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# Experimental study on the treatment of acid mine drainage by modified corncob fixed SRB sludge particles

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In view of the characteristics of high content of  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  in acid mine drainage and low pH value, based on the microbial immobilization technology, the single factor test and the orthogonal test were set respectively to determine the optimum alkaline  $\text{H}_2\text{O}_2$  modification conditions for corncob. Then combining with sulfate reducing bacteria sludge, the modified corncob immobilized SRB sludge particles were prepared to treat acid mine drainage. On this basis, three dynamic column test models, including Column 1 without corncob particles, Column 2 with unmodified corncob particles, and Column 3 with modified corncob particles, were constructed. Through dynamic experiments, the three dynamic columns were compared to study the efficacy of AMD and their ability to resist changes in pollution load. The results of the orthogonal experiment showed that: when the corncob modified time was 24 h, the concentration of NaOH was 6% and the concentration of  $\text{H}_2\text{O}_2$  was 1.5%, the prepared immobilized particles performed best. The results of the dynamic test showed that the treatment effect of Column 3 on AMD was better than that of Column 1 and 2. In the dynamic tests before and after the increase of pollution load, the highest removal percentages of  $\text{SO}_4^{2-}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$  in Column 3 were 72.65%, 56.72%, 62.47% and 62.58%, 30.07%, 46.87% respectively, the average COD emission was  $234 \text{ mg L}^{-1}$  and  $102.75 \text{ mg L}^{-1}$ , the effluent pH value was 6.96 and 6.65. In the dynamic tests before and after the increase of pollution load, the highest removal percentages of  $\text{SO}_4^{2-}$ ,  $\text{Mn}^{2+}$ ,  $\text{Fe}^{2+}$  in Column 2 were 52.94%, 46.93%, 72.55% and 48.92%, 26.43%, 43.23% respectively, the average COD emission was  $508.14 \text{ mg L}^{-1}$  and  $152.88 \text{ mg L}^{-1}$ , the effluent pH value was 6.56 and 6.36. The high COD value of Column 2 is due to the organic matter leakage and poor metabolic activity of SRB contained in immobilized particles. Therefore, it indicated that Column 3 could better treat pollutants and resist changes of pollution load.

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## Introduction

Acid Mine Drainage (AMD) has the pollution characteristics of acidic pH, high heavy metal ions and high sulfate concentration. Arbitrary disposal of AMD will cause great pressure on water resources, and will also cause extremely serious consequences for the ecological environment.<sup>1–3</sup> In view of the pollution problem caused by AMD, the relevant scholars have carried out a lot of research work on governance and restoration. At present, the basic methods of treatment are the neutralization method,<sup>4</sup> the artificial wetland method<sup>5</sup> and the microbial method.<sup>6</sup> The application of the neutralization method and the artificial wetland method has been limited due to the production of a large amount of solid waste or poisonous and harmful gases.<sup>7,8</sup> The Sulfate Reducing Bacteria (SRB)

immobilization technology overcomes the limitations of traditional repair technology. Allowing the SRB and nutrients to be highly aggregated, the SRB immobilization technology has the advantages of strong adaptability, low processing cost, less pollution caused by secondary pollution, recyclable element metal ions and simple operation,<sup>9</sup> so it has attracted much attention from academia, at present.

The research of SRB immobilization technology focuses on immobilized materials. Zhang Mingliang<sup>10</sup> prepared SRB immobilized particles with heavy metal resistance from polyvinyl alcohol (PVA), sodium alginate, iron powder and silica sand. The particles have a good removal effect on Fe, Cu, Zn and Cd. However, the particle still uses the conventional carbon source lactic acid as a carbon source for the growth of the immobilized microorganism, so it cannot effectively release the carbon source slowly. Therefore, selection and optimization of slow-release organic carbon source materials have become a difficult point in the research of SRB immobilization methods. Microbial growth requires a carbon source, and it is susceptible to external conditions.<sup>11</sup> In order to get a better treatment effect,

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cohesive carbon source is an ideal method. Jiang Fu<sup>12</sup> prepared SRB sludge immobilized particles with corncob as the carbon source to deal with AMD. According to the results, the removal percentages of  $\text{SO}_4^{2-}$  and  $\text{Mn}^{2+}$  in AMD were 94.13% and 84.39% respectively, and the pH value of effluent was 7.03. The results showed that this method had some effect on AMD treating. But the COD value of the effluent after treatment was still high, indicating that the corncob could not be fully utilized, and the corncob did not achieve the desirable sustained release effect.

Corn cob is mainly composed of cellulose, hemicellulose, lignin, ash and pectin.<sup>13</sup> Since the microbes can use the soluble carbon source and the easily decomposable substance in the corncob at the beginning of the reaction, the microbial metabolic rate is higher during the time. But hemicellulose is easy to hydrolyze, and the amount of carbohydrate released is large and concentrated, so it is unfavorable to slow release control of biomass carbon source.<sup>14</sup> The carbon source required for the microorganism at the later stage of the reaction must be obtained by decomposing insoluble cellulose. But because of the influence of lattice structure, and the cross-linking effect of cellulose, hemicellulose and lignin, which is not conducive to the decomposition of carbon source,<sup>15</sup> of the carbon supply is inappropriate in the later. Therefore, SRB metabolic rate decreased, and it is not conducive to treatment of AMD. Relevant scholars transformed organic matter components such as cellulose, hemicellulose and lignin in corncob with modified methods to improve the utilization rate of organic matter in corncob and control the release of carbon source. Until now they have achieved good results. R. L. Tseng<sup>16</sup> studied the adsorption properties of phenol and methylene blue by modifying the corncob with KOH solution as an activator. The results show that the regular surface shape of corncob honeycomb was favorable for the formation of micropores and increased surface area. So the performance of pollutants adsorption was significantly enhanced. Zhao Wenli<sup>17</sup> modified corncob by combining alkaline hydrogen peroxide (containing 1.5% of  $\text{H}_2\text{O}_2$  NaOH solution) and UV irradiation to investigate the carbon release from corncob, denitrification and microbial attachment. The results showed that the availability and denitrification efficiency of modified corncob were improved remarkably, and the removal percentage of nitrate still kept more than 90% after 41 days of static denitrification.

Based on the above, this paper proposed to modify corncob with alkaline  $\text{H}_2\text{O}_2$  to obtain an effective slow release carbon source. Alkaline  $\text{H}_2\text{O}_2$  modified corncob material was prepared by single factor test and orthogonal test. The modified corncob fixed SRB sludge particles were prepared by using modified corncob and SRB as materials. The above immobilized particles were used to treat AMD. The optimal modification conditions of modified corncob material required for the preparation of immobilized particles was determined by the changes of COD release,  $\text{SO}_4^{2-}$  removal percentage and  $\text{Fe}^{2+}$  removal percentage in the solution. By modifying the corncob material in the immobilized particles, the treatment efficiency of the immobilized particles is enhanced, the adaptability of the immobilized particles is improved, and the carbon source in the particles can

be released long-term, effectively and slowly. And on this basis, three groups of dynamic columns, containing non-corn cob, unmodified corncob and modified corncob were constructed. The effects of three particle systems on AMD treatment and their impact load resistance were compared and analyzed under different pollution load conditions. Finally, the internal mechanism of pollutant removal by the immobilized particles was further studied by XRD and SEM analysis of the particles before and after the reaction. The study is expected to provide some scientific theoretical basis for the practical application of modified corncob SRB sludge immobilized particles.

## Methods

### Experimental materials and water samples

Modified Starkey medium: 0.5 g  $\text{Na}_2\text{SO}_4$ , 1.0 g  $\text{NH}_4\text{Cl}$ , 0.5 g  $\text{K}_2\text{HPO}_4$ , 0.1 g  $\text{CaCl}_2 \cdot \text{H}_2\text{O}$ , 2.0 g  $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ , 1.2 g  $(\text{NH}_4)_2\text{Fe}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$ , 0.1 g ascorbic acid, 4.0 mL sodium lactate, 1.0 g yeast extract, 1 L distilled water, pH = 7.0, sterilized at 121 °C for 30 min. Among them,  $(\text{NH}_4)_2\text{Fe}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$  and ascorbic acid cannot be sterilized at high temperature, and were sterilized by a 0.22  $\mu\text{m}$  filter membrane.

SRB sludge: the thick active sediment of a river in Fuxin City was taken as seed mud. Then it was inoculated into sterilized modified Starkey medium for anaerobic culture. By analyzing the black sediment amount and  $\text{H}_2\text{S}$  smell in the medium, the SRB with strong activity was enriched and cultured for use.

Corn cob: the corncob in the new local farmland was taken. After drying and pulverizing, the corncob particles with a particle size of about 0.15 mm are sieved for use.

In the single factor and orthogonal experiments, the experimental water samples were simulated water samples. The mass concentrations of characteristic pollutants on  $\text{SO}_4^{2-}$ ,  $\text{Mn}^{2+}$  and  $\text{Fe}^{2+}$  were 816  $\text{mg L}^{-1}$ , 6  $\text{mg L}^{-1}$  and 14  $\text{mg L}^{-1}$ , respectively, and the pH value was 4.0.

In the dynamic column tests, the characteristic pollutants of AMD were simulated at low and high concentrations, which were carried out in two stages of Test I and II, and the concentration was increased on the 8th day. The characteristic pollutant concentration indexes of each stage are shown in Table 1.

### Experimental apparatus and method

Based on single factor experiment, orthogonal experiment, dynamic experiment and reaction kinetics experiment, the modified corncob SRB sludge immobilized particles were prepared by using alkaline  $\text{H}_2\text{O}_2$  modified corncob and SRB sludge as experimental materials. By analyzing the treatment effect of granule on AMD, the method of optimal preparation of modified corncob SRB sludge immobilized particles was determined. The specific experimental methods are as follows.

### Preparation method of modified corncob

5 g of corncobs were placed in different concentration of alkaline  $\text{H}_2\text{O}_2$  solutions at a solid-liquid ratio of 1 : 10 ( $m : V$ , g



$\text{mL}^{-1}$ ), which were modified by the oscillating reaction for a certain time. Then modified corncob material was formed.

### Preparation method of fixed SRB sludge particles

Single factor experiment, orthogonal experiment, dynamic experiment and reaction dynamics experiment all take modified corncob fixed SRB sludge particle as materials, so the preparation methods of immobilized particles are introduced as follows:

Preparation method of modified corncob fixed SRB sludge particles: based on the preliminary results of the research group, the preparation method of modified corncob fixed SRB sludge particles is as follows:<sup>12</sup> the mass fraction of 9% of PVA and 0.5% of sodium alginate were dissolved in distilled water and sealed at room temperature. After 24 hours of full swelling, it was placed in a constant temperature water bath and stirred at 90 °C until no bubbles were present. The modified corncob powder with a mass fraction of 5% was slowly added to the gel, stir well until it was evenly distributed, sealed at temperature to  $37 \pm 1$  °C. The cultured SRB sludge suspension was centrifuged at 3000 rpm for 10 min and the supernatant was removed. The 30  $\text{mg L}^{-1}$  SRB was added to the prepared gel mixture and stirred evenly. The gel mixture was dropped into 2%  $\text{CaCl}_2$  saturated boric acid solution with a specific syringe and the particles were removed after 4 hours of cross-linking and stirring with a 100 rpm agitator. Finally, the particles were washed with 0.9% of saline. Before the particles were used, they were activated in an anaerobic environment with an improved Starkey medium solution without organic ingredients for 12 h. The method was used in the follow-up experiments to prepare immobilized particles.

Preparation method of unmodified corncob fixed SRB sludge particles: compared with the modified corncob fixed SRB sludge particle preparation method, the modified corncob material was replaced by the unmodified corncob material, and the other steps were the same.

Preparation method of SRB sludge fixed particles without corncob: compared with the modified corncob fixed SRB sludge particle preparation method, no modified corncob material was added in the gel, and other steps were the same.

