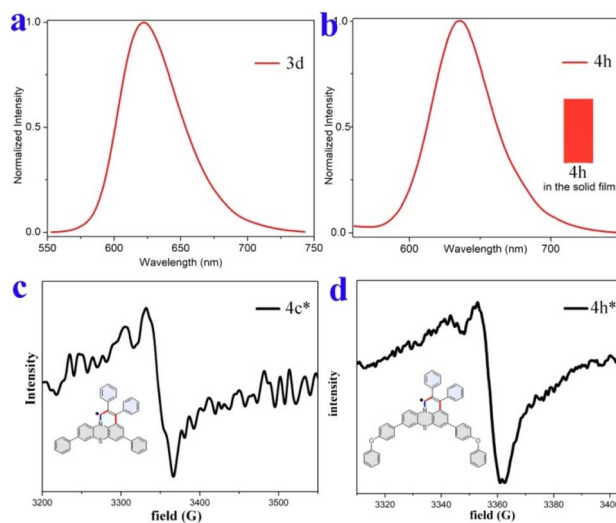


Fig. 7 Fluorescence spectra and EPR spectra. (a) Fluorescence emission spectrum of **3d** in the solid film. (b) Fluorescence image and emission spectrum of **4h** in the solid film. (c) and (d) EPR spectra of reduced **4c** and **4h** in the solid state at room temperature.

Conclusions

In conclusion, rhodium-catalyzed nondirected C–H activation/annulation has been developed to build a library of structurally diverse pyrido-phenothiazin/phenoxazin/phenoselenazin/

phenazin-12-iums for the first time. The protocol shows excellent regioselectivity and broad substrate scope. Deuterium-labeling experiments have indicated that the C(sp³)-H bond cleavage of the *N*-methyl group might be involved in the rate-limiting step in the catalytic process. Furthermore, this approach provides straightforward access to highly π -conjugated fused heterocyclic cations, which opens up a new avenue for rapid screening of single-molecular white-light-emitting and pure red-light-emitting materials. More importantly, novel π -conjugated fused heterocyclic cations displayed anti-Kasha doublet emissions and were used as an excellent white light source to rapidly fabricate robust organic and low-cost white LEDs. Additionally, these π -conjugated fused heterocyclic cations could readily be converted into stable π -conjugated radical materials. The highly efficient gateway toward single-molecular white-light-emitting and pure red-light emitting materials and π -conjugated radical materials developed herein has highlighted the exceptional charm of C-H functionalization for streamlining the lead-optimization phase in the discovery of organic functional materials.

Data availability

All experimental data associated with this work are provided in the ESI.†

Author contributions

J. Z., T. S., K. W., R. H. and C. Z. performed the experiments and analyzed the data. H. G. and B. L. designed and directed the project and wrote the manuscript. All authors contributed to discussions.

Conflicts of interest

There are no conflicts to declare.

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