### Single factor experimental method

Factors and levels of single-factor experiment: in the single factor test, alkaline  $\text{H}_2\text{O}_2$  modification time, NaOH concentration and  $\text{H}_2\text{O}_2$  concentration were selected as the single factors. The modification time was set to 6 h, 12 h, 18 h, 24 h and 30 h respectively. NaOH concentration was set at 3%, 4%, 5%, 6% and 7% respectively.  $\text{H}_2\text{O}_2$  concentration was set at 0.5%, 1%, 1.5%, 2% and 2.5% respectively.

Method: NaOH solution and  $\text{H}_2\text{O}_2$  solution with different concentrations were prepared according to the factors and levels of the single-factor experiment. Different concentration of NaOH solution was added to different concentration of  $\text{H}_2\text{O}_2$  solution to form different concentration of alkaline  $\text{H}_2\text{O}_2$  solution. According to the modified corncob preparation method mentioned above, the corncob was placed in different

concentrations of alkaline  $\text{H}_2\text{O}_2$  solution to form the modified corncob material required by single-factor experiment. Then modified corncob fixed SRB sludge particles were prepared according to the preparation method mentioned above. The corncob fixed SRB sludge particles prepared under different modification conditions were placed in 200 mL wastewater respectively, and the water quality indexes were measured by sampling after reaction for a period of time. With COD release and  $\text{SO}_4^{2-}$  removal percentage as the main evaluation indicators, the optimal conditions for corncob modification were analyzed and determined to guide orthogonal tests.

$$\text{Removal percentage} = [(C_0 - C_t)/C_0] \times 100\%$$

where,  $C_0$  and  $C_t$  are the initial concentration of ions and the concentration after treatment ( $\text{mg L}^{-1}$ ) respectively.

### Orthogonal experimental method

The optimal preparation conditions for the alkaline  $\text{H}_2\text{O}_2$  modified corncob were determined by  $L_9$  ( $3^3$ ) orthogonal test. Nine extractions were carried out at the modification time was 18 h, 24 h, 30 h, the NaOH concentration was 4%, 5%, 6%, and the  $\text{H}_2\text{O}_2$  concentration was 0.5%, 1%, 1.5% on the basis of the single-factor test. The optimal modification conditions of corncob were determined by using  $\text{SO}_4^{2-}$  removal percentage,  $\text{Mn}^{2+}$  removal percentage,  $\text{Fe}^{2+}$  removal percentage, COD release amount and pH value. Among them, the calculation formula for the removal percentage of  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  was the same as that of the single-factor experiment.

### Dynamic experimental method

Fixed SRB sludge particles without corncob, unmodified corncob fixed SRB sludge particles and modified corncob fixed SRB sludge particles were prepared by using the preparation method of immobilized particles mentioned above (modified corncob conditions: modification time was 24 h, NaOH concentration was 6%,  $\text{H}_2\text{O}_2$  concentration was 1.5%).

The dynamic tests adopted three groups of organic glass tube with inner diameter of 60 mm and height of 400 mm. The three dynamic tubes were constructed into 3 columns, including Column 1 filled with SRB sludge fixed particles without corncob, Column 2 filled with unmodified corncob fixed SRB sludge particles, Column 3 filled with modified corncob fixed SRB sludge particles. The immobilized particles were filled with 20 mm quartz sand of 3–5 mm in diameter on the top and bottom, which could fix and protect the particles. The water flew in from the bottom and out from the top. The water inlet was conducted according to Table 1. Flow rate was

Table 1 Two-stage water quality index of dynamic test

Project	$\text{SO}_4^{2-}/(\text{mg L}^{-1})$	$\text{Fe}^{2+}/(\text{mg L}^{-1})$	$\text{Mn}^{2+}/(\text{mg L}^{-1})$	pH
Stage I	800–850	14.5–15	8.5–9	4–4.5
Stage II	1450–1500	24.5–25	13.5–14	3–3.5



controlled with the peristaltic pump and the flow meter as  $1 \times 10^{-5} \text{ m}^3 \text{ s}^{-1}$ . The test device is shown in Fig. 1. The test was carried out continuously, and sampled regularly once a day to analyze and determine water quality indicators. At 8:00 every morning, the residual  $\text{SO}_4^{2-}$  concentration,  $\text{Fe}^{2+}$  concentration,  $\text{Mn}^{2+}$  concentration, COD concentration and pH value in the effluent water of the three dynamic columns were measured, and the removal percentage of  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  in the influent water of AMD were calculated by the three dynamic columns. Among them, the calculation formula of the removal percentage of  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  is the same as the single-factor experiment.

### Kinetic analysis of immobilized particles

1# modified corncob fixed SRB sludge particles (modification conditions: modification time: 24 h, NaOH concentration: 6%,  $\text{H}_2\text{O}_2$  concentration: 1.5%) and 2# unmodified corncob fixed SRB sludge particles were prepared by using the immobilized particle preparation method mentioned above. Modified corncob fixed SRB sludge particles (1#) and unmodified corncob fixed SRB sludge particles (2#) of 20 g were added to 200 mL simulated wastewater respectively according to the solid-liquid ratio of 1 : 10 ( $m : V$ ,  $\text{g mL}^{-1}$ ), and placed in a constant temperature shaker of 100 rpm and 25 °C. The residual  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  concentrations in the wastewater were measured by sampling at 24 h intervals, and the removal percentage and adsorption capacity of the immobilized particles to  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  were calculated. Among them, the calculation formula of the removal percentage of  $\text{SO}_4^{2-}$ ,  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  is the same as the single-factor experiment. The calculation formula of adsorption capacity  $q$  is as follows:

$$\text{Adsorbing capacity } q = [(C_0 - C_t) \times V] / m$$

where,  $q$  is the adsorption capacity of immobilized particles to ions ( $\text{mg g}^{-1}$ ),  $C_0$  is the initial concentration ( $\text{mg L}^{-1}$ ),  $C_t$  is the residual concentration after treatment ( $\text{mg L}^{-1}$ ),  $V$  is solution volume (L),  $m$  is the mass of immobilized particles (g).

### Water quality monitoring methods

$\text{SO}_4^{2-}$ : barium chromate spectrophotometry;  $\text{Mn}^{2+}$ : potassium periodate spectrophotometry;  $\text{Fe}^{2+}$ : phenanthroline spectrophotometry; COD: rapid digestion spectrophotometry; pH: glass electrode method.

## Results and discussion

### Single factor test

**Determination of the optimal modification time.** The test results are shown in Fig. 2. From Fig. 2, with the extension of the modification time, COD release increased first, then decreased and then increased, while residual concentration of  $\text{SO}_4^{2-}$  decreased first and then increased slightly. This indicates that the reaction time had a great effect on the transformation degree of hemicellulose, cellulose and lignin in corncob. The hemicellulose, cellulose and lignin in corncob could not be completely dissolved and transformed in a short modification time, which affects the performance of corncob as a slow-release carbon source for further hydrolysis.<sup>18</sup> However, the long time of modification will further destroy the internal structure of hemicellulose, cellulose and lignin in corncob, which will result in the poor hydrolysis of carbon source such as glucose and fructose, which are easy to be used by SRB, and affect the activity of SRB dissimilation  $\text{SO}_4^{2-}$ . From Fig. 2, the COD emission reached the maximum value when the modification time was 6 h, which is consistent with the study results of Su Yapeng *et al.*<sup>19</sup> Studies have shown that lignin in corncobs will detach from solution during the first 6 h in modification corncob with alkaline hydrogen peroxide.<sup>19</sup> Therefore, COD value reached the highest level at 6 h. When the reaction time was 24 h, both COD release amount and the  $\text{SO}_4^{2-}$  residual concentration reached the lowest level, indicating that hemicellulose, cellulose and lignin in corncob were dissolved and transformed in large quantities, which improved the further hydrolysis of the corncob to form small molecules organic matter and promoted the dissimilatory reduction activity of

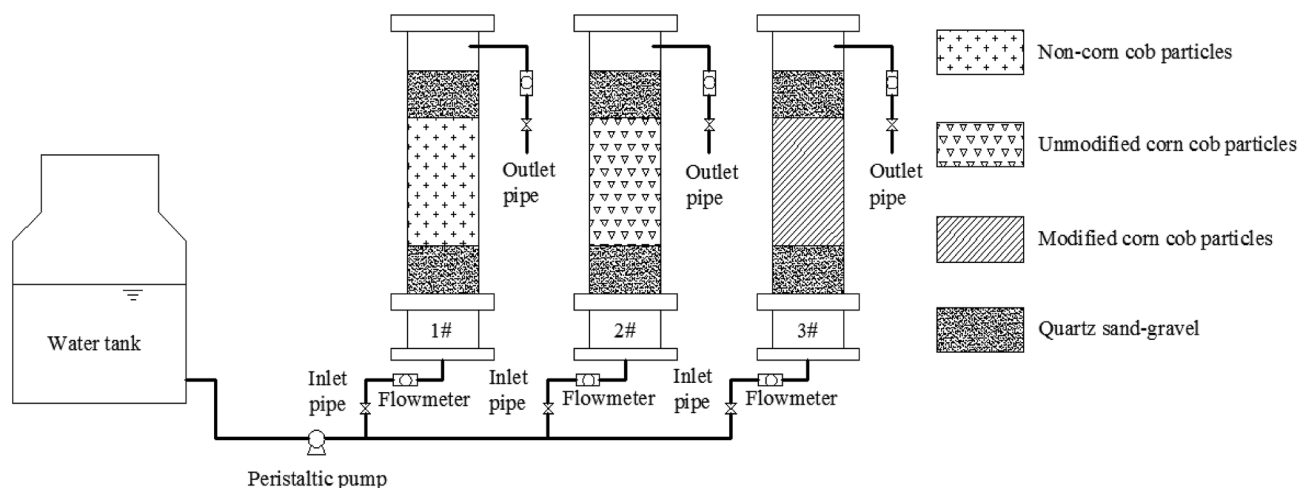


Fig. 1 System diagram of dynamic testing device.



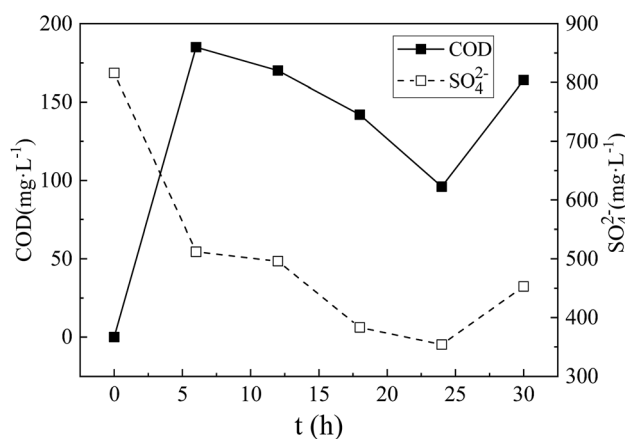


Fig. 2 COD release and  $\text{SO}_4^{2-}$  residual concentration at different modification time. The NaOH concentration was 5%, the  $\text{H}_2\text{O}_2$  concentration was 1.5%.

SRB. SRB used organic matter to carry out dissimilation and reduction metabolism, which reduced both COD release and  $\text{SO}_4^{2-}$  residual concentration in the solution. At 30 h, the biological activity of SRB was reduced. The ability of SRB to reduce  $\text{SO}_4^{2-}$  by organic metabolism was reduced, resulting in the increase of COD and  $\text{SO}_4^{2-}$  in the system. In summary, the optimal modification time for this experiment was 24 h.

**Determination of optimum NaOH concentration.** The test results are shown in Fig. 3.

From Fig. 3, with the increase of NaOH concentration, both COD release and  $\text{SO}_4^{2-}$  residual concentration decreased first and then increased. This may be due to the fact that alkali can use  $\text{OH}^-$  to break the lignin ether bond, and  $\text{OH}^-$  can saponify the ester bond between hemicellulose and lignin, which can dissolve most of lignin, and dissolve part of hemicellulose.<sup>20</sup> Thus the internal structure of the corncob was changed and the ability of further hydrolysis of corncob to form small molecular organic matter was improved too. However, when the NaOH concentration was too high, on the basis of further removal of

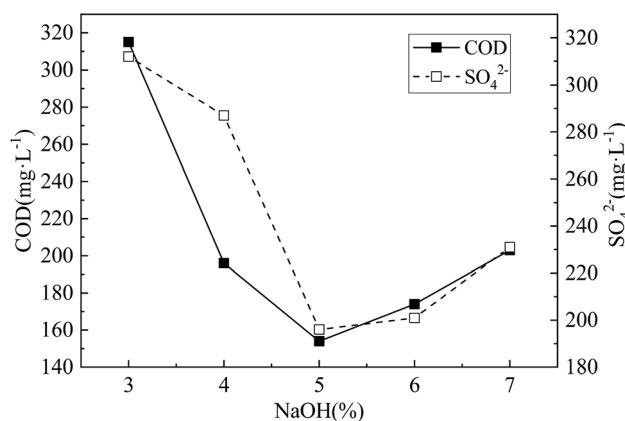


Fig. 3 COD release and  $\text{SO}_4^{2-}$  residual concentration under different NaOH concentration. The  $\text{H}_2\text{O}_2$  concentration was 1.5%, the modification time was 24 h.

lignin and hemicellulose, the degree of crystallinity and polymerization of the cellulose were further reduced or even decomposed, resulting in the loss of degradable substances, and insufficient carbon supply for SRB dissimilation reduction. Moreover, high concentration of NaOH inhibited the activity of SRB dissimilation  $\text{SO}_4^{2-}$  to a certain extent, resulting in the decrease of  $\text{SO}_4^{2-}$  removal percentage. The research of Nicolas Le Moigne<sup>21</sup> shows that lignocellulose can be dissolved by continuous dismantling and breaking in NaOH solution. When the concentration of NaOH is 5%, the amount of COD release and the residual concentration of  $\text{SO}_4^{2-}$  are both at the lowest level. This also shows that the hemicellulose, cellulose and lignin in the corncob are dissolved and transformed in large quantities, which enhances the ability of corncob to hydrolyze to form small molecular organic matter and promotes the dissimilatory reduction activity of SRB. In summary, the optimal NaOH concentration was 5%.

**Determination of optimal  $\text{H}_2\text{O}_2$  concentration.** The test results are shown in Fig. 4. From Fig. 4, with the increase of  $\text{H}_2\text{O}_2$  concentration, both COD release and  $\text{SO}_4^{2-}$  residual concentration decreased first and then increased. This is mainly because  $\text{H}_2\text{O}_2$  could degrade lignin with its own peroxidic ion oxidation and destroy the complex structure of lignin. However, due to the easy decomposition of  $\text{H}_2\text{O}_2$ , the effect of using  $\text{H}_2\text{O}_2$  along to treat waste water is not ideal. When alkali and  $\text{H}_2\text{O}_2$  are used simultaneously, the alkali can activate  $\text{H}_2\text{O}_2$ , enhance the removal of lignin, and improve the bioavailability of corncob,<sup>22</sup> which is beneficial to SRB heterogeneous reduction. Thus the amount of COD release and the residual concentration of  $\text{SO}_4^{2-}$  decreased with the increase of concentration of  $\text{H}_2\text{O}_2$ . When the content of  $\text{H}_2\text{O}_2$  was too high,  $\text{H}_2\text{O}_2$  and NaOH easily form  $\text{Na}_2\text{O}_2$ , which reduced the ability of NaOH to dissolve and the ability of  $\text{H}_2\text{O}_2$  to oxidize hemicellulose, cellulose and lignin in corncob, and formed a material structure that is not conducive to SRB utilization. The amount of COD release and the residual concentration of  $\text{SO}_4^{2-}$  increased as the concentration of  $\text{H}_2\text{O}_2$  increased. When the concentration of  $\text{H}_2\text{O}_2$  was 1%, both COD release and  $\text{SO}_4^{2-}$  residual concentration reached a lower level,

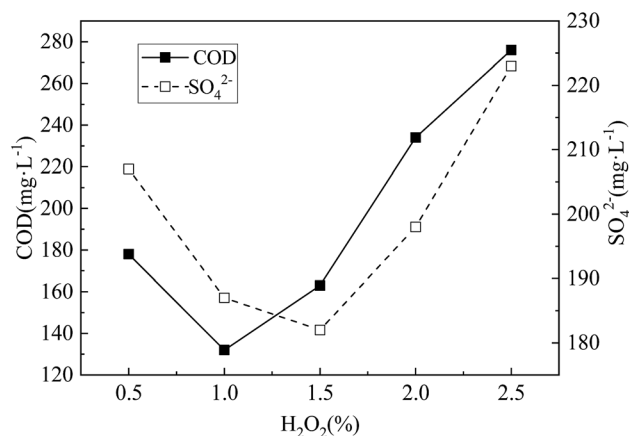


Fig. 4 COD release and  $\text{SO}_4^{2-}$  residual concentration under different  $\text{H}_2\text{O}_2$  concentration. The NaOH concentration was 5%, the modification time was 24 h.



indicating that the lignin in the corncob was fully oxidized, which improved the ability of the corncob to further hydrolyze to form small molecular organic matter, and promoted the dissimilatory reduction activity of SRB. In summary, the optimal  $\text{H}_2\text{O}_2$  concentration was 1%.

### Orthogonal experimental study

Based on the composition of the single factor tests and the test results, taking the modification time, NaOH concentration and  $\text{H}_2\text{O}_2$  concentration as three factors, according to the orthogonal experiment table, the orthogonal experiment of  $L_9(3^3)$  was carried out to determine the optimum modification condition of corncob.

Orthogonal test settings and test results are shown in Table 2.

According to Table 2, combined with visual analysis and analysis of variance, the best combination of  $\text{SO}_4^{2-}$  removal percentage,  $\text{Mn}^{2+}$  removal percentage,  $\text{Fe}^{2+}$  removal percentage, COD release amount, and pH adjustment ability were A3B1C2, A2B3C3, A3B1C1, A2B3C3, and A1B3C1, respectively. After comprehensive consideration, the optimal proportion of immobilized particles was A2B3C3, that is, the modification time was 24 h, the NaOH concentration was 6%, and the  $\text{H}_2\text{O}_2$  concentration was 1.5%. The immobilized particles prepared by the modified corncob under these conditions had the best AMD treatment effect.

### Dynamic experimental study

Combined with the results of the previous orthogonal test, the AMD dynamic experiment of immobilized particle treatment was carried out by constructing 3 dynamic columns. The experimental results and analysis are as follows. The "Influent" in Fig. 5–9 indicates the influent concentrations of various ions in AMD during the dynamic experiment, which were tested each morning at 8:00.

**Analysis on the change rules of  $\text{SO}_4^{2-}$ .** According to Fig. 5(a) and (b), in the first stage of the test, the residual amount of  $\text{SO}_4^{2-}$  in the dynamic Columns 1, 2 and 3 showed a downward trend. By the 5–6 d, the residual amount of  $\text{SO}_4^{2-}$  in the three columns all reached the lowest level. At the 1–7 d, the average residual amount of  $\text{SO}_4^{2-}$  were  $624.1 \text{ mg L}^{-1}$ ,  $500.8 \text{ mg L}^{-1}$  and

$416.3 \text{ mg L}^{-1}$  respectively, and the average removal percentages of  $\text{SO}_4^{2-}$  were 21.9%, 37.2% and 47.9% respectively. In the second stage of the test, with the increase of pollutant load, the residual amount of  $\text{SO}_4^{2-}$  in the Columns 1, 2 and 3 decreased first and then increased. At the 12–13 d, the residual amount of  $\text{SO}_4^{2-}$  in the three columns reached the lowest level. At the 8–15 d, the average residual amount of  $\text{SO}_4^{2-}$  were  $1002.4 \text{ mg L}^{-1}$ ,  $914.2 \text{ mg L}^{-1}$  and  $735.8 \text{ mg L}^{-1}$  respectively, and the corresponding  $\text{SO}_4^{2-}$  average removal percentages were 30.4%, 36.4% and 48.8% respectively.

In the first stage of the reaction, the  $\text{SO}_4^{2-}$  removal percentage increased rapidly in the three columns. This may be due to the fact that at the beginning of the reaction, with sufficient carbon source, SRB metabolic activity was strong, which was able to resist the low pollution load on its activity inhibition and can have a good dissimilation reduction on  $\text{SO}_4^{2-}$ .<sup>23</sup> So the removal effect was more obvious. In the second stage of the reaction, the removal percentage of  $\text{SO}_4^{2-}$  in three columns showed a trend of first increasing and then significantly decreasing. At the same time, the increase and decrease of  $\text{SO}_4^{2-}$  removal percentage in the three columns were different. This may be due to the fact that the increasing pollution load in the initial phase had not great impact on SRB activity, and SRB could still have a good dissimilation reduction on  $\text{SO}_4^{2-}$ . However, as the reaction progressing, the carbon source in the particles is continuously consumed by the SRB, and the COD/ $\text{SO}_4^{2-}$  ratio decreased below 0.67,<sup>24</sup> the optimum carbon and sulfur ratio required for microbial growth. In addition, the toxicity and inhibition of SRB by persistent high concentration  $\text{Mn}^{2+}$  and acidic pH resulted in a significant decrease in  $\text{SO}_4^{2-}$  removal percentage at the second stage of the reaction.<sup>25</sup>

The removal capacity of  $\text{SO}_4^{2-}$  and the resistance to pollution load change of three dynamic columns are Columns 3, 2 and 1 in order from strong to weak. The main reason was that after the corncob in Column 3 was modified, the alkaline  $\text{H}_2\text{O}_2$  could swell the cellulose, remove most of the lignin and part of the hemicellulose, increase the inner surface area of the corncob, and reduce the crystallinity and polymerization degree of the cellulose. The ability of the corncob to further hydrolyze to form small molecular organic substances such as glucose and fructose in the immobilized particles was improved, which

Table 2  $L_9(3^3)$  Orthogonal test design and results

Test number	Time A/h	NaOH B/%	$\text{H}_2\text{O}_2$ C/%	The removal percentage of $\text{SO}_4^{2-}$ /%	The removal percentage of $\text{Mn}^{2+}$ /%	The removal percentage of $\text{Fe}^{2+}$ /%	pH	COD/( $\text{mg L}^{-1}$ )
1	18	4	0.5	86.20	76.44	93.57	7.89	296
2	18	5	1	89.69	66.25	86.64	7.77	267
3	18	6	1.5	81.65	80.93	82.42	8.07	254
4	24	4	1	82.07	72.53	87.42	7.79	252
5	24	5	1.5	79.06	75.38	84.85	7.84	240
6	24	6	0.5	85.44	79.22	86.00	7.91	227
7	30	4	1.5	90.17	75.06	89.92	7.74	280
8	30	5	0.5	82.94	69.18	90.14	7.87	305
9	30	6	1	87.33	68.61	84.78	7.91	282



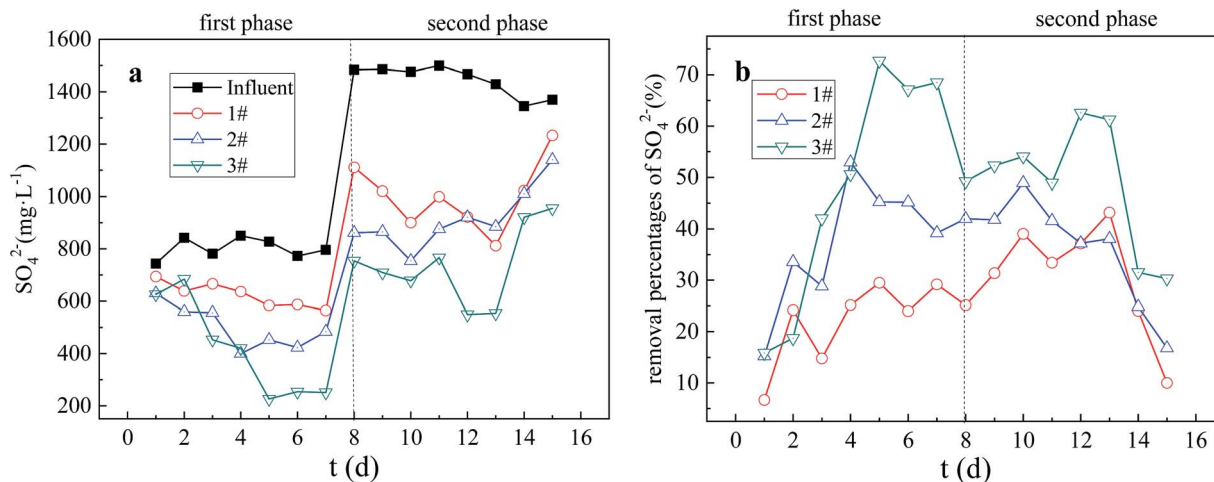


Fig. 5 (a)  $\text{SO}_4^{2-}$  residual concentration in Columns 1, 2, 3 (b) Removal percentage of  $\text{SO}_4^{2-}$  in Columns 1, 2, 3. Influent is the  $\text{SO}_4^{2-}$  concentration in influent water of AMD. 1#, 2# and 3# in (a) are the residual  $\text{SO}_4^{2-}$  concentration in AMD after repairing AMD with Column 1, Column 2 and Column 3 respectively. 1#, 2# and 3# in (b) are the removal percentage of  $\text{SO}_4^{2-}$  in AMD by the three dynamic columns corresponding to (a) respectively.

provided sufficient carbon source for SRB growth and metabolism, and could well resist the inhibition of high concentration pollution load on its activity. So the catalytic activity of SRB in dissimilation reduction of  $\text{SO}_4^{2-}$  was promoted. Therefore, Column 3 had the strongest ability to remove  $\text{SO}_4^{2-}$  and resist the change of pollution load. The unmodified corncob, in Column 2, contained lignin and cellulose, which were difficult to be decomposed and cannot form fermentable sugars. Only part of hemicellulose was easily degraded into small molecular organic substances, such as monosaccharides, to supply SRB growth. In Column 1, the particle activity was inhibited due to the absence of the external carbon source, so it had the worst ability to remove  $\text{SO}_4^{2-}$  and resist change of pollution load.

**Analysis on the change rules of  $\text{Fe}^{2+}$ .** It can be seen from Fig. 6(a) and (b) that in the first stage of the test, the residual amount of  $\text{Fe}^{2+}$  in the dynamic columns 1, 2 and 3 showed a downward trend. At the 1–7 d, the average residual amount of  $\text{Fe}^{2+}$  in the three columns were 10.95  $\text{mg L}^{-1}$ , 8.36  $\text{mg L}^{-1}$  and 7.28  $\text{mg L}^{-1}$  respectively, and the average removal percentages of  $\text{Fe}^{2+}$  were 22.51%, 43.08% and 50.49% respectively. In the second stage of the experiment, the residual amount of  $\text{Fe}^{2+}$  in the three dynamic columns showed an upward trend. At the 8–15 d, the average residual amount of  $\text{Fe}^{2+}$  were 21.90  $\text{mg L}^{-1}$ , 19.34  $\text{mg L}^{-1}$  and 18.46  $\text{mg L}^{-1}$  respectively, and the corresponding  $\text{Fe}^{2+}$  average removal percentages were 11.05%, 21.91% and 25.45% respectively.

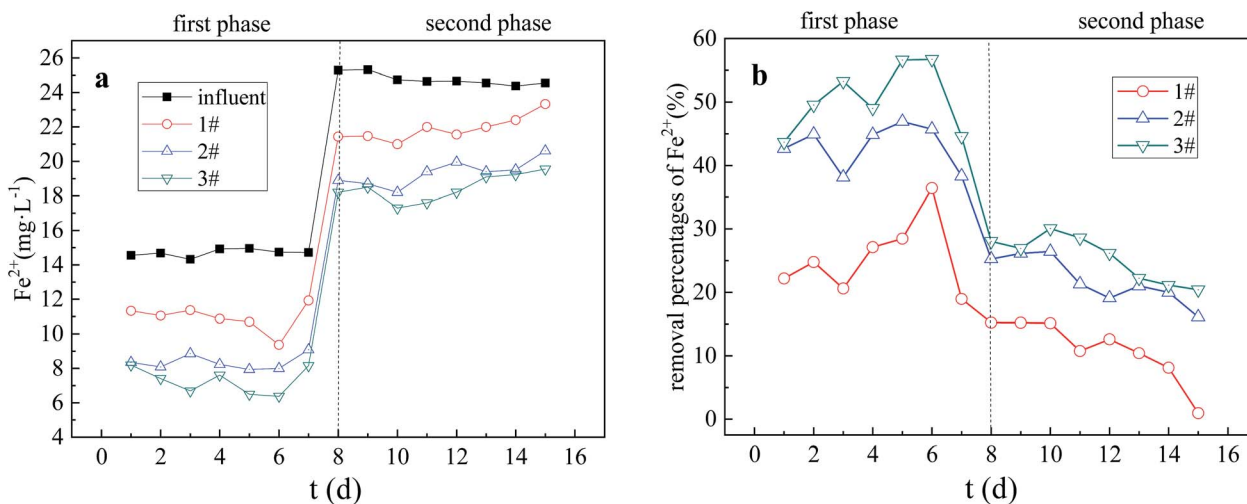


Fig. 6 (a)  $\text{Fe}^{2+}$  residual concentration in Columns 1, 2, 3. (b) The removal percentage of  $\text{Fe}^{2+}$  in Columns 1, 2, 3. Influent is the  $\text{Fe}^{2+}$  concentration in influent water of AMD. 1#, 2# and 3# in (a) are the residual  $\text{Fe}^{2+}$  concentration in AMD after repairing AMD with Column 1, Column 2 and Column 3 respectively. 1#, 2# and 3# in (b) are the removal percentage of  $\text{Fe}^{2+}$  in AMD by the three dynamic columns corresponding to (a) respectively.



In the first stage, the removal percentage curves of  $\text{Fe}^{2+}$  in the three columns showed an upward trend. This is because  $\text{Fe}^{2+}$  removal in the dynamic columns is mainly the results of the adsorption of immobilized particles, SRB dissimilation reduction  $\text{SO}_4^{2-}$  genesis of metal sulfide precipitation and SRB biosorption flocculation.<sup>26</sup> In the initial stage of the reaction, the particle adsorption capacity was large and the SRB biological activity was strong. Therefore, the  $\text{Fe}^{2+}$  removal effect was significant in the initial stage of the reaction. In the second stage, the  $\text{Fe}^{2+}$  removal percentage curves of the three columns decreased, which may be due to the increase of the pollution load in the second stage and the gradual saturation of the particle adsorption capacity. At the same time, high concentration of  $\text{Mn}^{2+}$  entered the SRB cells, destroyed the enzyme proteins and deactivated them, inhibited the biological activity of SRB and produced toxic effects on them, thus weakening the dissimilation and reduction process of  $\text{SO}_4^{2-}$  by sulfate-reducing bacteria, and affecting the precipitation of reduction product  $\text{S}^{2-}$  and the removal of  $\text{Mn}^{2+}$  by biosorption flocculation. So the removal percentage of  $\text{Fe}^{2+}$  decreased significantly.

The ability of the three dynamic columns to remove  $\text{Fe}^{2+}$  and resist pollution load changes was Columns 3, 2 and 1 in order from strong to weak. Due to the increase in the specific surface area of the corncob and the  $\text{OH}^-$  content of the granules in the immobilized particles of Column 3, which were modified by alkaline  $\text{H}_2\text{O}_2$ , the adsorption capacity of the particles increased, and it is easy to form hydroxides and carbonate precipitation, which enhanced the removal of  $\text{Fe}^{2+}$ . So Column 3 had the strongest ability to remove  $\text{Fe}^{2+}$  and resist pollution load change. In the Column 1, however, the biological activity was inhibited due to the absence of additional carbon source. At the same time, the process of  $\text{S}^{2-}$  precipitation and biosorption flocculation to remove  $\text{Fe}^{2+}$  was weak, and the adsorption capacity of the immobilized particles was low, so Column 1 had the worst ability to remove  $\text{Fe}^{2+}$  and resist pollution load changes.

**Analysis on the change rules of  $\text{Mn}^{2+}$ .** It can be seen from Fig. 7(a) and (b) that the  $\text{Mn}^{2+}$  residuals in the three dynamic columns showed different fluctuation in the first stage of the experiment. At the 1–7 d, the average residual amount of  $\text{Mn}^{2+}$  in the three systems were 6.69  $\text{mg L}^{-1}$ , 4.56  $\text{mg L}^{-1}$  and 4.18  $\text{mg L}^{-1}$  respectively. At this time, the corresponding  $\text{Mn}^{2+}$  removal percentages were 23.15%, 47.68% and 51.96% respectively. In the second stage of experiment (8–15 d), the residuals of  $\text{Mn}^{2+}$  in the three dynamic columns fluctuated and increased, and the average residual amount of  $\text{Mn}^{2+}$  in the three systems were 10.53  $\text{mg L}^{-1}$ , 9.19  $\text{mg L}^{-1}$  and 8.46  $\text{mg L}^{-1}$  respectively. At this time, the corresponding  $\text{Mn}^{2+}$  removal percentages were 23.46%, 33.16% and 38.48% respectively.

In the first stage, the  $\text{Mn}^{2+}$  removal percentage curves of Columns 1 and 2 showed a downward trend, while  $\text{Mn}^{2+}$  removal percentage curve of Column 3 showed an upward trend. In the second stage, the  $\text{Mn}^{2+}$  removal percentage curves of the three columns fluctuated and decreased. The SRB immobilized particles have the same mechanism of action for removing heavy metals  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  from AMD. However, the average removal percentage of  $\text{Fe}^{2+}$  in the first stage was slightly

higher than the average removal percentage of  $\text{Mn}^{2+}$ , while the average removal percentage of  $\text{Fe}^{2+}$  in the second stage was slightly lower than the average removal percentage of  $\text{Mn}^{2+}$ . Karathanasis<sup>27</sup> found through experiments that it is difficult to form  $\text{Mn}^{2+}$  sulfides, when other metal ions are present in the bioreactor. The removal of  $\text{Mn}^{2+}$  relies mainly on adsorption to form oxides, hydroxides or carbonates. Therefore, in the first stage of strong SRB biological activity, the average removal percentage of  $\text{Fe}^{2+}$  was slightly higher than the average removal percentage of  $\text{Mn}^{2+}$  due to the influence of  $\text{Fe}^{2+}$  on the  $\text{S}^{2-}$  precipitation and biosorption flocculation removal of  $\text{Mn}^{2+}$ . However, in the second stage, due to the increase of pollution load, gradual saturation of the particle adsorption capacity, inhibit of the biological activity, the weakening of biological removal process of  $\text{Fe}^{2+}$ , and the replace of  $\text{Fe}^{2+}$  with the  $\text{Mn}^{2+}$  entering the particle by ion exchange, the average removal percentage of  $\text{Fe}^{2+}$  was slightly lower than the average removal percentage of  $\text{Mn}^{2+}$ .

**Analysis on the change rules of COD.** It can be seen from Fig. 8 that in the first stage of the experiment, the COD release of effluent in the three dynamic columns increased first and then decreased (1–7 d), and the average releases were 146.14  $\text{mg L}^{-1}$ , 508.14  $\text{mg L}^{-1}$  and 234  $\text{mg L}^{-1}$  respectively. In the second stage of the experiment (8–15 d), the concentration of effluent COD in the three dynamic columns decreased steadily, and the average COD release of the three systems were 36.75  $\text{mg L}^{-1}$ , 152.88  $\text{mg L}^{-1}$  and 102.75  $\text{mg L}^{-1}$  respectively.

The concentration of effluent COD in the dynamic columns was related to the leakage of organic matters and its biological metabolites in the immobilized particles. In Column 1, the effluent COD was mainly caused by the leakage of SRB sludge and its metabolites in the immobilized particles. While in Columns 2 and 3, the effluent COD was mainly produced both by the leakage of the SRB sludge, and the corncob and its hydrolyzed products in the immobilized particles. In the initial stage of the first stage reaction, the SRB metabolic activity utilized less organic matters, resulting in a large amount of organic matter leakage in the immobilized particles, thus the effluent COD concentration gradually increased. However, when the SRB activity gradually increased, and the organic matter required for its metabolism also increased, the effluent COD concentration gradually decreased. In the second stage, with the increase in pollution load, the concentration of COD release gradually decreased, due to excessive consumption of organic matter in the immobilized particles.

The COD concentration values of the three columns were Columns 2, 3 and 1 in order. Since the immobilized particles in Column 1 only contained the SRB sludge carbon source, the organic mass of the leakage was small, and thus the effluent COD concentration of Column 1 was always the smallest. Column 3 had the strongest bioactivity, and the amount of organic matters needed for SRB metabolic growth was large. Therefore, the organic matter leakage was less. So the COD concentration of effluent of Column 3 was less than that of Column 2.

**Analysis on the change rules of pH.** It can be seen from Fig. 9 that in the beginning initial stage of the first stage of the test,





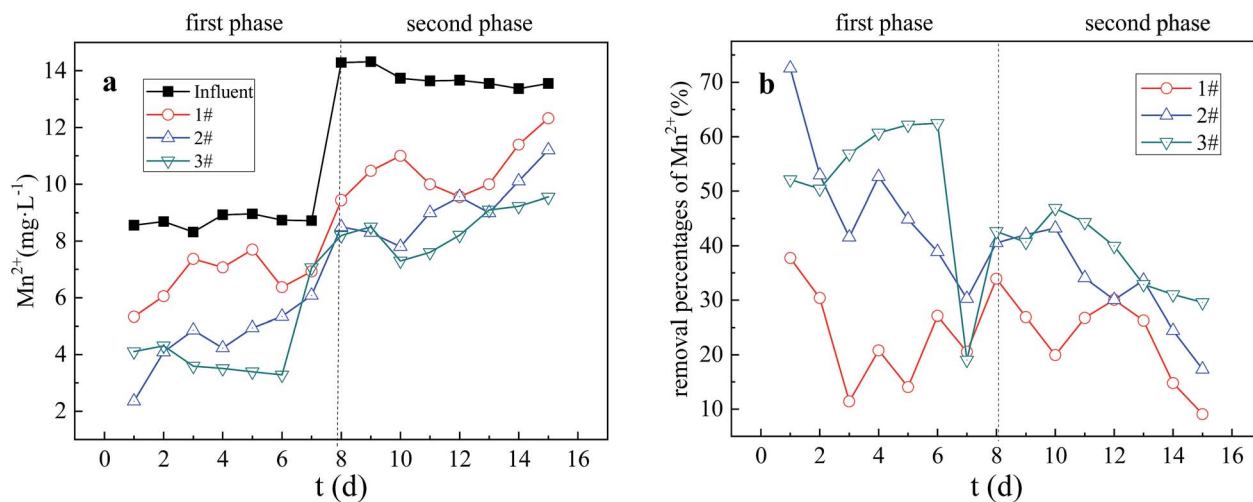


Fig. 7 (a)  $Mn^{2+}$  residual concentration in Columns 1, 2, 3 (b) The removal percentage of  $Mn^{2+}$  in Columns 1, 2, 3. Influent is the  $Mn^{2+}$  concentration in influent water of AMD. 1#, 2# and 3# in (a) are the residual  $Mn^{2+}$  concentration in AMD after repairing AMD with Column 1, Column 2 and Column 3 respectively. 1#, 2# and 3# in (b) are the removal percentage of  $Mn^{2+}$  in AMD by the three dynamic columns corresponding to (a) respectively.

the pH values of the effluent of the 3 dynamic columns all increased rapidly, and then decreased slightly. The first stage of the test (1–7 d), the average pH of the effluent were 6.32, 6.56 and 6.96 respectively. In the second stage of the experiment (8–15 d), the pH values of the effluent in the 3 dynamic columns all decreased to varying degrees, and the average effluent pH were 5.96, 6.36 and 6.65 respectively.

The process of raising the pH of the solution in the dynamic columns is mainly the effect of particles on the adsorption of  $H^+$  and the regulation of microbial growth and metabolism.<sup>28</sup> The main reason for the rapid increase of pH in the beginning of stage I is the adsorption of  $H^+$  by particles. As the reaction progresses to the second stage, the soluble carbon source continuously decreases, the SRB dissimilatory reduction activity

decreases as well, and the high-concentration heavy metal ions and high acidity inhibit the metabolic activity of SRB in the particles,<sup>29</sup> resulting in the amount of alkaline substances produced by microorganisms is reduced. Therefore, the pH of the effluent decreased to varying degrees.

The order of pH improvement capacity and resistance to load capacity of 3 columns was Column 3, 2 and 1. Column 3 had the strongest biological activity, the largest ability to adjust the acid, and the corncob in the particles was modified by alkaline  $H_2O_2$ , which contained a large amount of alkaline substances which could be neutralized with  $H^+$ . So Column 3 was strongest in improving pH of solution and resisting the impact load. While Column 1 had the weakest biological activity, so it performed the worst pH in improvement and impact load resistance.

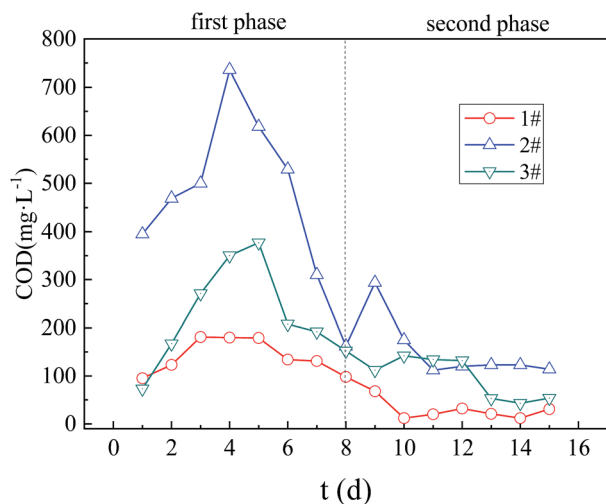


Fig. 8 The changes of COD in Columns 1, 2, 3. 1#, 2# and 3# are the COD concentration in AMD after repairing AMD with Column 1, Column 2 and Column 3 respectively.

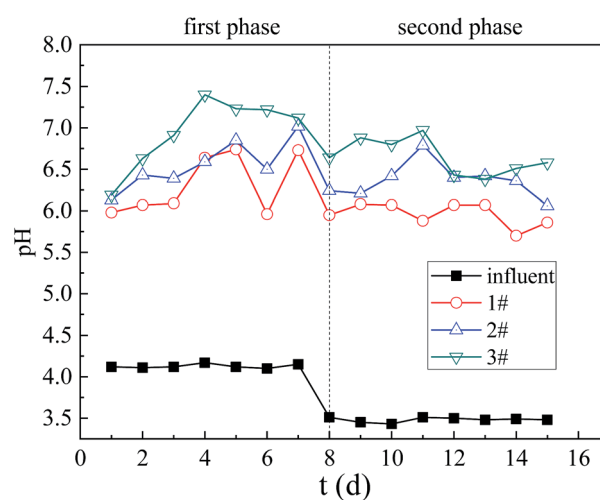


Fig. 9 The changes of pH in Columns 1, 2, 3. Influent is the pH value in influent water of AMD. 1#, 2# and 3# are the pH value in AMD after repairing AMD with Column 1, Column 2 and Column 3 respectively.



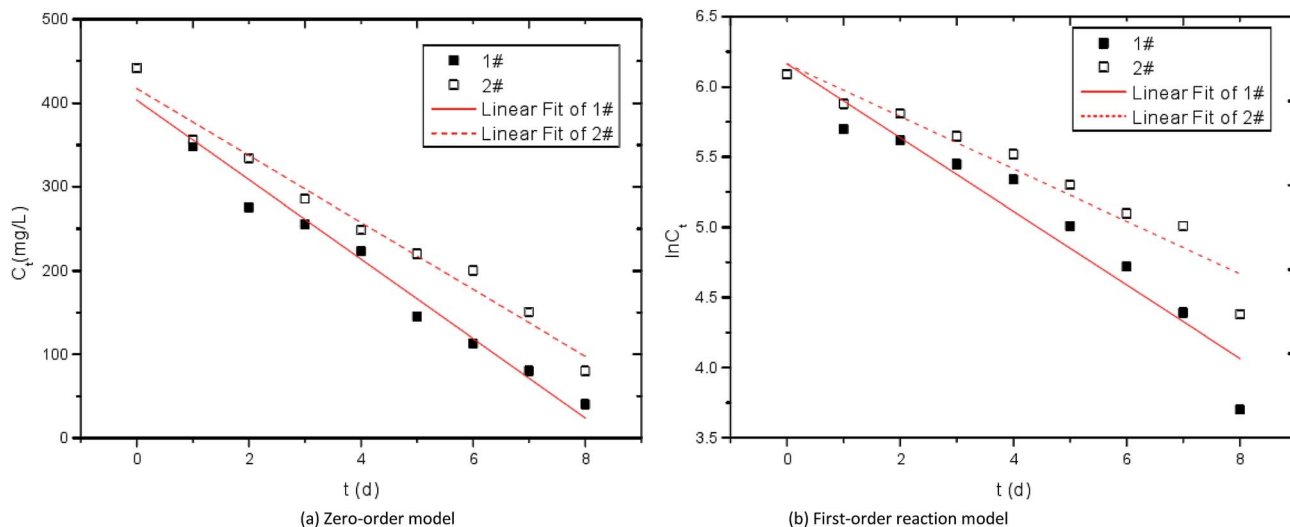


Fig. 10 Kinetic curves of  $\text{SO}_4^{2-}$  reduction by immobilized particles. 1# is modified corncob fixed SRB sludge particles, 2# is unmodified corn cob fixed SRB sludge particles.

Table 3 Coefficients of kinetic models

Sample	Zero-order reaction		First-order reaction	
	$k_0$	$R^2$	$k_1$	$R^2$
1#	47.46	0.9269	0.262	0.9706
2#	39.96	0.9268	0.187	0.9736

### Kinetic analysis of immobilized particles

**Kinetic analysis of reduction of  $\text{SO}_4^{2-}$  by immobilized particles.** Zero-order reaction kinetics equation  $c_t = c_0 - k_0 t$  and first-order reaction kinetics equation  $\ln c_t = \ln c_0 - k_1 t$  were used to fit the curves of  $\text{SO}_4^{2-}$  reduction by modified corncob

fixed SRB sludge particles (1#) and unmodified corncob fixed SRB sludge particles (2#). The fitting curves of  $\text{SO}_4^{2-}$  reaction kinetics were obtained as shown in Fig. 10 and Table 3. Where,  $c_0$  is the initial  $\text{SO}_4^{2-}$  concentration ( $\text{mg L}^{-1}$ ),  $c_t$  is the  $\text{SO}_4^{2-}$  concentration at a certain time ( $\text{mg L}^{-1}$ ),  $k_0$  is the zero-order reaction rate constant ( $(\text{mg L}^{-1} \text{h}^{-1})$ ),  $k_1$  is the first-order reaction rate constant ( $\text{h}^{-1}$ ).

It can be seen from Table 3 that the correlation coefficient  $R^2$  of the first-order reaction models of the two particle materials are both large, indicating that the reduction process of  $\text{SO}_4^{2-}$  by modified corncob fixed SRB sludge particles (1#) and unmodified corncob fixed SRB sludge particles (2#) conforms to first-order kinetics. As the first-order reaction kinetics process is mainly affected by electron acceptor, this model can well prove the process of SRB dissimilation reduction of  $\text{SO}_4^{2-}$ . In the first-

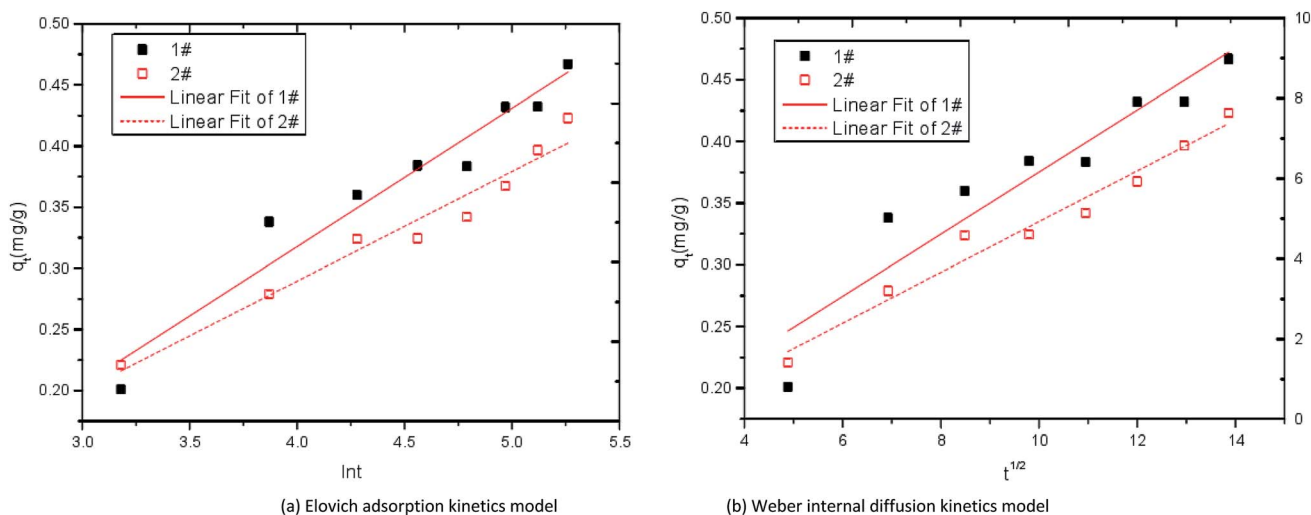


Fig. 11 Kinetic curves of adsorption  $\text{Fe}^{2+}$  by immobilized particles. 1# is modified corncob fixed SRB sludge particles, 2# is unmodified corn cob fixed SRB sludge particles.



**Table 4** Kinetic parameters of adsorption  $\text{Fe}^{2+}$  by immobilized particles

Sample	Elovich adsorption kinetics model				Weber internal diffusion kinetics model		
	$q_e$	$\alpha$	$\beta$	$R^2$	$q_t$	$k$	$R^2$
1#	0.642	0.034	0.012	0.925	0.849	0.023	0.874
2#	0.581	0.028	0.015	0.902	0.760	0.017	0.846

order reduction kinetic model, the reaction rate constant of 1# particle is greater than that of 2# particle, indicating that the dissimilation reduction rate of modified corncob particle to  $\text{SO}_4^{2-}$  is higher than that of unmodified corncob particle, and the synergistic effect of corncob and SRB in immobilized particles improves the dissimilation reduction rate of SRB to  $\text{SO}_4^{2-}$ .

#### Kinetic analysis of $\text{Fe}^{2+}$ adsorption by immobilized particles.

Elovich adsorption kinetics model  $q_e = (2.303/\beta)\ln[t + 1/(\alpha\beta)] - (2.303/\beta)\ln[1/(\alpha\beta)]$  and Weber internal diffusion kinetics model  $q_t = kt^{1/2} + c$  were used to fit the adsorption curves of  $\text{Fe}^{2+}$  by modified corncob fixed SRB sludge particles (1#) and unmodified corncob fixed SRB sludge particles (2#). The adsorption kinetics fitting curves of  $\text{Fe}^{2+}$  were obtained as shown in Fig. 11 and Table 4. Where,  $\alpha$  is the adsorption rate constant ( $\text{g} (\text{mg h}^{-1})^{-1}$ ),  $\beta$  is the desorption rate constant ( $\text{mg h}^{-1}$ ),  $q_e$  is the adsorption amount at adsorption equilibrium ( $\text{mg g}^{-1}$ ),  $q_t$  is the adsorption amount at adsorption time  $t$  ( $\text{mg g}^{-1}$ ),  $k$  is the rate constant of in-particle diffusion ( $\text{g} (\text{mg h}^{1/2})^{-1}$ ),  $c$  is a constant related to the boundary thickness ( $\text{mg g}^{-1}$ ).

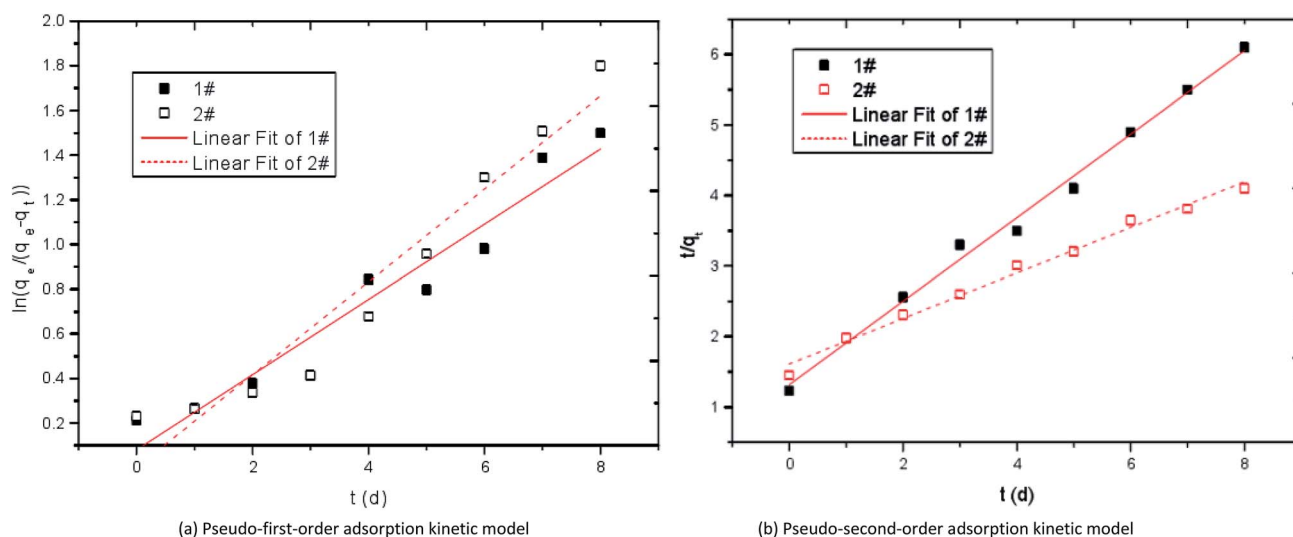
It can be seen from Table 4 that the correlation coefficient  $R^2$  of Elovich adsorption kinetics model is larger than that of Weber internal diffusion kinetics model in both the 1# particle and 2# particle, indicating that Elovich adsorption kinetics

model can better describe the  $\text{Fe}^{2+}$  adsorption process. In the Elovich adsorption kinetics model, the adsorption rate constant of 1# particle is higher than that of 2# particle, indicating that the adsorption rate of modified corncob particles to  $\text{Fe}^{2+}$  is higher than that of unmodified corncob particles.

**Kinetic analysis of  $\text{Mn}^{2+}$  adsorption by immobilized particles.** Lagergren pseudo-first-order adsorption kinetic model  $\ln(q_e - q_t) = \ln q_e - (k_1/2.323)t$  and McKay pseudo-second-order adsorption kinetic model  $t/q_t = 1/(k_2q_e^2) + t/q_e$  were used to fit the adsorption curves of  $\text{Mn}^{2+}$  by modified corncob fixed SRB sludge particles (1#) and unmodified corncob fixed SRB sludge particles (2#). The adsorption kinetics fitting curves of  $\text{Mn}^{2+}$  were obtained as shown in Fig. 12 and Table 5. Where,  $q_e$  is the adsorption amount at adsorption equilibrium ( $\text{mg g}^{-1}$ ),  $q_t$  is the adsorption amount at adsorption time  $t$  ( $\text{mg g}^{-1}$ ),  $k_1$  is the first-order kinetic rate constant ( $\text{h}^{-1}$ ),  $k_2$  is the second-order kinetic rate constant ( $\text{g} (\text{mg h}^{-1})^{-1}$ ).

It can be seen from Table 5 that the correlation coefficient  $R^2$  of the pseudo-second-order adsorption kinetic model is larger than that of the pseudo-first-order adsorption kinetic model in both the 1# particle and 2# particle, indicating that the pseudo-second-order adsorption kinetic model can better describe the adsorption process of  $\text{Mn}^{2+}$  and the adsorption of  $\text{Mn}^{2+}$  by particles is dominated by chemical adsorption. In the pseudo-second-order adsorption kinetic model,  $k_2$  of 1# particle is larger than that of 2# particle, indicating that the adsorption rate of modified corncob particles to  $\text{Mn}^{2+}$  is higher than that of unmodified corncob particles. The pore structure of corncob modified by alkaline  $\text{H}_2\text{O}_2$  is large, SRB has strong reducing activity, and alkaline substances exist in the particles, so the removal effect of 1# particle on  $\text{Mn}^{2+}$  is better than that of 2# particle.

By comparison with Tables 4 and 5, it can be seen that the adsorption equilibrium capacity  $q_e$  ( $0.642 \text{ mg g}^{-1}$  and  $0.581 \text{ mg g}^{-1}$ ) of # 1 and # 2 particles on  $\text{Fe}^{2+}$  is higher than that of  $\text{Mn}^{2+}$



**Fig. 12** Kinetic curves of adsorption  $\text{Mn}^{2+}$  by immobilized particles. 1# is modified corncob fixed SRB sludge particles, 2# is unmodified corncob fixed SRB sludge particles.



Table 5 Kinetic parameters of adsorption  $\text{Mn}^{2+}$  by immobilized particles

Sample	$q_e$	Pseudo-first-order adsorption kinetic		Pseudo-second-order adsorption kinetic	
		$k_1$	$R^2$	$k_2$	$R^2$
1#	0.425	0.168	0.93267	0.5920	0.99302
2#	0.386	0.208	0.93354	0.3229	0.98803

(0.425  $\text{mg g}^{-1}$  and 0.386  $\text{mg g}^{-1}$ ). In the solution where  $\text{Fe}^{2+}$  and  $\text{Mn}^{2+}$  coexist, immobilized particles will form a competitive adsorption effect of ions, so the removal effect of  $\text{Fe}^{2+}$  by particles is better than that of  $\text{Mn}^{2+}$ , and the adsorption equilibrium capacity of  $\text{Fe}^{2+}$  is higher than that of  $\text{Mn}^{2+}$ .

### Instrumental analysis

**XRD analysis.** The particles before and after the above dynamic test reaction were dried and ground to a 200 mesh powder, and subjected to XRD analysis. The results are shown in Fig. 13.

From Fig. 13, the particles (No. 1 to No. 6) before and after the treatment of the wastewater contained polyvinyl alcohol and sodium alginate. Among them, diffraction peaks of polyvinyl alcohol appeared when  $2\theta$  was  $9.6^\circ$ ,  $11.1^\circ$ ,  $16.2^\circ$ ,  $19.7^\circ$ ,  $20.2^\circ$ ,  $22.3^\circ$ ,  $27.4^\circ$ , and  $40.7^\circ$ . This result is consistent with the findings of Hai T. A. P. and R. Ricciardi. They found that the diffraction peaks of PVA at  $2\theta$  values of 19.68, 22.31, and  $40.65^\circ$ , which correspond to the (101), (200), and (111) planes of the monoclinic unit cell.<sup>30,31</sup> At the same time, M. Hema<sup>32</sup> found that there is a correlation between the diffraction peak height of pure vinyl alcohol and the crystallinity of vinyl alcohol at  $19.7^\circ$  of  $2\theta$ . The diffraction peak of sodium alginate appeared at  $13.7^\circ$  and  $23.0^\circ$  of  $2\theta$ ,<sup>33</sup> but the intensity of the diffraction peak was not strong, indicating that the crystallinity of sodium alginate was low. According to Xuan D. *et al.*, the peak associated with sodium alginate in the XRD pattern was not obvious due to the amorphous nature of sodium alginate.<sup>34</sup> Hemicellulose and lignin in corncobs are amorphous components, while cellulose is a crystalline component,<sup>35</sup> so the focus of the XRD pattern of corncobs is to analyze changes in cellulose. Comparing the XRD patterns of No. 2 and No. 5, the intensity of the diffraction peak at  $12^\circ$ ,  $15^\circ$ ,  $21^\circ$ ,  $23^\circ$ ,  $27^\circ$ , and  $34^\circ$  of  $2\theta$  after the treatment of AMD by the unmodified corncob particles decreased. The XRD patterns of No. 3 and No. 6 showed that the intensity of the diffraction peaks at  $12^\circ$ ,  $17^\circ$ ,  $21^\circ$ , and  $27^\circ$  decreased after the modified corncob particles were treated with AMD. Studies have shown that diffraction peak is a typical cellulose I-type structure<sup>36,37</sup> when occurred at  $2\theta = 15^\circ$  (101 planes),  $17^\circ$  (101 planes),  $21^\circ$  (021 planes),  $23^\circ$  (002 planes) and  $34^\circ$  (004 planes), the peak at  $12^\circ$  is a characteristic peak of cellulose II, and the peak at about  $27^\circ$  corresponds to the (002 plane) of carbon.<sup>38</sup> Therefore, after processing the AMD, the cellulose and C contained in the corncob in the particles had a decrease in crystallinity, indicating that the corncob could slowly release organic matter for

SRB growth and metabolism while the particles were repairing AMD. The difference in peak reduction might be related to the modification process of the corncob. At the same time, in the XRD patterns of No. 5 and No. 6, MnS and FeS peaks appeared at  $30^\circ$ ,  $34.2^\circ$ , and  $49.1^\circ$  of  $2\theta$ , indicating that SRB can utilize the carbon source released by corncob for self metabolism, and produce  $\text{S}^{2-}$ , which formed sulfide precipitation with the heavy metal ions in the AMD. In addition, there was  $\text{Mn}(\text{OH})_2$  in No. 6. It might be that the heavy metal ions in AMD react with NaOH in the modified corncob particles to form hydroxide precipitation, or that modified corncob promoted SRB metabolism to produce alkalinity, which converted heavy metal ions into hydroxide precipitation.

**SEM analysis.** The particles before and after the dynamic test were dried and the surface and the internal structure of the particles were scanned by SEM. The microstructures of the particles were observed before and after the reaction, and the mechanism of particle co-treatment of AMD was further revealed. SEM micrographs are shown in Fig. 14 and 15.

Fig. 14(a), (b) and (c) are SEM surface structure micrographs of SRB sludge fixed particles without corncob, unmodified corncob fixed SRB sludge particles and modified corncob fixed SRB sludge particles before the dynamic experimental reaction respectively. Fig. 14(d), (e) and (f) are SEM surface structure micrographs of SRB sludge fixed particles without corncob, unmodified corncob fixed SRB sludge particles and modified corncob fixed SRB sludge particles after the dynamic experimental reaction respectively.

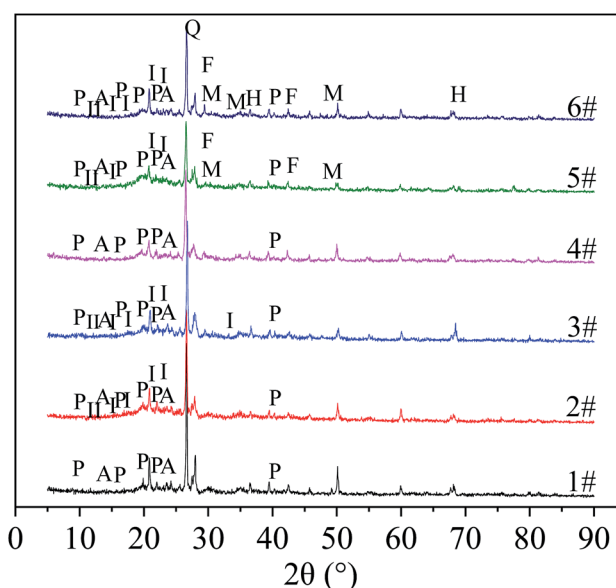
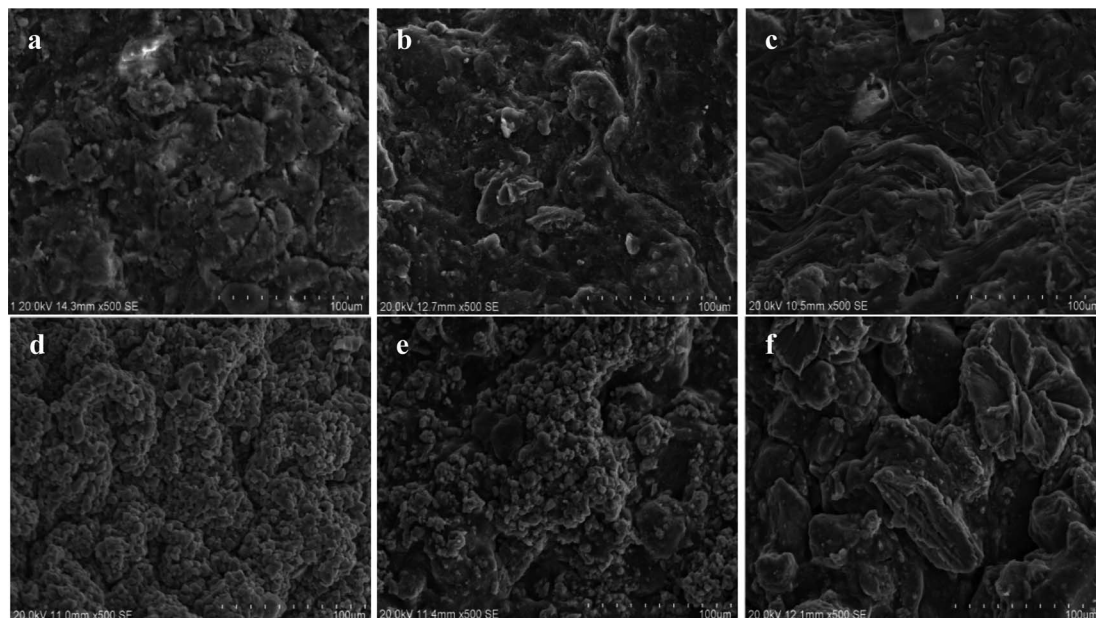


Fig. 13 XRD results of particles before and after treatment. No. 1 and No. 4 were XRD patterns before and after treatment of AMD without corncob particles, respectively. No. 2 and No. 5 were XRD patterns before and after treatment of AMD with unmodified corncob particles, respectively. No. 3 and No. 6 were XRD patterns of modified corncob particles before and after treatment of AMD. P: polyvinyl alcohol, A: sodium alginate, I: cellulose I, II: cellulose II, Q: quartz, F: FeS, M: MnS, H:  $\text{Mn}(\text{OH})_2$ .

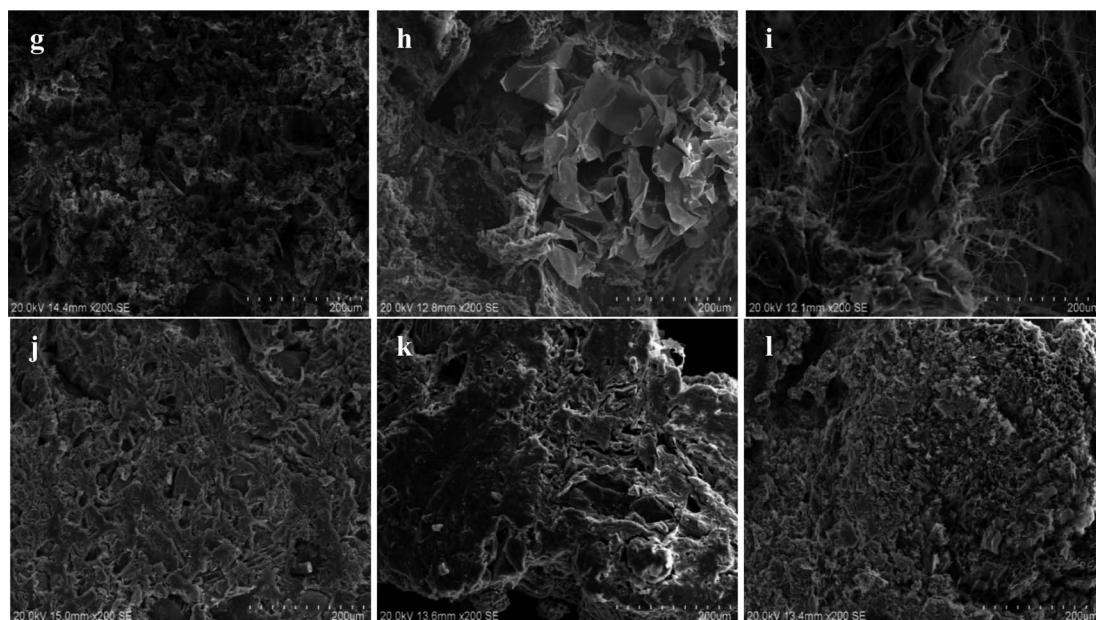




**Fig. 14** Three kinds of particle surface electron microscopy before and after the dynamic experimental reaction. (a) SEM surface structure micrographs of SRB sludge fixed particles without corncob before the dynamic experimental reaction. (b) SEM surface structure micrographs of unmodified corncob fixed SRB sludge particles before the dynamic experimental reaction. (c) SEM surface structure micrographs of modified corncob fixed SRB sludge particles before the dynamic experimental reaction. (d) SEM surface structure micrographs of SRB sludge fixed particles without corncob after the dynamic experimental reaction. (e) SEM surface structure micrographs of unmodified corncob fixed SRB sludge particles after the dynamic experimental reaction. (f) SEM surface structure micrographs of modified corncob fixed SRB sludge particles after the dynamic experimental reaction.

From Fig. 14(a–c), the surface of particles without corncob was rough, and the concavo-convex was obvious. There were large cracks and less irregular pores. The surface of the particles

with unmodified corncob was more rough, and the concavo-convex were more obvious. There were small pieces of cracks and a large number of irregular large pores. The surface of the



**Fig. 15** Three kinds of particle inside electron microscopy before and after the dynamic experimental reaction. (g) SEM inside structure micrographs of SRB sludge fixed particles without corncob before the dynamic experimental reaction. (h) SEM inside structure micrographs of unmodified corncob fixed SRB sludge particles before the dynamic experimental reaction. (i) SEM inside structure micrographs of modified corncob fixed SRB sludge particles before the dynamic experimental reaction. (j) SEM inside structure micrographs of SRB sludge fixed particles without corncob after the dynamic experimental reaction. (k) SEM inside structure micrographs of unmodified corncob fixed SRB sludge particles after the dynamic experimental reaction. (l) SEM inside structure micrographs of modified corncob fixed SRB sludge particles after the dynamic experimental reaction.



particles with modified corncobs had a complex, porous interstitial filamentous structure with large pores and fissures, and each gap penetrated and was closely connected. This suggested that the addition of corncobs increased the roughness and porosity of the particles. The modification process of alkaline  $H_2O_2$  destroyed a large amount of cellulose crystal structure and lignin in corncobs, resulting in a low molecular weight fibrous structure.

According to Fig. 14(d–f), a large amount of particles were deposited on the surface of particles without corncob after the reaction, and less were deposited on the surface of particles with unmodified corncob after the reaction. Nearly none were deposited on the surface of particles with modified corncob after the reaction, and the filamentous structure disappeared, resulting in larger pores and cracks. This indicates that due to the smaller pores and weaker SRB activity, the metal ions were adsorbed on the surface and deposited on the particles without corncobs. While particles containing modified corncobs had a large pore structure and strong SRB biological activity, metal ions can successfully enter the inside of particles and the reaction deposition occurred. And because the SRB used the organic carbon sources, the fibrous structure of the particle surface disappeared.

Fig. 15(g), (h) and (i) are SEM inside structure micrographs including of SRB sludge fixed particles without corncob, unmodified corncob fixed SRB sludge particles and modified corncob fixed SRB sludge particles before the dynamic experimental reaction respectively. Fig. 15(j), (k) and (l) are SEM inside structure micrographs of SRB sludge fixed particles without corncob, unmodified corncob fixed SRB sludge particles and modified corncob fixed SRB sludge particles after the dynamic experimental reaction respectively.

From Fig. 15(g–i), the pore distribution of the inside of particles without corncob was relatively uniform and the pores were small, and no obvious large block material structure. The pore distribution of particles with unmodified corncob was not uniform, the pores were large, and the block material structure existed. The distribution of pores in the particles with modified corncob was more uneven and the pores were larger, and the cross-network structure was obvious. This indicated that the addition of corncobs enlarged and increased the internal pores of the particles, and increased the inhomogeneity of the pore structure. The obvious block structure might be the added corncob. Alkaline  $H_2O_2$  modification process changed the structure of corncob, which presented the crossover material structure of particles with modified corncob.

From Fig. 15(j–l), we can see that the internal structure of particles without corncob did not change much. The massive material of particles with unmodified corncob disappeared and the pores became smaller and less. And the internal crossover material structure of the particles with modified corncob disappeared, and a large number of honeycomb small voids appeared. This indicated that particles with corncob (including both unmodified and modified corncob), were easier to exchange material within and outside the granular system due to the well-developed pores and good permeability, which could guarantee the free access of

nutrients and metabolic products needed by the growth of microorganisms and was conducive to promoting the adsorption of particles and SRB dissimilation and reduction activity. While the particles with modified corncob contained organic carbon sources, available to SRB, and richer pore channels. Therefore, the change of the internal structure of the particles with the modified corncob was the most significant, which further proved the superiority of the modified corncob as the internal carbon source.

## Conclusions

The single factor test was used to study the modification conditions of corncob. The optimum conditions for the determination of modified corncob were as follows: the modification time was 24 h, the concentration of NaOH was 6% and the concentration of  $H_2O_2$  was 1.5%.

Combined with single factor test results, using the orthogonal test of  $L_9(3^3)$ , the optimum modification conditions of corncob were determined by analysis of variance and range analysis. The optimum modification conditions were as follows: the modification time was 24 h, the concentration of NaOH was 6% and the concentration of  $H_2O_2$  was 1.5%. The modified corncob prepared under these conditions had the best effect on the immobilized SRB particles.

The three constructed dynamic columns had certain ability to remove and resist change of the impact load of  $SO_4^{2-}$ ,  $Mn^{2+}$  and  $Fe^{2+}$  in AMD solution. Dynamic Column 3 was strongest in removing ability and resisting impact load change, and its ability to control the effluent COD release and increase the solution pH value was better than those of Column 1 and 2. This suggests that modified corncobs are more suitable as internal carbon source for SRB to treat AMD than unmodified corncobs.

The reduction reaction process of particles to  $SO_4^{2-}$  in wastewater is mainly affected by electron acceptor, and the first-order reaction model can well describe the reduction process. The adsorption process of  $Fe^{2+}$  by particles is in accordance with Elovich adsorption kinetics model, and the adsorption process of  $Mn^{2+}$  is in accordance with pseudo-second-order adsorption kinetics model. In the waste where  $Fe^{2+}$  and  $Mn^{2+}$  coexist, immobilized particles will form a competitive adsorption effect of ions, so the removal effect of  $Fe^{2+}$  by particles is better than that of  $Mn^{2+}$ .

XRD and SEM analysis showed that the chemical composition and spatial structure of the three immobilized particles changed significantly before and after the reaction. The particles with modified corncob showed various precipitates such as  $MnS$ ,  $FeS$  and  $Mn(OH)_2$  after the reaction, which indicated the main mechanism of the removal of heavy metal ions. The particles with modified corncob had multi-porous cross-filament structure on the surface and inside before the reaction, and there were large pores and cracks as well. After the reaction, the surface and the internal filament structure disappeared, and larger pores and cracks appeared on the surface, while a large number of small honeycomb-like voids appeared inside, and the internal structure changed greatly.



## Conflicts of interest

The authors declare that there is no conflict of interests regarding the publication of this article.

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## Notes and references

- 1 S. K. Hwang and E. H. Jho, *Sci. Total Environ.*, 2018, **635**, 1308–1316.
- 2 K. K. Kefeni, T. A. M. Msagati and B. B. Mamba, *J. Cleaner Prod.*, 2017, **151**, 475–493.
- 3 A. J. Giachini, T. S. Sulzbach, A. L. Pinto, R. D. Armas, D. H. Cortez, E. P. Silva, E. B. Buzanello, A. G. Soares, C. Soares and M. J. Rossi, *Arch. Microbiol.*, 2018, **200**, 1227–1237.
- 4 A. Qureshi, Y. Jia, C. Maurice and B. Ohlander, *Environ. Sci. Pollut. Res. Int.*, 2016, **23**, 17083–17094.
- 5 A. S. Sheoran and V. Sheoran, *Miner. Eng.*, 2006, **19**, 105–116.
- 6 D. K. Nordstrom, D. W. Blowes and C. J. Ptacek, *Appl. Geochem.*, 2015, **57**, 3–16.
- 7 R. Perez-Lopez, D. Quispe, J. Castillo and J. M. Nieto, *Am. Mineral.*, 2011, **96**, 781–791.
- 8 P. J. Oberholster, A. R. De Klerk, L. De Klerk, J. Chamier and A.-M. Botha, *Ecol. Indic.*, 2016, **62**, 106–116.
- 9 T. T. Wei, Y. Yu, Z. Q. Hu, Y. B. Cao, Y. Gao, Y. Q. Yang, X. J. Wang and P. J. Wang, *Appl. Mech. Mater.*, 2013, **409–410**, 214–220.
- 10 M. Zhang and H. Wang, *Miner. Eng.*, 2016, **92**, 63–71.
- 11 M. Berggren and P. A. del Giorgio, *J. Geophys. Res.: Biogeosci.*, 2015, **120**, 989–999.
- 12 F. Jiang, MPhil, Liaoning Technical University, 2015.
- 13 N. Yu, Z.-Y. Zhu, Y. Liu, J.-Y. Zhang and Y.-M. Zhang, *Ind. Crops Prod.*, 2017, **95**, 163–169.
- 14 C. Huang, X. Y. Yang, L. Xiong, H. J. Guo, J. Luo, B. Wang, H. R. Zhang, X. Q. Lin and X. D. Chen, *Appl. Biochem. Biotechnol.*, 2015, **175**, 1678–1688.
- 15 X. Y. Guo, L. Zhang, S. T. Shu and J. Y. Hao, *Appl. Mech. Mater.*, 2014, **672–674**, 154–158.
- 16 R. L. Tseng and S. K. Tseng, *J. Colloid Interface Sci.*, 2005, **287**, 428–437.
- 17 W. Zhao, R. Hao, B. Li, W. Zhang and P. Du, *Environ. Sci.*, 2014, **35**(3), 987–994.
- 18 Z. Jiang, T. He, J. Li and C. Hu, *Green Chem.*, 2014, **16**, 4257–4265.
- 19 Y. Su, R. Du, H. Guo, M. Cao, Q. Wu, R. Su, W. Qi and Z. He, *Food Bioprod. Process.*, 2015, **94**, 322–330.
- 20 G. Li, J. Chen, T. Yang, J. Sun and S. Yu, *Water Sci. Technol.*, 2012, **65**, 1238–1243.
- 21 N. Le Moigne and P. Navard, *Cellulose*, 2009, **17**, 31–45.
- 22 E. I. Evstigneyev, O. S. Yuzikhin, A. A. Gurinov, A. Y. Ivanov, T. O. Artamonova, M. A. Khodorkovskiy, E. A. Bessonova and A. V. Vasilyev, *J. Wood Chem. Technol.*, 2016, **36**, 259–269.
- 23 M. Zhang, H. Wang and X. Han, *Chemosphere*, 2016, **154**, 215–223.
- 24 T. P. H. van den Brand, K. Roest, G.-H. Chen, D. Brdjanovic and M. C. M. van Loosdrecht, *Environ. Eng. Sci.*, 2015, **32**, 858–864.
- 25 K. Yoo, K. Sasaki, N. Hiroyoshi, M. Tsunekawa and T. Hirajima, *Mater. Trans.*, 2004, **45**, 2429–2434.
- 26 J. Di, W. An, M. Wang, W. Zhao, J. Guo and H. Han, *China Water Wastewater*, 2016, **32**, 120–124.
- 27 A. D. Karathanasis, J. D. Edwards and C. D. Barton, *Mine Water Environ.*, 2010, **29**, 144–153.
- 28 J. Di, W. An, N. Dai, Z. Zhu, F. Jiang, Y. Ren and Q. Zhao, *Chin. J. Environ. Eng.*, 2016, **10**, 1103–1108.
- 29 P. Kikot, M. Viera, C. Mignone and E. Donati, *Hydrometallurgy*, 2010, **104**, 494–500.
- 30 R. Ricciardi, F. Auriemma, C. D. Rosa and F. Lauprêtre, *Macromolecules*, 2004, **37**, 1921–1927.
- 31 T. A. P. Hai and R. Sugimoto, *Synth. Met.*, 2018, **240**, 37–43.
- 32 M. Hema, S. Selvasekarapandian, D. Arunkumar, A. Sakunthala and H. Nithya, *J. Non-Cryst. Solids*, 2009, **355**, 84–90.
- 33 Q. Wang, N. Zhang, X. Hu, J. Yang and Y. Du, *J. Biomed. Mater. Res., Part A*, 2007, **82**, 122–128.
- 34 D. Xuan, Y. Zhou, W. Nie and P. Chen, *Carbohydr. Polym.*, 2017, **155**, 40–48.
- 35 T. K. Ghose, *Pure Appl. Chem.*, 1987, **59**, 257–268.
- 36 H. A. Silvério, W. P. Flauzino Neto, N. O. Dantas and D. Pasquini, *Ind. Crops Prod.*, 2013, **44**, 427–436.
- 37 W. P. Flauzino Neto, H. A. Silvério, N. O. Dantas and D. Pasquini, *Ind. Crops Prod.*, 2013, **42**, 480–488.
- 38 I. M. A. Mohamed, A. S. Yasin, N. A. M. Barakat, S. A. Song, H. E. Lee and S. S. Kim, *Appl. Surf. Sci.*, 2018, **435**, 122–129.

